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Oxygen nanobubbles enhanced photodegradation of oxytetracycline under visible light: synergistic effect and mechanism

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Abstract

Nanobubbles have gained much attention due to their unusual properties such as large specific surface area, surface energy and internal energy, thereby, increase the surface reaction and improve the mass transfer efficiency. The long term stability of nanobubbles and generation of reactive oxygen species enables them to address various environmental issues for water purification without secondary pollution. However, there are few studies concerning the application of oxygen nanobubbles in the advanced oxidation processes (AOPs). In this study, first time, oxygen nanobubbles were introduced into the photoreaction system under visible light irradiation for enhanced photodegradation of oxytetracycline (OTC). Results indicated that oxygen nanobubbles significantly improved the photodegradation of OTC through high mass transfer and generation of reactive radicals during the nanobubble collapse. The oxygen nanobubble stability was confirmed to be pH dependent. The nanobubbles concentration increased from 0.76 to 3.78×10^8 particles/mL, while the mean size decreased from 205 nm to 138 nm as the pH value increased (3.2-11.0). The degradation efficiency of OTC was improved from 45% to 98% with the increase of pH (4.0-11.0). Oxygen nanobubbles hold promise to improve the AOPs, such as reducing catalysts and chemicals usage and costs. Quenching experiments demonstrated that $\bullet\text{OH}$ was the predominant active species induced the photodegradation of OTC. With

the development of nanobubbles technology, it is anticipated to realize totally chemicals free green technology. This study provides a novel nanobubble/AOPs technology, which will be more efficient and environmental friendly and will have broad application in waste water treatment and drinking water purification.

Keywords: Oxygen nanobubble; AOPs; Photodegradation; Free radicals; PPCPs; Oxytetracycline

1. Introduction

Nanobubbles are defined as gas cavities in the aqueous solution with diameters less than 1 μm . Interfacial / surface nanobubbles and free bulk nanobubbles are the two prevailing types of ultrafine/nanobubbles [1-3]. Atomic force microscope (AFM) micrographic evidence on the existence of nanobubbles established a significant milestone for nanobubble technology in physicochemical grounds [4, 5]. Since then, nanobubbles have attracted tremendous attention, which culminated in significant number of research outputs on nanobubbles generation methodologies [6-9], nanobubble nucleation [10], nanobubble dynamic stabilities as a function of different parameters [11, 12], nanobubbles properties [13-15] and mechanism of elucidating these properties [16, 17], characterization of nanobubbles [18-21] and potential applications in many fields, which include medical [22-24], industrial [25, 26], agricultural [27-29] and environmental applications [30, 31].

Nanobubbles possess large specific surface area, high surface energy and internal energy, which enhance the surface reaction and improve the efficiency for mass

transfer [32, 33]. The life span of nano-sized bubbles (hours to days) is longer than that of macro or microbubbles in the aqueous system [29]. It has been reported that the collapse of micro and nanobubbles generate reactive oxygen species (ROS), which include hydroxyl radicals, singlet oxygen, superoxide [34-37]. These properties were the impetus for the consideration of micro and nanobubbles as a useful technology for water treatment and disinfection [32, 38]. The cases in which micro and nanobubbles have been applied for pollution management include the remediation of dyes, phenol and alachlor polluted aqua system [33, 39]. Moreover, ozone nanobubbles have also showed the more effectiveness than air-nanobubbles in the decomposition of pollutants [40-42]. So far, there is paucity of research reports on the application of nanobubbles in the advanced oxidation processes (AOPs) for pollution remediation [43, 44]. Thus, synergistic effect of oxygen nanobubbles with AOPs and their detailed mechanisms is yet to be elucidated.

Photocatalytic degradation technology is one of the most efficient AOPs in the removal of organic pollutants from wastewater. It has been widely used in water treatment since the development of powdered photocatalyst. However, the use of powdered catalysts are hampered by low reusability in suspension system and less absorption in the visible light region. These factors have hindered the large-scale application and efficiency. In order to overcome these challenges, a more efficient method, to reduce the catalyst dosage or other chemical based oxidant is urgently required. Premised on the peculiar features of nanobubbles, which have been observed to influence the physicochemical properties of the medium in which they are

dispersed [45], it is considered a veritable tool that can be deployed to improve the current AOPs or enable new technologies for producing less secondary pollutants and achieving more efficient and ecofriendly treatment goals.

Tetracycline (i.e. tetracycline, chlortetracycline and oxytetracycline) is a class of antibiotics with broad-spectrum antimicrobial action against infectious bacteria. They are commonly used in animal feed as growth promoters and prophylactics for swine and cattle production [46]. Excessive usage and low biodegradability of tetracycline can potentially cause serious risk to human and environment. Traditional physical/chemical methods are ineffective in the management and complete removal of tetracycline. Photocatalytic degradation is considered an efficient and attractive method. Although, number of studies have focused on the development of high efficient photocatalysts, include $\text{TiO}_2@\text{GO}$ [47, 48], Au-CuS-TiO_2 [49], BiVO_4 [50], MWCNT/BiVO_4 [51], $\text{Ag/p-Ag}_2\text{S/n-BiVO}_4$ [52], ZnO/ZrO_2 [53], $\text{Ag}_3\text{PO}_4/\text{TiO}_2@\text{MoS}_2$ [54], $\text{Ag/Ag}_3\text{VO}_4/\text{g-C}_3\text{N}_4$ [55], but nano-catalysts preparation requires high chemical requirements and costs. The problem of recycling is also an issue, which has restricted their practical applications. Therefore, it is necessary to find more efficient and ecofriendly treatment processes to eliminate tetracycline and other antibiotics from the waste stream.

The purpose of this work is to develop a novel catalysts free nanobubble/AOPs technology and investigate the synergistic effect of oxygen nanobubbles on the photodegradation of oxytetracycline (OTC) under visible light. The oxygen nanobubbles in the reaction solution were characterized and the effects of pH,

nanobubbles size and concentration on the photodegradation process were also evaluated.

2. Materials and methods

2.1. Materials

Oxytetracycline (OTC) was purchased from Sigma-Aldrich and HPLC grade of methanol and formic acid were obtained from Fisher Chemicals (Beijing, China). 1, 4-benzoquinone (BQ), 2-propanol and NaN_3 were obtained from Aladdin (Shanghai, China). NaOH and HCl were obtained from Sinopharm Chemical Reagent Co., Ltd (China). Pure water was obtained from a Milli-Q system (18.2 M Ω , Millipore Corp., Boston, MA). All reagents were used without further treatment. High-purity oxygen (99.99%) was used as the gas source for preparing oxygen nanobubbles. Stock solution of (OTC, 10 mM) was prepared by dissolving the reagent in the ultrapure water and subsequently stored in refrigerator (4 °C) in the dark.

2.2. Nanobubbles preparation and characterization

Oxygen nanobubbles were prepared by the nanobubbles generator (Zhongnong Tianlu micro-nano bubble water technology Co. Ltd, China). Different sizes of nanobubbles were obtained by filtrating with various pore size of teflon membrane (0.45 μm and 0.22 μm). The bulk nanobubbles concentration and mean size were detected by the nanosight (Nanoparticle tracking analysis, NTA NS500, Malvern Instrument Ltd. UK). The zeta potential of nanobubbles were detected by Zetasizer Nano ZS (Malvern Instrument Ltd. UK). A fiber optic oxygen meter (PreSens, Fibox 4, Germany) was used to measure the variation of dissolved oxygen (DO) during the

photolysis reaction.

2.3. *Photolysis experiments*

The photolysis experiment was carried out in a 500 mL open cylindrical reactor, with a 250W tungsten halogen lamp (Philips, JLZ250KN, U4K-PSH, Shanghai Yaming, Co., Ltd), with a wavelength range of 400 nm to 780 nm. The reaction was conducted at a constant temperature of 25°C using a circulating thermostatically controlled water through the outer pyrex glass water-jacket of the reaction vessel (Fig. 1).

Reaction solutions were freshly prepared by dissolving OTC in various reaction system. The reaction system with oxygen nanobubbles was prepared by using nanobubble water, which was freshly generated in ultrapure water by nanobubbles generator and the ultrapure O₂ was used as gas sources. The reaction solution with ordinary oxygen aeration was prepared by ultrapure water bubbling with ultrapure O₂ about 5 min. Various concentrations of nanobubbles were obtained by diluted the higher concentration of nanobubbles water using the ultrapure water (DO 30mg/L). Thus, these solutions have the similar DO concentrations (30mg/L). The initial pH of reaction solution was adjusted with 1mM HCl and 1mM NaOH. All photoreaction experiments were repeated three times and carried out at the room temperature.

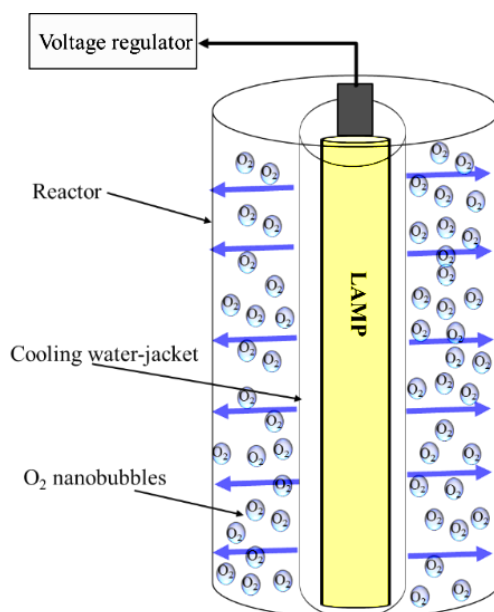


Fig. 1 Scheme of irradiation setup

2.4. HPLC analysis

At a given time interval, 3mL solution was sampled and filtered through a 0.45 μm cellulose acetate membrane for OTC determination. The concentrations of OTC was analyzed with an Agilent 1200 HPLC system with 2996 Diode Array Detector (Agilent, USA) and XTerra MS C18 column (3.5 μm , 150 mm \times 4.6 mm). The optimized mobile phase was 50% methanol and 50% formic acid (0.02%) aqueous solution. The photodiode array detector was set at 355 nm. The temperature was set at 25 $^{\circ}\text{C}$ and the flow rate was 0.5 mL/min. The photodegradation efficiency was expressed as C_t/C_0 , where C_0 is the initial concentration, and C_t is the concentration at reaction time t .

Radical quenching experiment was conducted to identify the active species in the photocatalytic reaction. 2-propanol (1 mM), NaN_3 (1 mM) and benzoquinone (1 mM) were used as scavengers for trapping hydroxyl radicals ($\cdot\text{OH}$), singlet oxygen ($^1\text{O}_2$) and superoxide radical ($\text{O}_2^{\cdot-}$), respectively. All the experiments were conducted in

triplicate.

3. Results and discussion

3.1. Nanobubbles preparation and characterization

After 10 min generation period, nanobubbles were detected in the solution by the nanoparticle-tracking analysis instrument. The time dependent variations of concentration and size distribution of nanobubbles are presented in Fig. 2. At 240 min, the size distribution of the nanobubble was still in the range of 100-500 nm and the concentration was around of $2.29\text{-}2.51 \times 10^8$ particles/mL. The mean size was 185-210 nm and these concentration and size range were maintained, even after 24 hours (Fig. 2 e-e1). Thus, oxygen nanobubbles stability in pure water was confirmed, which was in consistence with previous reports that nanobubbles have long lifetime from hours to months [29, 56-59]. The long term stability of nanobubbles and other properties such as large specific surface area, surface energy and internal energy, thereby, increase the surface reaction and improve the mass transfer efficiency, which enables them to address various issues for water purification and other applications.

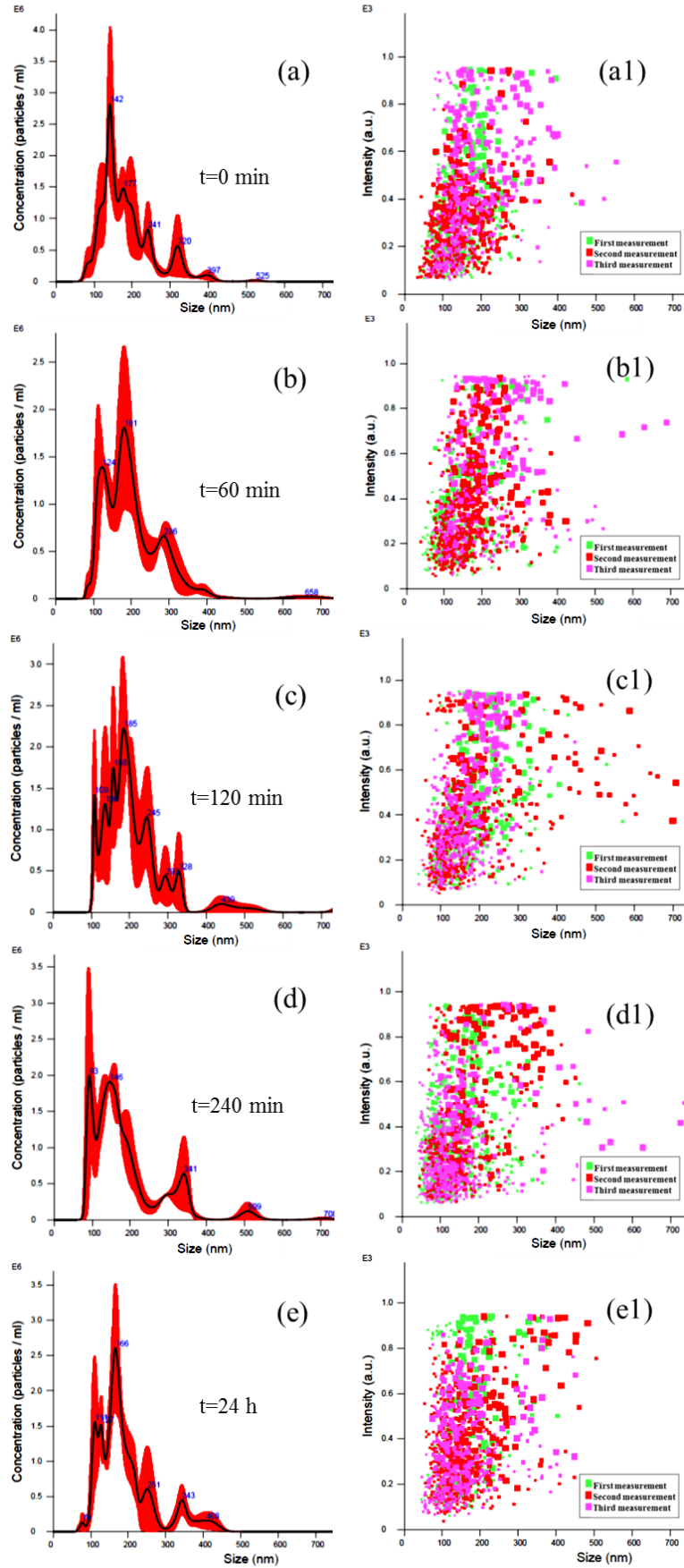


Fig. 2 Stability of oxygen nanobubbles according to time, (a-e) Concentration / Size

distribution graph; (a1-e1) Intensity / Size graph for nanobubbles and three colors
represent three times measurement.

3.2. Effect of oxygen nanobubbles on the photodegradation of oxytetracycline

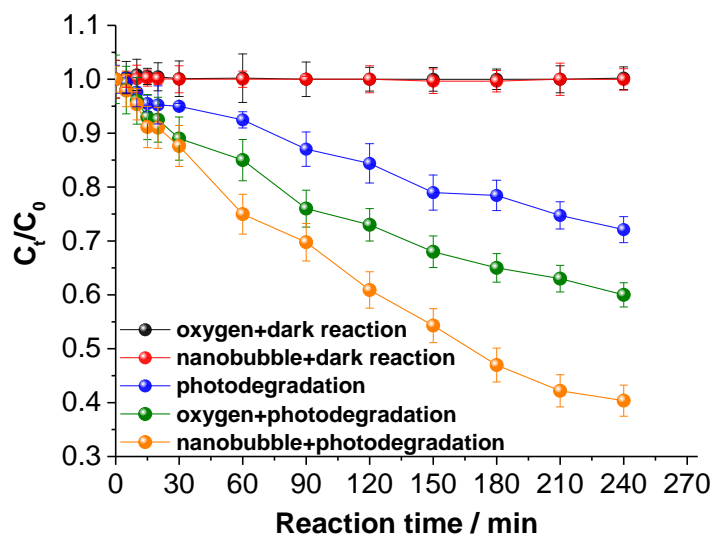


Fig. 3 Effect of oxygen nanobubbles on the photodegradation of OTC

In order to study the synergistic effects of oxygen nanobubbles on the photodegradation of OTC, experiments were carried out under different reaction conditions. As described in the experimental part, all reaction solutions were freshly prepared by using ultrapure water before irradiation. The initial concentration of oxygen nanobubbles and OTC in the reaction solution was 2.0×10^8 particles/mL and 20 μ M, respectively. All experiments were performed at pH 7.0. Dark reactions were carried out in systems with oxygen nanobubbles and ordinary oxygen aeration, respectively. The initial concentration of dissolved oxygen (DO) was 30 mg/L in both systems. The bubbles concentration in the ordinary oxygen aeration system was only 8.8×10^6 particles/mL (Fig. S1). The macro or micro-bubbles were easily dissolved into the reaction solution. As depicted in Fig. 3, no OTC degradation was detected in

the dark reaction systems. It was demonstrated that OTC could not be directly degraded by the oxygen nanobubbles or oxygen. The direct photodegradation of OTC was occurred under irradiation and the photodegradation efficiency was 28% after 4 h. This result was in agreement with the previous study that described the direct photodegradation of OTC under visible light was one of the removal pathways in water [60]. The effects of oxygen nanobubbles and ordinary oxygen aeration on the photodegradation of OTC were also compared at the same initial concentration of DO 30 mg/L. The results presented in Fig. 3 showed that with ordinary aeration of oxygen, the photodegradation efficiency was about 40% after 4 h reaction, but in the presence of the oxygen nanobubbles the photodegradation efficiency increased to 60%. The kinetics analysis showed that the OTC degradation fitted well with the pseudo-first-order correlation and the k value for OTC photodegradation in the presence of oxygen nanobubbles was 1.8 fold higher than that of oxygen aeration (Table 1, Fig. S2). These results showed that introduction of oxygen nanobubbles into the AOP operation for OTC degradation enhanced the photodegradation efficiency. The observed improvement can be ascribed to three factors viz: (a) the role of oxygen in the photodegradation process, which further helped the formation of reactive radicals that caused the degradation of OTC; (b) nanobubble presents higher stability and can provide oxygen into water slowly and last longer time when compare with the ordinary oxygen aeration (Fig. S3), which means that nanobubble can supply more oxygen and with higher efficiency during the photoreaction process; (c) nanobubbles surface effect, which can improve the efficiency of mass transfer during the

photodegradation process. It should be noted that during the nanobubble collapse, reactive radicals, such as hydroxyl radicals [34, 39], are also generated, which may be effective in the OTC degradation. However, the results presented in Fig. 3 showed that OTC degradation did not occur in the dark condition, thus only nanobubbles could not induced the degradation of OTC without irradiation.

Table 1. Kinetics analysis of the photodegradation of OTC

Reaction system	$K (\times 10^{-3}, \text{min}^{-1})$	Correlation coefficient R^2
Oxygen + dark reaction	0.01	0.20
Nanobubble + dark reaction	0.01	0.24
Photodegradation	1.34	0.99
Oxygen + photodegradation	2.14	0.98
Nanobubble + photodegradation	3.92	0.99

$k (\text{min}^{-1})$: reaction rate constant

3.3. Effect of pH

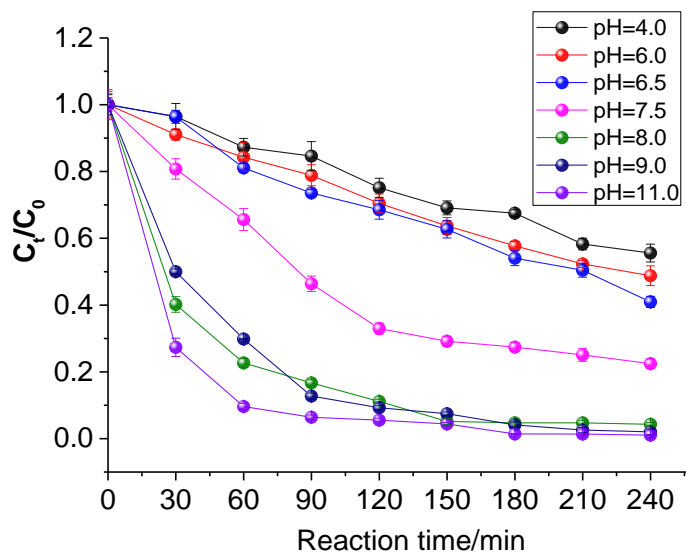


Fig. 4 Effect of pH on the photodegradation of OTC

In photochemical reactions, the pH value of the reacting mixture can play an important role. The influence of solution pH value on the process was determined at

constant concentrations of the OTC (20 μM) and nanobubbles (2.0×10^8 particles/mL, pH 7.0) but varied solution pH values from 4.0 to 11.0. The photodegradation efficiency of OTC was the initial solution pH dependent (Fig. 4). The value of the photodegradation efficiency increased from 45% to 98%, as the solution pH value increased from 4.0 to 11.0. Consequent upon the affirmed role of nanobubbles in the photodegradation of the OTC (Fig. 3), the stability of the nanobubbles was also monitored at the different operating reaction pH values. The nanobubbles concentration and size distribution fluctuated with the solution pH values. The nanobubbles concentration increased from 0.76 to 3.78×10^8 particles/mL with increasing pH value from 3.2 to 11.0, while the mean size decreased from 205 nm to 138 nm as the pH value increased (3.2-11.0). The higher efficiency of the process at higher pH value was ascribed to the increase in the concentration of the nanobubbles at higher pH values and the fact that smaller nanobubbles last a longer time and release more effective oxygen under higher pH during the reaction process. Zeta potential of oxygen nanobubble was also detected and it was also pH dependent. It had higher zeta potential (negative) and showed higher stability under higher pH conditions (Fig. 5c), which was in agreement with the previous report that bubbles with a high absolute value of zeta potential confers stability [11].

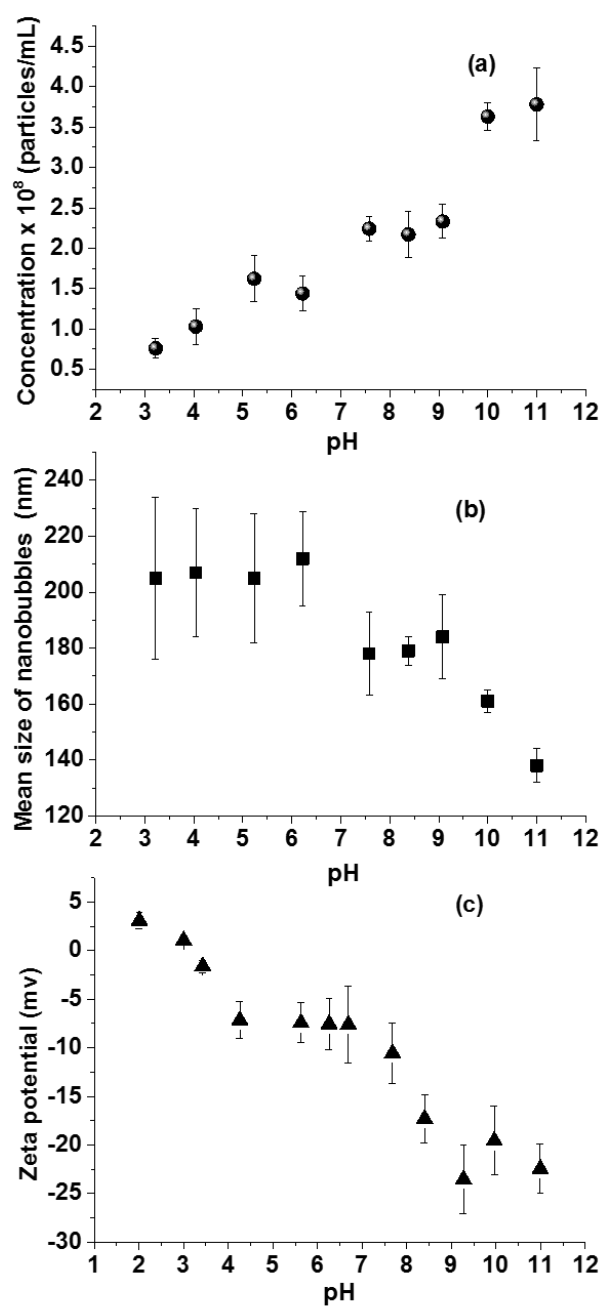


Fig. 5 Variation of concentration (a), size distribution (b), and zeta potential (c) of oxygen nanobubbles according to pH

The aqua phase species of OTC are pH dependent and the previously reported pKa values of OTC are 3.57, 7.49 and 9.88 [61]. Thus, it has four species, which include H_3OTC^+ , H_2OTC , $HOTC^-$ and OTC^{2-} , respectively. The OTC exists as protonated and deprotonated forms at pH 4.0 and 7.0, and almost fully deprotonated at

pH 8.0 and 11.0. As presented in Fig. 6, during the deprotonation process a red shift was found in the UV-vis absorption spectra. Probably, the redshift would have increased the spectrum overlap between the OTC light absorption and visible light, which has favored the direct photodegradation of OTC. Accordingly, the deprotonation of OTC was favorable to the direct photodegradation process [60].

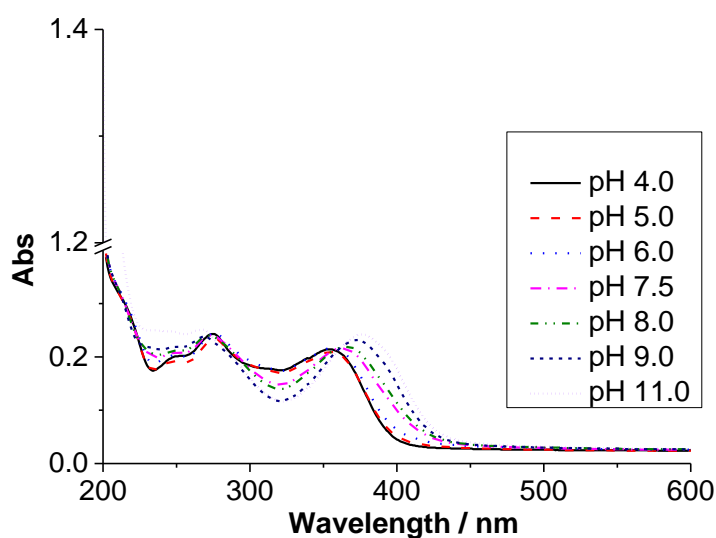


Fig. 6 UV-vis absorption spectra of OTC at various pH 4.0 to 11.0

3.4. *Effect of nanobubbles size and concentration*

In order to determine the effects of nanobubbles size and concentration on the reaction, different sizes of nanobubbles were generated by mechanical filtration with teflon membrane with different pore sizes (0.45 μm and 0.22 μm). The nanobubbles with mean size distributions of 259 nm (SD 121 nm), 202 nm (SD 74 nm) and 173 nm (SD 39 nm) were obtained separately (Fig. 7). Reaction solutions were prepared with the same initial nanobubbles concentration (1.0×10^8 particles/mL) and OTC, but with different bubbles sizes. OTC solution with smaller size nanobubbles showed higher photodegradation efficiency. It can be ascribed to the smaller bubbles have

higher interfacial area per volume of gas, surface properties and interface effect [32], and these features enhance the surface reaction and improve the efficiency of mass transfer during the photoreaction.

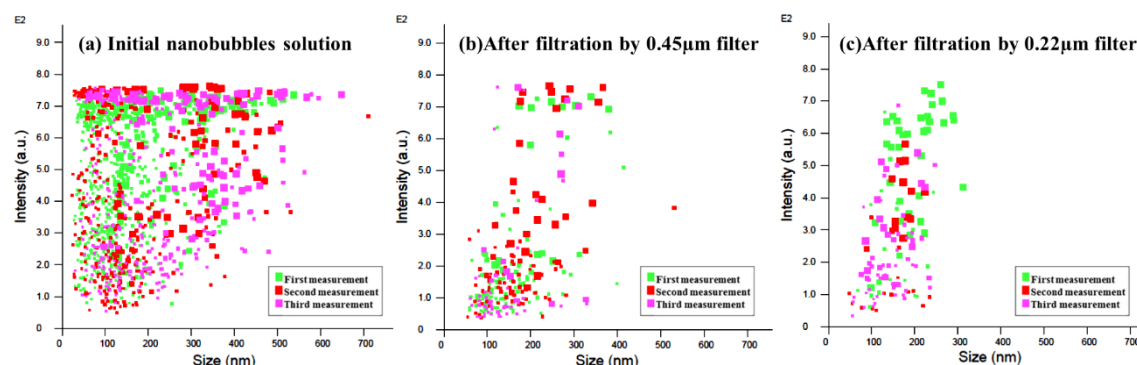


Fig. 7 Intensity / Size distribution graph for nanobubbles and three colors represent three times measurement.

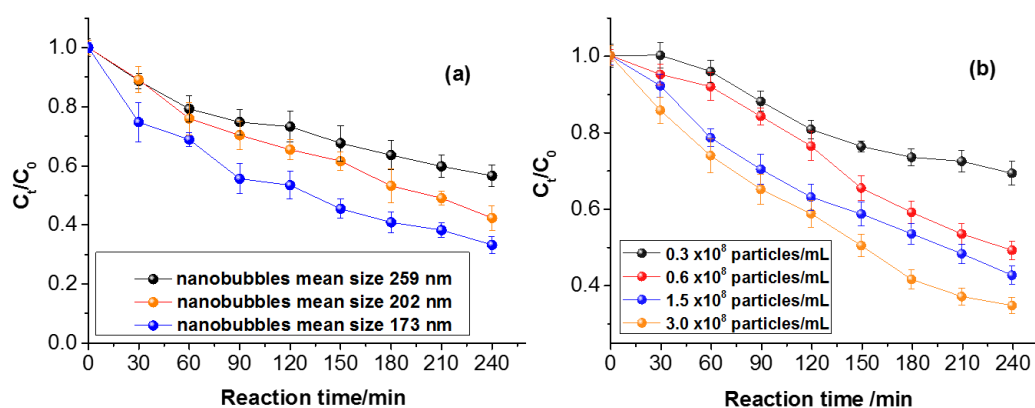


Fig. 8 Effect of nanobubbles size (a) and concentration (b) on the photodegradation of OTC (30µM)

The influence of initial nanobubbles concentration that ranged between 0.3 and 3.0×10^8 particles/mL was determined at fixed OTC concentration and nanobubbles mean size of 200 nm. The photodegradation efficiency of OTC increased with the increasing nanobubbles concentration (Fig. 8b). The higher concentration of nanobubbles not only support more oxygen during the reaction process but also

improved the interfacial effect, which strengthens the photodegradation process.

3.5. Proposed photodegradation mechanisms

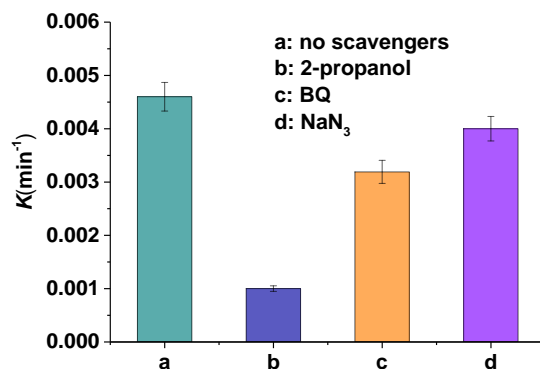


Fig. 9 Active species quenching experiments for degradation of OTC

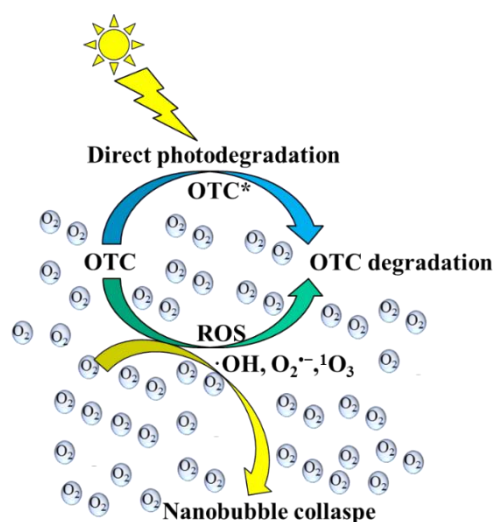


Fig. 10 The proposed mechanism of OTC photodegradation

The quenching experiment was performed to investigate the possible reactive species that induced the photodegradation of OTC. 2-propanol, benzoquinone (BQ) and NaN_3 were used as scavengers for trapping hydroxyl radicals ($\cdot\text{OH}$), superoxide radical ($\text{O}_2^{\cdot-}$) and singlet oxygen ($^1\text{O}_2$), respectively. All experiments were performed at pH 7.0. The kinetics of OTC photodegradation fit very well with the pseudo-first-order correlation. As shown in Fig. 9, the pseudo-first-order rate constant

(*K*) of OTC decreased to 0.001 min^{-1} when 2-propanol was added, thus indicating that $\bullet\text{OH}$ played the predominant role in the photodegradation reaction. With the addition of BQ, the *K* slightly increased to 0.003 min^{-1} compared with that of the 2-propanol group, which confirmed that $\text{O}_2^{\bullet-}$ was another major active species for OTC degradation. Whereas the presence of $^1\text{O}_2$ scavenger NaN_3 showed little effect, which suggested that $^1\text{O}_2$ slightly contributed to the inhibition of the photodegradation efficiency. This result was in agreement with the previous reports that $\bullet\text{OH}$, $\text{O}_2^{\bullet-}$ and $^1\text{O}_2$ coexisting within nanobubbles solution, while $\bullet\text{OH}$ was a more important oxidant in the system than $\text{O}_2^{\bullet-}$ or $^1\text{O}_2$ [35]. It was reported that the long term stability of nanobubbles (>hours) could continuously generate $\bullet\text{OH}$ [62]. Since $\bullet\text{OH}$ are nonselective and can oxidize many kinds of organic pollutants, it could be inferred that $\bullet\text{OH}$ is the predominant active species in the photodegradation of OTC.

The proposed mechanism of OTC photodegradation is shown in Fig. 10. Since the OTC species are pH dependent and the fractions of HOTC^- and OTC^{2-} have visible light absorption (Fig. 6), thus the OTC degradation can be easily triggered through formation of OTC excited state (OTC^*) at neutral or alkaline pH under visible irradiation. During the photoreaction process, oxygen is consumed and the more required oxygen can change the mass balance around oxygen nanobubbles and thus caused the collapse of oxygen nanobubbles to supply more active oxygen. As shown in Fig. S4, the oxygen nanobubble concentration was decreased from 2.0×10^8 to 0.9×10^8 particles/mL after 4 h reaction. ROS such as $\bullet\text{OH}$, $\text{O}_2^{\bullet-}$ and $^1\text{O}_2$ were generated during the collapse of nanobubbles, which can improve the removal of OTC. The

synergistic mechanism of oxygen nanobubble/photolysis process is that oxygen nanobubbles supply both dissolved oxygen and ROS to the reaction system, whereas no effect was found in the dark control system without the stimulus of photolysis degradation of OTC.

4. Conclusions

For the first time, oxygen nanobubbles were added into the photoreaction system to improve the photodegradation efficiency of OTC under visible light irradiation. From the results, oxygen nanobubbles can be stable exist in the pure water system hours to days, whereas its stability was pH dependent. Both concentration and size distribution were changed according to pH from 3.2 to 11.0. The smaller bubbles presented higher photodegradation efficiency, can be attributed to their higher interfacial area per volume of gas and high surface properties. OTC could not be directly degraded by the oxygen nanobubbles without irradiation, while during the photoreaction process, oxygen consumption improved the nanobubble collapse, which can help the generation of ROS. Quenching experiments results demonstrated that $\cdot\text{OH}$ was the predominant active species induced the photodegradation of OTC. Oxygen nanobubbles hold promise to improve the AOPs, such as reduce chemicals usage and costs. This study provides a novel principle and method for the improvement of traditional advanced oxidation technology and nanobubble/AOPs will have an important application in the removal of pollutants with high efficiency.

Acknowledgments

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Conflicts of Interest: The authors declare no conflict of interests.

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