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Atomistic modelling of electrocatalysis: Are we there yet?

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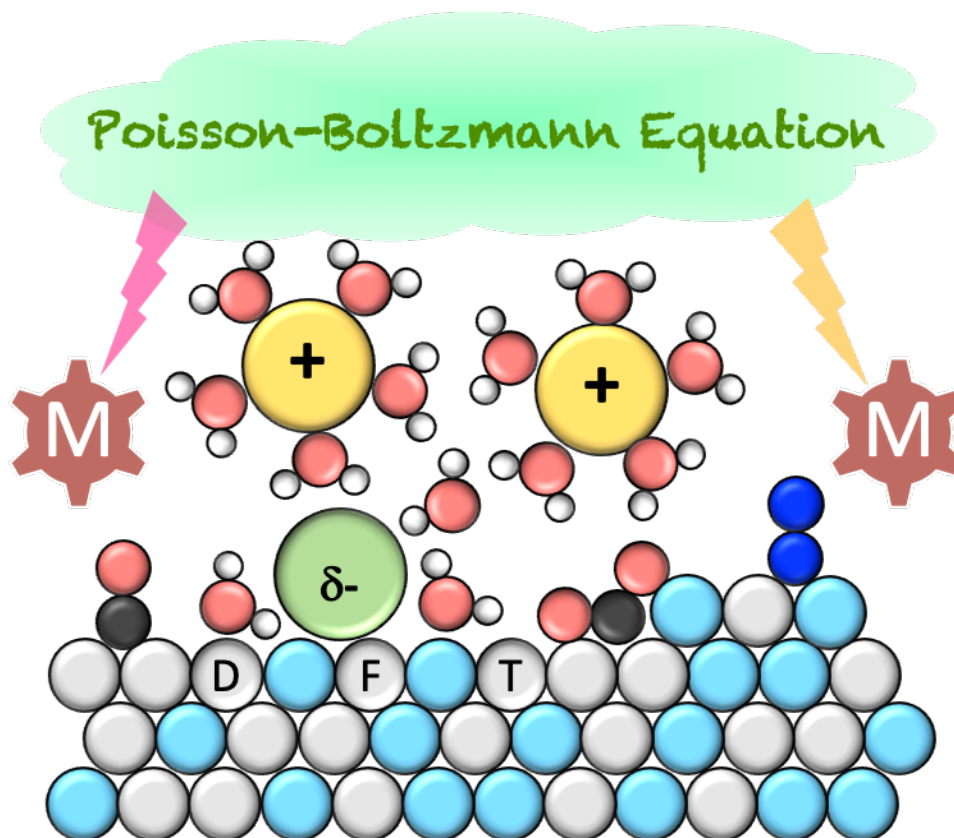
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Abstract

Electrified interfaces play a prime role in energy technologies, from batteries and capacitors to heterogeneous electrocatalysis. The atomistic understanding and modelling of these interfaces is challenging due to the structural complexity and the presence of the electrochemical potential. Including the potential explicitly in the quantum mechanical simulations is equivalent to simulating systems with a surface charge. For realistic relationships between the potential and the surface charge (i.e., the capacity), the solvent and counter charge need to be considered. The solvent and electrolyte description are limited by the computational power: either molecules and ions are included explicitly, but the phase-space sampling is at least 10 times too small to reach convergence or implicit solvent and electrolyte descriptions are adopted which suffer from a lack of realism. Both approaches suffer from a lack of validation against directly comparable experimental data. Furthermore, the limitations of density functional theory in terms of accuracy are critical for these metal/liquid interfaces. Nevertheless, the atomistic insight in electrocatalytic interfaces allow insights with unprecedented details. The joint theoretical and experimental efforts to design non-noble hydrogen evolution catalysts are discussed as an example for the success of theory to spur and accelerate experimental discoveries.

Graphical/Visual Abstract and Caption



1. INTRODUCTION

Electrochemistry is a key technology to overcome our fossil energy based economy and energy-management.(1,2) The common characteristic of electrochemical processes is the importance of interfaces: Even for molecular electrochemistry, where the reaction is conceptually occurring in bulk solution, the need to exchange electrons with the electrode limits the reaction to the diffusion layer. Indeed, in certain cases the impact of the electric double layer on such reactions has been evidenced. (3–5) The major efforts for understanding electrified interfaces is devoted to batteries(6,7) and electrocatalysis,(8,9) with the more historical applications in metal-plating and corrosion(10–12) being more sparse, despite enormous economic importance (corrosion generates costs of ~3% of the gross-domestic-product).(13) For batteries, the aim is to extend the cyclability of Li batteries and improve their safety via reduced dendrite formation propensity.(14) The development of Na or Mg based batteries, is also a hot topic.(15) At the intersection between batteries and electrocatalysis are the super- and pseudo-capacitors, which allow a short-term storage of electrical energy.(16) These electrochemical devices are essentially constituted of interfaces and still poorly understood due to their complex morphology.(17) Heterogeneous electrocatalysis, which is the main object of this review, combines the power of electrochemistry and heterogeneous catalysis to drive and accelerate chemical reactions, respectively. This catalytic activity is located at the electrode surface, where a (mostly liquid) electrolyte ensures the transport of ionic species.

The typical electrocatalyst is composed of a metallic surface immersed in an electrolytic solution, i.e., a solvent containing the dissolved electrolyte. However, there is a large diversity of electrocatalytic devices, ranging from gas-diffusion electrodes with only very small liquid layers at room-temperature(18) to solid-oxide fuel cells and electrolyzers(19) which operate at ~500 K and rely on a

solid electrolyte. Similarly, in the electrochemical promotion of catalysis,(20) the catalyst is supported on a solid electrolyte and it is the application of the electrochemical potential that tunes the chemical reactivity. Another axis of diversity opens when considering the nature of the electrocatalyst: while transition metals are most common and Pt being the prototypical electrocatalyst, sp-metals such as Sn or Bi have also attracted attention.(21) Furthermore, non-noble transition metal oxides, sulphides, phosphides etc. have a great potential to replace the rare and expensive Pt catalysts.(22–24) More exotic electrocatalysts exploit intermetallics,(25) MXenes(26) and transition metal doped carbon nitrides.(27,28) These materials generally undergo modifications upon their use as electrocatalysts and eventually they age and deactivate. This leads to challenges regarding the determination of the relevant surface state and interfaces.(29) However, since these challenges resemble very much the ones faced in heterogeneous catalysis in general,(30,31) we do not treat them herein.

Today, the main application of electrocatalysis is the electrolysis of water to generate ultra-pure hydrogen and, in the reverse direction, fuel cells.(32) The use of other fuels than hydrogen (e.g., ethanol or boron hydrides) is characteristic of the research of novel electrocatalysts for oxidation reactions,(33) while electro-reduction catalysts are mostly devoted to CO₂(34) and, increasingly, N₂ reduction.(35) Application of electrocatalysis for electrosynthesis of organic molecules is comparably rare.(36)

The large diversity of electrified interfaces in general and even when focusing on heterogeneous electrocatalysis constitutes the first major challenge for their computational modelling: Even with today computational resources, a single methodology is unlikely to be powerful enough to reliably describe the disorderd solid-electrolyte interphases found in batteries, the impact of morphology on super-capacitor performance and bond-breaking and formation processes at the active sites on electrocatalysts. Hence, the relevant scale (from atoms to centimeters) dictates the use of given modelling methods or imposes the use of mulit-scale models.

Even though electrocatalysis is intrinsically a multiscale process, with charge- and mass-transport being of significant technological importance,(37) we herein focus on the atomic scale: The atomic scale is where the intrinsic limitations of the catalysts are found and where the methods that allow to understand electrocatalysis can be almost directly transposed to other electrified interfaces. We refer to the appropriate reviews for the larger scale models (applied to fuel cells) and do not treat them herein.(38–40)

The remainder of this review is structured into five sections. The short first one discusses conceptual questions that haunt the atomistic modelling of electrocatalysis. The second section gives an extensive overview on techniques to model the electrified interface. The third section discusses outstanding challenges that wait for solutions and the fourth section presents the significant insights obtained by atomistic modelling of electrocatalysis for the hydrogen evolution reaction. The fifth section briefly discusses possibilities to more thoroughly validate the atomistic models.

2. VOLCANO CURVES, OVERPOTENTIAL AND THEIR POTENTIAL DEPENDENCE

Computational modelling complements experimental observations and is, in principle, able to suggest the most promising combinations of electrode material, electrolyte and electrochemical potential. However, given that the methodology required for these simulations is still in its infancy, it cannot come as a surprise that there is no consensus in the community and harsh criticism regarding the

usefulness of computational modelling to complement experimental studies is partially justified.(41,42) Here, we first introduce the notion of the volcano diagrams and their limitations and then tackle the overpotential and how to connect the experimental overpotential to the quantities from atomistic modelling.

Sabatier's principle of heterogeneous catalysis states that the optimal catalyst binds adsorbates and intermediates with an appropriate binding energy: if intermediates are bound too strongly, they block active sites and if they are too weakly adsorbed, they are not sufficiently present on the surface to allow for high reaction rates.(43) This principle still holds in electrocatalysis: indeed, it is almost trivial that intermediates need to be stabilized appropriately in order for the reaction to proceed without major kinetic limitations. In practice, for a family of catalysts, the Sabatier principle is most often represented as a two-dimensional graph, with a descriptor (possibly the binding energy of the limiting intermediate) serving as the x-axis and the y-axis representing the activity. These graphs are commonly called "volcano plot", as they tend to show a unique maximum activity. Following an earlier proposal by Parsons,(44) Trasatti was the first to publish an electrocatalytic volcano curve, applied to the case of the hydrogen evolution reaction (HER).(45) While in principle any descriptor that leads to an exploitable volcano curve works, the determination of this descriptor (x-axis) is not always possible experimentally. Therefore, most of the volcano curves in the literature rely on Density Functional Theory (DFT) computations(42) as pioneered by Nørskov and co-workers.(46) Nevertheless, as we will discuss in section 3, the binding energy of intermediates depend, in principle, on the full electrified electrode/electrolyte interface and even on the target potential. It turns out that in practice simple descriptors, such as the adsorption energy of atoms (typically H, O, N and C) are a close enough proxy for the determination of volcano curves. The advantage of these very simple descriptors is that they allow to screen a very large number of similar materials (e.g., surface alloys) at modest computational cost.(47,48,8,9,49)

The slopes of the volcano tend to be linear. This has led to a second highly important concept: the "universal" scaling relationships,(50–52) which describe the fact that the adsorption energy of various intermediates cannot be tuned independently. Overcoming these scaling relations has, therefore, been an important goal in the recent years, with the main approaches being (i) bifunctionality, e.g., alloys (ii) addition of promoters, (iii) functionalization of the surface via ligands (iv) electrolyte and solvent engineering, (v) interfacial sites, i.e., heterofunctional catalysts and (vi) confinement (see Fig. **Figure 1**). (53–55)

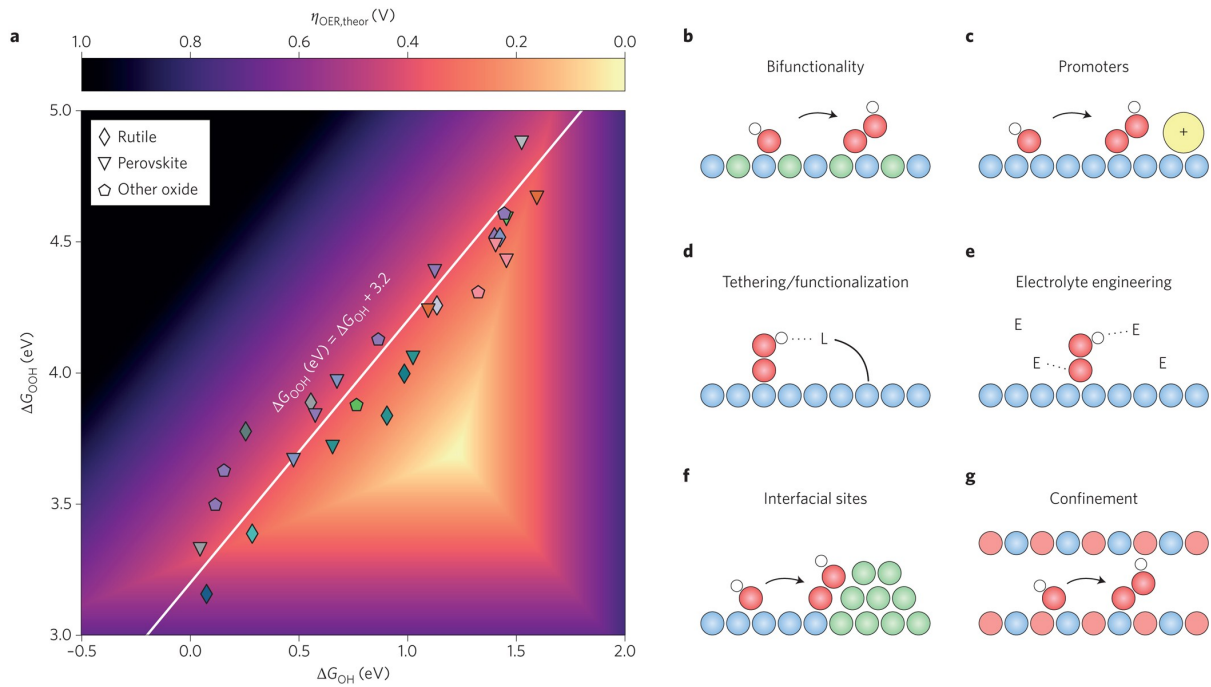


Figure 1 Left: Scaling relationship for OH* and OOH* over various oxides, superposed on a heat map of activity as a function of these two variables. Note that the optimum (brightest color) is does not lie on the scaling line (a). Right: Examples of how to overcome the scaling relation by introducing adsorbate specific stabilizations. Bifunctionality (b) on alloys or mixed metal/oxide/sulfide catalysts, promoters in terms of specifically adsorbed ions (c), ligands (d) or electrolyte molecules (e) can stabilize various adsorbates specifically. Three dimensional active sites such as interfaces between two materials (f) or confinement (g) can also provide opportunities to break scaling relationships. Differently colored circles indicate distinct elements. Reproduced from(54), copyright Nature Publishing Group.

Closely related to the volcano curves in electrocatalysis is the concept of overpotential and the exchange current density: In short, the overpotential (η) is the minimal voltage that is necessary to significantly accelerate an electrochemical reaction, while the exchange-current density (i_0) is the current density at the equilibrium potential U° , i.e. zero overpotential. The standard equilibrium potential is connected to the free reaction energy ΔG° via:

$$\Delta G^\circ = -nFU^\circ \quad \text{Eq. 1}$$

Where n is the number of electrons exchanged and F Faraday's constant. The more general form of this equation is known as the Nernst equation. For a half-cell reduction reaction it reads:

$$U_{red} = U_{red}^\circ - \frac{RT}{nF} \ln \frac{a_{red}}{a_{ox}} \quad \text{Eq. 2}$$

Where R is the universal gas constant, T the temperature and a_{red} and a_{ox} the activities of the reduced and oxidized species, respectively. **Eq. 1** and **Eq. 2** together determine the minimum cell potential that is required for the reaction to be spontaneous, i.e., that the reaction free energy is zero or negative. If a larger potential difference is applied, the reaction is driven by the external power source. In the reverse direction, the overpotential measures the minimal energy loss of a battery or fuel cell.

In the simplest systems, the Tafel equation provides the link between overpotential and the exchange-current density:

$$\eta = A \cdot \log_{10} \left(\frac{i}{i_0} \right) \quad \text{Eq. 3}$$

Where A is the so called Tafel slope and i the current density. **Figure 2** shows a model Tafel plot to illustrate the connection between the Tafel slope and the exchange current density. It also becomes

evident that the Tafel plot relies on a linear approximation in order to extract information out of the measured current densities.

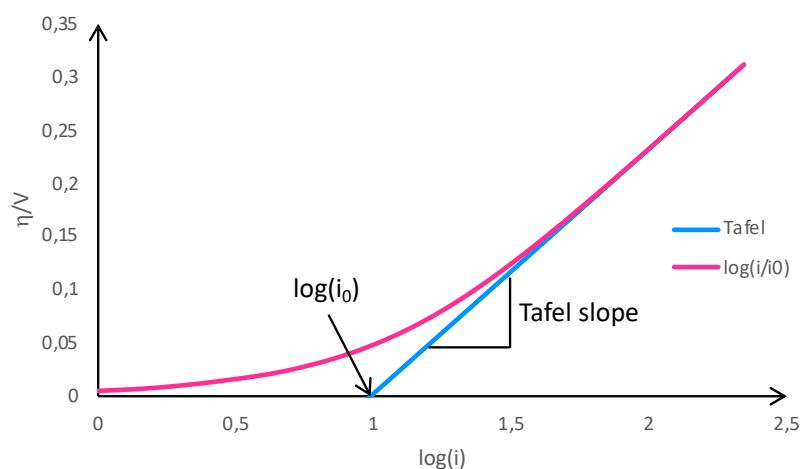


Figure 2 Schematic plot of the overpotential (η) vs the logarithm of the current density (pink). The Tafel slope and the exchange current density i_0 as determined by the linear fit (Tafel approximation) are indicated (blue).

Tafel slopes depend on the reaction mechanism, i.e., if the electrochemical step (i.e., an elementary steps where electrons are exchanged with the electrode) is rate determining or not and if it is preceded by chemical or electrochemical steps, as nicely summarized in ref (56). Here, chemical steps refer to elementary steps that preserve the formal number of electroactive species in the system. Typical values for Tafel slopes at room temperature vary between 30-120 mV/decade. Detailed microkinetic models show that the connection between the observed Tafel slope and the reaction mechanism is far from straightforward.(57–59) For atomistic modelling of electrocatalysis, the limitation is almost immediately clear: The overpotential as defined in **Eq. 3** is intrinsically a kinetic quantity and thus depends on the larger scales. The crudest approximations for obtaining these kinetic characteristics is the framework of microkinetics, where mean-field differential equations are solved.(60) Lattice based kinetic Monte Carlo simulations, where the coverages of intermediates and their spatial arrangements are explicitly simulated, are more realistic, but also computationally much more involved.(61,62) A variety of other multi-scale and multi-physics simulations have been developed in order to include the macroscopic properties such as the charge and mass-transfer, viscosity and so on. (38–40) Since performing these simulations exclusively from first principles information is unfeasible in practice, a number of approximations have been adopted. While some studies prefer to approximate prefactors in the mean-field differential equations which then are solved under steady-state conditions, the most common approximation is the most drastic one: neglecting all detailed kinetics and focusing exclusively on the thermodynamics of the reaction pathway. This leads to the concept of “thermodynamic overpotential” (η_{TD}). η_{TD} is the potential which is necessary to make all the electrochemical steps either athermic or exothermic.(60,63) Assuming that the reaction mechanism only involves electrochemical steps and no chemical steps, η_{TD} ensures that there are no thermodynamic barriers anymore. In contrast with the kinetic modelling, determining η_{TD} is well defined at an atomistic level. Furthermore, the influence of the level of theory (inclusion of the surface polarization, coverage and electrolyte effects and so on) can be assessed with greater ease than for a full kinetic study. However, this approximation is very crude: in particular, the reaction barriers and the thermodynamics of chemical steps, as opposed to electrochemical steps, are

difficult to include within this framework. Hence, while η_{TD} tends to be a relevant quantity for screening catalysts, it is clearly not enough for a detailed understanding of reaction mechanism.(64) The second key issue with η_{TD} is that it is such a crude approximation that it is not even clear what experimental quantity it should be compared or correlated to: Is it the overpotential based on the Tafel relation, i.e., when the linear approximation becomes valid? – The overpotential as determined experimentally based on a limiting current density, often taken as 10 mA/cm² and thus more commonly available than a complete Tafel analysis? – Or correlated to the exchange current density, which is most closely related to thermodynamic barriers?

The discussion of η_{TD} is closely related to the volcano curves, as it is one of the most frequently used theoretical measures for the activity of a given family of catalysts. To simplify the determination of theoretical volcano plots even further, the binding energy of intermediates in the absence of an applied potential can be used. This, however, is an even more dramatic approximation, as it assumes that the binding energy of all intermediates behave the same as a function of the potential. This later approximation is equivalent to assume that the Tafel slope does not change as a function of the potential, an assumption that is rarely fulfilled for effective electrocatalysts.(65–67)

3. APPROACHES TO MODEL THE ELECTRIFIED INTERFACE

3.1 Early days

In this section, we adopt a historical and largely chronological point of view. As very nicely described by Bockris in 1971, the existence of overpotentials, which is a key concept in electrocatalysis, was still very ill understood and even underrated by the community for many decades.(68) Therefore, it cannot come as a surprise that atomistic modelling via quantum chemistry was very slow to shed light on this tantalizing problem. Nevertheless, the group of Anderson pioneered the use of quantum-chemical computations to investigate electrocatalysis already in the late 70ies.(69–71) These early (semi-empirical) approaches already tried to take the electrochemical potential into account, for instance by shifting the ionization energy of metal atoms. This approach could even be validated against experimental changes in infra-red frequencies as a function of the electrode potential.(70) Similarly and only slightly later, the “dipped adcluster” model of Nakatsuij proposed to perform Hartree-Fock computations with an applied potential based on a cluster model.(72) However, the electronic structure of the cluster features a HOMO-LUMO gap in contrast with metals that have a continuous density of state around the Fermi level. This gap turned out to interfere strongly, so that the physics of a metallic electrode(73) could not be retrieved. As most first-principles studies, the golden days only started in the late 1990s. The work by Bureau and Lecayon introduced the use of grand-canonical DFT (GC-DFT) for simulating electrocatalysis in 1997.(74) To make it simple (for details see section 3.3.1), this approach consists in varying the number of electrons. Thus, the Fermi-level, i.e., the workfunction and thus the electrochemical potential is tuned. Even though the actual model was still based on a cluster and the treatment of the solvent and countercharge was not addressed, the theoretical framework and its connection with conceptual DFT, such as the importance of the Fukui function, was fully laid out. Around the same time, Anderson’s group introduced a second key concept in the atomistic modelling of electrocatalysis, which is the thermodynamic cycle that allows to link a given computation to the electrochemical potential. (75) The approach to deal with electrochemical steps closely follows the early arguments of Bockris(76) which avoids the need to know the energy of the (solvated) electron independently. Although performing computations in the gas-phase, Anderson

already assessed the potential dependence not only of electrochemical steps, but also for reaction barriers.(77) Coming from the molecular chemistry community, these pioneering computations basically disregarded the explicit influence of the electrode or approximated the electrode by single atoms. (78) Bredas and co-workers validated the use of grand-canonical DFT by determining Stark tuning slopes, for which a model charge-distribution was used to stabilize the surface charges.(79) Alternatively, explicit countercharges were used.(80) To simplify such computations that probe the effect of the local electric field, the use of a uniform electric field was proposed.(81) However, the relation between the field-dependent results and the actual electrochemical potential is not immediate and this approach was thus not very successful for electrocatalysis. (82) Along the lines of grand-canonical DFT, the group of Alavi and Sprik turned to the problem of using periodic computations for treating solution redox chemistry.(83) The periodicity introduces the delicate issue of how to neutralize the unit cell in order to avoid a diverging Coulomb potential of periodic charges. In this pioneering work, the authors relied on a homogeneous background charge without discussing the issue in details. The use of the homogeneous background charge was subsequently criticized since it introduces artificial energy terms and forces.(84) Nevertheless, even in this study the conclusion was that for some applications it remains the only practical solution, although the introduced errors are difficult to assess. Alavi and co-workers later provided examples for the use of a Gaussian countercharge sheet in periodic simulations instead of the homogeneous background charge. However, the scheme was still applied in vacuum.(85) For determining redox-potentials in solution, Sprik and co-workers heavily rely on the homogeneous background charge, last but not least since no clear, practical alternative exists for small, periodic simulation cells.(86,87) Turning back to heterogeneous electrocatalysis, it was again Anderson who, after summarizing his “reaction center” cluster approach (88) turned to the use of periodic slabs. (89,90) Compared to the now so popular method called computational hydrogen electrode (CHE), which we will discuss in the next subsection, Andersons scheme is virtually identical. The only substantial difference is that he did not advocate the use of the very convenient H₂ molecule as a reference compound but rather the OH radical. This scientifically small detail seems to have tipped the balance, so that today the concept of CHE and the surge of computations for heterogeneous catalysis is ascribed to the Norskov group instead of Anderson, who, at the very same time, had already two decades of experience with the atomistic modelling of electrocatalysis.

3.2 The CHE revolution

The field of atomistic modeling of electrocatalysis has made an enormous leap forward with the introduction of the computational hydrogen electrode (CHE) by Norskov et al.(60) in 2004. The CHE method enables the evaluation of the relative free energies of intermediates along an electrochemical reaction network. Each elementary step is formulated as a removal or addition of the electroactive species, (H⁺ + e⁻). The pathway of water oxidation, relevant to water electrolysis, provides a typical example:



The adsorbates are indicated via an asterisk and a single asterisk * denotes an empty reactive site on the catalyst. At each (electrochemical) step, (H⁺ + e⁻) are abstracted from the adsorbates. The free energy ΔG for each step in the oxidation reaction can then be expressed as below:

$$\begin{aligned}\Delta G_1 &= G_{\text{OH}^*} - G_{\text{H}_2\text{O}} - G^* + \mu_{\text{H}^+} + \mu_{\text{e}^-} \\ \Delta G_2 &= G_{\text{O}^*} - G_{\text{OH}^*} + \mu_{\text{H}^+} + \mu_{\text{e}^-} \\ \Delta G_3 &= G_{\text{OOH}^*} - G_{\text{OOH}^*} - G_{\text{H}_2\text{O}} + \mu_{\text{H}^+} + \mu_{\text{e}^-} \\ \Delta G_4 &= G_{\text{O}_2} + G^* - G_{\text{OOH}^*} + \mu_{\text{H}^+} + \mu_{\text{e}^-}\end{aligned}\tag{Eq. 5}$$

G_x is the free energy of each specie involved and it can be determined by DFT computations. It typically contains the electronic energy E , the entropy S , and a zero-point energy correction ZPE: $G = E + \text{ZPE} - TS$. μ_{H^+} and μ_{e^-} are the electrochemical potentials of the removed protons and electrons, respectively, which are not directly determined by ab initio computations. Instead, the beautifully simple method of Norskov is based on the observation that under standard conditions ($\text{pH} = 0$, $T = 298 \text{ K}$, $p = 1 \text{ bar}$), defining 0 V of the standard hydrogen electrode (SHE), there is an equilibrium ($\Delta G = 0$) between hydrogen molecules in the gas phase and solvated protons and electrons:



Hence, the free energy of the H_2 molecule in the gas phase, which is trivial to compute, can be used to determine the combined energy of ($\text{H}^+ + \text{e}^-$):

$$\mu_{\text{H}^+}(\text{U}=0 \text{ vs SHE}) + \mu_{\text{e}^-}(\text{U}=0 \text{ vs SHE}) = \frac{1}{2} G_{\text{H}_2} \tag{Eq. 7}$$

which is much easier than computing the solvation energy of the protons and electrons. When changing the potential away from 0 V vs SHE a term neU has to be added where n is the number of exchange electrons, e the elementary charge and U the potential, all in suitable units. This term is then added to each electrochemical step. Together with a free energy correction for the proton activity that depends on the pH of solution, the free energy of an electrochemical oxidation step according to CHE is written as:

$$\Delta G_x(\text{U}) = \Delta G_x - neU + k_B T \text{pH} \ln(10) \tag{Eq. 8}$$

Use of Eq. 8 naturally shifts the focus of CHE from SHE to the reversible hydrogen electrode (RHE), i.e., to a pH independent reference electrode. Indeed, CHE is not well adapted to study (explicit) pH effects and implicitly assumes that the neutral surface state computed by DFT is representative of 0 V vs RHE.

In order for the overall reaction to occur without thermodynamic barriers, $\Delta G_x(\text{U})$ has to be either zero or negative. The computation of free energy differences along a pathway allows to estimate this potential U_{min} , which, when compared to the equilibrium potential U° defines the thermodynamic overpotential η_{TD} , itself by definition always positive. The great simplicity of CHE relies on a second approximation: the solvation effects are assumed to be negligible, which means that only computations in the gas-phase are necessary to evaluate electrocatalytic reaction pathways. However, this second assumption is not necessary and some authors use the CHE in conjunction with solvation effects.(91,92)

A lot of DFT based research in electrocatalysis studies relies on the CHE approach as a tool to design the catalytic activities of different materials(47,49,93,94), determine phase diagrams(95) and estimate the free energy of various electroreduction and electrooxidation reactions such as the oxygen reduction reaction,(96) the oxygen evolution reaction,(97) the hydrogen evolution reaction (HER)(98–100), the CO_2 reduction reaction(101–103),and the N_2 reduction reaction(104,105).

Despite these exploitable results, CHE has several limitations and unanswered questions remain related to the approximations in establishing the model. In the original applications of the CHE, solvation effects were assumed to play a minor role for the energetics of neutral reaction intermediates and thus omitted. The interaction between electrode, adsorbates and the solvent is neglected. In addition, the electrode potential in this method is not an explicit variable in the quantum mechanical calculations. The electronic energies are only measured for electroneutral entities independently from the electrochemical potential as the electron transfer reaction is always coupled. This is problematic in view of the accumulating experimental evidence that decoupled pathways are

also key. (106–109) All these interrelated limitations of the CHE method are addressed by the “beyond the CHE” methods discussed in the next subsections, which explicitly model the electrochemical potential and the electrode/electrolyte interface.

To summarize, the CHE approach often provides information on adsorption trends but the role of the chemical electrolyte/electrode interface is outside the scope of this type of model. It only allows the evaluation of the thermodynamics of the elementary reaction steps. Generally, activation barriers for charge transfer reactions are ignored, even though some schemes have been developed to approximate the potential dependence of barriers in a CHE-like spirit. (110–112)

In the next section, we will review the methods to model the electrochemical with increased realism compared to CHE, which allow to understand the interaction between adsorbates and the electrified, solvated electrode. These methods are based on explicitly changing the surface charge density of the electrode surface, so that activation energies and chemical steps, can also be assessed at a given potential.

3.3 Beyond the computational hydrogen electrode

Any method that aims to go beyond the first order CHE approximation should describe the interactions between the adsorbates and the electrified electrode. This can be either achieved via potential dependent density functional theory computations (section 3.3.1) or model Hamiltonians (section 3.3.2). We here emphasize that to be considered an improvement beyond the CHE, the methods have to include the energetics as expressed by Eq. 7 and 8. Figure 3 summarizes the various possibilities to treat the electrochemical interface, ranging from the vacuum interface at zero charge to grand-canonical DFT coupled to molecular dynamics. The aim of the figure is to illustrate the panel of available models and that the timeline is scattered across the two-dimensional surface.

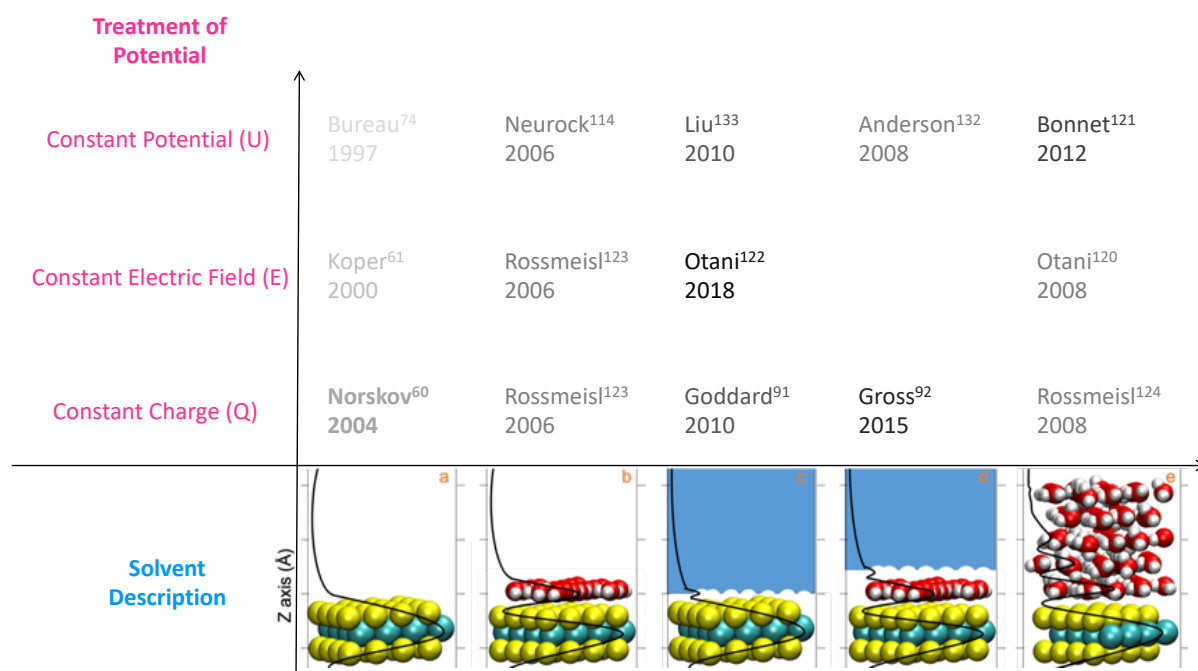


Figure 3 Representation of the key contributions introducing various descriptions of the electrocatalytic interface. The superscript refers to the reference to the corresponding article, while the year is given below the lead author name. The grey scale highlights the chronological ordering. Note, that other groups have applied these techniques in various contexts and

some of the methods have been developed several times independently with slight variations, but we have tried to highlight the earliest contributions. The horizontal axis is indicative of the treatment of the solvent, with the black line symbolizing the electrostatic potential. (a) corresponds to vacuum, (b) the combination of a small number of solvent molecules at the vacuum interface, (c) the use of an implicit solvent, (d) the combination of implicit solvent (light blue) and explicit solvent molecules and (e) the use of a fully explicit description of the liquid phase. The CHE cornerstone is highlighted in bold. The vertical axis represents the description of the electrochemical potential. Constant charge approaches typically work with neutral surfaces. Constant electric fields can be used as a proxy for the electrochemical potential under certain assumptions (typically the thickness of the double layer). Truly constant potential approaches are typically relying on grand-canonical DFT.

3.3.1 Grand-Canonical Density Functional Theory

As discussed in section 3.1, grand canonical DFT is the framework which allows to perform electronic structure computations at a given electrochemical potential to mimic the experimental conditions. Hence, the electronic grand-canonical (free) energy F now depends on the electrochemical potential U , instead of being evaluated at zero charge as in the CHE approximation:

$$F(U) = E(U) - q(U) * U \approx E(U_0) - \frac{1}{2}C(U - U_0)^2 \quad \text{Eq. 9}$$

Where $E(U)$ is the electronic energy at potential U , the (surface) charge q_{surf} is positive if electrons are removed and negative when electrons are added. U_0 is the workfunction of the system at zero charge and thus $E(U_0)$ is the energy which would be used in the CHE framework. The second equality is valid for metallic (gap-less) systems, for which the grand-canonical energy $F(U)$ can be expanded to second order in terms of the capacitance C of the system. In contrast to CHE, the explicit dependence of the electronic energies on the electrochemical potential capture the potential-dependence of formally chemical steps,^(113,64) e.g. the desorption of O_2 depicted in **Figure 4**. The last ten years or so have seen several computational approaches that all rely on **Eq. 9** to compute potential dependent adsorption energies on periodic surfaces. There are two main characteristics that distinguish them: how the surface charge is neutralized and how the solvent is described.

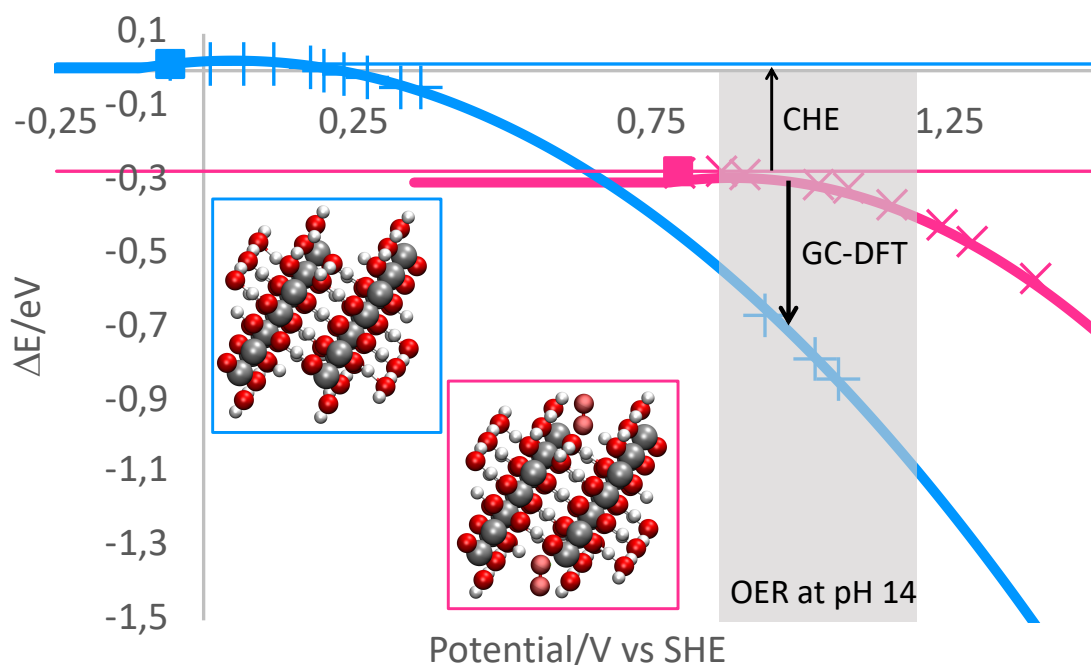


Figure 4 Comparison of CHE (horizontal lines and solid points) and grand-canonical DFT (empty symbols and thick lines) and grand-canonical DFT (GC-DFT) results for the energy difference of two intermediates during the oxygen evolution reaction (OER) over CoOOH (a semi-conductor) in basic conditions. The intermediate in pink (crosses) has O_2 bound to the catalyst, while O_2 is desorbed in the intermediate in blue (plus signs). The desorbing O_2 molecule is highlighted in pale red in the

structural representation, which also exemplifies the use of symmetric unit cells to avoid surface dipole moments and obtain unambiguous workfunctions. The data is taken from ref. (64)

Fundamentally, a charged periodic system has a divergent Coulomb energy. Hence, the periodic unit cell has to be neutralized. However, the countercharge can take various forms. Neurock and co-workers were among the first to apply grand-canonical DFT to electrocatalysis on a routine basis for electrocatalysis.(114–116) In their scheme, a homogenous background charge is used to neutralized the simulation cell. However, since part of the system is metallic, a simple energy correction with respect to the neutral system allows to obtain physically relevant electronic energies. In practice, they used static water (ice) layers to simulate the solvent, even though later the approach has been transposed to the use of implicit solvation models.(117,118) This model does not require any modification of the computational software, explaining its attractiveness from a user perspective. Performing different computations with varying numbers of electrons, a quadratic fit (right hand side of **Eq. 9**) allows to determine the energies as functions of the electrochemical potential at moderate computational cost. Finally, the equilibrium (redox) potential is calculated as an intersecting point between two curves.

Almost at the same time, Otani and Sugino(119–122) presented another model which avoids the use of the homogeneous background charge. This approach is called “effective screening medium” (ESM) and exploits Green’s function techniques to formally treat a non-repeated slab along the direction of the surface normal. The Green’s function allows to impose two different electron potentials across the unit cell, so that the “bottom” can be imposed to have the metal bulk potential, while the potential away from the surface mimics the counter-electrode. The model is most suitable to simulate situations with a linear electric field in the solvent region and requires extensive modifications of the electronic structure code, so that it is not widely used. Like for other approaches that explicitly rely on an electric field,(123) the assumptions about the thickness of the double layer introduces uncertainties which are require somewhat arbitrary decisions.

The homogeneous background charge,(114) ESM(119) or the Gaussian charge sheets(85) can only qualitatively represent the actual electric double layer and introduce, therefore, ad hoc assumptions. To avoid using these artificial treatments for countercharges, Rossmeisl and co-workers(124–127) proposed to control the electrode potential by changing the number of hydrogen atoms that dissociate into protons and electrons so that the total charge remains neutral. However, adding or removing one electron and one proton to or from the practical (i.e., relatively small) unit cells causes a drastic change in the calculated electrode potential, since fractional protons(128) are not considered as an option. Consequently, controlling the electrode potential within a desired narrow range is difficult. Hence, the reaction energies for electrochemical steps (e.g., adsorption of a proton) need to extrapolated to infinite supercell sizes. Furthermore, just like for the ice-layers used by Neurock,(114) the description of solvation effects via static solvent layers is very approximate, but highly important for these charged species. Note, that these arguments “against” the generalized CHE method of Rossmeisl and co-workers are only of practical nature. In principle, the generalized CHE method with varying numbers of explicit cations/anions is sound and if sufficiently large unit cells could be afforded, while achieving a representative sampling of the solvent (and electrolyte) phase-space, it would lead to a realistic description of the electrified interface (up to the errors of DFT itself, discussed in section 4.2). However, we need to keep in mind that such simulations imply a grand-canonical treatment of

the electrolyte and/or electroactive species. Based on the experience with constant pH simulations at the classical force field level(129,130), which remain a field of active research, and the practical and conceptual issues discussed in section 4.1, we expect that these simulations remain rather impractical for routine applications at least for another ten years or so. Nevertheless, promising results have been obtained for the simulation of cyclic voltammograms.(131)

An elegant solution is to modify the surface charge and describe the double layer according the solution of the (modified) Poisson-Boltzmann (mPB) equation.(132–138) The Poisson-Boltzmann equation provides a mean-field description, obtained by placing the countercharge in the solvent region according to the local effective electrostatic potential. Compared to the homogeneous background charge or the Gaussian charge-sheet, the resulting ion distributions are much more realistic. The formal theory behind this combination has also been called joint DFT (JDFT) by Arias and co-workers,(134,139) making reference to the fact that these ion- and solvent distributions can be formally generated by classical, “molecular” density functional theory, while the electronic structure of the slab is determined by the more conventional quantum mechanical, “electronic” DFT. The mPB equations do not automatically satisfy the charge neutrality constraint, which necessitates further corrections.(140,141) Thus, even the electronic DFT combined with mPB models need in practice either to use Lagrange multipliers in order to enforce the charge neutrality,(140) to employ simple background charges(142) or to utilize modified mPB boundary conditions(119) which mimic neutralization by image charges to treat the charged semi-periodic systems. This contrasts with the most popular linearized Poisson-Boltzmann equation, which automatically generates neutral unit cells. Since the linearized PB equation assumes that the countercharge is made of point charges, it neglects, however, all finite size effects, resulting in countercharge distributions that approaches the solute too closely.(143) Compared to the homogeneous background method, where an energy correction is needed that assumes that the slab is metallic, the (linearized) Poisson-Boltzmann equation is also readily compatible with semi-conductors (see **Figure 4**)(64) and layered electrocatalysts,(144,145) making this approach universally applicable. Indeed, the advantage of the combination of DFT with the linearized Poisson-Boltzmann equation is that it can be applied to any charged system, while producing physically meaningful electronic energies. Hence, we have also applied it to investigate the origin of electrochemical promotion of catalysis (EPOC) phenomenon(146,147) and to study the effect and the role of adding a base in heterogeneous catalysis (148).

3.3.2 Model Hamiltonians

While DFT is the most popular approach used to simulate the electrochemical reactions, it has several limits. As mentioned in section 3.1, several theories for electrochemical reactions already existed a long time before the advent of DFT, taking their origin in the theories of Marcus(149) and Hush(150) and its extended versions.(151) These treatments rely on a highly simplified representation of the valence orbital of the reactant, the electrons on the metal, their electronic interaction, and the coupling of the electron exchange to the solvent. The electrochemical potential is taken into account by shifting the reactant’s valence orbital with respect to the metal. The current use of model Hamiltonians in electrocatalysis is mainly the result of Schmickler and co-workers, (152–154) who have combined Marcus theory and the Anderson-Newns model,(155,156) within a tight-binding or Hückel approximation and exploiting Green’s function methods. This model Hamiltonian allows to account for bond-breaking and formation and leads to a thorough understanding of trends in electrocatalysis,

e.g. for the hydrogen evolution reaction.(157) A similar decoupling between electron transfer and the evolution of the nuclear coordinates can, in principle, be achieved via the block-localized wave function techniques(158) at the DFT level. On the one hand, the limitations of the model Hamiltonians are, of course, that they have to be set up for each system specifically, with certain terms nowadays being determined by DFT computations.(159) On the other hand, the major strength of the model Hamiltonians is the deep insight they provide and the ability to compute free-energy surfaces for electrochemical reactions as a function of the reactant's position and the solvent coordinate, a topic that we will discuss again in the next section.

4. SELECTED CHALLENGES WHEN MODELLING ELECTROCATALYSIS

4.1 The rise and fall of implicit solvents

Electrocatalysis occurs at the electrode–electrolyte interface and a very important part of the electrochemical system is the solvent and thus the associated solvation free energies.(160) Solvation free energies are statistically averaged interactions between solutes and solutions. Most of the schemes discussed in section 3 include either just a few solvent molecules(110) or a couple ice layers(161) but without performing any statistical average or, in the other extreme, they include only the statistical average, directly obtained from integral or differential equations. These two options define the typical “explicit” and “implicit” solvation models, although approaches that combine characteristic of the two are common.(162,163)

The downsides of explicit solvation models are the limited sampling of static structures, the difficulty of reaching a global energy minimum and the challenge to maintain a consistent solvation sphere for various adsorbates along a reaction pathway. Ab initio molecular dynamics (AIMD) (126,164) can alleviate these drawbacks, as it allows to sample a range of solvation structures at the electrode/electrolyte interface. However, AIMD does not reach the time and length scales necessary to assure a representative sampling of the structural ensemble. Indeed, given the size of the solid/liquid interface, brute force AIMD studies have only been applied to open-circuit model systems,(165–168) or small systems, where lateral ordering is imposed. Even for these rare case-studies, the accessible <50 ps of dynamics are an order of magnitude smaller than what is necessary to equilibrate a pure H₂O/Pt(111) interface, even in the absence of ions and adsorbates.(169) To circumvent this issue, the group of Goddard has been relying extensively on pre-equilibration at a reactive force field level of theory(170) before turning to the actual AIMD simulations.(168,171,172) This is only possible if a suitable empirical force fields is available, which is currently only the case for Cu(170) and for dominating noble metal surfaces(169,173). Furthermore, “suitable” is based on the assumption that the equilibration at the force field level brings the system into the vicinity of the DFT equilibrium ensemble, a hypothesis that is difficult to (in-)validate in the absence of rigorous DFT based sampling.

From a statistical mechanics point of view, the atomistic description of the interface is meaningless in the absence of a representative phase-space sampling. Typically $\sim 10^6$ time steps, corresponding to ~ 1 ns, are necessary to capture entropic effects appropriately. Alternatively, implicit solvation models(174,175) or the in practice basically equivalent(176) classical density functional theory for liquids(139,177) could in principle achieve a computationally much more efficient description of the liquid phase. These methods derive their efficiency from the mean-field treatment, i.e., by averaging

the fast solvent modes leading to the description of the liquid phase as an effective medium. This is particularly attractive at interfaces, where the diffusion coefficient is low,(178) requiring equilibration times of ~ 0.5 ns.(169,179) Hence, a fundamental questions arises:(180) The electrolyte relaxation times (~ 0.5 ns) are long compared to individual bond formation times (~ 1 ps), although fast compared to turn over frequencies where one reactive event per second and active site is generally acceptable. Therefore, comparatively rare electrolyte arrangements that lower the activation barrier could dominate the overall kinetics due to the exponential dependence of the rate on the activation energy, in analogy to the relation between nanoparticle shape and activity.(181) As well appreciated, Marcus theory identifies solvent reorganizations as one of the major contributions to the activation energy of electron transfer processes,(182) i.e., the electron transfer only occurs when the solvent fluctuations have generated a favorable environment. Furthermore, we have shown that the solvation-shell of cations indeed fluctuate significantly, and cation-specifically, at the electrified interface.(183) For electrocatalytic reactions, where both the electrolyte and the adsorbate have to rearrange, the situation is less clear: During the reaction, the adsorbate undergoes a bond-breaking/formation process (horizontal axis of **Figure 5**) and the electrolyte (vertical axis) has to adapt to the new arrangement. If the solvent rearrangement is fast (flat solvation free energy surface in y-direction) compared to the nuclear reorganization, then the adiabatic approximation applied to the stationary states can also be applied to the transition state (TS) and no special treatment is necessary. If, however, the electrolyte has to re-organize significantly, then such a procedure is inadequate.

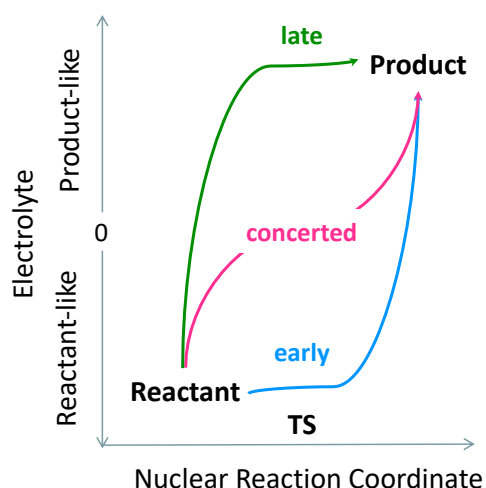


Figure 5 Model Free Energy surface in terms of the adsorbate, nuclear coordinate (x-axis) and the electrolyte arrangement coordinate (y-axis). Early and late transition states with respect to the electrolyte coordinate can be distinguished.

Although non-equilibrium continuum solvation models exist, it is unlikely that they will be able to resolve this fundamental question of coupling between the solution and adsorbate dynamics. The model Hamiltonians of section 3.3.2 are, however, ideally suited to address these questions.(159) For thermodynamic quantities, i.e., reaction energies, the implicit solvation models come with a significant advantage: in contrast with the explicit phase space sampling, which is associated with a statistical uncertainty, the effective medium reproducibly determines a single number. Additionally, electronic structure computations in conjunction with an implicit solvent are associated with only a very small computational overhead compared to vacuum, so that reaction pathways are efficiently and conveniently accessible.(184–186,138) However, these implicit solvation models have several shortcomings: (i) by construction, the implicit solvent and electrolyte will never actively participate in

the reaction and not even act as a relay for proton transfers. This is likely at the origin of their negligible influence on scaling relations.(187,163) Therefore, such “strong” or “direct” solvation effects need to be modelled via the explicit inclusion of the relevant solvent molecules.(163,188) (ii) Many models, and in particular the ones that belong to the broad class of polarizable continuum models, neglect the size of solvent or electrolyte molecules and ions. As a consequence, the effective medium is in too close contact with the quantum mechanical system. This leads to formally filled cavities and the unphysical presence of solution between weakly bound layers (e.g., in graphite and MoS₂), even if there is no space to accommodate solvent molecules.(145,189) Strategies to overcome the issue are actively explored.(190) (iii) All of these models reside on more or less extensive parametrization. The existence of comprehensive tables for solvation free energies in bulk solution(191) makes this parametrization rather unproblematic for solvated molecules and ions, even though the crude approximations in implicit solvent models prohibits the development of truly universal models, spurring the continuous development in the field.(192,193,184,194,186,195) Such experimental benchmark data is much more sparse and less diverse for the (electrified) surfaces, so that the current parametrizations are only based on bulk solvation properties of organic molecules and ions. Therefore, while being very attractive from a practical point of view, the accuracy of these models is very questionable.(196,197)

4.2 The electronic structure level of theory: Dominance and Limitation of Density Functional Approximations

As for any computational investigation using electronic structure methods, the accuracy and cost of the level of theory has to be balanced. For heterogeneous electrocatalysis, the choices are rather limited by seemingly trivial requirement: The electronic structure method needs to be able to describe (periodic) metallic systems. For example, this is not the case for the perturbative Møller-Plesset wave function methods.(198) MP2, which is widely used in main-group chemistry, diverges for gap-less systems. Even the most efficient, (near) linear scaling density matrix optimization algorithms(199,200) for Kohn-Sham density functional theory (DFT),(201) rely on the existence of a fundamental band-gap. Hence, they cannot be applied to metals. To overcome this technical limitation it has been proposed to neglect the partially occupied orbitals altogether.(202) While it technically works after an initialization using a different method, the consequence of simulating this somewhat unphysical electronic structure on the derived properties is hard to predict. At the expense of a very large prefactor, rigorous near-linear scaling can be achieved while preserving the physics of partially occupied orbitals.(203,204) This method only becomes competitive for large systems (> 1000 atoms) and supercomputers, which prohibits its routine application. The second requirement related to the electronic structure method is, of course, its robustness and accuracy. Since heterogeneous (electro-)catalysis deals with the activation of small molecules such as CO, CO₂, N₂ and O₂, on surfaces that range from transition metals to oxides, sulfides and TM-doped graphenes, the room for error cancellation is rather limited. Indeed, depending on the interaction of the atoms and molecular fragments with the catalyst, the change in electronic structure is significant.(205)

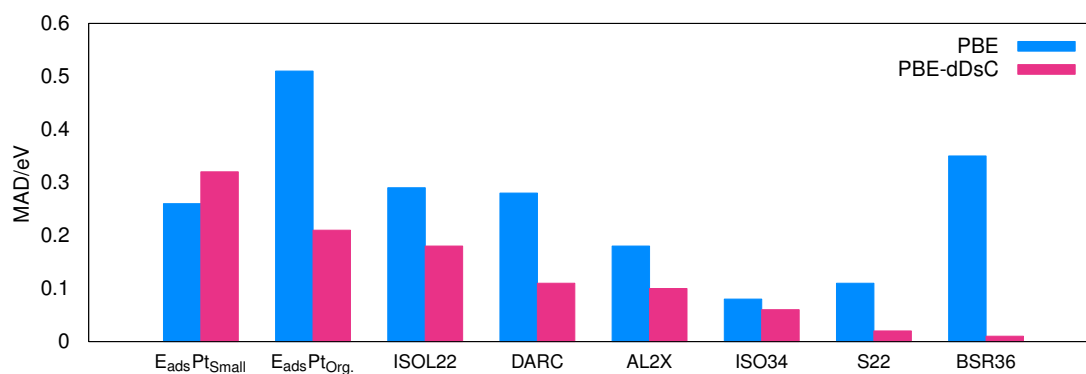


Figure 6 Mean average deviations (MADs) for eight diverse test sets. The performance for PBE (blue) and the dispersion corrected PBE-dDsC (pink) are shown. The data for surface adsorption is taken from ref (206). The $E_{adsPtSmall}$ refers to H, O and CO adsorption on Pt(111). $E_{adsPtOrg.}$ contains methane, ethane, ethylidyne, cyclohexene, benzene and naphthalene on Pt(111). The other test sets are extracted from GMTKN24 (207). The PBE-dDsC data was published in ref. (208). ISOL22 assesses isomerization energies of large organic molecules. DARC stands for Diels-Alder reaction energies, while AL2X contains dimerization energies of AlX3 species. ISO34 determines isomerization energies of small organic molecules while S22 is dedicated to weak intermolecular interactions. BSR36 measures bond separation energies of alkanes, i.e., hydrogenation reactions of alkanes to methane.

The challenges can be exemplified in the context of the development of density functionals according to the generalized gradient approximation (GGA). PBE,(209) was developed as a general-purpose functional. The functional has then been revised to yield more accurate molecular thermodynamics under the name revPBE.(210) Alternatively, RPBE(211) achieves a better compromise between metallic surfaces and molecular adsorption. Later, PBEsol,(212) was proposed for higher accuracy for solids. Even though these functionals share the philosophy of minimal parameterization, they illustrate the struggles one is facing when having to balance the accuracy of overall (gas-phase) thermodynamics of molecules, the behavior of adsorbates and the electronic structure of solids. Even reproducing the structure and density of liquid water with DFT is challenging,(213) illustrating the additional constraints the liquid component of the solid/liquid interface imposes on the robustness of density functionals for atomistic simulations of electrocatalysis.

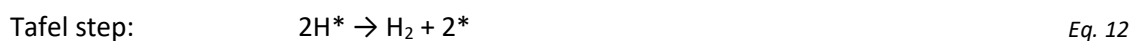
The consequences of this limited accuracy are exemplified by the performance for energetic quantities for the few systems where the corresponding benchmark data is well-established. Recently, the claim has appeared that an accuracy of 0.05 eV is routinely achievable in electrocatalysis.(214) However, the solid/gas interface benchmark data suggests that an accuracy of ~ 0.2 eV(206,215) is the best we can hope for in general. This is depicted in **Figure 6**, where the mono- and diatomic adsorbates on Pt(111) are separated from the organic adsorbates and typical data for gas-phase thermodynamics is shown as well. The graph also nicely illustrates the benefit of a dispersion correction at the metal/gas interface. Error cancellation or accumulation occurs unpredictably, so that the accuracy for relative energies is expected to be on the same order of ~ 0.2 eV for surface reaction and adsorption energies, which compares well with the 0.2-0.3 eV gas-phase thermodynamics accuracy for PBE or RPBE.(216) Solution data is much less available, but the accuracy of DFT in solution has no reason to be higher than in gas-phase. Due to the omnipresent delocalization error,(217) the situation at the metal/liquid interface is probably even worse, in particular for electrolyte solutions. Just as a reminder, the delocalization error denotes the observation that semi-local and hybrid density functionals artificially stabilize fractional charges. This shows up as an unphysical smearing of charges among atoms, ions and molecules. The prototypical systems for this fundamental failure are the dissociation limits of alkali metal halides, for which GGAs erroneously predict fractionally charged ions instead of neutral

atoms.(218) Similarly, the band-alignment at the metal/liquid interface is problematic: The intrinsically delocalized electronic structure of metals leads to generally reasonable predictions for workfunctions by GGAs (within ~ 0.2 eV).(219) In contrast, the band-gap of semi-conductors and insulators is too small by 40% on average.(220) As a consequence, the band alignment at the interface can be seriously wrong and leading to spurious results.(221)

5. RATIONALIZING EXPERIMENTAL HER ACTIVITY WITH ATOMISTIC MODELLING: SUCCESS STORIES

An ever increasing energy demand due to rapid industrialization has accelerated the need to develop environmentally and economically sustainable means to produce renewable energy.(1,9,222) Hydrogen (H_2) production from water electrolysis has emerged as a promising process to fuel a carbon-neutral energy economy due to its clean reaction by-products (H_2 and O_2), which can be stored and released in fuel cells on demand.(223) In this regard, significant progress has been made towards industrial-scale water electrolysis for H_2 production through combining atomistic modelling and experimental data in the development of efficient and durable hydrogen evolution reaction (HER) electrocatalysts.(9,48) In this section, we review the exemplary progress made by atomistic modelling of electrocatalysis to inform and drive the experimental discovery and optimization of new materials for HER.

Fundamentally, HER in acidic solution proceeds through either of two mechanisms beginning with the Volmer step (**Eq. 10**), followed by the Heyrovsky (**Eq. 11**) or Tafel step (**Eq. 12**). Since both mechanisms begin with the adsorption of a free H^+ ion in the electrolyte onto an active site on the HER electrocatalyst surface (*), the rate of HER depends on the Gibbs free energy of H adsorption (ΔG_H).(44,46) Using a simple kinetic model, Nørskov *et al.* established a theoretical correlation between ΔG_H and the HER exchange current density (i_0) for a series of transition metals, in good agreement with the experimentally-obtained volcano plot and Sabatier's principle (**Figure 7a**).(46) The findings were consistent with platinum (Pt) as the most active pure transition metal HER electrocatalyst due to its near-thermoneutral ΔG_H , as predicted earlier by Parson.(44) Very recent reports reveal that Pt is special beyond the $\Delta G_H \sim 0$ criterion, i.e., that it is the high-coverage differential hydrogen adsorption properties that make Pt particularly active for HER.(224) Nevertheless, the establishment of ΔG_H as a simple yet accurate atomic-scale descriptor for HER catalytic activity has motivated a subsequent computational screening of 736 bimetallic alloy combinations from 16 transition metals. This screening predicted that a BiPt surface alloy could match and potentially outperform Pt in HER activity, given that the calculated ΔG_H difference between Pt and BiPt was ≈ 0.04 eV.(47) The BiPt alloy was synthesized by annealing its precursor (Pt-Bi_{ir}) and its HER performance was evaluated (**Figure 7b**). From **Figure 7b**, it was evident that at 0 V, BiPt alloy produced at least 50% more current density than Pt, (47) suggesting that a good atomistic understanding of the HER mechanism can inform high-throughput computational screening and yield accurate predictions of undiscovered HER catalysts.



Beyond transition metals, efforts have also been made in the computational screening and theory-guided discovery of economically viable and Earth-abundant non-platinum group metal HER electrocatalysts for industrial implementation.(8,125) Hinnemann and co-workers calculated the ΔG_H for various H-binding sites in HER-active hydrogenase and nitrogenase enzymes which revealed that ΔG_H was almost thermoneutral in the enzyme's equatorial sulfur (S) sites, expanding the application of ΔG_H descriptor for HER activity from metals to now include enzymes.(225) Recognizing a close similarity in the coordination number and mode between the nitrogenase S active site and S in 2D MoS₂, the authors then verified the HER activity of MoS₂ experimentally to demonstrate the theory-guided discovery of MoS₂ as promising HER catalysts. 2H-MoS₂ was subsequently determined to be HER active on its edges,(226) prompting follow-up reports in nanostructuring to expose more 2H-MoS₂ edge sites for enhanced HER activity.(227,228,23) Separately, Li *et al.* evaluated the effect of introducing S vacancies and elastic strain on ΔG_H and electronic band structure of monolayer 2H-MoS₂.(229) The results indicated that various levels of strain and S vacancy combinations could theoretically yield $\Delta G_H = 0$ in 2H-MoS₂ (**Figure 7c**), since both processes separately affect H binding to fine-tune and optimize ΔG_H . Indeed, subsequent experimental results confirmed that a combination of both tensile strain and S vacancies (SV-MoS₂) significantly enhanced the intrinsic HER activity of 2H-MoS₂ (**Figure 7d**). Based on grand-canonical DFT computations, we have recently shown that the active sites of MoS₂ might change in the presence of water, but remain active through a change in mechanism.(145) Apart from vacancy and strain engineering, other research groups have also employed a variety of phase, defect, dopant and interfacial engineering of 2D transition metal dichalcogenides (TMDs) using ΔG_H as a guiding principle for HER activity prediction and catalyst discovery.(230–236)

Most recently, a newly discovered class of 2D transition metal carbides, nitrides and carbonitrides (MXenes) has emerged as promising HER candidates due to their high metallic conductivity, hydrophilicity and solution processibility.(237–240) MXenes possess a general formula of M_{n+1}X_nT_x, where n = 1 – 4, M = early-row transition metal such as Mo, Ti, Cr and V, X = C and/or N and T_x = surface terminations such as –O, –OH and –F formed during MXene synthesis.(241–243) Early works reported that different polar T_x functional groups altered the electronic structure and band gap of MXenes,(244) which would consequently affect their ΔG_H and predicted HER activity.(245) Seh *et al.* first demonstrated the HER activity of MXenes experimentally after computing the volcano plot and identifying Mo₂CT_x as a promising HER electrocatalyst (**Figure 7e**). (246) Unlike 2H-MoS₂ where HER activity was localized only at its edges, the –O terminated MXene basal planes of Mo₂CT_x were determined to be HER active,(246,247) garnering significant interest in exploring MXenes as 2D HER electrocatalysts. Subsequently, follow-up work from Handoko *et al.* reported that tuning the composition of T_x terminations on MXene basal planes affected HER activity.(247) By considering the –F coverage on the MXene basal planes, the authors discovered a good agreement between the experimental and theoretical overpotential (calculated from ΔG_H) at a current density of -10 mA cm⁻² for a variety of MXenes (**Figure 7f**), thus validating the HER atomistic model and ΔG_H as a predictor for HER activity in MXenes. Thereafter, further computational work informed successive MXene designs to enhance HER activity by modulating ΔG_H through adsorption and doping of transition metals(26,248–251) and non-metals(252,253) onto MXene basal planes and through precise -T_x control.(254,255) Currently, efforts have also been made in the computational screening of novel MXenes with promising HER activity using ΔG_H as an indicator for targeted experimental synthesis and validation.(256–259)

In summary, mechanistic insights into the HER mechanism (Volmer, Tafel and Heyrovsky elementary steps) resulted in the discovery of the correlation between ΔG_H and HER activity. Using ΔG_H as a predictor for HER activity has resulted in good agreement with experimental data for varying classes of materials from transition metals to TMDs and MXenes, and informed numerous successful theory-guided experimental designs with significantly enhanced HER activity in acidic solution. We note in passing that the atomistic origin of the pH dependence of HER remains a vividly debated topic.(260–265) With a strong understanding in atomistic modelling, high-throughput computational screening can yield predictions with high levels of accuracy for targeted experimental work and reconcile experimentally observed trends. Coupled with advanced in-situ and operando characterization, these catalytic insights can be extended beyond HER to other more complex reactions such as CO₂ and N₂ reduction in the future.(266–268)

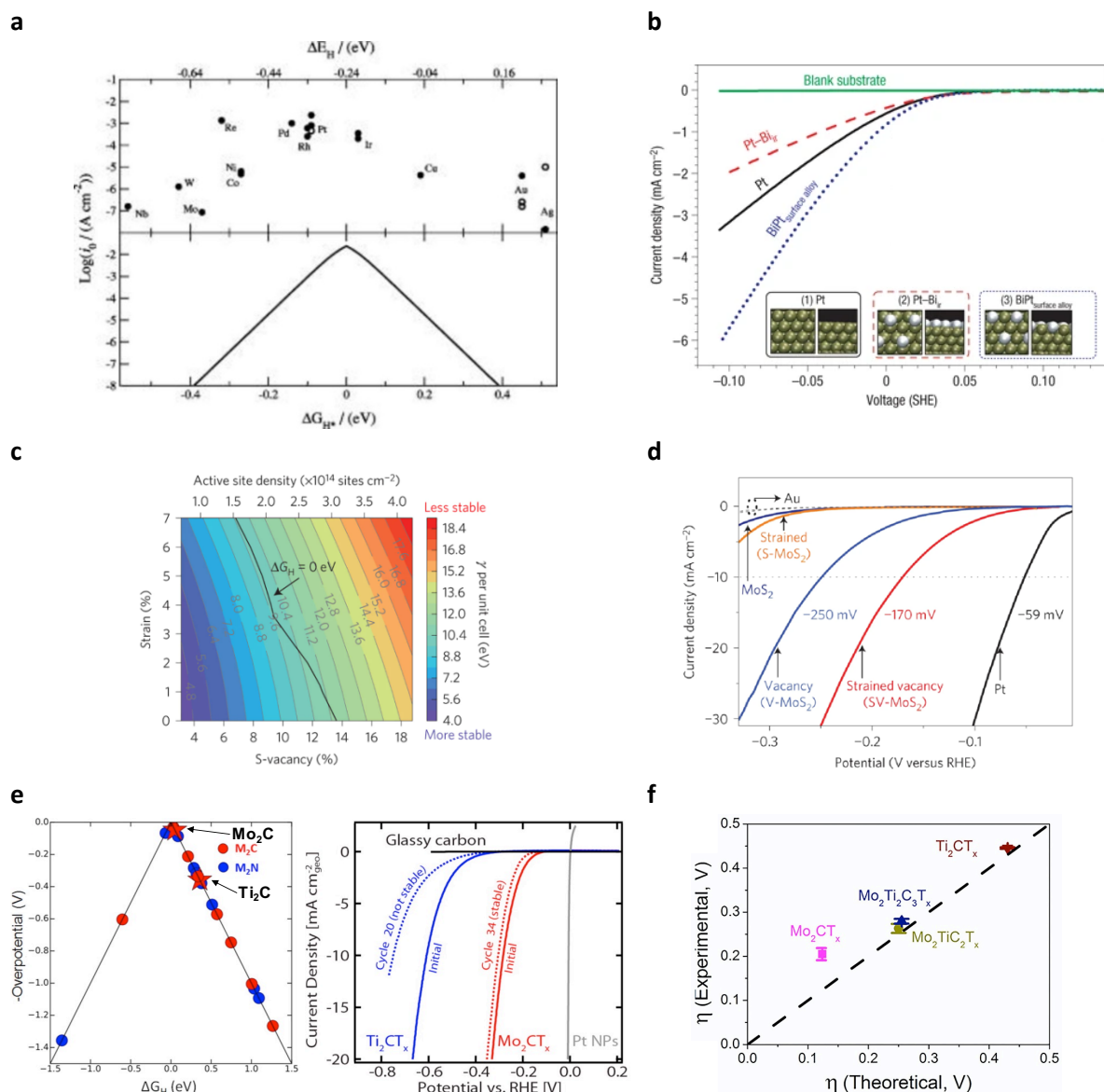


Figure 7 Reconciling theory with experiment results for HER electrocatalysts. (a) (top) Experimental i_0 plotted against theoretically calculated ΔG_H for various transition metal surfaces, compared against the proposed kinetic model (below). Open circles represent single crystal data. Reproduced from(46), copyright IOP science. (b) Experimental HER data of Pt (black), BiPt alloy precursor (Pt-Bi₁₀, red dash) and the synthesized BiPt alloy (blue dots). Modified from(47), copyright Nature Publishing Group. (c) Colour contour plot of stability (measured by surface energy per unit cell γ) with varying levels of strain and S vacancies. Black line illustrates combinations of strain and S vacancies for $\Delta G_H = 0$. (d) HER data of Au substrate, Pt control and 2H-MoS₂ (pristine and with various treatments). (c) and (d) reproduced from (229), copyright Nature Publishing

Group. (e) Theoretical volcano plot of MXenes (Ti_2C and Mo_2C marked out with red stars) and HER activity of Ti_2CT_x and Mo_2CT_x . Modified from(246), copyright American Chemical Society. (f) Comparison of theoretical and experimental overpotential at -10 mA cm^{-2} for various MXenes after considering $-F$ coverage on basal planes. Dashed line indicates complete agreement. Reproduced from(247), copyright American Chemical Society.

6. VALIDATION: WHAT MORE COULD BE DONE?

The development of advanced methods for the atomistic modelling of electrocatalysis is thriving. However, it is far from obvious to judge the general applicability and improvement of one approach compared to another. Furthermore, many of the most sophisticated approaches are basically designed ad hoc and used only in a couple of studies, which makes comparisons even harder. If, furthermore, the used software is either not public or system specific (e.g., only for Cu catalysts), the method should not be considered generally validated. From the experimental point of view, validation of theoretical (partial) current densities as a function of the potential are the most attractive quantities to reproduce. Such data is indeed increasingly available through online mass-spectrometry.(269) However, as argued in the introduction, these kinetic quantities necessitate the most complete theoretical models – or rely on the most drastic approximations. Hence, while appealing, these quantities are not the most suitable ones for the validation of the atomistic description of the processes at the interface. In contrast, for few systems the potential dependent atomistic structure of the interface is known. In this respect, EC-STM is most valuable, since it gives a direct depiction of the atomistic structure.(270–275) Alternative operando (potential dependent) spectroscopies give also precious information about surface state under operating conditions.(276,277) Reproducing (and interpreting) such data from first principles brings methods one step further towards validation. Nevertheless, most of these techniques do not directly probe the electrochemical reactivity. For instance, the Stark tuning slopes, which have been and still are, popular to validate theoretical models,(278) probe the local electrostatic field, the geometry of the adsorbate and the IR frequencies, but do not tell anything about the reactivity. The quantification of capacitance as a function of the electrochemical potential assesses, on the other hand, the realism of the description of the interface. Unfortunately, only few examples exist that are unequivocally associated with one single phenomenon, i.e., not affected by specific adsorption.(279,280) Again, while reassuring if well reproduced, the correct description of the capacitance in the double-layer regime does by no means validate the use of a given method to investigate the reaction mechanism. Over the decades, electrochemists have accumulated a fair quantity of potential dependent adsorption energies over single-crystal surfaces, e.g., for alcohols (281,282) halides, (283,284) sulfate, (285,286) phosphate, (287) uracil(288) and benzoic acid.(289) This data has so far not been exploited for method validation, last but not least because it puts stringent requirements on the compatibility and relative reliability of the solvation model and the description of the adsorbate—electrode interaction. This would, however, bring theoretical methods a significant step towards validation, as the description of reversibly bound adsorbates is an integral part of reaction mechanisms. Last but not least, the electrosorption values which are experimentally determined and give information on the exchange of electrons between the adsorbate and the electrode (electron reservoir) give direct access to the (re-)distribution of electrons during elementary steps (such as adsorption) and can, therefore, be considered fundamentally important quantities for electrocatalysis.(290) We suggest that a first principles-based method that describes potential dependent geometries, capacities and adsorption energies and electrosorption values quantitatively can be considered validated for the use of the elucidation of reaction mechanisms for electrocatalysis with increasing reliability in this order.

Conclusion

In this overview, we have highlighted the challenges that atomistic modelling of electrified interfaces and in particular electrocatalysis is facing. We have described the various methods that are currently applied to gain an atomistic understanding of reaction mechanisms at the electrode surfaces and reviewed the exemplary case of the hydrogen evolution reaction, where the complementary efforts of theory and experiment have proven most fruitful, leading to catalysts based on Earth-abundant materials that rival with the precious platinum benchmark catalysts.

Beyond the standard challenges in heterogeneous catalysis which are the limited accuracy of affordable density functional approximations for periodic, metallic systems and the choice of the active site model, the fundamental challenge in the field of first-principles based modelling of electrocatalysis is the difficulty to validate the models: experimentally, kinetic measurements are the most available data. However, between the atomistic picture of the electrode and the macroscopically measured kinetics, there is a long, hazardous connection, which can be seen like a novel: if we want to tell the entire story with all its details in order to provide a true understanding of the electrochemical mechanisms, we have to deploy extraordinary efforts. If, however, we are more interested in screening catalysts, we can skip through all but the first and last pages. These crude approximations are very powerful indeed, as they allow to explore the chemical space efficiently. In this review, we have highlighted the virtues and techniques of the complementary approach, which consists in the characterization of the protagonists (the processes at the electrified interface). We foresee a bright future for methods that combine extensive phase-space sampling with a control of the electrochemical potential. Assuming that the DFT computations provide sufficiently accurate energies for these complex systems, this combination will allow to clarify electrolyte and pH effects, but will also shed light on the fundamental principles of electrocatalysis, where the role and coupling between adsorbate and electrolyte solution dynamics remains to be uncovered.

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