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Interdiffusion behaviors observation in TiN/ZrO_xN_y bilayer by XAS and ToF-SIMS

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Abstract

Interface behavior plays an important role in multilayer structure. The deep understanding of interface diffusion and reaction is necessary to improve the performance of device with multilayer structure. In the present work, we perform a detailed systematic study on interdiffusion and interaction between the TiN and ZrO_xN_y layers in the TiN/ZrO_xN_y bilayer system. A series of TiN/ZrO_xN_y bilayers with reverse deposition orders or with varied composition have been prepared by radio-frequency reactive magnetron sputtering. Synchrotron radiation x-ray absorption spectroscopy (XAS) and time of flight secondary ion mass spectrometry (ToF-SIMS) are both used to understand well interface diffusion behavior. The XAS results indicate that interaction and interdiffusion extent between TiN and ZrO₂

(ZrO_xN_y) layers depend on deposition order of both layers. Herein the O atoms enter easily into the TiN layer by substituting the N atoms when the ZrO_2 or ZrO_xN_y layer is deposited onto the TiN layer. No significant interaction between TiN and ZrN layers takes places in the TiN/ZrN sample. Both XAS results and ToF-SIMS depth profiles reveal that the interaction between TiN and ZrO_xN_y layers depends on the deposition order of the layers. TiN is more easily oxidized when ZrO_2 or ZrON is deposited onto TiN. In addition, the resulting stacks are not significantly different if ZrON is used instead of ZrO_2 , revealing most probably the full oxidation of ZrON.

Keywords: TiN/ ZrO_xN_y ; X-ray absorption spectroscopy (XAS); Time of Flight secondary ion mass spectrometry (ToF-SIMS); Interdiffusion

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1. Introduction

Transition metal nitrides, such as TiN, ZrN, and CrN are widely used as protective coatings for cutting and forming tools due to their high hardness and wear resistance [1-4]. Some multicomponent coatings are also designed to further provide the performance of hard coatings, like TiYN, TiCrN, TiWAlN, TiMoSiN [5-10]. Recently periodic multilayer systems alternating (at least) two kinds of component materials have attract number of interests [11-13]. They are not only applied in protecting coatings but also in optics as well as other devices [14-19].

Nanoscale multilayer system can provide excellent performances due to their multi-interfaces structures, refined crystalline strengthening and periodic modulation. Regarding the hard protective coating, transition metal oxides and nitrides have been widely used to prepare periodic nano-multilayers owing to their excellent properties such as chemical stability and mechanical properties as well as oxidation resistance [20-25]. Among them TiN and ZrN crystallize in the face centered cubic (fcc) system and have similar physical properties such as high melting point, high hardness and chemical inertness. TiN-based and ZrN-based multilayer coatings have received amount of attentions [4, 26-32]. Multilayer structures consisting of transition metal oxynitrides show potential excellent performance and have been widely studied because they can combine or even exhibit better properties of both nitrides and oxides [33-36].

According to the hardening mechanism, the mechanical properties of hard coatings composed of two alternating materials are quite related to the coherent epitaxial growth of the two layers[26]. If the two layers can form coherent epitaxial growth with sharp interface, the hardness of coatings increases largely. However, the stable coherent epitaxial growth of oxides cannot be easily obtained meanwhile interfacial diffusion mechanisms between two layers has great influence on growth behavior.

The combination of transition metal nitrides and oxynitrides are proposed to form a periodic multilayer coating, with aim to not only improve the hardness but also the oxidation resistance at high temperature. It is important to know how the

physical-chemical reactive behaviors at the interfaces affect the mechanical properties of such a multilayer. Most published papers have mainly reported about their mechanical properties and phase structure [3, 15, 29, 37, 38]. The effects of layer growth order on the interface diffusion behaviors are rarely reported [39].

In this context, magnetron sputtering technique was used to prepare all samples. The present work focused on interface diffusion behaviors in the TiN/ZrO_xN_y multilayer coatings with reverse deposition order. In order to observe separately two kinds of interfaces, we have prepared a series of samples by varying oxygen and nitrogen mixture partial pressure during the deposition process. We have also prepared the same series of samples but with reverse orders: ZrO_xN_y-on-TiN and TiN-on-ZrO_xN_y. X-ray absorption near edge structure (XANES) spectra of all samples was measured to check their electronic and structural properties. Otherwise, time of flight secondary ion mass spectrometry (ToF-SIMS) elemental depth profiles are also recorded to analyze the interface information. This is of great significance for improving the preparation of the periodic multilayer coatings and their related properties of device.

2. Experimental details

2.1 Preparation of sample

To understand in detail the interaction of TiN and ZrO_xN_y layers, three series of bilayer samples: Si/TiN/ZrO₂, Si/TiN/ZrON, Si/TiN/ZrN, with corresponding reverse deposition order of layers, were prepared on Si (111) substrates using radio-frequency magnetron sputtering (JCP-500M3 Beijing Technol. Co. LTD., China) with targets of TiN (99.9%), ZrO₂ (99.95%) and ZrN (99.5%). The Si (111) substrates were sliced with an area of 20 mm×20 mm. The base pressure was 6.0×10⁻⁴Pa before deposition. The target–substrate distance was about 100 mm. The sputtering powers of the targets TiN, ZrO₂ and ZrN were set at 120, 100 and 100 W, respectively. The working gas was argon (99.999%) with a constant flow rate of 20.0 standard cubic centimeters per minute (sccm) when the TiN, ZrO₂ and ZrN layers were deposited. The ZrO_xN_y layers

were prepared using ZrO₂ target with argon (10 sccm) and nitrogen (10 sccm). TiN, ZrO₂ and ZrN single layers were also prepared as reference samples. Capping layers of B₄C were deposited on all the bilayer and single layer samples to prevent the oxidation of their surface. Regarding the bilayer, the expected thickness of each layer is around 2.0 nm. The thickness of the single layers (ZrO₂, TiN, ZrN) is around 15.0 nm. All samples were grown at room temperature.

2.2 X-ray absorption spectroscopy

X-ray absorption spectroscopy (XAS) is a powerful technique for probing the local electronic structure around a specific element in thin films and other structures[27]. X-ray absorption near edge structure (XANES) provides important information about the electronic and structural properties. Because XANES measurements probe the unoccupied density of states, the method is very sensitive to the local bonding environment, such as the number of valence electrons, their spin configurations, and the symmetry and coordination number of the structural unit cell of the transition metal compounds. In the case of buried interfaces, the phases are often formed in a narrow region, are highly disordered, or present a short-range order. Thus, the XAS technique with short-range order sensitivity provides a suitable tool for the study of interface phases, also called interphases.

XAS measurements of all the samples were performed on the BEAR beamline in the Elettra synchrotron radiation facility[40]. Before XAS measurements, we first performed sputtering on the samples for about 40 min using Ar⁺ (10⁻⁶ mbar, 3 keV, 45°, 1 μA), to remove the capping layer of B₄C. Then the B K edge was checked to be sure that the B₄C layer was removed enough. Thus, it was possible to get information from the buried interfaces. The absorption spectra were obtained by measuring the drain current of the samples, that is to say XAS measurements were performed in total electron yield mode. All the spectra were normalized (divided) by the I₀ curve obtained measuring the total electron yield on a clean gold sample and then to take into account features present in the spectra due to the contamination of the optical

elements (mirrors, gratings and so on) present on the beam line. This is a particularly important correction in the N K and O K ranges. We collected the Zr M_{2,3}, Ti L_{2,3}, N K, O K XANES spectra. The EXAFS measurements were measured above the O K edge, up to 650 eV.

2.3 Time of flight secondary ion mass spectrometry

ToF-SIMS has already been demonstrated to be a useful tool to obtain the elemental and chemical depth profiles in multilayers [19, 39, 41]. In this paper the elemental depth distributions of the bilayer samples, Si/TiN/ZrO₂, Si/ZrO₂/TiN, Si/TiN/ZrON and Si/ZrON/TiN, were investigated using a ToF-SIMS instrument (TOF.SIMS V spectrometer, ION-TOF GmbH) working in the dual-beam mode. The sputtering was performed using a 500 eV (25 nA) Cs⁺ ion beam, rastered over an area of 300 μm × 300 μm. A pulsed 25 keV Bi⁺ primary ion source at a current of 1.0 pA, rastered over a scan area of 100 μm × 100 μm was used as the analysis beam. Both ion beams were impinging the sample forming an angle of 45° with the surface normal and were aligned in such a way that the analyzed ions were taken from the center of the sputtered crater. Both positive and negative ions were recorded. The ToF-SIMS technique provides a qualitative chemical analysis as the intensity does not correspond directly to the number of atoms in the multilayer (the sputtering yields vary from one element to another as well as from one matrix to another).

3. Results and discussion

3.1 X-ray absorption spectroscopy

Figure 1 shows the XANES spectra of the Si/ZrO₂/TiN, Si/TiN/ZrO₂ bilayers and ZrO₂ and TiN single layers. As reported in Ref. [27] the features of XAS spectra in transition metal oxides and nitrides are related with number of *d* electrons, electronic configuration as well as crystal structure. In the following part, all these XAS spectra will be analyzed by considering these factors. Figure 1(a) presents the Zr M_{2,3} edge. The spectra of both bilayer samples show similar features but they are very different from that of ZrO₂ single layer. The bilayer samples have two main peaks around at

326.4 and 339.2 eV corresponding to the expected Zr $3p_{1/2}$ - $3p_{3/2}$ spin-orbit splitting of 13.7 eV. These two peaks around at 326.4 and 339.2 eV correspond to the Zr-O bond but two small peaks around at 341.6 and 344.7 eV cannot be identified directly. They should probably be ascribed to multiplet electron effect between the TiN and ZrO₂ layers. The Si/ZrO₂/TiN sample shows a strong intensity around these two small peaks compared with that of the Si/TiN/ZrO₂ sample. It means that interaction takes places when ZrO₂ layer is deposited onto the TiN layer or crystallization behavior of the TiN layer changes.

Regarding the N K edge, Figure 1(b), the three samples show similar features with an exception of the Si/ZrO₂/TiN sample which shows no peak around at 390.7 eV. According to the Ref. [27], the first two relatively sharp features A and B can be assigned to transitions to the t_{2g} (Ti3d + N2p π) and e_g (Ti3d + N2p σ) orbitals, respectively; while the higher energy broad feature D is related to the excitation to the (Ti4sp + N2p) states. The peak located around 390.7eV corresponds to the N-Ti2p bond.

Comparison of the Ti L_{2,3} XANES, Figure 1(c) obtained from two bilayers and TiN samples shows similar features but a peak shift of 1.4 eV toward higher energy for the Si/TiN/ZrO₂ sample. The Ti L_{2,3} spectra display a complex structure due to the combination of the atomic interactions and crystal field effect. The spectral features in the L_{2,3} region corresponds to O2p_{3/2}-Ti3d and O2p_{1/2}-Ti3d transitions, respectively. Furthermore, it can also be clearly seen that Ti L₃ and L₂ edges further split into t_{2g} (A1 and C1) and e_g (B1 and D1) states due to crystal field effect. Compared with the Ti L_{2,3} spectra of TiO₂ from the reference [42], the secondary peaks A1 and C1 around the main peaks in all samples are due to the oxidation of the TiN layer. However, the intensities of A1 and C1 relative to those of B1 and D1 in the Si/TiN/ZrO₂ sample is higher than that in the other two samples. It means that when the ZrO₂ layer is deposited onto the TiN layer the O atoms enter easily into the TiN layer by substituting the N atoms. Indeed, when the ZrO₂ layer is deposited onto the TiN layer, there are many oxygen energetic species in the chamber because the ZrO₂ target is in

operation. Therefore these energetic species easily interact with TiN layer to oxidize Ti. However in reverse mode, TiN layer is deposited onto ZrO₂ layer which is a very stable compound. Moreover the Zr and O atoms have a big electronegativity difference so O atoms in the ZrO₂ compound difficulty enter into the TiN layer.

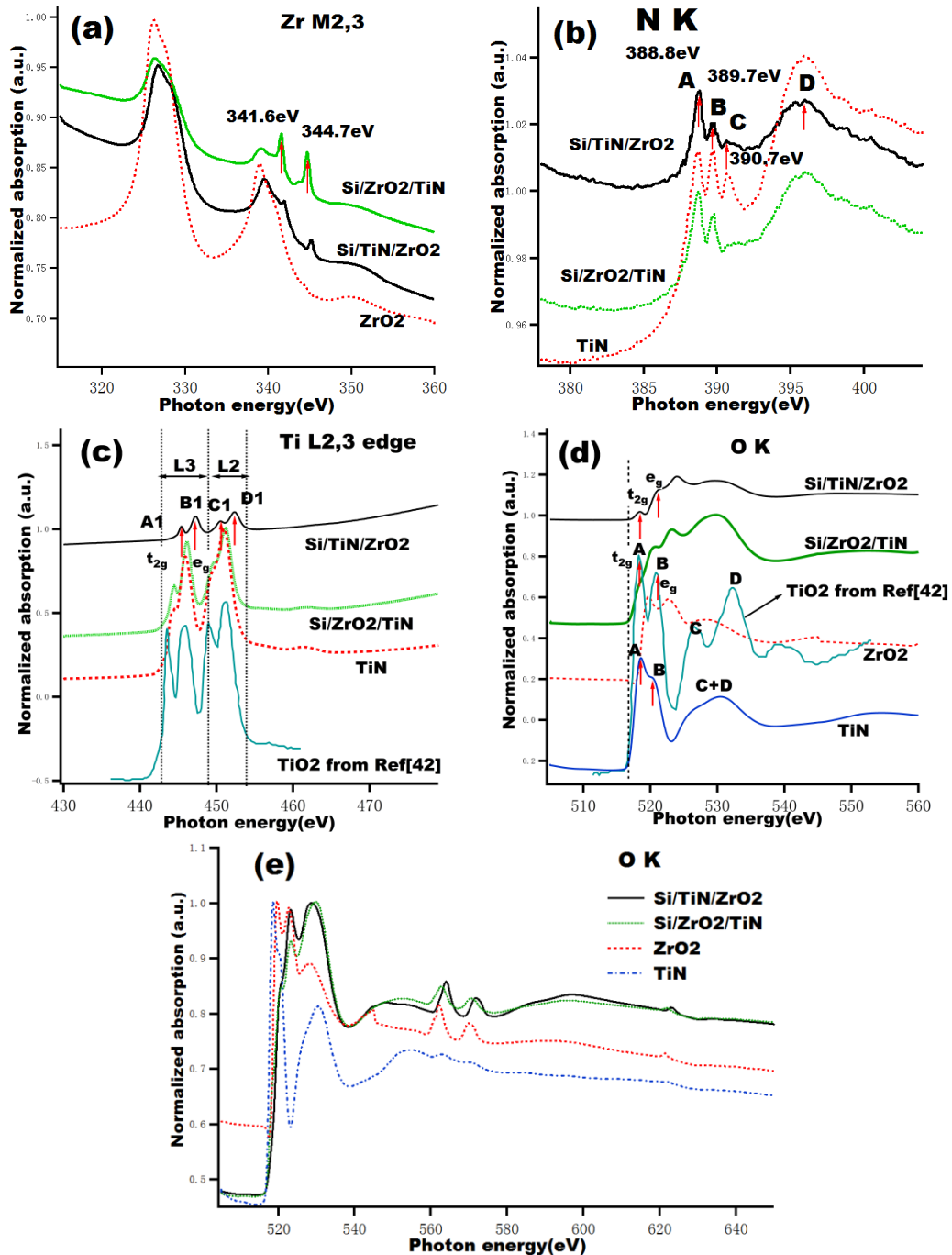
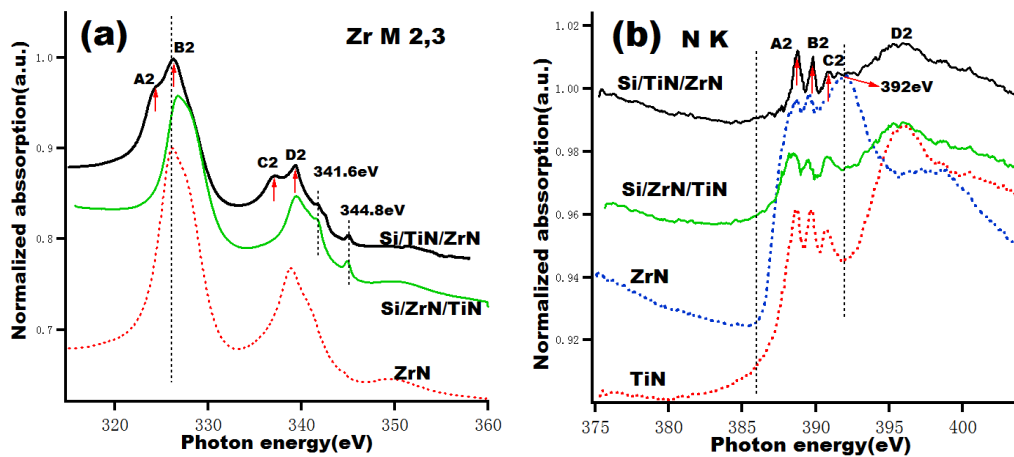


Figure 1. XANES (a-d) and EXAFS (e) spectra of the samples: Si/TiN/ZrO₂, Si/ZrO₂/TiN bilayers, TiN and ZrO₂ single layer samples. For the sake of clarity, some spectra are shifted vertically.

Figure 1(d) shows the O K spectra of two bilayer samples together with TiN and

ZrO₂ references and TiO₂ from Ref.[42]. Significant changes are obtained between single layer and bilayer samples. In principle, the TiN layer should not contain O atoms but here O K spectrum was observed because the bilayer sample was slightly oxidized. Ref. [27] reported that the number of O K edge features and their peak positions are very different from one metal oxide to another. It can be seen clearly that the features of two bilayer samples are neither same as those of TiN nor ZrO₂. It means that some interaction occurs at the interface between TiN and ZrO₂ layers. In the TiO₂ spectrum, peaks A and B can be assigned to transitions to the t_{2g} (Ti 3d + O2p π) and e_g (Ti3d + O2p σ) orbitals, respectively. The broad feature at higher energy is related to the excitation to the (Ti4sp + O2p) states. For TiN and ZrO₂ samples, the broad features are related to the excitation to the (Ti4sp + N2p) and (Zr4sp + O2p) states, respectively. Then, the broad features of two bilayer samples are related to the excitation to the (Ti4sp + N2p) and/or (Zr4sp + O2p) states.

Figure 1(e) shows the O K-edge EXAFS spectra. Two peaks located around at 562 and 570 eV appear in the two bilayer sample and ZrO₂ sample. They should be due to the Zr and O atoms interaction. It can be concluded from the comparison of the spectra of the bilayer and single layer samples that an interaction between TiN and ZrO₂ layers does take place. The bilayer with different deposition order presents different spectral changes indicating different interface behaviours.



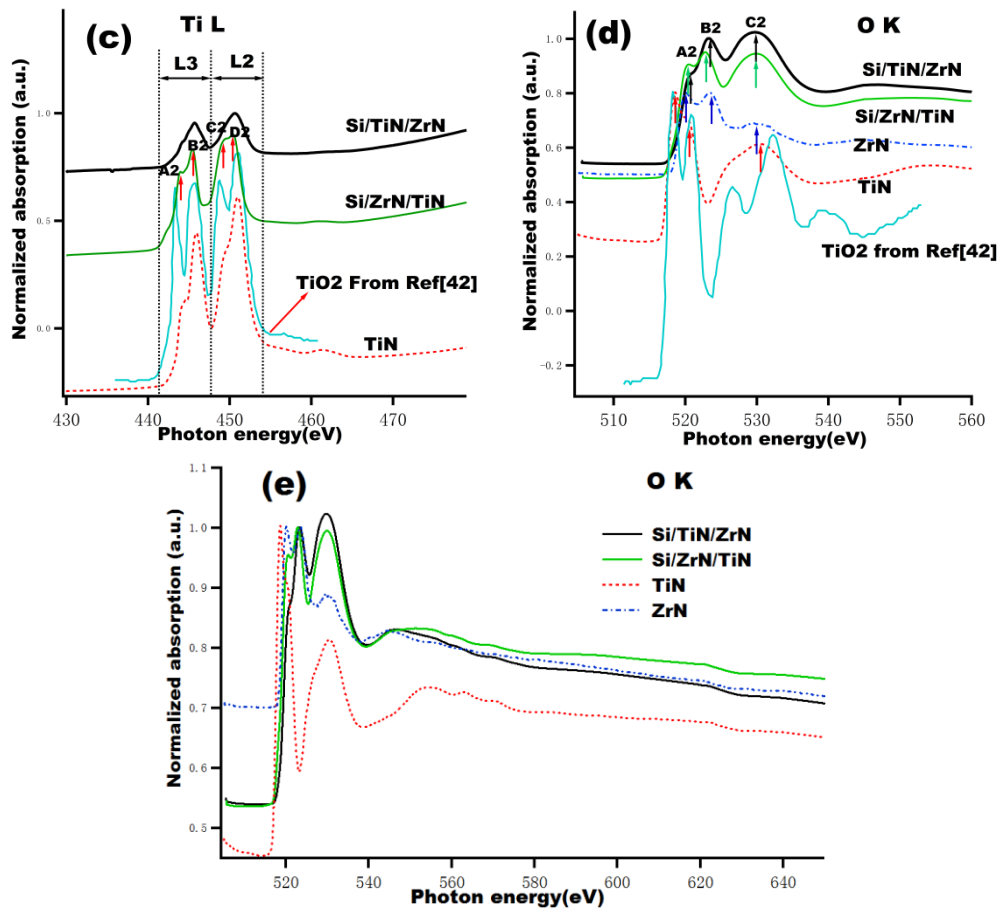


Figure 2. XANES (a-d) and EXAFS (e) spectra of the samples: Si/TiN/ZrN, Si/ZrN/TiN bilayers and TiN and ZrN single layer samples. For the sake of clarity, some spectra are shifted vertically.

Figure 2 shows the XAS spectra of the bilayers Si/TiN/ZrN, Si/ZrN/TiN as well as single layer TiN, ZrN samples. The Zr $M_{2,3}$ spectra, shown in Figure 2(a), of the bilayers is different from that of ZrN single layer. Some small peaks (341.6 and 344.8 eV) and shoulders are accompanied the two main peaks which probably correspond to the Ti-N bond. For the Si/TiN/ZrN sample, shoulders toward lower energy around two main peaks are observed. In addition, the positions of the two main peaks in the Si/ZrN/TiN sample shift to higher energy with respect to the two other samples. This indicates that the bonding environment of Zr atoms in the Si/ZrN/TiN sample is different from the other samples.

About the N K spectra, shown in Figure 2(b), they all have a similar shape with an exception of the ZrN sample. The features can be assigned to the excitation of N 1s electrons to the hybridized molecular orbitals of TiN or ZrN. As discussed earlier in the Figure 1(b) the first two relatively sharp features (A2 and B2) in the bilayer

samples and the TiN samples can be assigned to the N 2p orbitals hybridized with metal *nd* states, which are splitted into the t_{2g} (Ti/Zr3d + N2p π) and e_g (Ti/Zr3d + N2p σ) orbitals, respectively; the higher energy broad feature (D2) is related to the excitation to the (Ti/Zr4sp + N2p) states. The sharp feature located at round 390.7eV (C2) should correspond to the N-Ti/Zr 2p bond. However, for the ZrN sample, a new feature located at 392 eV appears with the narrow feature at higher energy. It is probably due to the bonding environment of N in the ZrN thin films. The broad feature D2 at higher energy of two bilayer samples has a lower intensity than that of TiN single film. This should be ascribed to the comprehensive bonding environment of N atoms in the ZrN and TiN layers. Perhaps some interaction takes place at the interface between ZrN and TiN layer.

Regarding the Ti L spectra in Figure 2(c), both bilayer samples show different features. As discussed in earlier, the L_3 and L_2 edges in the Si/ZrN/TiN sample further split into t_{2g} (A2 and C2) and e_g (B2 and D2) states due to crystal field effect.

About the O K spectra, Figure 2(d), the four samples present similar peak feature: A2, B2 and C2, but with different positions and intensities. The absorption edge position changes as a function of samples. The Si/ZrN/TiN sample shows two first sharp peaks similar as those of the ZrN single layer. The position of the first two features in the Si/TiN/ZrN samples shifts towards higher energy compared with TiN single layer. The two peak features in TiN layer are similar to the two first features of TiO₂ from Ref. [42]. It means that a little bit of oxidation occurs in the TiN layer. The three other samples behave in a similar way.

In Figure 2(e) no obvious changes of EXAFS spectra happen on the all samples meaning that the distance of interstitial atoms and coordination number in the ZrN and TiN layer do not change.

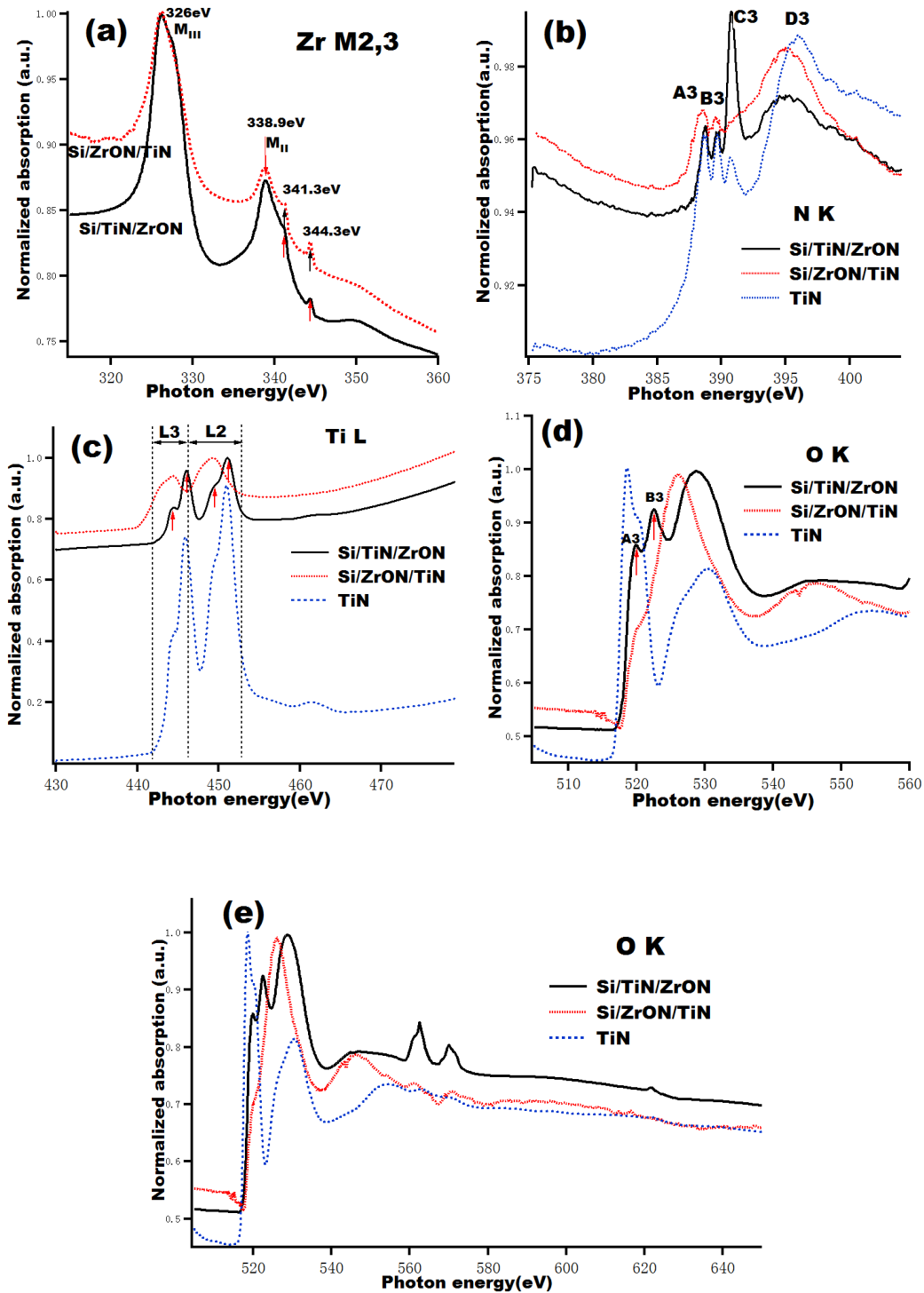


Figure 3. XANES (a-d) and EXAFS (e) spectra of the samples: Si/TiN/ZrON, Si/ZrON/TiN bilayers and TiN single layer samples.

Figure 3 shows the XAS spectra of the Si/TiN/ZrON, Si/ZrON/TiN and TiN samples. The Zr M_{2,3} edge of two bilayer samples, shown in Figure 3(a), have similar features but with different peak intensities. In addition, at higher energy, as discussed earlier, both bilayer samples show two weak peaks around at 341.3 and 344.3 eV

which can be ascribed to the multiplet electron effect between TiN and ZrON layers.

In Figure 3(b), each sample shows a similar first two A3 and B3 features. They are due to the transitions of the t_{2g} (Ti/Zr 3d + N 2p π) and e_g (Ti/Zr3d + N2p σ) orbitals, respectively. But the Si/TiN/ZrON sample possesses a strong C3 peak ascribed to the Zr-N bond [27]. Furthermore the crystal structure of both layers also affects the spectra shape because of multi-electron configuration interaction.

Figure 3(c) shows the Ti L spectra of the three samples. It can be seen clearly that the Si/TiN/ZrON sample gives rise to similar four features as TiO₂ shown in Figure 1(c), which indicates the slightly oxidation of TiN layer. The peak position of the Si/ZrON/TiN sample shifts towards lower energy indicating a different crystal structure of both bilayer samples.

In Figure 3(d), the Si/TiN/ZrON sample gives rise to significant sharp features of O K edge which is related to the hybridization of the O 2p orbital with metal 3d states, with a broad structure in higher energy. However, the Si/ZrON/TiN sample does not show a similar behavior. This means that these three features can be attributed to oxidation of the TiN layer.

The O K edge EXAFS spectra are presented in Figure 3(e). They greatly change from one sample to another. The O K absorption edge of the Si/ZrON/TiN sample shifts toward higher energy. The bilayer samples present different spectra. Furthermore, the peaks around at 562.5 and 570.0 eV should be due to the interaction between the Zr and O atoms. All these changes demonstrate again that the O atoms easily diffuse into the TiN layer only when ZrO_xN_y layers are deposited onto TiN layer.

From the above analysis, it can be seen that different deposition orders of the thin layers lead to different spectral features. This means that the interface atoms environment is affected by the deposition order of the two same layers. The first deposited layer can change the crystallization behavior of post-deposition layer as well as accompanies the multiplet electrons effects and electronic configuration change at interfaces.

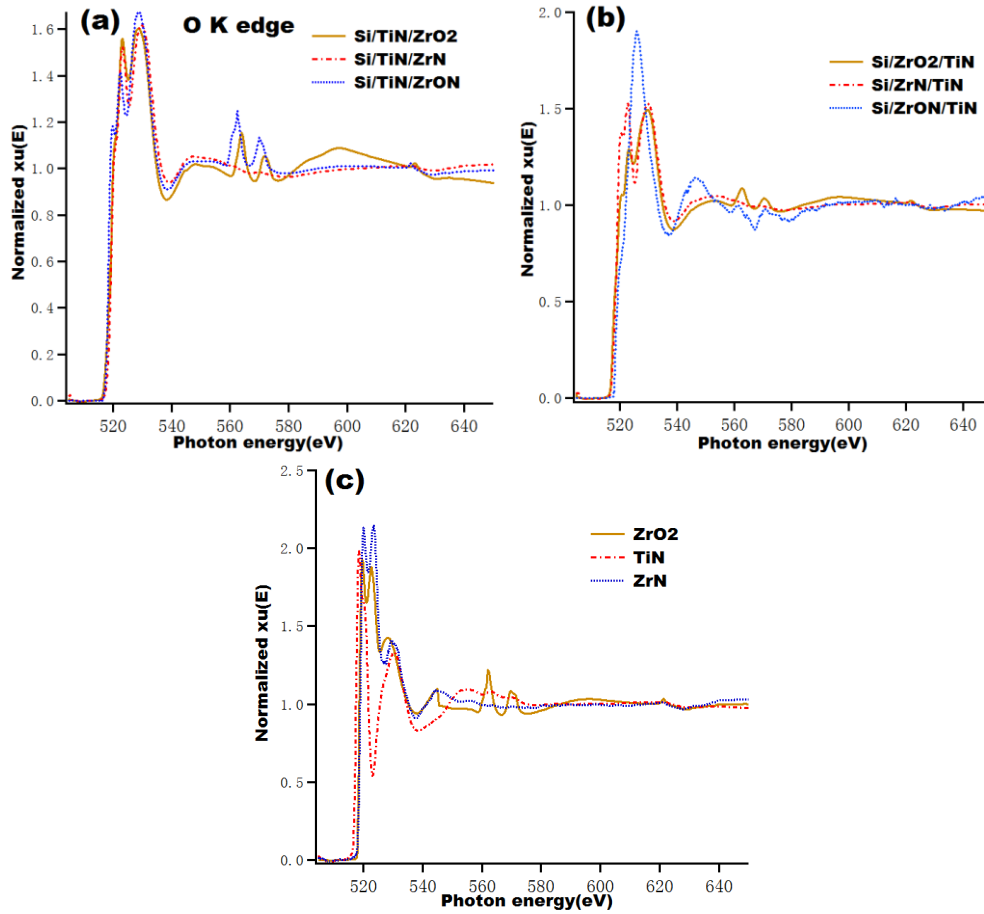


Figure 4. O K edge EXAFS of all samples, (a): Si/TiN/ZrO₂, Si/TiN/ZrON, Si/TiN/ZrN bilayer; (b) Si/ZrO₂/TiN, Si/ZrON/TiN, Si/ZrN/TiN bilayer; (c) ZrO₂, TiN and ZrN single layer.

Figure 4(a) shows the O K EXAFS spectra of three bilayer samples with the same deposition order of TiN layer, that is to say TiN being the first layer close to Si substrate. Near the edge, the three samples show similar features. But the O K absorption edge of the Si/TiN/ZrON sample shifts toward lower photon energy. In the extended edge area range at 550-640 eV, especially after 620 eV, the Si/TiN/ZrO₂ sample presents different spectra from the others. It means that an interaction between O and Ti or N atoms takes place. Figure 4(b) is EXAFS with the ZrO_xN_y layer close to the Si substrate. The feature of the Si/ZrON/TiN sample is different from the others, as it presents a strong peak around at 526 eV. After 560 eV no more features are observed. The Si/ZrO₂/TiN and Si/ZrN/TiN samples present similar features but with different intensities near the edge. After the 560 eV, only two small peaks appear for

the Si/ZrO₂/TiN sample. Figure 4(c) shows the O K EXAFS spectra of the three single layer samples. ZrO₂ and ZrN show similar spectra around their edge. But in the extended edge area only the ZrO₂ sample present two small peaks at 562 and 567 eV. Now let us go back to the EXAFS spectra of the bilayer samples: it is observed that all bilayer samples with ZrO₂ layer present two similar small peaks around at 562 and 567 eV. In addition, the Si/TiN/ZrON sample has also two similar peaks. It indicates that these two peaks are due to the Zr-O bonding environment.

3.2 Time of Flight Secondary Ion Mass Spectrometry

A selection of four samples has been profiled in depth by ToF-SIMS to get the direct view of the composition of the stacks. The focus has been put on the order of the stack as regards TiN and the Zr-containing compounds, as well as on the impact of using ZrON compared to ZrO₂. Figure 5 shows the negative ion depth profiles, which intensities are normalized, of the Si/TiN/ZrON sample. The vertical red dashed lines: L1, L2 and L3 arbitrarily materialize B₄C-on-ZrON, ZrON-on-TiN and TiN-on-Si substrate interfaces, for better visualizing the stack. The C⁻, BO⁻, ZrO₂⁻, TiO⁻, TiON⁻, Si₂⁻ and O₂⁻ are excellent chemical markers for the different layers (from the outermost part of the stack to bulk Si) which reveal :

- the outermost contamination layer of adventitious carbon (C⁻);
- the B₄C capping layer, well described by the BO⁻ and BO₂⁻ markers (BO₂⁻ line is not shown here for sake of clarity), which is at least partly (maybe totally) oxidized most probably in air during transfers;
- the ZrO₂⁻ ion profile (or ZrO⁻, not shown here) represents the Zr-rich layer, with an increase, a short plateau and a decrease at the interface with the following layer;
- the TiON⁻ and TiO⁻ depth profile, at the interface with the ZrON layer and Si substrate, respectively (the substrate is itself well described by the Si₂⁻ profile, representing bulk Si).

In addition, the depth profile of the O₂⁻ ion also presents an interesting shape, with a

maximum corresponding to the most oxidized part, which is most probably oxidized-ZrON into ZrO_2 , during deposition. It is also necessary to indicate that the $ZrON^-$ ion does not present any significant intensity along the Zr-rich layer, confirming the change in chemistry of the Zr-compound (into ZrO_2). This is also in agreement with the XANES conclusions.

As regards interfaces between the ultra-thin layers, Si/TiN/ZrON shows a significant overlap between the TiO^- and ZrO_2^- profiles, indicating a possible intermixing between the Zr-rich and Ti-rich layers. The internal interface between the Ti-rich layer and bulk Si is also mixed with the passive part of Si (oxidized Si into SiO_2) which has formed on top of the wafer, whatever the cleaning pre-treatment before deposition.

One important intermediate conclusion is that the ToF-SIMS depth profile proves the bi-layered structure in Si/TiN/ZrON, with additional proofs of the chemical information in and between the layers. After about 100 s of sputtering in our conditions, both ultra-thin deposited layers are sputtered.

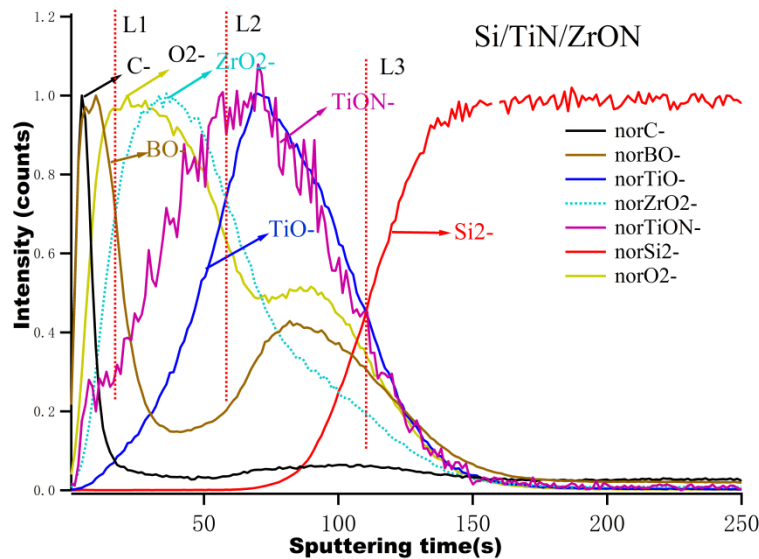


Figure 5. ToF-SIMS depth profiles of negative ions for the Si/TiN/ZrON sample (normalized intensities)

Figure 6 shows the negative ion depth profiles of the Si/ ZrO_2 /TiN and Si/ZrON/TiN samples. The C^- , BO^- , ZrO_2^- , TiO^- , $TiON^-$, Si_2^- and O_2^- ion profiles, which intensities are normalized, present a similar bi-layered structure as in Si/TiN/ZrON, but this time, the Ti-rich layer (represented by the TiO^- and $TiON^-$ markers) is at the interface with

the B_4C layer (represented by the BO^- marker), while the Zr-rich layer (represented by the ZrO_2^- marker) is an inner interface with the passivated silicon of bulk Si. Again, after about 100 s of sputtering, the full stack is sputtered. When comparing with $Si/TiN/ZrO_2$, it seems that there is no significant difference here between ZrO_2 and $ZrON$ compounds in the stacks, from the examination of the profiles. Most probably the reason is that they are too thin layers (2.0 nm) to get a bulk composition, having two close interfaces with oxidized compounds, oxidized silicon at the inner interface and oxidized Ti at the other interface.

The question is then to compare $Si/TiN/ZrON$ and $Si/ZrON/TiN$ samples. Between the two profiles (Figure 5 and Figure 6), we can compare both the shape of the TiO^- line and the intensity of the $TiON^-$ profile: the former is reasonably sharp and symmetric (the profile presents a full width at half maximum of about 35 s of sputtering) in $Si/ZrON/TiN$ (similar in $Si/ZrO_2/TiN$), and it is larger (about 40 s of sputtering) and less symmetric in $Si/TiN/ZrON$ (same observation can be made with the $Si/TiN/ZrO_2$ sample, which profile is not shown here). This increased “width” may reveal a mixed inner interface with the oxidized silicon present in the outermost layer of bulk Si.

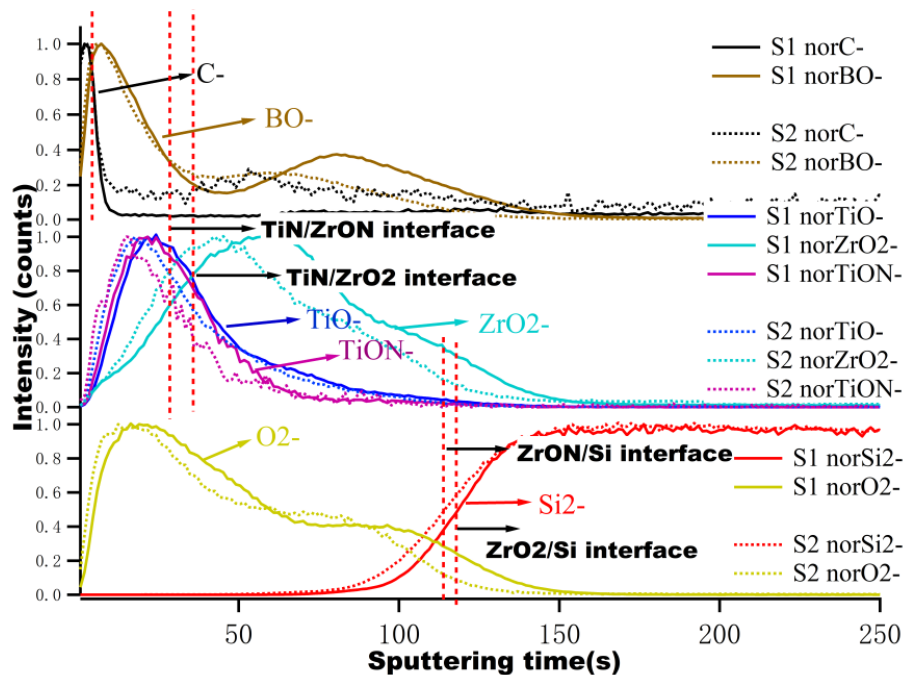


Figure 6. ToF-SIMS profiles (normalized intensities) of negative ions of the $Si/ZrO_2/TiN$ (labeled as S1) and the $Si/ZrON/TiN$ (labeled as S2) samples.

The shape of the TiON⁻ profile is the same as the TiO⁻ profile with a smaller intensity but still very significant in the case of Si/ZrON/TiN (similar in Si/ZrO₂/TiN), while the TiON⁻ signal is hardly detectable in the case of Si/TiN/ZrON and Si/TiN/ZrO₂ samples. This may indicate that the TiN layer is not fully oxidized into a TiO₂ layer, as stated above from the XANES measurements, when TiN is at the outer interface, while it is more oxidized at the inner interface.

4. Conclusion

The interdiffusion and interactions in TiN/ZrO_xN_y bilayers with reverse deposition orders were studied by synchrotron radiation x-ray absorption spectroscopy and time of flight secondary ion mass spectrometry. The interface diffusion and reaction between TiN and ZrO_xN_y layers is fully understood by using the samples with reverse layer order. No significant interaction between TiN and ZrN layers was found in the Si/TiN/ZrN sample. Both XAS results and ToF-SIMS depth profiles reveal that the reaction between TiN and ZrO_xN_y layers depends on the deposition order of layers. The TiN layer is more easily oxidized when the ZrO₂ or ZrON layer is deposited onto the TiN layer. In addition, the resulting stacks are not significantly different if ZrON is used instead of ZrO₂, revealing most probably the full oxidation of ZrON.

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