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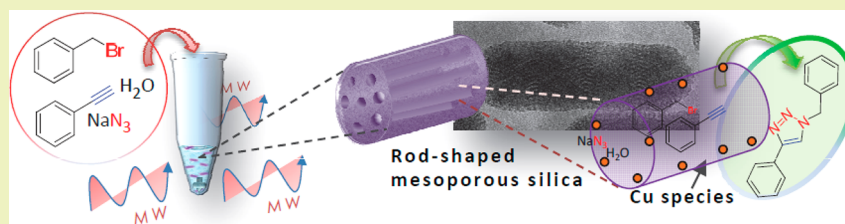
Copper-Containing Rod-Shaped Nanosized Silica Particles for Microwave-Assisted Synthesis of Triazoles in Aqueous Solution

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ABSTRACT: Cu-doped rod-shaped mesoporous silica nanoparticles (Cu-RMSN) were synthesized by a new one-step approach based on *in situ* functionalization procedure referring to the phenomenon of micellar solubilization. SEM and TEM studies revealed the rod-like morphology of uniformly sized particles with longitudinal mesopore channels. The BET specific surface areas were near $500 \text{ m}^2 \text{ g}^{-1}$ and the average pore diameter varied from 3 to 3.6 nm. The composite Cu-RMSN proved to be an efficient heterogeneous catalyst for a microwave-assisted three-component 1,3-dipolar cycloaddition reaction in aqueous solution. The one-pot preparation of 1,4-dibstituted-1,2,3-triazole derivatives was straightforward and high yielding, due to the high copper dispersion at the pore surface resulting in the high accessibility of the active sites. The efficiency of the catalyst was also demonstrated upon recycling, making such synthesis a truly green process with marked step and solvent economy and important waste reduction.

KEYWORDS: Heterogeneous catalysis, Multicomponent reactions (MCR), Click chemistry, Microwaves, Mesoporous silica, Nanoparticles, Micellar solubilization

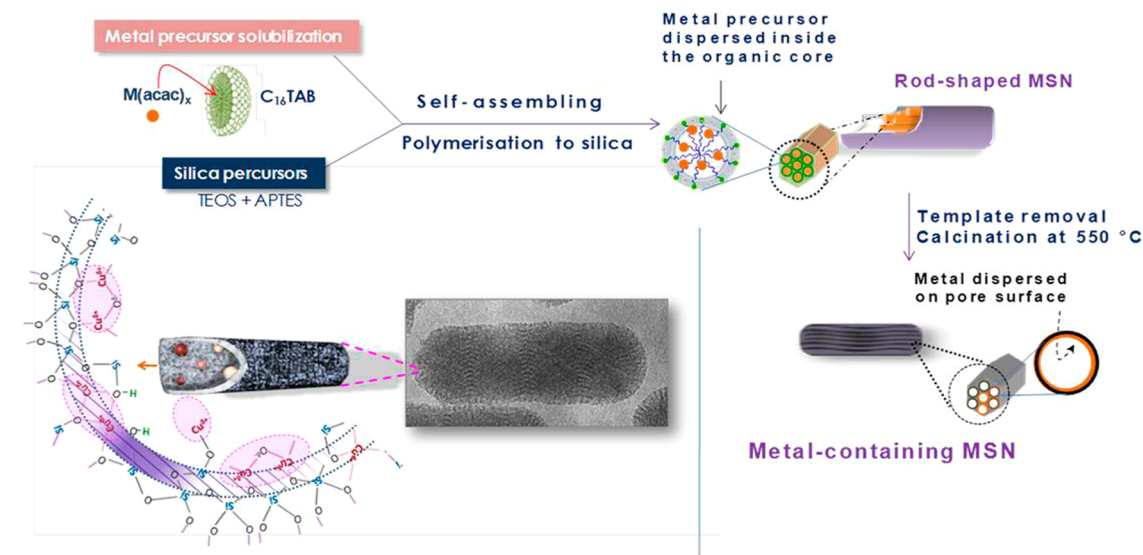
INTRODUCTION

The design of nanoporous materials with the simultaneous control of the particle size, morphology and porosity remains one of the key elements in the context of heterogeneous catalysis applications in solution.^{1,2} Indeed, the resistance to mass transfer of solutes is often large enough to impede the accessibility to the surface reactive sites occurring by a diffusion process at the solid-surface interface.³ Mesoporous silica nanoparticles (MSN) display great potential advantages in liquid-phase catalytic processes, because particle size and shape can noticeably modify the transport of the solutes from the bulk solution to the internal pore surface. It is thus expected that morphology be of great importance in the development of more efficient heterogeneous catalysts.^{4–6} Although, rod-shaped mesoporous silica nanoparticles maintaining a lengthwise alignment of uniform pore channels are of high current interest in the context of their potential applications in catalysis,^{1,7,8} separation,⁹ sensing,^{10,11} and nanomedicine,^{12–16} there are only few reports investigating the catalytic efficiency of rod-shaped MSN as supports of active catalytic

centers.^{1,2,8,17} It was recently reported that accessibility to acid sites in Al-mesoporous silica nanoparticles could be correlated with the particle size rather than with their morphology.⁸ Compared to mesoporous silica spheres of several hundred nanometers, rod-shaped MSN possessed in fact diameters below hundred nanometers, which resulted in a high accessibility to the active sites.

Moreover, another key focus of MSN applications as catalysts was the introduction of either chemical functionalities or heteroatoms into the silica framework to improve the weak reactivity of silica surfaces.^{18–22} Among the various strategies for the synthesis of small mesoporous silica particles,²³ the preparation of nanosized mesoporous silica spheres with a controlled size and morphology, as well as regular pore ordering includes the use of base-mediated conditions (i.e., modified Stöber route),^{24,25} which are not compatible with

Scheme 1. Synthesis Pathway of Metal-Rod Shaped Mesoporous Silica Nanoparticles



common *in situ* incorporation of ionic metal precursors. Transition metal inclusion in the MSN framework can be also achieved by a postsynthesis impregnation; nevertheless, the backfilling approach risks clogging of the matrix pores, thereby reducing the pore accessibility. In view of the one-step approach strategy, a new *in situ* functionalization procedure has been developed in our institute based on the phenomenon of micellar solubilization.^{26,27} This combined synthesis and functionalization approach, depicted in Scheme 1, was for the first time, to the best of our knowledge, adapted for preparing copper-containing rod-shaped MSN with perfect conservation of pore-ordering, shape and monodispersity, and high intrapore metal dispersion. For this purpose, templating micelles of a cationic surfactant *in situ* loaded with a given amount of copper precursor were expected to play a double role of “soft porogens” to create nanoscopic pore structure and precursor-delivery “capsules” for targeted functionalization of the future internal surfaces. Location of the copper precursor inside the micelles was conceived in a way to preclude the metal species from being inserted into the bulk of the solid during silica condensation.

It is well-known that heterogeneous catalysts supporting copper species may regioselectively catalyze the Huisgen 1,3-dipolar cycloaddition of (organic) azides and terminal alkynes (CuAAC)^{28–33} in pure solvents^{34,35} or solvent mixtures^{36,37} to achieve 1,2,3-triazole derivatives, an important class of bioactive compounds with, among others, antiallergenic,³⁸ antibacterial,^{39–41} or anti-HIV^{42,43} activities. In addition, immobilized copper has proven to be very efficient in catalyzing the three-component reactions of organic halides, sodium azide, and terminal alkynes, allowing the one-pot synthesis of triazoles.^{44–50} This reaction pathway has the advantage of eliminating the use and handling of explosive and hazardous organic azides, as they are generated *in situ* via a substitution mechanism usually catalyzed by a metallic Lewis acid.

The present paper reports on the one-pot synthesis of Cu-doped rod-shaped mesoporous silica nanoparticles (Cu-RMSN) with active sites located at the internal surface and thus easily accessible by the reactants, as required for specific surface reactivity in the liquid phase. The classical characterization methods were employed to determine the particle size

distribution, textural properties, degree of Cu incorporation into the silica framework, and accessibility to the active sites in the resulting materials. An eco-friendly microwave-assisted multicomponent reaction in water was subsequently tested, without addition of any other cosolvent or ligand. This “green” approach succeeded in copper-mediated cycloaddition of alkynes, benzyl or alkyl halides, and sodium azide making use of Cu-RMSN as a heterogeneous and recyclable catalyst in aqueous solution, for the first time.

■ MATERIALS AND METHODS

All reagents were commercially available and used without any further purification.

Materials Characterizations. Scanning electron microscopy (SEM) and transmission electron microscopy images (TEM) were obtained with a Hitachi S-2600N (using a secondary electron scintillator as detector) and a JEOL 1200 EX microscope, respectively. For the purpose of TEM analysis, the sample particles were dispersed in ethanol and then dropped onto copper grids with porous carbon films. The specific surface area and pore structure parameters of the materials studied were determined from the measurements of nitrogen adsorption–desorption at 77 K with the use of a Micromeritics ASAP 2020 V3.00 H unit. Prior to the sorption experiment, the sample (about 40 mg) was evacuated under vacuum at 523 K for 12 h. In view of applications of the present Cu-RMSN materials as catalysts, the number of acid sites was assessed and attributed to the number of accessible copper atoms at the solid surface by assuming that pure silica displayed a very limited number of acid sites.²⁶ Measurements of the two-cycle adsorption of gaseous ammonia at 373 K were achieved with the aid of a Micromeritics ASAP 2010 V2.02 and the results served to determine the number of acid sites in the material surface. The sample was outgassed prior to the measurements at 473 K for 2 h at a rate of 5 K·min⁻¹. Afterward, the two-cycle adsorption experiments were performed under equilibrium adsorption conditions at 373 K with gaseous NH₃ of ultrahigh purity. At the end of the first adsorption cycle, the sample was outgassed under vacuum at 373 K for 60 min and a second adsorption cycle was then performed at the same temperature. It is usually assumed that the linear section distinguished in both isotherms at higher equilibrium pressures may be ascribed to physical adsorption on nonreactive surface sites. Therefore, the two straight lines were extrapolated to zero pressure. The difference in the amount adsorbed for points at which the curves reach zero pressure was ascribed to the irreversible chemisorption of basic NH₃ at the solid–gas interface and thereby to the total number of acid sites at the solid surface per unit mass of the adsorbent.

Small-angle powder X-ray diffraction (XRD) measurements were performed with a PANalytical X'Pert MPD (Philips 1710) diffractometer controlled by a computer. The XRD patterns were collected using a Cu $K\alpha$ radiation, the 2θ diffraction angles being recorded at a rate of $0.5^\circ \text{ min}^{-1}$ in the range from 1° to 7° . For some selected samples containing copper, the XRD experiment was continued in the 2θ range up to 60° . Temperature-programmed reduction and metal dispersion experiments were also carried out in the temperature range from 35 to 800 $^\circ\text{C}$ (heating rate 2 K min^{-1}) under a flow of H_2 utilizing a Micromeritics AutoChem 2910 equipped with a Dewar flask containing cold medium to trap water. ^1H NMR spectra were obtained at 297 K using an AC 300L Bruker spectrometer operating at a Larmor frequency of 69.19 MHz (15 000 signal acquisitions, recycle delay of 8 s between $\pi/2$ pulses). NMR spectra were recorded at room temperature with the appropriate deuterated solvent (CDCl_3 or $\text{DMSO}-d_6$). Chemical shifts (δ) of ^1H NMR and ^{13}C NMR spectra are reported in ppm relative to residual solvent signals (CHCl_3 in CDCl_3 : $\delta = 7.27$ ppm for ^1H and CDCl_3 : $\delta = 77.04$ ppm for ^{13}C NMR); data are reported as br = broad, s = singlet, d = doublet, t = triplet, q = quartet, m = multiplet; J values are given in Hz. ^1H and ^{13}C NMR spectra were registered at 300 or 400 MHz. The nature of analytically pure final products was assessed by comparison of their ^1H NMR data previously described in the literature and by their fragmentation in the LC/MS equipment. Conversions and yields were calculated by adding CH_2Br_2 (0.25 mmol; 20 μL) as an internal standard. Analytical high performance liquid chromatography (HPLC) was performed using a CHROMOLITH RP18 column (50 \times 4.6 mm), variable wavelength diode, flow of 5 mL min^{-1} , linear gradient CH_3CN in water 0–100% (+0.1% TFA) in 4.5 min. The LC-MS analysis was performed by means of HPLC Waters Alliance 2695 (UV Waters 2489), column Onyx C18 (25 \times 4.6 mm), flow of 3 mL min^{-1} , linear gradient CH_3CN in water 0–100% (+0.1% HCO_2H) in 2.5 min. Mass spectra (electrospray ionization mode, ESI-MS) were recorded on a Micromass (Manchester, UK) Q-TOF quadrupole mass spectrometer fitted with an electrospray interface. The mass spectrometer was calibrated in the positive- and negative-ion ESI modes. Samples were dissolved in a $\text{H}_2\text{O}/\text{CH}_3\text{CN}$ (50/50 v/v) mixture. Microwave-assisted reactions were performed in sealed vessels with a Biotage Initiator 60 EXP instrument. The temperature was measured with an IR sensor at the outer surface of the reaction vial. Thermal gravimetric analysis (TGA) and differential thermal analysis (DTA) were carried out on a SETARAM Labsys apparatus in order to study changes in the chemical properties of the pristine samples and those containing Cu species as a function of increasing temperature. A given mass of solid sample was placed in a 80 μL alumina crucible. The TGA and DTA curves were obtained up to 873 K in flowing nitrogen using a constant heating rate of $10^\circ \text{ min}^{-1}$. Ultraviolet–visible (UV–vis) absorption spectra were collected with an UV–visible Jasco V-670 spectrophotometer. Leaching of Cu from Cu-RMSN was assessed by inductive coupled plasma (ICP-MS, Agilent 7700x) analysis of the aqueous phase; the instrument detection limit for copper was below 30 ppt.

Catalyst Preparation. The Cu-doped mesoporous silica rod-shaped nanoparticles (Cu-RMSN) were prepared by a one-step synthesis-functionalization procedure, based on the co-condensation of (3-aminopropyl)triethoxysilane (APTES) and tetraethyl orthosilicate (TEOS) as a silicon source. In a typical synthesis, 0.05 g of hexadecyltrimethylammonium bromide (CTAB) was dissolved in 50 mL of deionized water under continuous stirring at room temperature. Then, $\text{Cu}(\text{acac})_2$, previously dissolved in 200 μL of chloroform, was added to the above-mentioned mixture under vigorous stirring at 343 K. This step corresponded to the solubilization of the copper precursor in micellar aggregates. 575 μL TEOS and 80 μL APTES were subsequently added to the solution at 313 K under intensive stirring. Hydrolysis and polycondensation steps were carried out at pH adjusted to 12 with sodium hydroxide. The molar composition of the final solution was as follows: (i) CTAB:NaOH:TEOS:APTES:Cu-(acac) $_2 = 1.97 \times 10^{-4}$: 1.9×10^{-2} : 2.5×10^{-4} ; (ii) X ($X = 2.0 \times 10^{-5}$; 2.7×10^{-5} ; 3.4×10^{-5}). The colloidal mixture was stirred for 2 additional hours. Afterward, the resulting white precipitate was

recovered by centrifugation and dried in air at 323 K. The template was eliminated from Cu-RMSN by air calcinations at 723 K for 8 h with a heated rate of 1 K min^{-1} . A fine slightly blue powder of Cu-RMSN was finally obtained. The three samples were designated Cu-RMSN- X with X corresponding to the initial Si:Cu ratio (i.e., 190, 112, 84) in the gel synthesis. The pure siliceous material designated as RMSN was prepared using the previous procedure except that the Cu source was excluded from the gel synthesis.

General Procedure for the One-Pot Synthesis of 1,4-Disubstituted-1,2,3-triazoles under Microwave Irradiation.

Cu-RMSN catalyst (35 mg, 1.25 mol %) was charged in a microwave Pyrex vial and activated throughout gassing under vacuum at 473 K for 8 h. Then, the appropriate alkyl or benzyl halide (0.2 mmol), alkyne (0.2 mmol), sodium azide (0.2 mmol), and sodium ascorbate (0.2 mmol) were added, together with about 2 mL of water. The reaction mixture was heated under microwave irradiation⁵¹ with constant stirring (900 rpm) at 333 K for the specified time. At the end of the reaction, $\text{CH}_2\text{Cl}_2/\text{Et}_2\text{O}$ (1:4 v/v, 5 mL \times 3) was added into the vial and the heterogeneous system was again stirred for 1 min. Then, the upper organic layers were recovered by decantation, dried on MgSO_4 , filtered, and evaporated under vacuum to collect the corresponding 1,4-disubstituted-1,2,3-triazole as a white solid.

General Procedure for the Cu-RMSN Catalyst Recycling. To get rid of all water-soluble species from the catalyst surface, the catalyst was washed three times with $3 \times 5 \text{ mL}$ of water added under constant stirring to a given Pyrex vial containing the catalyst issued from the previous run. The aqueous phase was removed each time by centrifugation and at least the catalyst was lyophilized. The resulting solid catalyst was reactivated as described previously, before being reused in a subsequent run following the same operating procedure.

RESULTS AND DISCUSSION

As described previously,²⁷ the one-pot synthesis-functionalization approach allowing *in situ* metal incorporation in mesoporous silica led to Cu-containing nanosized silica objects having different morphologies. Figure 1 shows SEM micro-

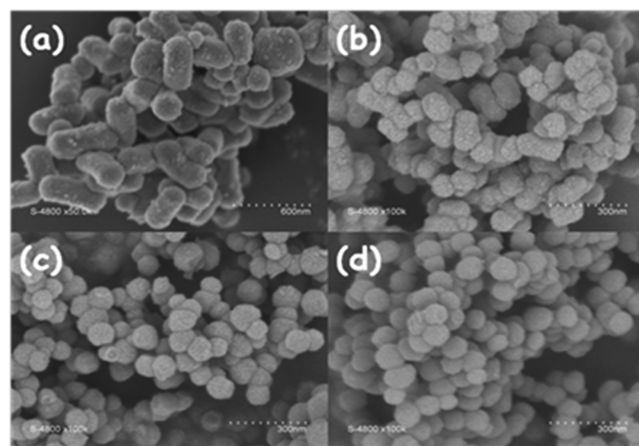


Figure 1. Scanning electron microscope images of nanosized silica objects prepared by a one-pot synthesis-functionalization method and containing different amounts of copper, as defined by the following molar Si:Cu ratio: (a) purely siliceous sample, (b) Si:Cu = 190, (c) Si:Cu = 112, and (d) Si:Cu = 84.

graphs of the resulting particles, which are either rod-like or spherical, depending on the amount of metal incorporation dissolved in chloroform. Apparently, the rod-to-sphere shape transition was favored by higher copper contents. Irrespective of their shape, the particles appeared quite homogeneous in size.

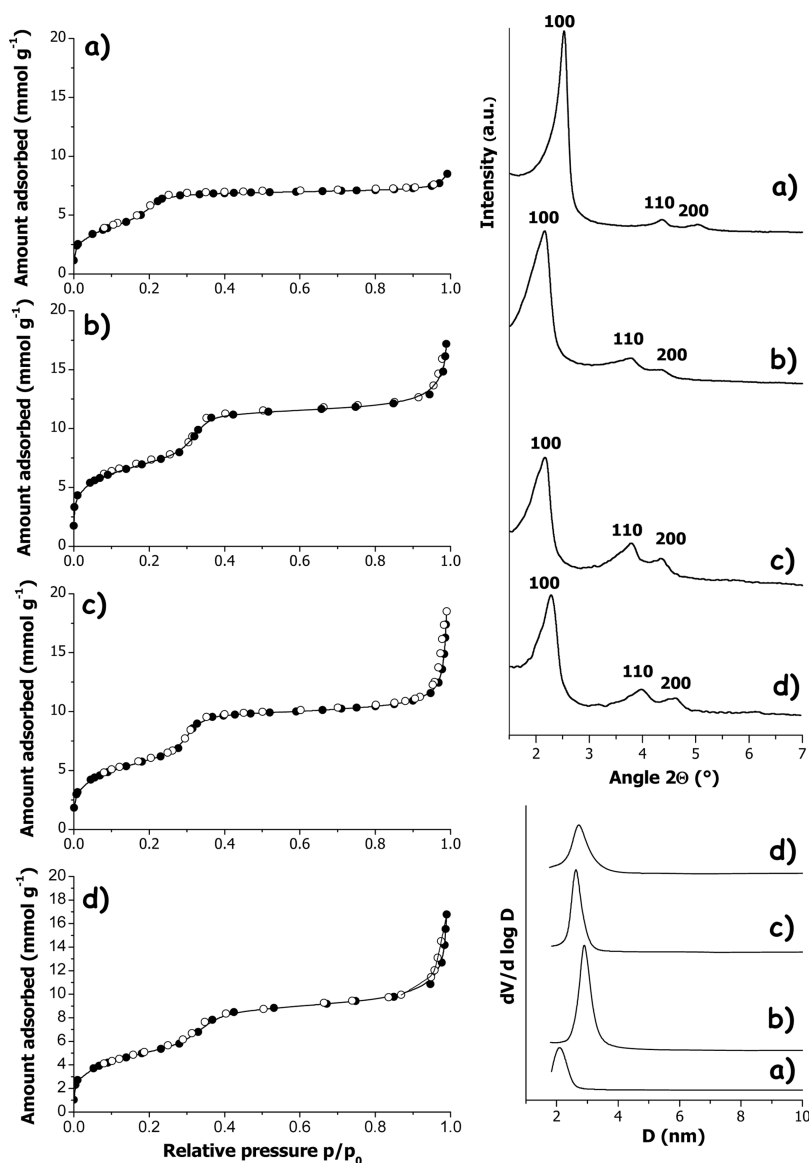


Figure 2. Adsorption–desorption isotherms of gaseous nitrogen at 77 K (4 left panels), powder XRD patterns (upper-right panel), and BJH pore size distributions based on the desorption branch of N₂ adsorption isotherms (lower-right panel) for silica materials containing different amounts of copper, as defined by the following molar Si:Cu ratio: (a) purely siliceous sample, (b) Si:Cu = 190, (c) Si:Cu = 112, and (d) Si:Cu = 84.

The appropriate image processing gave the following size parameters: purely siliceous particles were 450–500 nm in length and 160–190 nm in width; rod-like particles with the smallest Cu content were 160–200 nm in length and 80–100 nm in width; spherical particles were in the 75–110 and 75–100 nm diameter range with a Si:Cu ratio of 112 and 84, respectively. This indicated that the increased copper incorporation resulted in the concomitant decrease in the particle size, the effect being less pronounced within the spherical shape category.

Changes in the pore arrangement and porosity parameters could be followed by the adsorption–desorption isotherms analysis of gaseous nitrogen at 77 K, powder XRD patterns, and BJH pore size distributions as reported in Figure 2. The shape of the nitrogen adsorption curves was characteristic of silica materials possessing intra particle tubular mesopores at the lower end of the mesoporous regime, some residual inter particle porosity (Cu-containing spherical particles in particular), and a significant external surface area.^{26,52–54} A gradual

increase in the amount of gaseous nitrogen adsorbed in the capillary condensation region indicated the intra particle mesopores were somehow heterogeneous, especially for the sample, which contained the highest copper amount. This heterogeneity was illustrated by the width of the BJH distribution curves. An additional characterization procedure based on the α_s -method was performed to determine the “average” pore structure (see Figure S1 in the Supporting Information), thus providing the following parameters: the micropore volume, V_{mic} , the area of the mesopore walls, S_{mes} , the external surface, S_{ext} , and the volume of mesopores, V_{mes} , as well as the mean pore diameter, d_m , calculated assuming a cylindrical pore shape.^{53,55} The X-ray diffraction patterns clearly showed the first-order scattering peak (d_{100}), whereas the low intensity (110) and (200) peaks were still distinguishable from the instrumental noise.

These XRD results corresponded to a 2D hexagonal mesostructure with a $p6$ mm space group. Therefore, the porosity of the four silica samples might be qualified of the

MCM-41 type. All textural parameters inferred from the analysis of XRD patterns and N₂ adsorption–desorption isotherms are reported in Table 1. It was observed that copper

Table 1. Surface Area and Pore Structure Parameters of the Calcined Silica Particles Containing Various Initial Amounts of Copper^a

Si:Cu ratio	no Cu	190	112	84
d_{100} spacing (nm)	3.5	4.10	4.01	3.87
S_{BET} (m ² g ⁻¹)	410	485	570	420
d_{BJH} (nm)	2.06	2.90	2.61	2.69
V_{mic} (cm ³ g ⁻¹)	0.02	0.03	0.05	0.02
S_{ext} (m ² g ⁻¹)	15	47 ^a	40 ^a	54 ^a
S_{mes} (m ² g ⁻¹)	275	326 ^a	395 ^a	294 ^a
V_{mes} (cm ³ g ⁻¹)	0.20	0.30 ^a	0.32 ^a	0.26 ^a
d_{m} (nm)	3	3.6 ^a	3.3 ^a	3.6 ^a

^aBecause the ascending part of the adsorption isotherm and the hysteresis loop at relative pressures close to unity do not allow these parameters to be determined unambiguously by applying the α_s -plot procedure, the values reported in the table should be treated with caution and only general trends are reliable. ^aAs indicated by the molar Si:Cu ratio.

incorporation during the one-pot synthesis had a positive effect on promoting porosity formation. In any case, the pore size and the specific surface area increased, compared to the corresponding textural parameters characterizing the pristine sample containing no metal at all. These trends depended on both the overall metal content and the particle morphology. The greatest values of the surface area and pore volume parameters were obtained for the Si:Cu = 112 sample. The Cu-containing rod-like particles (i.e., Si:Cu = 190) had the largest porosity, providing easier access to the internal porous space. Obviously, the selection of materials for applications in heterogeneous catalysis cannot be limited only to textural parameters. Such other characteristics like nature, distribution, and accessibility of the surface reactive sites are also of great significance. Further catalytic tests in aqueous medium revealed the superior performance of the Si:Cu = 190 sample. Therefore, the surface properties of this material are described in more details in the following section.

Surface Properties of the Cu-containing Rod-like Particles. Further confirmation of the quasi-hexagonal order in the arrangement of tubular pores on the surface of silica particles could be obtained from the analysis of TEM micrographs (Figure 3). Different regions showing more or less marked alignment of pore channels with a uniform cross-sectional diameter were visible (Figure 3). For numerous rods orientated horizontally, pore channels parallel to each other were extended along the entire length of the particle. A few particles had adopted a vertical orientation on the TEM holder and a quasi-hexagonal alignment of uniform channels could be seen from the top. In some cases, partial “finger-print” structures were also observed indicating some disorders in the channel alignment. Moreover, no metal containing clusters could be revealed (Figure 3), which points out a good dispersion of the copper species in the silica matrix.

More information about the nature and location of the copper species was obtained from the diffuse reflectance UV–vis spectrum, powder XRD pattern in the wide-range region, H₂-TPR profile, as well as TGA/TDA profiles acquired from the Si:Cu = 190 sample (Figure 4). The UV–vis spectrum

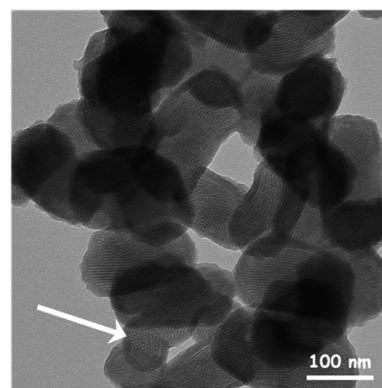


Figure 3. Transmission electron micrographs of nanosized rod-like silica objects containing copper (Si:Cu = 190). The white arrow indicates a region where a particle is oriented vertically thus showing a quasi-hexagonal alignment of pore channels.

displayed features similar to those observed previously:²⁷ a main band representing a right-skewed peak with a maximum at about 225 nm, a weak and very broad band at high wavelength values beyond 600 nm (Figure 4a). The main band could be considered as indicative of the ligand-to-metal charge transfer between Cu(II) ion and oxygen in mononuclear species located at the internal surface.^{27,56} The XRD pattern showed no peak in the 2θ range between 30° and 60° (Figure 4b), which argues against the presence of extra-framework metal oxide crystalline aggregates within the material.⁵⁷ The TGA profile of the calcined Cu-RMSN displayed no weight loss that could be ascribed to the decomposition of remaining surfactant (Figure 4d). The weight loss that occurred at temperatures close to 373 K could be assigned to the release of physisorbed water. The TPR profile showed a broad and asymmetric peak with a temperature at maximum hydrogen consumption T_{m} around 418 K (Figure 4c), usually attributed to the reduction of Cu(II) or finely dispersed CuO to copper metal in the channels of MCM-41 materials.^{19,20} Because the temperature of reduction was lower than that previously recorded for the spherical silica particles achieved by the same synthesis-functionalization procedure (i.e., 485 K²⁷), it appeared that, in the present case, copper oxide was located in readily accessible positions well dispersed at the pore surface. Moreover, on the basis of the results of EDX and those of two-cycle ammonia adsorption onto Cu-RMSN (Si:Cu = 190) (Figure 5), the accessibility of copper surface sites was evaluated by assuming that the ratio between the number of Cu atoms calculated from the number of acid sites, N_{acid} , and the theoretical number of Cu atoms calculated from the EDX results corresponded to the number of Cu atoms accessible to gaseous NH₃ molecules. In the case of RMSN (Si:Cu = 190), the N_{acid} value was of 70 $\mu\text{mol g}^{-1}$ (i.e., 4.4 mg of Cu per gram of catalyst). When ascribed to the number of accessible copper atoms, this gave 85% accessibility with respect to the total number of copper atoms.

Catalytic Performance of the Cu-containing Rod-like Particles. The rod-like mesoporous silica nanoparticles with copper functionalities at the pore surface (Cu-RMSN), achieved in a one-step approach, were evaluated as nanocatalyst in the eco-friendly version of the Huisgen 1,3-dipolar cycloaddition (click reaction) leading to 1,4-disubstituted-1H-1,2,3-triazoles. The regioselectivity was studied by making use of organic halides and terminal alkynes with sodium azide in water under microwave irradiation. This multicomponent

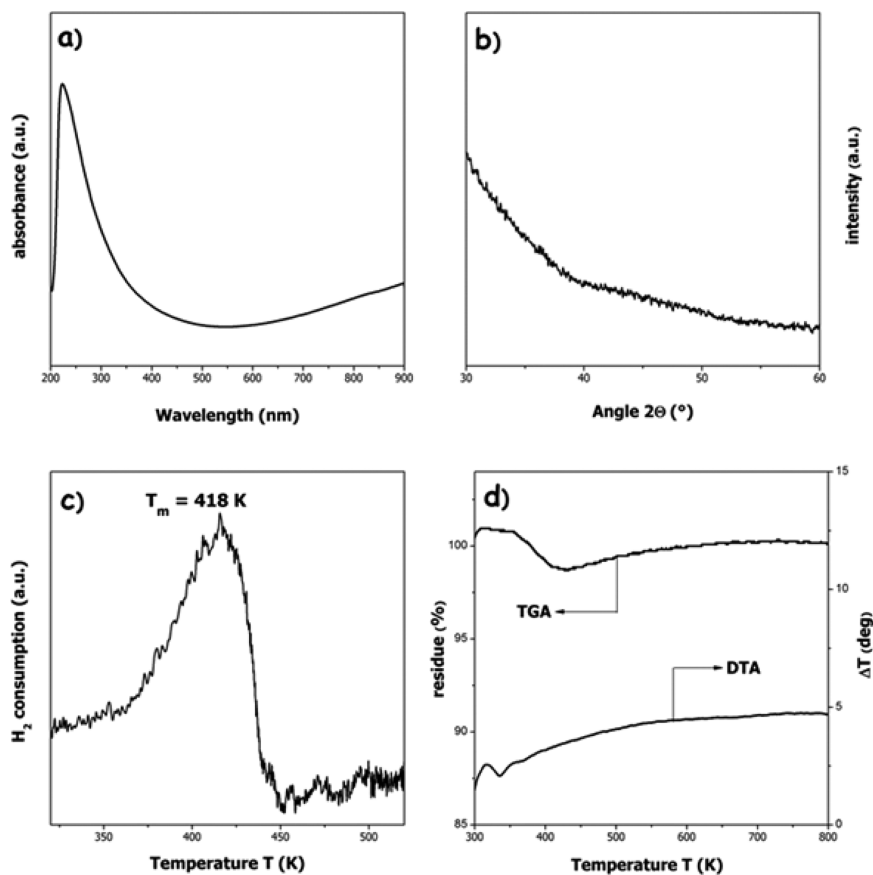


Figure 4. (a) Diffuse reflectance UV–vis spectrum; (b) wide-angle range X-ray powder diffraction pattern; (c) temperature-programmed desorption (H_2 -TPR) profile; (d) results of thermogravimetric analysis for the rod-like silica particles containing copper (Si:Cu = 190).

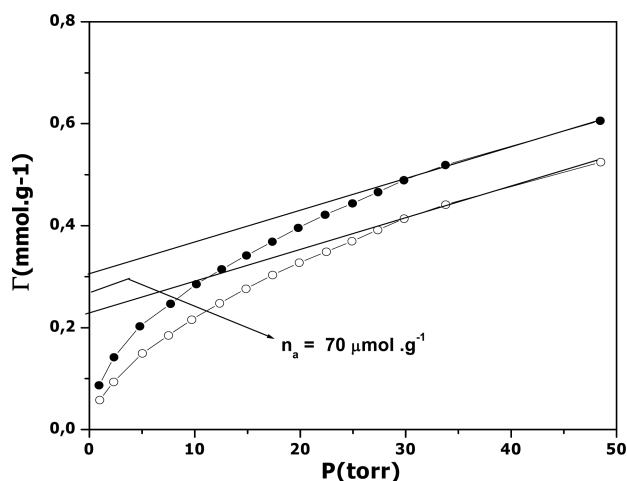
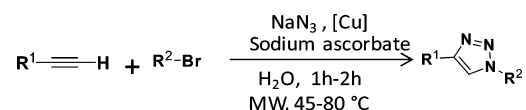


Figure 5. First-cycle (●) and second-cycle (○) adsorption isotherms for gaseous ammonia onto R-CuMSN (Si:Cu = 190) at 373 K.

reaction followed a two reactions cascade where the Cu-RMSN promoted first the nucleophilic substitution for the *in situ* formation of the organic azide and then catalyzed cyclo-addition.^{44,58} This avoided the isolation of organic azides (generally stable either in water or in the presence of oxygen under most reaction conditions), some lower alkyl azides decomposing rapidly with a risk of explosion.⁵⁹ As representative substrates, benzyl bromide and phenyl acetylene were added in stoichiometric amounts to optimize the reaction conditions with NaN_3 under microwave irradiation in aqueous

medium (Table 2). Because the reaction had to be performed in the absence of catalyst, no triazole product was observed (Table 2, entry 1). As expected,⁴⁴ the solvolysis product benzyl alcohol was formed. By irradiating the mixture for 30 min at 60 °C, in the presence of Cu-RMSN catalyst (1.25 mol %), the corresponding 1,4-disubstituted-1*H*-1,2,3-triazole was regioselectively obtained with a high yield (97%, entry 2). The Cu-RMSN catalyst played a double role: it catalyzed the reaction^{45,60} and avoided the formation of 1,5-disubstituted triazolyl regioisomer that had never been observed in the crude mixture,^{29,61} as previously reported. However, shortening the reaction time to 15 min decreased the yield (entry 3). In the absence of sodium ascorbate [Cu(II) reducing agent], the reaction did not proceed as previously reported.²⁹ This result ruled out the presence of Cu(I) entities dispersed in the silica matrix³⁴ during the oxidative homocoupling of alkynes leading to the diynes formation throughout the Glaser reaction,⁶² which had never been observed in the present study.

For comparison, the reaction was also carried out with the conventional catalyst CuBr salt at 7 and 49 mol % of Cu content (entries 4 and 5) under the same reaction conditions, leading respectively to 29% and 96% yield of triazole 1. This different behavior at low Cu concentrations compared to the conventional Huisgen copper-catalyzed reaction (CuAAC) in DMF solvent^{63,64} under microwave irradiation came from the biphasic character of the reactive medium. Indeed, copper ions were water-soluble, whereas the organic substrates were too hydrophobic to be soluble in aqueous solutions. As a result, the contact between the organic substrates and the copper species was limited and consequently the product yield was quite low.

Table 2. Cu-RMSN-Catalyzed One-Pot Synthesis of 1,4-Disubstituted-1H-Triazoles


entry ^a	R ¹ / R ²	product	T (°C)	t (h)	[Cu] catalyst			
					[CuBr] mol %	yield (%) ^b	[Cu-RMSN] mol %	yield (%) ^b
1	Ph/PhCH ₂	1	60	1	0	0	0	0
2	Ph/PhCH ₂	1	60	0.5	n.p. ^c	n.p. ^c	1.25	97
3	Ph/PhCH ₂	1	60	0.25	n.p. ^c	n.p. ^c	1.25	55
4	Ph/PhCH ₂	1	60	1	7.0	29	n.p. ^c	n.p. ^c
5	Ph/PhCH ₂	1	60	1	49	96	1.25	98
6	Ph/CH ₂ CO ₂ Me	2	60	1	49	86	1.25	96
7	<i>n</i> -Bu/PhCH ₂ ^d	3	60	1	49	78	1.25	79
8	Ph/CH ₂ CO ₂ Bn	4	70	2	49	67	1.25	86
9	Ph/2-I-C ₆ H ₄ CH ₂	5	45	1	49	63	1.25	86
10	Ph/ <i>n</i> -Bu ^e	6	80	2	49	61	1.25	69
11	Ph/ <i>sec</i> -Bu ^e	7	80	2	49	69	1.25	73

^aTypical procedure: terminal alkyne (1.0 equiv), benzyl or alkyl halide (1.0 equiv), sodium azide (1.0 equiv), sodium ascorbate (1.0 equiv), and the catalyst were suspended in water (2 mL) and heated under microwave for the indicated time and temperature. ^bYields were calculated by ¹H NMR using CH₂Br₂ (20 μL, 0.25 mmol) as internal standard. ^cn.p. = not performed. ^d4.0 equiv of benzyl bromide was used. ^e2.0 equiv of alkyl bromide was used.

When increasing the amount of CuBr to 49% (Cu content), the solubility limit was reached with the opportunity of substrate adsorption on the active CuBr particles dispersed in the aqueous solution. As a consequence, in order to compare the action of the Cu-RMSN catalyst with that of the conventional CuBr at 49 mol %, the selected reaction conditions were fixed at 1 h and 60 °C in aqueous solution, under microwave irradiation (entry 5).

The possibility of reacting phenyl acetylene with variation in the halide substrates compatible with the standard reaction conditions was subsequently explored. Aliphatic hex-1-yne reacted similarly to phenyl acetylene, yielding the corresponding triazoles with mostly the same yields and regioselectively (entry 7). The “click reaction” with 2-iodo-benzyl bromide was also efficient and not affected by the steric and electronic effects regarding the substitution on the *ortho*-position (entry 9) or by using functionalized alkyl bromides (entries 6 and 8). Primary or secondary alkyl bromides (entries 10 vs 11) reacted similarly, although the yields were slightly decreased. In all our study, Cu-RMSN showed a higher efficiency and performance than CuBr. Indeed, in all cases the yield of reactions performed with Cu-RMSN was higher compared to that using CuBr. This may be explained by the higher copper dispersion at the pore surface, accounting for the higher accessibility of the active sites, as inferred from the chemisorption measurements (Figure 5). Indeed, CuBr was dissolved in the aqueous phase and was not in direct contact with the nonaqueous miscible substrates. In the case of Cu-RMSN, the substrates were adsorbed at the pore surface of the mesoporous nanoparticles in the vicinity of the Cu catalytic active sites.

Reusability of the Cu-RMSN Catalyst. Considering the great economic and environmental added values of heterogeneous catalysts, it is important to examine their ease of separation, recoverability and reusability. In order to evaluate the reuse of the mesoporous Cu-RMSN catalyst we investigated the “click synthesis” of 1-benzyl-4-phenyl-1H-1,2,3-triazole 1 (Table 2, entry 5) and we adjusted the experimental conditions of the present system. After each cycle, the reaction was followed by extraction of the products and unreacted substrates,

and the catalyst was washed with water, lyophilized, and then reactivated before performing the next run, without any further purification. The performance of the recycled catalyst was tested up to ten successive runs without significant yield loss of the reaction products (Figure 6). These results pointed out the

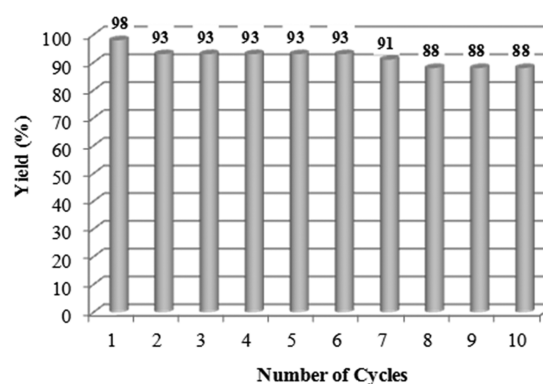


Figure 6. Recyclability of the Cu-RMSN catalytic system during repeated synthesis.

catalytic efficiency of the recovered mesoporous Cu-RMSN, retaining the same activity throughout all reaction cycles, thereby indicating the high stability of the copper species at the pore surface. Moreover, at the end of the first cycle, the reaction mixture contained only 523 ppb of Cu, as indicated by the ICP-MS analysis. These copper species in solution were released during the reaction and originated from leaching. This amount corresponded to 0.16% molar of the copper initially present in Cu-RMSN and did not disturb catalysis of the reaction, since 7% of the Cu content was globally ineffective (Table 2, entry 4). The copper amount after the second cycle was of 61 ppb and only of 9 ppb after 10 successive cycles, indicating the high stability of the Cu-RMSN catalyst (Figure 6). It seems that the copper species localized at the mesopore surface in the rod-like MSN help to prevent the metal leaching and to preserve the high efficiency of the catalyst recycling, while keeping the regioselectivity of the reactions.

■ CONCLUSIONS

The insertion of Cu into the mesopores of rod-shaped mesoporous silica nanoparticles has been achieved with good control of morphological and textural properties. The TPR analysis of Cu-RMSN showed that the well-dispersed Cu species were confined inside the mesopores. The Cu-RMSN sample with a low catalyst loading (1.25%) was used to prepare regioselectively 1,4-disubstituted-1H-1,2,3-triazoles with a high yield via an eco-friendly multicomponent approach, in aqueous media and using microwave heating. The intrinsic and optimal atom economy for the CuAAC reaction, the use of water as a sustainable solvent, the lower catalytic charge compared to the previous reports with similar systems, the solvent economy related to the recovery of the final pure product, together with the ability to reuse the catalyst for at least ten subsequent runs without a marked loss of activity, contributed to the “green” character of the present catalytic process. In the future, preparation of enlarged mesopores represents a challenge to be considered to generalize the catalytic applications of such rod-like mesoporous silica nanoparticles to other reagents and substrates.

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Notes

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■ REFERENCES

- (1) Tsai, C.-H.; Chen, H.-T.; Althaus, S. M.; Mao, K.; Kobayashi, T.; Pruski, M.; Lin, V. S.-Y. Rational Catalyst Design: A Multifunctional Mesoporous Silica Catalyst for Shifting the Reaction Equilibrium by Removal of Byproduct. *ACS Catal.* **2011**, *1*, 729–732.
- (2) Carrillo, A. I.; Serrano, E.; Luque, R.; García-Martínez, J. Microwave-assisted catalysis by iron oxide nanoparticles on MCM-41: Effect of the support morphology. *Appl. Catal., A* **2013**, *453*, 383–390.
- (3) Liu, D.-J.; Wang, J.; Ackerman, D. M.; Slowing, I. I.; Pruski, M.; Chen, H.-T.; Lin, V. S.-Y.; Evans, J. W. Interplay between Anomalous Transport and Catalytic Reaction Kinetics in Single-File Nanoporous Systems. *ACS Catal.* **2011**, *1*, 751–763.
- (4) Chen, Y.; Zhang, F.; Fang, Y.; Zhu, X.; Zhen, W.; Wang, R.; Ma, J. Phosphotungstic acid containing ionic liquid immobilized on magnetic mesoporous silica rod catalyst for the oxidation of dibenzothiophene with H₂O₂. *Catal. Commun.* **2013**, *38*, 54–58.
- (5) Yin, Y.; Liu, K.; Gao, M.; Zhang, L.; Su, H.; Zeng, S. Influence of the structure and morphology of CuO supports on the amount and properties of copper-cerium interfacial sites in inverse CeO₂/CuO catalysts. *J. Mol. Catal. A: Chem.* **2015**, *404–405*, 193–203.
- (6) Gao, D.; Duan, A.; Zhang, X.; Zhao, Z.; E, H.; Li, J.; Wang, H. Synthesis of NiMo catalysts supported on mesoporous Al-SBA-15 with different morphologies and their catalytic performance of DBT HDS. *Appl. Catal., B* **2015**, *165*, 269–284.
- (7) Carrillo, A. I.; Serrano, E.; Luque, R.; García Martínez, J. Introducing catalytic activity in helical nanostructures: microwave assisted oxathioacetalisation catalysed by Al-containing helical mesoporous silicas. *Chem. Commun.* **2010**, *46*, 5163–5165.
- (8) Carrillo, A. I.; Serrano, E.; Serrano-Ruiz, J. C.; Luque, R.; García-Martínez, J. Helical Al- and Ce-MCM-41 materials as novel catalyst for acid and redox processes. *Appl. Catal., A* **2012**, *435–436*, 1–9.
- (9) Minakuchi, H.; Nakanishi, K.; Soga, N.; Ishizuka, N.; Tanaka, N. Octadecylsilylated Porous Silica Rods as Separation Media for Reversed-Phase Liquid Chromatography. *Anal. Chem.* **1996**, *68* (19), 3498–3501.
- (10) Rambaud, F.; Vallé, K.; Thibaud, S.; Julián-López, B.; Sanchez, C. One-Pot Synthesis of Functional Helicoidal Hybrid Organic–Inorganic Nanofibers with Periodically Organized Mesoporosity. *Adv. Funct. Mater.* **2009**, *19* (18), 2896–2905.
- (11) Trewyn, B. G.; Giri, S.; Slowing, I. I.; Lin, V. S. Y. Mesoporous silica nanoparticle based controlled release, drug delivery, and biosensor systems. *Chem. Commun.* **2007**, No. 31, 3236–3245.
- (12) Huang, X.; Teng, X.; Chen, D.; Tang, F.; He, J. The effect of the shape of mesoporous silica nanoparticles on cellular uptake and cell function. *Biomaterials* **2010**, *31* (3), 438–448.
- (13) Huang, X.; Li, L.; Liu, T.; Hao, N.; Liu, H.; Chen, D.; Tang, F. The Shape Effect of Mesoporous Silica Nanoparticles on Biodistribution, Clearance, and Biocompatibility in Vivo. *ACS Nano* **2011**, *5* (7), 5390–5399.
- (14) Xu, C.; Niu, Y.; Popat, A.; Jambhrunkar, S.; Karmakar, S.; Yu, C. Rod-like mesoporous silica nanoparticles with rough surfaces for enhanced cellular delivery. *J. Mater. Chem. B* **2014**, *2* (3), 253–256.
- (15) Yang, P.; Gai, S.; Lin, J. Functionalized mesoporous silica materials for controlled drug delivery. *Chem. Soc. Rev.* **2012**, *41* (9), 3679–3698.
- (16) Meng, H.; Yang, S.; Li, Z.; Xia, T.; Chen, J.; Ji, Z.; Zhang, H.; Wang, X.; Lin, S.; Huang, C.; Zhou, Z. H.; Zink, J. I.; Nel, A. E. Aspect Ratio Determines the Quantity of Mesoporous Silica Nanoparticle Uptake by a Small GTPase-Dependent Macropinocytosis Mechanism. *ACS Nano* **2011**, *5* (6), 4434–4447.
- (17) Fernandes, C. I.; Saraiva, M. S.; Nunes, T. G.; Vaz, P. D.; Nunes, C. D. Highly enantioselective olefin epoxidation controlled by helical confined environments. *J. Catal.* **2014**, *309*, 21–32.
- (18) Bérubé, F.; Khadraoui, A.; Florek, J.; Kaliaguine, S.; Kleitz, F. A generalized method toward high dispersion of transition metals in large pore mesoporous metal oxide/silica hybrids. *J. Colloid Interface Sci.* **2015**, *449*, 102–114.
- (19) Kargol, M.; Zajac, J.; Jones, D. J.; Rozière, J.; Steriotis, T.; Jiménez-López, A.; Rodríguez-Castellón, E. Copper- and Silver-Containing Monolithic Silica-Supported Preparations for Selective Propene–Propane Adsorption on the Gas Phase. *Chem. Mater.* **2005**, *17* (24), 6117–6127.
- (20) Velu, S.; Wang, L.; Okazaki, M.; Suzuki, K.; Tomura, S. Characterization of MCM-41 mesoporous molecular sieves containing copper and zinc and their catalytic performance in the selective oxidation of alcohols to aldehydes. *Microporous Mesoporous Mater.* **2002**, *54* (1–2), 113–126.
- (21) Rascon, F.; Wischert, R.; Coperet, C. Molecular nature of support effects in single-site heterogeneous catalysts: silicav.alumina. *Chem. Sci.* **2011**, *2* (8), 1449–1456.
- (22) Perego, C.; Millini, R. Porous materials in catalysis: challenges for mesoporous materials. *Chem. Soc. Rev.* **2013**, *42* (9), 3956–3976.
- (23) Taguchi, A.; Schüth, F. Ordered mesoporous materials in catalysis. *Microporous Mesoporous Mater.* **2005**, *77* (1), 1–45.
- (24) Valtchev, V.; Tosheva, L. Porous Nanosized Particles: Preparation, Properties, and Applications. *Chem. Rev.* **2013**, *113* (8), 6734–6760.
- (25) Huh, S.; Wiench, J. W.; Yoo, J.-C.; Pruski, M.; Lin, V. S. Y. Organic Functionalization and Morphology Control of Mesoporous Silicas via a Co-Condensation Synthesis Method. *Chem. Mater.* **2003**, *15* (22), 4247–4256.
- (26) Nyalosaso, J. L.; Derrien, G.; Charnay, C.; de Menorval, L.-C.; Zajac, J. Aluminium-derivatized silica monodisperse nanospheres by a one-step synthesis-functionalization method and application as acid catalysts in liquid phase. *J. Mater. Chem.* **2012**, *22* (4), 1459–1468.

- (27) Derrien, G.; Charnay, C.; Zajac, J.; Jones, D. J.; Roziere, J. Copper-containing monodisperse mesoporous silica nanospheres by a smart one-step approach. *Chem. Commun.* **2008**, 27, 3118–20.
- (28) Huisgen, R. Kinetics and reaction mechanisms: Selected examples from the experience of forty years. *Pure Appl. Chem.* **1989**, 61 (4), 613–628.
- (29) Rostovtsev, V. V.; Green, L. G.; Fokin, V. V.; Sharpless, K. B. A Stepwise Huisgen Cycloaddition Process: Copper(I)-Catalyzed Regioselective “Ligation” of Azides and Terminal Alkynes. *Angew. Chem., Int. Ed.* **2002**, 41 (14), 2596–2599.
- (30) Tornøe, C. W.; Christensen, C.; Meldal, M. Peptidotriazoles on solid phase: [1, 2, 3]-triazoles by regioselective copper (I)-catalyzed 1, 3-dipolar cycloadditions of terminal alkynes to azides. *J. Org. Chem.* **2002**, 67 (9), 3057–3064.
- (31) Padwa, A. *1,3-Dipolar cycloaddition chemistry*; Wiley-Interscience: New York, 1984; Vol. 1.
- (32) Bock, V. D.; Hiemstra, H.; van Maarseveen, J. H. CuI-Catalyzed Alkyne–Azide “Click” Cycloadditions from a Mechanistic and Synthetic Perspective. *Eur. J. Org. Chem.* **2006**, 2006 (1), 51–68.
- (33) Mohammed, S.; Padala, A. K.; Dar, B. a.; Singh, B.; Sreedhar, B.; Vishwakarma, R. a.; Bharate, S. B. Recyclable clay supported Cu (II) catalyzed tandem one-pot synthesis of 1-aryl-1,2,3-triazoles. *Tetrahedron* **2012**, 68, 8156–8162.
- (34) Wang, Y.; Liu, J.; Xia, C. Insights into Supported Copper(II)-Catalyzed Azide-Alkyne Cycloaddition in Water. *Adv. Synth. Catal.* **2011**, 353 (9), 1534–1542.
- (35) Lal, S.; Diez-Gonzalez, S. [CuBr(PPh₃)₃] for Azide-Alkyne Cycloaddition Reactions under Strict Click Conditions. *J. Org. Chem.* **2011**, 76 (7), 2367–2373.
- (36) Soltani Rad, M. N.; Behrouz, S.; Doroodmand, M. M.; Movahediyani, A. Copper-doped silica cuprous sulfate (CDSCS) as a highly efficient and new heterogeneous nano catalyst for [3 + 2] Huisgen cycloaddition. *Tetrahedron* **2012**, 68 (38), 7812–7821.
- (37) Yamada, Y. M. A.; Sarker, S. M.; Uozumi, Y. Amphiphilic Self-Assembled Polymeric Copper Catalyst to Parts per Million Levels: Click Chemistry. *J. Am. Chem. Soc.* **2012**, 134 (22), 9285–9290.
- (38) Buckle, D. R.; Outred, D. J.; Rockell, C. J. M.; Smith, H.; Spicer, B. A. Studies on v-triazoles. 7. Antiallergic 9-oxo-1H,9H-benzopyrano-[2,3-d]-v-triazoles. *J. Med. Chem.* **1983**, 26 (2), 251–254.
- (39) Romero, A.; Liang, C.-H.; Chiu, Y.-H.; Yao, S.; Duffield, J.; Sucheck, S. J.; Marby, K.; Rabuka, D.; Leung, P. Y.; Shue, Y.-K.; Ichikawa, Y.; Hwang, C.-K. An efficient entry to new sugar modified ketolide antibiotics. *Tetrahedron Lett.* **2005**, 46 (9), 1483–1487.
- (40) Genin, M. J.; Allwine, D. A.; Anderson, D. J.; Barbachyn, M. R.; Emmert, D. E.; Garmon, S. A.; Graber, D. R.; Grega, K. C.; Hester, J. B.; Hutchinson, D. K.; Morris, J.; Reischer, R. J.; Ford, C. W.; Zurenko, G. E.; Hamel, J. C.; Schaadt, R. D.; Stapert, D.; Yagi, B. H. Substituent Effects on the Antibacterial Activity of Nitrogen-Carbon-Linked (Azolylphenyl)oxazolidinones with Expanded Activity Against the Fastidious Gram-Negative Organisms *Haemophilus influenzae* and *Moraxella catarrhalis*. *J. Med. Chem.* **2000**, 43 (5), 953–970.
- (41) Holla, B. S.; Mahalinga, M.; Karthikeyan, M. S.; Poojary, B.; Akberali, P. M.; Kumari, N. S. Synthesis, characterization and antimicrobial activity of some substituted 1,2,3-triazoles. *Eur. J. Med. Chem.* **2005**, 40 (11), 1173–1178.
- (42) Zhou, L.; Amer, A.; Korn, M.; Burda, R.; Balzarini, J.; De Clercq, E.; Kern, E. R.; Torrence, P. F. Synthesis and Antiviral Activities of 1,2,3-triazole Functionalized Thymidines: 1,3-dipolar Cycloaddition for Efficient Regioselective Diversity Generation. *Antiviral Chem. Chemother.* **2005**, 16 (6), 375–383.
- (43) Alvarez, R.; Velazquez, S.; San-Felix, A.; Aquaro, S.; Clercq, E. D.; Perno, C.-F.; Karlsson, A.; Balzarini, J.; Camarasa, M. J. 1,2,3-Triazole-[2,5-Bis-O-(tert-butyl)dimethylsilyl]-.beta.-D-ribofuranosyl]-3'-spiro-5''-(4''-amino-1'',2''-oxathiole 2'',2''-dioxide) (TSAO) Analogs: Synthesis and Anti-HIV-1 Activity. *J. Med. Chem.* **1994**, 37 (24), 4185–4194.
- (44) Bénétteau, V.; Olmos, A.; Boningari, T.; Sommer, J.; Pale, P. Zeo-click synthesis: CuI-zeolite-catalyzed one-pot two-step synthesis of triazoles from halides and related compounds. *Tetrahedron Lett.* **2010**, 51 (28), 3673–3677.
- (45) Radatz, C. S.; Soares, L. D. A.; Vieira, E. R.; Alves, D.; Russowsky, D.; Schneider, P. H. Recoverable Cu/SiO₂ composite-catalyzed click synthesis of 1,2,3-triazoles in water media. *New J. Chem.* **2014**, 38, 1410–1417.
- (46) Lipshutz, B. H.; Taft, B. R. Heterogeneous Copper-in-Charcoal-Catalyzed Click Chemistry. *Angew. Chem., Int. Ed.* **2006**, 45 (48), 8235–8238.
- (47) Sharghi, H.; Khalifeh, R.; Doroodmand, M. M. Copper Nanoparticles on Charcoal for Multicomponent Catalytic Synthesis of 1,2,3-Triazole Derivatives from Benzyl Halides or Alkyl Halides, Terminal Alkynes and Sodium Azide in Water as a “Green” Solvent. *Adv. Synth. Catal.* **2009**, 351 (1–2), 207–218.
- (48) Alonso, F.; Moglie, Y.; Radivoy, G.; Yus, M. Multicomponent Synthesis of 1,2,3-Triazoles in Water Catalyzed by Copper Nanoparticles on Activated Carbon. *Adv. Synth. Catal.* **2010**, 352 (18), 3208–3214.
- (49) Kale, S.; Kahandal, S.; Disale, S.; Jayaram, R. Conventional and microwave-assisted multicomponent reaction of alkyne, halide and sodium azide catalyzed by copper apatite as heterogeneous base and catalyst in water. *Curr. Chem. Lett.* **2012**, 1, 69–80.
- (50) Nador, F.; Volpe, M. A.; Alonso, F.; Feldhoff, A.; Kirschning, A.; Radivoy, G. Copper nanoparticles supported on silica coated maghemite as versatile, magnetically recoverable and reusable catalyst for alkyne coupling and cycloaddition reactions. *Appl. Catal., A* **2013**, 455, 39–45.
- (51) Baghbanzadeh, M.; Carbone, L.; Cozzoli, P. D.; Kappe, C. O. Microwave-Assisted Synthesis of Colloidal Inorganic Nanocrystals. *Angew. Chem., Int. Ed.* **2011**, 50 (48), 11312–11359.
- (52) Kruk, M.; Jaroniec, M. Gas Adsorption Characterization of Ordered Organic–Inorganic Nanocomposite Materials. *Chem. Mater.* **2001**, 13 (10), 3169–3183.
- (53) Roziere, J.; Brandhorst, M.; Dutartre, R.; Jacquin, M.; Jones, D. J.; Vitse, P.; Zajac, J. Effect of surfactant type, substitution by aluminium and additives on direct liquid crystal templated monolithic silica. *J. Mater. Chem.* **2001**, 11 (12), 3264–3275.
- (54) Tompsett, G. A.; Krogh, L.; Griffin, D. W.; Conner, W. C. Hysteresis and Scanning Behavior of Mesoporous Molecular Sieves. *Langmuir* **2005**, 21 (18), 8214–8225.
- (55) Gregg, S. J.; Sing, K. S. W. *Adsorption, Surface Area, and Porosity*, Second ed.; Academic Press: London, 1982.
- (56) Chmielarz, L.; Kuśtrowski, P.; Dziembaj, R.; Cool, P.; Vansant, E. F. Catalytic performance of various mesoporous silicas modified with copper or iron oxides introduced by different ways in the selective reduction of NO by ammonia. *Appl. Catal., B* **2006**, 62 (3–4), 369–380.
- (57) Trejda, M.; Tuel, A.; Kujawa, J.; Kilos, B.; Ziolk, M. Niobium rich SBA-15 materials – preparation, characterisation and catalytic activity. *Microporous Mesoporous Mater.* **2008**, 110 (2–3), 271–278.
- (58) Lee, C.-T.; Huang, S.; Lipshutz, B. H. Copper-in-Charcoal-Catalyzed, Tandem One-Pot Diazo Transfer-Click Reactions. *Adv. Synth. Catal.* **2009**, 351 (18), 3139–3142.
- (59) Alvarez, S. G.; Alvarez, M. T. A practical procedure for the synthesis of alkyl azides at ambient temperature in dimethyl sulfoxide in high purity and yield. *Synthesis* **1997**, 1997 (4), 413–414.
- (60) Veerakumar, P.; Velayudham, M.; Lu, K.-L.; Rajagopal, S. Highly dispersed silica-supported nanocopper as an efficient heterogeneous catalyst: application in the synthesis of 1, 2, 3-triazoles and thioethers. *Catal. Sci. Technol.* **2011**, 1 (8), 1512–1525.
- (61) Himo, F.; Lovell, T.; Hilgraf, R.; Rostovtsev, V. V.; Noodleman, L.; Sharpless, K. B.; Fokin, V. V. Copper(I)-Catalyzed Synthesis of Azoles. DFT Study Predicts Unprecedented Reactivity and Intermediates. *J. Am. Chem. Soc.* **2005**, 127 (1), 210–216.
- (62) Siemsen, P.; Livingston, R. C.; Diederich, F. Acetylenic coupling: a powerful tool in molecular construction. *Angew. Chem., Int. Ed.* **2000**, 39 (15), 2632–2657.
- (63) Ellanki, A. R.; Islam, A.; Rama, V. S.; Pulipati, R. P.; Rambabu, D.; Rama Krishna, G.; Malla Reddy, C.; Mukkanti, K.; Vanaja, G. R.;

Kalle, A. M.; Shiva Kumar, K.; Pal, M. Solvent effect on copper-catalyzed azide-alkyne cycloaddition (CuAAC): Synthesis of novel triazolyl substituted quinolines as potential anticancer agents. *Bioorg. Med. Chem. Lett.* **2012**, *22* (10), 3455–3459.

(64) Glasnov, T. N.; Kappe, C. O. Microwave-Assisted Click Chemistry for the Preparation of 3- and 4-Triazolyl-2(1H)-quinolones as Potential Fluorescent Probes. *QSAR Comb. Sci.* **2007**, *26* (11–12), 1261–1265.