



HAL
open science

Synergistic sorption of mixed solvents in wood cell walls: experimental and theoretical approach

Sonia Milena Aguilera-Segura, Julie Bossu, Stéphane Corn, Philippe Trens,
Tzonka Mineva, Nicolas Le Moigne, Francesco Di Renzo

► **To cite this version:**

Sonia Milena Aguilera-Segura, Julie Bossu, Stéphane Corn, Philippe Trens, Tzonka Mineva, et al..
Synergistic sorption of mixed solvents in wood cell walls: experimental and theoretical approach.
Macromolecular Symposia, 2019, 386 (1), pp.1900022. 10.1002/masy.201900022 . hal-02350667

HAL Id: hal-02350667

<https://hal.science/hal-02350667>

Submitted on 6 Nov 2019

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



**Synergistic sorption of mixed solvents in wood cell walls:
experimental and theoretical approach**

Journal:	<i>Macromolecular Symposia</i>
Manuscript ID	masy.201900022.R1
Wiley - Manuscript type:	Full Paper
Date Submitted by the Author:	n/a
Complete List of Authors:	Aguilera-Segura, Sonia Milena; Institut Charles Gerhardt Bossu, Julie; Institut Charles Gerhardt; Ecole des Mines d'Ales Trens, Philippe; Institut Charles Gerhardt Mineva, Tzonka; Institut Charles Gerhardt Di Renzo, Francesco; Institut Charles Gerhardt Le Moigne, Nicolas; Ecole des Mines d'Ales Corn, Stephane; Ecole des Mines d'Ales
Keywords:	composite materials, mixed solvents, sorption isotherms, viscoelastic properties, molecular dynamics

SCHOLARONE™
Manuscripts

Synergistic sorption of mixed solvents in wood cell walls: experimental and theoretical approach

Sonia Milena Aguilera-Segura, Julie Bossu, Stéphane Corn, Philippe Trens, Tzonka Mineva, Nicolas Le Moigne, Francesco Di Renzo**

S. M. Aguilera-Segura, Dr. J. Bossu, Prof. P. Trens, Dr. T. Mineva, Dr. F. Di Renzo
Institut Charles Gerhardt, UMR 5253 UM-CNRS-ENSCM, ENSCM, 240 Avenue Emile
Jeanbrau, 34296 Montpellier Cedex 05, France
direnzo@enscm.fr

Dr. S. Corn, Dr. N. Le Moigne
C2MA, IMT Mines Ales, Université de Montpellier, 6 avenue de Clavières 30319, Alès,
France
nicolas.le-moigne@mines-ales.fr

Keywords: composite materials, mixed solvents, sorption isotherms, viscoelastic properties, molecular dynamics

Abstract

The behavior of poplar wood and its components in water-ethanol mixtures is investigated through a coupled experimental and theoretical approach including physico-chemical and dynamic mechanical analyses and molecular dynamics simulations. Affinity for water-ethanol vapors, measured by isothermal gravimetric sorption experiments, features a maximum for mixed vapors. The longitudinal viscoelastic behavior of the same wood upon immersion in ethanol-water solutions, measured by dynamic mechanical analysis, features a minimum in storage modulus and a maximum damping upon immersion in solutions of intermediate composition. Optical microscopy observation of solvent-saturated samples evidences inter- and intra-cellular disbonding in pure ethanol and ethanol aqueous solutions. Molecular dynamics simulations provide information on interactions of water-ethanol solutions with models of cellulose microfibrils and lignin. The relative solvation of cellulose microfibrils by water and ethanol shows a nearly linear variation with the composition of the solution. In contrast, the accessibility of lignin dimers to the solvent presents a maximum at intermediate ethanol concentrations, in correspondence with a conformational transition of the dimer towards an extended conformation. The modelisation of the interactions of cellulose and lignin in water-ethanol solutions indicates a minimum of adhesion of the two components of wood in the presence of solutions with intermediate concentrations.

1. Introduction

1
2
3 How does the exposure to solvents affect the mechanical and chemical stability of composite
4 materials? The rise of polymer composites as structural materials in all technological fields
5 justifies a more and more general interest to the question. Indeed, the same synergy of
6 properties which is at the basis of the interest of composite materials multiplies possible weak
7 points, especially at the interfaces between components, where specific interactions and
8 sensitivity to solvents can bring to loss of cohesion detrimental for macroscopic properties of
9 the composite.

10
11 More complex occurrences can be observed when composite materials are exposed to mixed
12 solvents. A typical instance of the technological problems implied has been observed with the
13 introduction of oxygenates in transportation fuels, with the consequent exposition of the rich
14 variety of materials of vehicle engines to a mix of solvents with different polarities. Some
15 elastomers had to be banned from fuel-exposed components as, for instance, polyurethanes
16 underwent a three times larger swelling in methanol-gasoline blends than in the neat liquids^[1]
17 In the case of fiber-reinforced resin fuel tanks, a test fuel blend of octane, toluene and
18 methanol was found to produce much more swelling, surface crazing and cracking than the
19 component solvents alone.^[2] The microstructure of the composite materials can significantly
20 tune the solvent effects, as it was observed in the case of conducting carbon sensors
21 embedded in polymer matrices, when the trend of variation of the composite conductivity
22 with the composition of swelling solvent blends was inverted according to the level of the
23 electrical percolation point of the material.^[3]

24
25 Which effects can be expected when mixed solvents interact not with artificial composites but
26 with wood, the most widespread natural composite, whose cell walls structure is a complex
27 hierarchical assembly of biopolymers and biomolecules with different polarity? Swelling by
28 solvents is an important technological parameter for wood, either for its deconstruction and
29 further use of its different components in biobased chemistry, or for its use as structural
30 material (timber or wood-polymer composites), often implying the penetration of additives
31 and preservatives.

32
33 Non-linear evolution of the mechanical properties of wood with the composition of mixed
34 solvent solutions has early been noticed. Robertson observed that binary mixtures of liquids
35 can produce more swelling than each component alone and that this synergistic effect is
36 particularly pronounced when water is one of the liquids.^[4] Later, O'Leary and Hodges
37 mentioned that such systems behave "strangely" and reported that preferential adsorption of
38 certain components within binary systems can cause disproportionate tangential swelling.^[5]
39 In this study, we would like to address the specific case of water-ethanol solutions, which are

1
2
3 involved in a large range of industrial uses. Studying poplar and pinewood in green and dry
4 conditions, Chang *et al.* and Meier *et al.* reported that water-ethanol solutions generate
5 “hyperswelling”, namely a swelling more important than that observed for either pure water
6 or pure ethanol.^[6-8] More precisely, the most important deformations were reported for a 50%
7 volume fraction of ethanol. Similar results were observed on gels of acetylated lignin,
8 showing no swelling in ethanol but eight times the initial volume for swelling in water/ethanol
9 50%.^[9]

10 The sorption of solvents does not only alter the mass and volume of the material but also
11 strongly affects its mechanical properties. Alteration of both tensile and shear strength of
12 bamboo fibers composites has been evidenced upon aqueous treatments with ethanol.^[10-11]
13 Alterations in stress transfer mechanisms within cell walls and between wood cells in the
14 presence of solvents are probably related to selective extraction of some components of cell
15 walls and middle lamellae. Resulting modifications of the physico-mechanical and
16 hygroscopic properties of lignocellulosic fibers have been monitored by Meier, who reported
17 that ethanol-water mixtures generate local intercellular decohesion.^[12,13] Studies on other
18 lignocellulosic materials also identified specific effects of the cooperative sorption of water
19 and organics on middle lamellae and different layers of primary and secondary cell walls.^[14]
20 In the study of the behavior of (ligno-)cellulosic fibers in aqueous solutions of NMMO (*N*-
21 methylmorpholine *N*-oxide), Le Moigne *et al.* observed gradients in swelling and dissolution
22 mechanisms between the different cell walls, related to variations in their biochemical
23 composition.^[15,16] Acera Fernández *et al.* and Lefeuvre *et al.* showed that the removal of non-
24 cellulosic components from the cell walls of flax fibers with various aqueous and organic
25 solvents could strongly affect the transverse mechanical properties of the fibers and their
26 composites.^[17,18]

27 To unravel the underlying mechanism of non-linear effects of mixed solvents on wood
28 structure, we attempted a two-prongs approach including experimental monitoring of the
29 change of properties of poplar samples upon sorption of given amount of solutions or vapors
30 and a theoretical chemistry modelisation of the interaction of primary wood components with
31 mixed solvents. In the experimental approach, poplar was chosen as a model wood material
32 due to its low extractive content, thus preventing the issues related to solubility of low-
33 molecular weight components. Non treated veneers from poplar sapwood were used to
34 produce identical sample replicas. The progress of sorption of water-ethanol vapors as a
35 function of the relative pressure was monitored by isothermal gravimetric sorption and
36 provided information on the affinity of mixed vapors for the material. The effect of saturation
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

of wood by the solvents on viscoelastic properties and wood cell structure was studied by operando dynamic mechanical analysis and optical microscopy, respectively. Molecular dynamics (MD) simulations were carried out to model the interaction of water-ethanol mixed systems with the main wood components, cellulose and lignin.

2. Results and discussion

2.1. Sorption isotherms of water-ethanol vapours on poplar veneer

Sorption isotherms of vapor at the equilibrium with water, ethanol and an aqueous ethanol solution are represented in **Figure 1A**. The water vapor isotherm, with the sorption of nearly 16% water at p/p° 95, is typical of wood moisture sorption.^[14] The shape of the isotherm has often suggested its classification as a type IV IUPAC isotherm, typical of mesoporous solids. However, the hysteresis loop of condensation in mesopores should have a lower limit at the relative pressure at which capillary tension in the mesopore meniscus exceeds the tensile strength of the condensed liquid.^[19] The continuation of the hysteresis loop to relative pressure lower than 0.31, the limit value for water, indicates that sorption is not taking place by surface adsorption and pore condensation but by absorption in the material of the cell walls.^[20]

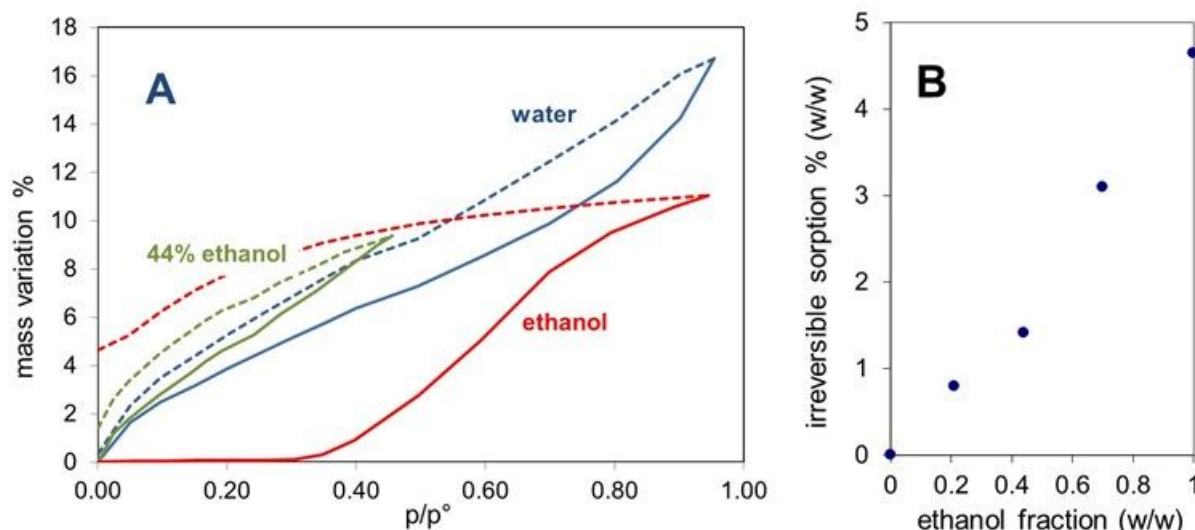


Figure 1. Sorption isotherms of water, ethanol and 44 wt% ethanol aqueous solution on poplar wood (A) and irreversible sorption as a function of the ethanol fraction in aqueous solution (B).

The sorption isotherm of ethanol vapour follows a very different pattern and features a sigmoidal shape with an important delay of uptake. Sorption begins only at p/p° higher than

0.3, indicating an activation energy barrier for the penetration of methanol in the material. The desorption curve presents a broad hysteresis and a significant amount of sorbate is retained also at $p/p^0=0$, when no more ethanol vapour can be measured. The sorption isotherm of the vapour at the equilibrium with a 44% w/w aqueous ethanol solution do not present any delay of uptake and features a shape similar to the water isotherm, albeit with a higher uptake at any given relative pressure. Moreover, some irreversible retention of sorbate is observed at the end of the desorption. The irreversibly-sorbed amount has been measured at the end of sorption loops for different concentrations of ethanol in solution (see Figure 1B) and is roughly proportional to the concentration of ethanol.

The different sorption behavior of water and ethanol can be globally interpreted by assuming an important role of the heterogeneity of the wood material. Water is clearly absorbed at very low partial pressure, showing a high affinity for the cell walls. The delayed sorption of ethanol can be interpreted as a more difficult penetration in the wood cell walls, which has to be related to their specific chemical composition and microstructure. Once the activity of ethanol is large enough, the molecule can go across the cell walls and favourable sorption in the bulk of the material is observed. In the case of water-ethanol mixtures, the sorption of water allows the sorption of ethanol at lower partial pressure, indicating that the water-sodden cell walls become more permeable to ethanol. The proportionality of the irreversible retention of sorbate with the concentration of ethanol strongly suggests chemisorption of ethanol in the bulk of the material.

2.2. Viscoelastic behavior of poplar veneer in water-ethanol solvents

The main properties measured by mechanical testing of viscoelastic materials are storage modulus E' , corresponding to the elastic energy stored in the material under test, and damping $\tan\delta$, corresponding to the ratio between the viscous energy dissipated as heat and the stored energy.^[21] The evolution of these parameters upon immersion in water, ethanol and 44 wt% ethanol aqueous solution are reported in **Figure 2**. Saturation of poplar veneer with any solution brings to decrease of storage modulus and increase of damping. However, the nature of the solvent highly affects the extent of these variations. Immersion in water brings to much larger decrease of storage modulus and increase of damping than immersion in ethanol. In the case of water, damping passes through a maximum at 5.5 times the initial value before decreasing to equilibrium values nearly 35% lower, whereas the lower values observed for ethanol are quite stable on a longer time. The decrease of damping after the transient maximum in water is accompanied by a slight recovery of E' .

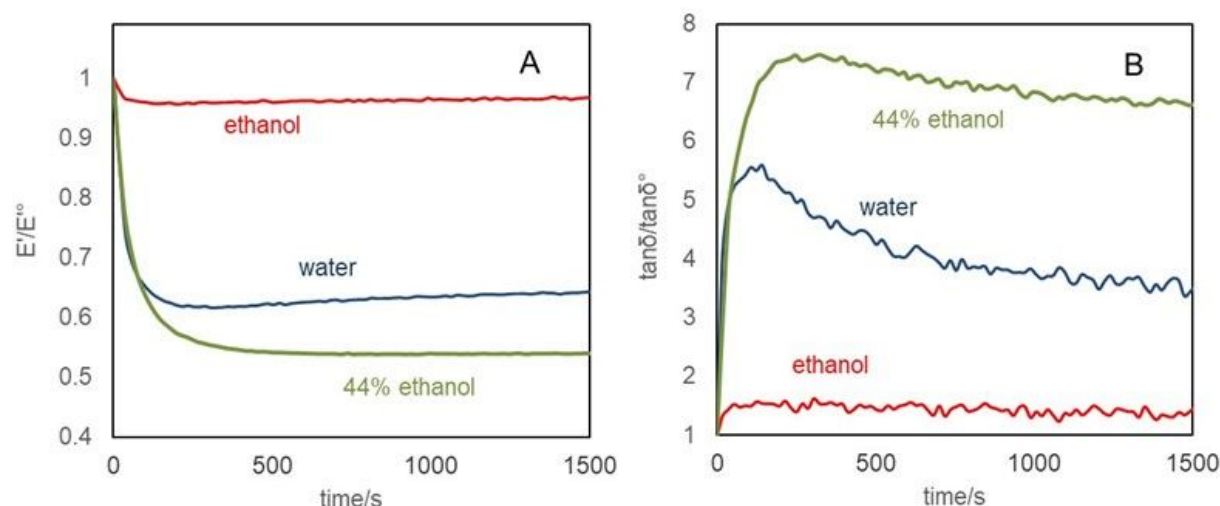


Figure 2. Viscoelastic properties of poplar veneers specimens immersed in water, ethanol and 44% ethanol aqueous solution. Evolution of relative storage modulus (A) and relative damping (B) after immersion at t^0 .

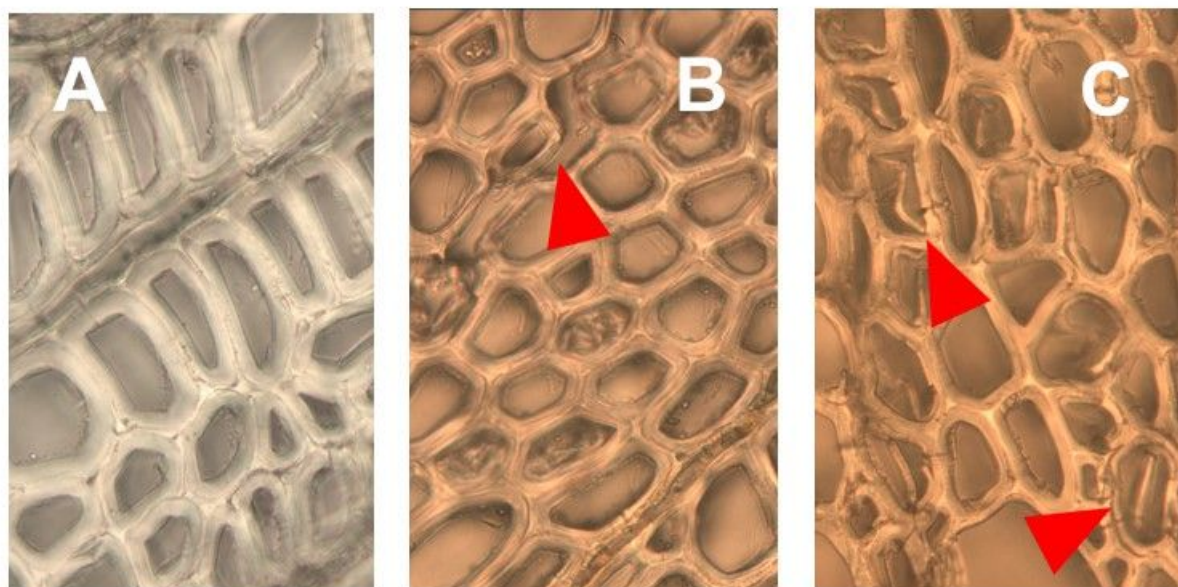
The strong variations in viscoelastic behavior observed for water saturation has been attributed to the diffusion of water molecules within the secondary cell wall, i.e. the main body of cell walls, at the origin of stiffness and longitudinal strength of wood.^[22] Navi et al. showed that the transient hydrogen bonding between crystalline cellulose and the surrounding amorphous polymers, owed to the introduction or removal of water, may accelerate shear slip between the two phases in the presence of an external load.^[23] Meier et al., working with different loading modes, already suggested that sorption of ethanol entails lower modifications of the mechanical properties than water.^[13] The large difference between mechanical properties after immersion in water or ethanol contrasts with the reasonably similar amounts of water and ethanol absorbed at saturation in the isothermal gravimetric sorption experiments. If ethanol molecules were absorbed within cell wall layers at the same absorption sites as water molecules, their larger steric hindrance would generate more swelling and higher modifications in viscoelastic behavior than water sorption, contrary to observed evidence. On these bases, it is assumed that the absorption sites of water and ethanol and the nature of their interactions with wood biopolymers are different and do not play the same role in the resulting viscoelastic behavior of poplar wood.

The variation of mechanical parameters upon immersion in water-ethanol solutions are not a linear combination of the effects observed for pure water and ethanol. Indeed, the observed decrease in storage modulus and increase in damping are higher than those observed with pure water (Figure 2). Some trends of mechanical properties upon saturation can be related to effects observed in sorption behavior (Figure 1). The observed order of relative E' decrease

1
2
3 and $\tan\delta$ increase, mixed solvent > water >> ethanol, corresponds to the ranking of initial
4 slopes of the sorption isotherms, itself related to the affinity of the solvents for the material.
5 The extent of decrease of damping after a transient peak follows instead the order water <
6 mixed solvent << ethanol and can be related to the extent of irreversible retention of sorbate at
7 the end of the desorption isotherm. The gradual decrease in relative $\tan\delta$ after reaching
8 extreme variations corresponds to a rearrangement of sorbate molecules and/or related
9 relaxation processes of biopolymers within the wood structure after initial swelling. The
10 penetration of ethanol molecules in the material seems to slow down the entire process of
11 desorption and re-equilibration after sorption. A possible mechanistic interpretation of these
12 results could be that irreversibly absorbed ethanol molecules modified the pristine
13 microstructure of wood cells and/or their intercellular cohesion.
14
15
16
17
18
19
20
21
22
23

2.3. Optical microscopy of water-ethanol sodden poplar veneer

24 Slices of 25 μm thickness were sampled from specimens used for dynamic mechanical
25 analysis in water, ethanol or mixed solvents, to be observed by optical microscopy (**Figure 3**).
26 Wood cells from the sample immersed in water (Figure 3A) are tightly packed against each
27 other and show thick cell walls.
28
29 Wood cells from the sample immersed in water (Figure 3A) are tightly packed against each
30 other and show thick cell walls.
31
32
33
34



35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
Figure 3. Cross-sections of poplar veneer immersed in water (A), 44 wt% ethanol aqueous
54 solution (B) and absolute ethanol (C). Width of each picture is 100 μm . Red arrows highlight
55 intercellular decohesion in B and interlayer disbonding in C.
56
57

58 The cell walls are compact and well bound to the middle lamella, the continuous lignin-rich
59 phase which connects cell walls through the tissues. On the contrary, in ethanol (Figure 3C),
60

1
2
3 cell walls swell to a lower extent, in agreement with previous reports.^[14,24] In addition, after
4 saturation with ethanol, primary wall (the cell wall layer nearer to the middle lamella) seems
5 less bounded to the middle lamella, where peeling can be observed, suggesting that
6
7 intercellular decohesion occurred in ethanol. It also appears that ethanol absorption can entail
8 disbonding within the cell walls, i.e. intracellular decohesion between the cell wall layers.
9
10 These phenomena can be the result of the weakening of interfaces between and within the cell
11 walls, and can be related to previous observations by Meier *et al.* on disbonding in pinewood
12 swollen in ethanol or by Clair *et al.* on the detachment of the G-layer in the case of poplar
13 tension wood.^[8,25]

14
15 Vapor sorption measurements have already shown that ethanol might be preferentially
16 absorbed by lignin rather than cellulose.^[26] The middle lamella, which is mainly composed of
17 lignin and other phenolic compounds, might then preferentially interact with ethanol, being
18 altered or even partially dissolved. This could explain the disbonding between primary wall
19 and middle lamella. Within the secondary wall (*viz.* the wall layers nearest to the cell lumen),
20 external layers contain greater amounts of lignin, and can thus possibly disbond from the
21 others.

22
23 In water–ethanol mixed solvent (Figure 3B), a combination of high swelling due to water and
24 intercellular decohesion due to ethanol can be observed. After immersion in 44 wt% ethanol
25 aqueous solution, both high swelling of the secondary wall and frequent local disbonding
26 between the middle lamella and the primary wall can indeed be noticed. Non-linearity effects
27 of aqueous solution of organics have already been observed in papermaking, where aqueous
28 solutions of ethanol and acetone led to higher lignin solubilization and extraction of
29 polyphenols than pure solvents.^[27,28]

30
31 The coupling of water sorption and intercellular disbonding by ethanol appears to be the best
32 combination to maximize microscopic swelling of wood cells. Indeed, in this situation, cells
33 are less mutually constrained and the cell walls swell more freely. If ethanol reacts
34 preferentially in lignin-rich regions as the middle lamella and the outer parts of the cell walls,
35 it is also likely that pre-swelling by water enables ethanol to get access to these sites. This
36 hypothesis is also supported by works of Hosseinpourpia *et al.* and Yang *et al.*, who observed
37 that the removal of lignin from pinewood samples allows higher water absorption and
38 stronger swelling of the polysaccharide matrix.^[29,30]

2.4. Elaboration of model structures of cellulose and lignin

39
40 The possibility of a detachment between lignin-rich middle lamella and cellulose-rich cell wall

1
2
3 upon impregnation of wood by mixed solvents supported the need for a modelisation of the
4 effects of water-ethanol solvents on the interactions between the surfaces of cellulose and
5 lignin. The first step was the determination of adequate models for these two main wood
6 components. A model of truncated cellulose crystal (cellulose nano-crystallite) was built with
7 seven cellulose chains, eight-monomers long (56 anhydroglucose units, **Figure 4A**), using the
8 crystallographic structure of cellulose I β reported by Nishiyama *et al.* and built with the
9 cellulose-builder tool of Gomes and Skaf.^[31,32] Oxygen terminal residues were capped with
10 hydrogen atoms and carbon terminal residues were capped with OH groups to obtain a finite
11 chain. The model reproduces correctly the intramolecular and intermolecular hydrogen bond
12 patterns expected for cellulose I β chains.^[33] Despite the limitation of our computational
13 analyses due to the small size of the model, some information can be drawn on the more
14 hydrophobic or hydrophilic behavior of different chains. The hydroxymethyl groups of
15 cellulose chains BDEG (Figure 4A) remain mainly buried inside the nanocrystal during the
16 simulations, therefore these chains display a more hydrophobic behavior than cellulose chains
17 AF, whose hydroxymethyl groups stand exposed to the surface.

18
19 A lignin dimer formed by two guaiacyl (G) monomer units (Figure 4B) was drawn including a
20 β -O-4 linkage, the most frequent linkage found in natural lignin. In order to study the lignin-
21 cellulose interactions, we modelled the assembly of the cellulose nanofiber model with 4
22 lignin dimers, one at each surface, as illustrated in Figure 4C. Each lignin dimer was placed
23 close to the cellulose surface, at a distance below 2 Å of the atomic centre of the closest atom,
24 in order to favour the initial cellulose-lignin interactions. The cellulose/lignin models were
25 centred in cubic boxes, leaving at least 1 nm to the longest side of the model to avoid
26 interactions between cellulose/lignin model and its images in the neighbouring boxes. This
27 rim spacing was used to determine the size of the box. Each cellulose/lignin system was
28 further solvated with the number of solvent molecules needed to fill the box size, as
29 summarized in the experimental section.
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

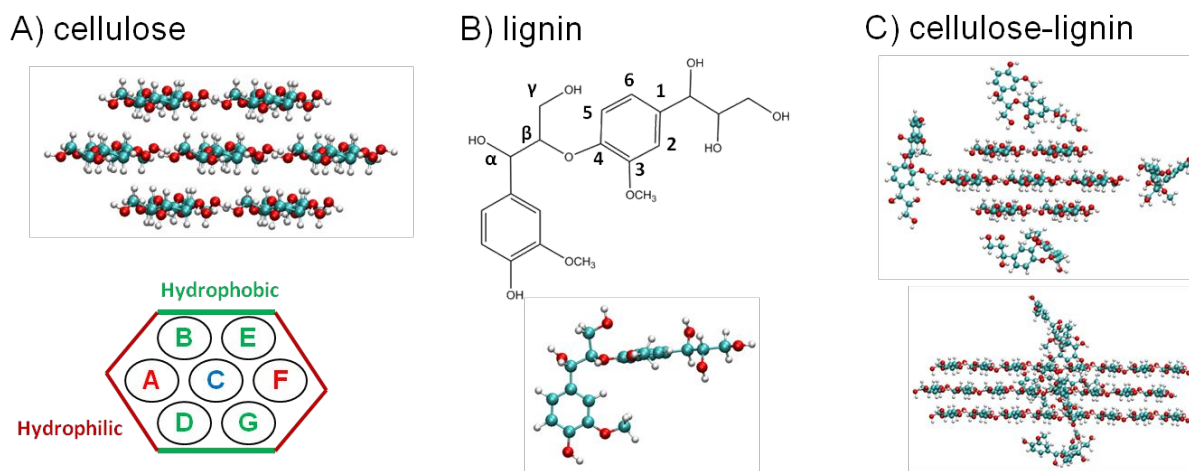


Figure 4. (A) Cellulose nanocrystal model with 7 chains (ABCDEFG). Chains BDEG represent the hydrophobic surface and chains AF represent the hydrophilic surface. Color code: oxygen (red), hydrogen (white), and carbon (cyan). (B) Lignin model: G-G dimer with β -O-4 linkage. Top: schematic representation and atom notation. Down: optimized structure in gas phase from quantum-chemistry DFT-PBE calculations. (C) Cellulose-lignin assembly. Top: front view; down: lateral view.

2.5. Modelling the interaction of water-ethanol solvents with a cellulose nanocrystal

The cellulose solvation is described by the radial distribution function (rdf), that describes the number of particles at distance r from a reference site. Due to the amphiphilic nature of cellulose chains in our model, we examined separately the hydrophobic and hydrophilic surfaces labeled as BDEG and AF chains, respectively, in Figure 4. The $O_{\text{cellulose}}-O_{\text{solvent}} g(r)$ s describe well-structured solvation layers at both hydrophilic and hydrophobic surfaces, in agreement with previous cellulose-water rdf studies.^[34] The integration of $g(r)$ from 0 to r , $cn(r)$, gives the cumulative number of particles within a distance r from the cellulose surface atoms. The computed $cn(r)$ of water and ethanol molecules within $r = 0.7$ nm from the glycosidic oxygen of cellulose, O_4 , as a function of ethanol concentration, are summarized in **Figure 5**. The number of water molecules surrounding the cellulose surfaces (Figure 5A) decreases with cosolvent concentration, as a result of the water displacement by ethanol.

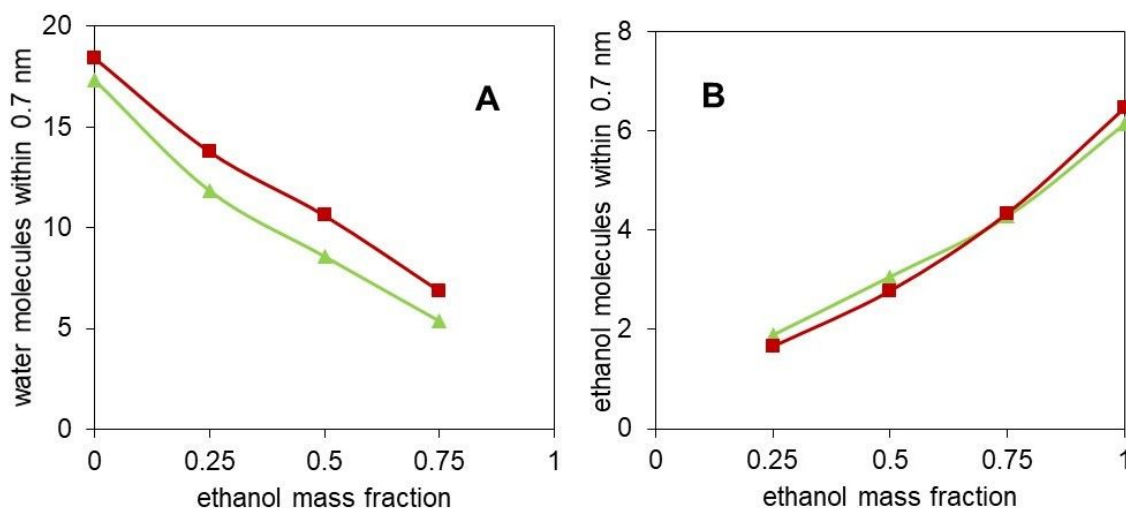


Figure 5. Cumulative number (cn) of molecules of water(A) and ethanol (B) within 0.7 nm from the cellulose glycosidic oxygen, as a function of ethanol concentration. Green triangles: hydrophobic surface (chains BEDG); red squares: hydrophilic surface (chains AF).

Moreover, water molecules are always more abundant around the hydrophilic surface than on the hydrophobic surface of cellulose. Correspondingly, in Figure 5B, the number of ethanol molecules surrounding the surface increases with the solvent concentration. At low ethanol fraction, there is a slight preference of ethanol for hydrophobic surface, due to competition with water. However, beyond 75 % ethanol, an inversion of the $cn(r)$ is observed, in which more ethanol molecules are found near the hydrophilic than the hydrophobic cellulose surface, suggesting an amphiphilic behaviour of ethanol. This effect is most likely due to the reconstruction of the H-bond network of the ethanol solvent, less disrupted by the remaining water molecules. We note that, keeping in mind that the volume of a H_2O molecule is about 1/3 of the volume of an ethanol molecule, at 100% ethanol concentration approximately three water molecules should be replaced by ethanol. Our analysis in Fig. 5, indeed, indicates that nearly 2.8 water molecules are replaced by one ethanol molecule.

The intermolecular cellulose-water and cellulose-ethanol hydrogen bonds (HBs) in each water-ethanol system are summarized in **Figure 6A** and 6B. The cellulose-water interactions play the major contribution to the intermolecular hydrogen bonding and more cellulose-water HBs are found on the hydrophilic surfaces than on the hydrophobic surfaces. The cellulose-water HBs in alcohol-water systems linearly decrease with the ethanol concentration, as a response to water displacement due to alcohol competition for H-bonding sites, as can be inferred from the cellulose-ethanol HBs in Figure 6B. The frequency of cellulose intramolecular H-bonds (Figure 6C) slightly increases from water to aqueous solutions of ethanol and strongly increases for pure ethanol. This suggests that, in the absence of strong

interactions with water, the surface cellulose chains are reoriented towards the core of the nanocrystal increasing the cohesion of the sub-surface layer.

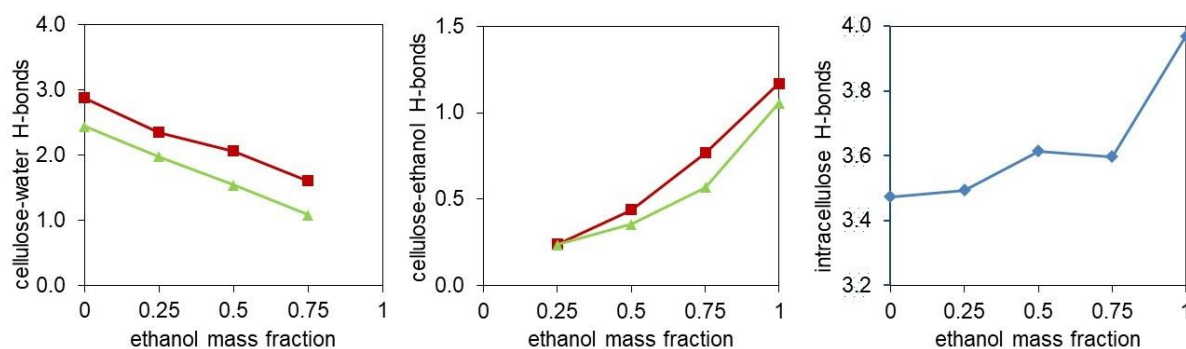


Figure 6. Average number of hydrogen bonds (HBs) per anhydroglucose monomer between cellulose and water (A) or cellulose and ethanol (B) at different ethanol concentrations. Green triangles: hydrophobic surface (chains BEDG); red squares: hydrophilic surface (chains AF). Average number of cellulose intra-molecular HBs per anhydroglucose monomer (C) at different ethanol concentrations.

2.6. Dynamics of lignin model dimer in water-ethanol solvents

MD simulations of a lignin dimer in pure water, ethanol and ethanol-water mixtures allow to calculate the distribution of the solvent accessible surface area (SASA). It can be observed in **Figure 7A** that, in water and the most ethanol-poor mixed solution, the lignin dimers display the lowest average SASA, as a hydrophobic response of the lignin dimer towards water, a bad solvent for lignin.

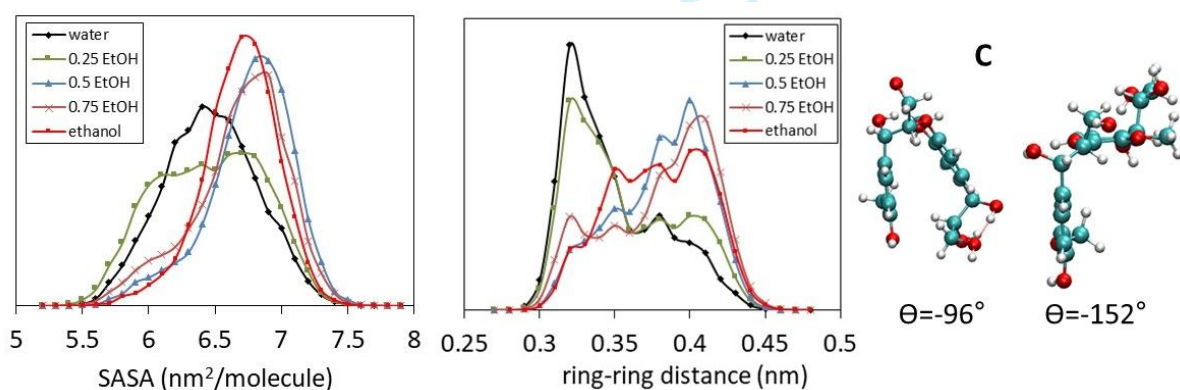


Figure 7. Distributions of solvent accessible surface area, SASA, of lignin model at different ethanol concentrations (A). Center-of-mass ring-ring distance distribution of guaiacyl monomers at different ethanol concentrations (B). Model lignin conformers (C) with different solvent accessibility and ring-ring coupling depending on the dihedral torsion angle $\theta = C'_4-O'_4-C_8-C_7$ (C). Conformers with $\theta = -96^\circ$ are found mostly in water and diluted ethanol mixture, whereas conformers with $\theta = -152^\circ$ are predominant in 50-75 wt% mixtures and pure ethanol.

In pure ethanol, the surface increases as the interactions between the solvent and the lignin improve. However, it is in the range of the mixed solvent composition where the largest

1
2
3 accessible surface is found. This effect corresponds to a change of conformation of the lignin
4 dimer, which assumes a more extended conformation, corresponding to a higher interaction
5 with the solvent. The distribution of the distance between the two aromatic rings of the
6 guaiacyl dimer, shown in Figure 7B, indicates the presence of a stacked conformation with
7 ring-ring distance at 0.32 nm, predominant in water and 0.25% ethanol. A less stacked
8 conformer with a ring-ring distance close to 0.40 nm is dominant in ethanol and ethanol-rich
9 mixtures. The torsion angle of the dimer θ ($\theta=C4'-O4-C\beta-C\alpha$) shows a dihedral conformation
10 distribution near $\theta=-96^\circ$ in pure water and diluted ethanol (Figure 7C). For the pure ethanol
11 and concentrated mixtures, the dihedral conformation is mostly found near $\theta=-152^\circ$.
12 Furthermore, the angle distribution near -152° is narrower in 75% ethanol solution than in
13 pure ethanol. This effect is in qualitative agreement with the results of Smith *et al.*, who
14 studied the conformation of lignin polymers of up to 60 units in water-tetrahydrofuran (THF)
15 mixtures.^[35] In their study, they found that mixed solvents facilitate the solvation of lignin
16 molecules, which, instead, in pure water adopt a crumbled globular-like shape, with a
17 considerably reduced SASA. Additionally, very recent studies of lignin solubility in water-
18 ethanol mixtures have identified a solubilisation maximum at 60 wt% ethanol.^[36]

2.7. Modelling the cellulose-lignin interactions in water-ethanol solvents

32
33 Molecular dynamics of a model of cellulose and four lignin dimers in the same simulation box
34 provided some information on the effect of solvent on the physical adhesion between
35 cellulose and lignin. The cellulose-lignin intermolecular energies reported in **Figure 8** are
36 weaker in mixed solvents.

37
38 This simulation suggests that a detachment between cellulose and lignin, an important
39 phenomenon in the separation of cellulose and dissolved lignin in pulping processes, is
40 favoured in the presence of mixed solvents. It is interesting to compare the relative intensity
41 of Van der Waals and electrostatic interactions. Van der Waals interactions modelled with the
42 Lennard-Jones potential are stronger than Coulomb energies (Figure 8). The similar behaviour
43 of the Lennard-Jones and Coulomb energy curves in the mixed solvent region indicates that
44 these two types of interactions are highly correlated.

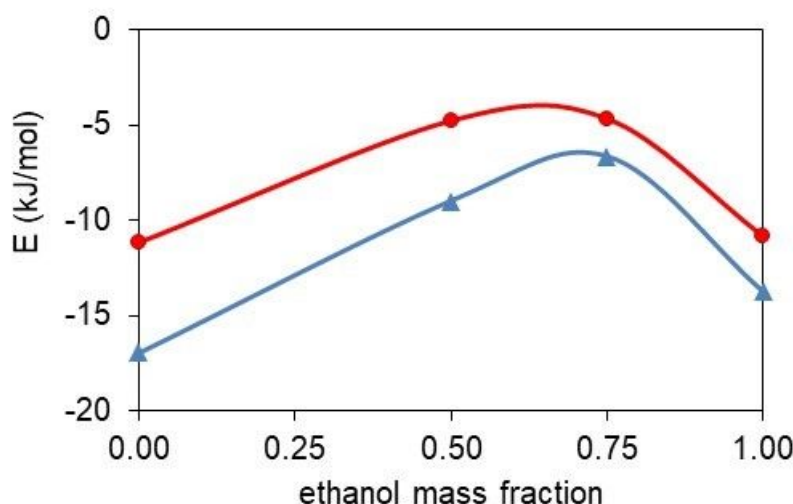


Figure 8. Lennard-Jones (blue triangles) and Coulomb (red circles) energies of lignin-cellulose molecular interaction normalized on number of lignin dimers in the system vs. composition of water-ethanol solutions.

Despite Lennard-Jones energy is predominant, the electrostatic interactions also bring an important contribution, which can be better evaluated by the evolution of hydrogen bonds between cellulose and lignin with the composition of the solvent. The H-bond probability distribution describes the probability of finding lignin bound to cellulose ($H\text{-bond} \geq 1$) or unbound ($H\text{-bond} = 0$). Moreover, a high probability at $H\text{-bond} = 0$ is a good descriptor of the capacity of the solvent of disturbing the lignin-cellulose electrostatic interactions, favouring the unbound system. Given the hydrophobic nature of lignin, we show the cellulose-lignin H-bond probability distribution with the cellulose hydrophobic surface (**Figure 9**).

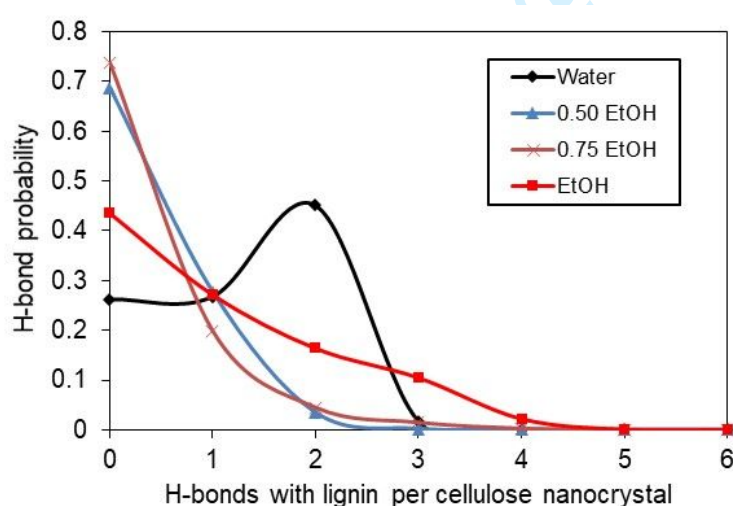


Figure 9. Probability distribution of hydrogen bonds between lignin dimers and hydrophobic cellulose chains (BDEG) in pure and aqueous ethanol mixtures. Lines are added to help the reader's eye despite the H-bond intervals being discrete.

1
2
3 In water alone, the probability distribution falls around H-bond=2. Thus, our results show that
4 in water (a bad solvent for lignin) the adhesion of lignin to cellulose is favoured. In pure
5 ethanol, the higher probability falls at H-bond=0, favouring the unbound state of lignin, but
6 also an important part of the distribution falls at the bounded state. Nevertheless, the H-bond
7 distribution confirms that water-ethanol mixture enhances lignin dissolution, as the unbound
8 state is favoured with the highest probability for the mixed water-ethanol solutions.
9
10
11
12
13

15 3. Conclusions

17 The study of the interaction of poplar wood and its components with water-ethanol solutions
18 and their vapors confirms that wood is an effective model for the study of the peculiar
19 behavior of composite materials exposed to mixed solvents. Two kinds of non-linearity play a
20 role in the behavior of composite materials exposed to mixed solvents: (i) the colligative
21 properties of ethanol aqueous solutions are themselves not linear with their composition;^[37]
22 (ii) moreover, the differential interactions of each component of the solvent with specific
23 components of the composite material creates further non-linearity of the sorption and
24 mechanical properties of the material.
25
26
27
28
29

31 Several instances of “hyperswelling” of wood in the ethanol-water system can be observed,
32 namely an effect of mixed solution exceeding the linear combination of the effects observed
33 with individual solvent components. The affinity of the material for the solvent, proportional
34 to the initial slope of the sorption isotherms, increases in the order ethanol << water <
35 ethanol-water solution, confirming a synergistic effect of water and ethanol. The amount of
36 solvent absorbed is an important property, strictly related to the swelling of the material. Also,
37 considering the viscoelastic properties, storage modulus decreases and damping increases in
38 the order ethanol << water < ethanol-water solution, attesting for higher softening in the case
39 of ethanol-water mixtures. Optical microscopy observations after saturation of wood by the
40 mixed solvents indicate that the differential affinity of the components of the solvent within
41 wood microstructure plays a significant role in the phenomena. Partial dissolution of specific
42 zones, possibly lignin-rich regions, within wood microstructure can favour the release of
43 constraints within and between the cell walls, and can allow more free swelling of hydrophilic
44 polysaccharide-rich zones.
45
46
47
48
49
50
51
52
53

55 Molecular dynamics have proved effective in simulating the non-linear effects of solvent
56 solutions on wood components. Once established the relative affinities of water and ethanol
57 for cellulose (more hydrophilic) and lignin (less hydrophilic), non-linear effects have been
58 observed in the solvation of lignin, which reaches a maximum between 50 % ethanol aqueous
59
60

1
2
3 solution and pure ethanol. The simulation of a cellulose-lignin complex has indicated that the
4 magnitude of interactions between the two wood components follows the order water >
5 ethanol > ethanol-water solution. The decrease of lignin-cellulose interaction in mixed
6 solvents can play a significant role in the purification of cellulose from dissolved lignin in
7 organosolv pulping processes.
8
9
10
11
12

13 **4. Experimental section**

14 **4.1 Materials and experimental methods**

15
16 Never-dried and untreated poplar veneers (*Populus tremula*) were cut with a hydraulic press
17 in the sapwood zone along fibers direction (R 5 mm, T = 0.5 mm, L= 45 mm) and stabilized
18 for 2 months at 25 °C; 45% RH.
19
20
21

22 Vapor sorption measurements were performed at 40 °C by dynamic vapor sorption (DVS) in
23 an isothermal gravimetric sorption instrument (DVS-Vacuum, Surface Measurement
24 Systems). Each wood specimen was loaded and dried in-situ at 40 °C under high vacuum ($2 \times$
25 10^{-6} Torr) for 10 h before admission of vapor in successive steps of increasing and decreasing
26 relative pressure (p/p°).
27
28
29

30
31 The effects of solvent sorption on the dynamic longitudinal mechanical behavior of wood
32 specimens were analyzed by coupling a Dynamic Mechanical Analyzer (DMA) (50N
33 Metravib) with a tank connected to a thermostatic bath regulated at 40 °C. Samples were
34 mounted in tensile mode before being immersed in the solvent bath. The variations of the
35 complex modulus E^* , which is composed of the storage modulus E' (real part) and loss
36 modulus E'' (imaginary part), and the damping ($\tan\delta = E''/E'$) were measured at a frequency of
37 0.05 Hz all along the absorption process. After mechanical testing, each sample was cut into
38 thin radial slices of 25 μm thickness with a WSL GSL1 sledge microtome. The slices, stored
39 in the same solution, were placed between two glass plates and the cell walls were observed
40 with a Leica DM LM/P optical microscope in transmission mode.
41
42
43
44
45
46
47
48
49

50 **4.2 Molecular dynamics simulation methods**

51
52 All-atom MD simulations of each system described in Table 1 were carried out using the
53 GROMACS package 2016.3,^[38-42] along with the 4-sites Transferable Intermolecular Potential
54 (TIP4) for liquid water,^[43-45] the CHARMM36 additive force field,^[46,47] and the CHARMM-
55 compatible force field for lignin.^[48] Solvent structure for ethanol was available at the
56 GROMACS molecule and liquid database.^[49] Each cellulose/lignin model was centered in a
57
58
59
60

cubic box and each macromolecule-solvent system was solvated with the number of solvent molecules dictated by the volume of the box, as summarized in **Table 1**.

Table 1. Configuration of simulated systems and equilibrium size of simulation boxes for cellulose nanocrystal (7 chains, 8 glucose monomers each) and lignin models studied in this work

Model wood component	Solvent system	Ethanol fraction (wt%)	Number of ethanol molecules	Number of water molecules	Cubic box side length (nm)	Volume (nm ³)
Cellulose nanocrystal (7 chains, 56 glucose monomers)	pure water	0	0	9764	6.739	306.03
	ethanol-water	25	1373	10559	7.698	456.32
		50	2607	6674	7.699	456.34
		75	3693	3151	7.699	456.40
pure ethanol	100	4321	0	7.547	429.92	
Lignin G-G β -O-4 dimer	pure water	0	0	1448	3.54	44.55
	ethanol-water	25	132	1012	3.515	43.45
		50	250	641	3.514	43.41
		75	355	302	3.51	43.55
pure ethanol	100	399	0	3.4	39.48	
Cellulose nanocrystal + 4 lignin dimers	pure water	0	0	10370	6.886	326.56
	ethanol-water	50	1725	4415	6.749	307.46
		75	2443	2084	6.749	307.48
pure ethanol	100	2985	0	6.711	302.31	

For each simulation box, energy minimization was performed using the steepest descent algorithm until convergence to a tolerance of $100 \text{ kJ mol}^{-1} \text{ nm}^{-1}$. After minimization, restrained simulations were performed for 200 ps at 298.15 K. Afterwards, 20-ns MD simulations were performed with a frame-saving rate (for analysis) of 1 ps, in order to study the interaction of the cellulose/lignin models with the solvent molecules. Temperature and pressure coupling were handled using the leap-frog stochastic dynamics integrator and the Parrinello-Rahman method, respectively. Initial velocities were generated from a Maxwell distribution at 298.15 K and the isothermal-isobaric (NPT) ensemble was considered for data collection. Neighbor searching and short-range non-bonded interactions were handled with the Verlet cut-off scheme. Electrostatics were treated with the Fast smooth Particle-Mesh Ewald (SPME) method, with a Coulomb cut-off of 1.2 nm, a fourth order interpolation and Fourier spacing of 0.12 nm. Van der Waals (vdW) interactions were treated using the Lennard-Jones potential with a cut-off distance of 1.2 nm. Simulations were carried out in an Intel Xeon CPU with 2.10Gz with 32 logical cores

The structure and dynamics of the macromolecules in the water-ethanol mixtures were characterized by site-to-site radial distribution functions $g(r)$, cumulative numbers $cn(r)$,

average number of hydrogen bonds (HBs), solvent accessible surface area (SASA), and intermolecular energies using the incorporated tools within GROMACS. Each of these descriptors was calculated for all the water-ethanol concentrations considered in this study. Site-to-site $g(r)$ s were computed for the water-macromolecule and ethanol-macromolecule pairs using the oxygens of water, ethanol, cellulose and lignin. Cumulative numbers were obtained from the integration of the $g(r)$ up to a 2 nm correlation distance. HBs were calculated using a geometrical criterion with a maximum donor-acceptor distance of 0.35 nm and a hydrogen-donor-acceptor angle of 30° , and they were accordingly further normalized, as described in the text. Solvent accessible surfaces were computed using the double cubic lattice method (DCLM) by using a solvent probe radius of 0.14 nm. Further details of the MD simulations and the description of the tools used for analysis can be found in our previous work on water-ethanol mixture interactions.^[37]

Acknowledgments

This work was partially funded by the Agence Nationale de la Recherche program Investissements d'Avenir through the contract ANR-10-LABX-05-01 (LabExCheMISyst) and by an ENSCM PhD grant in the framework of the SINCHEM Joint Doctorate program under the Erasmus Mundus Action 1 Programme (FPA 2013-0037). Access to the HPC resources of CCRT/CINES/IDRIS was granted under the allocation x2017087369 by GENCI. Poplar veneers were graciously provided by Maurice Orly, Danzer Company. The authors acknowledge the support of Surface Management Systems London and their team for DVS analyses, the technical support of Marc Longerey, Centre des Matériaux des Mines d'Alès (C2MA), for dynamic mechanical measurements and the assistance of Olivier Arnoult, Laboratoire de Mécanique et Génie Civil (LMGC) of the Université de Montpellier, in microscopy techniques.

Received: ((will be filled in by the editorial staff))

Revised: ((will be filled in by the editorial staff))

Published online: ((will be filled in by the editorial staff))

References

- [1] R. L. Bechtold, M. B. Goodman, T. A. Timbario. *Use of Methanol as a Transportation Fuel*, The Methanol Institute, Arlington, VA, USA **2007**.
- [2] L. S. Norwood, Chemical resistance of composites, <https://netcomposites.com/guide-tools/guide/coatings/chemical-resistance/>, accessed December 2018.
- [3] S. Frackowiak, M. Maciejewska, A. Szczurek, M. Kozłowski. *e-Polymers* **2011**, 032.
- [4] A. A. Robertson. *Pulp Paper Mag Canada* **1964**, 65, 171.
- [5] P. O'Leary, P. A. Hodges. *Wood Sci Technol* **2001**, 35, 217.
- [6] S. Chang, B. Clair, J. Gril, H. Yamamoto, F. Quignard. *Wood Sci Technol* **2009**, 43, 703.

- [7] S. Chang, F. Quignard, F. Di Renzo, B. Clair. *BioResources* **2012**, 7, 2418.
- [8] P. Meier, T. Kaps, U. Kallavus. *Mater Sci Medzg* **2005**, 11, 140.
- [9] M. Nishida, Y. Uraki, Y. Sano. *Bioresour Technol* **2003**, 88, 81.
- [10] Y. A. Hu, M. He, R. X. Zhu, Y. H. Zhang, Y. L. Yu, W. J. Yu. *J Trop For Sci* **2016**, 28, 112
- [11] L. Qin, W. J. Yu. *Adv Mater Res* **2009**, 79, 1395.
- [12] P. Meier, U. Kallavus, A. Rohumaa, T. Kaps. *Mater Sci Medzg* **2006**, 12, 25.
- [13] P. Meier, E. Stöör, T. Kaps, U. Kallavus. *Est J Eng* **2006**, 12, 125.
- [14] J. Bossu, N. Le Moigne, S. Corn, P. Trens, F. Di Renzo. *Wood Sci Technol* **2018**, 52, 987.
- [15] N. Le Moigne, E. Montes, C. Pannetier, H. Höfte, P. Navard. *Macromol Symp* **2008**, 262, 65.
- [16] N. Le Moigne, J. Bikard, P. Navard. *Cellulose* **2010**, 17, 507.
- [17] J. Acera Fernández, N. Le Moigne, A. S. Caro-Bretelle, R. El Hage, A. Le Duc, M. Lozachmeur, P. Bono, A. Bergeret. *Ind Crops Prod* **2016**, 85, 93.
- [18] A. Lefeuvre, A. Le Duigou, A. Bourmaud, A. Kervoelen, C. Morvan, C. Baley, *Ind. Crops Prod* **2015**, 76, 1039.
- [19] P. Trens, N. Tanchoux, A. Galarneau, D. Brunel, B. Fubini, E. Garrone, F. Fajula, F. Di Renzo. *Langmuir* **2005**, 21, 8560.
- [20] K.S.W. Sing, J.D. Madeley. *J Appl Chem* **1954**, 4, 365.
- [21] M.A. Meyers, K.K. Chawla, *Mechanical Behavior of Materials*, 2nd edition, Cambridge University Press, Cambridge, UK **2009**.
- [22] A. Bergander, L. Salmén. *J Mater Sci* **2002**, 37, 151.
- [23] P. Navi, V. Pittet, C. J. G. Plummer. *Wood Sci Technol* **2002**, 36, 447.
- [24] G. I. Mantanis, R. A. Young, M. Rowell. *Holzforschung* **1994**, 48, 480.
- [25] B. Clair, B. Thibaut, J. Sugiyama. *J Wood Sci* **2005**, 51, 218.
- [26] S. Lequin, D. Chassagne, T. Karbowiak, J.-P. Bellat. *J Agric Food Chem* **2013**, 61, 5391.
- [27] Y. Yang, F. Zhang. *Ultrason Sonochem* **2008**, 15, 308.
- [28] J. Domínguez-Robles, T. Tamminen, T. Liitiä, M. S. Peresin, A. Rodríguez, A.-S. Jääskeläinen. *Int J Biol Macromol* **2018**, 106, 979.
- [29] R. Hosseinpourpia, S. Adamopoulos, C. Mai. *Wood Sci Technol* **2016**, 50, 165.
- [30] T. Yang, H. Zhou, E. Ma, J. Wang. *Results in Physics* **2018**, 10, 61.
- [31] Y. Nishiyama, P. Langan, H. Chanzy. *J Am Chem Soc* **2002**, 124, 9074.
- [32] T. C. F. Gomes, M. S. Skaf. *J Comput Chem* **2012**, 33, 1338.
- [33] R.C. Pettersen, in *The Chemistry of Solid Wood* (Ed: R. Rowell), American Chemical Society, Washington, DC, USA **1984**, Ch. 2.
- [34] A.P. Heiner, O. Teleman. *Langmuir*, **1997**. 13, 511.
- [35] M. D. Smith, B. Mostofian, X. Cheng, L. Petridis, C. M. Cai, C. E. Wyman, J. C. Smith. *Green Chem.* **2016**. 18, 1268
- [36] W. M. Goldmann, J. Ahola, M. Mikola, J. Tanskanen, *Sep Purif Technol* **2019**, 209, 826.
- [37] S. M. Aguilera-Segura, F. Di Renzo, T. Mineva, *J Mol Model* **2018**, 24, 292.
- [38] M. J. Abraham, T. Murtola, R. Schulz, S. Pall, J. C. Smith, B. Hess, E. Lindahl. *SoftwareX*, **2015**. 1–2, 19.
- [39] H. Bekker, H. J. C. Berendsen, E. J. Dijkstra, S. Achterop, R. Vondrumen, D. Vanderspoel, A. Sijbers, H. Keegstra, M. K. R. Renardus, in *Physics Computing '92* (Eds: R. A. De Groot, J. Nadrchal) World Scientific Publishing, Singapore **1993**, p. 252.
- [40] H. J. C. Berendsen, D. van der Spoel, R. van Drunen. *Comput Phys Commun* **1995**. 91, 43.
- [41] D. Van Der Spoel, E. Lindahl, B. Hess, G. Grienhof, A. E. Mark, H. J. Berendsen. *J Comput Chem* **2005**. 26, 1701.
- [42] S. Páll, M. J. Abraham, C. Kutzner, B. Hess, E. Lindahl, in *Solving Software Challenges for Exascale* (Eds: S. Markidis, E. Laure) Springer International Publishing, Cham, Switzerland **2015**, p. 3.
- [43] W. L. Jorgensen, J. Chandrasekhar, J. D. Madura, R. W. Impey, M. L. Klein. *J Chem Phys* **1983**. 79, 926.
- [44] T. J. Dick, J. D. Madura, *Annu Rep Comput Chem* **2005**, 1, 59.
- [45] P. Mark, L. Nilsson. *J Phys Chem A*, **2001**, 10, 9954.
- [46] O. Guvench, S. N. Greene, G. Kamath, J. W. Brady, R. M. Venable, R. W. Pastor, A. D. MacKerell. *J Comput Chem* **2008**. 29, 2543.
- [47] O. Guvench, E. Hatcher, R. M. Venable, R. W. Pastor, A. D. MacKerell. *J Chem Theory Comput* **2009**, 5, 2353.
- [48] L. Petridis, J. C. Smith, *J Comput Chem* **2009**, 30, 457.
- [49] D. van der Spoel, P. J. van Maaren, C. Caleman, *Bioinformatics* **2012**, 28, 752.

The table of contents entry should be 50–60 words long and should be written in the present tense and impersonal style (i.e., avoid we). The text should be different from the abstract text.

Synergistic effects of mixed solvents on properties of composite materials are observed through the study of sorption of water-ethanol solutions in lignocellulose, a natural model composite. Non-linear effects of solution composition on absorption, viscoelastic properties and cohesion of poplar wood are investigated and modelled by molecular dynamics.

Keyword: Lignocellulose swelling by mixed solvents

S. M. Aguilera-Segura, J. Bossu, S. Corn, P. Trens, T. Mineva, N. Le Moigne*, F. Di Renzo*

Synergistic sorption of mixed solvents in wood cell walls: experimental and theoretical approach

ToC figure ((Please choose one size: 55 mm broad × 50 mm high or 110 mm broad × 20 mm high. Please do not use any other dimensions))

