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Structure and Magnetic Properties of an Original {Cu^{II}Mn^{II}W^V} Heterotrimetallic Coordination Polymer

Maria-Gabriela Alexandru,^[a] Diana Visinescu,^[b] Augustin M. Madalan,^[c] Rodolphe Clérac,^{*,[d]} and Marius Andruh^{*,[c]}

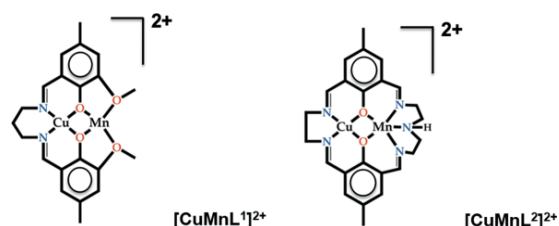
Dedicated to Professor Herbert W. Roesky, on the occasion of his 85th birthday

Abstract: A monodimensional heterotrimetallic coordination polymer of formula $\infty_1[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]_3\{(\mu\text{-NC})_3\text{W}^{\text{V}}(\text{CN})_5\}_2\cdot\text{H}_2\text{O}$ (**1**) has been assembled from pre-designed $[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]^{2+}$ cationic modules and homoleptic $[\text{W}^{\text{V}}(\text{CN})_8]^{3-}$ metalloligands (H_2L^2 is the compartmental Schiff-base resulting from the stepwise condensation of 2,6-diformyl-*p*-cresol with ethylenediamine and di-

ethylenetriamine). The magnetic properties of **1** have been investigated revealing a three-dimensional ordered antiferromagnetic ground state below 4 K resulting from different antiferromagnetic interactions between all the spin carriers ($S = 1/2$ Cu^{II}, $S = 5/2$ Mn^{II}, and $S = 1/2$ W^V centers).

Polynuclear complexes containing different paramagnetic metal ions are of high interest in molecule-based magnetism.^[1] The large majority of the heterometallic molecular systems are constructed from two different metal ions. Conversely, complexes with three different paramagnetic centers are significantly less common, mainly because of the difficulties related to the scrambling of the metal ions.^[2] A rational synthetic approach towards heterotrimetallic complexes consists of assembling preformed heterobimetallic modules and metalloligands.^[2] Anionic cyanido complexes, homo- and heteroleptic, are excellent metalloligands for building magnetic molecule-based materials containing two, or more, different metal ions.^[3] Our work in this field of research has been focused on the use of heterobimetallic tectons like phenoxido-bridged 3d-3d' and 3d-4f complexes, with the metal ions enclosed into different coordination sites generated by dissymmetric macrocyclic or

side-off compartmental ligands.^[4] Indeed, most of the known heterotrimetallic complexes are assembled using 3d-4f precursors, which are easily obtained by exploiting the differences between 3d and 4f metal ions [coordination numbers, HSAB – “hard and soft (Lewis) acids and bases” – behavior] and using side-off compartmental ligands.^[5] The synthesis of 3d-3d' magnetic building-blocks is more delicate, since the differences mentioned above are less prominent. One appropriate pair is composed of Cu^{II} and Mn^{II} ions, especially because Mn^{II} does not have a particular stereochemical preference, while Cu^{II} does. Robust binuclear Cu^{II}-Mn^{II} complexes have been generated using side-off compartmental ligands (like H_2L^1 see Scheme 1) with an open compartment that allows various coordination numbers and geometries for the Mn^{II} ion, while the inner compartment hosts a Cu^{II} metal ion with a square-planar or square-pyramidal geometry (Scheme 1).^[6]



Scheme 1.

For analogous bis-phenoxido-bridged Cu^{II}-Mn^{II} binuclear complex, $[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]^{2+}$ (Scheme 1), the metal ions are held together by a dissymmetric compartmental Robson-type ligand.^[7] Both binuclear systems are characterized by an $S = 2$ ground state, which results from a significant antiferromagnetic interaction between $S = 1/2$ Cu^{II} and $S = 5/2$ Mn^{II} centers, and are thus good building-block candidates for designing heterotrimetallic magnetic complexes and networks. For example, the self-as-

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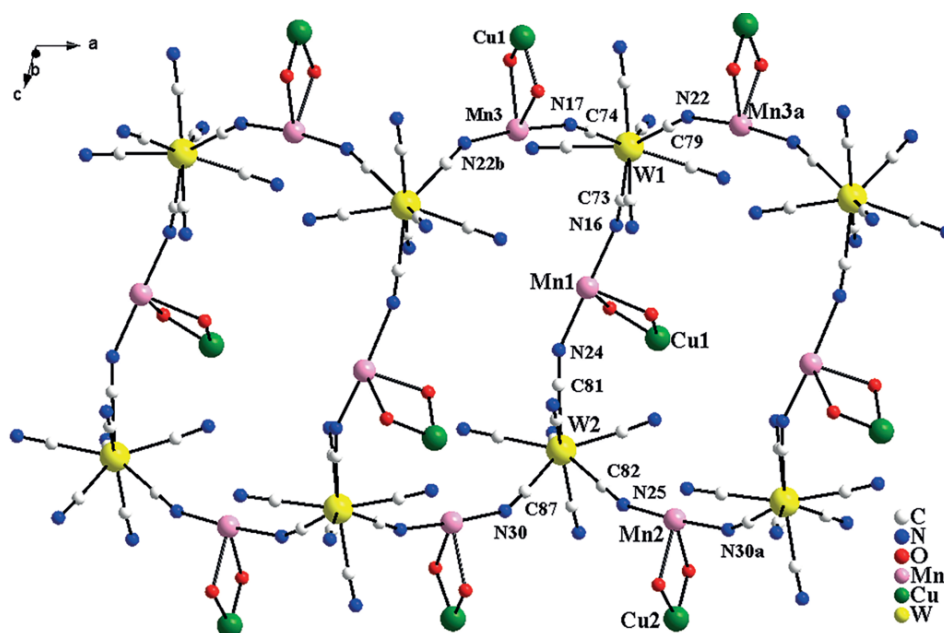


Figure 1. View of the crystal structure of **1** emphasizing the one-dimensional coordination polymer running along the *a* axis (symmetry operations: $a = 0.5 + x, 0.5 - y, z$; $b = -0.5 + x, 0.5 - y, z$).

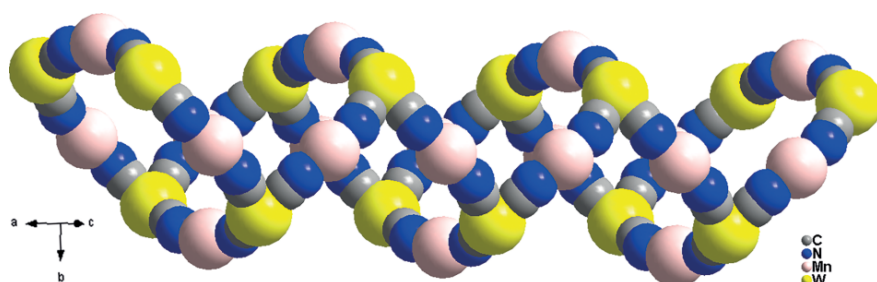


Figure 2. Perspective view of a wavy/accordion-like chain in **1** resulting from fused octanuclear $\{Cu^{II}-NC-W^V-CN\}_4$ metallacycles.

assembly of $[Cu^{II}Mn^{II}L^1]^{2+}$ and $[Cr^{III}(ampy)(CN)_4]^-$ modules (ampy = aminomethylpyridine) results in a 2-D 3d-3d'-3d'' network which possess a ferrimagnetic state below 4.3 K.^[6b] So far, only three heterotrimetallic chains, $[Cu^{II}Mn^{II}L^2M^{III}]_n$ ($M = Fe, Cr$)^[7a] and $[Cu^{II}Mn^{II}L^2W^V]_n$,^[7b] have been reported and obtained starting from $[Cu^{II}Mn^{II}L^2]^{2+}$ and heteroleptic cyanido metalloligands.

Herein, we are adding a new member to the small family of the 3d-3d'-3(4,5)d'' complexes. This coordination polymer, $\infty_1\{[Cu^{II}Mn^{II}L^2]_3[(\mu-NC)_3W^V(CN)_5]_2\} \cdot H_2O$ (**1**), which shows a very original one-dimensional topology, is obtained by layering equimolar acetonitrile solutions of $[Cu^{II}Mn^{II}L^2](ClO_4)_2$ and $(NHBu_3)_3[W^V(CN)_8]$ complexes (Supporting Information).

Compound **1** crystallizes in the monoclinic $P2_1/a$ space group. Its structure consists of neutral one-dimensional coordination polymer (Figure 1) and crystallization water molecules. The asymmetric unit is formed by three binuclear $\{Cu^{II}Mn^{II}L^2\}$ nodes and two $\{W^V(CN)_8\}$ spacers, each one connecting three binuclear units, through three cyanido groups (Figures 1 and S1, crystallographic details are given in Tables S1 and S2, CCDC 2000465). The resulting coordination network forms octanu-

clear $\{Cu^{II}-NC-W^V-CN\}_4$ metallacycles which are fused to form a wavy/accordion-like chain (Figure 2) running along the crystallographic *a* direction.

Both crystallographically independent W1 and W2 metal ions are coordinated by eight cyanido carbon atoms building a distorted square antiprism and a distorted dodecahedron geometry, respectively, according to SHAPE 2.1 calculations (Fig. S2, Table S3).^[8] The $\{W1(CN)_8\}$ spacer is connected, through the cyanido ligands, to Mn1, Mn3 and Mn3a atoms from the $\{Cu^{II}Mn^{II}L^2\}$ modules. The second cyanido metalloligand, $\{W2(CN)_8\}$, also adopts a *tris*-monodentate coordination mode to Mn1, Mn2, and Mn2b. Each of the three crystallographically independent binuclear modules contains one Cu^{II} ion, placed into the $\{N_2O_2\}$ compartment, and one Mn^{II} ion into the larger $\{N_3O_2\}$ compartment of the Schiff-base ligand. The copper(II) atoms are tetracoordinated by two imino nitrogen atoms [the shortest and longest bond lengths being 1.885(9) Å for Cu3-N12, and 1.929(10) Å for Cu3-N11], and by two phenoxido oxygen atoms [the shortest and longest bond lengths being 1.896(7) Å for Cu3-O5 and 1.928(8) Å for Cu1-O1] in a distorted square planar environment (Table S2).^[8] Focusing now on the

di-phenoxido bridged $\{\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\}$ fragment, the distances between the metal ions are: 3.249 Å for $\text{Cu1}\cdots\text{Mn1}$, 3.298 Å for $\text{Cu2}\cdots\text{Mn2}$, and 3.287 Å for $\text{Cu3}\cdots\text{Mn3}$. The angles at the bridging phenoxido atoms have values of 99.3(3)–101.8(3) for $\text{Cu1}-\text{O1}(\text{O2})-\text{Mn1}$, 101.6(3)–100.9(3) for $\text{Cu2}-\text{O3}(\text{O4})-\text{Mn2}$, and 101.6(3)–101.7(3)° for $\text{Cu3}-\text{O5}(\text{O6})-\text{Mn3}$. The Mn^{II} ions are heptacoordinated by two oxygen atoms [the shortest and longest bond lengths being 2.267(8) Å for $\text{Mn1}-\text{O2}$ and 2.351(8) for $\text{Mn2}-\text{O4}$], and by three nitrogen atoms belonging to the $\{\text{N}_3\text{O}_2\}$ compartment of the Schiff base ligand [the shortest and longest bond lengths being 2.244(11) Å for $\text{Mn2}-\text{N8}$ and 2.389(12) Å for $\text{Mn3}-\text{N14}$], and by two cyanido-nitrogen atoms from two $\{\text{W}^{\text{V}}(\text{CN})_8\}$ units. Their coordination geometries, determined using *SHAPE* 2.1 program, are capped trigonal prism for Mn1 , and pentagonal bipyramid for Mn2 and Mn3 (Fig. S3, Table S3). The $\text{Mn}-\text{N}_{\text{cyanide}}$ bond lengths are: 2.421(11) and 2.242(10) Å for $\text{Mn1}-\text{N16}$ and $\text{Mn1}-\text{N24}$, respectively; 2.276(10) and 2.279(10) Å for $\text{Mn2}-\text{N25}$ and $\text{Mn2}-\text{N30a}$; 2.286(10) and 2.272(10) Å for $\text{Mn3}-\text{N17}$ and $\text{Mn3}-\text{N22b}$ (symmetry operations: $a = 0.5 + x$, $0.5 - y$, z ; $b = -0.5 + x$, $0.5 - y$, z). The most relevant angles involving the cyanido bridging ligands are: 152.4(10) and 159.1(10)° for $\text{Mn1}-\text{N24}-\text{C81}$ and $\text{Mn1}-\text{N16}-\text{C73}$; 159.1(10) and 154.4(8)° for $\text{Mn2}-\text{N25}-\text{C82}$ and $\text{Mn2}-\text{N30a}-\text{C87a}$; 161.0(10) and 153.4(9)° for $\text{Mn3}-\text{N17}-\text{C74}$ and $\text{Mn3}-\text{N22b}-\text{C79b}$. Selected bond lengths and angles are gathered in Table S2. The geometrical parameters concerning the $\{\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}\}$ fragment are in agreement with similar reported examples.^[7b] The close inspection of the packing diagram reveals hydrogen bonds connecting the double chains as shown in Figures S4 and S5.

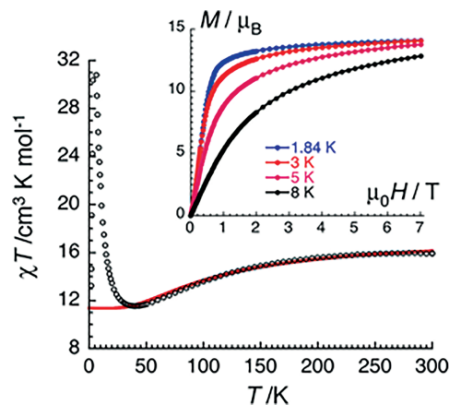


Figure 3. Temperature dependence of the χT product for **1** at 0.1 T applied dc fields (χ is defined as M/H per mole of **1**). Inset: Field dependence of the magnetization at indicated temperature with field sweep rates of 100–400 Oe min^{-1} for $\mu_0 H < 1$ T and 500–2500 Oe min^{-1} for $\mu_0 H > 1$ T.

The dc magnetic susceptibility has been measured for **1** between 1.85 and 300 K with magnetic field up to 7 T (Figure 3 and Figure 4). Consistent dc susceptibility at 0.1 T and in-phase ac susceptibility in zero-dc field (with an ac frequency of 1000 Hz) have been obtained between 1.85 and 15 K without detection of an out-of-phase ac signal, Figure S6. The χT product at room temperature is 15.9 $\text{cm}^3 \text{K mol}^{-1}$, which is perfectly consistent with the 3:3:2 $\text{Cu}^{\text{II}}/\text{Mn}^{\text{II}}/\text{W}^{\text{V}}$ stoichiometry of the $\infty_1[\{\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}_2\}_3\{(\mu\text{-NC})_3\text{W}^{\text{V}}(\text{CN})_5\}_2\cdot\text{H}_2\text{O}]$ complex, the three differ-

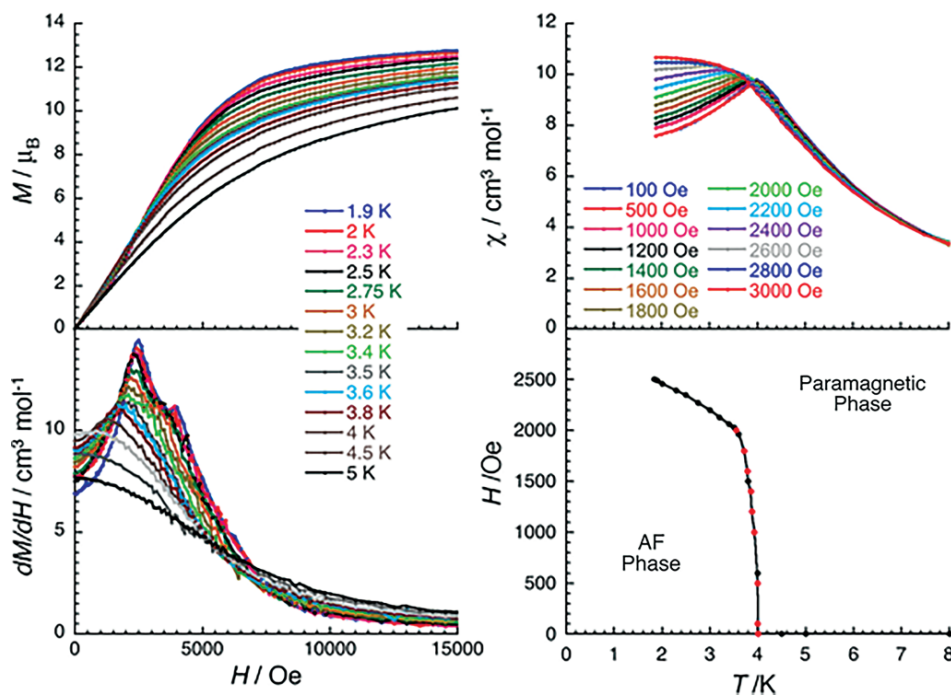


Figure 4. Low temperature magnetic properties for **1**: (top left) M vs. H and (bottom left) dM/dH vs. H plots with field sweep rates of 100–200 Oe min^{-1} ; (top right) χ vs. T plots; and (bottom right) magnetic phase diagram (H , T) constructed from χ vs. T (red) and M vs. H (black) data. Solid lines are visual guides for the eye.

ent types of paramagnetic centers ($S = 1/2$ / $S = 5/2$ / $S = 1/2$), and the expected Curie constant of $15 \text{ cm}^3 \text{ K mol}^{-1}$ with $C_{\text{Cu}} = 0.375 \text{ cm}^3 \text{ K mol}^{-1}$, $C_{\text{Mn}} = 4.375 \text{ cm}^3 \text{ K mol}^{-1}$ and $C_{\text{W}} = 0.375 \text{ cm}^3 \text{ K mol}^{-1}$ with an average g value of 2.0. The decrease of the χT product upon lowering the temperature down to $11.5 \text{ cm}^3 \text{ K mol}^{-1}$ at 40 K reveals the strongest antiferromagnetic interactions in **1** that are located between $S = 1/2 \text{ Cu}^{\text{II}}$ and $S = 5/2 \text{ Mn}^{\text{II}}$ magnetic centers within the $[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]^{2+}$ module as already observed with similar systems. The double-phenoxido-bridge is expected to mediate significant antiferromagnetic interaction, J/k_{B} , between Cu^{II} and Mn^{II} spins of the order of -22 to -53 K with $H = -2J_{\text{CuMn}} S_{\text{Cu}} S_{\text{Mn}}$.^[6b,6c,9] Using this Heisenberg heterospin Hamiltonian for the three $[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]^{2+}$ modules and treating the two $S = 1/2 \text{ W}^{\text{V}}$ centers as Curie-type spins, the experimental data down to 40 K can be well fitted allowing an estimation of J/k_{B} at $-27(1) \text{ K}$ with an average g factor of 2.17(5) (see solid red line in Figure 3). At lower temperatures, the χT product variation is the consequence of the smaller magnetic interactions which become relevant between the $S = 1/2 \text{ W}^{\text{V}}$ and $S = 5/2 \text{ Mn}^{\text{II}}$ spins through cyanido bridges, as well as weaker dipolar interactions. Below 40 K, the increase of the χT product (up to $31 \text{ cm}^3 \text{ K mol}^{-1}$ at 4.5 K) highlights the non-compensation of the metal ion spins and the growth of the magnetic correlations in the material. Below 4.5 K, the decrease of the χT product indicates the stabilization of an ordered antiferromagnetic state as demonstrated by the combined χ vs. T and M vs. H data shown in Figure 4.

The M vs. H data below 4 K show an inflection point at low fields below 3000 Oe, which is clearly detected on the dM/dH vs. H plots (Figure 4). This critical field, H_{C} (2500 Oe at 1.85 K), can be followed increasing temperature up to 4 K where it vanishes as expected in the case of a magnetic phase transition. Concomitantly, the dc magnetic susceptibility at a fixed dc field exhibits a maximum around 4 K corresponding to the Néel Temperature, T_{N} . The temperature dependence of H_{C} and the field dependence of T_{N} can be obtained from the dM/dH vs. H and χ vs. T data, respectively, in order to determine the (T, H) magnetic phase diagram shown in Figure 4. The topology of this phase diagram is typical of a metamagnetic material exhibiting a 3D ordered antiferromagnetic ground state. In the saturated paramagnetic state (Figure 3 inset) at high field, the magnetization at 1.85 K saturates above 5 T to about $14 \mu_{\text{B}}$ in agreement with the field alignment of the three $S = 2 [\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}^2]^{2+}$ units and two $S = 1/2 \text{ W}^{\text{V}}$ spins.

This molecule-based chemistry using pre-designed magnetic modules can be easily extended using other metalloligands, as well ancillary bridging ligands, in order to obtain new trimetallic networks with various dimensionalities and topologies incorporating larger magnetic interactions in order to stabilize high T_{C} magnets. Further work in these directions is currently in progress in our laboratories.

Deposition Number 2000465 (for **1**) contains the supplementary crystallographic data for this paper. These data are provided free of charge by the joint Cambridge Crystallographic Data Centre and

Fachinformationszentrum Karlsruhe Access Structures service www.ccdc.cam.ac.uk/structures.

Acknowledgments

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Supporting Information

Structure and Magnetic Properties of an Original $\{\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{W}^{\text{V}}\}$ Heterotrimetallic Coordination Polymer

Maria-Gabriela Alexandru, Diana Visinescu, Augustin M. Madalan,
Rodolphe Clérac,* and Marius Andruh*

Experimental part

Table S1. Crystal data and details of structure determination for compound **1**.

Table S2. Bond Lengths (Å) and angles (deg) of the environments of the copper(II), manganese(II) and tungsten(V) ions in **1**.

Figure S1. View of the asymmetric unit of **1** along with the atom numbering scheme.

Table S3. Results of the SHAPE1 analysis for the [W(CN)₈], [MnO₂N₅] and [CuO₂N₂] fragments.

Figure S2. Coordination sphere of the W metal ion illustrating the distorted square antiprism coordination geometry for W1 and trigonal dodecahedron for W2 in **1**.

Figure S3. Coordination sphere of the Mn1, Mn2 and Mn3 sites in **1**.

Figure S4. Supramolecular interactions: fragment of the crystal structure of **1** highlighting the hydrogen bond network established between the water molecules and the heterotrimetallic chains.

Figure S5. Supramolecular interactions: chains running along *a* axis, connected by hydrogen bonds shown in Figure S4.

Figure S6. Temperature dependence in zero-dc field of the in-phase and out-of-phase components of ac susceptibility (blue dots) for **1**.

References

Experimental part

Materials. The chemicals and solvents used were of reagent grade and they were purchased from commercial sources. The starting compounds, $[\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}](\text{ClO}_4)_2 \cdot \text{H}_2\text{O} \cdot \text{MeOH}$,¹⁻³ and $(n\text{-Bu}_3\text{NH})_3[\text{W}(\text{CN})_8]$,⁴ were synthesized by following the literature procedures ($n\text{-Bu}_3\text{NH}^+$ = tributyl ammonium cation).

Synthesis of $\infty_1[\{\text{Cu}^{\text{II}}\text{Mn}^{\text{II}}\text{L}\}_3\{\mu\text{-NC}\}_3\text{W}^{\text{V}}(\text{CN})_5\}_2] \cdot \text{H}_2\text{O}$ (1**).** A solution of $(n\text{Bu}_3\text{NH})_3[\text{W}(\text{CN})_8]$ (20 mg, 0.02 mmol) in acetonitrile (15 cm³) was poured slowly into an acetonitrile solution (20 cm³) containing 16 mg $[\text{CuMnL}](\text{ClO}_4)_2 \cdot \text{H}_2\text{O} \cdot \text{CH}_3\text{OH}$ (0.02 mmol). The resulting purple solution was left undisturbed for one week. X-ray quality purple crystals were obtained. Yield: *ca.* 70 %. Anal. Calcd for $\text{C}_{88}\text{H}_{83}\text{Cu}_3\text{N}_{31}\text{Mn}_3\text{O}_7\text{W}_2$ (**1**): C, 43.82; H, 3.44; N, 18.01%. Found: C, 43.45; H, 3.39; N, 18.37%. IR (KBr, cm⁻¹): 3340m, $\nu(\text{C}\equiv\text{N})$: 2120s, $\nu(\text{C}=\text{N})$: 1635m, 1607m, 1473s, 1304m, 1223s, 1073m, 742m.

Physical measurements

Elemental analysis was carried out using a EuroVector model EA3000 CHNS elemental analyzer. The IR spectrum of **1** was recorded on a FTIR Bruker Tensor V-37 spectrophotometer using KBr pellets in the 4000–400 cm⁻¹ range at room temperature. Magnetic measurements were carried out with a Quantum Design SQUID magnetometer MPMS-XL. This magnetometer works between 1.8 and 400 K for dc applied fields ranging from -7 to 7 T. Ac susceptibility measurements were performed with a Quantum Design PPMS-9 susceptometer working with an oscillating ac field of 1 Oe with a frequency between 10 and 10 000 Hz. These magnetic measurements were performed on a microcrystalline sample sealed in a polyethylene bag (3 Å ~ 0.5 Å ~ 0.02 cm; typical mass of 15–30 mg). Magnetic data were corrected for the sample holder and diamagnetic contributions for all magnetic measurements.

Crystallographic data collection and structure determination

Suitable single crystals of compound **1** were mounted on an IPDS II STOE diffractometer. Diffraction data for compound **1** were collected at 293(2) K using graphite-monochromated Mo-K α radiation ($\lambda = 0.71073$ Å). Multiscan absorption corrections were applied. The structures were solved by direct methods using SHELXS-2014 and refined by means of least squares procedures using SHELXL-2018. Atomic scattering factors were taken from the international tables for X-ray crystallography. The hydrogen atoms of the organic ligand were refined using a riding model, although most of the water hydrogen atoms were refined freely (geometrical restraints were used for the refinement of some of them). Drawings of the molecule were performed with the program Diamond 4. A summary of the crystallographic data and the structure refinement is given in Table S1. CCDC reference number: 2000465.

Table S1. Crystal data and details of structure determination for compound **1**.

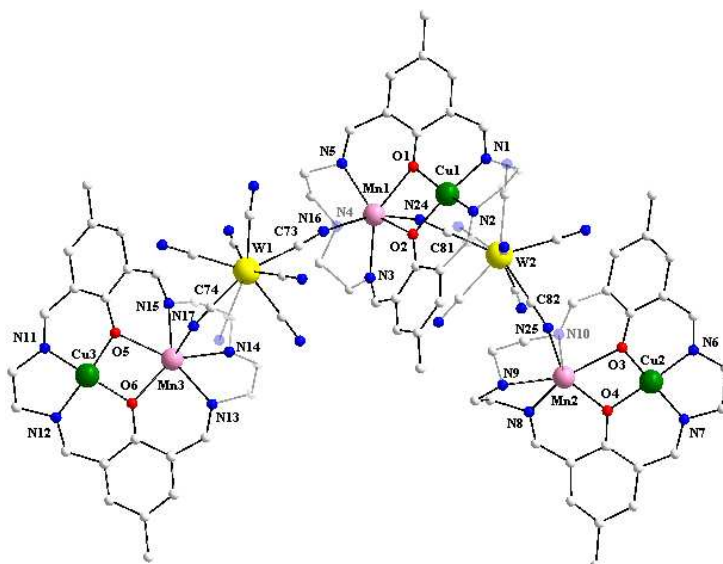
Formula	C ₈₈ H ₈₃ N ₃₁ O ₇ Cu ₃ Mn ₃ W ₂
Formula weight	2409.99
<i>T</i> (K)	293(2)
Crystal system	Monoclinic
Space group	<i>P</i> 2 ₁ / <i>a</i>
<i>a</i> (Å)	14.6629(6)
<i>b</i> (Å)	29.3757(14)
<i>c</i> (Å)	22.7532(9)
β (°)	105.685(3)
<i>V</i> (Å ³)	9435.6(7)
<i>Z</i>	4
<i>D</i> _{calcd} (g·cm ⁻³)	1.697
μ (mm ⁻¹)	3.548
Goodness-of-fit, <i>S</i>	0.960
^a Final <i>R</i> indices [<i>I</i> ≥ 2σ(<i>I</i>)] ^a	<i>R</i> ₁ = 0.0704 <i>wR</i> ₂ = 0.1469
^b <i>R</i> indices (all data) ^b	<i>R</i> ₁ = 0.1383 <i>wR</i> ₂ = 0.1682

$$^{[a]} R_1 = \sum (|F_0| - |F_c|) / \sum |F_c|, \quad ^{[b]} wR_2 = \{\sum [w(F_0^2 - F_c^2)^2] / \sum [w(F_0^2)^2]\}^{0.5}$$

Table S2. Bond Lengths (Å) and angles (deg) of the environments of the copper(II), manganese(II) and tungsten(V) ions in **1**.

1									
Cu1-N1	1.881(10)	Mn2-O4	2.351(8)	N7-Cu2-O4	94.8(4)	N24-Mn1-N3	85.7(4)	N9-Mn2-O4	139.3(4)
Cu1-N2	1.895(11)	Mn3-O5	2.330(8)	N11-Cu3-O5	95.2(4)	N24-Mn1-N4	81.5(4)	N10-Mn2-O3	76.5(4)
Cu2-N6	1.914(10)	Mn3-O6	2.314(7)	N11-Cu3-O6	166.3(4)	N24-Mn1-N5	107.7(4)	N10-Mn2-O4	143.1(4)
Cu2-N7	1.899(10)	W1-C73	2.191(12)	N12-Cu3-O5	167.4(4)	N25-Mn2-N30a	175.1(4)	N25-Mn2-O3	82.8(3)
Cu3-N11	1.929(10)	W1-C74	2.175(10)	N12-Cu3-O6	95.4(4)	N25-Mn2-N8	86.4(4)	N25-Mn2-O4	105.1(3)
Cu3-N12	1.885(9)	W2-C81	2.153(12)	Cu1-O1-Mn1	99.3(3)	N25-Mn2-N9	93.1(4)	N30a-Mn2-O3	100.2(3)
Cu1-O1	1.928(8)	W2-C82	2.173(13)	Cu1-O2-Mn1	101.8(3)	N25-Mn2-N10	82.5(4)	N30a-Mn2-O4	79.7(3)
Cu1-O2	1.909(7)	W1-C75	2.146(13)	Cu2-O3-Mn2	101.6(3)	N30a-Mn2-N8	93.7(4)	N13-Mn3-O5	140.8(3)
Cu2-O3	1.902(7)	W1-C76	2.201(14)	Cu2-O4-Mn2	100.9(3)	N30a-Mn2-N9	82.4(4)	N13-Mn3-O6	78.0(3)
Cu2-O4	1.911(7)	W1-C77	2.150(14)	Cu3-O5-Mn3	101.6(3)	N30a-Mn2-N10	94.4(4)	N14-Mn3-O5	146.6(5)
Cu3-O5	1.896(7)	W1-C78	2.152(13)	Cu3-O6-Mn3	101.7(3)	N17-Mn3-N22b	173.0(4)	N14-Mn3-O6	143.3(5)
Cu3-O6	1.912(7)	W1-C79	2.191(11)	N3-Mn1-N4	74.8(5)	N17-Mn3-N13	85.9(4)	N15-Mn3-O5	73.8(4)
Mn1-N3	2.260(12)	W1-C80	2.123(13)	N4-Mn1-N5	70.9(5)	N17-Mn3-N14	92.1(5)	N15-Mn3-O6	138.4(4)
Mn1-N4	2.376(13)	W1-C82	2.173(13)	N3-Mn1-N5	140.5(4)	N17-Mn3-N15	87.2(4)	N17-Mn3-O5	83.6(3)
Mn1-N5	2.251(10)	W2-C83	2.126(16)	N8-Mn2-N9	69.3(5)	N22b-Mn3-N13	91.0(4)	N17-Mn3-O6	105.5(3)
Mn2-N8	2.244(11)	W2-C84	2.131(12)	N9-Mn2-N10	74.3(5)	N22b-Mn3-N14	81.0(5)	N22b-Mn3-O5	102.7(3)
Mn2-N9	2.388(12)	W2-C85	2.165(12)	N8-Mn2-N10	141.2(4)	N22b-Mn3-N15	91.5(4)	N22b-Mn3-O6	80.0(3)
Mn2-N10	2.263(12)	W2-C86	2.132(16)	N13-Mn3-N14	71.2(5)	N3-Mn1-O1	143.6(4)	O1-Mn1-O2	70.0(3)
Mn3-N13	2.275(11)	W2-C87	2.153(11)	N14-Mn3-N15	73.0(5)	N3-Mn1-O2	77.6(4)	O3-Mn2-O4	69.0(3)
Mn3-N14	2.389(12)	W2-C88	2.145(17)	N13-Mn3-N15	143.2(4)	N4-Mn1-O1	134.1(5)	O5-Mn3-O6	68.8(3)
Mn3-N15	2.292(11)	N1-Cu1-N2	83.4(5)	Mn1-N24-C81	152.4(10)	N4-Mn1-O2	152.0(4)	W1-C73-N16	174.7(11)
Mn1-N16	2.421(11)	N6-Cu2-N7	83.7(5)	Mn1-N16-C73	159.1(10)	N5-Mn1-O1	76.0(3)	W1-C74-N17	177.5(10)
Mn1-N24	2.242(10)	N11-Cu3-N12	85.3(5)	Mn2-N25-C82	159.1(10)	N5-Mn1-O2	136.6(4)	W2-C81-N24	178.1(11)
Mn2-N30a	2.279(10)	N1-Cu1-O1	95.9(4)	Mn2-N30a-C87a	154.4(8)	N16-Mn1-O1	117.7(3)	W2-C82-N25	177.3(10)
Mn2-N25	2.276(10)	N1-Cu1-O2	171.5(4)	Mn3-N17-C74	161.0(10)	N16-Mn1-O2	86.8(3)	W1-C75-N18	175.8(11)
Mn3-N17	2.286(10)	N2-Cu1-O1	160.6(4)	Mn3-N22b-C79b	153.4(9)	N24-Mn1-O1	79.3(3)	W1-C76-N19	177.1(14)
Mn3-N22b	2.272(10)	N2-Cu1-O2	96.9(4)	N16-Mn1-N24	161.0(3)	N24-Mn1-O2	91.8(3)	W1-C77-N20	177.5(18)
Mn1-O1	2.323(7)	N6-Cu2-O3	96.3(4)	N16-Mn1-N3	75.5(4)	N8-Mn2-O3	138.8(4)	W1-C78-N21	175.4(12)
Mn1-O2	2.267(8)	N6-Cu2-O4	165.8(4)	N16-Mn1-N4	91.0(4)	N8-Mn2-O4	75.8(3)	W1-C79-N22	176.4(10)
Mn2-O3	2.338(8)	N7-Cu2-O3	167.5(4)	N16-Mn1-N5	85.9(4)	N9-Mn2-O3	150.7(4)	W1-C80-N23	177.9(14)

Symmetry codes: $a = 1/2+x, 1/2-y, z$; $b = -1/2+x, 1/2-y, z$



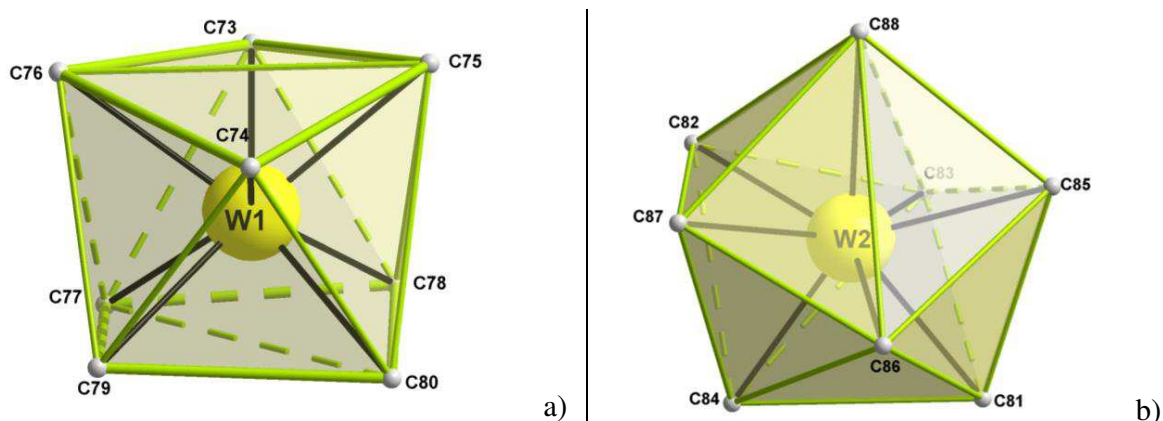


Figure S2. Coordination sphere of the W metal ion illustrating the distorted square antiprism coordination geometry for W1 and trigonal dodecahedron for W2 in **1**.

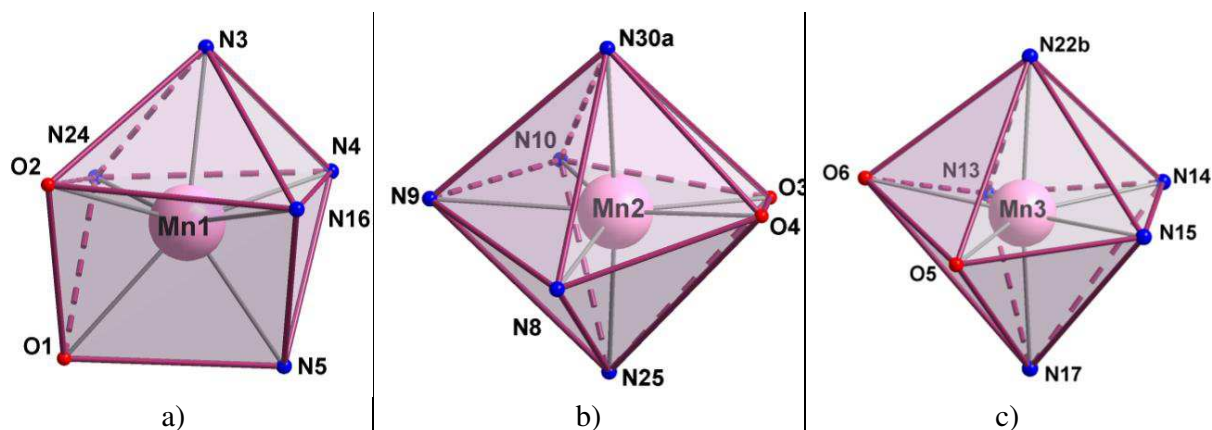


Figure S3. Coordination sphere of the Mn1, Mn2 and Mn3 sites in **1**. A distorted monocapped trigonal prism could best describe the coordination environment of Mn1 (Figure S3a). The imino N3 atom occupies the capping position of the rectangular face (O2-N24-N4-N16). The geometry of the Mn2 site is a distorted pentagonal bipyramid, with the pentagonal plane formed by O3, O4 phenoxido, N8, N10 imino, and N9 amino atoms, which belong to the second compartment of the organic ligand, the apical positions being occupied by cyanido nitrogen atoms, N25 and N30a (Figure S3b). The Mn2, N25 and N30a atoms are almost collinear, the N25-Mn2-N30a angle being of 175.1(3)°. where N16 and N24 atoms form the diagonal of the capped face. The geometry around Mn3 metal ion could also be described as a distorted pentagonal bipyramid, with N13, N14, N15, O5, O6 atoms, forming the basal plane, and the cyanido N17 and N22b atoms occupying the apical positions (Figure S3c). Symmetry codes: $a = 1/2+x, 1/2-y, z$; $b = -1/2+x, 1/2-y, z$

Supramolecular interactions

The water molecules establish hydrogen bonds with the nitrogen cyanido atoms from two neighboring chains, Figs S4 and S5.

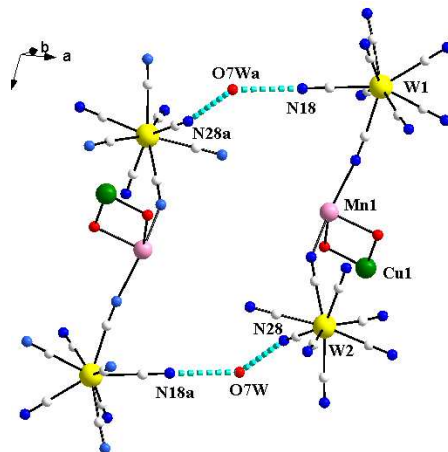


Figure S4. Supramolecular interactions: fragment of the crystal structure of **1** highlighting the hydrogen bond network established between the water molecules and the heterotrimetallic chains. The water molecules establish hydrogen bonds with the nitrogen cyanido atoms from two neighboring chains shown in Figure S5. Symmetry codes: $a = 1-x, -y, 2-z$.

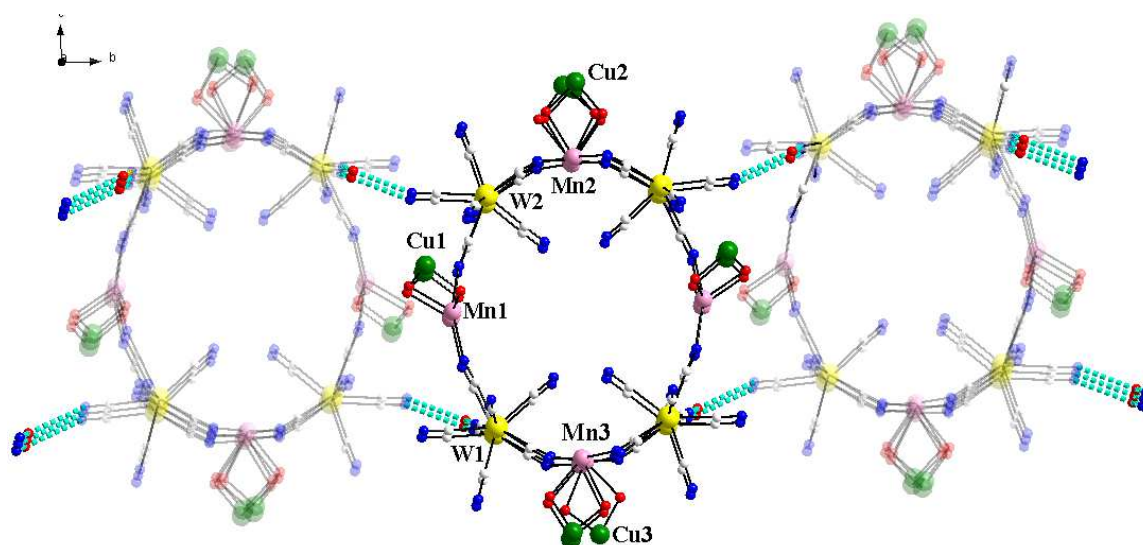


Figure S5. Supramolecular interactions: chains running along a axis, connected by hydrogen bonds shown in Figure S4.

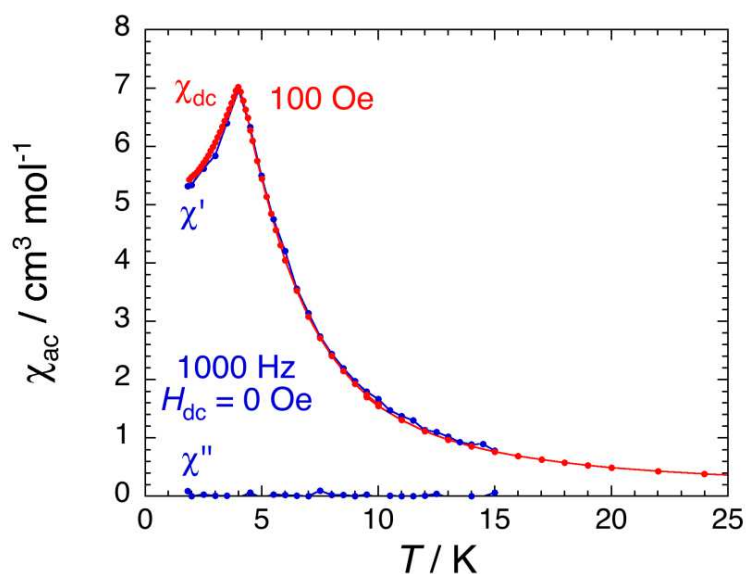


Figure S6. Temperature dependence in zero-dc field of the in-phase and out-of-phase components of ac susceptibility (blue dots) for **1** between 25 and 1.85 K at 1000 Hz with an amplitude of the ac field of 3 Oe. For comparison, the temperature dependence of the dc susceptibility at 100 Oe is shown in red.

References

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