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Measurement and modelling of solid apparition temperature for the CO₂ – H₂S – CH₄ ternary system

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Abstract:

In order to develop a new process to eliminate acid gases from natural gas, it is important to consider phase equilibrium properties. New solid-liquid equilibrium data were determined concerning the ternary system CO₂ + H₂S + CH₄. The experimental technique is based on visual synthetic method. The new experimental data were compared to predictions given by the Predictive Peng-Robinson equation of state (PPR78), PSRK UNIFAC, Peng-Robinson with Huron-Vidal mixing rules and GERG 2008 models coupling with classical approach or Jager and Span for solid phase. The results have shown that the model selected for the fluid phase has more impact than the model considered for the solid phase. New binary interaction parameters of Peng-Robinson equation of state were fitted only on our experimental data and it was found that the obtained values are very different from the value used for Vapor Liquid Equilibrium computation.

Key words: Solid formation; Ternary mixture; CO₂; H₂S; CH₄, Data treatment

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1. Introduction

In the past three decades the share of natural gas in the world panorama has been appreciably growing, to almost one quarter of the energy used worldwide. This trend is expected to increase in the future since a number of natural gases and gas condensates fields have been discovered around the world. In addition, natural gas is an energy carrier with a lower carbon footprint than coal or oil. However, this development will depend on the progress made in technologies allowing access to sour gases contaminated reserves (40% of the world's gas reserves). In some cases, the amount of hydrogen sulfide in natural gas can be greater than 30%, and natural gas should be clear of this compound and other acid gases (mercaptan, carbon dioxide, carbon disulfide, carbonyl sulfide) before its transport.

The technologies currently employed in most industrial plant to remove acid gases are based on chemical or physical absorption with solvents [1]. Since these solvents need heat to be regenerated between each absorption, the higher the concentration of acid gases is, the lower the profit. In other hand, some cryogenic process for bulk hydrogen sulfide (H_2S) or carbon dioxide (CO_2) removal from very sour gases has been developed. The kind of process, involving treating the raw gas by cryogenic distillation, has the advantage of producing sweetened gas. This process has high energetic efficiency and allows the minimisation of hydrocarbons losses. However, at low temperature, solid phase should appear. As an example, hydrates were found to develop under some operating conditions in the Lacq pilot unit [2]. Such solid phase formations are highly undesirable as they decrease the efficiency of the plant, but mainly they can plug the pipes and eventually damage the installations. In other hand, the crystallization of a compound may be desirable; for example, there are CO_2 capture processes that use fluid-solid equilibrium. Petroleum industry needs reliable and accurate vapor-liquid and liquid-solid equilibrium data for mixtures of hydrocarbons and acid gases in order to develop accurate models for calculating thermodynamic properties of natural gas requested for plants design.

MINES ParisTech fluid thermodynamic laboratory has previously developed specific apparatus and protocols in order to carry on phase equilibrium for systems with high level of H_2S . We can cite the work published by Dicko et al. [3] on the study of phase behaviour on binary systems composed with H_2S and propane, n-butane and pentane, and the work published by Coquelet et al. [4] on the phase diagram of the binary system $CH_4 + H_2S$ from 186 to 313 K.

In this paper, we present new experimental data concerning the apparition of solid phase for mixtures

composed with $\text{CO}_2 + \text{H}_2\text{S} + \text{CH}_4$ and their modelling with classical equations of state (EoS). The knowledge of this temperature is very important for the determination of optimal operating condition of pilot plants. The experimental principle is presented in section 2 and the results are presented and analysed in section 3.1. An attempt at modelling the data with different EoS was carried out with and without fitted binary interaction parameters (BIP). The method and analysis are presented in section 3.2.

2 Experimental principle

For the experiments in the present study, visual synthetic method was considered. Five compositions of the ternary mixture $\text{CO}_2 + \text{H}_2\text{S} + \text{CH}_4$ were considered. The measurements were performed using an in-house specially designed apparatus allowing the visual detection of a solid phase formation.

2.1 Materials

The considered mixtures, presented in Table 1, were purchased from AIR PRODUCTS. In the Table 1, uncertainties given by the supplier concerning the preparation of the mixtures are presented.

[Table 1]

Table 1. Chemical sample (Air Product certified values, standard ISO9001:2000).

Chemical/CAS number	Concentration (Mol %)	Relative Uncertainty/% (k=2)	Concentration (Wt %)	Uncertainty (Wt %) (k=2)	Method	P\MPa ^a
Mixture A						
Methane / 74-82-8	79.93	0.5	59.21	0.11	Gravimetric method	15.0
Carbon dioxide / 124-38-9	20.07	0.2	40.79	0.04		
Hydrogen sulphide / 7783-06-4	0	-	0	-		
Mixture B						
Methane / 74-82-8	76.03	0.2	54.81	0.04	Gravimetric method	8.0
Carbon dioxide / 124-38-9	18.99	0.2	37.56	0.04		
Hydrogen sulphide / 7783-06-4	4.98	1	7.63	0.22		
Mixture C						
Methane / 74-82-8	71.92	0.5	50.38	0.11	Gravimetric method	11.5
Carbon dioxide / 124-38-9	18.06	0.5	34.71	0.11		
Hydrogen sulphide / 7783-06-4	10.01	0.5	14.91	0.11		

Mixture D						
Methane / 74-82-8	68.02	0.5	46.42	0.12	Gravimetric method	6.1
Carbon dioxide / 124-38-9	17.01	0.5	31.85	0.12		
Hydrogen sulphide / 7783-06-4	14.97	0.5	21.72	0.12		
Mixture E						
Methane / 74-82-8	63.95	0.5	42.49	0.12	Gravimetric method	4.9
Carbon dioxide / 124-38-9	16.04	0.5	29.24	0.12		
Hydrogen sulphide / 7783-06-4	20.01	0.5	28.27	0.12		

^a Pressure of the mixture at 288.15 K

2.2 Apparatus description

An in-house developed apparatus based on a visual static-synthetic method is used in this work to perform the solid phase formation's temperature measurement (Figure 1). The equilibrium cell (length around 19mm, internal diameter around 12.7 mm and internal volume around 2.4 cm³) is comprised of a horizontal sapphire tube, which contains the mixture and a few agitation glass beads, sealed by two titanium flanges respectively on the left and the right. The left flange is connected to a platinum resistance thermometer probe (Pt100, 100 Ω). The right flange is connected a pressure transducer and a loading valve. The connected loading tube is linked, through a three-way valve, with the mixture storage bottle and a vacuum pump used for discharging, degassing and evacuation operations of the residual substance in the circuit. The equilibrium cell is immersed into a water thermoregulated liquid bath that provides and keeps the desired operating temperature within ± 0.01 K.

The temperature is measured by a platinum probe (100 Ω). The temperature probe was calibrated against a standard probe (25 Ω , TINSLEY, U.K., Model 5187 SA) connected to a precision Ohmmeter (AGILENT, Model 34420A). Temperature calibration uncertainty is estimated to be $U_{cal}(T) = 0.02$ K, in the temperature scale 188.48 K to 223.55 K.

The pressure is measured by a pressure transducer (SEDEME, Model MD200) connected to a display data acquisition unit (AGILENT, Model 34970A). During the experiments, the pressure transducer is thermoregulated at 308.15 K. The pressure transducer was calibrated using a deadweight tester (DESGRANGES & HUOT, Model 5202S CP). Pressure calibration uncertainties are estimated to be $U_{cal}(P) = 0.005$ MPa. Thermometers and pressure transducers are connected to a digital data display unit (Agilent, model 34420A) connected to a computer.

[Figure 1]

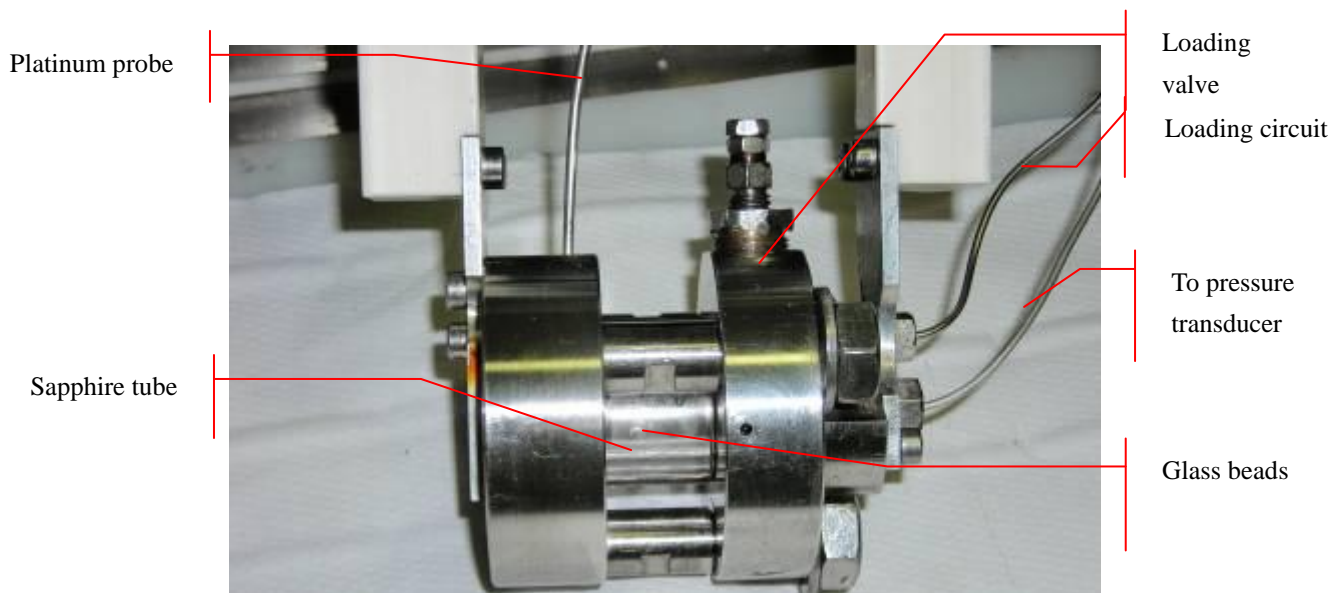


Figure 1: Picture of the equilibrium cell.

2.3 Experimental procedure

The cell and loading circuit are placed under vacuum at ambient temperature, then the cell is immersed in a thermostat liquid bath at 208.15 K. With the cell's loading valve closed, the mixture is allowed in the circuit by opening the safety valve on the mixture storage bottle and the valve connected to the loading tube. The cell's loading valve is then carefully opened to let the mixture fill the cell. The loading valve is closed when the cell is completely filled with the mixture.

When thermodynamic equilibrium is reached in the cell, the mixture is agitated by gently swinging the cell from left to right to move the glass beads inside the cell. The cell is maintained immersed in the thermostat liquid bath during the agitation process. Then the cell is extracted from the thermostat liquid bath for observation. If no solid phase is detected, the temperature of the thermostat liquid bath is lowered, or raised otherwise. The observation is done following an increase and a decrease of temperature to better identify the temperature of solid phase formation in the mixture. During the experiment, aside from the formation of the solid phase, a single phase was visually identified in the cell. Following our experimental

procedure, we have defined an uncertainty of temperature measurement which takes into account the variation of temperature between the moment where a solid phase is observed and the moment where there is no solid phase. The maximum “experimental procedure “uncertainty is less than $u(T)=1.2$ K.

Only for mixture E the detection of solid phase was not very clear, this is why we have repeated the measurement of solid formation only for this mixture. Repeatability measurement consists of removing the mixture in the equilibrium cell and loading a new one.

3. Results and discussion

3.1 Experimental results

The experimental temperature and pressure of solid phase formation for five compositions of the $\text{CO}_2 + \text{H}_2\text{S} + \text{CH}_4$ ternary system are reported in Table 2 and plotted in Figure 2. A second measurement was done with mixture E, as a repeatability test. The results are very similar, when taking into account the experimental uncertainties. Hereinafter, mixture E will refer to the average of these two measurements.

[Table 2]

Table 2: Experimental Temperature of solid phase formation for the different mixtures studied, $u_{\text{cal}}(\text{T})=0.02\text{K}$, $u_{\text{cal}}(\text{P})=0.005\text{ MPa}$.

Mixture identification	T/ K	P/ MPa	$u_{\text{EP}}(\text{T})^{\text{a}}/\text{K}$
A	210.04	2.224	1.2
B	202.60	2.186	0.4
C	197.03	1.848	0.6
D	194.47	1.974	0.6
E	192.45	2.123	1.0
	192.52	2.094	1.0

^a Uncertainty due to our experimental procedure

[Figure 2]

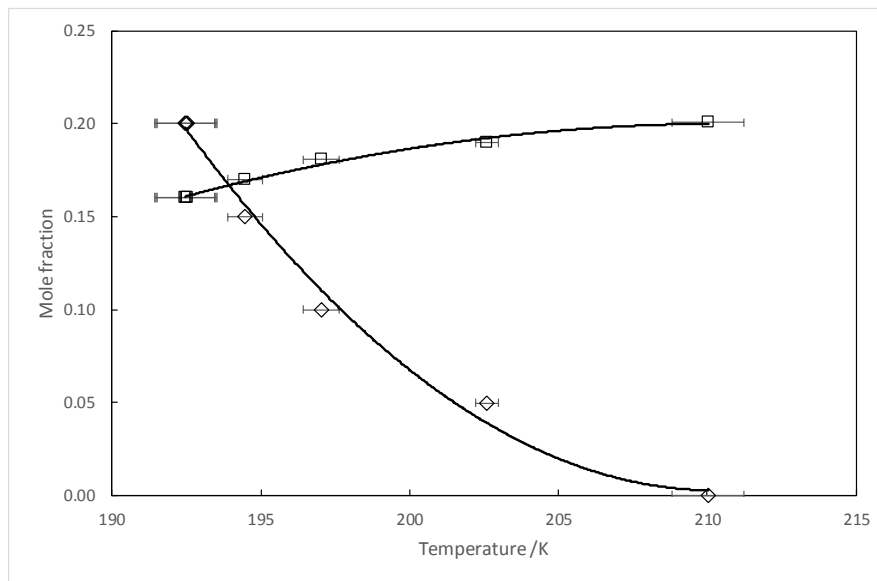


Figure 2 : Variation of acid gas composition as a function of the temperature of solid phase formation (□: CO₂, ◇: H₂S) in the CO₂ + H₂S + CH₄ ternary mixture. Solid lines: second degree polynomials fitted on the data. Error bar: values from Table 2.

In Figure 2, the temperature of solid phase formation for the mixture B ($z_{\text{CO}_2} = 0.1899$, $z_{\text{H}_2\text{S}} = 0.0498$) looks a bit out of trend, although the measurement for this mixture has not presented any particular problem. This lack of precision can be explained by a higher uncertainty on H₂S mole fraction in mixture B, as provided by AIR PRODUCTS (see Table 1).

This behaviour can be reproduced by correlating the temperature of solid phase formation with a ratio of the different mole fractions of the components of the system. In Table 3 are presented the composition and the ratios between each component. The ratios $\frac{n_{\text{CO}_2}}{n_{\text{CH}_4}}$ can be averaged over the different mixtures at

$\langle \frac{n_{CO_2}}{n_{CH_4}} \rangle = 0.2506$, with a relative deviation of $U_{rel} \left(\frac{n_{CO_2}}{n_{CH_4}} \right) = \frac{\sigma \left(\frac{n_{CO_2}}{n_{CH_4}} \right)}{\langle \frac{n_{CO_2}}{n_{CH_4}} \rangle} = 0.24\%$, so the ratio $\frac{n_{CO_2}}{n_{CH_4}}$ can be considered constant for the different mixtures. The Figure 3 shows that the temperature of solid phase formation can be well correlated (AAD = 0.38% and Bias = 0.003%) with the ratio $\frac{n_{H_2S}}{n_{CH_4}}$ by an exponential function (Eq. 1).

$$T = T_f + T_0 \exp \left(A \frac{n_{H_2S}}{n_{CH_4}} \right) \quad (1)$$

With $T_f = 190.81$ K, $T_0 = 19.27$ K and $A = -7.77$. The authors think that for other values of $\frac{n_{CO_2}}{n_{CH_4}}$, the same behavior will occur. This assertion could be tested with additional measurements.

[Table 3]

Table 3: Mixture compositions and mole fractions ratios

Mixture identification	Z_{CO_2}	Z_{H_2S}	Z_{CH_4}	$\frac{n_{H_2S}}{n_{CO_2}}$	$\frac{n_{CO_2}}{n_{CH_4}}$	$\frac{n_{H_2S}}{n_{CH_4}}$
A	0.2007	0	0.7993	0.000	0.251	0.000
B	0.1899	0.0498	0.7603	0.262	0.250	0.066
C	0.1806	0.1001	0.7192	0.554	0.251	0.139
D	0.1701	0.1497	0.6802	0.880	0.250	0.220
E	0.1604	0.2001	0.6395	1.248	0.251	0.313

[Figure 3]

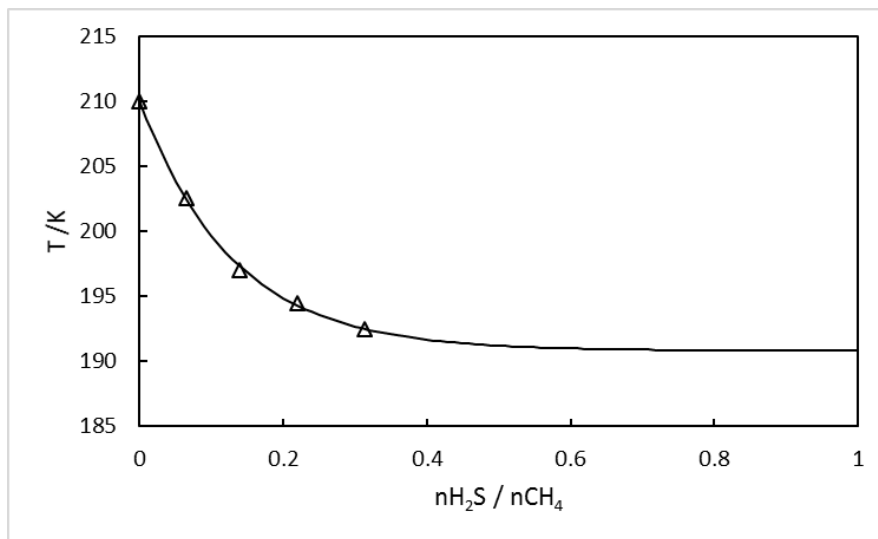


Figure 3 : Temperature of solid phase formation for different values of $\frac{n_{H_2S}}{n_{CH_4}}$. Δ : our data. Solid line: our correlation.

3.2 Thermodynamic modelling

Model for solid-liquid equilibrium calculation

The Figure 4 represents the classical PT envelop of a system including solid phase. The chosen approach for modelling the temperature of the solid phase formation is the equilibrium condition for a solid-liquid system written in terms of the fugacities of the solid-forming component, which can be assumed as pure CO₂ (Eq. 2). For the other chemical species, we assume that they are in a fluid state.

$$f_{\text{CO}_2}^F(T, P, x) = f_{\text{CO}_2}^S(T, P) \quad (2)$$

where T is the temperature, P is the pressure, x is the fluid phase concentration. The fugacity of the solid phase can be calculated (Eq. 3) from the model developed for the fluid-solid equilibrium described by Prausnitz et al. [5], page 635.

$$f_{\text{CO}_2}^S(T, P) = f_{\text{CO}_2}^{0,L}(T, P) \exp \left[\frac{\Delta h_{f,\text{CO}_2}^0}{RT_{f,\text{CO}_2}} \left(1 - \frac{T_{f,\text{CO}_2}}{T} \right) \right] \quad (3)$$

Where the suffix (0) is the reference state, $\Delta h_{f,\text{CO}_2}^0 = 192.464 \text{ kJ.kg}^{-1}$ is the molar fusion enthalpy of pure CO₂ in its reference state given by Gas Encyclopedia [6] and T_{f,CO_2} is the fusion temperature of pure CO₂.

[Figure 4]

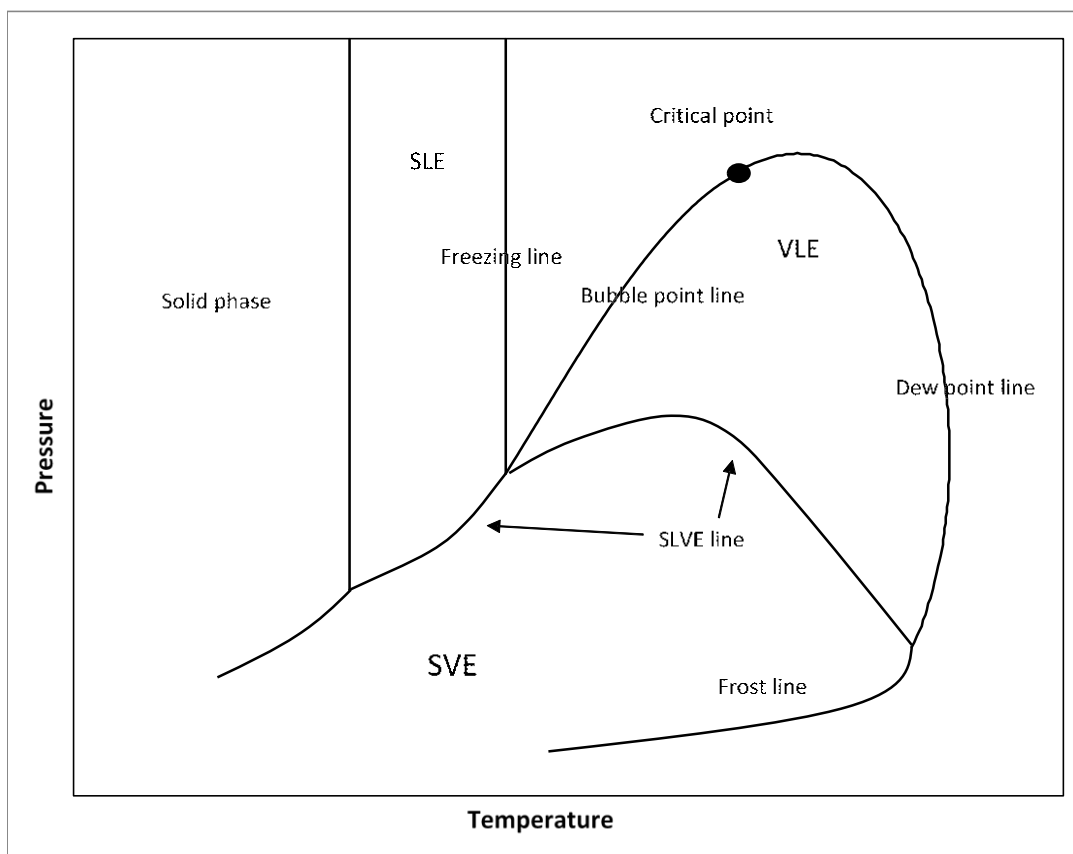


Figure 4: Classical PT envelop of a mixture including solid phase. VLE: Vapor-liquid Equilibrium, SLE: Solid-liquid equilibrium, SVE: Solid-vapor equilibrium, SLVE: Solid-liquid-vapor equilibrium.

It has been chosen to compute the fugacity of the fluid phase with different predictive EoS: Predictive Peng-Robinson EoS (PPR78) [7], Predictive Soave-Redlich-Kwong EoS UNIFAC (PSRK UNIFAC) [8], implemented in Simulis® Thermodynamics Software from Prosim (France), Peng-Robinson with Huron-Vidal mixing rules (PR-HV) [9] and GERG 2008 EoS [10] implemented in REFPROP (version 10.0) from NIST. The PR-HV EoS used Huron-Vidal mixing rules with the NRTL-V activity model for the Gibbs free energy; the parameters of this activity approach have been obtained using both VLE and VLLE data (binary and ternary subsystems). The details of the model have been given in [11] in the form of application case 1.

One objective of this work is to compare the performance of these four predictive models to correlate the experimental data presented in this paper. For each system composition and an initial temperature, the liquid phase fugacity at bubble point pressure and the liquid CO₂ fugacity at bubble point pressure are

calculated using one of the chosen EoS, i.e. PPR78, PSRK, PR-HV and GERG 2008 EoS. The solid CO₂ fugacity is calculated using the fluid-solid equilibrium model (Eq 2). Then, for each system composition, the temperature that satisfies the isofugacity condition is obtained through the minimization of the difference between the fluid phase fugacity and the solid CO₂ fugacity, using the non-linear Generalized Reduced Gradient (GRG) algorithm.

Results are presented on Table 4 and plotted in Figure 5. Mean Absolute Deviation (MAD) and Mean Relative Deviation (MRD), applied on temperatures, are defined by Eqs. (4) and (5) respectively.

$$\text{MAD} = \frac{1}{N} \sum |U_{\text{cal}} - U_{\text{exp}}| \quad (4)$$

$$\text{MRD} = \frac{100}{N} \sum \left| \frac{U_{\text{cal}} - U_{\text{exp}}}{U_{\text{exp}}} \right| \quad (5)$$

As we can see, the PR-HV EoS gives the best results in comparison with the prediction obtained using the two cubic EoS, PPR78 and PSRK, and the GERG approach. The binary parameters of the Huron-Vidal mixing rules have been obtained previously from regression work on binary systems: CH₄ + CO₂, CH₄ + H₂S, CO₂ + H₂S [12]. Among the other three models, the GERG approach is the one that gives the lowest deviations (mean deviation close to 4.1 K), the PSRK EoS shows a deviation about 10 K and PPR78 gives the larger deviation (21 K). The two later show also the larger deviations when the H₂S content increases. To confirm the impact of the fluid model on these phase equilibria, we also verified the ability of the PR-HV equation to reproduce the fluid phase diagrams of the different binary as well as the ternary. The figures 6, 7 and 8 below illustrate this ability. In the case of the ternary system, the PR-HV model makes it possible to predict the classical liquid-vapor behavior as well as the liquid-liquid situation, which is not the case with a classical PR model.

[Figure 5]

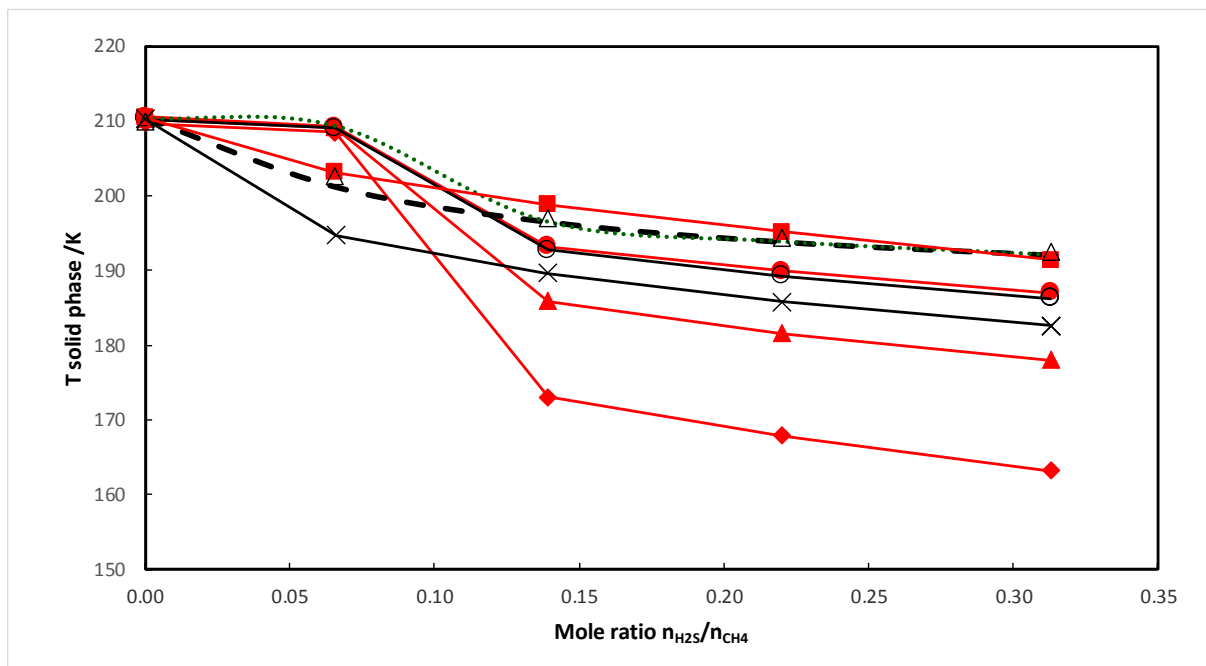


Figure 5 : Temperature of solid phase formation for different values of $\frac{n_{H_2S}}{n_{CH_4}}$.

Δ : our experimental data. \times : Calculated using PR EoS with Sandler [16] parameters. \blacklozenge : calculated data using PPR78 EoS. \blacktriangle : calculated data using PSRK UNIFAC EoS. \blacksquare : calculated data using PR-HV EoS. \bullet : calculated using GERG 2008 EoS with Eq. 3 for the calculation of solid CO_2 fugacity. \circ : calculated using GERG 2008 EoS with Jager and Span model [17] for the calculation of solid CO_2 fugacity. Green dotted line: calculation using Peng Robinson EoS with adjusted BIP. Black dashed line: calculation using Peng Robinson EoS with adjusted BIP and new composition of mixture B.

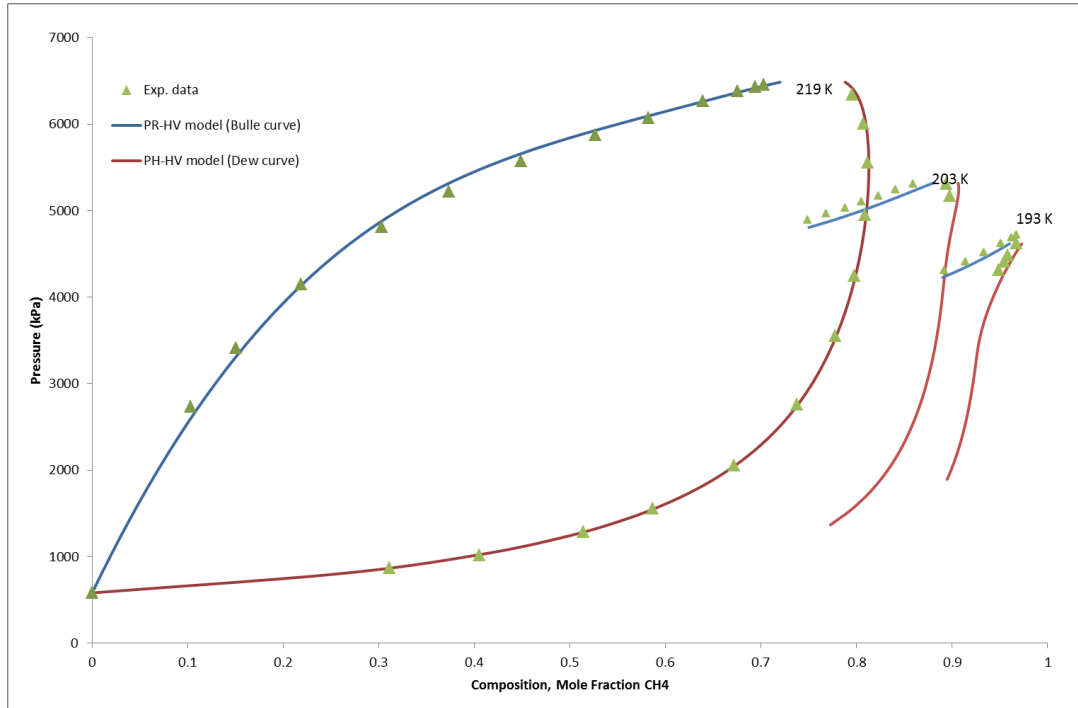


Figure 6 : Fluid phase diagrams of CH₄ + CO₂ system at low temperature [13]. Symbols are experimental data and lines are PR-HV model.

[Figure 6]

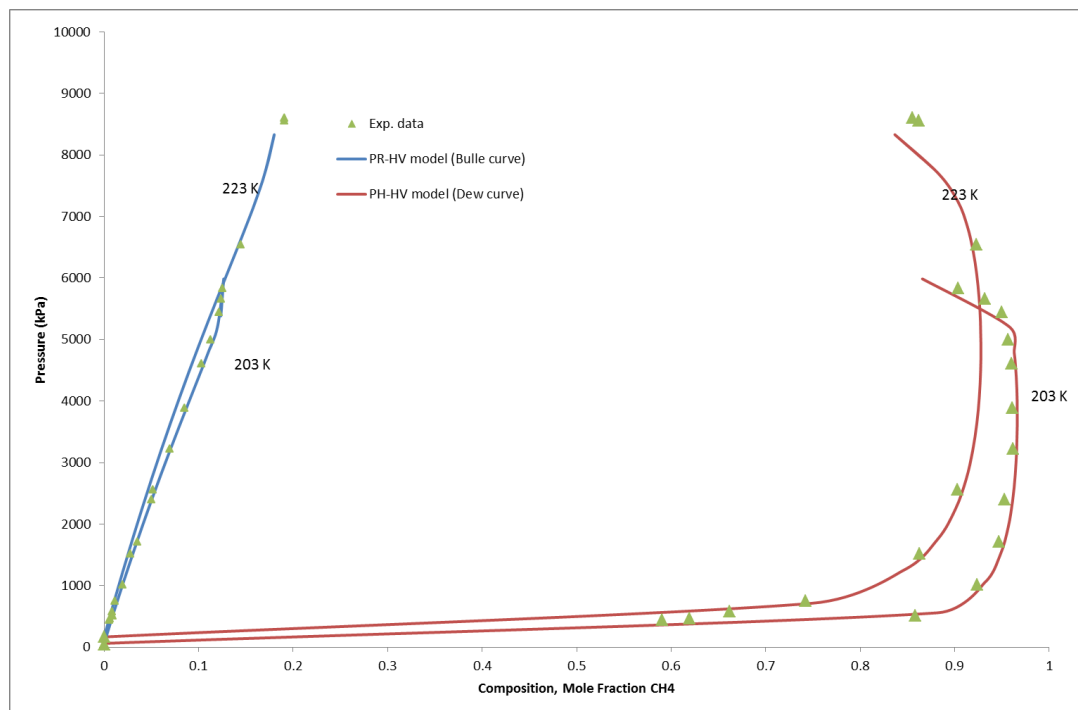


Figure 7 : Fluid phase diagrams of CH₄ + H₂S system at low temperature. Symbols are experimental data [12] and lines are PR-HV model.

[Figure 7]

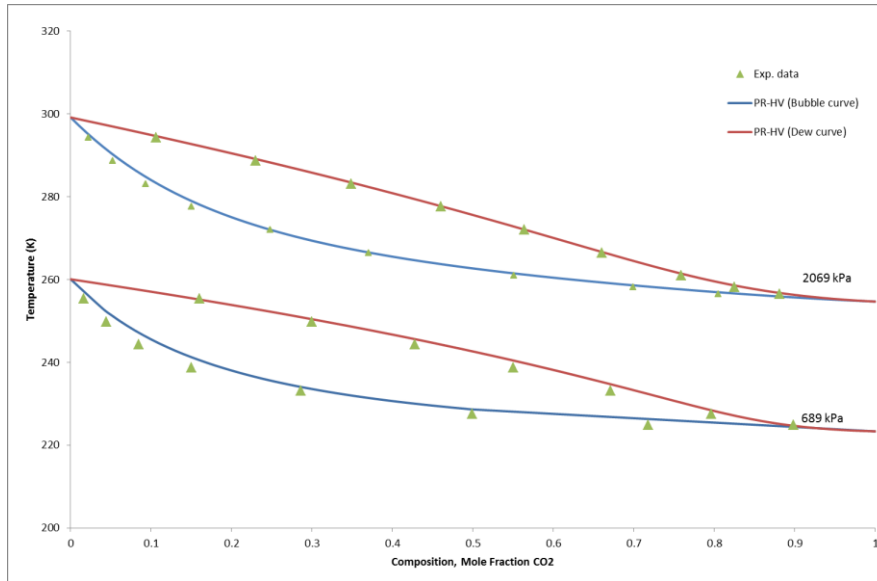


Figure 8 : Fluid phase diagrams of CO₂ + H₂S system at low temperature. Symbols are experimental data [14] and lines are PR-HV model.

[Figure 8]

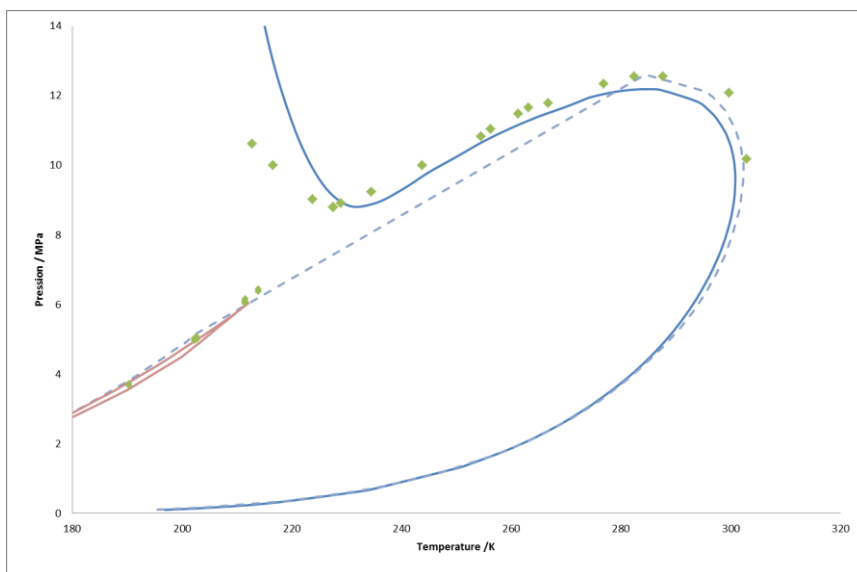


Figure 9 : Fluid phase diagrams of CH₄ + CO₂ + H₂S system at low temperature. Symbols are experimental data [15], solid lines are PR-HV model and dot line are PR model with conventional binary parameters [16].

[Figure 9]

Sensibility of the solid model

We can also observe that the prediction of the temperature of solid formation for the mixture without H₂S is not very good. Consequently, we have decided to change the model for the calculation of fugacity of the CO₂ at its solid state. Jäger and Span [17] has proposed in 2012 an EoS to compute directly the Gibbs energy of the CO₂ for temperature below its triple point temperature. The advantage is that we do not have to extrapolate the prediction of CO₂ liquid fugacity below CO₂ triple point temperature. Eq. (6) gives the expression of the solid CO₂ fugacity.

$$f_{\text{CO}_2}^{\text{S}}(T, P) = P^0 \exp\left(\frac{g_{\text{CO}_2}^{\text{S}}(T, P) - g_{\text{CO}_2}^0(T, P^0)}{RT}\right) \quad (6)$$

The reference state (0) for the Gibbs Energy model considered here is the ideal gas at $P^0 = 101325$ Pa and $T^0 = 150$ K. It is calculated using GERG 2008 EoS. The results are presented in Table 4 and Figure 5. We can conclude that it does not improve the prediction of the temperature of solid formation. Consequently, in order to improve the data treatment, we have decided to use the PR EoS for the liquid phase and to adjust the BIP using our experimental data. PR-HV model was consider to define the fluid state for each experimental data.

[Table 4]

Table 4: Prediction of temperature of solid formation by the four predictive models using Eq. 2 for the calculation of CO₂ fugacity at its solid state.

Mixture identification	Pressure	Fluid state	PPR78	PSRK UNIFAC	PR-HV NRTL-V	GERG 2008 (REFPROP)	GERG 2008 with Jäger and Span model
	P/ MPa		T/ K	T/ K	T/ K	T/ K	T/K
A	2.224	V	209.67	210.31	210.45	210.54	210.26
B	2.186	V	208.50	209.18	203.15	209.28	209.03
C	1.848	L	173.06	185.91	198.85	193.25	192.75
D	1.974	L	167.93	181.54	195.15	189.91	189.26
E	2.123	L	163.18	177.99	191.45	186.98	186.27
MAD/ K			17	9	0.9	4.2	4.5
MRD/ %			9	4.6	0.5	2.1	2.3

In conclusion, we can consider that there is no impact of the model considered for the solid phase but the selection of the equation of state for the fluid phase has a non-negligible impact.

Tentative of improvement of modelling by binary interaction parameter fitting for PR EoS

In this section we will use our new experimental data and adjust parameters of a cubic EoS used to calculate the fugacity of fluid phase. BIP were considered in the Peng Robinson EoS [18] used to compute the fugacity of the fluid phase and were optimized to minimize the deviation between experimental and calculated temperatures, using an algorithm based on simulated annealing methods (Salamon et al. [19] and van Laarhoven et al. [20]), which flowchart is presented in Figure 6. This algorithm was selected because of its ability to avoid getting stuck in local optima, thus to better explore the solutions space. Eq. (7) presents the selected objective function.

$$F = \sum_i^{N_{\text{data}}} (T_{\text{exp}} - T_{\text{cal}})^2 \quad (7)$$

The optimal values of the three BIP are presented in Table 5 and the resulting temperatures are presented in Table 6 and plotted in Figure 5.

[Table 5]

Table 5: Optimal BIP values

BIP	New value	Sandler [16]
$k_{\text{CO}_2\text{-H}_2\text{S}}$	0.1781	0.097
$k_{\text{CO}_2\text{-CH}_4}$	0.1212	0.09
$k_{\text{H}_2\text{S-CH}_4}$	0.0104	0.08

[Table 6]

Table 6: Results for EoS with optimal BIP

Mixture identification	PR EoS T /K
A	210.45
B	209.38
C	196.57
D	193.89
E	192.12
MAD/ K	1.7
MRD/ %	0.85

As expected the results are better after adjustment of BIP using our experimental data and close to the results obtained using the PR-HV EoS. We observe a strange calculation for the mixture B. Regarding the certificate given by AIR PRODUCTS concerning this chemical, we can see that the relative uncertainty on H₂S mole fraction is the most important. So we try to repeat the calculation by considering a new composition of the mixture. If we consider the mixture B with a new molar composition ($z_{\text{CO}_2} = 0.1887$, $z_{\text{H}_2\text{S}} = 0.0593$, $z_{\text{CH}_4} = 0.7520$), and if we consider that the fluid phase is a liquid phase, we calculate a temperature close to the temperature measured (201.21 K, see Figure 3). The deviation from original H₂S mole fraction is around 20% which seems to be very important. One reason of such deviation can be the inability of the PR EoS to generate accurate PT envelop for such mixture composition and consequently calculates temperature of solid formation with some deviation. This may explain why there is a non-negligible deviation between our experimental data and the model calculation also after adjustment of BIP. It is important to notice that the PR HV model predict very well the temperature of solid apparition. Certainly one reason is that the PR HV is more accurate to predict the phase diagram of the fluid phase (i.e. VLE) as it is composed with Huron Vidal mixing rules involving NRTL activity coefficient model. This point is confirmed by analysis of BIP values after their adjustment. In effect, as can be seen on Table 5, the values of the BIP are different to the values available in the literature (Sandler [16], values from DECHEMA) and one can notice their high values. It is due to the fact that classical mixing rule is not adapted for such condition of phase equilibrium. Also, we have presented in Figure 5 the results obtained using Sandler's parameters and they are very close to the values calculated using PSRK model. It confirms that PPR78 model prediction and the temperature dependency of the BIP of this model is not adapted for

such phase equilibrium calculation. Consequently, if classical mixing rule is selected, it may be wise to use BIP adjusted from vapor-liquid equilibrium data and to adjust BIP with a mixing rule on molar co-volume if users wants to remain constant the value of BIP for attractive parameter.

[Figure 10]

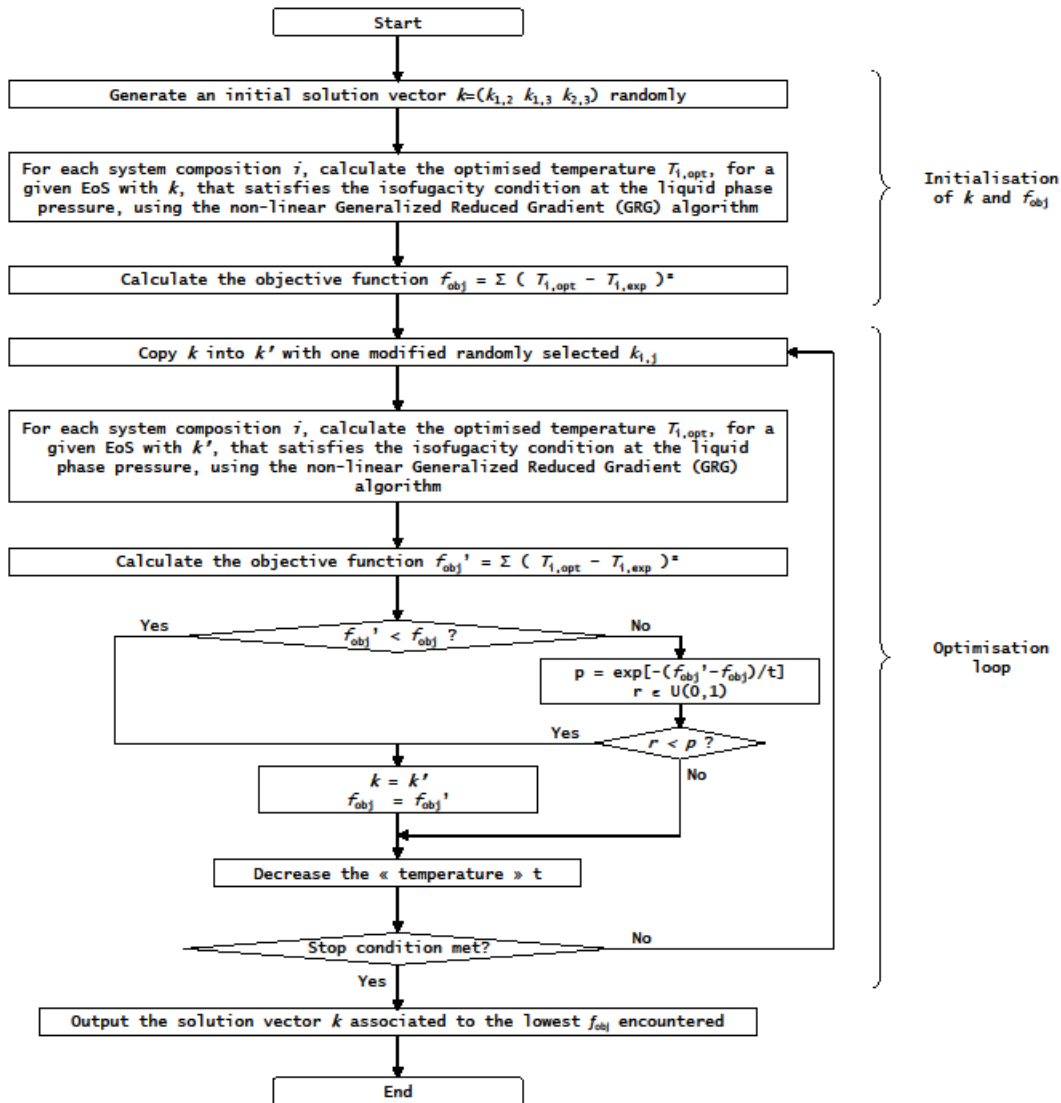


Figure 10 : Flowchart of the simulated annealing algorithm used to find the optimal 3-tuple of BIP

4. Conclusion

New temperatures of solid formation were measured for 5 mixtures composed with CO₂, H₂S and CH₄. The results seem to be self-coherent, and tend to show a good correlation between the temperature of solid phase formation and the ratio $\frac{n_{\text{H}_2\text{S}}}{n_{\text{CH}_4}}$ for a constant ratio $\frac{n_{\text{CO}_2}}{n_{\text{CH}_4}}$, and vice-versa. Two EoS based on Group Contribution approach (PPR78, PSRK) were tested to predicted the CO₂ freezing temperature and the results reveals that deviations are important when the mixture contains H₂S. We have also used a semi-empirical EoS developed for natural gas (GERG-2008 EoS) and the results are improved. The best results have been obtained with the PR-HV EoS which has been developed to well suit the behaviour of CH₄/CO₂/H₂S fluid phase equilibrium.

We have examined the influence of the solid fugacity model of the CO₂ with the Jäger and Span mode however the impact is negligible.

Adjustment of cubic EoS BIP leads to a better representation of the data but not for mixture B (calculations are strongly dependent on the mixture composition) but the values of BIP adjusted are very different from the one obtained by considering the VLE data. Strategy for SLE computation using cubic EoS for mixture has to be thought.

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Figure 5: Temperature of solid phase formation for different values of $\frac{n_{H_2S}}{n_{CH_4}}$.

Δ : our experimental data. \times : Calculated using PR EoS with Sandler [16] parameters. \blacklozenge : calculated data using PPR78 EoS. \blacktriangle : calculated data using PSRK UNIFAC EoS. \blacksquare : calculated data using PR-HV EoS. \bullet : calculated using GERG 2008 EoS with Eq. 3 for the calculation of solid CO₂ fugacity. o : calculated using GERG 2008 EoS with Jager and Span model [17] for the calculation of solid CO₂ fugacity. Green dotted line: calculation using Peng Robinson EoS with adjusted BIP. Black dashed line: calculation using Peng Robinson EoS with adjusted BIP and new composition of mixture B.

Figure 6: Fluid phase diagrams of CH₄ + CO₂ system at low temperature. Symbols are experimental data

[13] and lines are PR-HV model.

Figure 7: Fluid phase diagrams of CH₄ + H₂S system at low temperature. Symbols are experimental data [12] and lines are PR-HV model.

Figure 8 : Fluid phase diagrams of CO₂ + H₂S system at low temperature. Symbols are experimental data [14] and lines are PR-HV model.

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Figure 10 : Flowchart of the simulated annealing algorithm used to find the optimal 3-tuple of BIP