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Article

High Catalytic Activity of Heterometallic (Fe₆Na₇ and Fe₆Na₆) Cage Silsesquioxanes in Oxidations with Peroxides

Alexey I. Yalymov¹, Alexey N. Bilyachenko^{1,2,*}, Mikhail M. Levitsky¹, Alexander A. Korlyukov^{1,3}, Victor N. Khrustalev^{1,2}, Lidia S. Shul'pina¹, Pavel V. Dorovatovskii⁴, Marina A. Es'kova¹, Frédéric Lamaty^{5,*}, Xavier Bantreil⁵, Benoît Villemejeanne⁵, Jean Martinez⁵, Elena S. Shubina¹, Yuriy N. Kozlov^{6,7} and Georgiy B. Shul'pin^{6,7,*}

¹ A. N. Nesmeyanov Institute of Organoelement Compounds, Russian Academy of Sciences, Vavilov Str., 28, Moscow 119991, Russia; alexyalymov@gmail.com (A.I.Y.); levitsk@ineos.ac.ru (M.M.L.); alex@xrlab.ineos.ac.ru (A.A.K.); vnkhrustalev@gmail.com (V.N.K.); shulpina@ineos.ac.ru (L.S.S.); marinaeskovskaya@gmail.com (M.A.E.); shu@ineos.ac.ru (E.S.S.)

² Inorganic Chemistry Department, People's Friendship University of Russia (RUDN University), Miklukho-Maklay Str., 6, Moscow 117198, Russia

³ Chemistry department, Therapeutic faculty, Pirogov Russian National Research Medical University, Ostrovitianov Str., 1, Moscow 117997, Russia

⁴ National Research Center "Kurchatov Institute", Akademika Kurchatova pl., 1, Moscow 123182, Russia; paulgemini@mail.ru

⁵ Institut des Biomolécules Max Mousseron (IBMM) UMR 5247, CNRS, Université de Montpellier, ENSCM, Université de Montpellier Campus Triolet Place Eugène Bataillon, Montpellier CEDEX 5, 34095, France; xavier.bantreil@umontpellier.fr (X.B.); benoit.villemejeanne@enscm.fr (B.V.); martinez@univ-montp1.fr (J.M.)

⁶ Semenov Institute of Chemical Physics, Russian Academy of Sciences, Ulitsa Kosygina, dom 4, Moscow 119991, Russia; yunkoz@mail.ru

⁷ Chair of Chemistry and Physics, Plekhanov Russian University of Economics, Stremyannyi pereulok, dom 36, Moscow 117997, Russia

* Correspondence: bilyachenko@ineos.ac.ru (A.N.B.); frederic.lamaty@umontpellier.fr (F.L.); shulpin@chph.ras.ru (G.B.S.); Tel.: +7-499-135-9369 (A.N.B.); +33-4-6714-3847 (F.L.); +7-495-939-7317 (G.B.S.)

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Abstract: Two types of heterometallic (Fe(III),Na) silsesquioxanes—[Ph₅Si₅O₁₀]₂[Ph₁₀Si₁₀O₂₁]₂Fe₆(O²⁻)₂Na₇(H₃O⁺)(MeOH)₂(MeCN)_{4.5}.1.25(MeCN), **I**, and [Ph₅Si₅O₁₀]₂[Ph₄Si₄O₈]₂Fe₆Na₆(O²⁻)₃(MeCN)_{8.5}(H₂O)_{8.44}, **II**—were obtained and characterized. X-ray studies established distinctive structures of both products, with pair of Fe(III)-O-based triangles surrounded by siloxanolate ligands, giving fascinating cage architectures. Complex **II** proved to be catalytically active in the formation of amides from alcohols and amines, and thus becoming a rare example of metallasilsesquioxanes performing homogeneous catalysis. Benzene, cyclohexane, and other alkanes, as well as alcohols, can be oxidized in acetonitrile solution to phenol—the corresponding alkyl hydroperoxides and ketones, respectively—by hydrogen peroxide in air in the presence of catalytic amounts of complex **II** and trifluoroacetic acid. Thus, the cyclohexane oxidation at 20 °C gave oxygenates in very high yield of alkanes (48% based on alkane). The kinetic behaviour of the system indicates that the mechanism includes the formation of hydroxyl radicals generated from hydrogen peroxide in its interaction with di-iron species. The latter are formed via monomerization of starting hexairon complex with further dimerization of the monomers.

Keywords: alkanes; amides; hydrogen peroxide; dinuclear complexes; iron complexes; metallasiloxanes

1. Introduction

Heterometallic complexes and clusters are among the most popular objects of contemporary chemistry, due to several remarkable features. First of all, the use of different metal ions is known as a key to controlled design of high-nuclearity complexes of picturesque molecular architecture [1–5]. Also, acting together, different metal ions provide catalytic activity in a wide range of processes [6–8], as well as intriguing magnetic properties [9]. The most popular synthetic approaches to such complexes involve use of organic ligands as well as application of “complex-as-ligand” tactics [10–13].

In turn, high reactivity and flexibility of siloxane ligands allow us to evaluate them as promising potential components of heterometallic complexes. Indeed, several heterometallic metallasiloxanes of attractive cage-like molecular geometry were described [14–17]. Importantly, Fe-containing siloxanes may be regarded as the most attractive representatives of metallasiloxanes, being artificial models of catalytically prospective silicates, zeolites, and iron oxides. Surprisingly, such complexes are still scarce in literature [18]. In this context, we were interested in the synthesis of new types of (Fe, M) siloxane geometries. As a pair of metal ions, a Fe/Na combination has been chosen because of the following reasons. It is known that sodium containing heterometallic cage siloxanes provide exceptional varieties of architecture [14–17], as well as catalytic activities [16,17,19,20] and magnetic (spin glass) properties [21–24]. It is explained by the participation of specific siloxanolate [RSi(O)ONa] ligand in cage construction, giving rise to multiple metallasiloxane architectures [17]. It is also noteworthy that several reports have discussed in detail the influence of reactants ratio and/or choice of solvent system on structural features of cage-like metallasiloxanes [14–17,25–27]. This tactic has been rarely used for Fe, Na-silsesquioxane design. Furthermore, a unique example of such architecture, namely Fe_6Na_8 compound featuring a Lantern shape, was synthesized as its butanol/toluene complex by some of us very recently [23]. The first results regarding the application of the approach “ratio/solvent choice” towards the synthesis of Fe, Na-silsesquioxanes are reported herein, along with catalytic studies of the obtained complex under oxidation and amidation conditions.

2. Results and Discussion

2.1. Syntheses and Structures of Catalysts

The synthesis of target Fe,Na-silsesquioxanes was performed by transformation of $\text{PhSi}(\text{OEt})_3$ into intermediate siloxanolate $[\text{PhSi}(\text{O})(\text{ONa})_n]$ species. Reactions of sodium siloxanolate with iron(III) chloride were carried out in various media (DMF, THF, DMSO, or 1,4-dioxane). All of these solvents already proved to be proper solvating ligands for metallasiloxane design [14–17,25–28]. Despite our expectations, isolation of a crystalline product in these reactions failed. However, the use of acetonitrile as a medium for synthesis/crystallization led to the isolation of unusual Fe,Na-phenylsilsesquioxane $\{[\text{Ph}_5\text{Si}_5\text{O}_{10}]_2[\text{Ph}_{10}\text{Si}_{10}\text{O}_{21}]\text{Fe}_6(\text{O}^{2-})_2\text{Na}_7(\text{H}_3\text{O}^+)(\text{MeOH})_2(\text{MeCN})_{4.5}\} \cdot 1.25(\text{MeCN})$ **I** in 16% yield (Figure 1).

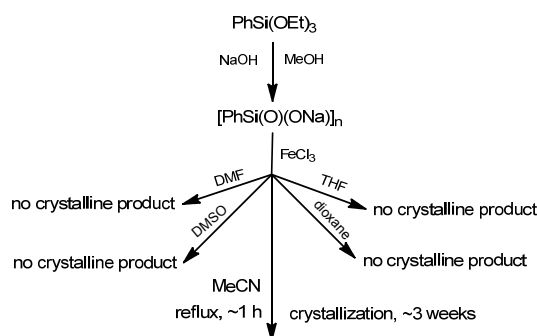


Figure 1. Scheme of solvent variation during synthesis of Fe,Na-phenylsilsesquioxanes. Use of MeCN provided **I**.

Taking this observation in mind, we focused on acetonitrile-containing media. Varying the ratio between reactants was found to be effective for changing the product framework. Formation of complex **II** (26% yield, Figure 2) was observed when a ~1/1.4/0.33 ratio between interacting silane/NaOH/FeCl₃ was used; while in the synthesis of **I**, ratio between reactants was ~1/1/0.25. The composition of product **II** differs from complex **I** and could be described as [Ph₅Si₅O₁₀]₂[Ph₄Si₄O₈]₂Fe₆Na₆(O²⁻)₃(MeCN)_{8.5}(H₂O)_{8.44}. Single crystal X-ray diffraction study revealed fascinating cage-like structures for both products (Figures 3 and 4).

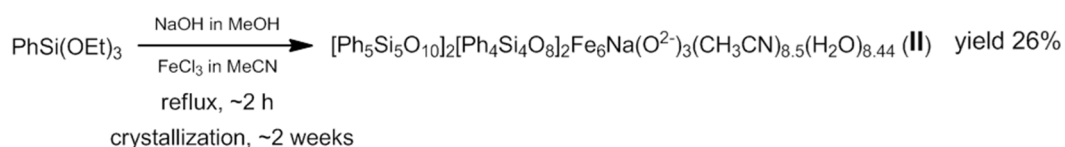


Figure 2. General scheme for the synthesis of **II**.

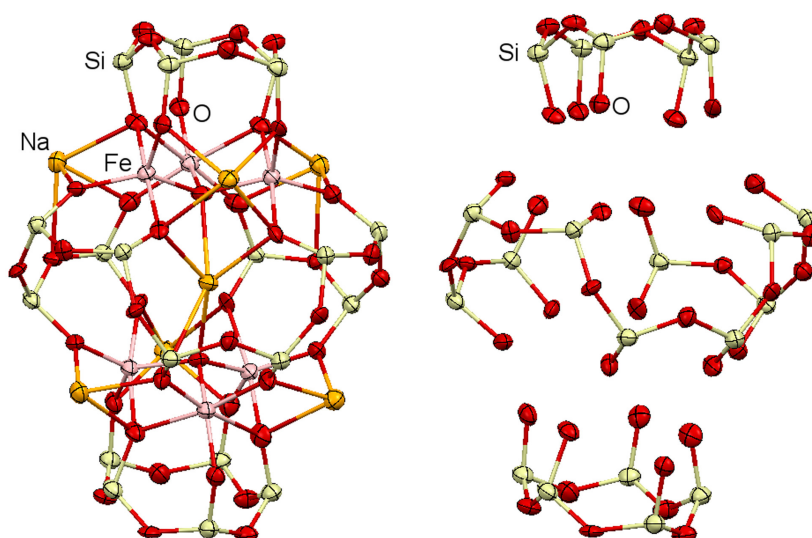


Figure 3. Molecular structures of $[\text{Ph}_5\text{Si}_5\text{O}_{10}]_2[\text{Ph}_{10}\text{Si}_{10}\text{O}_{21}]\text{Fe}_6(\text{O}^{2-})_2\text{Na}_7(\text{H}_3\text{O}^+)(\text{MeOH})_2(\text{MeCN})_{4.5} \cdot 1.25(\text{MeCN})$ **I** (left) and its silsesquioxane ligands (right).

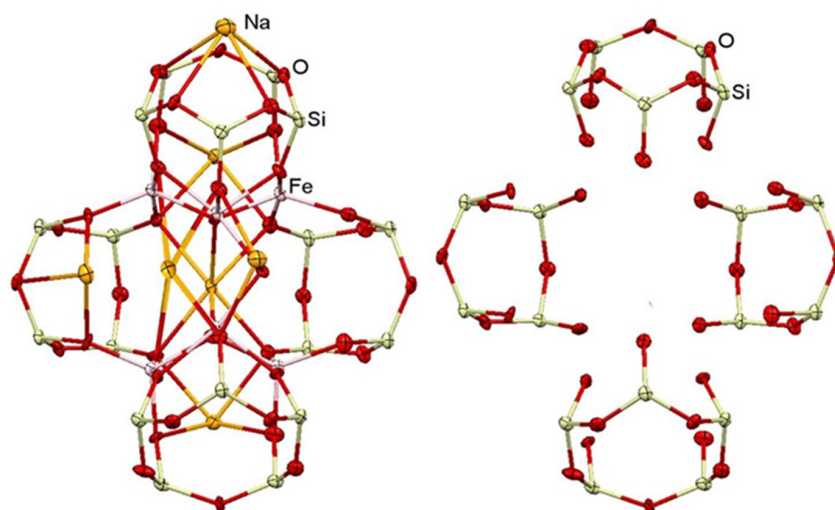


Figure 4. Molecular structures of $[\text{Ph}_5\text{Si}_5\text{O}_{10}]_2[\text{Ph}_4\text{Si}_4\text{O}_8]_2\text{Fe}_6\text{Na}_6(\text{O}^{2-})_3(\text{CH}_3\text{CN})_{8.5}(\text{H}_2\text{O})_{8.44}$ **II** (left) and its silsesquioxane ligands (right).

The most attractive feature of compounds **I** and **II** is the nature of their silsesquioxane ligands. Compound **I** includes two five-membered cyclic ligands of composition $[\text{Ph}_5\text{Si}_5\text{O}_{10}]$ and an acyclic ten-membered belt of composition $[\text{Ph}_{10}\text{Si}_{10}\text{O}_{21}]$ (Figure 3 right). Symptomatically, the same combination of ligands was observed in a previously reported Fe_6Na_8 compound [23]. In turn, compound **II** includes only cyclic ligands, two four- and two five-membered ones, with compositions $[\text{Ph}_4\text{Si}_4\text{O}_8]$ and $[\text{Ph}_5\text{Si}_5\text{O}_{10}]$, respectively (Figure 4 right). The presence of four-membered cyclic ligand is an extremely rare feature for cage metallasilsesquioxanes. To the best of our knowledge, we could cite only di- and tetranuclear Ti(IV)-containing compounds, obtained from cyclotetrasiloxanetetraols [29]. Compound **II** is thus the first instance of the simultaneous presence of four- and five-membered cyclic ligands in a metallasilsesquioxane structure.

A common feature of complexes **I** and **II** is the presence of six iron(III) centers. These are combined into two trinuclear metal oxo clusters $[\text{Fe}_3\text{O}_{12}]^{15-}$, including two penta- and one hexacoordinated Fe(III) ions (Figure 5). The first observation of such clusters in metallasilsesquioxane structure was reported by some of us [23]. It is noteworthy that the locations of trinuclear units in cages of **I** and **II** are quite different. In the case of compound **I**, these clusters are “independent”, connected through siloxane bonds (Figure 5), with the shortest contact between iron ions from different trinuclear units equal to 5.66 Å. On the other hand, trinuclear fragments of **II** are connected straight through bridging oxygen atoms (Figure 5), with the shortest Fe-Fe contact equal to 3.13 Å. In the case of compounds **I** and **II**, such rearrangement results in the formation of $[\text{Fe}_3\text{O}_{12}]^{15-}$ units; this is most probably explained by the high stability of such trinuclear geometry. To some extent, that statement could be confirmed by the observation of the same clusters in the composition of some other complexes [30–34].

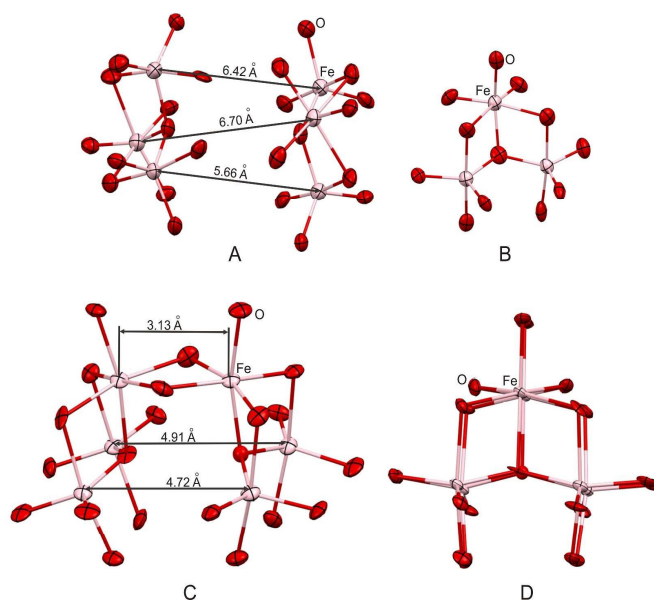


Figure 5. Pairs of trinuclear clusters in the frameworks of **I** (A) and **II** (C); General view of trinuclear iron oxo cluster in **I** (B) and **II** (D).

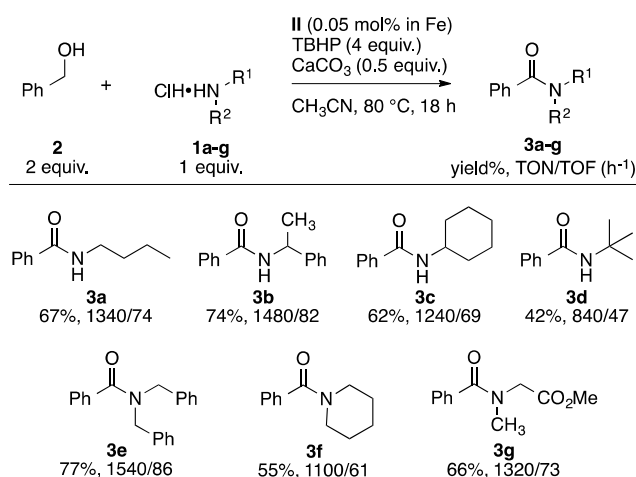
In our opinion, the appearance of such trinuclear $[\text{Fe}_3\text{O}_{12}]^{15-}$ clusters in the structures of **I** and **II** deserves additional discussion. Formal logic of metallasilsesquioxane synthesis implies formation of Si-O-M units by the reaction of silanolate Si-O-Na with metal chloride M-Cl functional groups. Thus, formation of iron oxo M-O-M fragments could not be explained just by a reactants interaction. We suggest that such (M-O-M) structural units arise as a consequence of metallasilsesquioxane skeleton rearrangement in solution. Several examples of such processes for individual and oligomeric metallasilsesquioxanes have been summarized by some of us [35–37].

On the other hand, the $[\text{Fe}_x\text{O}_y]_n$ species could be formed in solution from FeCl_3 and the base that is always present in such media (OH^- in equilibrium with OEt^-). These ironoxo units might then be trapped by the siloxane species. It is clear that the mechanism of metallasilsesquioxane cage formation is still a questionable subject, and many factors—including the newest results on siloxanolate and silanols reactivity [38–40], as well as DFT estimated influence of solvents on CLMS' formation [41]—should be taken in consideration.

2.2. Catalytic Transformation of Alcohols and Amines into Amides

Only isolated examples of copper and iron catalysis involving silsesquioxane complexes were already reported by ourselves [23,42] and others [43] in the literature. Complex **II** featuring an innovative structure was thus evaluated in the direct formation of amides from alcohols and amines. Thanks to the good solubility of **II** in organic solvent, stock solution could be prepared and allowed us to work at low iron loading. Reactions of benzyl alcohol with various ammonium chlorides were performed using as low as 500 ppm of iron in the presence of *tert*-butylhydroperoxide (TBHP) as oxidant and calcium carbonate, in refluxing acetonitrile (Scheme 1). To our delight, primary and secondary amines reacted accordingly and corresponding amides could be obtained in yields up to 77% (compound **3e**). Steric hindrance had a strong influence on the amidation reaction, since *N-tert*-butyl benzamide **3d** was isolated in only 42% yield. Importantly, the turnover number (TON) and turnover frequency (TOF) values obtained herein—up to 1540 and 86 h^{-1} , respectively—outmatched the values reported in the literature, with $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$ in refluxing acetonitrile ($\text{TON} \leq 9$ and $\text{TOF} \leq 2.2 \text{ h}^{-1}$) [44], or under microwave irradiation ($\text{TON} \leq 16.8$ and $\text{TOF} \leq 33.6 \text{ h}^{-1}$) [45].

Optimization of the reaction conditions was performed in a previous publication dealing with a different iron complex [23]. The conditions were thus directly adapted to this new complex, in order to allow comparison between the results obtained with two different complexes. The minimum amount of TBHP to obtain satisfactory results in terms of reaction time and yield is 4 equivalents. Since 2 equivalents are required for oxidation, the other 2 equivalents might decompose during the course of the reaction at $80 \text{ }^\circ\text{C}$. In addition, since only 500 ppm of iron are used, recyclability was not envisioned on this scale (0.5 mmol of ammonium salt). Scaling-up of the reaction, in order to study the potential changes in structure of complex **II** during the catalytic cycle, as well as its recyclability, are currently ongoing in the laboratory.



Scheme 1. Evaluation of complex **II** in amide bond formation. Turnover number (TON) = (mmol of product)/(mmol of Fe). Turnover frequency (TOF) = TON/(reaction time), given in h^{-1} .

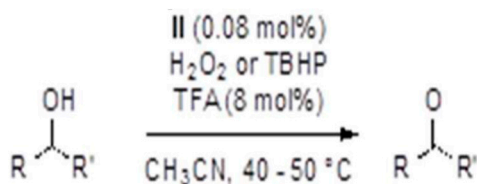
2.3. Catalytic Oxidation of Alcohols and Hydrocarbons








Even though several groups studied the iron-catalyzed oxidation of alcohols into amides, the precise mechanism is not yet completely elucidated. After oxidation of the alcohol into corresponding aldehyde, addition of the amine—generated in situ by reaction of the ammonium salt with poorly soluble calcium carbonate—onto the aldehyde and further oxidation of the hemiaminal would yield the desired amide. Due to the presence of iron and peroxide, it seems rather rational that radical species might be involved during the oxidation steps.

In turn, many mono or polynuclear iron-based compounds are known to be good catalysts for the oxidation of benzene, alcohols [46], and saturated and aromatic hydrocarbons [47–53] with peroxides. Oxygen-activating proteins, and especially enzymes containing polynuclear iron sites, attract a great deal of interest. Synthesized iron complexes are models of some enzymes with di-iron sites [54–56]. Methane monooxygenase (MMO) from methane-utilizing bacteria converts alkanes into the corresponding alcohols. Such enzymes oxidize regioselectively *n*-alkanes to afford predominately (in the case of *n*-heptane, even exclusively) 2-alcohols [57]. Compound **II** containing a polynuclear iron complex with chelating oxo-ligands exhibits some features similar to that of binuclear alkane oxygenases, and thus can be considered as an “inorganic alkane oxygenase”. In this context, behavior of compound **II** might be compared to our newest results concerning catalytic activity of Fe(III)-based silsesquioxane [23] and germaniumsesquioxane [58].

Complex **II** was found to be a very good catalyst in oxidations of alcohols, benzene, and alkanes with TBHP and H₂O₂. It is important to note that the reaction does not occur in the absence of trifluoroacetic acid. Gratifyingly, 1-phenylethanol and cyclooctanol could be converted into corresponding ketones in yields up to 92% and 85%, respectively, with only 0.08 mol% of catalyst **II** (Table 1).

Table 1. Oxidation of alcohols catalyzed by compound **II**.



Entry	Alcohol	Oxidant	Time (h)	Ketone Yield (%)
1		H ₂ O ₂	3	30
2		H ₂ O ₂	5	50
3		TBHP	4	80
4		TBHP	5	92
5		TBHP	3	71
6		TBHP	4	80
7		TBHP	5	85

Reaction conditions: Alcohol (0.6 M), oxidant (1.7 M for H₂O₂ or 1.5 M for *tert*-butylhydroperoxide [TBHP]), CF₃COOH (0.05 M), **II** (5 × 10^{−4} M), CH₃CN, 40 °C (entries 1, 2) or 50 °C (entries 3–7).

The oxidation of benzene into phenol, utilizing H₂O₂ (50%) in the presence of compound **II** in catalytic amounts, was also highly efficient. A rapid optimization regarding catalyst loading and temperature showed that the best conditions required 0.11 mol % of **II** in acetonitrile at 50 °C (Figure 6). Interestingly, under these conditions, maximum TON of 385 could be obtained in 6 h. The inset B of Figure 7 shows a saturation profile for the initial rate of phenol production vs. catalyst concentration. This behavior is typical of an enzyme-like mechanism involving a rapid binding of the substrate.

The oxidation of cyclohexane, which is especially attractive and challenging, was studied in more detail and followed by the GC. Moreover, cyclohexane gives a minimum number of oxidation products which are easily identified by the GC method. As demonstrated previously

in other oxidations [19,59–68], if the direct injection of a reaction sample into the chromatograph gave comparable amounts of cyclohexanol and cyclohexanone, the reduction of the sample with PPh_3 (or certain sulfides) prior to GC analysis led to the noticeable predominance of the alcohol in many cases (Figure 8). The comparison of the results obtained before and after the reduction clearly indicated that cyclohexyl hydroperoxide was formed as the main primary product. The oxidation by the **II**/ H_2O_2 / CF_3COOH system was very efficient because it gave alkane oxidation products in a high GC yield of 48% (TON = 440) after 3 h at 20 °C.

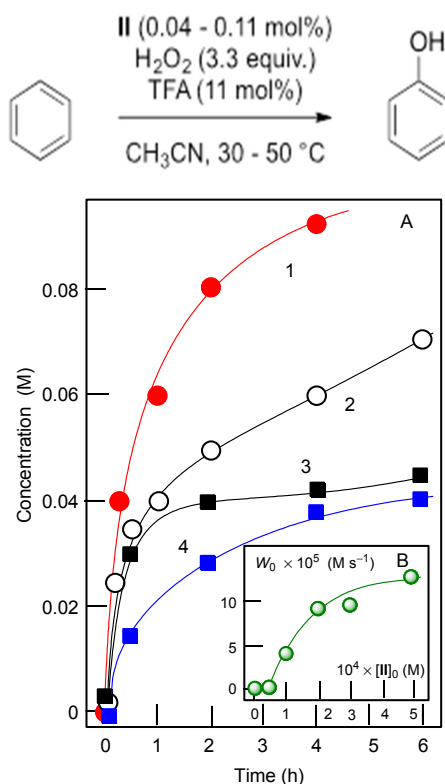


Figure 6. (Graph A) Accumulation of phenol with time. Benzene (0.46 M), H_2O_2 (50%, aqueous, 1.5 M), catalyst **II** (5×10^{-4} M for curve 1; 2×10^{-4} M for curves 2, 3 and 4), CF_3COOH (0.05 M) in CH_3CN (total volume of the reaction solution was 5 mL); temperature was 30 °C (curves 3 and 4) or 50 °C (curves 1 and 2); (Graph B) Dependence of initial phenol accumulation rate W_0 on initial concentration of catalyst $[\text{II}]_0$.

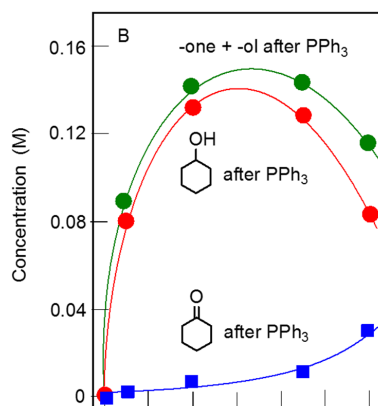


Figure 7. Cont.

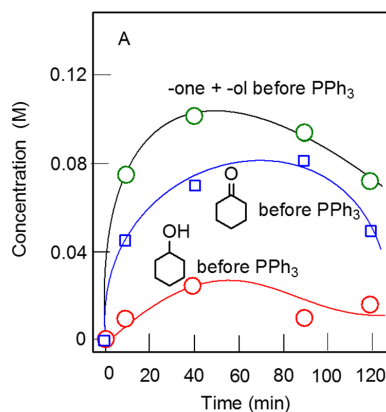


Figure 7. Accumulation of cyclohexanol and cyclohexanone with time. Cyclohexane (0.46 M), H_2O_2 (50%, aqueous, 1.5 M), catalyst **II** (5×10^{-4} M), CF_3COOH (0.05 M), CH_3CN (total volume of the reaction solution was 5 mL), 40°C . Concentrations measured by the GC method before (Graph A) and after (Graph B) reduction of the samples with PPh_3 are shown (for this method, see References [19,59–68]).

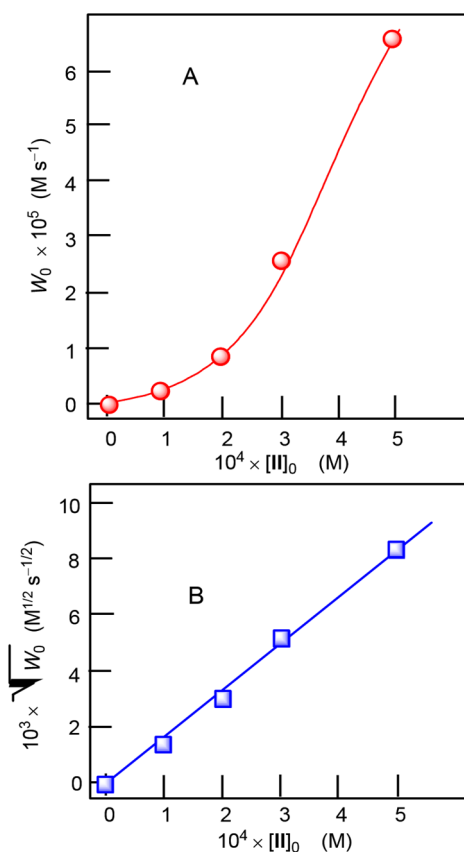


Figure 8. Dependence of the initial rate of formation cyclohexane oxidation products (the sum cyclohexanol + cyclohexanone) W_0 on initial concentration of catalyst **II** in the oxidation of cyclohexane (0.46 M) with hydrogen peroxide (50% aqueous, 1.5 M), catalyzed by compound **II** in the presence of CF_3COOH (0.05 M) in MeCN at 40°C (Graph A). Graph B: this dependence in coordinates “ $W_0^{1/2}$ vs. initial concentration of catalyst **II**”. Concentrations of cyclohexanone and cyclohexanol were determined by the GC method after reduction of the aliquots with solid PPh_3 .

The mode of dependence of the initial cyclohexane oxidation rate W_0 on concentration of catalyst **II** (Figure 9, Graph A) in the oxidation with hydrogen peroxide indicates that the rate of dependency is second order with respect to the initial concentration of **II**. Indeed, the proportional dependence of parameter $W_0^{1/2}$ on $[\mathbf{II}]_0$ is presented by the straight line (Figure 9, Graph B). On the one hand, it is not probable that the quadratic dependence of W_0 on $[\mathbf{II}]_0$ is due to the dimerization of starting complexes containing six iron ions. It is important to note that dependence of W_0 on $[\mathbf{II}]_0$ for the oxygenation of benzene to phenol (see Figure 7, Graph B) is also a quadratic one.

The role of added TFA in oxidations of alkanes and alcohols apparently was to split some bonds in the precatalyst cage, resulting in the formation of coordinatively unsaturated species active in H_2O_2 decomposition. In order to get additional insight into the mechanism of the alkane oxidation with the system under consideration, we carried out two experiments with cyclohexane. In the first experiment, we studied absorption spectra in the UV-visible region (30×10^3 – $13 \times 10^3 \text{ cm}^{-1}$) under conditions that were similar to conditions of the kinetic oxidation experiments. Figure 10, Graph A demonstrates the absorption of complex **II** (in CH_3CN ; $[\mathbf{II}] = 2.7 \times 10^{-4} \text{ M}$; curve 1). This absorption grows significantly when CF_3COOH is added to the solution ($[\text{TFA}] = 0.05 \text{ M}$, curve 2). If H_2O_2 (total concentration 1.3 M, containing $[\text{H}_2\text{O}] = 2.4 \text{ M}$) is added to this acidified solution, the absorption decreases (curve 3). Addition of cyclohexane (0.46 M) remains virtually the same spectrum (curve 4). Figure 10, Graph B corresponds to the spectrum obtained for higher concentration of initial complex **II** ($5.3 \times 10^{-4} \text{ M}$, curve 1). In the presence of TFA (0.05 M), the absorption is stronger (curve 2). Addition of water ($[\text{H}_2\text{O}]_{\text{added}} = 4.9 \text{ M}$) shifts curve 2 to the field of lower wavelength (curve 3). Thus, obtained data indicate that addition of an acid strongly affects the absorption of the starting complex which can be due to certain changes of its structure, particularly the monomerization of initial hexameric complexes. The effect of the addition of H_2O_2 is similar to the influence of the additive of H_2O (compare Figure 9, Graph A, curve 3, and Figure 9, Graph B, curve 3). It may be concluded that the changes in the presence of H_2O_2 are mainly due to the water which is introduced into the reaction solution simultaneously with hydrogen peroxide (50% aqueous). As expected, addition of cyclohexane does not affect the catalyst (compare curves 3 and 4 in Figure 9, Graph A). Absorption of the catalyst is not practically changed in the course of the oxidation reaction; at least, in the first 90 min (compare curves 1 and 2 in Figure 9, Graph C). A small difference can be detected only after 180 min when H_2O_2 is practically deceased (see below, Figure 10).

We carried out the second experiment in order to determine stability and activity of the complex **II** during the course of cyclohexane oxidation. At the moment corresponding to the maximum concentration of formed oxygenates, when the oxidant concentration is low (denoted by an arrow in Figure 10), an additional portion of hydrogen peroxide was added. We see that the oxidation of cyclohexane restarts with the rate equal to the rate noticed in the beginning of the reaction. It can be concluded that, in accordance with the kinetic scheme given above, the dimeric iron complexes generated in the system from monomers take part in the catalytic decomposition of hydrogen peroxide.

Finally, it is necessary to note that complex **II** containing siloxane ligands is a much more efficient catalyst in the cyclohexane oxidation, in comparison with simple iron salts. Thus, if the GC yield of 48% (TON = 440) was attained after 3 h at 20 °C in the **II**-catalyzed reaction, the oxygenate GC yield in the presence of $\text{Fe}(\text{NO}_3)_3$ under the same conditions was not higher than 1%–3%. Complex **II** catalyzes the oxidation of normal heptane and methylcyclohexane (Figures S1 and S2).

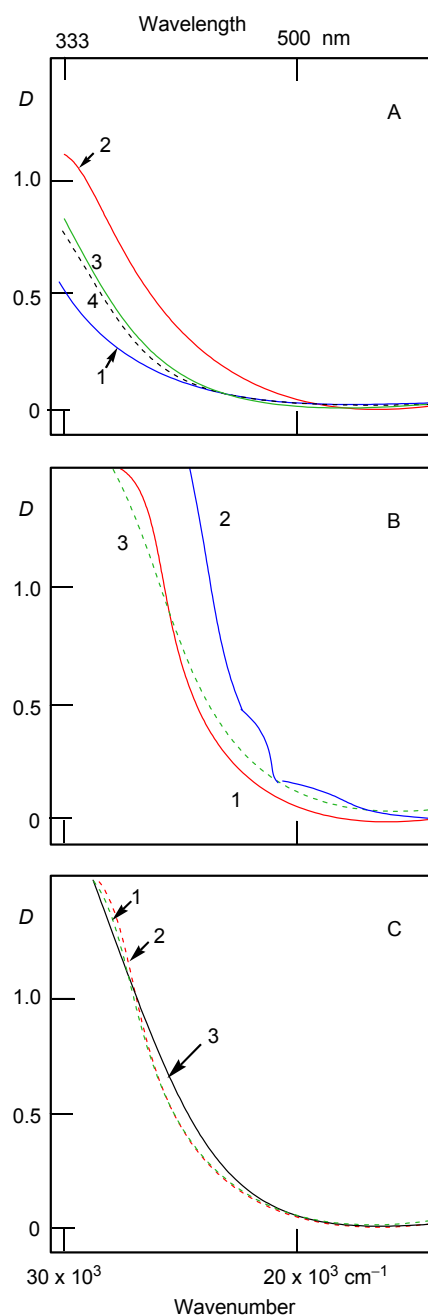


Figure 9. Electronic spectra of the pre-catalyst **II** solution in acetonitrile at 25 °C in the presence of various additives. (Graph **A**) Curve 1: a solution of **II** (2.7×10^{-4} M), H_2O_2 , and cyclohexane in MeCN (2.5 mL). Curve 2: the same solution after addition of TFA in MeCN (concentration of TFA 0.05 M). Curve 3: water ($[\text{H}_2\text{O}]_{\text{added}} = 4.9$ M) was added to the solution corresponding to curve 2. Curve 4: cyclohexane (0.46 M) was added to the solution corresponding to curve 3; (Graph **B**) Curve 1: a solution of **II** (5.3×10^{-4} M) in MeCN (2.5 mL). Curve 2: the same solution after addition of TFA in MeCN (concentration in the final solution was 0.05 M). Curve 3: hydrogen peroxide (50% aqueous, 1.3 M containing $[\text{H}_2\text{O}] = 2.4$ M) was added to the solution corresponding to curve 2; (Graph **C**) Curve 1: a solution of **II** (4.7×10^{-4} M), H_2O_2 and cyclohexane under conditions depicted by Figure 10, (Graph **A**) curve 4 in MeCN at the moment 5 min after preparation. Curve 2: The same solution after 90 (curve 2) and 180 min (curve 3) of incubation under conditions of cyclohexane oxidation experiments.

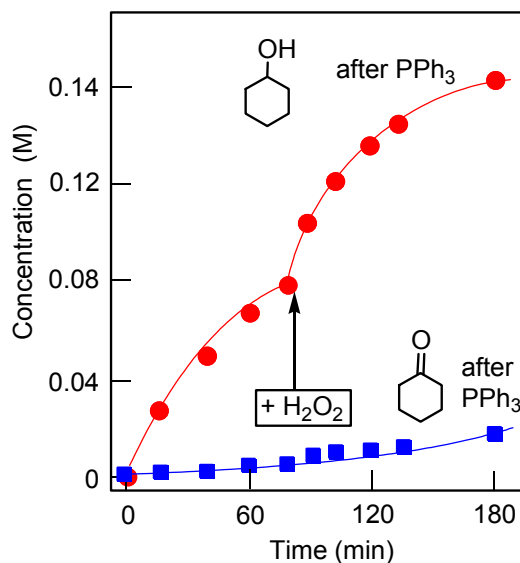


Figure 10. Accumulation of cyclohexanol and cyclohexanone with time. Conditions: Cyclohexane (0.46 M), H₂O₂ (50%, aqueous, 1.5 M), catalyst II (2.5×10^{-4} M), CF₃COOH (0.05 M), CH₃CN (total volume of the reaction solution was 5 mL), 40 °C. At the moment denoted by an arrow, an additional portion of hydrogen peroxide (the same amount as in the beginning of the reaction) was added. Concentrations were measured after reduction of the samples with PPh₃.

3. Materials and Methods

Starting compound PhSi(OEt)₃ and all solvents were purchased from Sigma Aldrich and were used as received. IR spectra were recorded on FTIR Shimadzu IR Prestige-21. IR spectrum in Nujol for solids and liquid solution in thin film were obtained using KBr discs.

3.1. Synthesis of Compound I

Compound PhSi(OEt)₃ (2 g, 8.32 mmol) and sodium hydroxide (0.333 g, 8.32 mmol) were dissolved in 45 mL of MeOH. After complete dissolution of sodium hydroxide, the mixture was heated at reflux for 2.5 h, and then iron(III) chloride (0.338 g, 2.08 mmol) in 60 mL of acetonitrile was added. The resulting brick-colored solution was additionally heated at reflux for 1 h and then cooled down to room temperature. Formation of a crystalline product with single crystals, useful for X-ray diffraction analysis (see below), was observed in solution after approximately three weeks. After ceasing of the crystal fraction growth, the solution was decanted and the solid fraction was dried in a vacuum without heating. Product I (0.22 g; 16% yield) was obtained.

Elemental analysis calcd. [(PhSiO_{1.5})₂₀(FeO_{1.5})₆(NaO_{0.5})₇]: Fe, 10.22; Na, 4.91; Si, 17.13. Found: Fe, 10.19; Na, 4.82; Si, 17.04.

3.2. Synthesis of Compound II

Compound PhSi(OEt)₃ (4 g, 16.64 mmol) and sodium hydroxide (0.96 g, 24 mmol) were dissolved in 30 mL of MeOH. After complete dissolution of sodium hydroxide, the mixture was heated at reflux for 2.5 h, and then iron(III) chloride (0.90 g, 5.55 mmol) in 100 mL of acetonitrile was added. The resulting brick-colored solution was heated at reflux for 2 h, then cooled down and filtered. Formation of a crystalline brick-colored product was observed in approximately two weeks. Several single crystals were used for X-ray diffraction analysis (see details below). After ceasing of the crystal fraction growth, the solution was decanted and the solid fraction was dried in a vacuum without heating. Product II (0.73 g, 26% yield) was obtained.

Elemental analysis calcd. [(PhSiO₂)₁₈Fe₆Na₆(O)₃]: Fe, 11.20; Na, 4.61; Si, 16.91. Found: Fe, 10.81; Na, 4.41; Si, 16.63.

3.3. X-ray Studies

The X-ray diffraction intensities of single crystals of compound **I** were measured at the Kurchatov Centre for Synchrotron radiation, while the dataset for **II** was collected with Bruker APEX DUO diffractometer. The structures were solved by direct method and refined in anisotropic approximation. Hydrogen atoms were calculated from a geometrical point of view, and were then refined with restraints applied for their displacement parameters and C–H (O–H) bond length. The crystal data for compounds **I** and **II** are summarized in Table 2 (see also the ESI).

Crystallographic data for **I** and **II** were submitted to CSD (CCDC 1481141 and CCDC 1481142) and can be obtained free of charge using web request from <http://www.ccdc.cam.ac.uk/request>.

Table 2. Results of X-ray experiments for complexes **I** and **II**.

Compound	I	II
Brutto formula	C _{133.50} H _{128.25} Fe ₆ N _{5.75} Na ₇ O ₄₆ Si ₂₀	C ₁₂₅ H _{117.96} Fe ₆ N _{8.5} Na ₆ O _{44.36} Si ₁₈
Formula weight	3606.99	3427.79
Wavelength, Å	0.96600	1.5418
T, K	100	120
Space group	P2 ₁ /n	Pn
Z	4	2
a, Å	27.130(5)	17.7227(10)
b, Å	18.150(4)	15.9140(9)
c, Å	32.830(7)	28.4984(17)
β, °	91.75(3)	100.136(3)
V, Å ³	16158(6)	7912.2(8)
ρ _{calc} , g·cm ^{−3}	1.483	1.439
μ, cm ^{−1}	17.86	64.35
F(000)	7402	3513
2θ _{max} , °	70.84	135.72
Reflections collected	234,993	20,450
Independent reflections	22,897	20,450
Independent reflections with I > 2σ(I)	18,123	17,065
Parameters	1681	1825
R ₁ [I > 2σ(I)]	0.1274	0.0925
wR ₂ (all reflections)	0.2763	0.2347
GOF	1.065	0.973
Residual electron density, e·Å ^{−3} (ρ _{min} /ρ _{max})	1.72/−1.10	2.08/−1.15

The main difficulty in the refinement of compounds **I** and **II** was to reveal the exact chemical composition of these structures. At first sight, positive and negative charges in these structures are imbalanced. This could be due to localization of counterions and hydrogen atoms attached to oxygen or nitrogen atoms due to disorder. In the case of **I**, we decided to treat the water molecule coordinated to Na as an oxonium cation H₃O⁺.

3.4. Oxidation of Alcohols and Hydrocarbons with Peroxides

The reactions of alcohols and hydrocarbons were usually carried out in air in thermostated Pyrex cylindrical vessels with vigorous stirring, using MeCN as solvent. Typically, catalyst **II** and the co-catalyst (acid) were introduced into the reaction mixture in the form of stock solutions in acetonitrile. The substrate (alcohol or hydrocarbon) was then added and the reaction started when hydrogen peroxide or TBHP was introduced in one portion. (CAUTION: The combination of air or molecular oxygen and H₂O₂ with organic compounds at elevated temperatures may be explosive). The reactions with benzene and 1-phenylethanol were analyzed by ¹H NMR method (solutions in acetone-*d*₆; “Bruker AMX-400” instrument, 400 MHz). For the determination of concentrations of phenol and quinone,

signals in the aromatic region were integrated using added 1,4-dinitrobenzene as a standard. Areas of methyl group signals were measured to quantify oxygenates formed in oxidations of 1-phenylethanol. In order to determine concentrations of all cyclohexane oxidation products, the samples of reaction solutions after addition of nitromethane as a standard compound were in some cases analyzed twice (before and after their treatment with PPh_3) by GC (LKhM-80-6 instrument, columns 2 m with 5% Carbowax 1500 on 0.25–0.315 mm Inerton AW-HMDS; carrier gas, argon) to measure concentrations of cyclohexanol and cyclohexanone. This method (an excess of solid triphenylphosphine was added to the samples 10–15 min before the GC analysis) was proposed by one of us earlier [19,59–68]. Attribution of peaks was made by comparison with chromatograms of authentic samples. Blank experiments with cyclohexane showed that, in the absence of catalyst **II**, no products were formed.

3.5. Selectivity in the Alkane Oxidations

In order to get an insight into the nature of oxidizing species, we measured the selectivity parameters in oxidations of certain linear, branched, and cyclic saturated hydrocarbons with H_2O_2 . The regioselectivity parameter [relative normalized reactivities of H atoms at carbon atoms C(1), C(2), C(3), and C(4) of *n*-octane chain] determined for the oxidation of *n*-octane is relatively low, i.e., C(1):C(2):C(3):C(4) = 1.0:6.7:6.6:6.1. It can be seen that hydrogen atoms in position 4 possess lower activity, apparently due to some sterical hindrance [50,57,60]. The bond-selectivity parameter ($1^\circ:2^\circ:3^\circ$; the relative normalized reactivities of hydrogen atoms at the primary, secondary, and tertiary carbons) in the oxidation of methylcyclohexane (1.0:6.7:17.5) is close to the corresponding values found for the systems oxidizing alkanes with hydroxyl radicals (see, for example, References [69–75]). The oxidation of *cis*-1,2-dimethylcyclohexane proceeds non-stereoselectively, because the *trans/cis* ratio [the ratio of isomers of *tert*-alcohols with mutual *trans*- and *cis*-orientation of two methyl groups] of isomeric alcohols (after reduction with PPh_3) was 0.8. The oxygenation of methylcyclohexane (MCH) with H_2O_2 proceeds mainly at the tertiary carbon atom with formation of 1-methylcyclohexanol after reduction with PPh_3 (product P5; see Supplementary Materials Figures S1 and S2). The GC profile of the products obtained in the **II**-catalyzed oxidation is very similar to the profiles reported previously for some other systems which oxidize with the participation of hydroxyl radicals (see Figure S2). All these data testify that an oxidizing species generated by the system exhibits a low selectivity typical for hydroxyl radicals.

3.6. General Procedure for Catalytic Amide Formation

In a sealed tube were added successively amine hydrochloride (0.5 mmol), CaCO_3 (25.0 mg, 0.25 mmol), CH_3CN (1 mL), **II** (50 μL of a solution of 2.8 mg of **II** in 1 mL of CH_3CN), benzylic alcohol (104 μL , 1.0 mmol), and TBHP (70% in H_2O , 140 μL , 1.0 mmol). The mixture was stirred at 80 °C for 2 h, and TBHP (70% in H_2O , 140 μL , 1.0 mmol) was again added to the mixture. After 16 h at 80 °C, the mixture was cooled to room temperature, and 1N HCl and AcOEt were added. The mixture was extracted twice with AcOEt, and the combined organic phase was washed with a saturated solution of NaHCO_3 , brine, and concentrated under reduced pressure. To remove the excess of benzylic alcohol, 80 mL of H_2O was added and evaporated under reduced pressure. Crude product was then purified using silica gel chromatography using gradients of cyclohexane/AcOEt to yield the pure compounds. The spectra of prepared amides are presented in the supplementary file ESI.

4. Conclusions

Two heterometallic (Fe_6Na_7) silsesquioxanes— $[\text{Ph}_5\text{Si}_5\text{O}_{10}]_2[\text{Ph}_{10}\text{Si}_{10}\text{O}_{21}]\text{Fe}_6(\text{O}^{2-})_2\text{Na}_7(\text{H}_3\text{O}^+)(\text{MeOH})_2(\text{MeCN})_{4.5} \cdot 1.25(\text{MeCN})$, **I**, and $[\text{Ph}_5\text{Si}_5\text{O}_{10}]_2[\text{Ph}_4\text{Si}_4\text{O}_8]_2\text{Fe}_6\text{Na}_6(\text{O}^{2-})_3(\text{MeCN})_{8.5}(\text{H}_2\text{O})_{8.44}$, **II**—were prepared using acetonitrile as a key reaction media. X-ray studies established the presence of Fe–O–Fe units in the composition of both products, which could be explained by additional rearrangement of metallasilsesquioxane skeletons before crystallization. A scheme of rearrangement is proposed. Compound **II** was found to be a highly active precatalyst in the oxidative amidation of

alcohols and amines. Amides could be isolated with TON/TOF values up to 1540/86 h⁻¹. Experiments on the oxidation of alcohols, benzene, and cyclohexane with the II/H₂O₂/CF₃COOH system were also very efficient. Importantly, this oxidative system was revealed to be particularly efficient for the oxidation of cyclohexane, yielding oxygenate derivatives in yield 48% and TON up to 440. Thus, this oxidation system is superior one because its activity much higher than the efficiency of oxidation catalyzed, for example, by nonanuclear Cu(II)-silsesquioxane, [(MeSiO_{1.5})₁₈(CuO)₉], reported by us very recently (total yield was 20%, TON 184) [76].

Supplementary Materials: The following are available online at www.mdpi.com/2073-4344/7/4/101/s1, X-ray Studies; Figure S1: Isomeric products formed in the methylcyclohexane oxidation; Figure S2: A chromatogram of products obtained in oxidations of methylcyclohexane by the “H₂O₂-II-CF₃COOH” system; Kinetic analysis of cyclohexane oxidation; Description of amides; References for the ESI.

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Conflicts of Interest: The authors declare no conflict of interest.

Abbreviations

TON	turnover number
TBHP	<i>tert</i> -butylhydroperoxide
GC	gas liquid chromatography
TFA	trifluoroacetic acid

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