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Achieving visible light-driven hydrogen evolution at positive bias with a hybrid copper–iron oxide|TiO₂-cobaloxime photocathode†

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 C. Tapia,^a E. Bellet-Amalric,^b D. Aldakov,^c F. Boudoire,^d K. Sivula,^d L. Cagnon^e and V. Artero^{*a}

H₂ is an environmentally-friendly fuel that would allow for a circular economy but its sustainable production, e.g. from solar energy and water, remains a challenge. A hybrid CuFe_xO_y|TiO₂-CoHEC (CoHEC = chloro([4,4'-bipyridine]-2,6-dicarboxylic acid)bis(dimethylglyoximate)cobalt(III)) photocathode for hydrogen evolution reaction (HER), with faradaic efficiencies of 54–88%, is described, the preparation of which uses only non-toxic and Earth-abundant elements, avoids etching treatments and limits the use of organic solvents. The semi-conducting CuFe_xO_y light absorber is obtained by sol–gel synthesis followed by calcination at moderate temperature. Grafting the CoHEC cobaloxime catalyst at its surface results in H₂ evolution with an onset photocurrent potential of +860 mV; this process being stabilized by the presence of a thin layer of amorphous TiO₂ deposited onto CuFe_xO_y.

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Introduction

Hydrogen production *via* the water splitting reaction is a promising solution to store massive amounts of renewable energy and allow the essential ecological transition.¹ However, direct sunlight-driven, *i.e.* photoelectrochemical, H₂ evolution must overcome significant challenges in terms of performance, stability and scalability in order to be considered as a credible alternative to current technologies,² which rely on photovoltaics coupled to water electrolyser. Water-splitting photoelectrochemical cells (water-splitting PECs) hold promise to exceed current technology since they are based on photoelectrode(s) that combine solar energy absorption, energy conversion and catalysis for hydrogen evolution or water oxidation.^{3,4} The choice of the photoelectrode material(s) determines the efficiency of the PEC. Light absorption is generally achieved by a semiconductor material that should combine (i) suitable bandgap energy and band positions to allow catalysis to proceed, (ii) efficient charge carrier separation and transpor-

tation and (iii) good stability in contact with aqueous electrolytes. Here, we focus on the design of an efficient photocathode for hydrogen evolution, for which fewer solutions exist. Indeed, the large majority of studies have focused on Cu₂O^{5–8} and p-doped Si^{9,10} as photocathode candidates due to their narrow band-gaps, adequate band positions, and Earth abundance. However, the flat-band potential of p-Si combined with the low photovoltage associated to its low band-gap directly affects the onset photocurrent potential, which is not positive enough for standalone photoelectrochemical water-splitting. In addition, both p-Si and Cu₂O suffer from the lack of photostability and require ultra-efficient protective coating to avoid oxidation and photo reduction, respectively. Other solutions exploit critical elements such as indium and gallium.^{11,12} Silicon and III–V semiconductors exhibit limited stability in aqueous solutions and must be thoroughly passivated after HF etching of native oxide layer.^{9,10,13–15} Metal oxides hold promises in that prospect.^{5,16} Unfortunately, copper oxides suffer from photocorrosion if they are not efficiently protected. To overcome this limitation, the interest has shifted to ABO_x ternary metal oxides, in which the insertion of a B cation (such Fe, V, Bi, Cr) can perturb the electronic band structure enough to stabilize copper oxides against photocorrosion.¹⁷ In particular CuFeO₂ delafossite, exclusively based on Earth-abundant elements, has a 1.5–1.6 eV bandgap suitable for absorption of visible light and a conduction band positioned negative enough to allow for hydrogen evolution.^{18–21} Only few studies were reported so far toward the use of this material in a photocathode for H₂ evolution,

^aUniv. Grenoble Alpes, CNRS, CEA IRIG, Laboratoire de Chimie et Biologie des Métaux, F-38000 Grenoble, France. E-mail: vincent.artero@cea.fr

^bUniv. Grenoble Alpes, CNRS, CEA IRIG, PHELIQS, F-38000 Grenoble, France

^cUniv. Grenoble Alpes, CNRS, CEA IRIG-SYMMES, F-38000 Grenoble, France

^dLaboratory for Molecular Engineering of Optoelectronic Nanomaterials, École Polytechnique Fédérale de Lausanne (EPFL), Station 6, 1015 Lausanne, Switzerland

^eUniv. Grenoble Alpes – CNRS UPR 2940, Institut NEEL, F-38000 Grenoble, France

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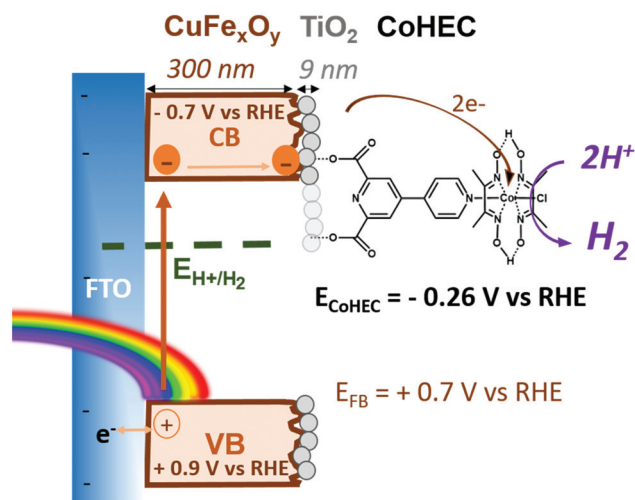


Fig. 1 Scheme of electron pathway on $\text{CuFe}_x\text{O}_y/\text{TiO}_2$ -CoHEC photocathode and energy potentials.

and combining this photoactive material with a surface catalyst to enhance its photochemical properties has not been explored so far. Recently, we and others have demonstrated that cobalt diimine-dioxime^{22–25} or cobaloximes,^{25–29} with proper anchoring groups can be used as molecular electrocatalysts for H_2 evolution in metal oxide-based photocathodes. Therefore, we decided to exploit the CoHEC cobaloxime catalyst (Fig. 1) to decorate photoactive copper-iron oxide materials towards the elaboration of a H_2 -evolving photocathode and found that good performance could be reached with the deposition of a thin layer of amorphous TiO_2 as a protective coating.^{23,26} Fig. 1 displays the structure and electronics of such a hybrid architecture. The protection of the Cu-Fe oxide layer with a thin layer of amorphous TiO_2 allows to photoproduce H_2 with faradaic efficiencies of 54–88% and with onset potential as high as 0.4 V vs. RHE in neutral aqueous electrolyte under pure visible light irradiation. Such a hybrid approach³⁰ thus proves a competitive and greener alternative to dye-sensitized photoelectrodes^{25,31,32} with the advantage of avoiding the dye preparation and sensitizing steps.

Experimental section

Materials

Fluorine-doped tin oxide (FTO) coated glass substrates were purchased from Solems (Palaiseau, France). Acetone, absolute ethanol, copper nitrate trihydrate, iron nitrate nonahydrate, ethylene glycol, methanol, poly(ethylene glycol)-*block*-poly(propylene glycol)-*block*-poly(ethylene glycol) (F-108) and potassium phosphate were purchased from Sigma-Aldrich. Chloro([4,4'-bipyridine]-2,6-dicarboxylic acid)bis(dimethylglyoximate) cobalt(III) (CoHEC) was purchased from Dyenamo (Sweden). Titanium(IV) isopropoxide (TTIP) was purchased from Strem

Chemicals Inc. All the reagents were used as received without further purification.

Photocathode preparation

Cu-Fe oxide (CuFe_xO_y) deposition on FTO glass substrates. The procedure was adapted from previously reported studies.^{18,33} Glass|FTO substrates ($20 \times 40 \text{ mm}^2$) were cleaned first with soap and water, then with ethanol and finally with acetone, each step during 15 minutes in an ultrasonic bath. Then, they were treated in an UV-ozone cleaner during 15 minutes. A mixture of 0.2 M $\text{Cu}(\text{NO}_3)_2 \cdot 3\text{H}_2\text{O}$, 0.2 M $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ and 0.45 M ethylene glycol in ethanol was spin coated on glass|FTO substrates at 3300 rpm during 1 minute followed for 600 rpm during 30 seconds, following a protocol previously reported.²¹ The substrates were placed in a flat titanium furnace and heated up to 200 °C in the air during 30 minutes to stabilize the film. The spin coating and the heating processes were repeated 5 times for each glass|FTO substrate. A last annealing step at 450 °C for 30 minutes and then at 600 °C for 2 hours in a tubular furnace under argon flow was performed to eliminate all organic residues, reduce Cu(II) to Cu(I) and crystallize the Cu-Fe oxide films ($\text{CuFe}_x\text{O}_y\text{FLAT}$).^{18,33} In some cases, the sixth layer was structured in order to increase the surface area of the CuFe_xO_y electrode. Here, F-108 polymer (120 mg mL^{-1}) was added to the solution used for fabricate the sixth layer.²⁵ We will refer to these layers as $\text{CuFe}_x\text{O}_y\text{STRUCT}$.

Deposition of amorphous TiO_2 by atomic layer deposition (ALD). TiO_2 was deposited on $\text{CuFe}_x\text{O}_y\text{FLAT}$ electrodes by ALD (Savannah 100 ALD system, Cambridge Nano Tech Inc.) with precursor temperature at 95 °C, pump line temperature at 150 °C and sample chamber temperature at 255 °C in continuous mode under 5 sccm nitrogen flow. TTIP was used as Ti source and water as O source. 355 cycles were applied, each cycle being composed of 1 s pulse of TTIP, 5 s pause, 0.015 s pulse of water and 5 s pause. The resulting electrodes are referred to as $\text{CuFe}_x\text{O}_y/\text{TiO}_2$.

CoHEC grafting on $\text{CuFe}_x\text{O}_y\text{STRUCT}$ and $\text{CuFe}_x\text{O}_y/\text{TiO}_2$. $\text{CuFe}_x\text{O}_y\text{STRUCT}$ and $\text{CuFe}_x\text{O}_y/\text{TiO}_2$ electrodes were immersed in 0.7 mM CoHEC methanol solution overnight, then rinsed with methanol and dried in air before use.

Characterization techniques

UV-visible absorption spectra of CuFe_xO_y and $\text{CuFe}_x\text{O}_y/\text{TiO}_2$ (before and after CoHEC grafting) were recorded with an Agilent Technologies Cary 60 UV-Vis spectrometer. The baseline corresponding to the glass|FTO support was recorded and automatically subtracted from the spectrum of the electrode.

Scanning electron microscopy (SEM) images showing the electrode morphology were recorded with a scanning electron microscope Zeiss Ultra 55 operating at 4 kV.

XRD measurement were performed on a Smartlab Rigaku diffractometer in glancing incidence: the incident beam was fixed at low angle (1 degree) to maximize the signal of the thin layer. The measurements correspond to a 2θ scan.

Film-thickness measurements were performed with a laser scanning microscope VK-X120 (Keyence), using VK Viewer and VK Analyser softwares.

An ellipsometer (M 2000 model) from J. A. Woollam Co., In. was used to measure the thickness of the ALD-TiO₂ substrate, using the CompleteEASE software (V4.72). It was used in the “off-sample” configuration with an incident angle of attack of 75° with respect to the film plane normal, and the light wavelength was scanned between 200 nm and 1700 nm. The obtained data were fitted with the CompleteEASE software (V4.72).

X-ray photoelectron spectroscopy (XPS) analyses were carried out with a Versa Probe II spectrometer (ULVAC-PHI) equipped with a monochromated Al K α source ($h\nu = 1486.6$ eV). The core level peaks were recorded with constant pass energy of 23.3 eV. The XPS spectra were fitted with CasaXPS 2.3 software using Shirley background. Binding energies are referenced with respect to the adventitious carbon (C 1s BE = 284.6 eV).

Inductively coupled plasma atomic emission spectroscopy (ICP-AES) was performed using a Shimadzu ICPE-9000 instrument with a mini plasma torch in axial reading mode. The sample preparation involves the digestion overnight of the samples CuFe_xO_y, CuFe_xO_y-CoHEC, CuFe_xO_y|TiO₂-CoHEC and CuFe_xO_y-CoHEC before and after chronoamperometry treatment in 1 mL of HNO₃ (64%).

Photoelectrochemical and hydrogen production measurements

Electrochemical data were acquired with a Biologic VSP 300 (Linear scan voltammetry), an IVIUM compact potentiostat (chronoamperometry) and Autolab potentiostat with FRA module (electrochemical impedance spectroscopy – EIS). All measurements were carried out in a three electrode (photo) electrochemical cell, filled with pH 6.7 0.2 M phosphate buffer as electrolyte. A titanium wire was used as counter electrode (CE) and a Ag/AgCl (3 M KCl) electrode as reference electrode (RE). Potentials are referred to RHE by using the equation ($E = 0.200 \text{ V} + 0.059 \times \text{pH}$). Calibration of the reference electrode was realized with [K₄Fe(CN)₆] in 0.1 M potassium phosphate buffer at pH 7.³⁴ Argon or nitrogen was bubbled to the electrolyte for 20 minutes to remove the oxygen before any electrochemical measurement. The geometrical active surface area was 0.5 cm². The electrical contact in the working electrode was from the FTO part. EIS measurements were conducted in 0.2 M phosphate pH 6.7 at 10 kHz and 0.4–0.7 V vs. RHE potential range.

For photoelectrochemical measurements, a 280 W Xe lamp was used to irradiate the electrodes with 1 sun filtered so that only the visible part of the spectrum reaches the photoelectrode (400–780 nm, 65 mW cm⁻²). UV irradiation was filtered with a Spectra-Physics 59472 ($\lambda < 400$ nm) cut off filter and IR irradiation was filtered by a Spectra-Physics 6123NS circulating water filter kept at 25 °C. Chronoamperometric measurements were performed in a previously described cell with 4 mL electrolyte volume and 2.5 cm² geometrical active surface area.³⁵

The working electrode was connected with conductive silver paste to a wire, and the cell was completely isolated from air. Hydrogen production was determined by a hydrogen probe from Unisense, polarized at 1 V vs. Ag/AgCl, and calibrated following the procedure described in our previous publication.³⁵ It was placed inside the photo-electrochemical cell before and after the chronoamperometry to measure the amount of hydrogen in solution.³⁵

Results and discussion

Different photocathodes for the hydrogen evolution reaction (HER) were prepared following the three-step strategy displayed in Fig. 2. First, the copper-iron oxide (CuFe_xO_y) thin layers were prepared by sol-gel method from a mixture of copper and iron salts and spin-coated on glass|FTO substrates, following previous reports^{18,33} as described in the Experimental section. Then, amorphous TiO₂ was deposited on CuFe_xO_y layers by ALD. Afterwards, the CoHEC cobaloxime catalyst was grafted on the top of CuFe_xO_y|TiO₂ samples by overnight soaking into methanolic solution. On another batch of samples, CuFe_xO_y layer was nanostructured (CuFe_xO_ySTRUCT) using the F108 triblock co-polymer template and the catalyst is grafted at the surface of CuFe_xO_ySTRUCT without amorphous TiO₂ deposition.

Characterization of the Cu-Fe oxide (CuFe_xO_y) layers

Brown colored CuFe_xO_yFLAT six-layered samples were obtained after several spin-coating deposition and annealing cycles as described in the Experimental section. Observation by SEM revealed a single thin film of 221 ± 30 nm thickness (Fig. 3A) without nanostructuring (Fig. 3Ac). The low active surface area of the electrode could result in poor catalyst loading in further steps.

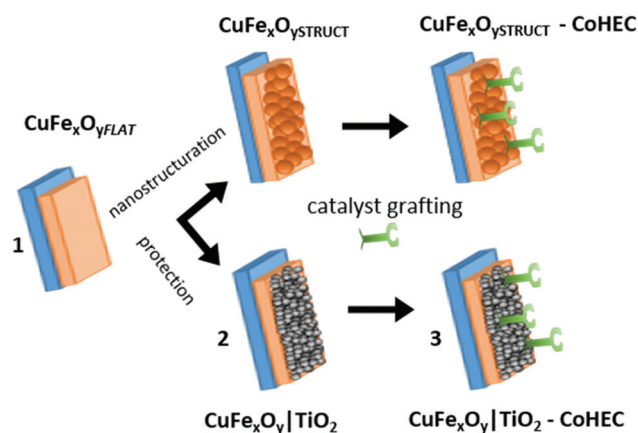


Fig. 2 Three-step preparation of the CuFe_xO_ySTRUCT-CoHEC and CuFe_xO_y|TiO₂-CoHEC photocathodes: (1) synthesis of multi-layered CuFe_xO_y light-harvesting semiconductor samples; (2) protection of the semiconductor with TiO₂ or CuFe_xO_y nanostructuring, and (3) grafting of the CoHEC molecular catalyst.

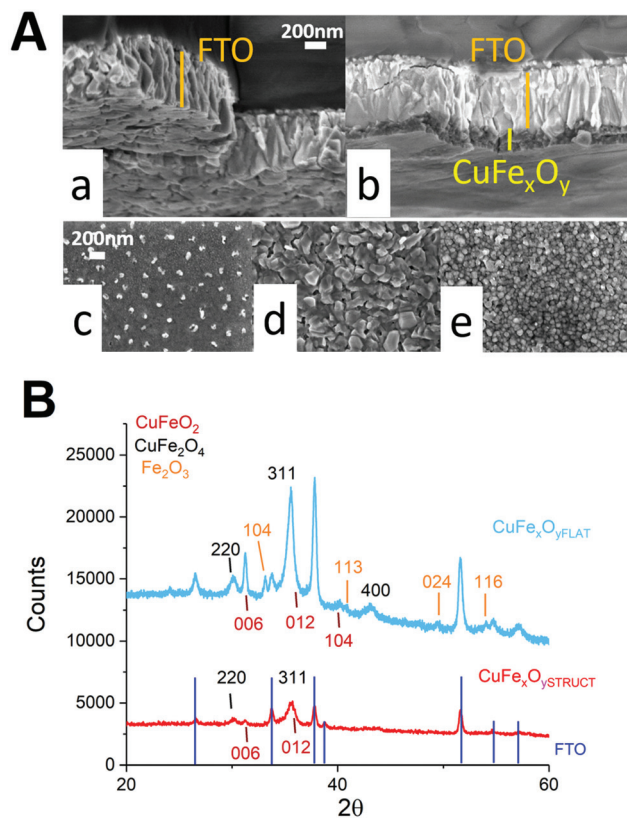


Fig. 3 (A) SEM sectional images from (a) clean FTO and (b) CuFe_xO_y layer synthesized on FTO (same scale). SEM front images from (c) CuFe_xO_yFLAT, (d) CuFe_xO_ySTRUCT and (e) CuFe_xO_y/TiO₂ (same scale). (B) XRD pattern of CuFe_xO_yFLAT in blue, CuFe_xO_ySTRUCT in red and FTO in dark blue. All peaks are referenced to delafossite CuFeO₂ (JCPDS 04-015-2087) (brown); spinel CuFe₂O₄ (JCPDS 01-074-8585) (black) and Fe₂O₃ (JCPDS 04-003-2900) (orange).

The addition of a triblock polymer known as F-108 in the sol-gel used for the sixth and last cycle resulted in higher nanostructuring (CuFe_xO_ySTRUCT, Fig. 2 and 3Ad). This top layer with a thickness of 137 ± 9 nm (Fig. S1†) displayed a grain size around 100–200 nm in diameter (Fig. 3Ad). The total thickness of CuFe_xO_ySTRUCT was estimated to 320 ± 40 nm.

The Cu : Fe atomic ratio measured by EDX on CuFe_xO_yFLAT thin film was 1 : 1.3 (Fig. S2†), indicating a copper-deficient delafossite CuFeO₂ structure. XRD revealed that CuFe_xO_yFLAT electrodes contain a rhombohedral delafossite CuFeO₂ (JCPDS 04-015-2087) phase together with cubic spinel CuFe₂O₄ (JCPDS 01-074-8585) and Fe₂O₃ (JCPDS 04-003-2900) parasitic phase (Fig. 3B). These results were confirmed by Raman spectroscopy performed in CuFe_xO_yFLAT and CuFe_xO_ySTRUCT (Fig. S3†). CuFe₂O₄ is first formed from CuO and Fe₂O₃, which are formed from the same sol-gel precursor solution before reaching 180 °C and around 300 °C, respectively.^{33,36} Afterwards, the spinel CuFe₂O₄ phase can be converted with CuO during the last annealing step and form CuFeO₂ delafossite with no need of additional reductant agent.³⁷ The presence of CuFe₂O₄ in our films suggests this last step is not fully achieved. The

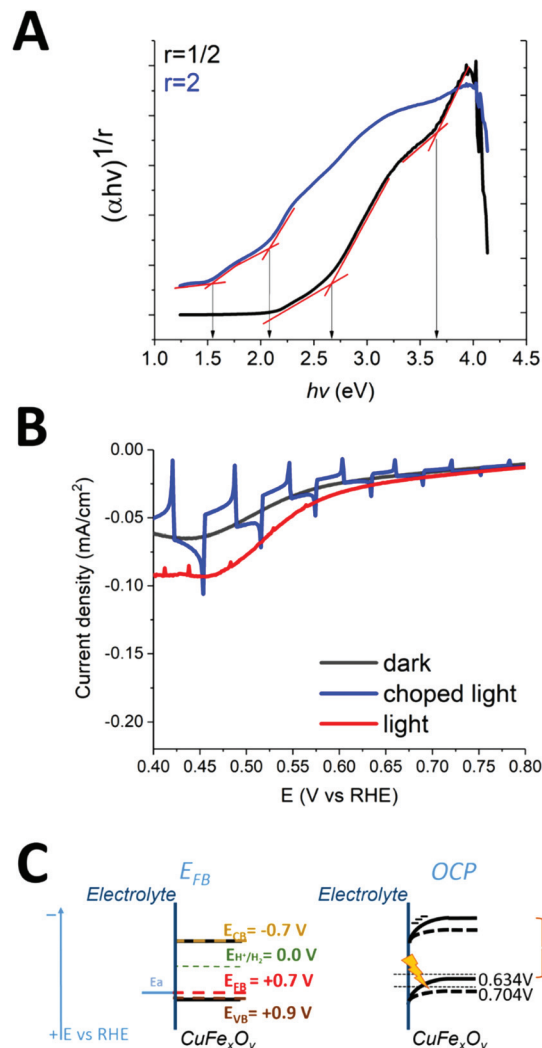


Fig. 4 (A) Tauc plots from CuFe_xO_yFLAT UV-visible spectrum for direct ($r = \frac{1}{2}$) and indirect ($r = 2$) bandgap calculation. (B) Linear sweep voltammogram (LSV) of CuFe_xO_ySTRUCT films under chopped light irradiation (blue), constant irradiation (red) and in the dark (black) in 0.2 M phosphate buffer pH 6.7 at 10 mV s⁻¹ scan rate, under argon; irradiation with 1 sun visible light (400–780 nm). (C) Scheme of the energy bands of CuFe_xO_yFLAT films, where E_{FB} , E_{VB} and E_{CB} are the flat-band, valence band and conduction band potentials, respectively, while E_{H^+/H_2} is the redox potential corresponding to the proton reduction: left, the electrode is poised at the flat-band potential; right: the electrode is in equilibrium with the solution at the open circuit potential (OCP).

XRD spectrum of CuFe_xO_ySTRUCT samples revealed a higher ratio between CuFe₂O₄ and CuFeO₂ phases. Nevertheless, CuFe₂O₄ has been also reported as a good candidate for HER under visible light, with a similar bandgap and band position as CuFeO₂ delafossite.^{38,39}

Bandgap energy values (E_g) for direct and indirect allowed transitions were determined by UV-visible absorbance spectrum. Tauc plots were measured based on the UV-visible absorbance data of CuFe_xO_yFLAT corrected for FTO absorption and taking into account the thickness of the layers. As shown in Fig. 4A, CuFe_xO_yFLAT thin film exhibited an indirect bandgap

of 1.6 eV and two direct bandgaps of 2.65 eV and 3.6 eV, in accordance to the literature for delafossite CuFeO_2 ^{20,21,33} and spinel ferrite CuFe_2O_4 ³⁹ confirming that this material is capable of absorbing visible light. However, an indirect bandgap at 2.0 eV was also observed, confirming the presence of Fe_2O_3 parasitic phase (which bandgap is 2.0–2.2 eV) already observed by XRD.⁴⁰

High resolution XPS analysis of $\text{CuFe}_x\text{O}_y\text{STRUCT}$ revealed Cu(I) and Cu(II) oxidation states which suggests a mixture of delafossite CuFeO_2 and ferrite CuFe_2O_4 phases, shown in Fig. S4.† The presence of Cu(II) is easily identified by intense peaks around 933 eV together with satellite shake-up signals above 940 eV, while Cu(I) is evidenced by a single peak at 932 eV.

The flat band potential (E_{FB}) was determined through Mott–Schottky plot by performing EIS measurements. The p-type nature of the CuFe_xO_y layers on FTO was confirmed by the negative slope of the Mott–Schottky plot. E_{FB} was estimated at 0.72 V vs. RHE in 0.2 M phosphate buffer (pH 6.7) and 10 kHz (Fig. S5†).⁴¹ This value is more negative than the already reported value of 1 V vs. RHE for delafossite CuFeO_2 ¹⁸ which again suggested the presence of impurities. The flat-band potential also corresponds to the potential at which the semiconductor starts to show photocurrents (onset potential)³ and a similar value of E_{FB} 0.70–0.75 V vs. RHE is determined from LSV measurements on $\text{CuFe}_x\text{O}_y\text{STRUCT}$ back illuminated with 1 sun visible light (65 mW cm⁻², 400–780 nm) in the same electrolyte (Fig. 4B). This electrolyte was selected based on previous studies on Cu-based oxides conducted in neutral pH.^{42,43} The open circuit potential (OCP) of the CuFe_xO_y layers when illuminating the sample was 0.70 V vs. RHE, and a shift in the photopotential towards less positive potential was recorded in dark (0.63 V vs. RHE), shown in Fig. S6.† When illuminating, the bands tended to flatten and the OCP became more positive.³ This behaviour also confirms the p-type nature of the semiconductor and its flat band potential measured by EIS. We acknowledge that OCP should be monitored for increasing light intensities until saturation is reached. Indeed if the light intensity is not high enough to flatten the bands, the flat band potential value may be underestimated.⁴⁴ We note however that the OCP value measured under 1 sun illumination matches both onset photocurrent potential and E_{FB} value determined using Mott–Schottky analysis.

When potentials more negative than the onset potential are applied to the $\text{CuFe}_x\text{O}_y\text{STRUCT}$ electrodes, the energy bands at the surface of the semiconductor in contact with the electrolyte bend towards more negative potential allowing electrons to travel to the electrolyte. According to the reported data,¹⁸ the diffusion length of the charge carriers of CuFeO_2 delafossite is ≈ 300 nm. Therefore, the thickness of the CuFe_xO_y layers was convenient to allow both the electrons from the hole–electron pair created near the interface with the FTO and the holes created at the surface with the electrolyte to reach the electrolyte and the FTO back contact, respectively, before they recombine. Indeed, LSV measured for $\text{CuFe}_x\text{O}_y\text{STRUCT}$ electrodes displayed cathodic photocurrents (Fig. 4B) under visible light irradiation. At 0.4 V vs. RHE photocurrent density reached

85 $\mu\text{A cm}^{-2}$ (Fig. 4B) which is close to the performances reported for similar delafossite CuFeO_2 samples in the absence of other electron acceptor apart from protons.⁴⁵

The energy bands position of the CuFe_xO_y layers and its behaviour in equilibrium with the electrolyte are represented in Fig. 4C. Assuming that the valence band is located 0.1–0.2 V below the flatband potential ($E_{\text{FB}} = +0.7$ V vs. RHE),⁴⁶ we determined $E_{\text{VB}} = +0.8$ to $+0.9$ V vs. RHE. Considering the indirect band gap (E_{g}) of 1.6 eV, the conduction band is therefore placed at -0.7 to -0.8 V vs. RHE ($E_{\text{CB}} = E_{\text{VB}} - E_{\text{g}}$). Therefore, the conduction band is negative enough for protons reduction, as illustrated in the scheme in Fig. 4C.

Catalyst grafting and photoelectrochemical performance for HER

The CoHEC catalyst is a cobaloxime bearing an axial 4,4'-pyridine ligand functionalized with two carboxylic acid anchoring groups. Following its use in dye-sensitized photoelectrodes,^{25,28} it was selected as a candidate to enhance the efficiency of the photoelectrocatalytic proton reduction at the surface of the CuFe_xO_y electrodes. The grafting of CoHEC on $\text{CuFe}_x\text{O}_y\text{STRUCT}$ electrode was achieved by soaking $\text{CuFe}_x\text{O}_y\text{STRUCT}$ electrodes in a methanolic CoHEC (0.7 mM) solution overnight. The presence of CoHEC grafted on CuFe_xO_y was evidenced by the increase in the absorbance around 450–350 nm in the UV-visible absorption spectrum (Fig. S7†), as previously reported.⁴⁷ The presence of cobalt in the $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC samples was confirmed by ICP with Co emission observed at 228.616 eV (Fig. S8†). The ICP signal is within the detection limit of this technique, indicating a surface concentration of 0.085 nmol cm⁻². XPS analysis of $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC showed characteristic peaks of Co(II) 2p_{3/2} with a multiplet splitting due to the coupling with the unpaired electrons and plasmon losses (779.9, 782.3, and 787.7 eV) likely corresponding to a single Co complex,⁴⁸ and two nitrogen atoms (N 1s at 399.0 eV for pyridinic N and 401.1 eV for oxime N), confirming the presence of the cobalt center of the CoHEC as well as nitrogen present in CoHEC ligands (Fig. S9†). Co:N ratio obtained from the XPS studies is 6.3, which is very close to 6 expected from the stoichiometry of the complex.

Catalytic onset potential for hydrogen evolution mediated by CoHEC at a bare FTO electrode in the dark was measured at -0.26 V vs. RHE in 0.2 M phosphate buffer (pH 6.7) (Fig. 5 and S10†), in accordance to the values reported in neutral pH.²⁵ Sensitization with $\text{CuFe}_x\text{O}_y\text{STRUCT}$ shifts this onset potential to $+0.7$ V vs. RHE, in accordance to the E_{FB} value previously determined for CuFe_xO_y films. In other words, the $\text{CuFe}_x\text{O}_y\text{STRUCT}$ material provides a photovoltage of ~ 960 mV to the H₂ evolution reaction, which is definitely much higher than that provided by p-Si or Cu₂O. Indeed, the onset potentials reported for p-Si,^{23,49} Cu₂O⁶ or WSe₂⁵⁰ photocathodes are 0.4–0.6 V vs. RHE. As shown in Fig. 5 and Fig. S11,† photocurrents measured at $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC are twice higher than those measured at unfunctionalized $\text{CuFe}_x\text{O}_y\text{STRUCT}$ photocathodes in the absence of CoHEC, showing the benefit of

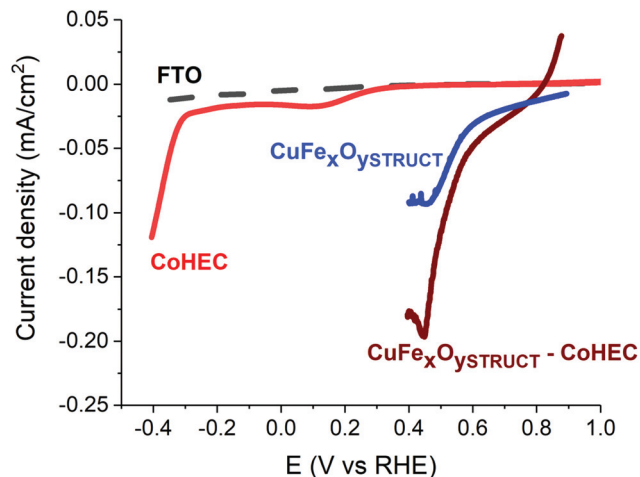


Fig. 5 LSV for bare FTO substrate (dashed line in black), bare FTO substrate with 0.7 mM CoHEC in solution (straight red line), $\text{CuFe}_x\text{O}_y\text{STRUCT}$ modified FTO substrate (straight blue line) and $\text{CuFe}_x\text{O}_y\text{STRUCT}$ decorated with CoHEC (straight brown line). Irradiation of 1 sun visible light (400–780 nm) in 0.2 M phosphate buffer (pH 6.7) at 10 mV s^{-1} scan rate under nitrogen.

surface functionalization to enhance charge extraction and catalysis.

Such an improvement was confirmed during chronoamperometric measurements performed at 0.4 V vs. RHE under irradiation (Fig. 6A). The photocurrent for $\text{CuFe}_x\text{O}_y\text{STRUCT}$ photocathode reached just $-14 \mu\text{A cm}^{-2}$. However, when CoHEC was grafted on the surface of $\text{CuFe}_x\text{O}_y\text{STRUCT}$, photocurrents were up to 4 times higher, varying from -60 to $-20 \mu\text{A cm}^{-2}$ during the first minute and remaining higher than those measured at $\text{CuFe}_x\text{O}_y\text{STRUCT}$ photocathode during the first 8 min.

The amount of hydrogen produced in solution was measured with an hydrogen probe, calibrated following the procedure described in previous publication.³⁵ The calculated faradaic yields (FY) for both $\text{CuFe}_x\text{O}_y\text{STRUCT}$ and $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC after 20 min were 2% and 6%, respecti-

Table 1 Hydrogen production, FY, TON and TOF values for $\text{CuFe}_x\text{O}_y\text{STRUCT}$, $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC, $\text{CuFe}_x\text{O}_y|\text{TiO}_2$ and $\text{CuFe}_x\text{O}_y|\text{TiO}_2$ -CoHEC photocathodes. 0.2 M phosphate buffer (pH 6.7), 1 sun irradiation (65 mW cm^{-2} , 400–700 nm), at 0.4 V vs. RHE applied bias for 20 min

Electrode	Q (mC cm^{-2})	H_2 (nmol cm^{-2})	FY (%)	TON _{Co}	Average TOF (s^{-1})
$\text{CuFe}_x\text{O}_y\text{STRUCT}$	70 ± 30	4.8 ± 2	2		
$\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC	16 ± 7	4.8 ± 2	6	57 ± 30	0.05
$\text{CuFe}_x\text{O}_y \text{TiO}_2$	4 ± 1	4.3 ± 1	22		
$\text{CuFe}_x\text{O}_y \text{TiO}_2$ -CoHEC	3 ± 1	7.6 ± 2	54	90 ± 30	0.08

vately (Table 1). Based on the estimated Co surface concentration, the turnover number (TON) was 57 ± 30 for CuFe_xO_y -CoHEC, corresponding to an average turnover frequency (TOF) of 0.05 s^{-1} . However, longer chronoamperometries did not result in higher amount of hydrogen produced. These results clearly show that the efficiency of catalytic hydrogen production is enhanced with CoHEC grafted on the surface of CuFe_xO_y . However, desorption of the catalyst occurred within minutes of operation as shown in Fig. 6A. The photocurrent decreased along time in the case of CuFe_xO_y -CoHEC, reaching the values similar to $\text{CuFe}_x\text{O}_y\text{STRUCT}$ after 8 minutes, and even lower photocurrents afterwards, indicating some degradation of the photocathode materials. Partial loss of cobalt from the surface is confirmed by ICP (Fig. S8†) and XPS (Fig. S9†) measurements after 20 min chronoamperometric measurement. For example, the Co : Cu surface atomic ratio drops from initial 3.7 to 0.14. It is also consistent with a former observation made for similar cobaloxime catalysts immobilized onto surfaces through axial pyridine ligand^{27,51} for cathodes operated in the dark and for CoHEC grafted onto dye-sensitized NiO.²⁵ The CoHEC grafting did not affect apparently the $\text{CuFe}_x\text{O}_y\text{STRUCT}$ composition, since the observed Cu 2p and Fe 2p XPS peaks are similar for both samples (Fig. S4†). However, after applying bias potential to the sample in presence of grafted CoHEC, the Cu visible by XPS was almost completely

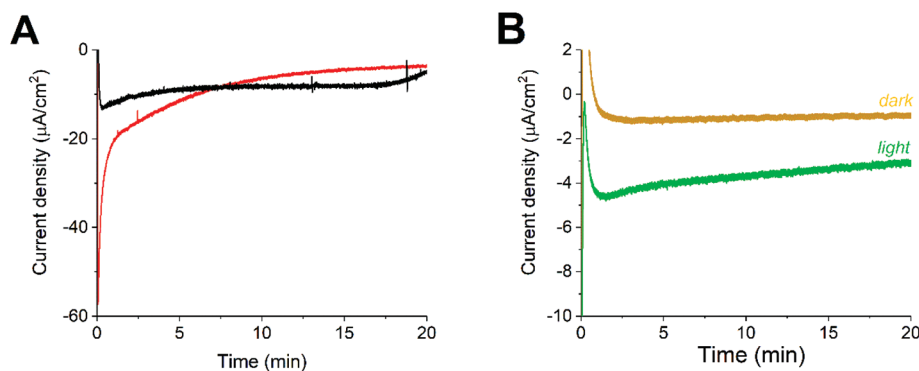


Fig. 6 (A) Chronoamperometries at +0.4 V vs. RHE on $\text{CuFe}_x\text{O}_y\text{STRUCT}$ (black) and $\text{CuFe}_x\text{O}_y\text{STRUCT}$ -CoHEC (red) photoelectrodes, irradiated by 1 sun visible light (400–780 nm) under nitrogen. (B) Chronoamperometries at 0.4 V vs. RHE on $\text{CuFe}_x\text{O}_y|\text{TiO}_2$ -CoHEC under dark (brown) and light (green) conditions. Measurements in 0.2 M phosphate buffer (pH 6.7).

reduced to Cu(I), which can explain the limited faradaic efficiency for H₂ evolution. This reduction phenomenon does not occur in the absence of CoHEC grafting but we observed a significant modification of the CuFe_xO_y material upon turnover. Indeed, the Cu:Fe atomic ratio measured by EDX changed from 1:1.3 before electrochemical treatment to 1:5 after treatment, indicating a significant loss of copper during operation (Fig. S12†). The degradation of the surface of the CuFe_xO_ySTRUCT electrode was confirmed by SEM (Fig. S13†). XRD performed on CuFe_xO_ySTRUCT electrodes after treatment revealed new unknown peaks which correspond neither to CuFe₂O₄, CuFeO₂, Cu₂O nor to Fe₂O₃ (Fig. S14†). To overcome such a degradation, we decided to protect the CuFe_xO_y material by depositing a thin TiO₂ layer on top.

Optimisation through atomic layer deposition of TiO₂

The instability of semiconductors under photoelectrochemical conditions is a general problem that can be solved by deposition of metal oxides protecting^{23,52,53} or passivating layers.⁵⁴ A classical protection strategy is the deposition of a few nm of thin amorphous TiO₂ layer by atomic layer deposition (ALD),^{23,26,49,55,56} allowing to control the thickness and improve the homogeneity of the surface layer.

We therefore deposited an amorphous TiO₂ layer, with a 8.3 nm thickness measured by ellipsometry (Fig. S15†), onto a flat CuFe_xO_y film by ALD following the procedure reported elsewhere.²³ Fig. 3Ae shows the top view of amorphous TiO₂ deposited on CuFe_xO_yFLAT electrode, revealing nanostructured TiO₂. It has been previously shown that a ~10 nm layer of amorphous TiO₂ is thick enough to passivate a semiconductor.⁵⁶ Since the resistance increases and the photoelectrochemical response drops with the thickness of the TiO₂ layer,⁵⁶ we limited the ALD process on CuFe_xO_yFLAT to 355 cycles, corresponding to a thickness of 8.3 nm measured by ellipsometry.

XRD was performed on CuFe_xO_y|TiO₂ electrodes. The main phase was CuFe₂O₄ (Fig. S16†), indicating a structural change of the bulk materials during ALD deposition. No peaks corresponding to any crystalline phase of TiO₂ were observed in the CuFe_xO_y|TiO₂ electrodes diffractogram (Fig. S16†) confirming the amorphous character of the TiO₂ layer, while the presence of Ti was confirmed by EDX (Fig. S17†) and XPS (Fig. S9B†). LSV under chopped light of CuFe_xO_y|TiO₂-CoHEC is shown in Fig. S18,† revealing lower photocurrent values and a shift of the onset photocurrent potential to ~0.6 V vs. RHE. This cathodic shift can be explained by the full oxidation of the CuFe_xO_y material into CuFe₂O₄ during the ALD process. Although only approximate bandgap values can be determined for composite films,⁵⁷ Tauc plots indicate higher (2 eV) indirect bandgap for CuFe_xO_y|TiO₂ layers compared to CuFe_xO_y films (Fig. S19†), in line with the shift of the onset photocurrent potential.

CoHEC catalyst was grafted by immersing CuFe_xO_y|TiO₂ electrodes in a methanolic solution of 0.7 M CoHEC overnight. The presence of CoHEC grafted on CuFe_xO_y and CuFe_xO_y|TiO₂ was confirmed by UV-visible spectroscopy (Fig. S7†). XPS analysis of the latter reveals a Co species manifested as three 2p_{3/2}

peaks at 781.8, 786.3, and 789.9 eV, while nitrogen 1 s appears poorly resolved at 400.3 eV (Fig. S9†). Here again, the amount of cobalt catalyst was estimated by ICP at the detection limit of the technique, *i.e.* close to 0.085 nmol cm⁻². EDX analysis confirmed the presence of Ti on the photoelectrode surface both before and after the photoelectrochemical experiment.

Photocurrents recorded for CuFe_xO_y|TiO₂-CoHEC photocathodes proved to be more stable over time since no decrease in the photocurrent was observed after 20 minutes (Fig. 6B). Within 20 min, CuFe_xO_y|TiO₂-CoHEC photocathodes produced 7.6 ± 2 nmol H₂ per cm² with 54% FY. A control experiment of CuFe_xO_y|TiO₂ in the absence of the CoHEC catalyst was launched recording 4.3 ± 1 nmol H₂ per cm² and 22% FY. The TON based on estimated Co surface concentration was 90 ± 30 for CuFe_xO_y|TiO₂-CoHEC. The TOFs for CuFe_xO_y|TiO₂-CoHEC were estimated to 0.08 s⁻¹. All these values are summarized in Table 1.

Additional measurements were carried out under the same conditions but with slightly higher visible-light intensity (up to 100 mW cm⁻²; 1.3 sun). CuFe_xO_ySTRUCT and CuFe_xO_ySTRUCT-CoHEC displayed 3% and 6% FY respectively, with similar behaviour regarding the desorption of catalyst and poor stability (Fig. S20†). However, under these conditions, the protected CuFe_xO_y|TiO₂-CoHEC photocathode showed 88% FY and 90 TON achieved within 20 min.

The hybrid photocathodes were evaluated up to two hours without increase in the amount of hydrogen detected, suggesting the photocathodes lost their activity following desorption of the catalyst which was already confirmed by the absence of peaks in the ICP and XPS measurements after 20 min CA (Fig. S8 and S9†).

Clearly, the presence of TiO₂ significantly and positively affects the FY for H₂ evolution. Although the photocurrent values are significantly lower for TiO₂-protected photoelectrodes, the amount of hydrogen produced remains similar (4.8 nmol cm⁻² for CuFe_xO_y and 4.3 nmol cm⁻² for CuFe_xO_y | TiO₂) or largely enhanced (4.8 nmol cm⁻² for CuFe_xO_ySTRUCT-CoHEC and 7.6 nmol cm⁻² for CuFe_xO_y|TiO₂-CoHEC). This indicates that TiO₂ has no significant effect in the catalysis for hydrogen evolution but that it prevents at least partially the photocorrosion of the semiconductor. The protected photocathodes showed an average Cu:Fe ratio of 1:1.8, unchanged after photoelectrochemical measurement. Any modification in the phase of the material was not observed by XRD after chronoamperometry (Fig. S11†). Additionally, the presence of TiO₂ significantly limits the decrease in the photocurrent intensity over time, indicating a better stability of the anchoring of the carboxylate functions of CoHEC onto TiO₂ compared to CuFe_xO_y.

When compared with other hybrid photocathodes from the literature, *i.e.* the p-GaInP₂/TiO₂-cobaloxime architecture described by Turner *et al.* (100% FY, TON = 1 400 000),²⁶ p-Si/mesoTiO₂-NiP system described by Reisner (80% FY, TON = 646)⁴⁹ and the p-Si/mesoTiO₂-Co_{C11P} photocathode reported recently by us (>80% FY, TON > 200),²³ the performance of the hybrid CuFe_xO_y|TiO₂-CoHEC (54–88% FY, TON = 90) may

appear limited. However, previous literature data were measured at an applied potential of 0–0.1 V vs. RHE, *i.e.* with a much larger driving force compared to data reported here, which were measured at 0.4 V vs. RHE. Besides stability and selectivity, important metrics for the photocathode performances are indeed its onset photopotential and the maximum photocurrent density. While p-Si and Cu₂O-based photocathodes display limited onset photocurrent potentials (<0.4 V vs. RHE), the CuFe_xO_y-based ones reported here perform better with 0.7 V vs. RHE onset photocurrent potential. Although not as efficient as those displayed by WSe₂⁵⁸ or (CuGa_{1-y}In_y)_{1-x}Zn_{2x}S₂⁵⁹ photocathodes, such a positive value vs. RHE compares well with the onset photocurrent potentials of dye-sensitized NiO^{35,60–62} and CuCrO₂⁶³ or CuGaO₂²⁵ photocathodes. However, the maximum photocurrent reported here is already one order of magnitude higher compared to dye-sensitized molecular photocathodes and there is still room for improvement through the composition optimization and nanostructuring. The cost and scalability of CuFe_xO_y materials are similar to those of the above mentioned transparent semiconducting oxides, while there is no need to sensitize them as they naturally absorb in the visible region. Compared with expensive III–V semiconductors such as InP, GaP or GaInP₂, such ternary oxide-based semiconductors therefore prove promising for implementation in photoelectrochemical cells.^{7,25,27,64}

Conclusion

We successfully designed a hybrid photocathode for H₂ evolution exclusively based on Earth-abundant elements. 300 nm thick CuFe_xO_y films indeed behave as a visible light absorber, perform charge separation and transport electron to the interface with the electrolyte where molecular catalyst can be positioned, for H₂ evolution to occur at a potential 860 mV more positive compared to the dark conditions. Atomic layer deposition of amorphous TiO₂ layer on top of the CuFe_xO_y film enhances the efficiency and improves the stability of the photoelectrode. Nevertheless, the stability of CuFe_xO_y material remains challenging and must be further improved for PEC applications, calling for a better understanding of the modifications of CuFe_xO_y films upon heat, irradiation or electrochemical treatment.

Conflicts of interest

Q4 ■■■■

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