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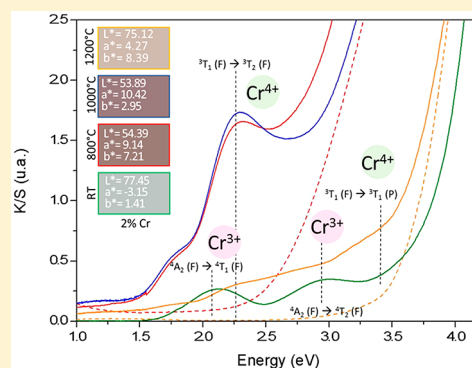
# Tuning the Cr<sup>IV</sup>/Cr<sup>III</sup> Valence States in Purple Cr-Doped SnO<sub>2</sub> Nanopowders: The Key Role of Cr<sup>IV</sup> Centers and Defects

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**ABSTRACT:** A low content of chromium ( $\leq 5$  mol %) has been incorporated into a SnO<sub>2</sub> cassiterite by a coprecipitation route in a basic medium, followed by an annealing step under an O<sub>2</sub> flow at  $T = 800$  °C and  $T = 1000$  °C. Accurate UV-vis and EPR spectroscopy investigations show the coexistence of isolated Cr<sup>4+</sup> and Cr<sup>3+</sup> ions as well as ferromagnetic Cr<sup>4+</sup>–Cr<sup>3+</sup> and antiferromagnetic Cr<sup>3+</sup>–Cr<sup>3+</sup> interactions. The strong purple hue is related to the isolated Cr<sup>4+</sup> ions stabilized in a distorted octahedral site. This is thanks to the second-order Jahn–Teller (SOJT) effect with a crystal field splitting  $10Dq$  value around 2.4 eV, whereas the  $10Dq$  value is around 2 eV for isotropic Cr<sup>3+</sup> ions, partially substituted for Sn<sup>4+</sup> ions in cassiterite. Just after the coprecipitation process, only Cr<sup>3+</sup> species are stabilized in this rutile network with a poor crystallinity. The isolated Cr<sup>4+</sup> content remains high after annealing at 800 °C for 2 days especially for the highest Cr rate (2 and 5 mol %), leading to a darker purple color, but unfortunately the Cr<sup>3+</sup> content also increases for a higher Cr concentration. A lighter purple hue can be reached after calcination at a higher temperature ( $T = 1000$  °C) for a shorter time (4 h) but with a lower Cr content to avoid Cr clusters. This is due to stabilizing a high content of isolated Cr<sup>4+</sup> species and limiting the Cr<sup>4+</sup>–Cr<sup>3+</sup> ferromagnetic interactions, which are optimal for a 2% Cr content and also cause the color to darken. The key roles of the Cr<sup>4+</sup> rate and the Cr<sup>4+</sup>–Cr<sup>3+</sup> clusters create local defects whose concentration strongly varies with a total Cr content, which have then been demonstrated to strongly influence the optical and magnetic properties.



## 1. INTRODUCTION

The control of the optical band gap (BG) and the nature and the rate of defects in a transparent oxide semiconductor and transition-metal-doped semiconducting materials are of high interest in optoelectronics, spintronic devices, sensors, transistors, electrode materials, and solar cells.<sup>1</sup> Titanium (TiO<sub>2</sub>), zinc (ZnO), indium (In<sub>2</sub>O<sub>3</sub>), and tin (SnO<sub>2</sub>) binary oxides as n-type oxide semiconductors with a wide band gap have been largely explored as suitable candidates for this vast field of applications. Among these four oxides, stannic oxide SnO<sub>2</sub> exhibits the highest direct band gap around 3.6 eV, with a refractive index around  $n = 2$  in the visible range.<sup>2</sup> The synthesis route followed by annealing at high temperatures and the average crystallite size of tin dioxide that adopts a rutile-type structure strongly affects the optical band gap.<sup>1,2</sup> The occurrence of defects within the crystals, including interstitial tin atoms, oxygen vacancies, or hydroxyl groups, contributes to create various impurity energy levels. Stannic oxides doped with low rates ( $<5$  at. %) of 3d transition metals such as Cr<sup>3+</sup>(3d<sup>3</sup>), Mn<sup>2+</sup>(3d<sup>5</sup>), Fe<sup>3+</sup>(3d<sup>5</sup>), Co<sup>2+</sup>(3d<sup>7</sup>), or Ni<sup>2+</sup>(3d<sup>8</sup>) can be considered dilute magnetic semiconductors (DMS).<sup>3,4</sup> At higher doping levels, antiferromagnetism and superparamagnetism may coexist due to the formation of dopant clusters.

Furthermore, the size of nanoparticles also influences their magnetic properties.<sup>3,4</sup>

In the case of Cr<sup>3+</sup> in a SnO<sub>2</sub> matrix,<sup>5–9</sup> depending on the synthesis route, such as self-combustion, a solubility limit of 2 mol % is achieved, whereas for samples prepared by the coprecipitation route, followed by annealing at high temperatures, ferromagnetism is observed for % Cr  $< 2.5$ , and cluster formation at Cr concentration  $> 2.5\%$  was identified by electron paramagnetic resonance (EPR). EPR experiments have revealed various isolated paramagnetic Cr<sup>3+</sup> centers as well as broad signals associated with antiferromagnetic Cr<sup>3+</sup>-clusters and ferromagnetic resonance lines mainly due to Cr<sup>4+</sup>–Cr<sup>3+</sup> or Cr<sup>4+</sup>–Cr<sup>4+</sup> exchange-coupled pairs. On the basis of DFT and the formation energy calculations taking into account a supercell, the presence of two Cr<sup>4+</sup> defects in oxygen-rich conditions gives rise to ferromagnetic coupling, whereas the occurrence of Cr<sup>3+</sup> pairs associated with oxygen vacancies in oxygen-poor conditions leads to paramagnetic interactions.<sup>10</sup> To the best of our knowledge, the occurrence of ferromagnetic coupling in Cr-doped cassiterite<sup>11–14</sup> can only be explained by the presence of Cr<sup>4+</sup> or electrons trapped in the lattice due to

oxygen defects, in interaction with other paramagnetic centers such as  $\text{Cr}^{3+}$  or  $\text{Cr}^{4+}$ .<sup>15–18</sup> The low solubility of  $\text{Cr}^{3+}$  may be related to its lower ionic size (0.62 Å in an octahedral site) compared to that of  $\text{Sn}^{4+}$  (0.69 Å), generating a more ionic bond and inducing oxygen defects into the rutile lattice, with octahedral sites sharing edges or faces, creating constraints.

The Cr doping of tin dioxide also contributes to tuning of the optical absorption properties, and chrome tin orchid cassiterite of the general formula  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  exhibits a purple hue.<sup>11,12,14</sup> However, the origin of this color is still unresolved but may be mainly due to the occurrence of  $\text{Cr}^{3+}/\text{Cr}^{4+}$  valence states in the  $\text{SnO}_2$  matrix. In malayaite  $\text{CaSnSiO}_5$  doped with chromium (about 2%) as ceramic pigment, which develops deep burgundy to pink shades, optical absorption spectroscopy investigations confirm the occurrence of both  $\text{Cr}^{4+}$  and  $\text{Cr}^{3+}$  ions substituting for  $\text{Sn}^{4+}$  in octahedral sites.<sup>11,15,19–24</sup> Neutron and X-ray diffraction analyses confirm the low Cr ion solubility (<3.5 mol %) in pyrochlore  $\text{Y}_2\text{Sn}_2\text{O}_7$ , and diffuse reflectance spectroscopy reveals the occurrence of  $\text{Cr}^{4+}$  in octahedral sites, also substituting for  $\text{Sn}^{4+}$  and leading to a magenta color.<sup>25</sup> On the basis of the Tanabe–Sugano diagram, the crystal field of a  $\text{Cr}^{4+}$  ( $3d^2$ ) ion stabilized in an octahedral site in malayaite and pyrochlore structures is high and estimated to be around 2.5 eV, with a low Racah parameter  $B$  varying between 550 and 650  $\text{cm}^{-1}$  and a deduced nephelauxetic effect  $\beta = B/B_{\text{free}}$  around 0.55–0.62, showing the highly covalent character of the  $\text{Cr}^{4+}-\text{O}^{2-}$  bond. X-ray photoelectron spectroscopy analysis of Cr-doped malayaite  $\text{CaSnSiO}_5$ – $\text{CaTiSiO}_5$ , perovskite  $\text{CaSnO}_3$ – $\text{CaTiO}_3$ , or rutile  $\text{SnO}_2$ – $\text{TiO}_2$  shows that the  $\text{Cr}^{3+}$  content increases as  $\text{Ti}^{4+}$  ions are substituted for  $\text{Sn}^{4+}$  and the color changes from reddish purple to brown.<sup>17</sup> To the best of our knowledge, despite the lower ionic size of  $\text{Ti}^{4+}$  in a 6-fold coordination (0.61 Å) compared to that of  $\text{Sn}^{4+}$  (0.69 Å), but closer to  $\text{Cr}^{3+}$  (0.62 Å) and *a fortiori* to  $\text{Cr}^{4+}$  (0.55 Å), a stronger hybridization and covalence in  $\text{SnO}_2$  involving 5p orbitals instead of 3d orbitals in  $\text{TiO}_2$  may explain the preferential stabilization of  $\text{Cr}^{4+}-\text{O}^{2-}$  covalent bonds in tin-based oxides. This point will be further discussed.

Thus, cassiterite  $\text{SnO}_2$  is a simple network where chromium induces a purple coloration, but the origin of this hue is still unclear. Moreover, the stabilization of isolated  $\text{Cr}^{4+}$  and/or  $\text{Cr}^{3+}$  species, as well as the occurrence of chromium clusters creating ferromagnetic or antiferromagnetic interactions in this network and their impact on hue change investigated, must be clarified.

In order to get information about the Cr valence states as well as the formation of Cr clusters in cassiterite and to follow their evolution versus the Cr rate and synthesis conditions,  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  ( $x = 0, 0.005, 0.01, 0.02, 0.05, \text{ and } 0.1$ ) complex oxides have been prepared by coprecipitation in a basic medium, and then annealed under  $\text{O}_2$ , air, or Ar at three different temperatures (800, 1000, and 1200 °C) for various heating durations. From the analysis of X-ray diffractograms, optical absorption, and EPR spectra, the influence of the Cr-doping rate and the synthesis conditions on the coloration of cassiterite will then be discussed and related to the distribution of Cr valence states and the occurrence of Cr clusters.

## 2. EXPERIMENTAL SECTION

$\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  complex oxides have been prepared starting from chloride precursors  $\text{CrCl}_3 \cdot 6\text{H}_2\text{O}$  (Aldrich) and  $\text{SnCl}_2 \cdot 5\text{H}_2\text{O}$  (Aldrich, 98%), coprecipitated in a basic medium (pH = 10), and matured at room temperature for 1 h. The chromium amount  $x$  is between  $x = 0$  and  $x$

= 0.10 as the solubility limit is reached. Precipitates were washed three times with water and dried in an oven at  $T = 80$  °C under air for one night. All these oxides were annealed under pure oxygen at different temperatures, 800 and 1000 °C, for 48 and 4 h, respectively. In order to follow the Cr valence state in 2% Cr– $\text{SnO}_2$ , the material was annealed under pure Ar at 1000 °C for 4 h. The pure phases crystallize with the rutile tetragonal type structure (space group (SG):  $P4_2/mnm$ ).

Powder X-ray diffraction patterns have been collected on a Philips X'Pert MPD X-ray diffractometer with Bragg–Brentano geometry, using  $K\alpha_1/K\alpha_2$  radiations ( $8^\circ < 2\theta < 80^\circ$  step of  $0.017^\circ$  and counting time of 60 s). Diffractograms have been refined using the full pattern profile matching method, with a Fullprof software package and conventional reliability factors ( $R_p$  and  $R_{wp}$ ). For the diffraction peak profiles, a full width at half maximum ( $H^2$ ) has been refined with a Pearson VII function, and the estimation of  $H^2$  determines the crystallite size by applying the Sherrer equation. No instrumental function and Lorentzian or Gaussian shapes for strains have been considered in these refinements. The unit cell and all the diffraction patterns have been indexed on the basis of rutile-type structures (SG:  $P4_2/mnm$ ).

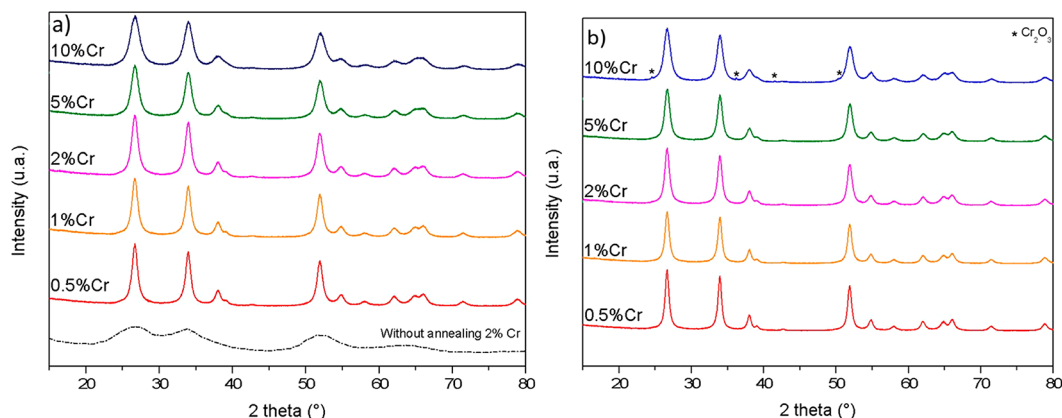
The morphology of  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  nanoparticles was evaluated by high resolution transmission electronic microscopy (HRTEM). HRTEM was performed on a TEM Jeol 2200FS operated at 200 kV with a field emission gun. The theoretical resolution (fwhm) was nominally 1.9 Å, and energy dispersive X-ray spectroscopy (EDX) was used to control the Sn/Cr distribution in nanoparticles.

The UV–vis–NIR diffuse reflectance spectra  $R(\lambda)$ , presented here after a Kubelka–Munk transformation,  $K/S = (1 - R)^2/2R$ , were recorded at room temperature from 200 to 2500 nm, with a step of 1 nm and a band length of 2 nm on a Cary 17 spectrophotometer, using an integration sphere. Halon polymer was used as a white reference for the blank. A mathematic treatment of the obtained spectra allowed the determination of the  $L^*a^*b^*$  space parameters (where  $L^*$  is the luminosity,  $a^*$  is the green-to-red axis coefficient, and  $b^*$  is the blue-to-yellow axis coefficient). The first step of the treatment consists of obtaining the XYZ tristimulus values (defined by the CIE in 1964) from the integration (on the visible range, i.e., from  $\lambda = 380$  nm to 800 nm) of the product of  $x(\lambda)$ ,  $y(\lambda)$ , or  $z(\lambda)$  functions (CIE, 1964), with the diffuse reflectance spectra function:  $X = \int x(\lambda) R(\lambda) d\lambda$ . Then, the transfer equations defined by the CIE, 1976 from XYZ space to the  $L^*a^*b^*$  space were used in order to obtain the  $L^*a^*b^*$  chromatic parameters. Diffuse reflectance spectra were also fitted considering the arctangent function for the band gap and the Gaussian curves for the d–d ( $\text{Cr}^{4+}/\text{Cr}^{3+}$ ) electronic transitions. On the basis of a Tauc plot for direct transition with the following equation,  $(ah\nu)^2 = A(h\nu - E_g)$ , the  $E_g$  BG can be estimated for pure  $\text{SnO}_2$  thermally treated at 800 or 1200 °C.

Electron paramagnetic resonance (EPR) spectra were recorded from room temperature down to 77 K using a Bruker ESP300E spectrometer operating at an X-band frequency (9.45 GHz) and equipped with a liquid nitrogen dewar. For the presented spectra, the main spectroscopic parameters are microwave power 10 mW, spectral resolution 0.6 mT/pt, and magnetic field modulation frequency 100 kHz and amplitude 1 mT. DPPH was used as an external reference to calibrate the spectrometer. For the estimation of isolated paramagnetic  $\text{Cr}^{3+}$  ion concentration, the  $\text{Fe}^{3+}$  impurity signal as well as the large resonance line centered at  $g \sim 2$  and associated with Cr clusters were first subtracted from the experimental spectra. The value obtained by the double-integration of the remaining first-derivative curve was then compared to the one calculated from the EPR spectra of the reference compound ( $\text{KAl}(\text{SO}_4)_2$ ) doped with known amounts of  $\text{Cr}^{3+}$  ions.

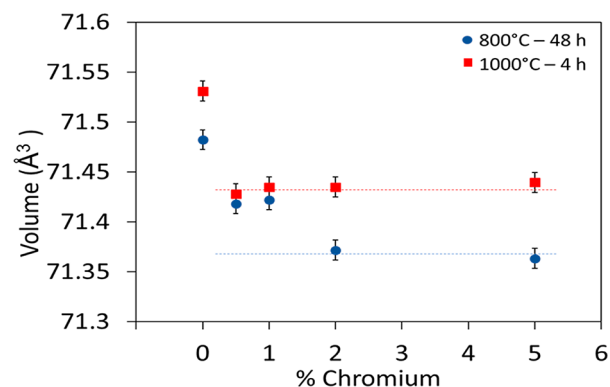
## 3. RESULTS AND DISCUSSION

**3.1. XRD and TEM Analysis.** The  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  complex oxides, where  $x$  represents the chromium rate, were prepared from the coprecipitation route under a basic medium, followed by annealing under an oxygen atmosphere at 800 °C for 48 h



**Figure 1.** XRD patterns of  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  compounds (with  $x = 0.005, 0.01, 0.02, 0.05,$  and  $0.1$ ) prepared from the coprecipitation route, followed by postannealing at  $800\text{ }^\circ\text{C}$ , 48 h (part a) and  $1000\text{ }^\circ\text{C}$ , 4 h (part b).  $\text{Cr}_2\text{O}_3$  (\*) second phase is identified. XRD pattern of  $x = 0.02$  sample before annealing is also represented in part a.

and at  $1000\text{ }^\circ\text{C}$  for 4 h. One should notice that the main diffraction lines corresponding to the rutile structure are already identified on the phase obtained just after precipitation and before annealing. The X-ray diffraction patterns of the as-prepared powders are reported in Figure 1. A single rutile ( $P4_2/mmm$  space group) phase is identified for all samples annealed at  $800\text{ }^\circ\text{C}$  (Figure 1a) and for an  $x$  value inferior or equal to 5 mol % of chromium for powders annealed at  $1000\text{ }^\circ\text{C}$  (Figure 1b). For an  $x = 0.10$  sample annealed at  $T = 1000\text{ }^\circ\text{C}$ , the second phase corresponds to  $\text{Cr}_2\text{O}_3$ , and this powder will be not considered further in this study. The solubility limit of chromium in  $\text{SnO}_2$  is reached between 0.05 and 0.10 for the samples annealed at  $T = 1000\text{ }^\circ\text{C}$ , and it is over 5 mol % for the oxides annealed at  $T = 800\text{ }^\circ\text{C}$ . The unit-cell parameters and the crystallite size deduced from full pattern profile matching have been reported in Table 1, and the unit cell



**Figure 2.** Variation of the unit-cell volume versus  $x$  chromium content in  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  oxides with various thermal treatments under  $\text{O}_2$  (red squares,  $800\text{ }^\circ\text{C}$ , 48 h and blue dots,  $1000\text{ }^\circ\text{C}$ , 4 h).

**Table 1. Unit-Cell Parameters and Crystallite Size of  $\text{SnO}_2$  Compounds<sup>a</sup>**

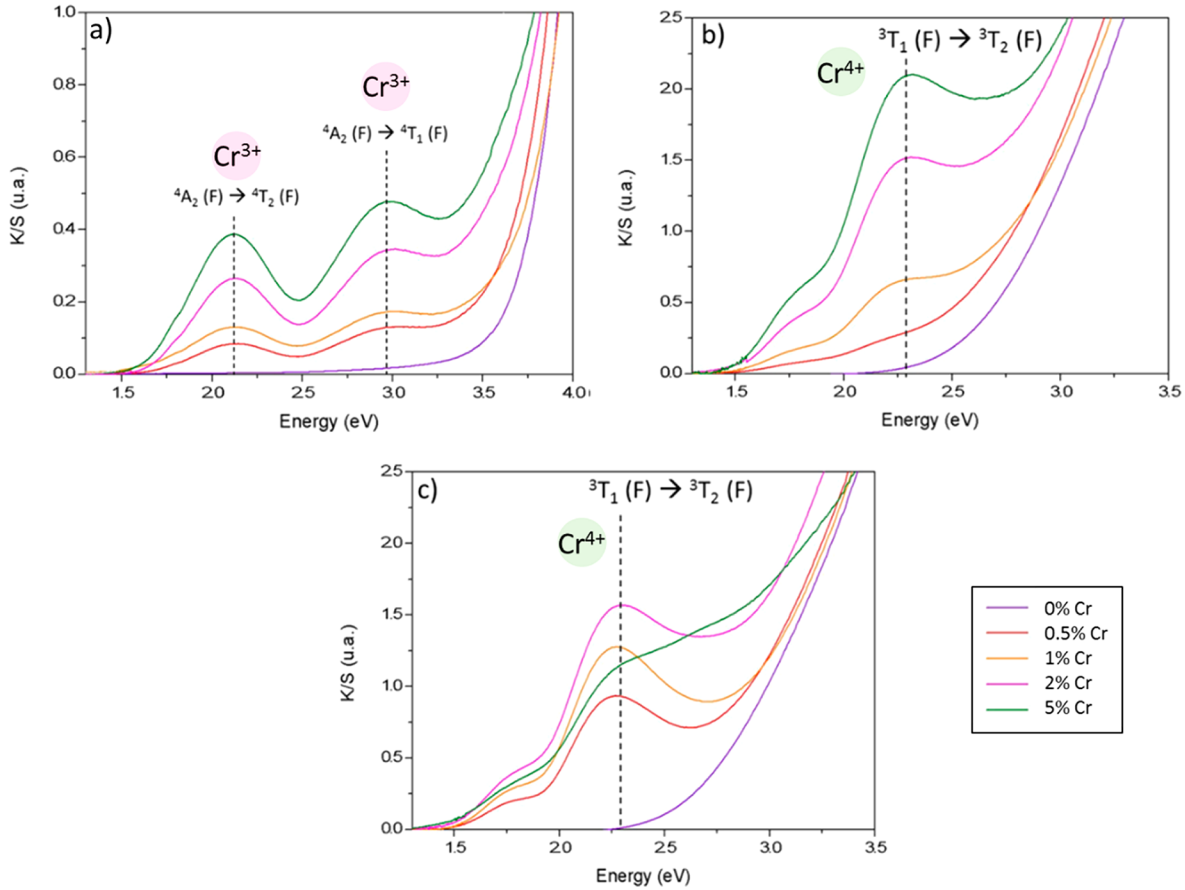
thermal treatment	% Cr	$a$ (Å)	$c$ (Å)	$c/a$	crystallite size (nm)
800 °C, 48 h	0	4.7373(6)	3.1884(6)	0.672	8(1)
	0.5	4.7357(7)	3.1846(7)	0.672	8(1)
	1	4.7353(6)	3.1852(6)	0.673	8(1)
	2	4.7341(7)	3.1848(7)	0.673	7(1)
	5	4.7338(8)	3.1846(7)	0.673	6(1)
1000 °C, 4 h	0	4.7364(3)	3.1866(2)	0.673	13(1)
	0.5	4.7355(4)	3.1852(4)	0.673	11(1)
	1	4.7358(4)	3.1851(4)	0.673	10(1)
	2	4.7358(5)	3.1851(5)	0.673	9(1)
	5	4.7361(5)	3.1849(5)	0.672	9(1)

<sup>a</sup>Undoped and Cr-doped with different molar Cr concentrations, annealed at  $800\text{ }^\circ\text{C}$  for 48 h and at  $1000\text{ }^\circ\text{C}$  for 4 h under  $\text{O}_2$ .

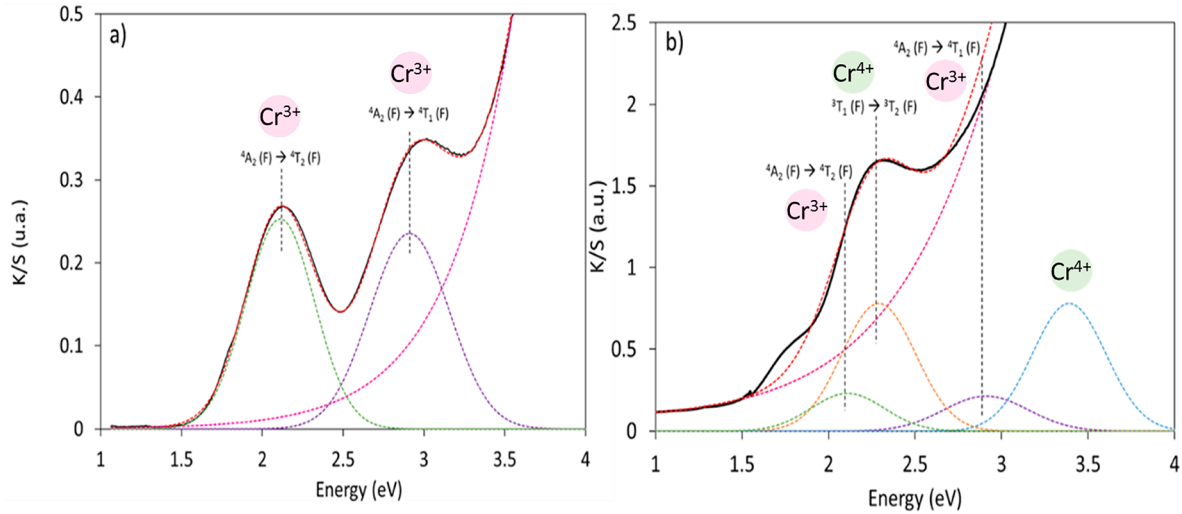
volume evolution versus  $x$  chromium rate is plotted in Figure 2. An example of a full pattern matching of diffractogram of  $\text{Sn}_{0.95}\text{Cr}_{0.05}\text{O}_2$  composition annealed at  $800\text{ }^\circ\text{C}$  under  $\text{O}_2$  for 48 h is proposed on Figure S1. In a cassiterite  $\text{SnO}_2$  network, the  $c/a$  ratio related to the octahedron distortion is around 0.673(1), which is superior to 0.645 for  $\text{TiO}_2$  (0.659 for  $\text{CrO}_2$  and 0.653 for  $\text{MnO}_2$ )<sup>26–28</sup> and also larger than the ideal value 0.586 ( $2 - \sqrt{2}$ ), corresponding to a perfect octahedral site. Even if the six Sn–O bond distances are almost equal around 2.055 Å, the  $\text{SnO}_6$  octahedron is more distorted along the 3-

fold axis than the  $\text{TiO}_6$  polyhedron. Whatever the  $x$  concentration, the  $c/a$  ratio remains identical and equal to 0.673. However, the  $\text{Cr}^{4+}$  ( $t_2^2$ ) ion can be considered as a second-order Jahn–Teller (SOJT) ion stabilized in distorted octahedral sites, whereas the  $\text{Cr}^{3+}$  ( $t_2^3$ ) ion gives rise to an isotropic or slightly distorted octahedron. Thus, concerning the octahedron distortion, there is specific competition between  $\text{Cr}^{4+}$  and  $\text{Cr}^{3+}$  cations. Because of the lowest ionic size of  $\text{Cr}^{3+}/\text{Cr}^{4+}$  (0.62/0.55 Å in 6-fold coordination vs 0.69 Å for  $\text{Sn}^{4+}$ ), a decrease of both  $a$  unit-cell parameter and volume is expected. However, the presence of oxygen defects associated with the  $\text{Cr}^{3+}$  aliovalent doping leads to a decrease of the Madelung energy and an expansion of the unit cell. Thus, a competition can take place between the incorporation of dopants with a lower ionic size and the creation of oxygen defects, which contribute to enlarge the rutile network. The largest decreases of the  $a$  unit-cell parameter and volume are observed for samples annealed at  $T = 800\text{ }^\circ\text{C}$ . One should mention that the  $c$  parameter changes abruptly for the lowest Cr content (0.5 mol %) and then remains constant no matter the Cr rate. Between 2 and 5% of Cr, the unit-cell parameters are constant and tend to show the formation of Cr clusters. Concerning the samples annealed at  $1000\text{ }^\circ\text{C}$ , the unit-cell parameters and the volume seem almost identical no matter the chromium-doping rate. For both these annealing temperatures, the higher the chromium content is, the lower the crystallite size (extracted from Rietveld refinement) is in a first





**Figure 3.**  $K/S$  optical absorption spectra of  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  compounds (with  $x = 0.005, 0.01, 0.02,$  and  $0.05$ ), just after coprecipitation (part a), annealed at  $800\text{ }^\circ\text{C}$ , 48 h (part b) and at  $1000\text{ }^\circ\text{C}$ , 4 h (part c) under  $\text{O}_2$ . The d-d transitions corresponding to  $\text{Cr}^{3+}$  and  $\text{Cr}^{4+}$  in the octahedral site are labeled in both these figures.



**Figure 4.** Deconvolution of  $K/S$  absorption spectra with only allowed transitions (Laporte selection rules) for  $\text{Sn}_{0.98}\text{Cr}_{0.02}\text{O}_2$  compound before annealing (part a) and after a thermal treatment at  $800\text{ }^\circ\text{C}$  for 48 h (part b).

approximation considering the error bars. The phases with the highest Cr content exhibit the smallest crystallites. In order to explain the Cr effect on the crystallite sizes, one should consider a partial Cr segregation at the particles surface. The stabilization of the unit-cell parameters is correlated to this segregation and/or to the oxygen defects associated with  $\text{Cr}^{3+}$  ions. The unit-cell volume (Figure 2) remains constant starting

from 0.5% Cr for samples annealed at  $T = 1000\text{ }^\circ\text{C}$  and 2% Cr for oxides calcined at  $T = 800\text{ }^\circ\text{C}$ . Such an evolution is related to more reducing conditions at  $T = 1000\text{ }^\circ\text{C}$  (higher  $\text{Cr}^{3+}$  content and associated defects) and to a faster segregation phenomena, leading to Cr clusters due to crystal growth.

TEM micrographs of  $x = 0.01$  composition are depicted in Figure S2 versus the thermal treatment. Powders obtained just

after coprecipitation and drying are monodispersed in size and in shape. The crystallite size varies between 2 and 5 nm (Figure S2a,d). Micrographs on the powders annealed at 800 °C for 48 h (Figure S2b,e) reveal a slight increase of the grain size, which is between 5 and 25 nm. (*hkl*) crystallographic planes of the rutile phase (Figure S2b) are well detected, and EDX analyses show the rather high homogeneity between chromium and tin (Figure S2g). The one annealed at 1000 °C for 4 h exhibits larger aggregates around 70 nm (Figure S2c,f).

**3.2. Optical Absorption Analysis.** The optical absorption spectra of  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  compounds obtained after coprecipitation in a basic medium and annealed at  $T = 800$  °C (48 h) or  $T = 1000$  °C (4 h) under  $\text{O}_2$  are depicted in Figure 3a–c. Pictures of powders with  $\text{Sn}_{0.95}\text{Cr}_{0.05}\text{O}_2$  composition annealed under  $\text{O}_2$  at 800 °C for 48 h and 1000 °C for 4 h are represented in Figure S3. Clear differences appear between the as-precipitated and the postannealed samples. Two Gaussian peaks around 2.1 eV (590 nm) and 2.9 eV (427 nm) in the visible region with almost the same magnitude are identified for oxides before annealing, whereas one major Gaussian function centered at 2.3 eV (539 nm) is detected for Cr-based cassiterite calcined at high temperatures. The spectrum of the  $x = 0.02$  compound obtained just after coprecipitation was deconvoluted into two bands with a Gaussian shape associated with d–d transitions of  $\text{Cr}^{3+}(3d^3)$  in an octahedral coordination and a charge transfer band (CTB) at higher energies (arctan( $h\nu$ ) function) (Figure 4a). On the basis of the Tanabe–Sugano diagram<sup>29–32</sup> ( $d^3$  electronic configuration), the two bands correspond to the d–d transitions  ${}^4\text{A}_2(\text{F}) \rightarrow {}^4\text{T}_2(\text{F})$  at 2.1 eV (585 nm) and  ${}^4\text{A}_2(\text{F}) \rightarrow {}^4\text{T}_1(\text{F})$  at 421 nm. The crystal field splitting and the *B* Racah parameter were thus estimated around 2 eV and 680  $\text{cm}^{-1}$ , respectively, in good agreement with the literature.<sup>20</sup> These energetic features are reported in Table 2.

**Table 2. *B* Racah Parameters and 10*Dq* Crystal Field Splitting Associated to  $\text{Cr}^{4+}$  and  $\text{Cr}^{3+}$  in Cr-Doped  $\text{SnO}_2$ <sup>a</sup>**

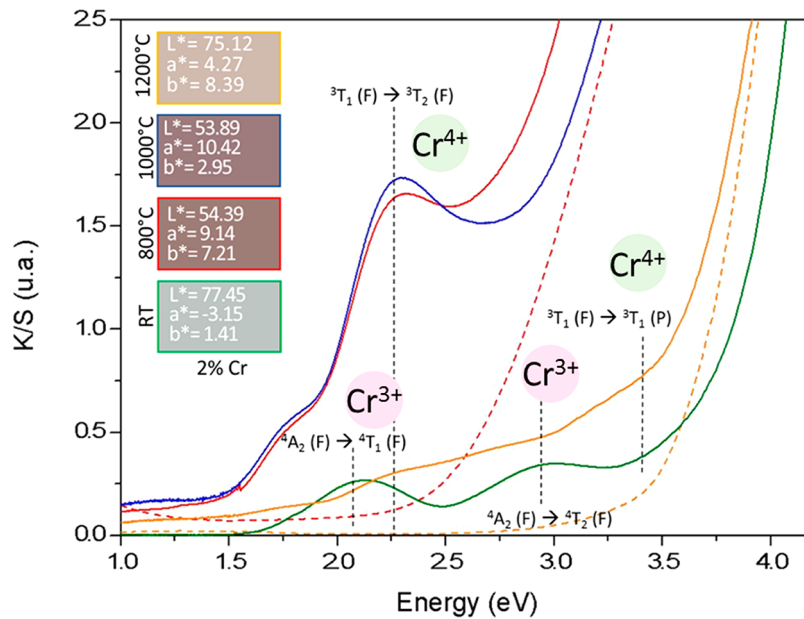
ion	<i>B</i> ( $\text{cm}^{-1}$ )	10 <i>Dq</i> (eV) ( $\text{cm}^{-1}$ )	$E_1$ (eV/nm)	$E_2$ (eV/nm)
$\text{Cr}^{3+}$	680	1.98	${}^4\text{A}_2(\text{F}) \rightarrow {}^4\text{T}_2(\text{F})$	${}^4\text{A}_2(\text{F}) \rightarrow {}^4\text{T}_1(\text{F})$
		15 980	2.12/585	2.94/421
$\text{Cr}^{4+}$	650	2.4	${}^3\text{T}_1(\text{F}) \rightarrow {}^3\text{T}_2(\text{F})$	${}^3\text{T}_1(\text{F}) \rightarrow {}^3\text{T}_1(\text{P})$
		19 470	2.29/540	3.38/366

<sup>a</sup>The two d–d transitions energies are also reported.

After it was annealed, the intense band centered at 540 nm (2.3 eV) was attributed to a  ${}^3\text{T}_1(\text{F}) \rightarrow {}^3\text{T}_2(\text{F})$  transition associated with  $\text{Cr}^{4+}$ . Taking into account the Tanabe–Sugano diagram for  $d^2$  electronic configuration and a *B* Racah parameter of 650  $\text{cm}^{-1}$  estimated in  $\text{Y}_2\text{Sn}_2\text{O}_7\text{Cr}^{4+}$  pyrochlore phase,<sup>25</sup> the 10*Dq* crystal field parameter becomes equal to 2.4 eV. With both these energetic parameters (10*Dq* and *B*), a second band should appear in the UV range at 3.39 eV (366 nm) and correspond to a  ${}^3\text{T}_1(\text{F}) \rightarrow {}^3\text{T}_1(\text{P})$  transition. The *B* Racah parameter value is also consistent with that obtained for  $\text{Cr}^{3+}$  (680  $\text{cm}^{-1}$ ) and associated with a more polarizing cation with a higher nephelauxetic effect. The larger 10*Dq* crystal field is in good agreement with the literature<sup>19–25</sup> but also higher than the 10*Dq* ( $\text{Cr}^{3+}$ ) previously estimated value. Nevertheless, a deconvolution is depicted in Figure 4b with the two d–d bands (with the same intensity) of  $\text{Cr}^{4+}$ , but without excluding the two d–d bands related to  $\text{Cr}^{3+}$  (the same magnitude for

the two d–d bands in the visible range as well as the same energy position identified in Figure 4a have been considered) and CTB at a higher energy. Thus, it seems that the d–d band intensity related to  $\text{Cr}^{3+}$  species remains much lower than the ones related to  $\text{Cr}^{4+}$ . Furthermore, the local environment of ( $d^2$ )- $\text{Cr}^{4+}$  ions is probably more distorted or more acentric than that of ( $d^3$ )- $\text{Cr}^{3+}$ . The  $t_2^2$  electronic configuration of  $\text{Cr}^{4+}$  may indeed induce a pseudo-Jahn–Teller effect (SOJT) of the octahedron involving  $t_2$  orbitals. However, a small band around 1.6 eV (775 nm) has not been taken into account in this deconvolution (Figure 4b). This last contribution may be related to the  $\text{Cr}^{4+}$ – $\text{Cr}^{3+}$  intervalence band. The intensity of the d–d bands related to  $\text{Cr}^{3+}$  and  $\text{Cr}^{4+}$  increases regularly versus the Cr-doping rate (Figure 3). However, only for the  $x = 0.05$  sample annealed at  $T = 1000$  °C, the d–d band  ${}^3\text{T}_1(\text{F}) \rightarrow {}^3\text{T}_2(\text{F})$  at 2.3 eV (540 nm), related to  $\text{Cr}^{4+}$ , strongly drops despite the increase of Cr content. It is the signature of the decrease of  $\text{Cr}^{4+}$  isolated species in favor of  $\text{Cr}^{4+}$ – $\text{Cr}^{3+}$  and  $\text{Cr}^{3+}$ – $\text{Cr}^{3+}$  interactions in clusters (Figure 4a–c). Various  $\text{Cr}^{3+}/\text{Cr}^{4+}$  and charge transfer bands have been identified in the optical absorption spectra of  $x = 0.02$  samples annealed at  $T = 1000$  °C for 4 h under  $\text{O}_2$ , and these bands drastically drop following the last anneal ( $T = 1200$  °C, 48 h under air) (Figure 5). This partial deconvolution (Figure S4) leads us to consider that the other bands associated with electronic transfer, such as intervalence exchanges into Cr clusters, have to be taken into account, especially for oxides calcined at  $T = 1200$  °C.

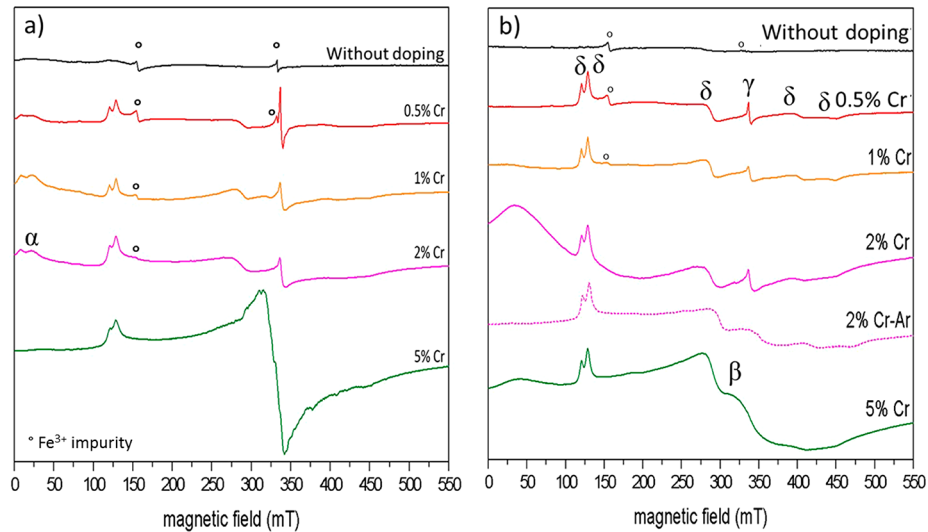
Finally, the comparison of pure and Cr-doped ( $x = 0.02$ ) cassiterite after coprecipitation, annealed at  $T = 800$  °C ( $\text{O}_2$ , 48 h),  $T = 1000$  °C ( $\text{O}_2$ , 4 h), and  $T = 1200$  °C (air, 48 h), allows for identifying qualitatively the evolution of  $\text{Cr}^{3+}/\text{Cr}^{4+}$  content with thermal treatment (unit-cell parameters and crystallite size deductions from full pattern profile matching have been reported in Table S1). The annealing temperature also has a clear impact on the band gap (Figure 5). After treatment at  $T = 800$  °C and  $T = 1000$  °C, the contribution at 540 nm (2.3 eV) related to  $\text{Cr}^{4+}$  is high, whereas at  $T = 1200$  °C, the intensity of this band is low and the electronic transfer into clusters dominates. Moreover, Tauc plots ( $((K/S)h\nu)^2 = f(h\nu)$ ) considering a direct band gap (Figure S5) give band gap values of about 3.7 eV for pure  $\text{SnO}_2$  annealed at  $T = 1200$  °C and 3 eV for  $\text{SnO}_2$  calcined at  $T = 800$  °C. Thus, after treatment at  $T = 800$  °C, a large content of defects into the cassiterite network at the surface (because of the small crystallite size around 7 nm and in approximation of spherical particles, the surface contribution versus bulk is close to 80%) contributes to destabilize the valence band and/or stabilize the conduction band and probably to form  $\text{Cr}^{4+}$  species in Cr-doped samples. Moreover, the stronger  $5p^05s^0$  ( $\text{Sn}^{4+}$ )– $2p^6$  ( $\text{O}^{2-}$ ) orbitals, overlapping more than the  $3d^0$  ( $\text{Ti}^{4+}$ )– $2p^6$  ( $\text{O}^{2-}$ ) ones, give rise to  $\text{Cr}^{4+}$  stabilization as compared to the preferentially more ionic  $\text{Cr}^{3+}$  stabilization in rutile  $\text{TiO}_2$ . The creation of  $\pi$  nonbonding states in the top of the valence band as impurity energy levels should raise the oxidizing power of  $\text{SnO}_2$  when annealed at 800 °C (much lower band gap energy around 3 eV), and it should also contribute to stabilize  $\text{Cr}^{4+}$  species. Furthermore, in Figure S6, the comparison of the optical absorption spectra of Cr-doped  $\text{SnO}_2$  ( $x = 0.02$ ) annealed at  $T = 1000$  °C under  $\text{O}_2$  or Ar (Figure S6) shows a drastic decrease of the intense band at 540 nm, which is attributed to  $\text{Cr}^{4+}$  after treatment under Ar due to the reduction phenomena associated with the powder color changes from magenta ( $\text{O}_2$ ) to light brown (Ar).



**Figure 5.**  $K/S$  absorption spectra of  $\text{Sn}_{0.98}\text{Cr}_{0.02}\text{O}_2$  (solid line) and  $\text{SnO}_2$  (dotted line) compounds, at room temperature (green line), that were annealed at 800 and 1000 °C under oxygen (red and blue lines) and at 1200 °C under air (yellow lines).

	$\text{SnO}_2$	$\text{Sn}_{0.995}\text{Cr}_{0.005}\text{O}_2$	$\text{Sn}_{0.99}\text{Cr}_{0.01}\text{O}_2$	$\text{Sn}_{0.98}\text{Cr}_{0.02}\text{O}_2$	$\text{Sn}_{0.95}\text{Cr}_{0.05}\text{O}_2$
<b>Powder without annealing</b>	$L^*=96.27$ $a^*=-0.27$ $b^*=3.1$	$L^*=85.22$ $a^*=-1.91$ $b^*=2.21$	$L^*=83.31$ $a^*=-2.22$ $b^*=1.52$	$L^*=77.45$ $a^*=-3.15$ $b^*=1.41$	$L^*=73.62$ $a^*=-3.92$ $b^*=1.37$
<b>800°C - 48h</b>	$L^*=81.90$ $a^*=-1.85$ $b^*=27.79$	$L^*=75.10$ $a^*=3.03$ $b^*=24.21$	$L^*=66.08$ $a^*=6.56$ $b^*=12.73$	$L^*=54.39$ $a^*=9.14$ $b^*=7.21$	$L^*=47.91$ $a^*=8.73$ $b^*=3.87$
<b>1000°C - 4h</b>	$L^*=78.16$ $a^*=-2.54$ $b^*=19.91$	$L^*=62.92$ $a^*=10.78$ $b^*=2.76$	$L^*=58.17$ $a^*=11.86$ $b^*=-0.47$	$L^*=53.89$ $a^*=10.42$ $b^*=2.95$	$L^*=54.58$ $a^*=6.61$ $b^*=8.15$

**Figure 6.**  $L^*a^*b^*$  experimental color parameters of each compound extracted from the  $K/S$  optical absorption spectra.



**Figure 7.** Room temperature X-band EPR spectra of  $\text{Sn}_{1-x}\text{Cr}_x\text{O}_2$  oxides annealed under oxygen at 800 °C, 48 h (part a) and at 1000 °C, 4 h (part b). (° corresponds to iron(III) impurity.) The different contributions ( $\beta$ ,  $\delta$ , and  $\gamma$ , see text) related to  $\text{Cr}^{3+}$  ( $\delta$ )/ $\text{Cr}^{5+}$  ( $\gamma$ ) isolated species and  $\text{Cr}^{3+}$  clusters ( $\beta$ ) are labeled on this figure.

From UV–vis optical absorption spectra, colorimetric parameters  $L^*a^*b^*$  can be calculated and are presented in Figures 5 and 6. One can observe a green color for powders that were obtained without annealing, which is typical of  $\text{Cr}^{3+}$  in an octahedral site, whereas  $\text{Cr}^{4+}$  in an octahedral site is

associated with a magenta-purple hue. For Cr-doped  $\text{SnO}_2$  annealed at  $T = 800$  °C, pigments with purple color are obtained with 2 and 5 mol % concentration of chromium and are yellow for lower doping. Then, the magenta color is the typical hue for  $\text{Cr}^{4+}$  in an octahedral site, whereas brown and

yellow colors may correspond to other species involving Cr interactions. Thus, after annealing at  $T = 800$  °C for 48 h, a low doping rate around  $x = 0.005$  should be associated with this kind of species. Annealing at 1000 °C has another impact on the powder color. Indeed, magenta powder is obtained as SnO<sub>2</sub> is doped with 0.5–2% of chromium with a majority of isolated Cr<sup>4+</sup> ions, whereas 5% Cr doping leads to a brown powder consistent with the presence of Cr–Cr interactions, as previously mentioned. It is worth mentioning that, for a low Cr-doping rate ( $x = 0.005$ ), clear differences can be highlighted between calcination at  $T = 800$  °C, where Cr–Cr interactions seem to dominate, and annealing at  $T = 1000$  °C, where the optical absorption spectra provide evidence for Cr<sup>4+</sup> species. The formation of Cr clusters is promoted for low Cr-doping rates and extended annealing at 800 °C (2 days), whereas isolated Cr<sup>4+</sup> species predominate after a short thermal treatment at a higher temperature ( $T = 1000$  °C for 4 h).

**3.3. EPR Analysis.** Among all the chromium valence states, only the paramagnetic Cr<sup>3+</sup> and Cr<sup>5+</sup> ions are detectable by EPR. However, the EPR spectra of Cr-doped oxide compounds (TiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>, SiO<sub>2</sub>, SnO<sub>2</sub>) exhibit various signals (commonly designated as  $\alpha$ ,  $\beta$ ,  $\delta$ , and  $\gamma$ ) in different proportions depending on the Cr-doping level and the chosen synthesis route.<sup>9,33–38</sup>

The room temperature EPR spectra of Sn<sub>1-x</sub>Cr<sub>x</sub>O<sub>2</sub> compounds, with  $x = 0, 0.005, 0.01, 0.02,$  and  $0.05$ , that were annealed under oxygen at 800 °C for 48 h and 1000 °C for 4 h are depicted in Figure 7. Spectra were also recorded at  $T = 77$  K (Figure S7a,b). The various signals associated with the Cr species are labeled in Figure 7a,b and Figure S7a,b. In addition to the Cr-related resonance lines, weak EPR signals are observed at  $g_{\text{eff}} = 4.29$  and  $2.003$  due to a low amount of Fe<sup>3+</sup> ions, as an impurity.

The signal labeled  $\delta$  is characterized by a fine structure with at least five resonance lines ( $g = 5.53, 5.18, 2.32, 1.66,$  and  $1.48$ ) within the investigated magnetic field range (Figure 7b and Figure S7). It is attributed to isolated Cr<sup>3+</sup> ions (d<sup>3</sup>) and can be described by a  $S = 3/2$  spin system, with an effective  $g$  value of about 1.975 and high values of  $D$  and  $E$  zero-field splitting parameters, which are related to the local crystal field and the distortion of the [CrO<sub>6</sub>] octahedron.<sup>39,40</sup> For each Cr-doped tin oxide, the isolated Cr<sup>3+</sup> ions concentration was estimated by double integration of the EPR spectrum after the subtraction of all other observed signals. The values reported in Table 3 are indeed slightly underestimated because some weak resonance lines of the Cr<sup>3+</sup> EPR signal are expected at a higher magnetic field (i.e., out of the investigated range);

**Table 3. Isolated Cr<sup>3+</sup> Content versus Cr-Doping Rate and Thermal Treatment (Temperature and Duration)<sup>a</sup>**

thermal treatment	% Cr	$N_{\text{spins}} \times 10^{19}$	mol % (isolated Cr <sup>3+</sup> ) $\pm 0.02$	[isolated Cr <sup>3+</sup> ]/doping rate
800 °C, 48 h, O <sub>2</sub>	0.5	1.21	0.10	20%
	1	1.22	0.10	10%
	2	5.66	0.47	24%
	5	4.15	0.35	7%
1000 °C, 4 h, O <sub>2</sub>	0.5	2.42	0.20	40%
	1	2.97	0.25	25%
	2	2.14	0.18	9%
	5	5.77	0.49	10%

<sup>a</sup>Estimated from EPR spectra recorded at room temperature.

nonetheless, they are directly proportional to the real isolated Cr<sup>3+</sup> concentrations.

The narrow axial signal, labeled  $\gamma$ , with  $g_{\perp} = 1.974$  and  $g_{\parallel} = 1.945$  is assigned to the Cr<sup>5+</sup> (d<sup>1</sup>) paramagnetic species located at the surface of the tin oxides.<sup>33</sup> After annealing Cr-doped SnO<sub>2</sub> ( $x = 0.02$ ) under Ar at  $T = 1000$  °C, the intensity of this signal is significantly reduced, confirming the above assignment (Figure 7b).

The  $\beta$  signal ( $g_{\text{eff}} = 1.98$ ) is usually associated with exchange-coupled or antiferromagnetically coupled Cr<sup>3+</sup> ions ( $S = 3/2$ ), i.e., Cr<sup>3+</sup>-clusters of various sizes (two or three units).<sup>9,34</sup> This broad and symmetric resonance line is especially observed for the compounds with  $x = 0.05$  annealed at  $T = 800$  °C or  $T = 1000$  °C.

Clusters involving ferromagnetic interactions between Cr<sup>4+</sup> (t<sub>2</sub><sup>2</sup>) and Cr<sup>3+</sup> (t<sub>2</sub><sup>3</sup>) ions are characterized by a broad and intense signal (labeled  $\alpha$  in Figure S7). As the temperature decreases, this ferromagnetic resonance line becomes broader and less intense and shifts toward low magnetic fields. Such an evolution was especially observed for the  $x = 0.02$  compound annealed at  $T = 1000$  °C under O<sub>2</sub> (Figure 7 and Figure S7), attesting to the occurrence of ferromagnetic clusters in this sample in good agreement with the literature illustrated by the magnetic hysteresis loops.<sup>4,35–37</sup> For samples with Cr-doping rates up to 2 mol % that were annealed at 800 °C under O<sub>2</sub>, a smaller FM signal associated with Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions is observed at a low magnetic field (Figure 7a) and disappears at  $T = 77$  K (Figure S7a). Such an RT signal is absent for  $x = 0.005$  and  $x = 0.01$  compositions annealed at  $T = 1000$  °C, leading us to consider only isolated Cr species (Figure 7b). As mentioned previously concerning the optical absorption spectra of oxides with a low Cr-rate (0.5 and 1 mol % Cr), Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions are observed for samples annealed at  $T = 800$  °C, whereas EPR analyses confirm the occurrence of isolated species for these compositions annealed at  $T = 1000$  °C. The absence of this low-field signal for  $x = 0.02$  composition annealed at  $T = 1000$  °C under Ar is clearly the signature (Figure S7b) of the strong decrease of the Cr<sup>4+</sup> amount in this sample as well as the broadening of the  $\beta$ -signal around  $g \sim 2$  related to Cr<sup>3+</sup> clusters.

For low Cr-doping levels (0.5 and 1 mol %), the content of isolated Cr<sup>3+</sup>, reported in Table 3, is clearly higher for the samples annealed at  $T = 1000$  °C, and fewer exchange-coupled Cr<sup>3+</sup>–Cr<sup>4+</sup> species are evidenced, which is in good agreement with the optical absorption curves showing a stronger absorption at 540 nm (2.3 eV), related to the Cr<sup>4+</sup> d–d transition (Figure 3b,c). For 2 mol % Cr-doped samples, a higher content of isolated Cr<sup>3+</sup> (24%) is detected for samples annealed at  $T = 800$  °C for 48 h under O<sub>2</sub> compared to the one calcined at 1000 °C under O<sub>2</sub> for 4 h (only 9%). In this latter compound, on the basis of EPR spectra recorded at RT and  $T = 77$  K, one should note the occurrence of more ferromagnetic Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions with a lower content of isolated Cr<sup>3+</sup> than in the sample annealed at  $T = 800$  °C. Starting from a 2 mol % Cr-doping rate, annealing at  $T = 1000$  °C clearly gives rise to an enhancement of Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions for a 5 mol % Cr concentration. Furthermore, it is worthwhile noting that, as expected, the highest concentration of isolated Cr<sup>3+</sup> (almost 1 mol %) was detected for the Sn<sub>0.98</sub>Cr<sub>0.02</sub>O<sub>2- $\delta$</sub>  complex oxide annealed at  $T = 1000$  °C for 4 h under Ar, whereas the same sample that was calcined at  $T = 1000$  °C for 4 h under O<sub>2</sub> contains only 0.2 mol % of isolated Cr<sup>3+</sup> ions (Figure S7).



A competition between the stabilization of Cr<sup>4+</sup> and Cr<sup>3+</sup> isolated ions as well as the occurrence of Cr<sup>3+</sup>–Cr<sup>3+</sup> and Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions, inducing defects into SnO<sub>2</sub> matrix, occurs as a function of annealing temperature and atmosphere. Taking into account the content of isolated Cr<sup>3+</sup> (estimated by EPR), isolated Cr<sup>4+</sup> (deduced qualitatively from optical absorption), and Cr<sup>3+</sup>–Cr<sup>4+</sup> interactions (revealed by EPR and optical absorption), the following tendencies can be proposed. A high content of isolated Cr<sup>4+</sup> as well as isolated Cr<sup>3+</sup> ions can be reached after annealing under O<sub>2</sub> at T = 1000 °C for a very low Cr-doping rate (0.5 and 1 mol %). With the Cr concentration increase (up to 5 mol %) or with the annealing temperature increase, the formation of Cr–Cr interactions increases. Also the increased annealing temperature has a reductive effect favoring the isolated Cr<sup>3+</sup> concentration and Cr<sup>3+</sup>–Cr<sup>4+</sup> interactions. Then, it is better to anneal coprecipitated samples at T = 800 °C under O<sub>2</sub> for a long time (2 days) with a high Cr rate (from 2 to 5 mol %) in order to induce a high amount of isolated Cr<sup>4+</sup> ions, to reduce the number of defects due to Cr<sup>3+</sup> species, and to get a strong purple color. Furthermore, calcination at T = 1000 °C under O<sub>2</sub> with a shorter time (4 h instead of 48 h) allows for reducing the Cr content (≤2 mol %) to obtain the lighter purple hue associated with a high content of isolated Cr<sup>4+</sup> entities. Finally as far as the magnetic interactions are concerned, in the case of the intermediate Cr concentration (2 mol %) with the highest rate of ferromagnetic Cr<sup>3+</sup>–Cr<sup>4+</sup> interactions for samples annealed at T = 1000 °C, the isolated Cr<sup>3+</sup> content decreases with the annealing temperature, whereas the formation of Cr<sup>3+</sup>–Cr<sup>3+</sup> interactions is promoted for a larger Cr rate (5 mol %) along with the reduction of Cr<sup>3+</sup>–Cr<sup>4+</sup> ferromagnetic interactions.

#### 4. CONCLUSION

The study of new inorganic pigments with a pink-magenta-purple color scale led us to focus on Cr-doped Sn-based oxides and especially Sn-doped cassiterite. Actually, the hue strongly varies with the Cr content, and the chromium valence state is related to the calcination temperature under an oxygen flow. The control of the isolated Cr<sup>4+</sup>-doping rate at the origin of visible absorption band around 540 nm is tricky because of the occurrence of ferromagnetic Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions and cluster formation as well as the more reduced isolated Cr<sup>3+</sup> species. Furthermore, the intensity of this peculiar band and the shape of the UV–vis absorption spectra are strongly dependent on the Cr-doping rate as well as the temperature of the thermal treatment under an O<sub>2</sub> atmosphere, related to Cr valence states and Cr cluster stabilization. In order to go further into detail, an accurate EPR analysis has been carried out with the estimation of isolated Cr<sup>3+</sup>-doping rate versus the total Cr-doping molar concentration and the annealing temperature. Thus, an optimal amount of ferromagnetic Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions in 2 mol % Cr-doped SnO<sub>2</sub> annealed at T = 1000 °C under O<sub>2</sub> for 4 h has been evidenced, whereas lower ferromagnetic Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions appear for the same sample calcined at T = 800 °C for 2 days, with also a higher rate of isolated Cr<sup>3+</sup> ions (approximately 3 times more than a sample annealed at T = 1000 °C). Cr<sup>4+</sup>–Cr<sup>3+</sup> and Cr<sup>3+</sup>–Cr<sup>3+</sup> interactions that form into clusters are strongly dependent on the calcination temperature but also on the Cr-doping rate. A higher annealing temperature and a shorter calcination time limit the formation of isolated Cr<sup>3+</sup> species and preferentially stabilize the Cr<sup>4+</sup>–Cr<sup>3+</sup> interactions into clusters. Furthermore, the isolated Cr<sup>3+</sup>-doping rate decreases gradually versus the

total Cr content for mol % Cr ≤ 2%, leading to a strong increase of isolated Cr<sup>4+</sup> species and/or Cr<sup>4+</sup>–Cr<sup>3+</sup> cluster content. This latter key point perfectly illustrates the more covalent character of the SnO<sub>2</sub> cassiterite network than that of the rutile TiO<sub>2</sub>, with the stronger 5p(Sn)–2p(O) orbital overlap and the hybridization stabilizing a high content of highly oxidative Cr<sup>4+</sup> species in this network, with a lower band gap around 3 eV and in a first approximation, “oxidative” energy impurity levels (p-type) on the top of the valence band associated with oxygen species.

#### Supporting Information

The Supporting Information is available free of charge at <https://pubs.acs.org/doi/10.1021/acs.inorgchem.9b02943>.

Figures of full pattern profile matching, HRTEM micrographs, pictures of powders, deconvolution of K/S absorption spectra, Tauc plots, K/S optical absorption spectra, and X-band EPR spectra, and tables of unit-cell parameters and crystalline size (PDF)

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#### Author Contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

#### Notes

The authors declare no competing financial interest.

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