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**Spontaneous galvanic deposition of nanoporous Pd on microfibrous-structured  
Al-fibers for CO oxidative coupling to dimethyl oxalate**

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## Abstract

Nanoporous palladium (np-Pd) is deposited spontaneously on microfibrillar-structured Al-fibers (np-Pd/Al-fiber catalyst) via the galvanic replacement between  $\text{Pd}^{2+}$  and aluminium metal ( $2 \text{ Al} + 3 \text{ Pd}^{2+} \rightarrow 2 \text{ Al}^{3+} + 3 \text{ Pd}$ ). The np-Pd/Al-fiber catalyst combines the high activity and good stability of the np-Pd with the excellent heat conductivity and large permeability of the microfibrillar-structured Al-fibers. The np-Pd/Al-fiber catalyst shows four times higher intrinsic activity (i.e., turnover frequency) than the conventional Pd/ $\alpha$ - $\text{Al}_2\text{O}_3$  catalyst. The unique nanoporous palladium with interpenetrating network of nanoligaments is responsible for the improved activity in the CO oxidative coupling to dimethyl oxalate ( $2 \text{ CH}_3\text{ONO} + 2 \text{ CO} \rightarrow 2 \text{ NO} + \text{CH}_3\text{OCOCOOCH}_3$ ).

**Keywords:** nanoporous palladium; galvanic deposition; microfibrillar-structured catalyst; aluminium fiber; dimethyl oxalate.

## 1. Introduction

Nanoporous palladium (np-Pd) is of great fundamental and technological importance for wide applications in catalysis, energy storage, sensors and fuel cells [1-3]. Various unsupported np-Pd samples have been successfully prepared by dealloying, consolidation, combustion, hydrothermal growth and template synthesis [4,5]. The fabricated np-Pd possesses high surface-to-volume ratio, interpenetrating nanostructure and large characteristic lengths in nanoligaments [1]. Owing to the special surface effect and resistance to sintering, the np-Pd exhibits good activity and stability in heterogeneous reactions [2,3]. Nevertheless, the high cost of the unsupported np-Pd hinders its practical application, and the as-made powder cannot satisfy the requirements of conventional fixed-bed reactors.

Aluminium metal is inexpensive, highly thermal-conductive, and is so soft that it could be easily made into different microstructured packings such as fibers and foams. These packings are opening avenues to “ideal” catalytic reactors with advantages of high throughput, low pressure drop and improved heat/mass transfer [6,7]. However, the real application of the microstructured Al metal as reactor packings is hampered by the inert surface and low surface area [8]. Recently, boehmite nanosheets were in-situ generated on microfibrillar-structured Al-fibers by the hydrothermal oxidation ( $4 \text{ H}_2\text{O} + 2 \text{ Al} \rightarrow 3 \text{ H}_2 + 2 \text{ AlOOH}$ ), and then incipiently impregnated with palladium precursor to obtain the Pd catalyst [9]. Therefore, it is important to find a direct and an effective way to functionalize aluminium metal.

The vapor phase carbonylation of methyl nitrite (MN) to dimethyl oxalate (DMO), a process converting inorganic C1 to organic C2 plays a key role from syngas to ethylene glycol (product of DMO hydrogenation) [10,11]. This process involves two main reactions: (1) Pd-catalyzed CO

oxidative coupling:  $2 \text{CH}_3\text{ONO} + 2 \text{CO} \rightarrow 2 \text{NO} + \text{CH}_3\text{OCOCOOCH}_3$ ; and (2) MN regeneration without catalyst:  $2 \text{NO} + 2 \text{CH}_3\text{OH} + 1/2 \text{O}_2 \rightarrow 2 \text{CH}_3\text{ONO} + \text{H}_2\text{O}$ . Although this process has been applied at industrial scale, the commercial Pd/ $\alpha$ -Al<sub>2</sub>O<sub>3</sub> catalyst needs a high Pd content of 2 wt% (the-state-of-the-art) [12]. Decreasing the Pd content is an on-going effort which could be achieved by tuning the Pd shape/size and adding of a promoter on the CO oxidative coupling to DMO [13,14]. However, little attention is paid to the  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> support used, which suffers from poor mass/heat transfer and non-regular flow pattern for the strongly exothermic and fast reaction.

Herein, we report the preparation of a supported catalyst using nanoporous palladium embedded on microfibrus-structured Al-fibers. We also reveal the top-down, versatile and effective method of the catalyst fabrication through the spontaneous galvanic deposition between Pd<sup>2+</sup> and Al ( $2 \text{Al} + 3 \text{Pd}^{2+} \rightarrow 2 \text{Al}^{3+} + 3 \text{Pd}$ ). The np-Pd/Al-fiber catalyst combines the activity and stability of the np-Pd and improved heat/mass transfer, large permeability originated from the microfibrus-structured Al-fibers.

## 2. Experimental

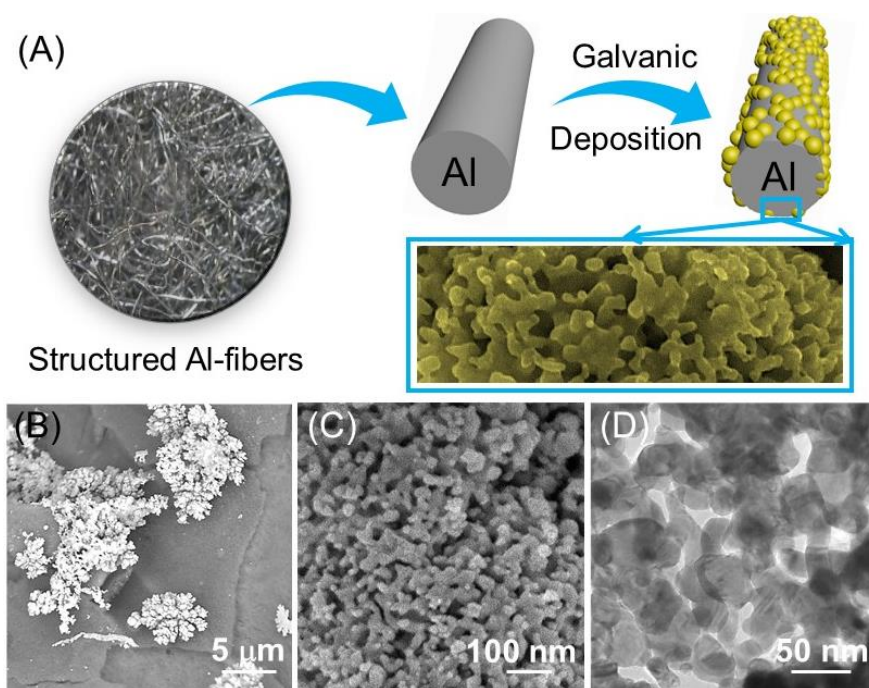
A large thin-sheet fibrofelt (2.0 m length  $\times$  1.0 m width  $\times$  ~1.3 mm thickness) with Al-fibers (60  $\mu\text{m}$  diameter, 99.9 wt% Al; Shanghai Xincal Network-microstructured Material Co. Ltd., China) was employed as pristine catalyst support [9]. Typically, circular shaped Al-fibers (16 mm diameter; Fig. S1) were tailored from the purchased fibrofelt and pre-treated with a NaOH solution (0.1 wt%) at 25 °C to remove the oxide layers (Al<sub>2</sub>O<sub>3</sub>) from the surface. After that, palladium was deposited onto the pre-treated Al-fibers with the aid of galvanic replacement between Pd<sup>2+</sup> and Al ( $2 \text{Al} + 3 \text{Pd}^{2+} \rightarrow 2 \text{Al}^{3+} + 3 \text{Pd}$ ). This reaction proceeded spontaneously at room temperature when the pre-treated Al-fibers were impregnated with an acetic acid (CH<sub>3</sub>COOH) solution containing appointed amount

of palladium acetate ( $\text{Pd}(\text{CH}_3\text{COO})_2$ ; Sinopharm Chemical Reagent Co. Ltd., China). During the galvanic process, the orange solution gradually became clear. The resulting Al-fibers were dried at  $100\text{ }^\circ\text{C}$  for 2 h, followed by calcining at  $300\text{ }^\circ\text{C}$  (air) for 2 h to obtain the np-Pd/Al-fiber catalyst. The nominal Pd content in the np-Pd/Al-fiber catalysts was in the range of 0.1 - 2.0 wt%. In addition, we prepared a reference catalyst of Pd/ $\alpha$ - $\text{Al}_2\text{O}_3$  (nominal Pd content of 0.3 wt%) using incipient-wetness impregnation with  $\text{Pd}(\text{CH}_3\text{COO})_2$  as a precursor and  $\alpha$ - $\text{Al}_2\text{O}_3$  (Alfa Aesar (China) Chemical Co. Ltd.) as a support, followed by drying at  $100\text{ }^\circ\text{C}$  for 2 h and calcination at  $300\text{ }^\circ\text{C}$  in air for 2 h. The details of catalyst characterizations and the calculations of the turnover frequency (TOF), conversion and selectivity are available in the Supporting Information (SI).

### 3. Results and discussion

The strategy for preparation of the microfibrillar-structured np-Pd/Al-fiber catalyst is schematically presented in Fig. 1A. The Al-fibers with irregular three-dimensional network and 90 % void volume, are used as pristine catalyst support (Fig. 1A and Fig. S1). The original surface of the Al-fibers is smooth, as shown in the SEM images (Fig. S2). After the spontaneous deposition carried out at room temperature through the galvanic replacement between  $\text{Pd}^{2+}$  and Al ( $2\text{ Al} + 3\text{ Pd}^{2+} \rightarrow 2\text{ Al}^{3+} + 3\text{ Pd}$ ), tree-like np-Pd is epitaxially formed on the Al-fibers surface (Fig. 1B). The circular shape of Al-fibers remains intact (Fig. S1). Both SEM and TEM images show that the np-Pd is composed of 10-30 nm ligaments surrounded by irregular nanopores having a wide size distribution (Fig. 1C,D). The EDX spectra of the np-Pd contain three peaks corresponding to Pd, Al and O with an average content of 90.8, 3.0 and 6.2 wt%, respectively (Fig. S3). The presence of Al and O is attributed to the Al substrate and the oxidation of the np-Pd and Al, respectively. The XRD patterns show that only the characteristic diffraction peaks of Al (JCPDS 04-0787) are identified in the

mother Al-fibers support (Fig. S4). For the np-Pd/Al-fiber catalyst, the diffraction peaks of Pd(200) (JCPDS 01-1201) and PdO(101) (JCPDS 88-2434) are weak, possibly due to the low Pd content (Fig. S4). The PdO phase is formed unavoidably due to the oxidation of the np-Pd during the calcination (see Experimental section), which is consistent with the EDX elemental mapping results (Fig. S3 and Fig. S4). Such PdO can be completely reduced below 25 °C due to the undetectable H<sub>2</sub> consumption above 25 °C studied by H<sub>2</sub>-temperature programmed reduction (Fig. S5). The nominal Pd content is 0.5 wt%, while the actual content is 0.28 wt% measured by ICP-AES analysis. The loss of Pd is occurred during the galvanic deposition; a visible black powder at the bottom of the galvanic solution was observed. The black powder has nanoporous structure similar to the np-Pd observed on the np-Pd/Al-fiber catalyst (Fig. 1C and Fig. S6). Medium dispersion (*D*) of Pd is measured to be 0.09 by CO pulse chemisorption measurements, and the particle size (*d*) is determined to be ~12 nm ( $d = 1.12 / D$  [15]), which is smaller than the size obtained by TEM (average particle size of 19 nm; Fig. 1D).



**Fig. 1.** (A) Strategy applied for preparation of the np-Pd/Al-fiber catalyst by spontaneous galvanic

deposition. (B-C) SEM and (D) TEM images of the np-Pd/Al-fiber catalyst (0.28 wt% Pd, ICP-AES).

As noted previously, the process of spontaneous galvanic deposition includes the following two steps: (1)  $\text{Al} \rightarrow \text{Al}^{3+} + 3 \text{e}^-$  and (2)  $\text{Pd}^{2+} + 2 \text{e}^- \rightarrow \text{Pd}$ . The galvanic deposition proceeds automatically at room temperature once the Al-fibers are impregnated with the  $\text{CH}_3\text{COOH}$  solution of  $\text{Pd}(\text{CH}_3\text{COO})_2$ . The driving force of this reaction is the large difference in electrode potential between  $\text{Al}^{3+}/\text{Al}^0$  (-1.66 V) and  $\text{Pd}^{2+}/\text{Pd}^0$  (0.91 V) pairs. Palladium is first nucleated and deposited at energy-favorable sites or dislocation centers on the Al-fibers surface [16]. With prolonged time of galvanic deposition, the palladium is epitaxially formed on the Al-fiber and assembled as np-Pd, which is confirmed by SEM-EDX analysis (Fig. S7).

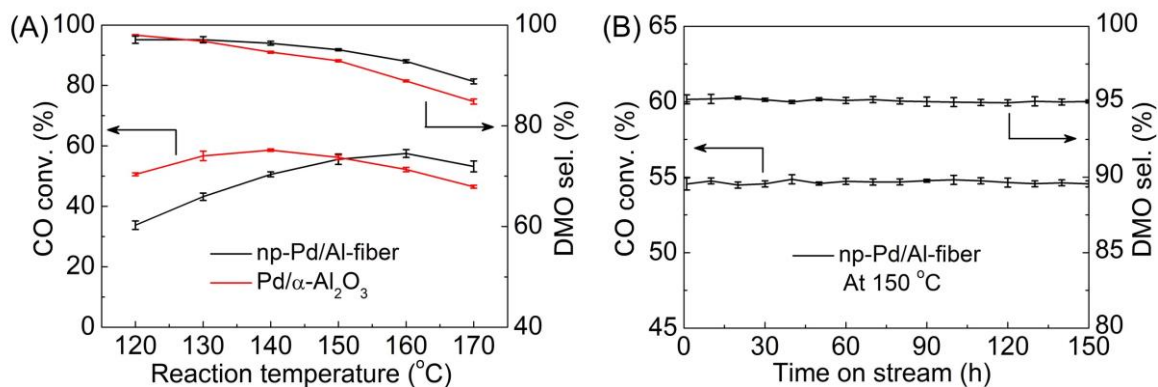
The effect of the nominal Pd content on the catalytic activity of the np-Pd/Al-fiber in the CO oxidative coupling to DMO was studied (Fig. S8). Clearly, with the increase of Pd content from 0.1 to 0.5 wt%, at 150 °C, the CO conversion is increased from 31.3% to 55.6% and gradually increased to 65.5% with further rising of the Pd content up to 2.0 wt%. While the DMO selectivity (~95%) is almost independent of the Pd content. Based on the balance between the CO conversion and the catalyst cost, the optimal Pd content of 0.5 wt% was identified, while the actual one is 0.28 wt%.

As shown in Fig. 2A, the np-Pd/Al-fiber catalyst shows higher DMO selectivity than the conventional Pd/ $\alpha$ - $\text{Al}_2\text{O}_3$  catalyst. Both catalysts show a volcano-type variation of CO conversion by elevating the reaction temperature from 120 to 170 °C. The np-Pd/Al-fiber catalyst reaches a maximal CO conversion at 160 °C while the Pd/ $\alpha$ - $\text{Al}_2\text{O}_3$  catalyst at 140 °C, which indirectly indicates the effect of the structure for np-Pd on Al-fibers and Pd nanoparticles on  $\alpha$ - $\text{Al}_2\text{O}_3$ . The actual Pd content and Pd oxidation state are similar for both catalysts, but Pd dispersion and surface Pd atoms



of the np-Pd/Al-fiber catalyst are much lower than for the Pd/ $\alpha$ -Al<sub>2</sub>O<sub>3</sub> (Table 1, XPS spectra in Fig. S9). However, the np-Pd/Al-fiber catalyst exhibits a TOF of 1.2 s<sup>-1</sup>, which is four times higher than for the Pd/ $\alpha$ -Al<sub>2</sub>O<sub>3</sub> (0.2 s<sup>-1</sup>), suggesting that the former possesses far larger intrinsic activity on the CO oxidative coupling to DMO. The np-Pd with well-connected network of nanoligaments, possesses nearly equivalent concave and convex surface (negative and positive curvature), thereby leading to lattice contraction and expansion based on the Laplace equation [17,18]. The HRTEM images show lattice fringes of 0.208-0.231 nm that correspond to Pd(111) facet on some disordered areas for the np-Pd/Al-fiber catalyst (Fig. S10); this result indicates that some lattices are expanded while some are contracted. Such lattice disorder of the np-Pd has a remarkable influence on the surface properties and adsorption performance, as confirmed by first-principles calculations [19,20]. It is thus inferred that the unique nanoporous structure of the np-Pd is possibly responsible for the enhanced intrinsic activity.

The np-Pd/Al-fiber catalyst exhibits a CO conversion of ~55%, and a DMO selectivity of ~95% for 150 hours test (Fig. 2B). The diffraction peak of Pd(200) (JCPDS 01-1201) is almost intact for the used catalyst after 150 hours test (XRD pattern in Fig. S11) suggesting the anti-sintering of the np-Pd. These results clearly indicate that the np-Pd/Al-fiber catalyst is very stable. Moreover, in comparison with the catalysts reported in the literature, the np-Pd/Al-fiber has low Pd content and medium catalytic performance (Table S1). However, the used support of our microfibrillar-structured Al-fiber possesses enhanced heat/mass transfer and low pressure drop compared to the common  $\alpha$ -Al<sub>2</sub>O<sub>3</sub> support.



**Fig. 2.** (A) Conversion and selectivity as a function of reaction temperature (np-Pd/Al-fiber: 0.28 wt% Pd, Pd/α-Al<sub>2</sub>O<sub>3</sub>: 0.26 wt% Pd) for the CO oxidative coupling to DMO. (B) Conversion and selectivity of the np-Pd/Al-fiber for 150 hours test. Reaction conditions: GHSV = 3000 mL g<sup>-1</sup> h<sup>-1</sup>, CO/CH<sub>3</sub>ONO/N<sub>2</sub> =14/10/76 (volume), 1 bar. Note that each point is an average of three measurements (Fig. 2A) and five measurements (Fig. 2B) performed.

**Table 1** Catalyst characteristics of the np-Pd/Al-fiber and Pd/α-Al<sub>2</sub>O<sub>3</sub> catalysts<sup>a</sup>.

Sample	<i>D</i> <sup>b</sup>	<i>d</i> (nm) <sup>b</sup>	Pd (wt%) <sup>c</sup>	Surface Pd (mol g <sub>cat.</sub> <sup>-1</sup> ) <sup>d</sup>	TOF (s <sup>-1</sup> ) <sup>e</sup>
np-Pd/Al-fiber	0.09	12	0.28	2.4 × 10 <sup>-6</sup>	1.2
Pd/α-Al <sub>2</sub> O <sub>3</sub>	0.60	2	0.26	1.5 × 10 <sup>-5</sup>	0.2

<sup>a</sup> Reaction conditions: 140 °C, GHSV = 3000 mL g<sup>-1</sup> h<sup>-1</sup>, CO/CH<sub>3</sub>ONO/N<sub>2</sub> =14/10/76 (volume), 1 bar.

<sup>b</sup> Dispersion (*D*) of Pd determined through pulse chemisorption of 10 vol% CO with a stoichiometry of CO/Pd = 1 and nanoparticle size (*d*, nm) obtained by the formula ( $d = 1.12 / D$ ) [15].

<sup>c</sup> Actual Pd content measured by ICP-AES analysis.

<sup>d</sup> Surface Pd atoms estimated from actual Pd content and Pd dispersion (*D*).

<sup>e</sup> TOF according to the CO conversion at 140 °C and surface Pd atoms.

#### 4. Conclusions

We report the preparation of nanoporous palladium (np-Pd) catalyst on microfibrillar-structured

Al-fibers for the CO oxidative coupling to DMO. The microfibrillar-structured np-Pd/Al-fiber catalyst from macro- to nanoscale is fabricated by a tailored spontaneous deposition via the galvanic replacement between Pd<sup>2+</sup> and Al ( $2 \text{ Al} + 3 \text{ Pd}^{2+} \rightarrow 2 \text{ Al}^{3+} + 3 \text{ Pd}$ ). For the catalyst with a low Pd content (0.28 wt%), a CO conversion of ~55% and a DMO selectivity of ~95% are obtained for a feedgas of CO/CH<sub>3</sub>ONO/N<sub>2</sub> (14/10/7.6, volume) at GHSV of 3000 mL g<sup>-1</sup> h<sup>-1</sup>. Notably, the catalyst provides stable activity and selectivity through the 150 hours lifetime test. The np-Pd/Al-fiber catalyst shows four times larger intrinsic activity (i.e., turnover frequency) than the Pd/α-Al<sub>2</sub>O<sub>3</sub> catalyst, which is attributed to the structure difference between the np-Pd with nanoligaments and the Pd nanoparticles present in the Pd/α-Al<sub>2</sub>O<sub>3</sub> catalyst. Moreover, compared to the catalysts reported in the literature, our np-Pd/Al-fiber catalyst has the unique structure of np-Pd with improved intrinsic activity, and shows enhanced heat/mass transfer, large permeability originated from the microfibrillar-structured Al-fibers.

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