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# The influence of the recycling stress history on LDPE waste pyrolysis

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## A B S T R A C T

It is elementary to recognize the benefits and the negative impacts of the use of plastic materials on modern societies. Polyethylene (PE) is the major plastic component present in the municipal solid waste. In this paper, two types of low-density PE (LDPE) waste with different mechanical recycling stress histories were used to investigate the influence of recycling cycles on pyrolysis. The kinetic triplet and thermal degradation study were obtained using TGA data. To determine the sample composition and hydrocarbon arrangements, ultimate, proximate and X-ray diffraction analyses were carried out. Taking advantage of these analyses and combining them with differential scanning calorimetry (DSC) data, a series-parallel pyrolysis pathway was formulated. The waste of recycled polyethylene presented low enthalpy of pyrolysis, at about 205 J/g against 299 J/g for a virgin PE. The DSC analyses evidenced a multi-step reaction behavior of the pyrolysis, confirmed by the kinetic study using different isoconversional methods: the waste of recycled polyethylene presented a higher variation of activation energies as a function of the fraction reacted. The main conclusion is that the results suggest that the recycling stress history promotes the increase of long carbon chains while weakening the boundary among the compounds. This explains the fact that recycled waste needs less activation energy than other samples to degrade thermally. Finally, different categories of low-density polyethylene wastes must be considered when dealing with either kinetics or modeling of the product recovery process.

## 1. Introduction

It is elementary to recognize the benefits and, of course, the negative impacts of the use of plastic materials on modern societies. The global production of plastic materials has grown drastically. According to a European report [1], world plastic production has increased by around 40% in the last 10 years and polyethylene (PE) is the major plastic component present in municipal solid waste (MSW) streams. In general, polyethylene accounts for 43% of the total plastic fraction present in MSW [2]. This is the reason for the rising processes of converting polyethylene waste into useful products [3–9]. Since these processes are supposed to convert waste plastics, some characteristics such as density and additives are well known to affect i.e. the rate of conversion and yields of products recovered [10–12]. Another characteristic might play an important role in the conversion processes: the physical stress history experienced by the waste in mechanical recycling. This process typically involves mechanical (separation, grinding, washing, agglutination) and thermal stress (drying, extrusion, and quenching) [3,13,14].

The polyethylene materials entering the mechanical recycling process can be described by the stream in Fig. 1. The virgin PE (VPE) pellets are used as raw material where noble plastic products are manufactured. Inside the industry is generated plastic residue that is well suited for the mechanical recycling process. The main part of plastic wastes is mostly generated by the final consumer. Depending on the infrastructure of each country, the plastic wastes can follow different paths (solid waste management cycle, landfills or waste pickers). The recycled material returns to manufacturing as feedstock producing less noble plastic products such as garbage bags and various packaging. Therefore, two categories of wastes are commonly generated: waste of virgin polyethylene (WVPE) and waste of recycled polyethylene (WRPE). In that way, the mechanical recycling process continues until the plastic waste previously recycled cannot be reused anymore.

Isoconversional methods are commonly used to evaluate the decomposition kinetics of polyethylene. The best known methods recommended by the International Confederation for Thermal Analysis

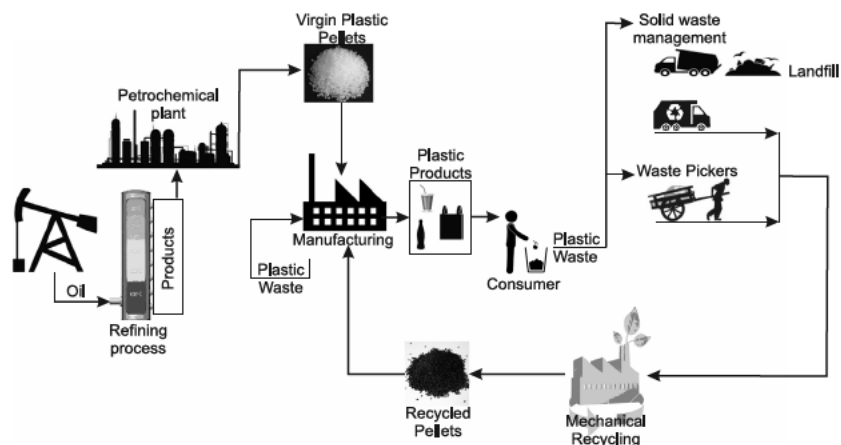


Fig. 1. Lifecycle of PE materials.



Fig. 2. Polyethylene virgin and recycled samples.

Calorimetry (ICTAC) to capture the kinetic triplet (the activation energy,  $E$ , the pre exponential factor,  $A$ , and the reaction model,  $f(\alpha)$ ) and widely used are Friedman (FR) [15], Ozawa Flynn and Wall (OFW) [16–19], Kissinger Akahira and Sunose (KAS) [20,21], and Starink (ST) [22].

The previous studies successfully explained polyethylene kinetics through the evaluation of variable  $E$  and distributed  $A$  with the extent of the progress of the reaction [23–28]. These authors were concerned about verifying the best statistical fit for the kinetic parameters. However, until now, there is no study concerned about how different categories of LDPE wastes can impact both kinetics values and pathways.

The literature oriented mainly consider that LDPE pyrolysis can be modeled as a single step reaction, and it is true in some cases [29,30], but some studies mention complex reaction mechanisms [26,31,32]. According to [33], a significant variation of the kinetics values indicates that a process is kinetically complex, therefore, the consideration of a single step mechanism cannot be applied. The major problem might be the fact that some studies modeled wastes using virgin polyethylene as a sample whereas, in others, there is no clear concern about the origin of these wastes.

The aim of the present study is to understand if the physical stress undergone by two categories of low density polyethylene waste (WVPE, WRPE) within the mechanical recycling cycles has an influence on the kinetics and pyrolysis pathway. The quantitative aspects such as enthalpy of conversion and hydrocarbon arrangement are also investigated. In the entire study process, the low density VPE is used as a reference for establishing a comparison with PE wastes in terms of thermochemical behavior.

The differences between virgin/waste and waste/waste samples are characterized by the x ray diffraction technique and complemented via differential scanning calorimetry. Both techniques have been widely employed to characterize qualitatively and quantitatively the crystalline structures and enthalpy of melting of polyethylene [34,35]. For these reasons, they are applied in the present work also for pyrolysis.

## 2. Material and methods

Independently of the thermal conversion process employed during the treatment of any type of waste, it is important prior knowledge of the ultimate and proximate composition, as well as some other basic thermochemical behavior such as mass loss and heat flux absorbed or released. Thereby, a series of experiments were performed on samples of PE to address the problem in a critical and effective way.

### 2.1. Polyethylene samples

The plastic materials used were samples of LDPE, the virgin from the petrochemical industry and the recycled from municipal solid waste provided by LukPlast Ind. (ES Brazil). VPE has regular geometry (but ton type) with an average diameter of around 4 mm and does not have any pigmentation (transparent color). The recycled samples are a mix of irregular geometries, 2 mm average particle size, and are organized into two types: WVPE, with a type of pigmentation resulting from the first and unique mechanical recycling process of waste VPE, and WRPE which underwent from two to undetermined numbers of mechanical recycling cycles, and carries all types of pigmentation (see Fig. 2).

### 2.2. Methods

The thermal degradation study was carried out by using thermogravimetric analysis (TGA) then the data were processed to obtain the kinetic triplet. X ray diffraction (XRD), ultimate and proximate analyses were applied to investigate if LDPE samples could have a certain degree of differentiation either on composition and hydrocarbon arrangement. The XRD results combined with the analysis of differential scanning calorimetry (DSC) under an inert atmosphere and complementary under an oxidative one were used to formulate the pyrolysis pathway. From DSC curve integration, the average enthalpies of melting and pyrolysis were established. The configurations of each method are presented in Table 1.

**Table 1**  
Experimentation details of PE samples.

Name of experiment	Make	Operating conditions	Reference
TGA	STA 449F3 — NETZSCH	Sample mass: 10–15 mg Crucible: Al <sub>2</sub> O <sub>3</sub> Carrier gas: N <sub>2</sub> , Air, 60 ml/min Temperature range: 20–670 °C Heating rates $\beta$ : 5, 10, 15, and 20 °C/min	ISO 11358 DIN 51006
XRD	ULTIMA IV — RIGAKU	CuK alpha radiation: average length $\lambda = 0.15419$ Theta/2-theta Bragg–Brentano geometry 2-theta data range: 5.00° to 125.00° Step size: 0.02° CuK alpha doublet: wavelengths of 0.154056 (65%) and 0.154439 (35%) nm. Radiation detector: LiF monochromator	Souza et al. [36]
Ultimate analysis	EA 1110 — CE	Default settings	<sup>a</sup> ASTM D5373-16
Proximate analysis	Muffle furnace	Moisture and ash: Weight loss Volatile matter: <sup>b</sup> by difference	ASTM D3173/D3173M-17a/D3174-12 –
DSC	SDT Q600 — TA	Sample mass: 20 mg Crucible: Al <sub>2</sub> O <sub>3</sub> Carrier gas: N <sub>2</sub> , Air, 50 ml/min Temperature range: 20–570 °C Heating rate $\beta$ : 20 °C/min	ASTM E-2160

<sup>a</sup>There is no specific standard for PE.

<sup>b</sup>Based on the assumption that all the carbon and hydrogen, except for the moisture and ash contents, form the volatile matter matrix.

### 2.3. Data processing

The rate of non isothermal solid decomposition,  $d\alpha/dt$ , registered by TGA can be expressed in terms of the rate constant,  $k(T)$ , and the reaction model,  $f(\alpha)$ , Eq. (1). The rate constant represents the dependence of the process rate on temperature and is parametrized through the Arrhenius equation, Eq. (2). The reaction model represents the dependence on the fraction reacted,  $\alpha$ , Eq. (3), which is determined experimentally as a fraction of the total mass loss during the process.

$$\frac{d\alpha}{dt} = k(T)f(\alpha) \quad (1)$$

$$k(T) = A \exp\left(-\frac{E}{\mathfrak{R}T}\right) \quad (2)$$

$$\alpha = \frac{m_0 - m_i}{m_0 - m_f} \quad (3)$$

where  $A$  and  $E$  are kinetic parameters, the pre exponential factor, and the apparent activation energy, respectively,  $\mathfrak{R}$  is the universal gas constant,  $T$  is the absolute temperature, and  $m_i$ ,  $m_0$ , and  $m_f$  represent the instantaneous, initial, and final masses of the sample, respectively.

Combining Eqs. (1) and (2) leads to Eq. (4), that is limited to describe the rate of decomposition of a single step process [33,37]. For constant heating rate non isothermal conditions ( $\beta = dT/dt$ ), Eq. (4) can be written in the form of Eq. (5).

$$\frac{d\alpha}{dt} = A \exp\left(-\frac{E}{\mathfrak{R}T}\right) f(\alpha) \quad (4)$$

$$\beta \frac{d\alpha}{dT} = A \exp\left(-\frac{E}{\mathfrak{R}T}\right) f(\alpha) \quad (5)$$

The degradation characteristics of VPE, WVPE, and WRPE were quantified through several parameters. Onset and offset temperatures,  $T_{on}$  and  $T_{off}$ , related to the start and end of the sample thermal conversion, were obtained from the TGA, its derivatives, DTG, and second time derivatives curves as the scheme demonstrated in Fig. 3.

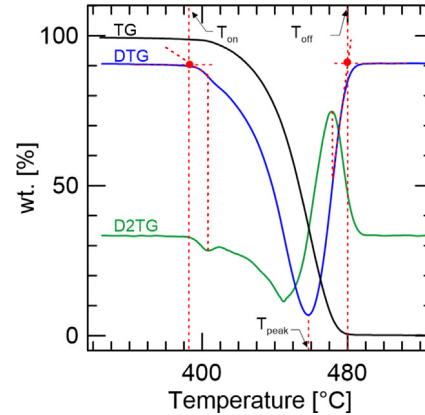


Fig. 3. Method used to determine the onset, offset, and peak temperatures.

The onset temperature was determined by extrapolating the slope of the DTG curve in correspondence with the first peak in second time derivatives curves and up to the zero levels of the DTG axis. The same procedure was used to determine the offset temperature, in this case, picking up the last peak in second time derivatives curves [28,38–40]. The peak temperatures,  $T_{peak}$ , related to the maximum rate of sample degradation, were obtained from DTG curves and correspond to the temperature of their peaks. The total mass losses,  $\Delta_m$ , were obtained from the difference between the initial mass and the mass after pyrolysis.

From the DSC data, the enthalpies of reactions were calculated by integrating the areas of the curves (heat flux versus time interval) using scientific graphing and data analysis software [41]. For that, an isothermal baseline was constructed based on the second derivative of the DSC signal; ergo, the onset and end temperatures of melting and pyrolysis were determined [38].

The diffraction patterns were analyzed using Fityk software, version 0.9.8.57 [41]. Phase identification was performed through the



Hanawalt method [42–44] using the Crystallography Open Database (COD) [45] and PDF 2 database [46]. The quantification of the phases was performed with Rietveld refinement using the FullProf program [47].

## 2.4. Isoconversional methods

Isoconversional methods are an effective tool that can be used to solve Eq. (4) or Eq. (5) and determine the kinetic parameters as a function of the fraction reacted. These methods are based on the assumption that the rate of decomposition, at a constant value of  $\alpha$ , is only a function of the temperature [37]. Taking into account the methodology adopted, the isoconversional methods can be separated into two main categories: differential and integral.

A widely used differential isoconversional method, based on the logarithm of Eq. (4) or Eq. (5), is the Friedman method [15], Eq. (6). By plotting  $\ln(\beta d\alpha/dT)$  against  $1/T$ , the kinetic parameters can be calculated from the slope at each given  $\alpha$ . It is important to point out that, although the differential methods do not use any approximation, they are very sensitive to experimental noises.

$$\ln\left(\frac{d\alpha}{dt}\right) = \ln\left(\beta \frac{d\alpha}{dT}\right) = \ln[Af(\alpha)] - \frac{E}{\mathfrak{R}T} \quad (6)$$

On the other hand, the variables in Eq. (5) can be separated and on integration gives Eq. (7)

$$g(\alpha) \equiv \int_0^\alpha \frac{d\alpha}{f(\alpha)} = \frac{A}{\beta} \int_0^T \exp\left(-\frac{E}{\mathfrak{R}T}\right) dT \quad (7)$$

where  $g(\alpha)$  is the integral form of the reaction model. The integral in Eq. (7) does not have an analytical solution and several approximations can be found in the literature [33,37]. Generally, the approximate solutions are represented by the general equation:

$$\ln\left(\frac{\beta}{T^B}\right) = \text{const} - C\left(\frac{E}{\mathfrak{R}T}\right) \quad (8)$$

where  $B$  and  $C$  are the parameters determined by the type of temperature integral approximation. Worth noting that the approximation used is directly related to the accuracy of the integral method.

Three different approaches, integral isoconversional methods, were considered in the present work: Starink, Eq. (9) [22]; Kissinger Akahira Sunose, Eq. (10) [20,21]; and Ozawa Flynn Wall, Eq. (11) [16–19].

$$\ln\left(\frac{\beta}{T^{1.92}}\right) = \text{const} - 1.0008\left(\frac{E}{\mathfrak{R}T}\right) \quad (9)$$

$$\ln\left(\frac{\beta}{T^2}\right) = \ln\left(\frac{A\mathfrak{R}}{Eg(\alpha)}\right) - \frac{E}{\mathfrak{R}T} \quad (10)$$

$$\ln(\beta) = \ln\left(\frac{AE}{\mathfrak{R}g(\alpha)}\right) - 5.331 - 1.052\frac{E}{\mathfrak{R}T} \quad (11)$$

The left hand side of each method was plotted against  $1/T$  for each experimental TGA data and the kinetic parameters were calculated from the slope of the straight lines fitted at each common value of  $\alpha$ .

## 2.5. Reaction model

The dependence of the decomposition process on the fraction reacted can be expressed by using a wide variety of reaction models, each of these with a characteristic reaction profile [37]. The Criado method [28,33,48,49] was used to determine the kinetic model able to describe the pyrolysis of each material investigated. The method is based on an approximate solution of the integral in Eq. (7) and leads to the construction of a generalized master plot. The relationship between the experimental data recorded under non isothermal conditions, the generalized reaction rate, and the differential and integral forms of

**Table 2**

Algebraic expressions of the most widely used theoretical solid-state reaction models.

Model	Code	$f(\alpha)$	$g(\alpha)$
Power-law	$P_n^a$	$n\alpha^{(n-1)/n}$	$\alpha^{(1/n)}$
Zero-order	$F0$	1	$\alpha$
First-order	$F1$	$1 - \alpha$	$-\ln(1 - \alpha)$
Second-order	$F2$	$(1 - \alpha)^2$	$[1/(1 - \alpha)] - 1$
Third-order	$F3$	$(1 - \alpha)^3$	$[1/(1 - \alpha)^2] - 1$
Avrami–Erofeev	$A_n^b$	$n(1 - \alpha)[- \ln(1 - \alpha)]^{1-1/n}$	$[- \ln(1 - \alpha)]^{1/n}$
Contracting geometry	$R_n^c$	$n(1 - \alpha)^{1-1/n}$	$1 - (1 - \alpha)^{1/n}$
One-dimensional diffusion	$D1$	$(1/2)\alpha^{-1}$	$\alpha^2$
Two-dimensional diffusion	$D2$	$[- \ln(1 - \alpha)]^{-1}$	$(1 - \alpha)\ln(1 - \alpha) + \alpha$
Three-dimensional diffusion	$D3$	$(3/2)(1 - \alpha)^{2/3}[1 - (1 - \alpha)^{1/3}]^{-1}$	$[1 - (1 - \alpha)^{1/3}]^2$
Ginstling–Brounshtein	$D4$	$(3/2)[(1 - \alpha)^{-1/3} - 1]^{-1}$	$1 - (2\alpha/3) - (1 - \alpha)^{2/3}$
Prout–Tompkins	$B1$	$\alpha(1 - \alpha)$	$\ln[\alpha/(1 - \alpha)]$
Exponential-law	$E1$	$\alpha$	$\ln(\alpha)$

<sup>a</sup> $n = 2/3, 2, 3, 4$ .

<sup>b</sup> $n = 2, 3, 4$ .

<sup>c</sup> $n = 2, 3$ .

the theoretical solid state reaction models, Table 2 [33,37,50], can be described by Eq. (12).

$$\left(\frac{T}{T_{0.5}}\right)^2 \frac{(d\alpha/dt)}{(d\alpha/dt)_{0.5}} = \frac{f(\alpha)g(\alpha)}{f(0.5)g(0.5)} \quad (12)$$

The equation above is normalized taking  $\alpha = 0.5$  as a reference, being  $T_{0.5}$  and  $(d\alpha/dt)_{0.5}$ , at the left hand side, and  $f(0.5)$  and  $g(0.5)$ , at the right hand side, the temperature, the reaction rate and the conversion functions corresponding to 50% conversion.

The most suitable kinetic model is determined by the best match between the plot of the generalized reaction rate vs. the fraction reacted and the plots of the theoretical models. The Root Mean Square Error (RMSE), Eq. (13), was used to measure how much error there is between the curves, assisting identify the best correspondence.

$$RMSE = \sqrt{\frac{\sum_{i=1}^n [(d\alpha/dt)_{exp} - (d\alpha/dt)_{model}]^2}{n}} \quad (13)$$

## 3. Results and discussion

### 3.1. Thermogravimetric analysis (TGA)

TGA and DTG results under inert atmosphere are summarized in Fig. 4. In this figure, each column contains the respective decomposition profile of VPE, WVPE, and WRPE at four different heating rates.

The heating rates were strategically chosen to cover the influence of a wide range of temperatures on PE degradation. The samples of virgin and recycled PE when subjected to pyrolysis, at all heating rates, lose almost all the mass in a single step. The heating rate increase resulted in a lateral shift of the curves along the abscissa without changing its shape. The initial mass, onset temperature, offset temperature, peak temperature at maximum degradation rate, and the total mass losses after pyrolysis are presented in Table 3.

### 3.2. X ray diffraction analysis

Fig. 5 shows the X ray powder diffraction patterns for all samples. The compounds quantification was performed based on the volumetric proportion of the structures found in a sample volume of 1 mm x 10 mm x 0.005 mm. The mass proportion of the phases depends on the determination of the true density (crystallographic) of the phases

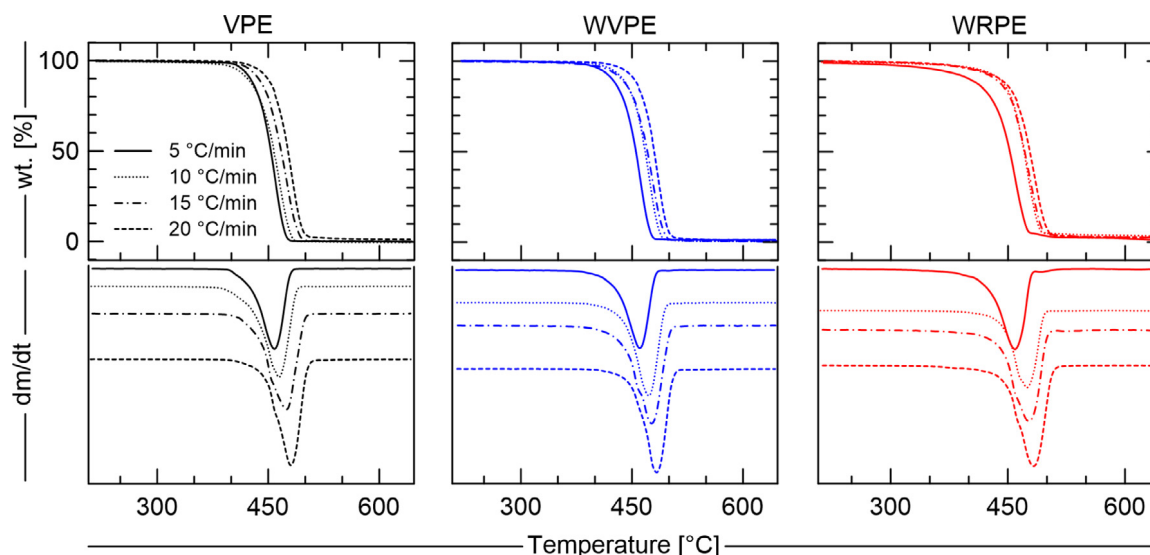


Fig. 4. Polyethylene thermogravimetric analysis under inert atmosphere and its derivatives.

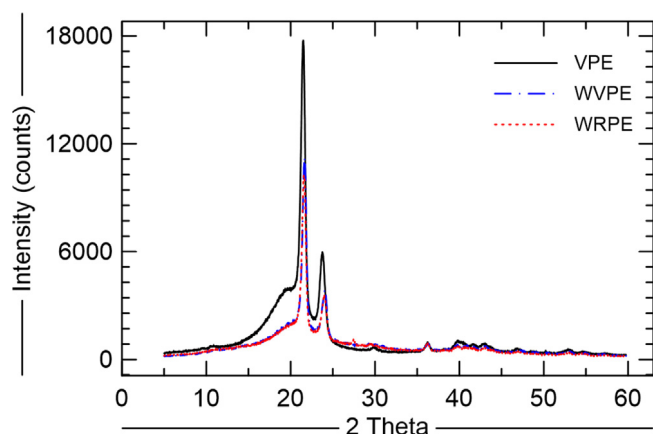


Fig. 5. X-ray powder diffraction pattern.

Table 3

Polyethylene thermogravimetric analysis: initial mass, onset, offset, and peak temperatures and total mass losses.

Sample	$\beta$ [°C/min]	$m_i$ [mg]	$T_{on}$ [°C]	$T_{off}$ [°C]	$T_{peak}$ [°C]	$\Delta_m$ [wt.%]
VPE	5	13.3	392	480	458	99.9
	10	12.3	387	488	465	99.9
	15	13.2	396	500	475	99.9
	20	12.7	405	504	481	99.9
WVPE	5	14.6	376	484	461	98.8
	10	10.9	384	496	473	99.7
	15	14.0	388	502	477	99.9
	20	14.1	394	506	483	99.8
WRPE	5	13.4	357	480	459	98.9
	10	13.1	367	497	475	96.8
	15	13.4	367	502	477	98.0
	20	12.9	367	507	483	97.7

present. The uncertainty volume value proposed for each phase found, as considered the volume measured by XRD, is limited to 2% taken into account the copper radiation (CuK $\alpha$ , 0.15419 nm), the Bragg Brentano geometry, 0.5° divergent slit, 0.5° scattering slit and LiF monochromator.

It was possible to identify and quantify the chemical compounds of all samples without considering the amorphous phase. The average crystallinity was 48%, 51%, and 54% for VPE, WVPE, and WRPE

Table 4

All samples content available by X-ray diffraction pattern analysis, values in vol.%.

Compound name	VPE	WVPE	WRPE
n-Paraffin (CH <sub>2</sub> ) <sub>x</sub>	10.1	–	–
Paraffin wax (C - H)	4.9	–	–
n-Heptadecane (C <sub>17</sub> H <sub>36</sub> )	–	2.3	–
n-Tricosane (C <sub>23</sub> H <sub>48</sub> )	–	–	4.8
n-Nonacosane (C <sub>29</sub> H <sub>60</sub> )	4.2	15.8	16.4
Pentatriacontane (C <sub>35</sub> H <sub>72</sub> )	80.8	–	–
n-Tetracontane (C <sub>40</sub> H <sub>82</sub> )	–	81.9	78.8

Uncertainty volume value is limited to 2%.

respectively. The analysis indicated that the main chemical compounds of crystalline phase of VPE were pentatriacontane (C<sub>35</sub>H<sub>72</sub>) and n paraffin (CH<sub>2</sub>)<sub>x</sub>, while the main compounds of the recycled samples were n tetracontane (C<sub>40</sub>H<sub>82</sub>) and n nonacosane (C<sub>29</sub>H<sub>60</sub>) (Table 4). The data collection details can be shown as supplementary material, Appendix A.

As one can observe in Table 4, the results suggest that the recycling stress history promotes the increase of long carbon chains in WVPE and WRPE compared to VPE. The cause might be the combination of thermal and extrusion stress experienced by the PE in the cycles of recycling.

### 3.3. Ultimate and proximate analyses

The results indicate mainly the presence of carbon and hydrogen for all samples. Small percentages of other elements, apart from CHNS, are present in the recycled samples: 0.3 and 2.0 wt% in WVPE and WRPE, respectively. The samples were found to be non hygroscopic materials and only a small amount of moisture, around 0.2 wt%, was found in WVPE and WRPE. When compared to VPE, the WRPE samples presented an approximately threefold relative increase in ash content, from 0.3 to 1.0 wt%. The results of ultimate and proximate analyses are listed in Table 5.

### 3.4. Differential Scanning Calorimetry (DSC)

Analyzing the DSC curves under pyrolysis conditions, Fig. 6, two endothermic peaks are observed for all samples. The first one is assigned to the melting point and confirmed by checking TGA under nitrogen where no mass loss is identified in the temperature range of the first

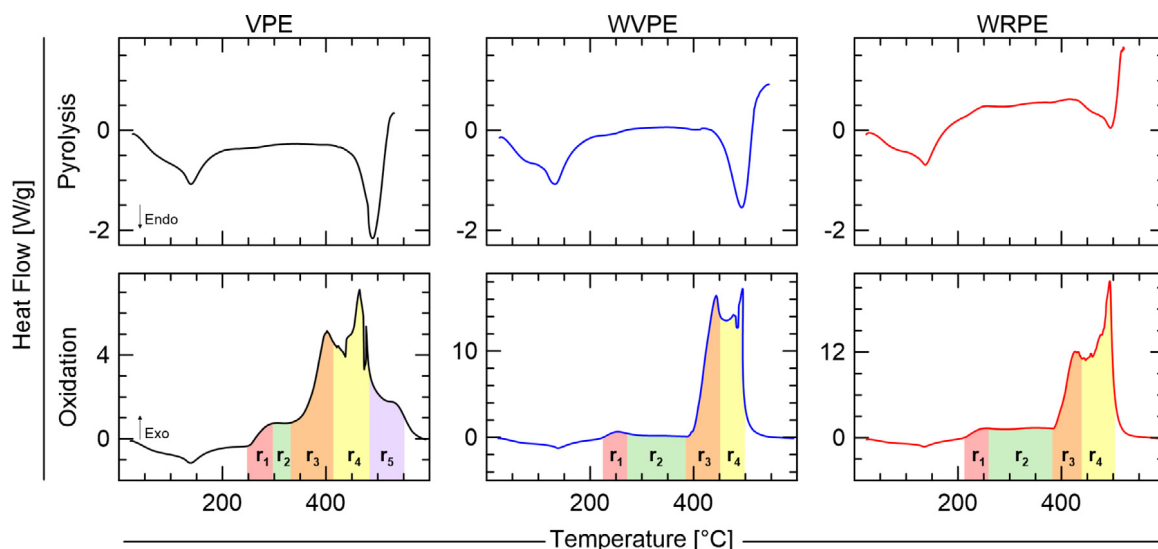


Fig. 6. Polyethylene differential scanning calorimetry under inert and oxidative atmospheres at 20 °C/min, the heat flows.

Table 5

Polyethylene samples ultimate and proximate analyses, values in wt.%.<sup>a</sup>

Component	VPE	WVPE	WRPE
C	85.7 ± 0.8	85.4 ± 0.1	84.3 ± 0.1
H	14.3 ± 0.2	14.3 ± 0.3	13.7 ± 0.3
N	0.0	0.0	0.0
S	0.0	0.0	0.0
Other elements <sup>a</sup>	0.0	0.3 ± 0.3	2.0 ± 0.3
Moisture	0.0	0.2	0.2
Ash	0.3	0.4	1.0
Volatile matter <sup>a</sup>	99.7	99.4	98.8

<sup>a</sup>Obtained by difference/Proximate analysis (mean ± 0.1).

endothermic peak. The second peak is attributed to PE degradation reactions, which are associated with the complete mass loss on TGA.

The integration of the DSC curves can be used to estimate the magnitude order of the enthalpies of melting and pyrolysis. The values of enthalpies of melting were 75, 74, and 65 J/g for VPE, WVPE, and WRPE respectively. The values of 299, 332, and 205 J/g for enthalpies of pyrolysis of VPE, WVPE, and WRPE were respectively obtained. Since the majority of the work does not categorize their type of residues, the values of enthalpies reported in [35,51] for melting at about 90 J/g, and in [52] for pyrolysis at about 258 J/g provide a rough indication of the expected order of magnitude. Looking at the effect of the physical stress on these categories of polyethylene waste, the results suggest that WRPE suffered some modification in its structure leading to weakening of the material, see Section 3.2.

To complement the result analysis of the inert atmosphere, DSC under air was carried out. Through the oxidation of the PE samples, it can be observed one endothermic peak related to the one found under an inert atmosphere, and various exothermic peaks indicating that competing reactions are taking place. These reactions with some overlapping degree can be attributed to the key compounds forming the samples, see Table 4. Therefore, by simplifying the analysis, one can attribute mainly four oxidation reactions for VPE ( $r_1, r_3, r_4$ , and  $r_5$ ) and three for WVPE and WRPE ( $r_1, r_3$ , and  $r_4$ ) according to the number of the compounds of each sample. The reaction  $r_2$  is attributed to endothermic phenomena such as phase change and pyrolysis reactions. As the degree of recycling increases (WVPE and WRPE), the endothermic phenomena  $r_2$  become stronger extending through a wider range of temperatures. The TGA under the oxidative atmosphere can be visualized as supplementary material, Appendix B.

By integrating the DSC curves under air for VPE, WVPE, and WRPE, in the temperature range of reactions ( $r$ ), the combustion enthalpies of the respective samples were about 2765, 3237, and 3665 J/g.

### 3.5. Kinetic analysis

The plots obtained by the left hand side of the equation of each isoconversional method considered against  $1/T$  are shown in Fig. 7 for VPE, WVPE, and WRPE. The corresponding lines are obtained at different fraction reacted and different heating rates. It can be seen from Fig. 7 that two groups of straight lines are formed depending on the degree of fraction reacted and method: one with smaller slope and lower fraction reacted degree ( $\alpha < 0.25$ ) and the other steeper lines and  $0.25 < \alpha < 0.95$ . A different tendency was observed when Friedman's isoconversional method was used.

For comparison, the values of the activation energy were determined in a wide range of  $\alpha = 0.05 - 0.95$  with a step of 0.05 and the dependence was reported in a plot  $E$  vs.  $\alpha$ , Fig. 8. This dependence is a warning about multi step kinetics existence in a given process [33]. As can be seen from the curves presented in Fig. 8, from 0.05 to 0.95 of fraction reacted, the average activation energy ( $E_{avg}$ ), varies in a range of  $216.0 \pm 46.3$  kJ/mol,  $225.5 \pm 42.9$  kJ/mol, and  $185.5 \pm 61.6$  kJ/mol for VPE, WVPE, and WRPE, respectively. In percentage terms, the difference between the maximum ( $E_{max}$ ) and minimum ( $E_{min}$ ) values of the average are respectively 42.8%, 38.0%, and 66.4%.

Although the literature shows that a strong dependence of  $E$  within a wide range of  $\alpha$  means that a given process cannot be considered as a single step one, just one work pointed out the acceptable degree of this dependence, [53]. That author considers that the ratio  $(E_{max} - E_{min})/E_{avg}$  should be less than 10%, which constitutes noteworthy impasse because many authors have been using a single step reaction to describe multi step ones with a relative degree of success [23-28] in rebuild thermogravimetric curves. On the other hand, if the preliminary results from isoconversional methods are used as a tool of diagnostic toward detailed kinetics related to the necessity of controlling the products yields at different reaction velocities, a multi step treatment should be performed.

The generalized master plots are shown in Fig. 9. At the top part of the figure, each column contains the respective master plot of VPE, WVPE, and WRPE with all theoretical models investigated. At the bottom part, only the most suitable models are presented. The RMSE between the experimental data and the theoretical reaction models, used to identify the best correspondence, are shown in Fig. 10.



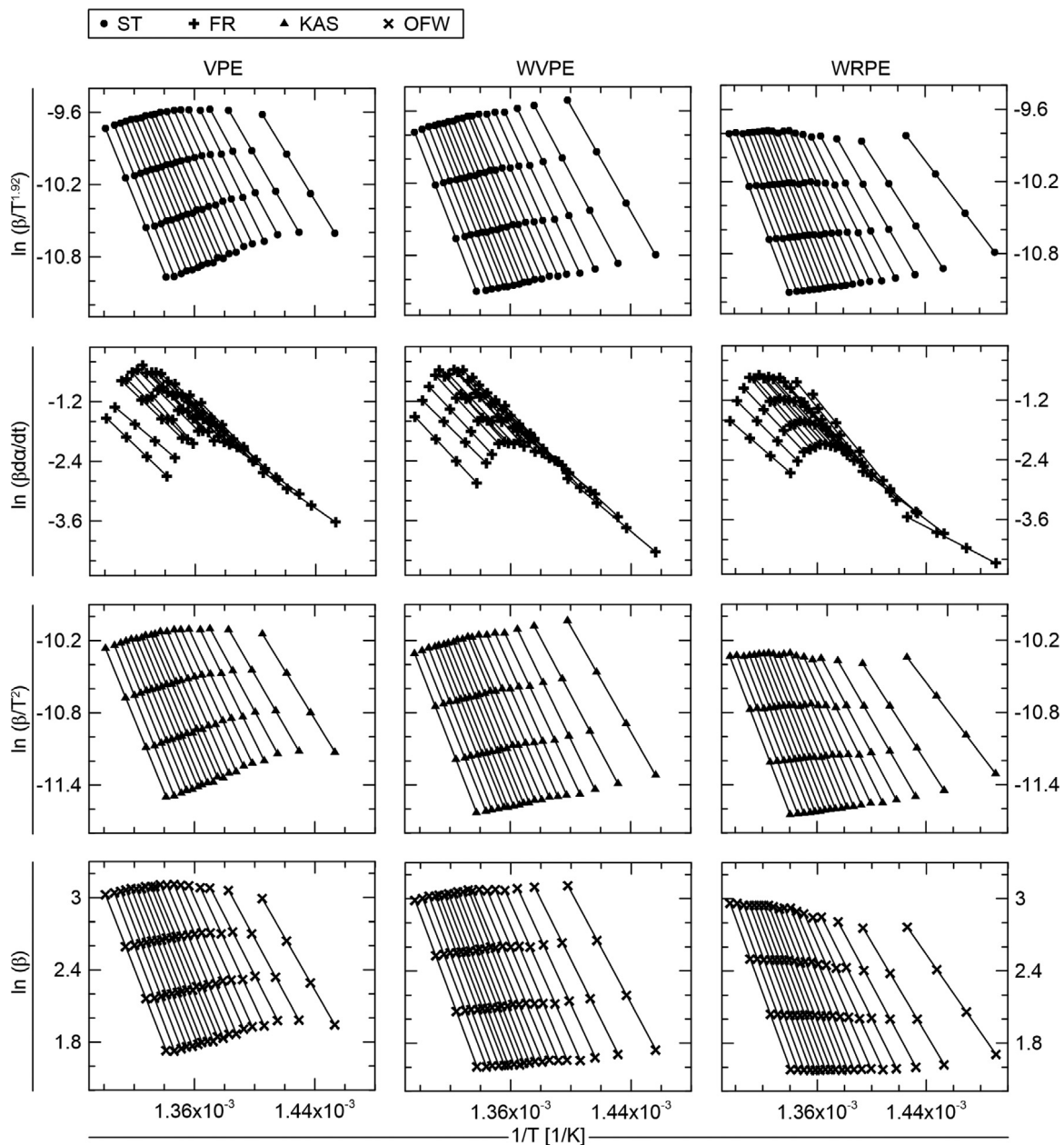


Fig. 7. Isoconversional plot of Starink, Friedman, Kissinger–Akahira–Sunose and Ozawa–Flynn–Wall methods in a range of  $\alpha = 0.5 - 0.95$  - from the TG under inert atmosphere data.

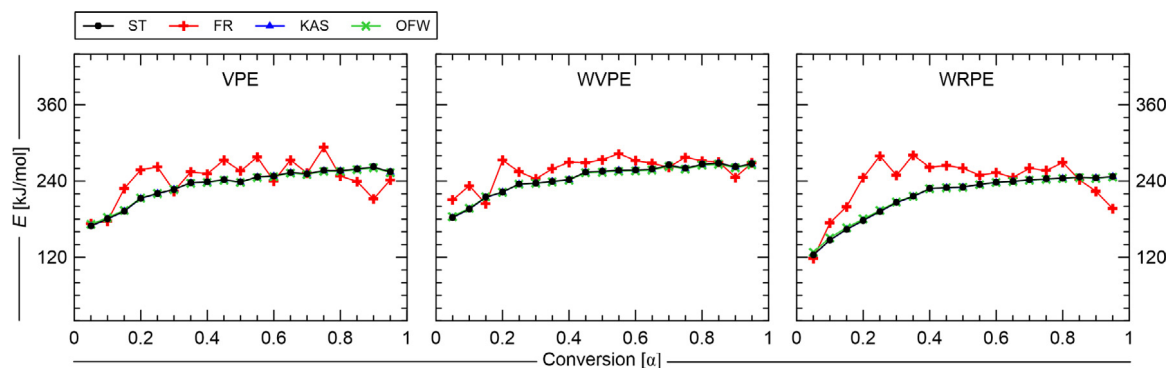


Fig. 8. Distribution of activation energy from isoconversional methods in a range of  $\alpha = 0.5 - 0.95$ .

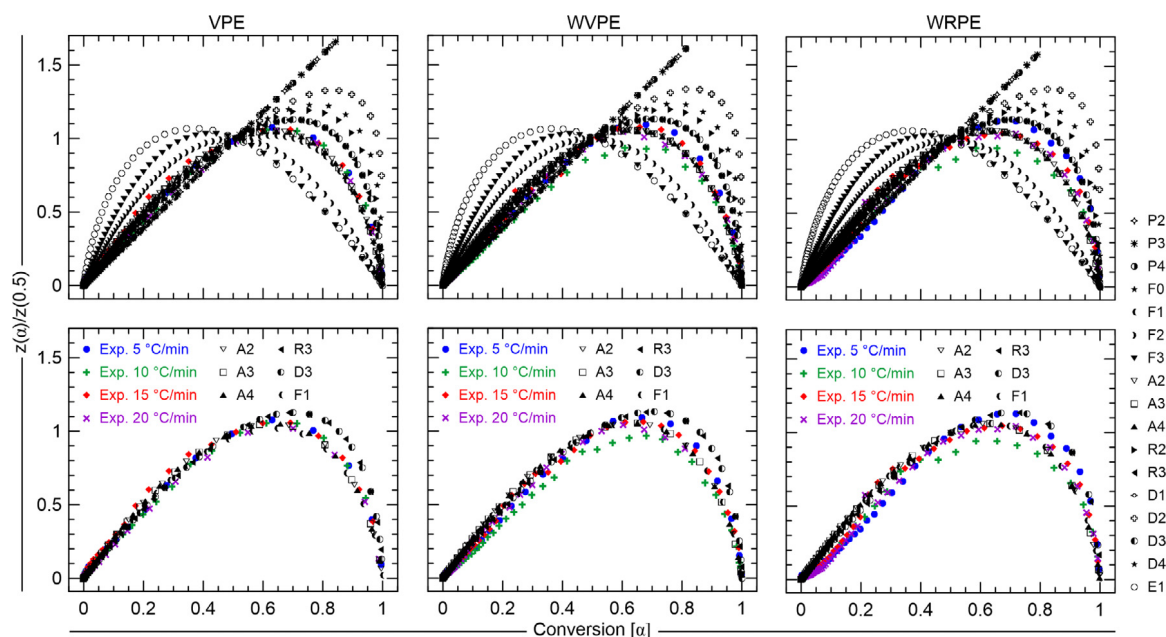


Fig. 9. At the top: theoretical masterplots of different reaction models; bottom: the best master plots vs. experimental data calculated by Eq. (12).

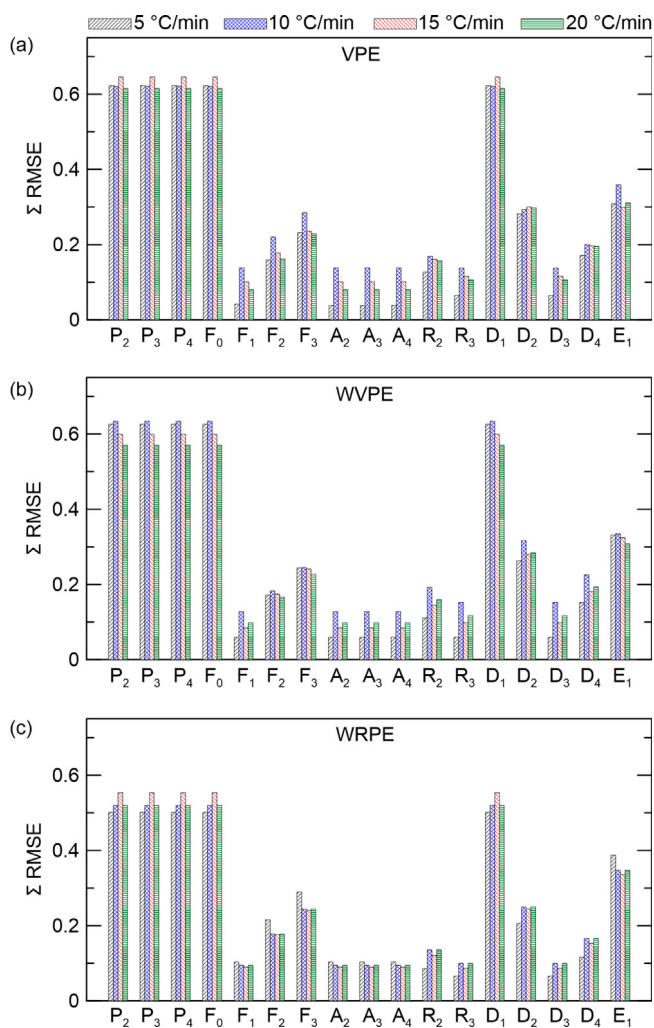


Fig. 10. Sum of the RMSE between the experimental TG curve and each reaction model listed in Table 2. (a) for VPE, (b) for WVPE, and (c) for WRPE.

The master plots results suggest that the thermal degradation of VPE could be governed by Avrami Erofeev models ( $A_2$ ,  $A_3$ , and  $A_4$ ), as well as, by a first order model ( $F_1$ ). The same trend was obtained for the recycled samples, which, in addition, can also be approximated by the contracting geometry ( $R_3$ ) and three dimensional diffusion ( $D_3$ ) models. These model equivalences were already expected since according to [48], it is not possible to distinguish between the models  $R_3$  and  $D_3$ , and also is impossible to discern the differences between a first order and Avrami Erofeev models.

It is noteworthy that, although a number of experimental curves at different heating rates result in a series of master plots, a single dependence on  $\alpha$  can be observed, which is practically independent of  $\beta$ , see in Fig. 9.

The values of the pre exponential factor were determined using the reaction model predicted by Criado master plots and Eq. (4). The range of calculated values of  $E$  and  $A$  are summarized in Tables 6 and 7, organized by sample and by the method. The dependence of the pre exponential factor on the conversion is reported in Fig. 11. In general, the dependence of  $\ln A$  on  $\alpha$  increase as more recycled is the sample, which follows the same trend of  $E(\alpha)$ .

The kinetic parameters obtained from the different isoconversional methods were used to reconstruct the conversion profiles in comparison with the experimental data, Fig. 12, and then the RMSE was calculated to check the accuracy of each method, Fig. 13. From this analysis, the Starink method showed to be the most accurate approach to determine the kinetic parameters of the pyrolysis process of polyethylene virgin and wastes. On the other hand, the reconstruction of the conversion profiles using KAS and OFW methods did not match well the experiments, as already reported by [27].

For the values obtained from the Starink method, the logarithmic form of the pre exponential factor was plotted against  $E$ . A linear relationship is observed between them, Fig. 14. The correlation coefficient, close to 1, indicates that for all heating rates the values of  $A$  are in accordance with the distributed activation energy along the conversion process.

### 3.6. Pyrolysis pathway

According to the results presented above, the significant variation of the kinetics values indicated that polyethylene pyrolysis is indeed

**Table 6**

Range of activation energy,  $E$  [kJ/mol], for the pyrolysis of VPE, WVPE, and WRPE obtained from isoconversional methods at various heating rates.

Sample	$\beta$ [°C/min]	ST	FR	KAS	OFW
VPE	5	169.7 – 262.2	173.1 – 293.0	169.3 – 261.9	172.0 – 260.9
	10				
	15				
	20				
WVPE	5	182.6 – 268.3	204.1 – 283.1	182.3 – 268.1	184.3 – 266.7
	10				
	15				
	20				
WRPE	5	123.9 – 247.1	118.4 – 280.9	123.6 – 246.7	128.3 – 246.5
	10				
	15				
	20				

**Table 7**

Range of pre-exponential factor,  $A$  [1/min], for the pyrolysis of VPE, WVPE, and WRPE obtained from isoconversional methods at various heating rates.

Sample	$\beta$ [°C/min]	ST	FR	KAS	OFW
VPE	5	$2.51 \times 10^{12} - 1.15 \times 10^{18}$	$7.65 \times 10^{12} - 2.35 \times 10^{20}$	$5.52 \times 10^{10} - 9.86 \times 10^{17}$	$9.79 \times 10^{10} - 8.66 \times 10^{17}$
	10				
	15				
	20				
WVPE	5	$2.16 \times 10^{13} - 2.98 \times 10^{18}$	$4.30 \times 10^{14} - 5.93 \times 10^{19}$	$5.03 \times 10^{11} - 2.06 \times 10^{18}$	$7.76 \times 10^{11} - 1.73 \times 10^{18}$
	10				
	15				
	20				
WRPE	5	$9.61 \times 10^8 - 9.57 \times 10^{16}$	$3.90 \times 10^8 - 9.87 \times 10^{19}$	$1.98 \times 10^7 - 7.32 \times 10^{16}$	$6.16 \times 10^7 - 7.30 \times 10^{16}$
	10				
	15				
	20				

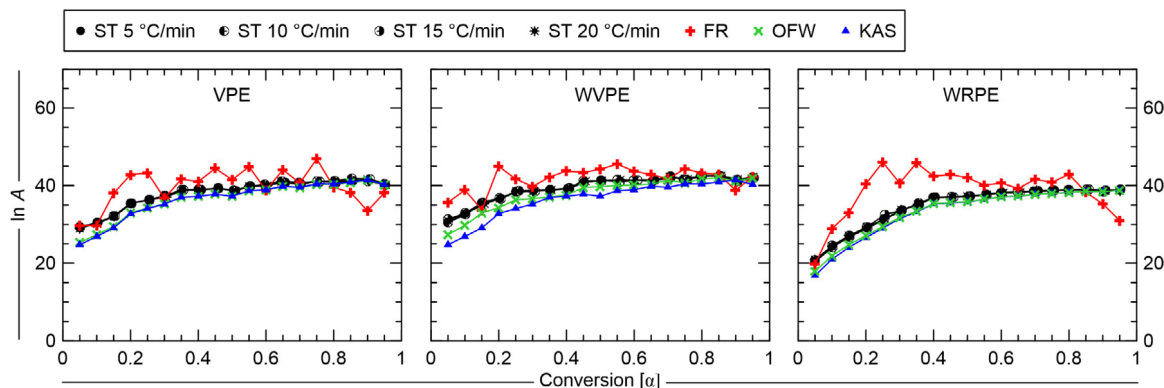


Fig. 11. Distribution of pre-exponential factor from isoconversional methods in a range of  $\alpha = 0.5 - 0.95$ .

kinetically complex. The evidence was also observed using the characterization results, where the XRD analysis shows that PE is composed of bonded compounds, and a strong degree of overlapping was identified in DSC analysis. Therefore, a multi step pathway is delineated below:

- Pre heating. The stage in which the sample absorbs heat. At this stage, only increases in the sample temperature take place.
- Melting. The heat promotes changes from the solid phase,  $PE_{(s)}$ , to the liquid phase,  $PE_{(l)}$ , in an endothermic process.
- Bond breaking. After the melting, the bonds linking the compounds forming the PE are thermally broken in an endothermic process. The result is a multi compound liquid mixture in which compounds with diverse molecular properties are present [54]. From this stage onwards, the liquid phase pyrolysis of the main compounds can take place.

- Evaporation. Once the boiling point of each compound is reached, the liquid phase starts to be vaporized.
- Pyrolysis reaction. Once a compound starts evaporation, it can be thermally cracked in series with respect to the evaporation, and in parallel with respect to the other compounds pyrolysis. The pyrolysis products formed in the individual reactions are grouped in  $P_p$ .

Note that, the phase change  $PE_{(s)}$  to  $PE_{(l)}$  does not account for mass losses; negligible mass losses can be attributed to the bond breaking [55]; the multi compound mixture in the liquid phase is formed only by the compounds identified by XRD. Even if the compounds n paraffin<sub>(g)</sub> and paraffin wax<sub>(g)</sub> are mixtures of hydrocarbons, they are modeled by their main constituents [56,57]. The amorphous

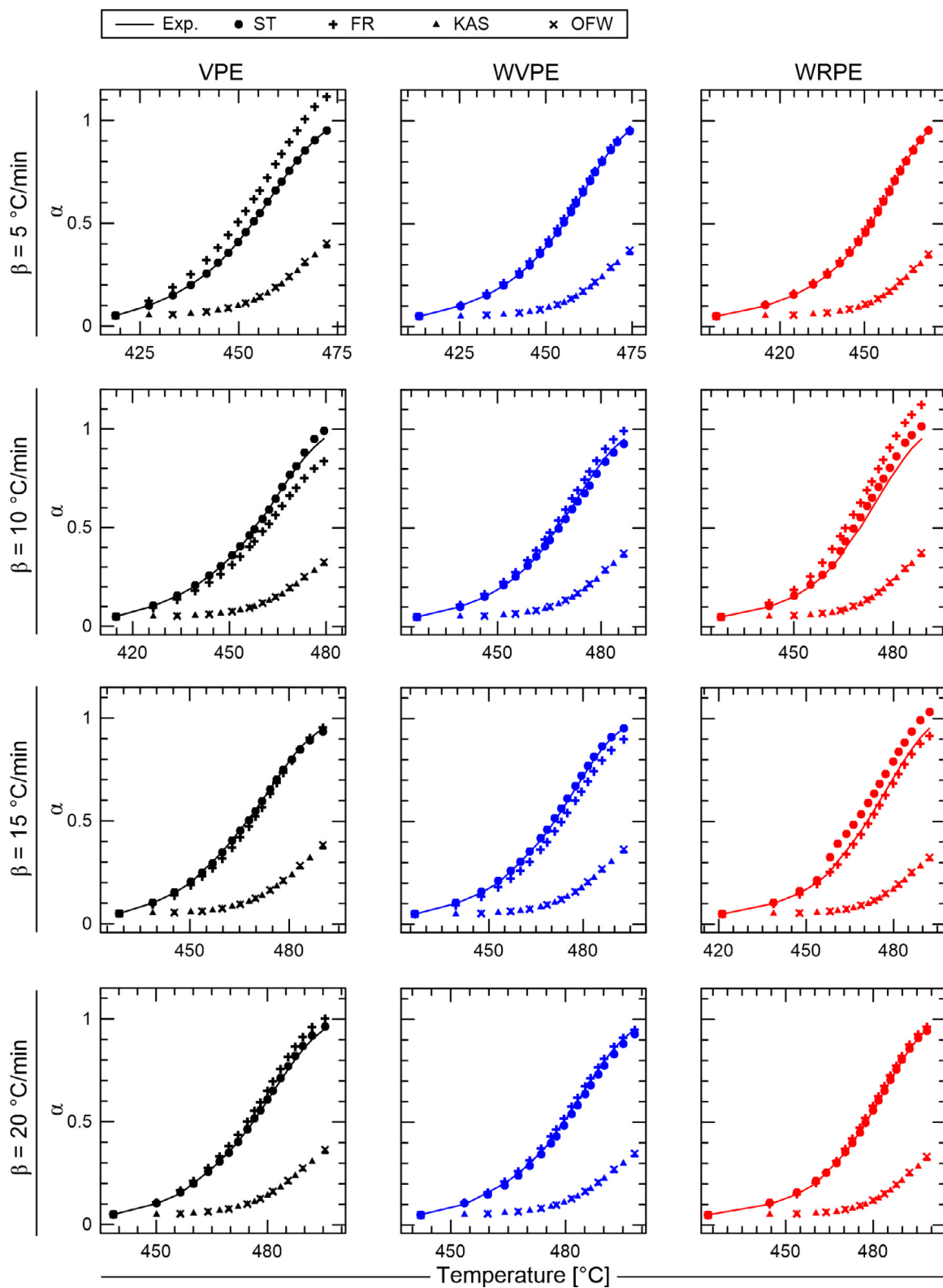


Fig. 12. Fitness of the thermogravimetric analysis (fraction reacted,  $\alpha$ ) under inert atmosphere using the kinetic parameters obtained from isoconversional methods at various heating rates.

phases can be considered fractions of irregular groups remaining at attached in each separated rigid structure (crystalline phase) identified [58]. Also, the pyrolysis products may include oil, wax, and gas [25,29,32,59], and their yields are not investigated in the present work.

The pyrolysis pathways for VPE, WVPE, and WRPE are shown in the schemes in Table 8.

To implement the reaction mechanism shown in Table 8 in numerical modeling, the stoichiometric and kinetic parameters of each reaction must be quantified, as well as, the amorphous phase may be



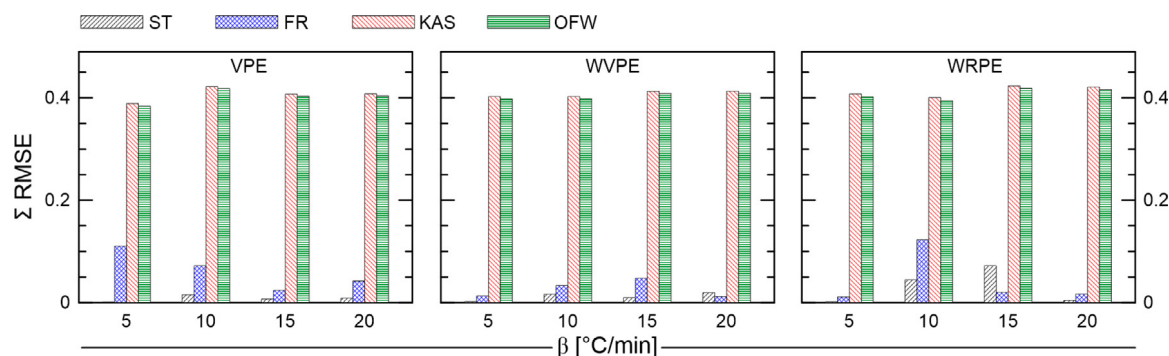


Fig. 13. Comparison of different isoconversional methods based on the sum of the RMSE between the experimental and reconstructed fraction reacted ( $\alpha$ ).

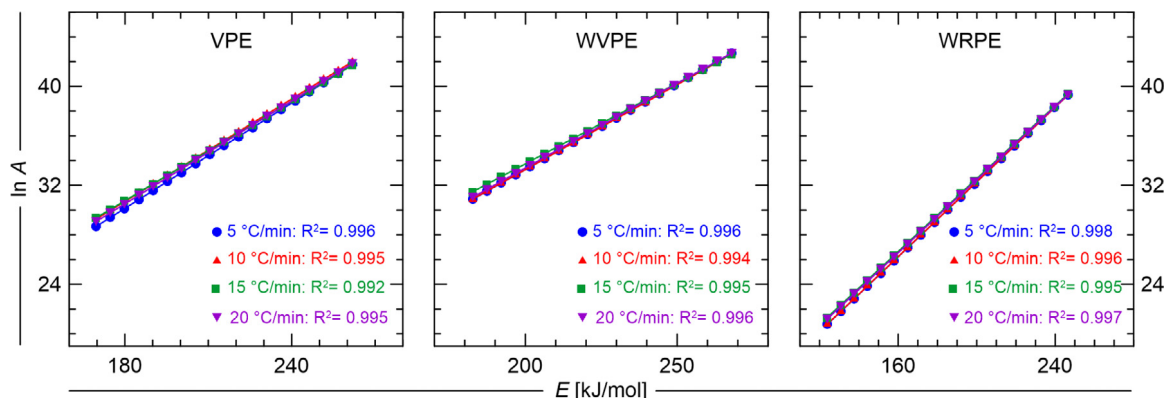


Fig. 14. Estimated  $\ln A$  versus  $E$  from the Starink method.

incorporated in the solid phase definition. It is necessary to reformulate the kinetic parameter estimation process as an optimization problem in terms of an objective function [32,60,61]. This procedure is beyond the scope of the present study, and for this reason, the kinetic parameters are not determined.

#### 4. Conclusion

In this paper, two types of PE wastes with different degrees of recycling were used to investigate the influence of recycling stress history on the pyrolysis. The main results suggest that the thermal and mechanical stress promotes the increase of long carbon chains while weakening the boundaries among the compounds. This might explain the fact that recycled waste needs less activation energy than recycled samples to degrade thermally. The waste of recycled polyethylene presented the lowest enthalpy of pyrolysis, at about 205 J/g.

The DSC trials evidenced a multi step reaction behavior. This behavior was confirmed by using different isoconversional methods for the kinetic study. Large variations around the average activation energy values were observed. The waste of recycled polyethylene presented a higher variation of about 66% in the range of  $185.5 \pm 61.6$  kJ/mol. Also, it needs less activation energy than VPE and WVPE to degrade thermally. The degradation of all samples can follow Avrami Erofeev ( $A_2$ ,  $A_3$ , and  $A_4$ ) and first order ( $F_1$ ) models, also, the recycled can be described by contracting geometry ( $R_3$ ) and three dimensional diffusion ( $D_3$ ) models.

The pyrolysis pathway is a series parallel phenomenon dominated mainly by the competition of evaporation and pyrolysis reactions.

We recommend that the physical stress history between different categories of polyethylene waste should be considered for kinetics purposes, as well as, for the modeling of the product recovery process.

#### Data availability

The raw/processed data required to reproduce these findings cannot be shared at this time as the data also forms part of an ongoing study.

#### Declaration of competing interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

#### CRediT authorship contribution statement

**João Vitor F. Duque:** Conceptualization, Methodology, Software, Validation, Formal analysis, Resources, Data curation, Writing original draft, Visualization. **Márcio F. Martins:** Conceptualization, Methodology, Data curation, Writing review & editing, Supervision, Project administration, Funding acquisition. **Gérald Debenest:** Conceptualization, Software, Methodology, Data curation, Writing review & editing, Supervision. **Marcos Tadeu D'Azeredo Orlando:** Software, Methodology, Data curation.

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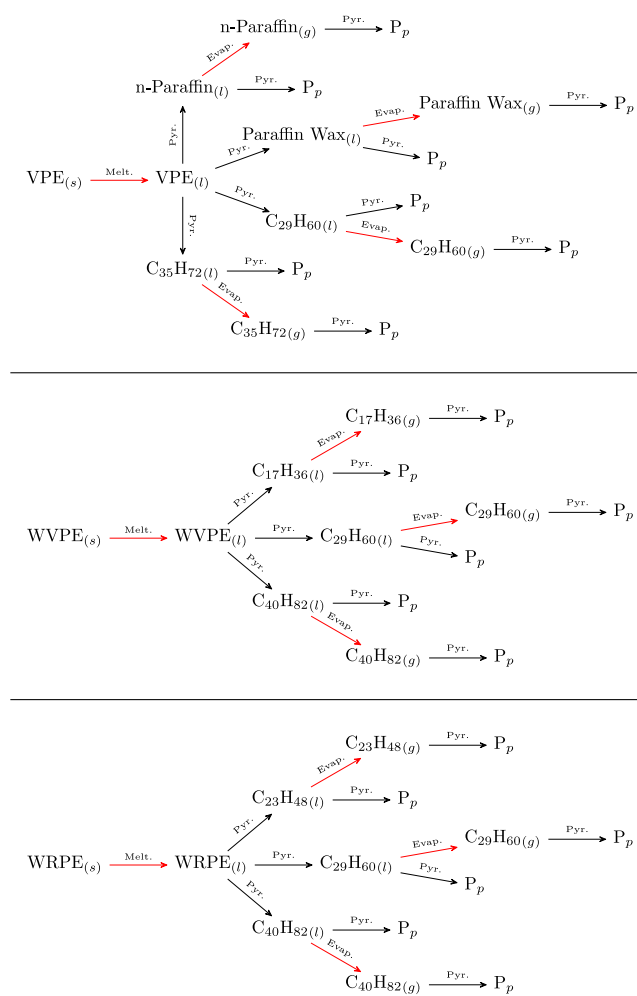
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#### Appendix A. Supplementary data

Supplementary material related to this article can be found online at <https://doi.org/10.1016/j.polymertesting.2020.106460>.



**Table 8**  
Schemes of the pyrolysis pathways for VPE, WVPE and WRPE.



Melt. and Evap. means melting and evaporation phase changes/  
Pyr. means pyrolysis reactions/P<sub>p</sub> may include oil, wax, and gas.

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