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Thermodynamic properties of mixed-layer illite-smectite by calorimetric methods: Acquisition of the enthalpies of mixing of illite and smectite layers

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ABSTRACT

The stability of illite-smectite interstratified with respect to discrete illite and smectite minerals was investigated by measuring their thermodynamic properties. The standard thermodynamic properties of the illite-smectite ISCz-1 mineral (G , H , S , C_p , and V) were determined between 298.15 K and 375 K by using calorimetric methods. Moreover, the enthalpies of mixing between the illite and smectite layers were measured at 298.15 K by acid solution calorimetry from a complete series of illite-smectite interstratified minerals (Shinzan area, Japan). The measured values were slightly negative, with a minimum value of -3.7 kJ mol^{-1} . This contributed to the stability of the mixed-layer with respect to a mechanical mixture of illite and smectite. In addition, the model from Blanc et al. (2015) was implemented to estimate the thermodynamic properties of the interstratified illite-smectite ISCz-1 mineral. The predicted values were consistent with the experimental ones. However, the estimates were slightly improved by considering the thermodynamic properties of the mixture of the illite and smectite components and then adding the terms of energies of mixing. This could be confirmed by establishing the stability domains of ISCz-1 and those of the corresponding illite and smectite end-members according to Meunier and Velde's determination of the smectite to illite reaction pathways (Meunier and Velde, 1989).

Key words: illite-smectite, enthalpies, heat capacities, entropies, Gibbs free energies, calorimetry.

1. Introduction

Illite/smectite mixed layer clay minerals (I/S) are widely present in the natural environment and occur in various contexts, notably during diagenetic processes in sedimentary basins ([3,4], among others) or resulting from the hydrothermal alteration of volcanic or plutonic rocks ([5–7], among others). Understanding the processes associated with the transformation of smectite into illite/smectite and their consequences on the mechanical and geochemical behaviors of clayey formations is of great importance in some major application fields, such as oil exploration or nuclear waste disposal. For instance, this transformation modifies the porosities of oil and gas reservoirs and consequently the oil reserves. It was and is still widely used as a geothermometer for the early diagenesis stage (possibly combined with other tools such as the Na/K or Si geothermometer or stable isotope analyses) in order to assess the maturity, in terms of the P-T-X reaction pathway, of a given reservoir [8]. In addition, the smectite-to-illite conversion may generate overpressures in sedimentary basins due to the dehydration of smectite and the precipitation of secondary minerals [9–11] that can be problematic for the functioning of oil wells. Furthermore, according to Gaucher and Blanc's review [12], the transformation of smectite to illite-smectite and then illite could be possibly expected at cement/clay interfaces after long periods of reaction times within the framework of deep disposal. Alkaline disturbance will generate mineralogical transformations at the cement/clay interface, in particular that of bentonite, thus implying a decrease of the smectite content. The latter will affect the swelling properties and the retention capacities of the clayey barrier near the cement/clay interface. Such reactions have to be integrated in the geochemical calculations used for performance assessments.

The thermodynamic stability of illite/smectite remains an unsolved issue. Numerous studies based on the mineralogical characterization of clays in natural systems ([13–18], among others) and in laboratory experiments ([19–23], among others) were performed in order to better understand the reaction mechanisms of the smectite illitization during hydrothermal alteration and diagenetic processes. Several authors ([19] and references therein, [21–23]) showed the coexistence of several phases (smectite, illite and one or more intermediate illite/smectite phases) during the transformation of smectite to illite.

By collecting the chemical analyses of natural I/S samples, Meunier and Velde [2] described the interstratified minerals using a ternary solid-solution model with one illite and two montmorillonite end-

members. Based on this approach, Blanc et al. [24] developed a thermodynamic model that allowed for the calculation of the energies of the mixing of illite and smectite layers depending on the degree of ordering of the stacking sequences. The resulting Gibbs free energies of mixing are then slightly negative, thus suggesting the stability of illite-smectite mixed-layers with respect to the mechanical mixture of the pure illite and smectite layers.

The present work focuses on the measurement of thermodynamic properties of I/S minerals with two main objectives.

- To acquire a complete set of thermodynamic properties (G , H , S , and C_p) for an illite-smectite mineral (ISCz-1), based on calorimetric methods ([25,26]), which could be used to supply the thermodynamic databases (e.g., ThermoChimie [27] and Thermoddem [28]) used for geochemical modeling calculations. Until now, the lack of thermodynamic data for illite/smectite mixed layers is a limitation to properly model the smectite-to-illite conversion.
- To acquire the enthalpies of the mixing of the illite-smectite layers for a series of interstratified illite-smectite (hydrothermal alteration minerals coming from the Shinzan Area; [7,29]).

In addition, the present work allows for the checking of the reliability of the predictive model from Blanc et al. [1] for the thermodynamic properties of interstratified illite-smectite. Based on the properties measured here, the influence of the thermodynamic parameters of mixing can be discussed, including the configurational entropy associated with the cationic disordering in crystallographic sites of clay minerals and the specific interactions between smectite and illite layers within the I/S stacking sequence.

2. Materials and methods

2.1. Clay samples

The ISCz-1 sample is a mixed-layer illite-smectite, originating from Slovakia. It was provided by the Source Clays Repository of the Clay Minerals Society (Chipera and Bish [30]).

After sieving the ISCz-1 sample to a size of $< 20 \mu\text{m}$, the $< 2 \mu\text{m}$ fraction was separated by centrifugation. This was then twice dispersed in CaCl_2 0.5 M solution for 24 h and centrifuged. Therefore, the sample was washed with a CaCl_2 0.001 M solution by using a dialysis bag for 3 days. In the present study,

Ca^{2+} was preferred to Na^+ for the saturation of the smectitic interlayers in order to facilitate the washing of the clay sample by avoiding the osmotic swelling of smectite. Eventually, the sample was dried at 40°C and then 60°C for several days. In the following, the sample referred to as ISCz-1 represents this treated sample.

A series of fourteen illite-smectite mixed-layers samples coming from the Shinzan area in northeast Japan was also used for this study. This series, which is well described in the literature (Inoue et al. [7,29]), is noticeable since it consists of a complete mixed-layer series with chemical compositions that vary from montmorillonite to illite end-members. The series results from the hydrothermal alteration of dioctahedral micas/smectites. The continuous conversion of smectite into illite was described by Inoue et al. [29]. The main transformation mechanism suggested by the authors was based on the dissolution of smectite and the crystallization of illite. A solid-state transformation mechanism was also shown from high resolution transmission electron microscopy (HRTEM) observations ([31,32]).

2.2. Analysis and characterization of the samples

2.2.1. ISCz-1 sample

The chemical analyses of Si, Al, Ti, Fe (total), Mn, Ca, Mg, K, Na and P were performed using X-ray fluorescence spectrometry. The results of these analyses are presented in Table 1. The amounts of total carbon and total sulfur were determined by infrared spectroscopy after burning the samples at 900°C in an oxygen atmosphere. The amount of Fe^{2+} was determined by ^{57}Fe Mössbauer spectrometry, as described further in the text (Section 2.2.1).

The chemical composition of the ISCz-1 sample (Ca-saturated) differs slightly from that of the previous Na-saturated ISCz-1 sample that was used for the heat capacity measurements in Gailhanou et al. [33] (Table 1). Both samples were supplied by the Clay Minerals Society with several years of intervals. The following study of the thermodynamic properties of the present ISCz-1 sample is complementary to the previous work from Gailhanou et al. [33].

Table 1. Chemical analyses of the sample ISCz-1 (< 2 μ m clay fraction). Mass fractions w_i , XRF = X-ray fluorescence spectrometry, ^{57}Fe Möss. = ^{57}Fe Mössbauer spectrometry, IR = infra-red spectroscopy.

	ISCz-1 (this study) /10 ² w_i per wt. %	Na-sat. ISCz-1 ([33]) /10 ² w_i per wt. %	Analytical technique
SiO ₂	53.8 ^a	52.6	XRF
Al ₂ O ₃	24.4 ^a	26.7	XRF
CaO	1.1 ^a	< 0.1	XRF
Fe ₂ O ₃	1.39 ^a	0.97	XRF, ^{57}Fe Möss.
FeO	0.17 ^a	0.30	^{57}Fe Möss.
K ₂ O	5.07 ^a	5.96	XRF
MgO	2.2 ^a	2.1	XRF
Na ₂ O	< 0.2	1.0	XRF
TiO ₂	0.13 ^a	0.09	XRF
P ₂ O ₅	< 0.05	0.13	XRF
Loss on ignition	11.9	10.0	Weight loss (up to 1000°C)
C total	0.07	0.04	IR
S total	< 0.01	0.06	IR
Total	100.25	99.95	

^a The relative expanded uncertainties $U_i(w_i)$ are two standard deviations of the mean with level of confidence 0.95 and are about $U_i(w_i) = 0.02$.

The cationic exchange capacity (CEC) of the sample was measured after saturating the sample with a chloride cobalt hexammine solution according to the method described in Gaucher et al. [34]. The total CEC was measured by colorimetry, and the exchangeable cations (Na⁺, Ca²⁺, K⁺, Mg²⁺) were analyzed by ionic chromatography. The results are provided in Table 2. The obtained CEC value is consistent with the one measured for the ISCz-1 sample at 39.0 meq.100 g⁻¹ [33].

Table 2. Equivalent fractions of exchangeable cations E_{Na} , E_{Ca} , E_{Mg} , E_K in meq/100 g of dry sample and cationic exchange capacity calculated $CEC_{calc.} = E_{Na} + E_{Ca} + E_{Mg} + E_K$ and measured by colorimetry $CEC_{meas.}$ for the ISCz-1 sample (< 2 μ m clay fraction) in meq per 100 g of dry sample.

Sample	E_{Na}	E_{Ca}	E_{Mg}	E_K	$CEC_{calc.}$	$CEC_{meas.}$
ISCz-1	0.4 ^a	39.7 ^a	< 0.5	3.8 ^a	43.9 ^a	44.1 ^a

^a The relative expanded uncertainties for the equivalent fractions of exchangeable cations and for CEC values are about $U_i(E_i) = 0.02$ and $U_i(CEC) = 0.02$ with level of confidence 0.95 ($k = 2$).

Moreover, XRD analyses were conducted on the ISCz-1 sample on powders and on oriented deposits using the method described by Gailhanou et al. [33]. For the powdered samples, both the identification of the

nature of impurities and their semi-quantification were performed using Modpattern [35] with an internal corundum standard (Fig. 1). In the present case, only rough estimates of the proportions of the impurities could be obtained by modeling due to the lack of a reference pattern for Ca-saturated illite-smectite.

Moreover, the modeling of the XRD pattern on the glycolated oriented deposit was performed by using the Arquant code [36], which allowed for the identification and quantification of a small amount of kaolinite. Finally, the amounts of impurities in the purified sample estimated by XRD modeling are 4.2, 3.8 and 0.45 wt.% of quartz, K-feldspar and kaolinite, respectively. Moreover, as a result of the modeling of the XRD pattern on the glycolated deposit and in agreement with Chipera and Bish's results [30], the interstratified ISCz-1 was ordered with a degree of interstratification ordering $R = 1$ (Reichweite index) and the illite layer content was estimated at 70%.

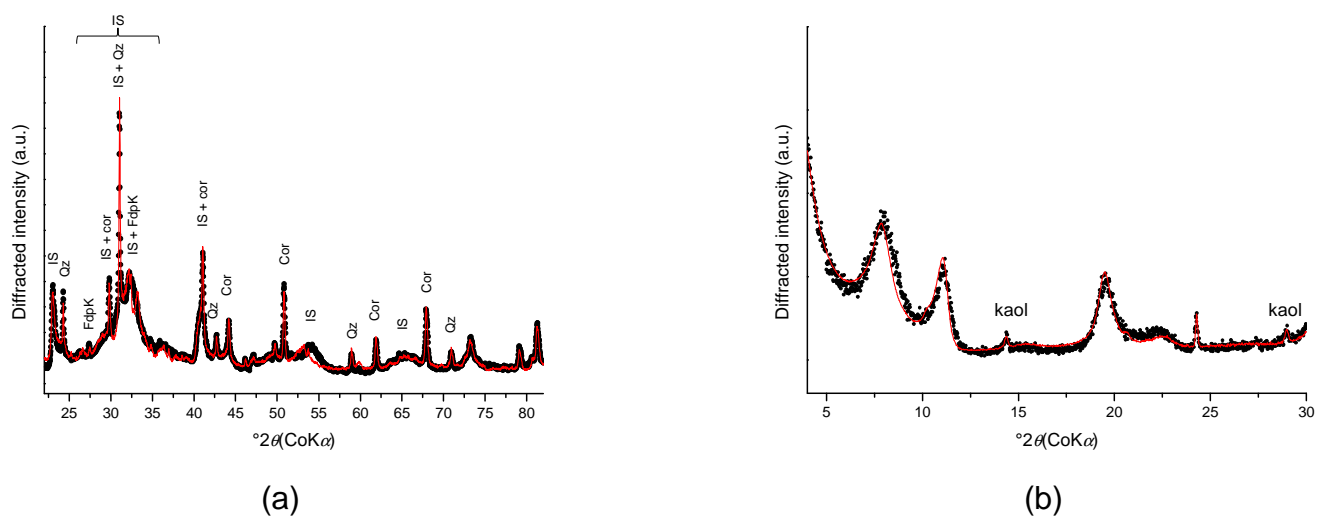


Fig. 1. Modeled XRD patterns obtained for the ISCz-1 sample ($< 2 \mu\text{m}$, Ca-saturated) on (a) the powdered sample and (b) the glycolated oriented deposit. IS = mixed-layer ISCz-1, Qz = quartz, Cor = corundum (internal standard of 9.19 wt.%), FdpK = K-feldspar, and Kaol = Kaolinite (in black, experimental; in red, modeling).

In addition, some investigations were performed by using ^{29}Si and ^{27}Al NMR in order to complete the characterization of the silicates and aluminosilicates impurities. A ^{29}Si NMR spectrum was acquired with a

Bruker Advance 300 MHz ($B_0 = 7.0$ T) spectrometer operating at a 79.5 MHz frequency for ^{29}Si (Fig. 2a). Notably, this analysis revealed, in addition to illite-smectite component (between -80 and -100 ppm), the presence of quartz (-107.52 ppm) and the absence of other silica impurities (cristobalite or amorphous silica, notably). The amount of quartz was estimated to 5.4 wt.% in the ISCz-1 sample. This result is in good agreement with the XRD estimate of 4.2 wt.%.

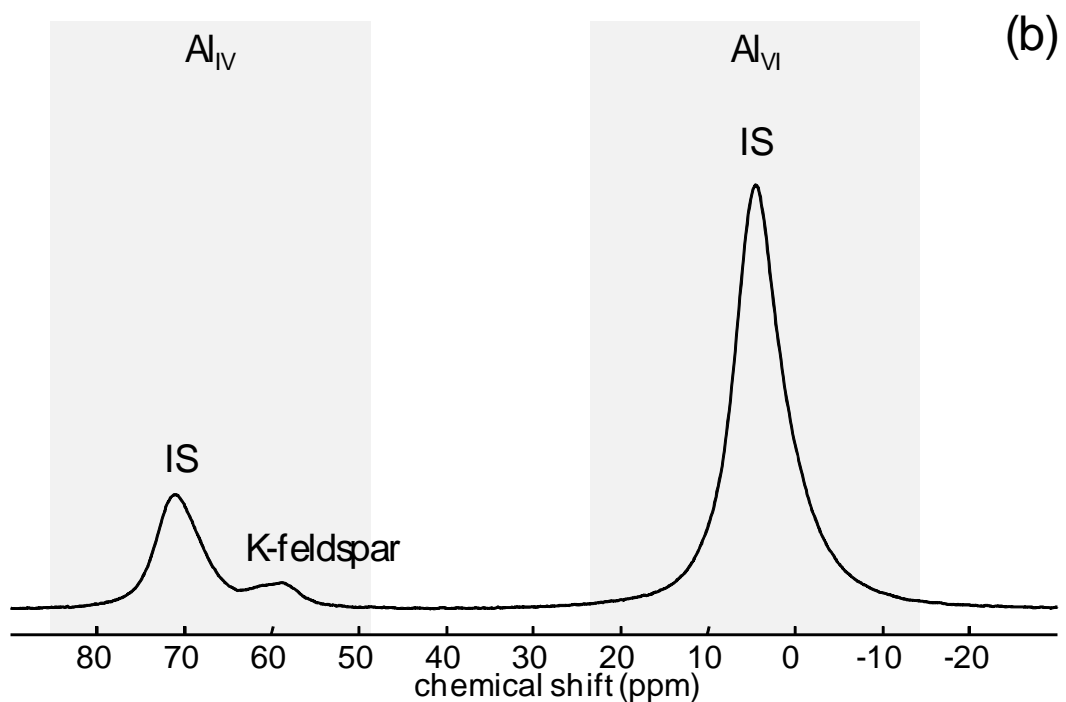
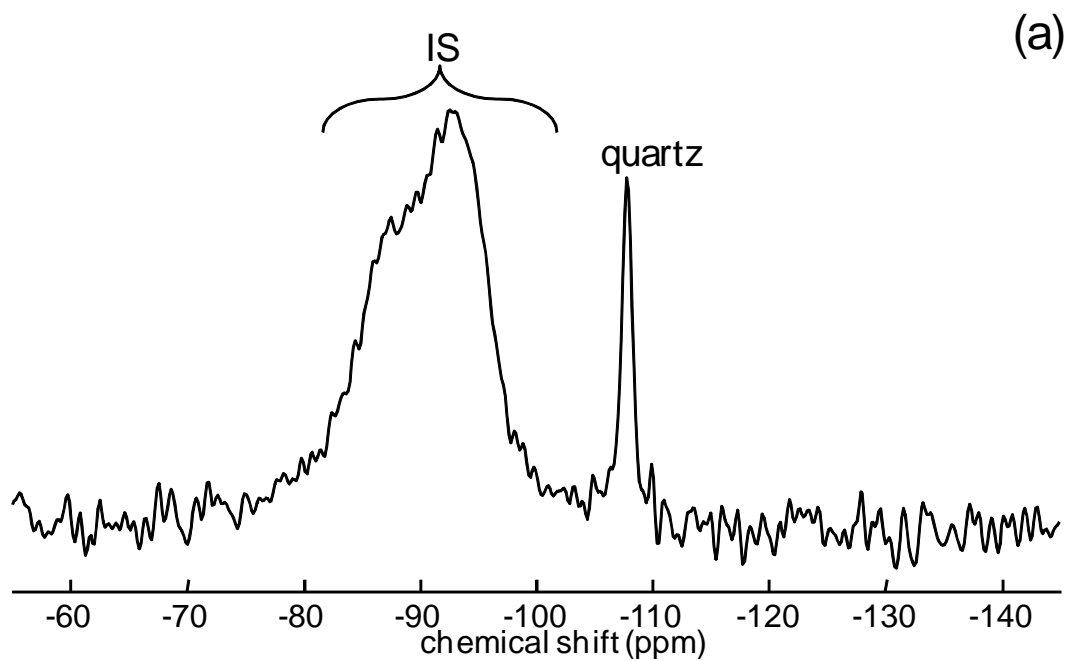


Fig. 2. Identification and quantification of impurities in the ISCz-1 sample (< 2 μm clay fraction) by NMR: (a) ^{29}Si and (b) ^{27}Al NMR spectra. Note: 1 ppm = 10^{-6} .

For its part, the ^{27}Al NMR spectrum (Fig. 2b) acquired with a Bruker Advance 750 MHz spectrometer ($B_0 = 17.7\text{ T}$, $\nu_{\text{RF}}(^{27}\text{Al}) = 195.5\text{ MHz}$) exhibits three main signals for one six-fold coordinated aluminum ($\delta_{\text{CG}} = 4.6\text{ ppm}$) and two four-fold coordinated aluminum ($\delta_{\text{CG}} = 59.4$ and 71.1 ppm). Both $\delta_{\text{CG}} = 4.6$ and 71.1 ppm sites are commonly attributed to Al^{VI} and Al^{IV} in illite-smectite [37,38], while the signal at $\delta_{\text{CG}} = 59.4\text{ ppm}$ is compatible with Al^{IV} in K-feldspar as detected by XRD (Cuadros and Linares [37]). The spectrum was modeled using the Dmfit software [39]. The component at $\delta_{\text{CG}} = 59.4\text{ ppm}$ corresponds to 2.9 at % of the aluminum of the sample. The resulting proportion of K-feldspar impurities in the ISCz-1 sample was then estimated to be 3.8 wt.%, which is in agreement with the value obtained by XRD modeling.

The proportions of ferric and ferrous ions in the clay minerals and, if present, amounts of ferrous or ferric impurities can be determined by ^{57}Fe Mössbauer spectrometry. These analyses were performed using a conventional constant acceleration Mössbauer spectrometer in transmission geometry that was equipped with a $^{57}\text{Co}(\text{Rh})$ source. The isomer shift values were corrected based on the calibration of the velocity scale made from $\alpha\text{-Fe}$ at 300 K. The amount of Fe in the sample was approximately 5 mg cm^{-2} . The spectra obtained at 300 K and 77 K were fitted using a discrete number of independent quadrupolar doublets and magnetic sextets composed of lorentzian lines, where the line width at half-height Γ (mm s^{-1}), the isomer shift δ (mm s^{-1}) and the quadrupole splitting ΔE_{Q} (mm s^{-1}) were refined using a least-square fitting procedure.

The ^{57}Fe Mössbauer spectra for the ISCz-1 sample (< 2 μm) are displayed in Fig. 3. The 300 K spectrum (inset Fig. 3) exhibits a quadrupolar broadened asymmetrical feature and a single line that can be described using two components that are clearly attributed to ferric and ferrous species. However, one can see that the baseline remains distorted (as confirmed by a spectrum obtained at 12 mm s^{-1} , which is not shown here). This feature results from the presence of superparamagnetic Fe species, which is the presence of ultra-fine Fe-bearing grains. Consequently, the hyperfine structure was investigated at 77 K to get the blocked magnetic states of some Fe species, as confirmed in Fig. 3. The description unambiguously requires

several magnetic sextets and quadrupolar doublets, but a single broadened line was added to better describe the baseline. The refined values of the Mössbauer parameters that were extracted from spectral modeling are listed in Table 3. The proportion of Fe in the clay mineral was estimated to be 42 at.% of the total Fe content and the ratio $\text{Fe}^{3+}/(\text{Fe}^{2+} + \text{Fe}^{3+})$ was estimated to be 72 %. It is important to note the good agreement between the Fe^{2+} contents estimated at 300 K and 77 K, assuming that the proportions are proportional to the relative absorption areas, i.e. assuming the same values of the Lamb-Mössbauer recoilless factors. The modeling allowed one to distinguish three magnetic phases, i.e., goethite ($\alpha\text{-Fe}^{3+}\text{O}(\text{OH})$) as fine grains (magnetic sextets), ultra-fine superparamagnetic grains (single line) and likely bernalite (ideally $\text{Fe}^{3+}(\text{OH})_3$, [40]) or a bernalite/hematite mixture.

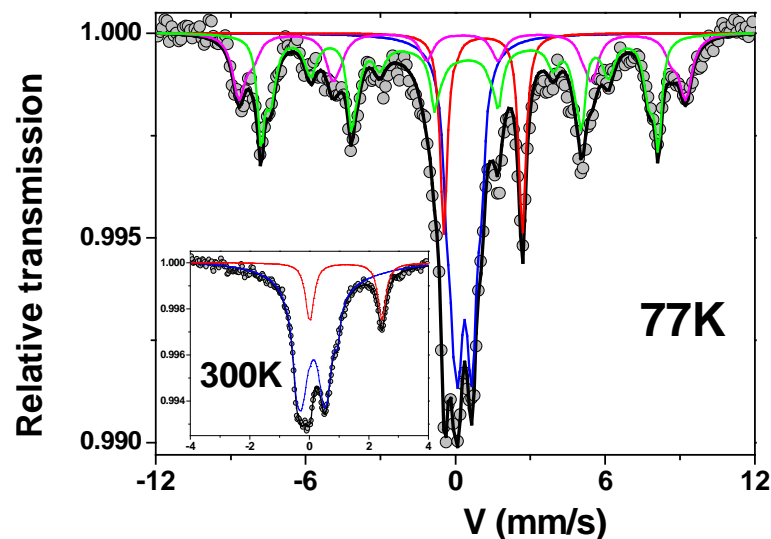


Fig. 3. ^{57}Fe Mössbauer spectra for the ISCz-1 sample ($< 2 \mu\text{m}$ clay fraction). The 77 K spectrum and its decomposition into 4 components: red for Fe^{2+} in clay, dark blue for Fe^{3+} in clay, green for the magnetic component of Fe^{3+} in goethite, and pink for the magnetic component of Fe^{3+} in bernalite. The inset represents the 300 K spectrum with the blue and red components attributed to ferric and ferrous species, respectively (see text).

Table 3. ^{57}Fe Mössbauer parameters at 77 K for the ISCz-1 sample (< 2 μm clay fraction): isomer shift value relative to that of the $\alpha\text{-Fe}$ at 300 K *I.S.* (mm s^{-1}), quadrupolar shift 2ϵ or quadrupolar splitting ΔE_Q (mm s^{-1}), magnetic field B_{hf} (T) and mole fraction of Fe for each component X_c .

Mineral	<i>I.S.</i> ^a	$2\epsilon^a$ or ΔE_Q^a	B_{hf}^a	$10^2 X_c^a$	Attributions
Illite-smectite	0.47	0.57		27	Fe^{3+} associated with clay minerals (structural and/or sorbed ions)
	0.42	1.32		4	Fe^{3+} associated with clay minerals
Illite-smectite	1.30	3.06		12	Fe^{2+} associated with clay minerals
Goethite	0.46	-0.27	49.2	28	Fe^{3+} in goethite
	0.47	-0.27	46.7		
	0.47	-0.27	36.8		
Bernalite	0.40	0.00	51.8	17	Fe^{3+} in bernalite
	0.40	0.00	55.3		
Goethite (ultra-fine grains)	0.46	0.00		12	Fe^{3+} in goethite

^a The expanded uncertainties with level of confidence 0.95 ($k = 2$) are about $U(I.S.) = 0.02 \text{ mm s}^{-1}$, $U(2\epsilon) = 0.01 \text{ mm s}^{-1}$, $U(\Delta E_Q) = 0.01 \text{ mm s}^{-1}$, $U(B_{hf}) = 0.5 \text{ T}$ and $U(X_c) = 0.02$.

Thus, based on these results and considering the total amount of Fe_2O_3 in the sample (Table 1), the amounts of goethite and bernalite were estimated to 0.70 wt.% and 0.36 wt.%, respectively.

From the preceding results, the final amounts of impurities were calculated and are reported in Table 4. The mean structural formula of illite-smectite was eventually calculated (Table 4) by subtracting the contributions of impurities that were obtained at ambient R.H. to the global chemical analyses from Table 1.

Table 4. Mass fractions of impurities in the ISCz-1 sample at ambient relative humidity (R.H.) $w_{i,r}$ and in dry sample $w_{i,dry}$, mean structural formula and molar mass of illite-smectite M .

Impurity	Mass fraction at ambient R.H.	Mass fraction in dry sample	Analysis method
	$10^2 w_{i,r}$	$10^2 w_{i,dry}^a$	
Kaolinite	0.45	0.49	XRD
Quartz	5.40	5.85	^{29}Si NMR
K-feldspar (orthoclase)	3.80	4.12	^{27}Al NMR
goethite	0.70	0.76	^{57}Fe Mössbauer
bernalite	0.36	0.39	^{57}Fe Mössbauer
Sample	Mean structural formula		$M/\text{g mol}^{-1}$
ISCz-1 (this work)	$(\text{Ca}_{0.092}\text{K}_{0.440})(\text{Si}_{3.562}\text{Al}_{0.438})(\text{Al}_{1.721}\text{Mg}_{0.255}\text{Fe}^{3+}_{0.029}\text{Fe}^{2+}_{0.011}\text{Ti}_{0.008})\text{O}_{10}(\text{OH})_2$		382.01
ISCz-1 (Gailhanou, 2007)	$(\text{Na}_{0.134}\text{K}_{0.530})(\text{Si}_{3.565}\text{Al}_{0.435})(\text{Al}_{1.717}\text{Mg}_{0.218}\text{Fe}^{3+}_{0.050}\text{Fe}^{2+}_{0.017})\text{O}_{10}(\text{OH})_2$		385.04

^a Calculated from the mass fractions at ambient R.H. by removing the amount of molecular water from the total mass of the samples.

Finally, the mineralogical and chemical compositions of the present ISCz-1 sample were compared to those obtained for the previous Na-saturated ISCz-1 sample, which had been used for heat capacity measurements in Gailhanou et al. [33]. The amounts of impurities are significantly higher in the present ISCz-1 sample than in the previous one, which contained approximately 0.75 wt.% quartz and 1.5 wt.% kaolinite. Moreover, the mean structural formula of the Ca-saturated ISCz-1 sample (this study) exhibits a slightly lower total layer charge (0.62 eq./O₁₀(OH)₂) compared with the one previously obtained for the Na-saturated ISCz-1 sample at 0.66 eq./O₁₀(OH)₂. The decrease of the layer charge is probably correlated to the increase of the total CEC, which were 0.44 and 0.39 eq. kg⁻¹ for the present Ca-saturated ISCz-1 sample and the previous Na-saturated ISCz-1 sample, respectively. Both indicated amounts of smectite layers slightly higher in the present case than for the previously sample analyzed.

2.2.2. *Illite-smectite series from the Shinzan Area*

The illite/smectite samples cover the entire range of compositions from 0 % to 100 % smectite (Table 5) as determined by the X-ray diffraction analyses of ethylene-glycol solvated samples (Inoue et al. [29]). According to these authors, samples containing more than 50 % smectite layers were randomly interstratified (Reichweite index $g = 0$), while the degrees of interstratification ordering increased progressively to $g = 1$, $g = 2$ and $g \geq 3$ with increasing illite contents.

The chemical compositions of all studied samples were determined by atomic absorption spectroscopy [41]. The mean structural formulae for the illite/smectite samples are given in Table 6.

Table 5. Illite/smectite interstratified minerals from the Shinzan hydrothermal alteration area in Japan. Drill holes, depth (m), proportions of smectite layers respectively to the total number of layers, illite N_{il} and smectite N_{sm} , and g = Reichweite index, according to [29].

Sample	Drill Hole	Depth z/m	$10^2 N_{sm}/(N_{sm} + N_{il})$	g
A (smectite)	Ws-2	183,5	100	0
B	Ws-4	329,5	85	0
C	Ws-4	383,0	73	0
D	Ws-2	377,8	70	0
E	Ws-2	407,4	68	0
F	Ws-4	392,8	50	0
G	Ws-2	423,2	40	1
H	Ws-10	199,9	35	1
I	Ws-4	440,6	30	1
J	Ws-5	240,1	20	2
K	Ws-7	123,6	15	≥ 3
L	Ws-7	115,5	15	≥ 3
M	Ws-7	197,0	5	≥ 3
N (illite)	Ws-8	420,9	0	

Table 6. Mean structural formulae of the illite/smectite mixed-layers from the Shinzan area on the basis of 12 oxygen atoms, calculated from the chemical analyses in [41]. Fe^{3+} and Ti were assumed to be in octahedral sheets.

	A	B	C	D	E	F	G	H	I	J	K	L	M	N
Si	3.886	3.822	3.786	3.714	3.692	3.728	3.586	3.524	3.560	3.483	3.542	3.473	3.397	3.329
Al	1.685	1.658	1.722	1.847	1.908	1.738	2.079	2.059	2.036	2.268	2.101	2.214	2.316	2.401
Fe^{3+}	0.124	0.116	0.096	0.086	0.069	0.130	0.078	0.070	0.096	0.033	0.017	0.029	0.033	0.065
Fe^{2+}	0.020	0.035	0.025	0.025	0.042	0.037	0.019	0.038	0.028	0.040	0.017	0.023	0.029	0.055
Mg	0.314	0.397	0.390	0.340	0.304	0.388	0.258	0.314	0.301	0.190	0.332	0.264	0.231	0.171
Ti	0.012	0.006	0.009	0.020	0.011	0.005	0.003	0.045	0.009	0.008	0.015	0.018	0.003	0.005
Mn	0.000	0.000	0.000	0.000	0.000	0.001	0.000	0.000	0.000	0.001	0.002	0.003	0.007	0.001
Ca	0.072	0.063	0.049	0.022	0.025	0.038	0.029	0.027	0.026	0.023	0.013	0.009	0.022	0.001
Na	0.162	0.223	0.211	0.258	0.226	0.152	0.208	0.101	0.103	0.004	0.004	0.004	0.005	0.011
K	0.008	0.148	0.222	0.235	0.292	0.388	0.353	0.481	0.516	0.621	0.686	0.702	0.769	0.802

A statistical analysis of the chemical compositions of the I/S minerals was performed using the Principal Component Analysis method (XLSTAT[®]). The main trend of chemical dispersion is described by the eigenvector F_1 and corresponds to the substitution $Si \rightarrow KAl$ that occurs during the transformation of montmorillonite to illite (Fig. 4). A second, less important trend seems to emerge along the F_2 axis. Regarding the composition of the eigenvector F_2 , it represents the chemical dispersions for the elements Si, Mg and Al and shows a correlation between Si and Mg, whereas these latter ones are anti-correlated with Al. Thus, the dispersion of compositions along the F_2 axis is probably associated with the Tschermak substitution $Al^{IV}Al^{VI} \rightarrow Si^{IV}Mg$, which corresponds to muscovite \rightarrow celadonite. Finally, the main chemical dispersion occurs along

the montmorillonite A – illite N axis, as expected. Nevertheless, because some compositions deviate from this axis, it is also necessary to take into account a supplementary dimension, along the F_2 axis in order to describe all the compositions of the I/S minerals. In the following section, we propose to describe the compositions of these illite-smectite series using a ternary solid-solution that is composed of illite N, montmorillonite A and interstratified F end-members. The interstratified F end-member was selected since it exhibits the stronger celadonite-type component along F_2 (Fig. 4).

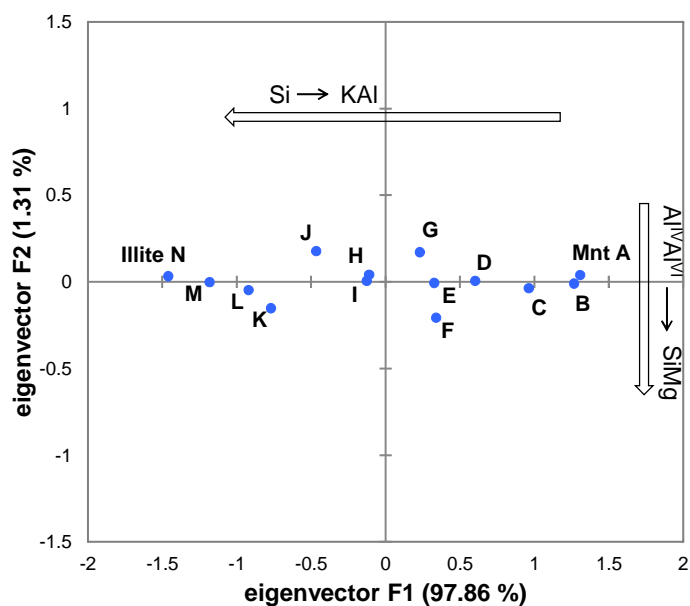


Fig. 4. Statistical treatment by PCA of the 14 compositions of I/S from the Shinzan area. The chemical trends shown by arrows were depicted from the components of the main eigenvectors F_1 and F_2 . Mnt A = Montmorillonite A.

2.2.3. Hydration state of samples for calorimetric measurements

Calorimetric measurements were performed on the dehydrated ISCz-1 sample. The dehydrated state was obtained by heating the sample at 150°C for 20 h. The sample was then stored in a glove box under a dry Ar atmosphere ($H_2O < 4$ ppm).

For the interstratified samples from the Shinzan hydrothermal area, the enthalpies of dissolution were acquired by Perbost et al. [41]. The samples were stored at room relative humidity before measurements. Their hydration states were not determined. Nevertheless, slight variations in relative humidities and therefore in the amounts of water adsorbed for a given sample were not expected to significantly affect the

enthalpy of the dissolution of the sample because the contribution of the enthalpy of dilution is negligible compared to the total measured enthalpy of the dissolution and its associated uncertainty.

2.3. Calorimetric methods

2.3.1. Heat capacity measurements

For the dehydrated ISCz-1 sample, the heat capacities were measured by heat-pulse calorimetry between 2 and 330 K using a commercial apparatus (PPMS model 6000, Quantum Design Inc.; [42]). Measurements were performed twice in the temperature range by cooling the sample from 50 K to 2 K and heating it from 50 K to 330 K, respectively. They were conducted on two pressed pellets weighing 4.16 and 14.93 mg, respectively, in order to check the reproducibility of the measurements. The relative uncertainties of the standard molar heat capacities $C_{p,m}^{\circ}$ were estimated to be $\pm 0.05 C_{p,m}^{\circ}$ for $2 \text{ K} < T < 50 \text{ K}$ and $\pm 0.02 C_{p,m}^{\circ}$ for $T > 50 \text{ K}$ according to the results from Dachs and Benisek [43] that were obtained for several standard reference powder samples that were encapsulated in Al-pans.

Moreover, the heat capacities of the dehydrated Ca-saturated ISCz-1 sample were also measured by using differential scanning calorimetry (DSC), between 300 and 520 K. The apparatus was a Calvet DSC111 from Setaram. DSC calibration consisted in a commonly used procedure for temperature and enthalpy calibration and heat capacity calibration. Standard elements indium (Goodfellow, purity 99.999%), zinc (Goodfellow, purity 99.999%) and aluminium (Goodfellow, purity 99.999%) were used for temperature and enthalpy calibration [44], and synthetic sapphire (National Bureau of Standards, [45]) for the heat capacity calibration, in the same conditions of nitrogen flow rate and heating rate as for the heat capacity measurements on the clay sample. The amount of each dehydrated sample was 140.03 mg. The heat capacity measurements were performed at 2 K min^{-1} heating rate, with temperature increase steps of 2.5 K and followed by temperature stabilization periods of 600 s allowing the heat flow to recover the baseline value.

The validation of the experimental data relies on the use of two heat capacity measurement techniques with a temperature overlapping domain, performed in two different laboratories (Tokyo Institute of Technology and IM2NP for PPMS and DSC measurements, respectively). Moreover, the measured data are

compared with previously measured on a dehydrated ISCz-1 sample saturated by sodium [33], using both low-temperature adiabatic calorimetry between 10 K and 380 K, and DSC between 300 K and 520 K. Such verification procedure was already used successfully in previous studies for the same type of minerals ([25,26,33]).

2.3.2. Acid solution calorimetry at 298.15 K

The enthalpies of the dissolution of the dehydrated ISCz-1 sample and of the fourteen illite/smectite samples from the Shinzan area were measured by isothermal dissolution calorimetry at 298.15 K in HF-HNO₃ solutions. The apparatus was a highly sensitive Tian-Calvet calorimeter, i.e., the Calsol calorimeter [46]. The calorimeter was calibrated by dissolving THAM (tris(hydroxymethyl)aminomethane) crystal in a solution of hydrochloric acid at a molality of 0.12 mol kgw⁻¹ (kgw = kilogram of water) under the same conditions as those used in the experiments of Ganteaume et al. [46]. The experimental procedure for the measurements was described in detail in Gailhanou et al. [25]. All the acids and THAM reagent used for calibration of the apparatus and the calorimetric measurements are analytical grade reagents, described in Table 7. The acid solution compositions were HF(12.0 mol kgw⁻¹)-HNO₃(1.9 mol kgw⁻¹) for the ISCz-1 sample and HF(3.0 mol kgw⁻¹)-HNO₃(1.9 mol kgw⁻¹) for the illite/smectite samples from the Shinzan area, respectively. A higher HF concentration was necessary in the case of the ISCz-1 sample compared with the other samples because of the presence of a relatively high amount of quartz (5.85 wt.% at dry state, Table 4) in the ISCz-1 sample. The compositions of the acid solutions were optimized to obtain the complete dissolution of the samples after stirring at 540 rpm for approximately 3 h. For the series of I/S samples from the Shinzan area, approximately 20–30 mg of samples were dissolved in 50 mL of the HF-HNO₃ acid solution. For the dehydrated ISCz-1 sample, the amount of sample was approximately 40 mg for 50 mL of the HF-HNO₃ acid solution. The weighing of this latter sample and the filling of the sample vessel were conducted in a glove box under a dry argon atmosphere in order to prevent the hydration of the sample.

Table 7. Origin and purity of the chemical reagents and the secondary references used for the solution calorimetry experiments.

Name	Formula	CAS number	Source	Titration 10 ² w ^a	Mole fraction purity	Analysis methods
<i>Acid solutions</i>						
Hydrochloric acid 37%, AnalaR Normapur	HCl	7647-01-0	VWR	36.3		N/A ^b
Hydrofluoric acid 40%, AnalaR Normapur	HF	7664-39-3	VWR	40.4		N/A ^b
Nitric acid 69%, AnalaR Normapur	HNO ₃	7697-37-2	VWR	69.7		N/A ^b
THAM ^c , ACS reagent	NH ₂ C(CH ₂ OH) ₃	77-86-1	Alfa Aesar		≥ 0.998	N/A ^b
<i>Secondary references</i>						
Gibbsite	Al(OH) ₃	14762-49-3	Baker		≥ 0.99	PXRD ^d
Quartz-α	SiO ₂	14808-60-7	Natural BRGM's collection ^e		≥ 0.99	PXRD ^d , Rietveld analysis
Goethite	FeOOH	20344-49-4	Natural BRGM's collection ^e		≥ 0.99	PXRD ^d
Iron(II) oxide	FeO	1345-25-1	Sigma-Aldrich		0.999	N/A ^b
Calcium hydroxide	Ca(OH) ₂	1305-62-0	Merck		≥ 0.96	PXRD ^d
Magnesium hydroxide	Mg(OH) ₂	1317-43-7	Fluka		≥ 0.99	N/A ^b
Potassium nitrate	KNO ₃	7757-79-1	Sigma-Aldrich		≥ 0.99	N/A ^b
Rutile	TiO ₂	13463-67-7	ChemPur		0.995	N/A ^b

^a Mass fraction of acid in solution.

^b Titration or purity as stated by the manufacturer.

^c THAM: Tris(hydroxymethyl)aminomethane (dried basis).

^d Powder X-ray diffraction.

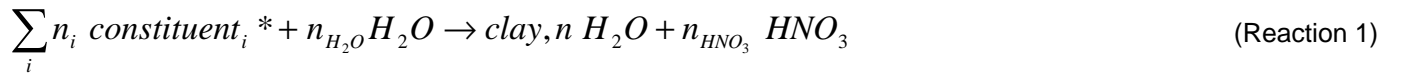
^e BRGM: The French Geological Survey.

2.3.3. Standard enthalpy of formation of a clay mineral from constituents at 298.15 K

The experimental method described in Gailhanou et al. [25,26] was used for the determination of the enthalpy of formation of the illite/smectite in the ISCz-1 sample. This method was based on a set of dissolution reactions that were performed in order to obtain the enthalpy of formation of the clay mineral from secondary references, i.e. constituents chosen among oxides, hydroxides and nitrates, at 298.15 K. The source and purity of each secondary reference used for the solution calorimetry experiments are described in Table 7. The reactions involved were (i) the dissolution of the clay sample containing the clay mineral and impurities, (ii) the dissolution of a mixture composed of the constituents and impurities with the same composition as those present in the sample and (iii) dilution type reactions. The masses of the constituents (except H₂O) and impurities are calculated to present the elemental stoichiometry of 40 mg of the anhydrous

clay sample. The impurities were added to the mixture of constituents in the same amounts as those present in the clay sample. The weightings were realized in an air atmosphere for better stability of the balance with a maximum deviation of 10 µg. Thus, approximately 46 mg of a mechanical mixture of constituents and impurities was made for each dissolution experiment to accurately determine the amounts of each of these compounds present in the mixture. Four or five measurements of enthalpies of dissolution were performed on each clay sample and constituent and impurity mixtures. The relative uncertainties for the enthalpy values are lower than 0.5%. For a given sample, slight variations in the final solution composition existed due to the slight variations in the sample mass. However, no dependency of the enthalpy of dissolution (per unit mass) on these variations of solution composition was observed.

The reaction of the formation of a clay mineral from the constituents is expressed by the following (Gailhanou et al. [26]):



The superscript (*) refers to all constituents except H₂O, n_i is the amount of the constituent i (in mol/mol of clay), n_{H_2O} is the amount of the constituent H₂O, obtained indirectly from the balance on the O and H elements, n is the amount of water adsorbed in the clay, and n_{HNO_3} is the amount of HNO₃ calculated from the balance of nitrates.

Therefore, the enthalpy of formation of the clay mineral from constituents $\Delta H_{f/constit}^0$, expressed in kJ mol⁻¹ O₁₀(OH)₂, may be calculated according to the following formula (Eq. (1)) described in Gailhanou et al. [26]:

$$\begin{aligned} \Delta H_{f/constit}^0 = & \left(\frac{M_{\min}}{\eta} + (n - n_{H_2O}) M_{H_2O} + n_{HNO_3} M_{HNO_3} \right) \Delta H_{diss,mixt} - \left(\frac{M_{\min}}{\eta} + n M_{H_2O} \right) \Delta H_{diss,sample} \\ & + n_{H_2O} M_{H_2O} \Delta H_{dil} - n_{HNO_3} M_{HNO_3} \Delta H_{sol} \end{aligned} \quad (1)$$

where M_{\min} is the molar mass of the dehydrated clay mineral, η is the mass fraction of the dehydrated clay in the dehydrated sample, $\Delta H_{diss,mixt}$ is the enthalpy of dissolution of the mixture of constituents and impurities (in J g⁻¹), and $\Delta H_{diss,sample}$ is the enthalpy of dissolution of the clay sample (in J g⁻¹). ΔH_{dil} and ΔH_{sol} are expressed in J g⁻¹ and are the enthalpy of the dilution of the HF-HNO₃ solution and the integral enthalpy of the mixing of pure HNO₃ in the HF-HNO₃ solution, respectively.

Finally, the standard enthalpy of formation ΔH_f^0 of the illite/smectite mineral was determined by considering $\Delta H_{f/constit}^0$ and the standard enthalpies of formation of the chosen secondary references (see Appendix C for these latter data).

3. Results

3.1. Thermodynamic properties of anhydrous illite-smectite ISCz-1

3.1.1. Heat capacities

The heat capacities measured for the Ca-saturated ISCz-1 sample are plotted between 2 and 330 K for the PPMS data and between 300 and 520 K for the DSC data (Fig. 5). All the raw data are given in Appendix A. A significant discrepancy between the measurements performed by both techniques appears in the common temperature range 300 – 330K. By comparison with the C_p data previously obtained for the Na-saturated ISCz-1 sample [33], the C_p values measured by the DSC increase similarly, whereas the C_p values from the PPMS tend to drastically depart from the general trend between approximately 230 K and 330 K. This comparison with the Na-saturated ISCz-1 sample reveals the inaccuracy of the C_p values measured by the PPMS at the highest temperature range. This could possibly be due to the low thermal conductivity of the sample. Indeed, Kennedy et al. [47] tested the reliability of the PPMS technique on several phases and noticed that the C_p data could be precise but inaccurate at high temperatures ($T > 150$ K), particularly in the case of poor thermal conductor phases. Therefore, the heat capacities were fitted between 200 K and 520 K based on PPMS data from 170 K and 200 K and DSC data from 300 K and 520 K, using a least-square fit method and the polynomial function $C_p(T) = 2.7070 - 3.1540 \cdot 10^{-4}T + 1.2571 \cdot 10^4 T^{-2} - 33.827 T^{-0.5}$. The relative errors for the least square fit method are lower than 1.5% at $T > 300$ K.

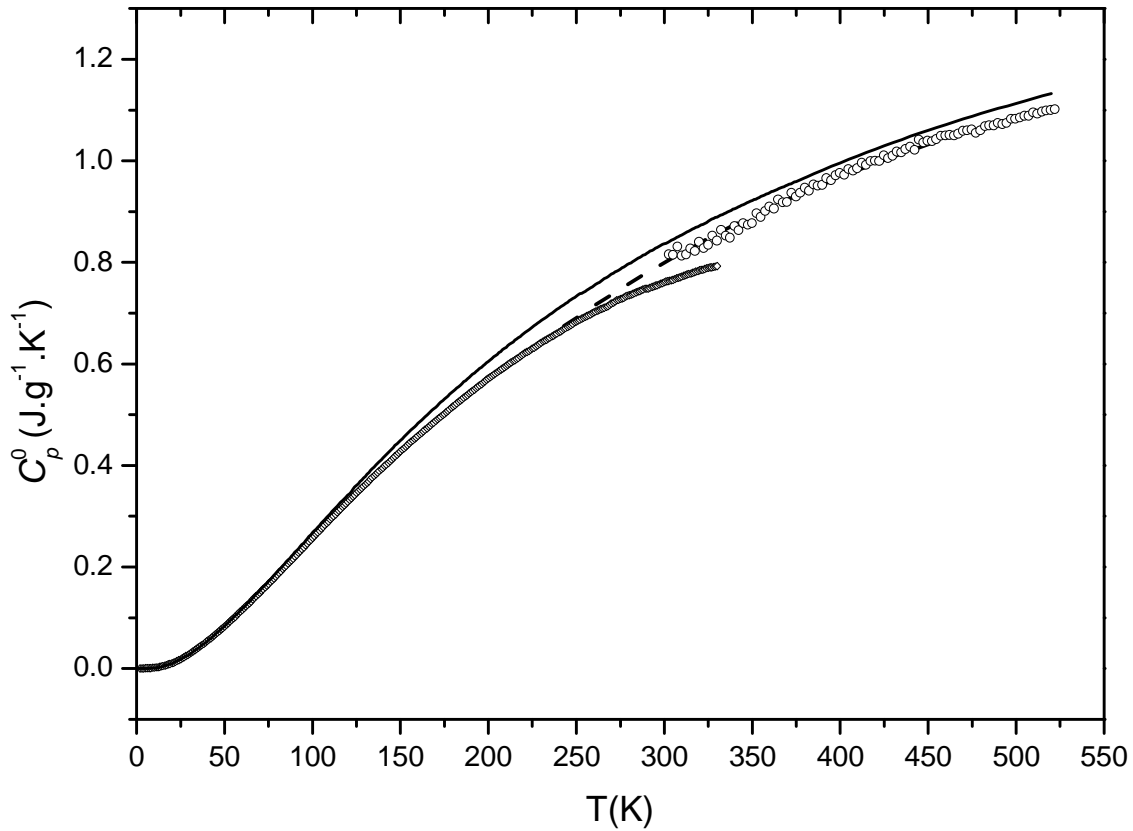


Fig. 5. Heat capacities measured for the ISCz-1(Ca) sample (diamonds are values obtained by PPMS for sample 2, circles are values obtained by DSC, and the dashed line represents interpolated data) and the ISCz-1(Na) sample (line represents data from Gailhanou et al., 2007) between 2 K and 520 K.

The heat capacities of the sample are linearly interpolated at every degree between 0 K and 520 K. The heat capacities of the interstratified mineral are then calculated from the heat capacities of the sample by subtracting the contributions of the impurities (Table 4, Amounts in dry sample) according to the method described in [33]. The heat capacity data for the impurities are taken from the literature (see references in Appendix B). Due to the lack of thermodynamic data for the bernalite ($\text{Fe}(\text{OH})_3$) mineral, this latter was assimilated to goethite (FeOOH) for the correction calculations. Therefore, the amount of goethite considered in the following text was adjusted to 0.98 wt.%, with respect to the total amount of Fe in goethite and bernalite. Furthermore, the heat capacity data for goethite provided by Majzlan et al. [48] are limited to the temperature range 0–375 K due to the occurrence of the magnetic Neel transition at approximately 375 K.

Consequently, the correction of the heat capacities of the sample is performed between 0–375 K. The heat capacities of the interstratified mineral $C_{p,m}^0$ expressed in $\text{J K}^{-1} \text{mol}^{-1}$ are given in Table 8 for this temperature range.

3.1.2. Heat contents

The heat contents of the mineral, $H^0(T) - H^0(298.15 \text{ K})$, are calculated at any temperature by the integration of the $C_{p,m}^0$ function according to (Eq. (2))

$$H^0(T) - H^0(298.15 \text{ K}) = \int_{298.15}^T C_{p,m}^0 dT \quad (2)$$

using the trapezoid method at every degree (Table 8).

3.1.3. Entropy

The third-law entropy of the interstratified mineral S^0 is expressed by (Eq. (3))

$$S^0 = \int_0^T \frac{C_{p,m}^0}{T} dT + S^{0(\text{resid.})} \quad (3)$$

where $S^{0(\text{cal})} = \int_0^T \frac{C_{p,m}^0}{T} dT$ is the calorimetric entropy obtained directly from heat capacity measurements, and $S^{0(\text{resid.})}$ is the residual entropy. This latter term is not available from the calorimetric measurements and requires additional structural information about the mineral (Ulbrich and Waldbaum [49]). The chemical site configurational entropy $S^{0(\text{conf})}$ that is associated with cationic disorder in the tetrahedral, octahedral and interlayer crystallographic sites is a significant part of the residual entropy. It is assessed according to the method described by Blanc *et al.* [1]. The residual entropy may also include a magnetic term $S^{0(\text{mag})}$ that is associated with the disordering of the magnetic spins of the transition metals in the clay minerals that occurs at very low temperatures (e.g., $T < 10 \text{ K}$ for iron-bearing chlorites, [50]). Indeed, for clay minerals, a recent study [51] has shown that the magnetic contribution is difficult to detect and directly measure. The maximum value of the magnetic entropy is assessed by using the method described by Ulbrich and Waldbaum [49]. It is based on the maximum number of spin configurations as follows:

$$S^{0(\text{mag})} = R \sum_i x_i \ln(2S_i + 1) \quad (4)$$

where R is the gas constant, x_i the amount of element i and S_i its spin number (for Fe^{3+} , $S = 5/2$; for Fe^{2+} , $S = 2$). Finally, in the case of the illite-smectite ISCz-1 mineral, the configurational entropy $S^{0(\text{conf})}$ is estimated at $29.53 \text{ J K}^{-1} \text{ mol}^{-1}$ and the maximum magnetic entropy is estimated at $0.58 \text{ J K}^{-1} \text{ mol}^{-1}$. The calorimetric entropy data are given in Table 8.

Table 8. Standard molar heat capacities and derived thermodynamic functions of the dehydrated Ca-saturated illite-smectite ISCz-1 mineral (corrected for impurities) at selected temperatures.

T/K^a	$C_{p,m}^0(T)^b$ /J K ⁻¹ mol ⁻¹	$S^{(0)cal\ b}$ /J K ⁻¹ mol ⁻¹	$H^0(T) - H^0(298.15)^b$ /kJ mol ⁻¹
0	0.00	0.00	-45.596
10	0.65	0.42	-45.594
20	3.81	1.65	-45.574
30	10.61	4.37	-45.505
40	19.85	8.64	-45.354
50	30.53	14.20	-45.103
60	43.01	20.85	-44.736
70	55.96	28.44	-44.242
80	69.39	36.78	-43.616
90	83.08	45.75	-42.853
100	97.33	55.24	-41.951
110	111.25	65.17	-40.908
120	124.93	75.43	-39.728
130	138.00	85.95	-38.413
140	150.81	96.66	-36.968
150	163.30	107.49	-35.397
160	175.05	118.41	-33.704
170	186.45	129.37	-31.896
180	197.55	140.34	-29.976
190	208.35	151.31	-27.946
200	216.66	162.25	-25.812
210	226.42	173.06	-23.597
220	236.19	183.82	-21.284
230	245.88	194.53	-18.873
240	255.41	205.20	-16.367
250	264.72	215.81	-13.766
260	273.81	226.38	-11.073
270	282.64	236.88	-8.291
280	291.20	247.31	-5.421
290	299.49	257.67	-2.468
298.15	306.05	266.06	0.000
300	307.51	267.96	0.568
310	315.23	278.17	3.682
320	322.66	288.30	6.871
330	329.83	298.34	10.134
340	336.73	308.29	13.467
350	343.37	318.14	16.868
360	349.78	327.91	20.334
370	355.95	337.58	23.862
375	358.95	342.37	25.650

^a Standard uncertainties u are $u(T) = 0.01$ K ($2 < T/K < 300$) and $u(T) = 0.05$ K ($300 < T/K < 375$).

^b The relative expanded uncertainties with level of confidence 0.95 ($k = 2$) are: $U_r(C_{p,m}^0) = 0.05$ ($2 < T/K < 50$), $U_r(C_{p,m}^0) = 0.02$ ($50 < T/K < 298$) and $U_r(C_{p,m}^0) = 0.015$ ($298 < T/K < 375$), $U_r(S^{(0)cal}) = 0.05$ ($2 < T/K < 75$), $U_r(S^{(0)cal}) = 0.04$ ($75 < T/K < 110$) and $U_r(S^{(0)cal}) = 0.03$ ($110 < T/K < 375$), $U_r(H^0(T) - H^0(298.15)) = 0.05$ ($2 < T/K < 50$) and $U_r(H^0(T) - H^0(298.15)) = 0.02$ ($50 < T/K < 375$).

3.1.4. Standard enthalpy of formation

The enthalpies of dissolution of the Ca-saturated ISCz-1 sample and the constituents* mixture (superscript * means: except H₂O constituent) are given in Table 9. The uncertainties associated with the mean enthalpies are calculated from the measured values using the Student's *t*-distribution with a 95 percent confidence interval.

Table 9. Enthalpies of dissolution of the dehydrated ISCz-1 sample and of the corresponding mixtures of constituents* and impurities (* = except H₂O constituent) measured in 50 mL of a HF (12.0 mol kgw⁻¹) – HNO₃ (1.9 mol kgw⁻¹)^a solution, at 298.15 K and at ambient pressure of 101 kPa^b.

Sample	Sample mass m _s /mg	ΔH_{diss} /J g ⁻¹	Mean ΔH_{diss} /J g ⁻¹
Dehydrated sample ISCz-1	39.86	-2256.60	-2257.82 ± 1.19 ^c
	40.01	-2257.62	
	40.00	-2258.31	
	40.20	-2258.74	
Mixture of constituents* + impurities	46.58	-1845.16	-1848.08 ± 3.28 ^c
	46.55	-1849.41	
	46.55	-1850.65	
	46.57	-1845.16	
	46.55	-1850.00	

^a kgw = kilogram of water.

^b Standard uncertainties are $u(T) = 0.01$ K and $u(p) = 3$ kPa.

^c The expanded uncertainties are calculated using the Student's *t*-distribution with level of confidence 0.95 ($k = 2$).

The enthalpy of dilution ΔH_{dil} and the solution enthalpy ΔH_{sol} associated with the mixing of pure HNO₃ in HF (12.0 mol kgw⁻¹) – HNO₃ (1.9 mol kgw⁻¹) were determined by Gailhanou *et al.* [25]. Their values were - 1.57 ± 0.30 kJ mol⁻¹ H₂O and - 27.7 ± 7.5 kJ mol⁻¹ HNO₃, respectively.

Afterwards, the enthalpy of formation of the mineral from the constituents was then calculated using (Eq. (1)) (Table 10).

Table 10. Reaction of the formation of illite-smectite ISCz-1 from the secondary reference constituents and the associated enthalpy $\Delta H_{f/constit.}^{\circ}$ in kJ mol^{-1} (at 101 kPa and 298.15 K)^a.

Reactions of formation of mixed-layer mineral		$\Delta H_{f/constit.}^{\circ}$ /kJ mol ⁻¹
Illite/smectite ISCz-1 (n = 0)	3.562 SiO _{2,quartz} + 2.159 Al(OH) _{3,gibbsite} + 0.255 Mg(OH) ₂ + 0.029 FeOOH _{1,goethite} + 0.011 FeO + 0.008 TiO _{2,rutile} + 0.440 KNO ₃ + 0.092 Ca(OH) ₂ = (Ca _{0.092} K _{0.440})(Si _{3.562} Al _{0.438})(Al _{1.721} Mg _{0.255} Fe ³⁺ _{0.029} Fe ²⁺ _{0.011} Ti _{0.008})O ₁₀ (OH) ₂ +0.440 HNO ₃ +2.380 H ₂ O	62.36 ^b ±3.49 ^c

^a Standard uncertainties are $u(T) = 0.01$ K and $u(p) = 3$ kPa.

^b $\Delta H_{f/constit.}^{\circ}$ is calculated using Eq. (1) and the following parameter values: $M_{min}(\text{ISCz-1}) = 382.01$ g mol⁻¹, $\eta(\text{ISCz-1}) = 0.88456$, $M_{\text{H}_2\text{O}} = 18.02$ g mol⁻¹, $M_{\text{HNO}_3} = 63.01$ g mol⁻¹, $n_{\text{H}_2\text{O}} = -2.380$, $n = 0$, and $n_{\text{HNO}_3} = 0.440$.

^c The combined uncertainty is calculated from propagation from enthalpies of dissolution, dilution and solution with level of confidence 0.95 ($k = 2$).

Eventually, the standard enthalpy of formation of the illite-smectite ΔH_f^0 was obtained from the enthalpy of formation of the mineral from the secondary reference constituents and from the standard enthalpies of formation of these reference constituents (Table 11). These latter values are selected consistently with the Thermochemie database [27] at 298.15 K and are tabulated in Appendix C.

The expanded uncertainty associated with ΔH_f^0 at approximately 95 percent confidence limits was calculated using the error propagation method [52]:

$$\Delta(\Delta H_f^0) = \sqrt{\left[\Delta(\Delta H_{f/constit.}^0)\right]^2 + \sum_i \nu_i^2 \left[\Delta(\Delta H_f^{constit.i})\right]^2} \quad (5)$$

where ν_i is the stoichiometric coefficient associated with the constituent i , and $\Delta H_f^{constit.i}$ is the standard enthalpy of formation of the constituent i . The expanded uncertainties associated with $\Delta H_{f/constit.}^0$ and $\Delta H_f^{constit.i}$ correspond to two standard deviations of the mean.

3.1.5. Standard Gibbs free energy of formation

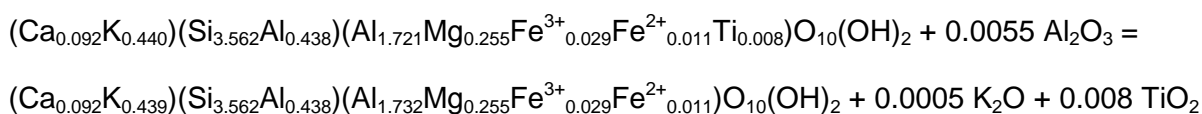
The standard Gibbs free energy of formation ΔG_f^0 is expressed by the following:

$$\Delta G_f^0(T) = \Delta H_f^0(T) - T\Delta S_f^0(T) \quad (6)$$

where $\Delta S_f^0(T)$ is the standard entropy of formation of the interstratified mineral that was calculated considering the reaction of the formation of the mineral from the elements in their stable thermodynamic form in standard conditions. The value for the ISCz-1 mineral was obtained at 298.15 K from the third-law entropy of the mineral S^0 (Table 11) and from the entropies of the elements at 298.15 K. These latter data were taken from Cox et al. [53] for Si_(cr), Al_(cr), Mg_(cr), Ti_(cr), K_(cr), Ca_(cr), O_{2(g)}, and H_{2(g)} and from Chase [54] for Fe_(cr).

3.2. Synthesis of the thermodynamic properties of anhydrous illite-smectite ISCz-1

The complete set of thermodynamic properties of the interstratified ISCz-1 mineral at 298.15 K and ambient pressure of 101 kPa is given in Table 11. Since its structural formula contains a small amount of TiO₂, a simplified structural formula is also proposed in order to ease the introduction of the phase into geochemical databases. The simplification is based on the fictive solid-solid reactions with properties of reaction equal to 0. In the case of the interstratified ISCz-1, the following fictive reaction was considered:



(Reaction 2)

The thermodynamic data for alumina- α , rutile and K₂O are taken from the Thermoddem database (<http://thermoddem.brgm.fr>, [28]).

Table 11. Standard thermodynamic properties of the dehydrated illite-smectite ISCz-1 mineral at 298.15 K.

T = 298.15 K ^a									A	B*10 ³	C*10 ⁻⁵	T range for C _p functions ^b
Mineral	ΔG_f^0 /kJ mol ⁻¹	ΔH_f^0 /kJ mol ⁻¹	S ^{(cal)0} /J K ⁻¹ mol ⁻¹	S ^{(cont)0} /J K ⁻¹ mol ⁻¹	S ^{(mag)0} /J K ⁻¹ mol ⁻¹	S ⁰ /J K ⁻¹ mol ⁻¹	V ^{0e} /cm ³ mol ⁻¹	C _p ⁰ (298.15 K) /J K ⁻¹ mol ⁻¹	/J K ⁻¹ mol ⁻¹	/J K ⁻² mol ⁻¹	/J K mol ⁻¹	T/K
Illite/smectite ISCz-1	-5417.38 ± 7.72 ^c	-5785.74 ± 7.46 ^c	266.06 ± 6.00 ^d	29.53	0.58	296.17 ± 6.00 ^d	137.13	306.05 ± 4.59 ^d	228.34	438.31	-47.04	298 – 375
Properties recalculated for a simplified structural formula^f												
Illite/smectite ISCz-1*	-5418.62 ± 7.72 ^c	-5787.22 ± 7.46 ^c	265.89 ± 6.00 ^g	28.89	0.58	295.36 ± 6.00 ^g	137.13	306.00 ± 4.59 ^g	236.14	422.80	-49.92	298 – 375

^a The standard uncertainty is $u(T) = 0.05$ K.

^b Between 298 and 375 K, the $C_p(T)$ functions are expressed according to $C_p(T) = A + B 10^{-3} T + C 10^5 T^{-2}$, where A , B and C are Maier-Kelley coefficients obtained by fitting the C_p curves.

^c The combined uncertainties $U(\Delta H_f^0)$ and $U(\Delta G_f^0)$ are calculated from propagation from enthalpy (Eq. 5) and from enthalpy and entropy (Eq. 6), respectively, with level of confidence 0.95 ($k = 2$).

^d The relative expanded uncertainties in the property values are $U_r(C_{p,m}^0) = 0.015$ and $U_r(S^{(0)cal}) = 0.03$ at approximately 95 percent confidence limits.

^e The molar volume V^0 is obtained from the values of unit-cell parameters a , b and c that are determined from the XRD patterns according to the method described in [25]. For the present dehydrated ISCz-1 sample, $a = 5.20$ Å, $b = 9.01$ Å and $c = 10.12$ Å.

^f The simplified formula is introduced to ease the predictive calculations and their implementation into databases. The simplification process is detailed in the text.

Illite/smectite ISCz-1: $(Ca_{0.092}K_{0.440})(Si_{3.562}Al_{0.438})(Al_{1.721}Mg_{0.255}Fe^{3+}_{0.029}Fe^{2+}_{0.011}Ti_{0.008})O_{10}(OH)_2$

Illite/smectite ISCz-1*: $(Ca_{0.092}K_{0.438})(Si_{3.562}Al_{0.438})(Al_{1.732}Mg_{0.255}Fe^{3+}_{0.029}Fe^{2+}_{0.011})O_{10}(OH)_2$.

^g The combined uncertainties $U(C_{p,m}^0)$ and $U(S^{(0)cal})$ are calculated from propagation from reaction (2) with level of confidence 0.95 ($k = 2$).

3.3. *Enthalpies of mixing for mixed-layers from Shinzan area samples*

The analysis of the chemical dispersions among the illite-smectite compositions (Fig. 4) shows that all of these can be described by a ternary solid solution with illite N, montmorillonite A and illite-smectite F end members. Although the chemical dispersions mostly occur along the illite – smectite join, the ternary solid solution is preferred to the binary illite-smectite solid solution for accuracy reasons. Since the enthalpies of mixing are expected to be very low and consequently difficult to determine accurately given their associated uncertainties, it is preferable to perform the measurements for all samples without adding a supplementary source of uncertainty induced by the consideration of the binary solid solution. In the following, the enthalpies of mixing for the illite-smectite compositions from the Shinzan area samples are determined by using isothermal dissolution calorimetry at 298.15 K.

3.3.1. *Enthalpies of dissolution at 298.15 K*

The enthalpies of dissolution of all the samples in HF (3.0 mol kg⁻¹) – HNO₃ (1.9 mol kg⁻¹) at 25 °C were measured by Perbost et al. [41]. The data are given in Table 12. The uncertainties associated with the mean enthalpies are calculated from the measured values using the Student's *t*-distribution with a 95 percent confidence interval. The relative expanded uncertainties are lower than 1%, except for sample E, for which it is 2.4%.

Table 12. Mean values of the enthalpies of dissolution of the I/S samples from the Shinzan area measured by [41] in a HF (3.0 mol kg⁻¹) – HNO₃ (1.9 mol kg⁻¹) solution at 298.15 K and at ambient pressure of 101 kPa^a. Mnt A = Montmorillonite A.

Sample	Mean value of $\Delta H_{diss}/\text{kJ mol}^{-1}$
Mnt A	-846.80 ± 7.64 ^b
B	-844.03 ± 1.45 ^b
C	-859.94
D	-847.84 ± 2.95 ^b
E	-842.04 ± 20.2 ^b
F	-840.77 ± 2.52 ^b
G	-859.50
H	-848.02 ± 6.14 ^b
I	-860.21
J	-866.77
K	-854.33
L	-858.14
M	-863.52
Illite N	-863.57 ± 6.39 ^b

^a Standard uncertainties are $u(T) = 0.01$ K and $u(p) = 3$ kPa.

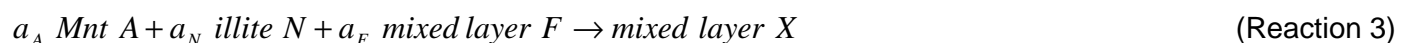
^b The expanded uncertainties are calculated using a Student's *t*-distribution with a 95 percent-confidence interval.

The opposite values of the enthalpies of dissolution are plotted in Fig. 6a in the 3-D space (F_1 , F_2 , $-\Delta H_{diss}$). The values were then modeled by a polynomial surface by using the least-square fit method. The polynomial function is expressed by $P(z_1, z_2) = 852.124 - 6.029 z_1 + 41.159 z_2 + 1.938 z_1^2 - 4.027 z_1 z_2 + 45.824 z_2^2$, where z_1 and z_2 correspond to the coordinates of the I/S samples in the (F_1 , F_2) plane.

The discrepancies between the measured and modeled values are lower than 1%, which correspond to the relative uncertainties for the measured enthalpies of dissolution. The modeled values are then considered to correctly represent the measured enthalpies.

3.3.2. Enthalpies of mixing at 298.15 K

The illite-smectite compositions may be expressed by a linear combination of the compositions of the three end-members A, F and N. The barycentric coordinates (a_A , a_F , a_N) for all I/S samples are given in Table 13. With respect to these end-members, the enthalpy of mixing for an interstratified X from the Shinzan series is the enthalpy of the reaction of formation:



which practically represents the enthalpy of mixing of illite and smectite (montmorillonite) layers. The enthalpy of mixing is then obtained from the combination of the reactions of dissolution of the minerals and is expressed by the following (Eq. (7)):

$$\Delta H_m(\text{mixed layer } X) = a_A \Delta H_{diss}(A) + a_N \Delta H_{diss}(N) + a_F \Delta H_{diss}(F) - \Delta H_{diss}(\text{mixed layer } X) + \Delta H_{mix,sol} \quad (7)$$

where $\Delta H_{mix,sol}$ is the enthalpy of mixing of solutions resulting from the dissolution reactions of the minerals, which is neglected in the following calculations. The enthalpies of mixing are then calculated for all illite-smectite compositions from the modeled values of the enthalpies of dissolution (Table 13) and are plotted in the (F_1 , F_2 , ΔH_m) space (Fig. 6b).

Table 13. Enthalpies of mixing of the I/S minerals from the Shinzan area, at 298.15 K and 101 kPa^a, with respect to the three end-members Mnt A, illite N and interstratified F. a_A , a_N and a_F are the barycentric coordinates of the I/S minerals that are associated with the corresponding end-members.

Samples	a_A	a_N	a_F	$\Delta H_m/\text{kJ mol}^{-1}$ ^b
Mnt A	1.00	0.00	0.00	0.00
B	0.85	-0.06	0.20	0.20
C	0.68	0.02	0.30	-0.99
D	0.66	0.21	0.13	-2.64
E	0.53	0.29	0.18	-3.18
F	0.00	0.00	1.00	0.00
G	0.97	0.58	-0.55	-2.98
H	0.50	0.52	-0.02	-3.70
I	0.40	0.47	0.12	-3.62
J	0.74	0.85	-0.59	-2.41
K	-0.25	0.48	0.77	-1.47
L	-0.02	0.69	0.33	-2.19
M	0.01	0.85	0.14	-1.39
Illite N	0.00	1.00	0.00	0.00

^a Standard uncertainties are $u(T) = 0.01$ K and $u(p) = 3$ kPa.

^b Expanded uncertainties are obtained from error propagation calculation method using Eq. (7) and are estimated at 5 kJ mol^{-1} with level of confidence 0.95.

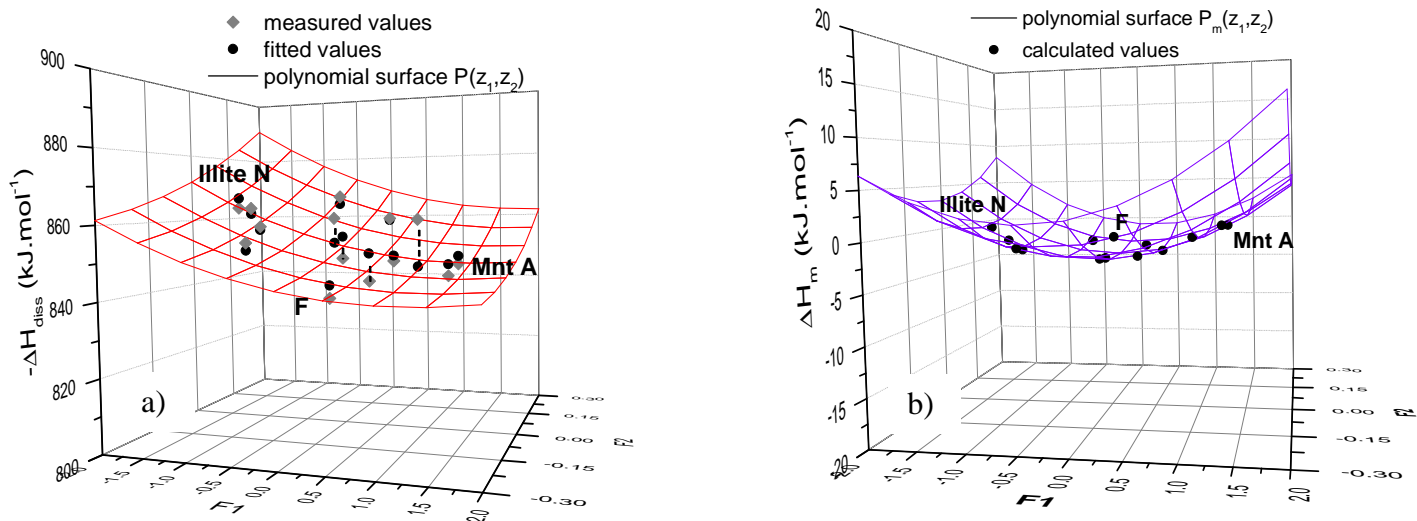


Fig. 6. 3-D plots of (a) the polynomial surface $P(z_1, z_2)$ obtained by fitting the experimental values of $-\Delta H_{diss}$ for the I/S samples and (b) the enthalpies of mixing of the illite and smectite layers (per mol $O_{10}(OH)_2$) for the I/S minerals and the modeled polynomial surface $P_m(z_1, z_2)$. F_1 and F_2 are the main eigenvectors representing the main trends of chemical dispersion (see text).

The resulting enthalpies of mixing are slightly negative, with a minimum value of approximately -3.7 kJ mol^{-1} for the H and I samples. According to XRD modeling (Inoue and Utada [7]), these latter samples are first-order R1 types and should contain approximately 30–35 % smectite layers. A model predicting ΔH_m was proposed by Blanc *et al.* [24] and resulted in $\Delta H_m = -3.6 \text{ kJ mol}^{-1}$ for illite/smectite that displayed such composition and ordering types.

4. Discussion: comparison with estimates and stability domains

The standard thermodynamic properties of the illite-smectite ISCz-1 obtained experimentally are compared with the estimated values calculated from the model of Blanc *et al.* [1] that includes a consistent set of models for predicting the standard thermodynamic properties (H , S , C_p , V) of phyllosilicates. First, the model from Blanc *et al.* [1] is applied merely by considering the whole structural formula of the illite-smectite.

The simplified structural formula $(\text{Ca}_{0.092}\text{K}_{0.439})(\text{Si}_{3.562}\text{Al}_{0.438})(\text{Al}_{1.732}\text{Mg}_{0.255}\text{Fe}^{3+}_{0.029}\text{Fe}^{2+}_{0.011})\text{O}_{10}(\text{OH})_2$ is considered in order to ease the predictive calculations. The results are given in Table 14. In addition, a second approach is conducted that separately predicts the thermodynamic parameters of the illite and smectite components and combines them according to the required composition of illite-smectite. This approach aims to differentiate the chemistry of the illite and smectite layers in the interstratified, consistently with the structure of the illite-smectite mineral.

The method for calculating the chemical compositions of illite and smectite end-members relies on the work of Meunier and Velde [2], which is based on the graphical representation of the $\text{M}^+-4\text{Si}-\text{R}^{2+}$ ternary system for natural sample compositions. Plotting ISCz-1 coordinates in such a diagram (Fig. 7) reveals that illite-smectite ISCz-1 can be considered as a binary solid solution with a high-charge illite end-member (0.87/ $\text{O}_{10}(\text{OH})_2$), namely, IllitePP, and a low-charge montmorillonite end-member (0.21), namely, MontmorillonitePP. Their coordinates in the $\text{M}^+-4\text{Si}-\text{R}^{2+}$ diagram provide some chemical constraints but these are not sufficient to determine their detailed compositions, *i.e.* interlayer cations (K, Ca) and octahedral cations (Al, Mg, Fe^{2+} , Fe^{3+}) occupancies. A complementary approach was then developed, assuming ideal mixing between elements in the illite and smectite structural sites and based on the least-square fit method to adjust the compositions, by constraining the ratios between the elements in the I/S structural sites given from the structural formula of ISCz-1 mineral. The compositions of the end-members were determined considering 70% illite and 30% smectite, which is consistent with the proportions obtained from the XRD analyses. Finally, the thermodynamic properties of the illite and smectite end-members are calculated by using the model from Blanc *et al.* [1] for each end-member and then considering a linear combination of both. The dataset is completed by considering the mixing enthalpy measured for the I/S sample with 70% illite layers (sample I, Table 5) at -3.6 kJ mol^{-1} .

Regarding the results (reported in Table 14), the prediction obtained by considering the mixture of illite and smectite components falls closest to the experimental values, except for those of heat capacity and volume. For these latter properties, both estimate methods provide similar results and, on the whole, the method based on a mixture of layers slightly improves the accuracy of the predictions for the properties of the illite-smectite ISCz-1 mineral.

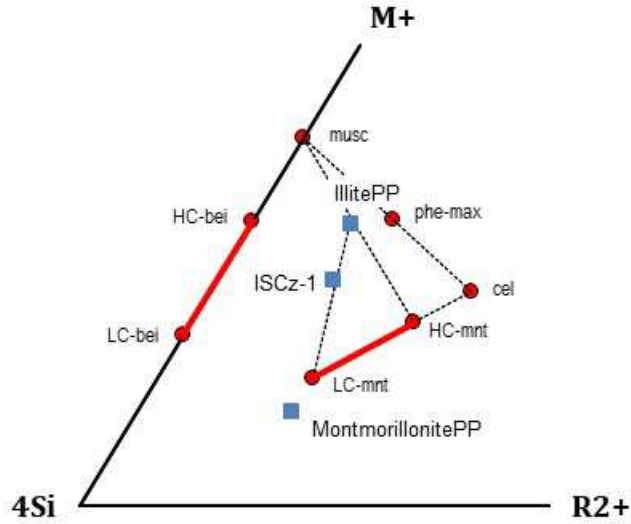


Fig. 7. Composition of clay minerals in the $M^+-4Si-R^{2+}$ system (M^+ = layer charge of ideal mica, $4Si$ = Si content of the tetrahedral sheet, and R^{2+} = amount of bivalent cations in the octahedral position). Musc = Muscovite, phe-max = Phengite, cel = celadonite, IllitePP = illite end-member (layer charge 0.87), HC-mnt = high-charge montmorillonite, LC-mnt = low-charge montmorillonite, and beid = beidellite. IllitePP and MontmorillonitePP are the predicted compositions of the illite and smectite components in the mixed-layer ISCz-1, respectively.

Table 14. Estimated thermodynamic properties of ISCz-1 at 298.15 K and 101.325 kPa according to the model from [1]. Data are compared to the experimental values obtained for the simplified structural formula that was given in Table 11. $\bar{\mathcal{E}}_{exp}$ and $\bar{\mathcal{E}}_{pred}$ refers to experimental and predicted thermodynamic properties (G , H , S , C_p , V), respectively.

Predicted	ΔG_f^0 /kJ mol ⁻¹	ΔH_f^0 /kJ mol ⁻¹	$S^{0(tot)}$ /J K ⁻¹ mol ⁻¹	$C_{p,m}^o$ (298.15 K) /J K ⁻¹ mol ⁻¹	V^o /cm ³ mol ⁻¹
Mean composition	-5412.63	-5779.31	301.78	317.63	137.94
$\bar{\mathcal{E}}_{exp} - \bar{\mathcal{E}}_{pred}$	-5.99	-7.91	-6.41	-11.58	-0.81
$(\bar{\mathcal{E}}_{exp} - \bar{\mathcal{E}}_{pred})/ \bar{\mathcal{E}}_{exp} *100$	-0.11	-0.13	-2.17	-3.78	-0.59
Mix illite/smectite layers ^a	-5415.19	-5785.52	290.03	317.79	138.00
Mixing terms	-3.6	-3.6			
Final estimate	-5418.89	-5789.22	290.03	317.79	138.00
$\bar{\mathcal{E}}_{exp} - \bar{\mathcal{E}}_{pred}$	0.27	2.00	5.33	-11.74	-0.87
$(\bar{\mathcal{E}}_{exp} - \bar{\mathcal{E}}_{pred})/ \bar{\mathcal{E}}_{exp} *100$	0.00	0.03	1.80	-3.84	-0.63

^a illite component: $K_{0.683}Ca_{0.084}(Si_{3.319}Al_{0.681})(Al_{1.777}Mg_{0.234}Fe^{2+}_{0.016})O_{10}(OH)_2$
smectite component: $Ca_{0.107}Si_4(Al_{1.652}Mg_{0.292}Fe^{3+}_{0.081}Fe^{2+}_{0.003})O_{10}(OH)_2$.

As stated by Blanc *et al.* [24], the mixing of layers introduces an additional entropy term, which is not directly measurable by calorimetry. The authors proposed a prediction model based on nearest neighbor interaction up to the fourth neighbor. In a simpler approach, we can first assume that the mixing entropy term is estimated for an ideal solid solution between illite and smectite end-members. In the case of the ISCz-1 sample with 70% illite, $\Delta S_m = 5.38 \text{ J K}^{-1} \text{ mol}^{-1}$.

In Fig. 8, the stability fields for the ISCz-1, Illite and Montmorillonite end-members are drawn at 25°C by considering two different sets of thermodynamic parameters for ISCz-1:

- a set including ΔH_m for the ISCz-1 properties of formation (Fig. 8a), and
- a set excluding ΔH_m for the ISCz-1 properties of formation (Fig. 8b).

Comparing Fig. 8a and b highlights the importance of mixing terms. Indeed, the ISCz-1 stability field just disappears when the mixing terms are not integrated to the thermodynamic function, whereas it lies as expected between the illite and montmorillonite end-members when the mixing terms are added to the ISCz-1 formation properties. Overall, in the calculation process followed here to predict the thermodynamic properties of illite/smectite, we consider a mixture of layers with the composition of the montmorillonite component ranging between high- and low-charge montmorillonite, which is consistent with Meunier and Velde [2]. As seen from Fig. 8, a simple mixture is not sufficient, and additional mixing terms are necessary to result in significant stability, ranging as expected between illite and montmorillonite end-members.

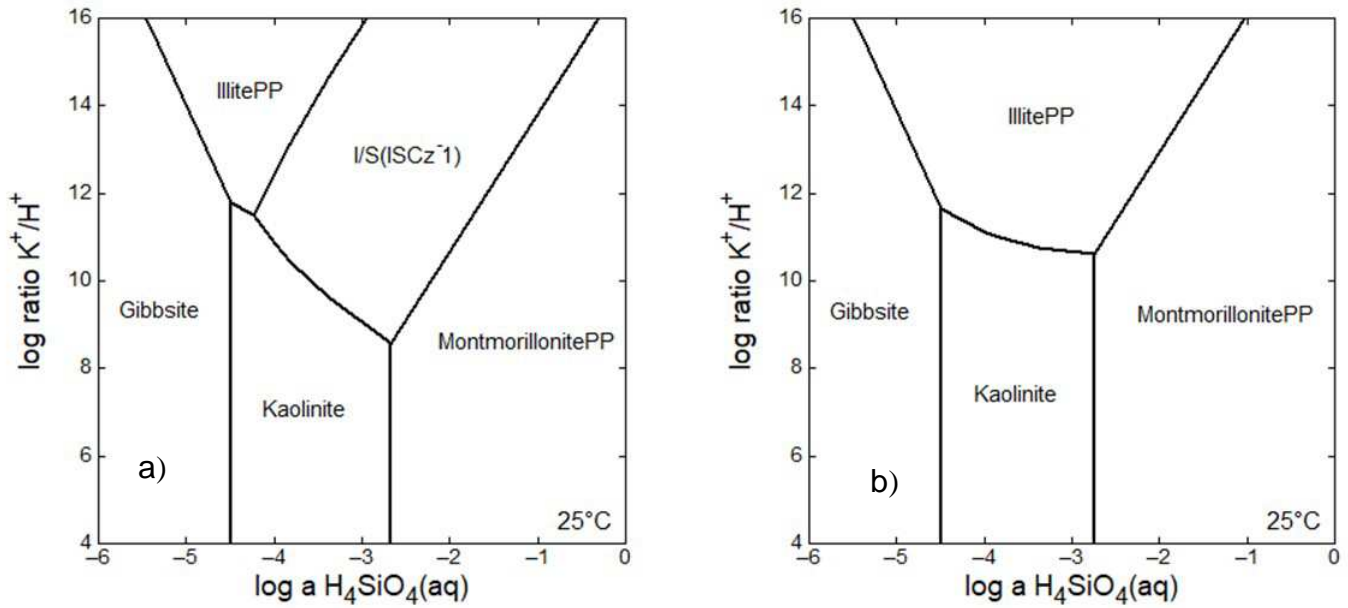


Fig. 8. Stability fields for the I/S ISCz-1 mineral at 298.15 K and ambient pressure of 101.325 kPa: a) including enthalpy of mixing and b) without enthalpy of mixing.

5. Summary and conclusions

In early diagenesis and underground storage studies, the transformation of smectite into illite has long been recognized to involve the formation of interstratified illite/smectite as an intermediate product ([2,55]). However, the lack of thermodynamic properties has prevented it from being implemented into geochemical calculations. This study presents the first direct calorimetric measurements of such properties.

These measurements were conducted using a purified, Ca-saturated ISCz-1 sample from Slovakia, which was provided by the Source Clays Repository of the Clay Minerals Society. ΔH_f° was obtained from acid solution calorimetry experiments, and S° and $C_{p,m}^\circ$ were measured by the PPMS and DSC techniques. In addition, a specific mixing enthalpy was measured by acid solution calorimetry, considering samples from the Shinzan hydrothermal series [7]. For the entire set of measured thermodynamic properties, the comparison with a previously published estimation method (Blanc *et al.* [1]) was favorable, provided that the estimate is calculated by distinguishing between the respective contributions of the illite and smectite layers. The respective compositions of the latter are derived from the compositional relations established by Meunier and Velde [2].

Among the concluding points, comparing the different estimating methods shows that the energetic terms for layer mixing are necessary for the mineral stability, even though their magnitudes are small with respect to the formation terms and may even fall within the uncertainties obtained for the formation terms. This implies that further improvements and advances could be made by considering weak energy terms such as hydration or exchange reaction contributions [56].

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Appendix A. Heat capacity raw data measured on the (Ca-saturated) illite-smectite ISCz-1 sample

Table A-1. Heat capacities at constant pressure^a of the illite-smectite ISCz-1 sample, at dehydrated state, between 2 K and 520 K. Series 1 and 2: PPMS on sample 1 (4.156 mg); Series 3 and 4: PPMS on sample 2 (14.926 mg); Series 5: DSC (140.03 mg).

T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹
Series 1		Series 1 (continued)		Series 2	
50.62	0.0835	12.40	0.0029	51.98	0.0899
49.00	0.0786	11.99	0.0027	53.47	0.0930
47.37	0.0737	11.60	0.0025	55.00	0.0979
45.81	0.0692	11.21	0.0022	56.52	0.1029
44.31	0.0649	10.84	0.0021	58.04	0.1079
42.85	0.0608	10.49	0.0019	59.57	0.1133
41.51	0.0571	10.14	0.0018	61.07	0.1183
40.14	0.0533	9.81	0.0016	62.59	0.1234
38.82	0.0498	9.49	0.0015	64.10	0.1284
37.52	0.0464	9.18	0.0014	65.61	0.1334
36.27	0.0433	8.88	0.0013	67.12	0.1385
35.07	0.0402	8.59	0.0012	68.63	0.1437
33.91	0.0373	8.31	0.0011	70.14	0.1490
32.80	0.0346	8.03	0.0011	71.65	0.1533
31.71	0.0320	7.76	0.0010	73.15	0.1585
30.66	0.0296	7.51	0.0009	74.66	0.1639
29.62	0.0273	7.26	0.0009	76.18	0.1692
28.64	0.0252	7.02	0.0008	77.69	0.1746
27.69	0.0232	6.78	0.0008	79.21	0.1801
26.77	0.0214	6.56	0.0007	80.72	0.1856
25.89	0.0196	6.35	0.0007	82.24	0.1910
25.06	0.0180	6.14	0.0006		
24.24	0.0166	5.72	0.0006		
23.44	0.0152	4.86	0.0005		
22.67	0.0139	4.11	0.0004		
21.86	0.0127	3.96	0.0004		
21.19	0.0117	3.83	0.0004		
20.45	0.0107	3.70	0.0004		
19.83	0.0098	3.58	0.0004		
19.17	0.0090	3.46	0.0004		
18.54	0.0083	3.35	0.0004		
17.91	0.0076	3.24	0.0004		
17.33	0.0070	3.13	0.0004		
16.76	0.0064	3.03	0.0004		
16.21	0.0059	2.93	0.0004		
15.66	0.0054	2.84	0.0004		
15.15	0.0049	2.75	0.0004		
14.66	0.0045	2.66	0.0004		
14.16	0.0041	2.57	0.0004		
13.70	0.0038	2.49	0.0004		
13.25	0.0034	2.41	0.0004		
12.82	0.0032	2.33	0.0004		

T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹
Series 2 (continued)		Series 2 (continued)		Series 2 (continued)	
83.75	0.1966	175.31	0.5030	289.10	0.7514
85.27	0.2018	177.30	0.5079	291.10	0.7537
86.78	0.2070	179.29	0.5152	293.12	0.7570
88.28	0.2125	181.30	0.5204	295.12	0.7594
89.80	0.2179	183.31	0.5269	297.14	0.7623
91.31	0.2232	185.31	0.5318	299.17	0.7628
92.82	0.2286	187.30	0.5369	301.17	0.7628
94.34	0.2342	189.29	0.5438	303.19	0.7668
95.86	0.2406	191.29	0.5481	305.22	0.7678
97.37	0.2461	193.29	0.5537	307.22	0.7710
98.88	0.2515	195.30	0.5589	309.24	0.7719
100.42	0.2571	197.29	0.5644	311.28	0.7804
101.91	0.2628	199.29	0.5699	311.56	0.7796
103.44	0.2681	201.27	0.5759	313.53	0.7790
104.95	0.2736	203.28	0.5814	315.76	0.7841
106.47	0.2790	205.28	0.5870	317.98	0.7940
107.98	0.2845	207.27	0.5917	320.22	0.7908
109.48	0.2897	209.26	0.5965	322.45	0.7999
111.01	0.2952	211.26	0.6041	324.70	0.8043
112.51	0.3005	213.25	0.6071	326.92	0.8055
114.04	0.3060	215.25	0.6103	329.20	0.8060
115.53	0.3113	217.24	0.6162	331.44	0.8086
117.05	0.3166	219.26	0.6200		
118.57	0.3217	221.26	0.6247	Series 3	
120.08	0.3277	223.26	0.6313	51.39	0.0869
121.59	0.3324	225.25	0.6324	49.77	0.0819
123.09	0.3379	227.25	0.6378	48.18	0.0772
124.60	0.3430	229.24	0.6430	46.62	0.0726
126.11	0.3484	231.26	0.6475	45.11	0.0682
127.62	0.3531	233.26	0.6531	43.65	0.0640
129.15	0.3582	235.26	0.6560	42.26	0.0600
130.65	0.3635	237.23	0.6624	40.89	0.0561
132.17	0.3681	239.23	0.6648	39.56	0.0526
133.69	0.3733	241.22	0.6693	38.29	0.0492
135.22	0.3788	243.22	0.6725	37.08	0.0461
136.74	0.3836	245.23	0.6786	35.88	0.0430
138.25	0.3884	247.22	0.6861	34.74	0.0400
139.77	0.3936	249.22	0.6881	33.62	0.0372
141.28	0.3984	251.21	0.6939	32.54	0.0346
142.81	0.4031	253.18	0.6969	31.50	0.0321
144.32	0.4082	255.17	0.7014	30.51	0.0299
145.84	0.4127	257.17	0.7050	29.55	0.0277
147.34	0.4180	259.17	0.7094	28.60	0.0256
148.85	0.4229	261.17	0.7102	27.69	0.0237
150.35	0.4284	263.16	0.7157	26.81	0.0219
151.88	0.4323	265.16	0.7193	25.96	0.0202
153.38	0.4374	267.14	0.7227	25.14	0.0186
155.44	0.4436	269.13	0.7251	24.34	0.0171
157.34	0.4499	271.13	0.7289	23.58	0.0158
159.33	0.4550	273.13	0.7331	22.85	0.0146
161.32	0.4617	275.14	0.7330	22.16	0.0134
163.34	0.4650	277.14	0.7373	21.48	0.0124
165.33	0.4724	279.13	0.7445	20.82	0.0115
167.33	0.4793	281.12	0.7455	20.18	0.0106
169.31	0.4850	283.10	0.7458	19.57	0.0097
171.30	0.4912	285.10	0.7514	18.96	0.0090
173.32	0.4962	287.11	0.7489	18.38	0.0083

T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹	T/K^b	C_p^c /J K ⁻¹ g ⁻¹
Series 3 (continued)		Series 4 (continued)		Series 4 (continued)	
17.81	0.0077	58.08	0.1090	142.72	0.4042
17.29	0.0071	59.60	0.1142	144.22	0.4094
16.77	0.0066	61.11	0.1193	145.73	0.4141
16.26	0.0061	62.63	0.1244	147.23	0.4191
15.78	0.0056	64.13	0.1294	148.74	0.4240
15.32	0.0052	65.66	0.1346	150.25	0.4285
14.86	0.0048	67.14	0.1396	151.76	0.4331
14.44	0.0045	68.65	0.1448	153.26	0.4380
14.03	0.0041	70.16	0.1501	155.35	0.4446
13.63	0.0038	71.66	0.1554	157.20	0.4501
13.24	0.0036	73.18	0.1602	159.18	0.4558
12.88	0.0033	74.69	0.1655	161.17	0.4616
12.53	0.0031	76.19	0.1709	163.16	0.4678
12.19	0.0029	77.70	0.1764	165.13	0.4736
11.87	0.0027	79.22	0.1819	167.12	0.4790
11.55	0.0025	80.72	0.1872	169.10	0.4851
11.26	0.0023	82.24	0.1928	171.11	0.4909
10.97	0.0022	83.76	0.1983	173.09	0.4964
10.69	0.0021	85.26	0.2037	175.07	0.5022
10.42	0.0019	86.77	0.2090	177.06	0.5078
10.16	0.0018	88.29	0.2143	179.05	0.5138
9.91	0.0017	89.80	0.2195	181.06	0.5193
9.66	0.0016	91.30	0.2251	183.05	0.5252
9.42	0.0015	92.82	0.2305	185.02	0.5307
9.21	0.0015	94.33	0.2364	187.02	0.5362
9.00	0.0014	95.84	0.2420	189.01	0.5417
8.78	0.0013	97.35	0.2475	191.01	0.5470
8.59	0.0012	98.87	0.2531	193.00	0.5522
8.40	0.0012	100.39	0.2585	195.00	0.5575
8.22	0.0011	101.90	0.2641	196.98	0.5627
8.05	0.0011	103.41	0.2694	198.98	0.5684
6.45	0.0007	104.92	0.2749	201.00	0.5737
4.16	0.0004	106.44	0.2804	203.01	0.5791
4.02	0.0004	107.96	0.2857	205.02	0.5840
3.89	0.0004	109.46	0.2912	206.99	0.5889
3.77	0.0004	110.98	0.2965	209.01	0.5935
3.64	0.0004	112.49	0.3017	211.02	0.5985
3.53	0.0003	114.00	0.3071	213.04	0.6034
3.41	0.0003	115.50	0.3123	215.05	0.6075
3.30	0.0003	117.01	0.3177	217.05	0.6118
3.19	0.0003	118.53	0.3230	219.06	0.6173
3.08	0.0003	120.04	0.3286	221.05	0.6221
2.98	0.0003	121.54	0.3337	223.06	0.6253
2.88	0.0003	123.04	0.3388	225.07	0.6307
2.79	0.0003	124.54	0.3441	227.11	0.6343
2.69	0.0003	126.07	0.3492	229.13	0.6401
2.60	0.0003	127.57	0.3543	231.14	0.6442
2.51	0.0003	129.07	0.3593	233.12	0.6475
2.43	0.0003	130.59	0.3642	235.11	0.6523
2.34	0.0003	132.10	0.3696	237.11	0.6559
		133.63	0.3749	239.11	0.6601
		135.15	0.3798	241.12	0.6644
Series 4		136.67	0.3845	243.12	0.6688
51.98	0.0889	138.16	0.3897	245.12	0.6736
53.52	0.0939	139.68	0.3944	247.09	0.6764
55.04	0.0989	141.21	0.3995	249.08	0.6818
56.57	0.1039				

^c The relative expanded uncertainties with level of confidence 0.95 ($k = 2$) are: $U_r(C_p) = 0.05$ ($2 < T/K < 50$), $U_r(C_p) = 0.02$ ($50 < T/K < 270$), $U_r(C_p) = 0.05$ ($270 < T/K < 300$) and $U_r(C_p) = 0.015$ ($302 < T/K < 520$). Between 200 and 300 K, the uncertainties associated with PPMS measurements are estimated by taking into account the discrepancy between measured values and fitted data using the polynomial function given in Section 3.1.1. (see text for explanation).

Appendix B. References for heat capacity data of mineral impurities

Table B-1. References for C_p data of mineral impurities, used for the calculations of C_p of the studied clay minerals.

Mineral	References for C_p data	T range
Quartz	[57]	50 K -520 K
Kaolinite	[58]	7 K – 298 K
	[59]	298.15 K – 520 K
K-Feldspar (microcline)	[60]	5 K – 298 K
	[59]	298.15 K – 520 K
Goethite	[61]	0.4 K – 375 K

Appendix C. Standard enthalpies of formation of the secondary references (constituents) at 298.15 K and 101.325 kPa, used in the solution calorimetry experiments

Table C-1. Standard enthalpies of formation ΔH_f° of the secondary references, at 298.15 K and 101.325 kPa.

Name	Formula	ΔH_f° /kJ mol ⁻¹	Reference
Gibbsite	Al(OH) ₃	-1293.13 (\pm 1.2)	[62]
Quartz	SiO ₂	-910.7 (\pm 1.7)	[57]
Goethite	FeOOH	-560.7 (\pm 1.2)	[61]
Iron(II) oxide	FeO	-272.04 (\pm 2.1)	[54]
Calcium hydroxide	Ca(OH) ₂	-984.55 (\pm 1.3)	[63]
Magnesium hydroxide	Mg(OH) ₂	-924.14 (\pm 0.4)	[64]
Potassium nitrate	KNO ₃	-494.5 (\pm 0.4)	[59]
Rutile	TiO ₂	-944.0 (\pm 0.8)	[53]

All the thermodynamic data for the secondary references come from Thermoddem database (<http://thermoddem.brgm.fr>) except for KNO₃, not available in the database.

Appendix D. Supplementary data

Supplementary data to this article can be found online at <https://doi.org/10.1016/j.jct.2019.06.004>.

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