



**HAL**  
open science

## Experimental redox transformations of uranium phosphate minerals and mononuclear species in a contaminated wetland

Lucie Stetten, Pierre Lefebvre, Pierre Le Pape, Arnaud Mangeret, Pascale Blanchart, Pauline Merrot, Jessica Brest, Anthony Julien, John R Bargar, Charlotte Cazala, et al.

### ► To cite this version:

Lucie Stetten, Pierre Lefebvre, Pierre Le Pape, Arnaud Mangeret, Pascale Blanchart, et al. Experimental redox transformations of uranium phosphate minerals and mononuclear species in a contaminated wetland. *Journal of Hazardous Materials*, 2020, 384, pp.121362. 10.1016/j.jhazmat.2019.121362 . hal-02343440

**HAL Id: hal-02343440**

**<https://hal.science/hal-02343440>**

Submitted on 27 Feb 2020

**HAL** is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.



Distributed under a Creative Commons Attribution - NonCommercial - NoDerivatives 4.0 International License

1           **Experimental redox transformations of uranium phosphate**  
2           **minerals and mononuclear species in a contaminated wetland**

3  
4 Lucie Stetten<sup>1,2,3\*</sup>, Pierre Lefebvre<sup>1</sup>, Pierre Le Pape<sup>1</sup>, Arnaud Mangeret<sup>2</sup>, Pascale  
5 Blanchart<sup>2</sup>, Pauline Merrot<sup>1</sup>, Jessica Brest<sup>1</sup>, Anthony Julien<sup>2</sup>, John R. Bargar<sup>4</sup>,  
6 Charlotte Cazala<sup>2</sup>, Guillaume Morin<sup>1</sup>  
7

8  
9 <sup>1</sup> Institut de Minéralogie, de Physique des Matériaux et de Cosmochimie (IMPMC), UMR  
10 7590 CNRS-Sorbonne Université-IRD-MNHN, case 115, 4 place Jussieu, 75252 Paris Cedex  
11 5, France  
12

13 <sup>2</sup> Institut de Radioprotection et de Sûreté Nucléaire, IRSN, 31 avenue de la Division Leclerc,  
14 92262 Fontenay-aux-Roses, France  
15

16 <sup>3</sup> University of Vienna, Centre for Microbiology and Environmental Systems Science,  
17 Department of Environmental Geosciences, Althanstraße 14, UZA II, 1090 Vienna, Austria  
18

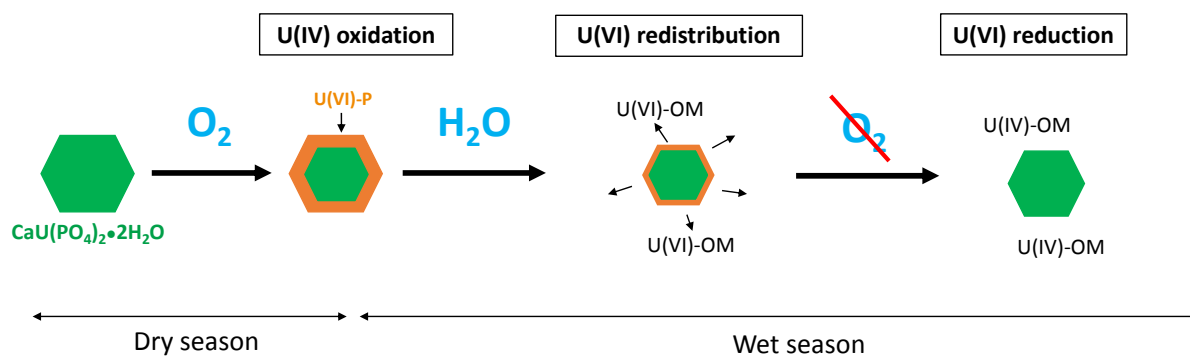
19 <sup>4</sup> Stanford Synchrotron Radiation Lightsource (SSRL), SLAC National Accelerator National  
20 Laboratory, MS 69, 2575 Sand Hill Road, Menlo Park, CA 94025, USA  
21

22 \*luciestetten@gmail.com  
23

24 Revised version to be submitted to *Journal of hazardous materials*  
25  
26  
27  
28  
29

30  
31  
32  
33  
34  
35  
36  
37  
38  
39  
40  
41  
42  
43  
44  
45  
46  
47  
48  
49  
50

## GRAPHICAL ABSTRACT



51  
52  
53  
54  
55  
56  
57  
58  
59  
60  
61  
62  
63  
64  
65  
66  
67  
68  
69  
70  
71  
72  
73  
74

**ABSTRACT**

Reducing conditions and high organic carbon content make wetlands favorable to uranium (U) sequestration. However, such environments are subjected to water-table fluctuations that could impact the redox behavior of U and its mobility. Our previous study on U speciation in a highly contaminated wetland has suggested a major role of water-table redox fluctuations in the redistribution of U from U(IV)-phosphate minerals to organic U(VI) and U(IV) mononuclear species. Here, we investigate the mechanisms of these putative processes by mimicking drying or flooding periods *via* laboratory incubations of wetland samples. LCF-XANES and EXAFS analyses show the total oxidation/reduction of U(IV)/U(VI)-mononuclear species after 20 days of oxic/anoxic incubation, whereas U-phosphate minerals appear to be partly oxidized/reduced. SEM-EDXS combined with  $\mu$ -XRF and  $\mu$ -XANES analyses suggest that autunite  $\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$  is reduced into lermontovite  $\text{U}(\text{PO}_4)(\text{OH}) \cdot \text{H}_2\text{O}$ , whereas oxidized ningyoite  $\text{CaU}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$  is locally dissolved. The release of U from this latter process is observed to be limited by U(VI) adsorption to the surrounding soil matrix and further re-reduction into mononuclear U(IV) upon anoxic cycling. Analysis of incubation waters show, however, that dissolved organic carbon enhances U solubilization even under anoxic conditions. This study brings important information that help to assess the long-term stability of U in seasonally saturated organic-rich contaminated environments.

75 **Keywords:** wetland, uranium, redox transformations, soil incubation, X-ray absorption  
76 spectroscopy

## 77 **1. INTRODUCTION**

78 Uranium is a toxic element [1] that naturally occurs in the continental crust at an  
79 average concentration of  $2.7 \text{ mg.kg}^{-1}$  [2]. It locally concentrates in economic deposits *via*  
80 reductive precipitation and/or hydrothermal processes [3] and has been mined for decades in  
81 several countries worldwide. Uranium mining operations and legacies have been shown to be  
82 responsible for local increases of U contaminations in surrounding surface environments. For  
83 instance, U concentrations exceeding the geochemical background were found in aquifer  
84 sediments [4], lacustrine sediments [5], soils [6] and wetlands [7-10] in the vicinity of former  
85 U mining sites.

86 Organic-rich wetlands are favorable to develop reducing conditions and have been  
87 recognized to accumulate uranium both *via* organic complexation of U(VI) and U(VI)  
88 reduction to less soluble U(IV) species. These environments are thus considered as being able  
89 to naturally attenuate U migration from nearby mining sites [7,10]. However, the long-term  
90 fate of U in wetlands remains unclear, especially when they exhibit intermittent oxidizing  
91 conditions due to water-table fluctuations [10]. Indeed, non-crystalline U species that have  
92 been identified as major U species in wetlands, such as mononuclear U(VI) and U(IV)  
93 complexes bound to organic matter [10,11-13], are known to be potentially mobile under both  
94 oxic [14-17] and anoxic conditions [7,18]. In addition, U(IV)-phosphate minerals may also  
95 play an important role in U retention in mining-contaminated wetlands. Indeed, such phases  
96 have been recently recognized as important U hosts in contaminated environments [10,19]  
97 and as possible products of microbial U(VI) reduction [16,20-21]. However, the geochemical  
98 reactivity of these phases when submitted to redox fluctuations is still scarcely documented  
99 and remains debated. For instance, Latta et al. (2016) [22] reported that the oxidation of a

100 synthetic amorphous U(IV)-phosphate resulted in the formation of a low-solubility U(VI)-  
101 phosphate phase. In contrast, oxidation of a ningyoite-like mineral (ideal formula  
102  $\text{CaU}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$ ) has been shown to lead to significant U remobilization [16]. To this regard,  
103 the reactivity of U(IV)-phosphate minerals such as ningyoite needs to be further investigated.  
104 In this context, our recent study of U speciation in a highly contaminated wetland in Brittany,  
105 France [10] has suggested that low water-table level conditions can lead to oxidative  
106 dissolution of U(IV)-phosphate minerals, especially ningyoite, and to subsequent sorption of  
107 U(VI) to soil organic matter, while flooding conditions could favor further reduction of U(VI)  
108 to U(IV) organic complexes. Based on these findings, the present study aims at elucidating  
109 the micro- and molecular-scale mechanisms underlying such U redistribution processes, by  
110 performing incubations experiments that mimics typical redox changes induced by water-  
111 table fluctuations in a seasonally saturated wetland.

112

## 113 **2. MATERIALS & METHODS**

114

### 115 **2. 1. Wetland soil samples**

116

117 The wetland soil samples selected for the present incubation experiments were collected in  
118 the highly contaminated wetland previously studied by Stetten et al. (2018b) [10], located  
119 near the former U-mine Ty Gallen in Brittany, France. Four types of soil layers were  
120 distinguished: a humus layer almost devoid of crystalline mineral components (O-type),  
121 organic-rich or organic-poor soil layers (A<sub>1</sub>- and A<sub>2</sub>-type layers respectively), and mine water  
122 deposit layers inherited from former mining operations (Figure SI-1b) [10].

123 For oxic incubations, three pristine reduced soil samples were selected below the water-table  
124 (Figure 1): an organo-mineral soil sample from a A<sub>2</sub>-type layer (C2-25 cm sample) and two

125 samples from mine water deposit layers (C2-30 cm and C6-33 cm samples). These samples  
126 were obtained from the C2 and C6 soil cores that were collected, vacuum-dried in a desiccator  
127 to a vacuum limit of 0.5 mbar and preserved under anoxic conditions [10].

128 For anoxic incubations, a humus layer sample (S1 sample; Figure 2) and a mine water deposit  
129 layer sample (S2 sample) were collected with a spade in the wetland in November 2015 and  
130 were stored under aerobic conditions at ambient temperature (21 °C), before the incubation  
131 experiments. These two latter samples were considered as representative of wetland soil  
132 samples exposed to a prolonged period of dryness and were thus used as starting samples for  
133 anoxic incubation experiments. In addition, a pristine organic-rich soil sample from the C2  
134 core corresponding to the A<sub>1</sub>-type layer above the water-table level (C2-5 cm sample) was  
135 also used as starting sample for anoxic incubation.

136

## 137 **2. 2. Incubation experiments**

138

139 For oxic incubation experiments, ~300 mg of dry soil was supplemented with ~1 ml of O<sub>2</sub>-  
140 free milli-Q water and were put in 50 ml tinted glass vials. The glass vials were then left open  
141 to air on a lab bench during 20 days in order to simulate a period of dryness. As these pristine  
142 soil samples exhibit similar uranium speciation [10], they were used as triplicates for this  
143 experiment (C2-25cm\_oxic; C2-30cm\_oxic; C6-33cm\_oxic). Anoxic incubation experiments  
144 were conducted in an anaerobic glove box (O<sub>2</sub> < 20 ppm vol. in N<sub>2</sub>) at IMPMC in sealed glass  
145 vials, in order to simulate a flooding period. Anoxic incubation of the S1 and S2 samples  
146 (S1\_anox\_bio\_1ml\_t20; S2\_anox\_bio\_1ml\_t20) were carried out with ~1 g of air-stored  
147 sample in 1ml of O<sub>2</sub>-free milli-Q water in 50 ml tinted glass vials sealed with a butyl rubber  
148 stopper for 20 days. Similar conditions were applied for the C2-5cm sample (C2-5

149 cm\_anox\_bio\_1ml\_t20) with ~300 mg of vacuum dry solid sample. No duplicates were  
150 carried out for these anoxic incubations.

151 Additional anoxic incubations of the S1 and S2 samples were carried out in duplicate, with a  
152 lower solid:liquid ratio (1:10) during 14 days (S1\_anox\_bio\_10ml\_t14\_A;  
153 S1\_anox\_bio\_10ml\_t14\_B; S2\_anox\_bio\_10ml\_t14\_A; S2\_anox\_bio\_10ml\_t14\_B; A and B  
154 corresponding to duplicates). Abiotic control incubations were conducted similarly but with  
155 samples previously autoclaved at 120°C for 20 min (S1\_anox\_abio\_10ml\_t14;  
156 S2\_anox\_abio\_10ml\_t14). For these 1:10 incubations, ~1 g of wet soil sample was  
157 supplemented by 10 ml of an anoxic 9 mg.L<sup>-1</sup> NaCl solution, put in 50 ml tinted glass vials,  
158 sealed by a butyl rubber stopper and tumbled for 14 days. For these latter experiments, NaCl  
159 was used as background electrolyte in order to avoid osmotic stress of the microbial biomass.  
160 At the end of these experiments, the incubation solutions were collected by centrifugation,  
161 filtered through 0.2 µm and an aliquot was acidified to pH~1 using 67 % HNO<sub>3</sub> for further  
162 dissolved U analysis.

163 At the end of all incubation experiments solid samples were vacuum dried in a desiccator  
164 placed in a glove box and stored under anaerobic conditions for further analyses.

165

### 166 **2. 3. Chemical Analyses procedures**

167

168 The initial solid soil samples of the C2 and C6 cores were previously analyzed for majors and  
169 traces elements [10]. For S1 and S2 initial solid samples, U content was measured by gamma  
170 spectrometry using a Ge HP type N - GEN27 detector at IRSN laboratory. For this purpose,  
171 solid samples were dried at 105 °C until a constant mass was obtained. They were 2mm  
172 sieved and ~40 g of the fraction inferior to 2mm was conditioned under vacuum in aluminized  
173 bag. Each sample was measured for 24 hours. Dissolved uranium in the 1:10 incubation



174 experiments were measured on acidified aliquots using an ICP-MS (X7 série 1  
175 ThermoFisher<sup>®</sup>). Dissolved Organic Carbon (DOC) and Dissolved Inorganic Carbon (DIC)  
176 were measured following the method described in Stetten et al. (2018b) [10].

177

## 178 **2. 4. X-ray Absorption Spectroscopy**

179

180 X-ray Absorption Near Edge Structure (XANES) and Extended X-ray Absorption Fine  
181 Structure (EXAFS) spectra at the Uranium L<sub>III</sub>-edge were collected on the 11-2 wiggler  
182 beamline at the Stanford Synchrotron Radiation Lightsource (SSRL) and on the CRG-FAME  
183 bending-magnet beamline at the European Synchrotron Radiation Facility (ESRF). The S1,  
184 S2 and anoxic incubation samples were analyzed on the 11-2 beamline at liquid N<sub>2</sub>  
185 temperature in fluorescence mode detection using a Si(220) double crystal monochromator  
186 and a 100-element solid state Ge array fluorescence detector. The oxic incubation samples  
187 were analyzed on the FAME beamline at liquid He temperature using a Si(220) double crystal  
188 monochromator and a 30 elements Ge fluorescence detector (ESRF) [23]. The incident beam  
189 energy was calibrated using a Y-foil in double transmission setup, with first inflection point  
190 set to E=17038 eV. XAS spectra collected at BL 11-2 were deadtime corrected and average  
191 using SIXPACK [24], and those collected at FAME-CRG beamline were averaged using  
192 ATHENA [25]. Energy calibration, normalization and background subtraction were  
193 performed with ATHENA.

194 In order to quantitatively determine the fractions of U(VI) and U(IV), U-L<sub>III</sub> edge XANES  
195 spectra were least-squares fit by linear combinations (LCF) of pure U(IV) and U(VI) model  
196 compounds XANES spectra, using a home-built code [10,19,26]. The chosen model  
197 compounds were a U(IV) humus sample and a synthetic U(VI)-humic acid model compound  
198 [10]. Uranium L<sub>III</sub>-edge EXAFS spectra of the incubated samples were least-squares fit by

199 linear combinations (LCF) of model compounds EXAFS spectra using a home-built code  
200 [10,19,26]. The model compounds chosen for this procedure were already used by Stetten et  
201 al. (2018b) [10] to fit the U L<sub>III</sub>-edge EXAFS spectra of the C2 and C6 soil samples and were  
202 previously interpreted using shell-by-shell fitting of the  $k^3$ -EXAFS spectra. This latter  
203 procedure yielded a description of the first and second neighbor shells around the U atom in  
204 these model compounds, as detailed in Stetten et al. (2018a) [5] and Stetten et al. (2018b)  
205 [10].

206

## 207 **2. 5. $\mu$ -XRF and $\mu$ -XAS data collection and analysis.**

208

209 Microfocused X-ray fluorescence ( $\mu$ -XRF) and U L<sub>III</sub>-edge micro-X-ray Absorption Near  
210 Edge Structure ( $\mu$ -XANES) spectroscopy analyses were conducted at the 2-3 beamline  
211 (SSRL) on polished sections of epoxy resin-embedded samples. The chosen samples were the  
212 mine water deposit sample (S2) and the product of its 20-day incubation under anoxic  
213 conditions (S2\_anox\_bio\_1ml\_t20). The  $\mu$ -XRF maps were collected at room temperature  
214 and an incident energy of 17200 eV, using a Vortex silicon drift detector, a beam spot size of  
215  $\sim 2 \times 2 \mu\text{m}$ , a step of  $10 \mu\text{m}$  and a counting time of 50ms. In order to limit U(IV) oxidation, a  
216 continuous flow of N<sub>2</sub> gas was applied on the surface of the samples during the  
217 measurements. Points of interest were selected on the  $\mu$ -XRF maps for U L<sub>III</sub>-edge  $\mu$ -XANES  
218 analyses using a Si(111) monochromator calibrated with elemental Y (K-edge at 17038 eV).  
219 The spectra were averaged using SIXPACK [24], normalized with ATHENA and analyzed by  
220 LCF with a custom-built software based on Levenberg-Marquardt least-squares minimization  
221 algorithm, using U(VI)-humic acid and U(IV)-humus model compound spectra as fitting  
222 components [5,10].

223

## 224 **2. 6. SEM-EDXS data collection and analysis**

225

226 Scanning electron microscopy analyses were performed at IMPMC, using a Zeiss ultra 55  
227 equipped with a Field Emission Gun. Backscattered images and EDX Spectra were collected  
228 at 15 keV with a working distance of 7.5 mm. The beam intensity was calibrated using the Cu  
229 K-alpha emission from a TEM copper grid. Semi-quantification of the EDXS spectra was  
230 performed using the Bruker<sup>®</sup> Esprit program and the Phi(rho,z) method using mineral  
231 standard spectra. The atomic content in U, Ca and P obtained with this method were then  
232 plotted in ternary diagrams using TERNPLOT [27].

233

## 234 **3. RESULTS AND DISCUSSION**

235

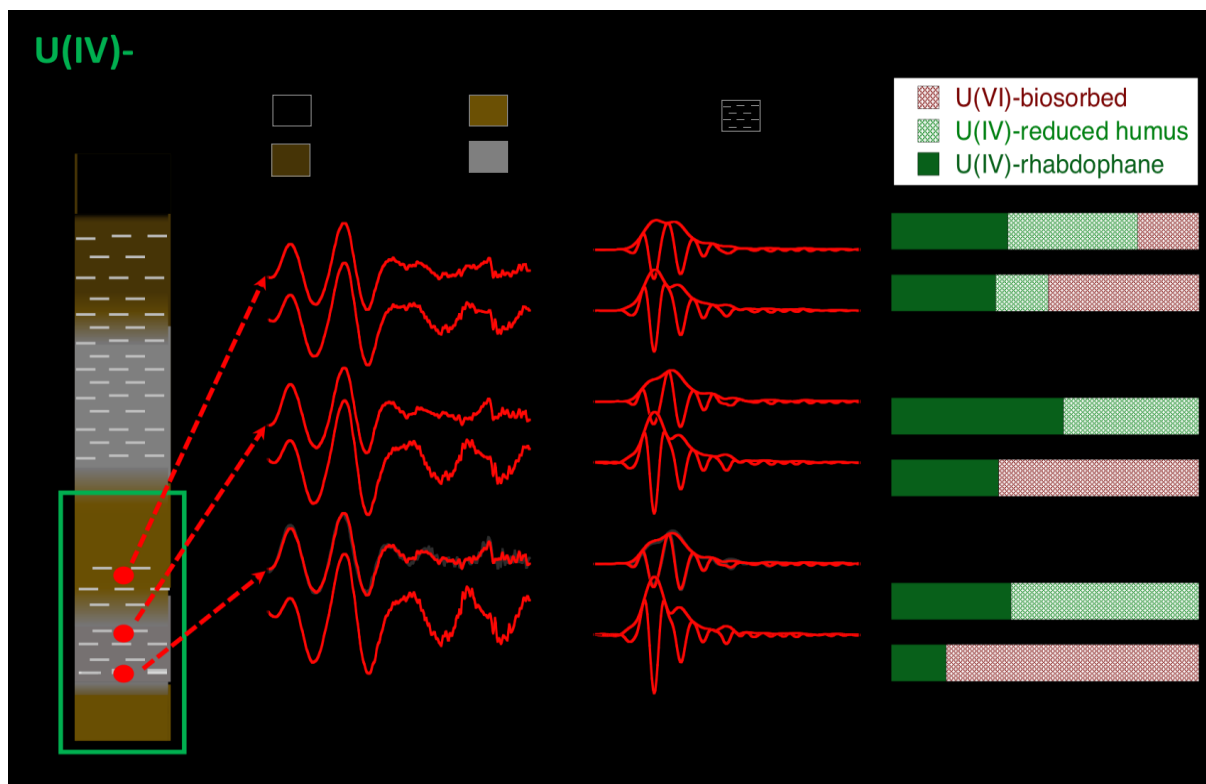
### 236 **3. 1. Effect of oxic incubations**

237

238 In the pristine wetland soil samples (C2-25 cm, C2-30 cm and C6-33 cm samples), ~40-58%  
239 U(IV) is present as mononuclear U(IV) species that were fitted with a U(IV)-humus  
240 component (Figure 1, Table SI-2). In this model compound, the U-P at 3.1 Å and U-C at 3.7  
241 Å EXAFS scattering paths are interpreted as bidentate and monodentate binding of U to  
242 phosphate and carboxylate groups, respectively [10]. The remaining ~35-55 % U in the soil  
243 samples was fitted with a U(IV)-rhabdophane component, used as proxy for U(IV) phosphate  
244 minerals identified as ningyoite and lermontovite [10].

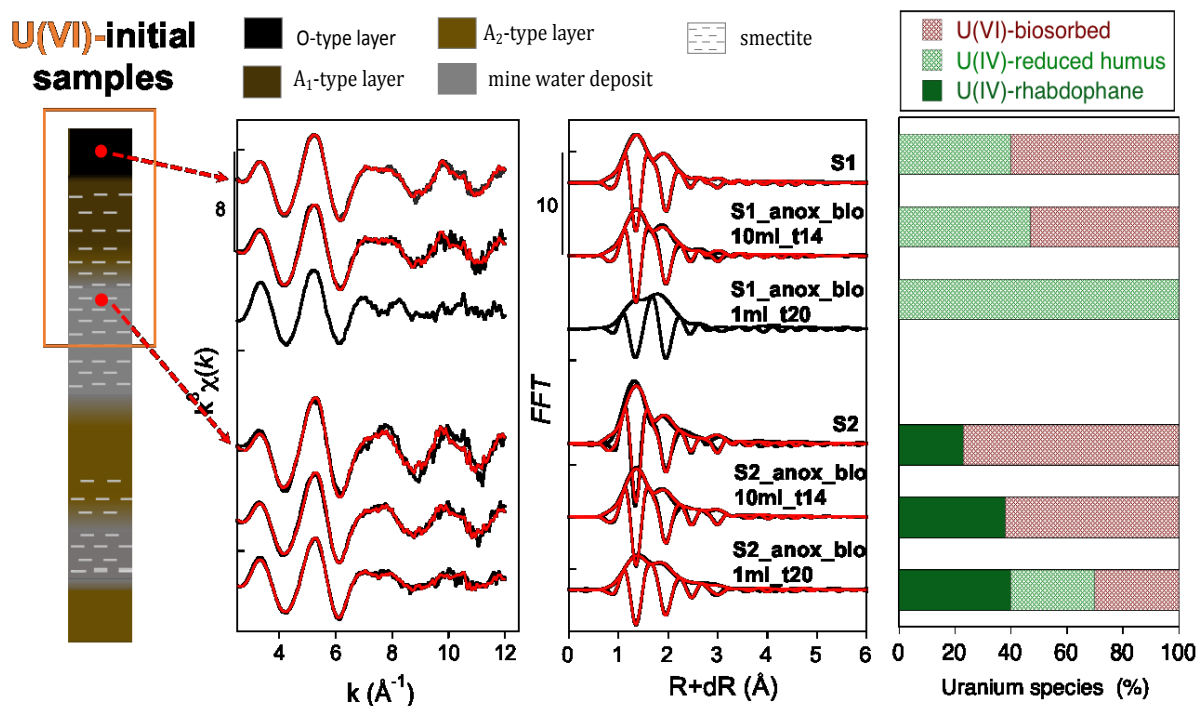
245 After 20 days of incubation under oxic conditions, the proportion of mononuclear U(IV)  
246 decreased down to 0-17% and that of U(IV)-phosphate mineral decreased down to 20-38 %,  
247 while mononuclear U(VI) increased to 51-89 %, as fitted with a U(VI)-biosorbed component  
248 (Figure 1; Table SI-2). In this latter model compound, U(VI) is bound to both phosphate in

249 monodentate geometry (U-P path at  $\sim 3.6$  Å; [10,28,29]) and carboxylate in bidentate  
250 geometry ( $\sim 2$  U-C paths at 2.9 Å; [10]). These data are consistent with the extent of U  
251 oxidation measured by LCF-XANES analysis of the same samples (Figure SI-3, Table SI-2).  
252 These results indicate complete oxidation of mononuclear U(IV) (Figure 1), except for sample  
253 C2-25 cm in which 17 % of mononuclear U(IV) persisted, which is consistent with the  
254 reported higher sensitivity of mononuclear species to redox change than crystalline U(IV)  
255 species [15-18]. In addition, except in sample C2-25 cm, 30-45 % of the U(IV)-phosphate  
256 minerals oxidized, likely into U(VI)-mononuclear species that were accounted by an increase  
257 of the U(VI)-biosorbed component in the LCF. These oxidized species could correspond to  
258 mononuclear U(VI) bound to organic phosphate groups of humic substances or biomass  
259 [28,30], as well as to inorganic phosphate groups [28,33] for instance at the surface of the  
260 oxidized U(IV)-minerals. Finally, 55-100 % of the U(IV) mineral phases were resistant to  
261 oxidation. The study of Newsome et al. (2015a) [16] also highlighted that U(IV)-phosphate  
262 minerals were partly resistant to oxidation. In this previous study, only  $\sim 20$  % of U(IV) was  
263 reoxidized from a biologically precipitated ningyoite-like phase after 90 days of air exposure.  
264  
265



266

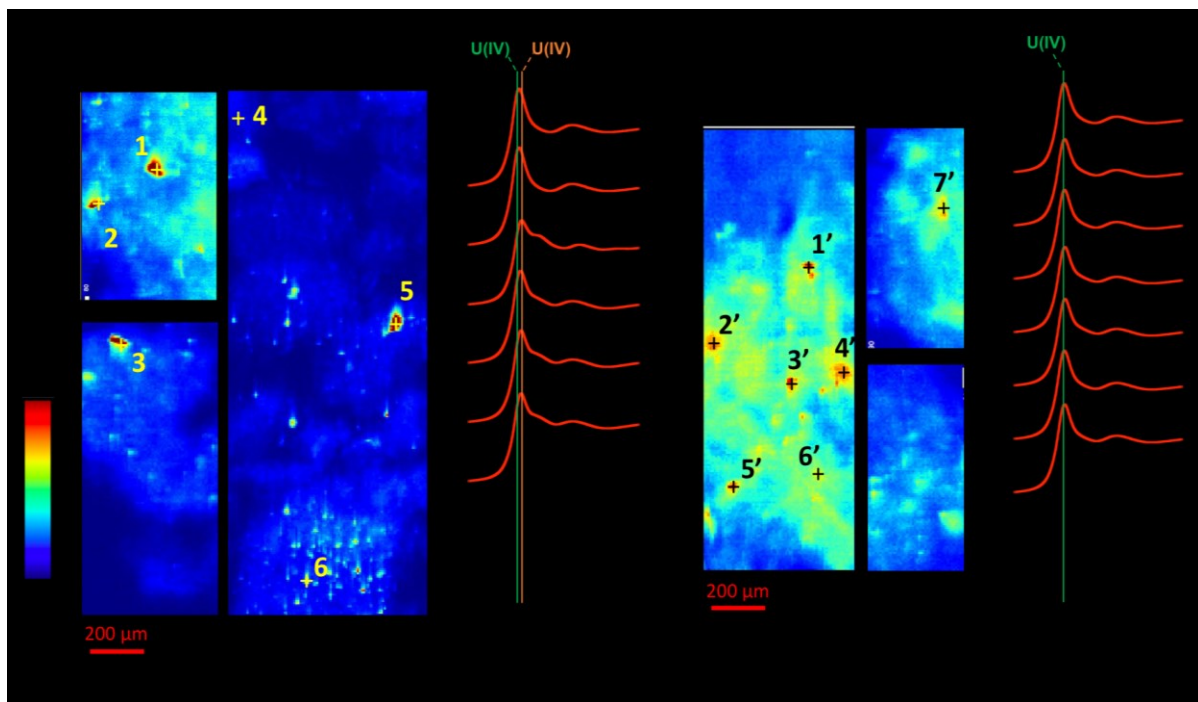
267 **Figure 1.** LCF analysis of U L<sub>III</sub> EXAFS data of pristine reduced soil samples from the A<sub>2</sub>-type and  
 268 mine water deposit layers and same samples after incubation under oxic conditions. Experimental and  
 269 fit curves are displayed in black and red colors respectively. Modulus and Imaginary part of the Fast  
 270 Fourier Transforms of the experimental and fit curves are also reported. The proportions of the fitting  
 271 components, represented by the bar-diagram, are normalized to 100 %. Non-normalized results and  
 272 uncertainties are given in Table SI-2.  
 273



274

275 **Figure 2.** LCF analysis of U L<sub>III</sub> EXAFS data of initial samples (S1 and S2 samples) and incubated  
276 samples under anoxic conditions. Experimental and fit curves are displayed in black and red colors  
277 respectively. Modulus and Imaginary part of the Fast Fourier Transforms of the experimental and fit  
278 curves are also reported. The proportions of the fitting components, represented by the histograms, are  
279 normalized to 100 %. Non-normalized results and uncertainties are given in Table SI-2.  
280

281 In order to evaluate the consequences of the oxidation of U(IV)-mineral phases on the  
282 potential U redistribution in the soil matrix, the spatial distribution of U was investigated by  
283  $\mu$ -XRF mapping. The sample studied was a polished cut of the air dried S2 mine water  
284 deposit sample that exhibited the same U speciation as C2-30cm\_oxic and C6-33 cm samples  
285 (Figures 2 and 1; Table SI-2). Uranium was mainly found by  $\mu$ -XRF as hot spots (Figure 3a)  
286 suggesting the presence of U minerals grains. LCF analysis of  $\mu$ -XANES spectra collected at  
287 hotspots #1 and #2 showed only U(IV) (Table SI-4). These grains could thus be interpreted as  
288 corresponding to abundant ningyoite  $\text{CaU}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$  (Figures 4d,4a, SI-4a), and to a lesser  
289 extent lermontovite  $\text{U}(\text{PO}_4)(\text{OH}) \cdot \text{H}_2\text{O}$  (Figures 4c; SI-4b) grains identified by SEM-EDXS in  
290 the same S2 sample.



291

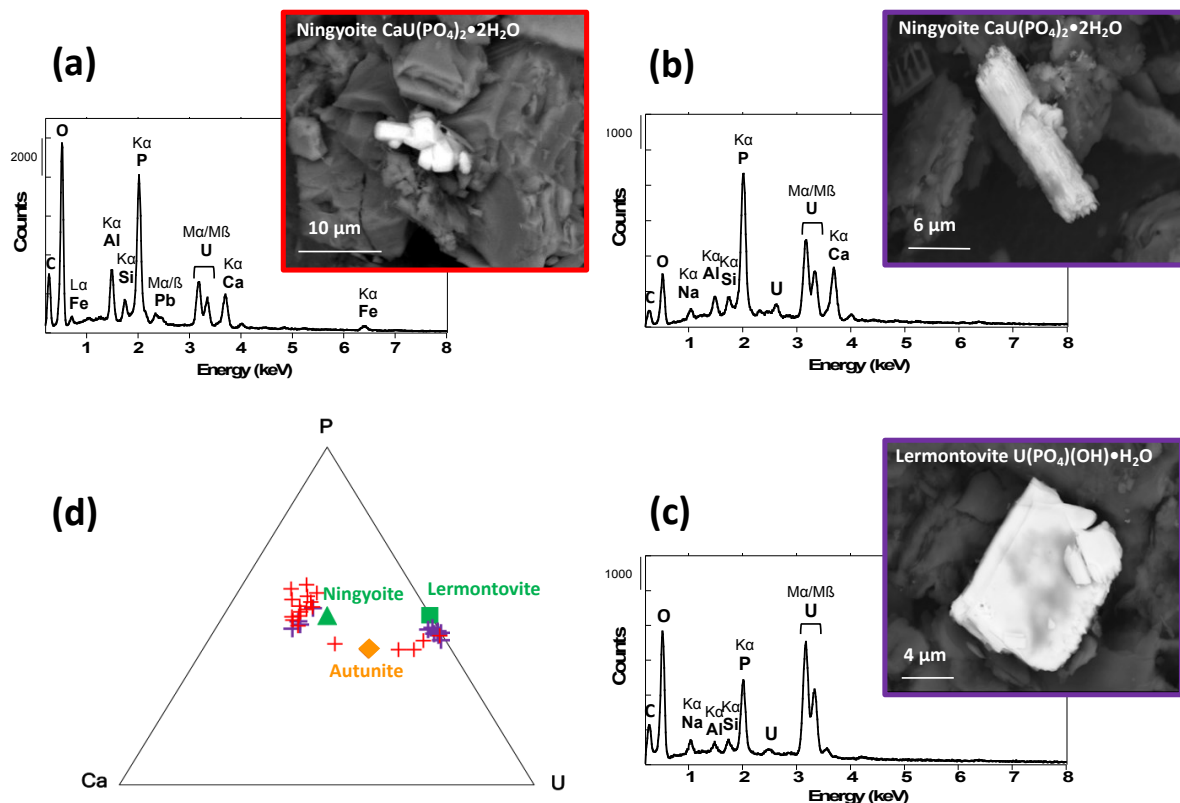
292 **Figure 3:**  $\mu$ -XRF mapping of U distribution and  $\mu$ -XANES analyses of U redox state in a mine water  
293 deposit sample collected in the wetland (Sample S2) and the product of its incubation under anoxic  
294 conditions (S2\_anox\_bio\_1ml\_t20); The data were obtained by analyzing polished sections of epoxy-

295 resin embedded samples. The  $\mu$ -XRF elemental maps are displayed in the same color intensity scale  
296 and show the spreading of U around U-rich grains in the incubated sample. The color scale of the  $\mu$ -  
297 XRF elemental maps corresponds to the background subtracted area of the U L $\alpha$  emission line,  
298 increasing from blue to red with U content; U oxidation state was determined by LCF fit of the micro-  
299 XANES spectra collected on selected points in the XRF maps using U(VI)-humic acid and U(IV)-  
300 humus model compound spectra as fitting components. Spectrum of these model compounds are  
301 plotted in black lines. Experimental and calculated spectra are plotted as black and red solid lines,  
302 respectively. No oxidation of U(IV) was reported during the time of the measurement as shown in  
303 Figure SI-6. Non-normalized results and uncertainties are given in Table SI-4.

304

305 Spot #3 exhibited only U(VI) could be attributed to crystalline autunite  
306  $\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$  that was scarcely observed in this sample (Figure SI-4c). Since no  
307 U(VI)-bearing mineral was detected by EXAFS analysis of this sample, autunite likely do not  
308 account for more than 10 % of total U in the S2 sample. The formation of an autunite-like  
309 mineral after oxidation of an amorphous U(IV)-phosphate was reported by Latta et al. (2016)  
310 [22]. By contrast, in the mine water deposit S2 sample, the scarce and massive autunite  
311 crystals (Figure SI-4c) differs from the reticulated microcrystals observed by Latta et al.  
312 (2016) [22], and thus did not likely formed by the oxidation of ningyoite. Instead, our  $\mu$ -XAS  
313 dataset and the non-detectability of U(VI)-mineral by EXAFS suggest that U(IV) minerals  
314 grains oxidized into mixed U(IV)/U(VI) grains, in which U(VI) may be mainly present as  
315 sorbed mononuclear U(VI). Indeed, the observation of U(VI) in hot spots #4, 5 and 6  
316 suggested that these U-rich grains contain sorbed U(VI) that could coat or replace former  
317 U(IV)-phosphate phases. Finally, the  $\mu$ -XRF map of the initial S2 sample (Figure 3) displayed  
318 a diffuse distribution of U, which was consistent with the occurrence of U(VI) mononuclear  
319 complexes determined by LCF-EXAFS (Figure 2; Table SI-2). Altogether, these results  
320 confirmed that a fraction of U is distributed in the soil matrix, which might account for the  
321 non-crystalline U(VI) species observed by EXAFS analysis.

322



323

324

325 **Figure 4.** Backscattered electrons SEM images and semi-quantitative EDXS analyses of U minerals in  
 326 the mine water deposit (S2 sample; red crosses) and in the same sample after anoxic incubation  
 327 (S2\_anox\_bio\_1ml\_t20; purple crosses). (a) Ningyoite in the S2 sample, (b) Ningyoite with prismatic  
 328 morphology [34] and (c) Lermontovite with platy morphology [37] in the S2\_anox\_bio\_1ml\_t20  
 329 incubated sample. (d) Ternary diagram showing P, Ca and U compositions of the U-mineral grains  
 330 observed in these samples. Ideal stoichiometric composition of Autunite  $\text{Ca}(\text{UO}_2)_2(\text{PO}_4)_2 \cdot 11\text{H}_2\text{O}$   
 331 (orange diamond), ningyoite  $\text{CaU}(\text{PO}_4)_2 \cdot 2\text{H}_2\text{O}$  (green triangle) and lermontovite  $\text{U}(\text{PO}_4)(\text{OH}) \cdot \text{H}_2\text{O}$   
 332 (green square) minerals are also reported.

333

334

### 335 3. 3. Effect of anoxic incubations

336

337 Two types of samples were incubated under anoxic conditions: S1 and C2\_5cm humus  
 338 samples in which U mainly occurs as mononuclear species and S2 mine water deposit sample  
 339 in which U is partitioned between U-phosphate minerals and mononuclear species. Full  
 340 reduction of U(VI) was observed for S1 and C2\_5 cm, whereas partial reduction was  
 341 observed for S2 (Figures 2 and SI-3; Tables SI-2 and SI-3) after 20 days. The extent of U



342 reduction was systematically lower at shorter incubation times (14 days) and at lower  
343 solid:liquid ratio (1:10) (Figure 2; Table SI-2). We may hypothesize that a lag phase in the  
344 growth of the microbial biomass could have delayed the effective microbial-mediated U  
345 reduction during the first 14 days of incubation. The reduction of U(VI) to U(IV) species  
346 could be enhanced by both enzymatic and abiotic U(VI) reduction pathways [5, 37-40]. Here,  
347 the role of microorganisms in U reduction was confirmed by the absence of U reduction in the  
348 abiotic incubations conducted for 14 days at solid:liquid ratio (1:10) (Figure SI-3; Table SI-  
349 3). The higher reduction rate observed for S1 samples could be explained by the high organic  
350 carbon content of these samples that would have enhanced biotic reduction processes [37],  
351 likely *via* direct or indirect dissimilatory reduction of U(VI) [38-40]. Even if it has been  
352 shown that direct enzymatic U reduction pathways dominate in sediments [36, 39], we cannot  
353 exclude that U is reduced by indirect microbial pathways, for example *via* the release of  
354 abiotic reductants such as Fe(II), itself possibly produced by the bioreduction of Fe(III)-  
355 bearing minerals [5, 40].

356 For the S1 sample, the best EXAFS LCF was obtained with 60% of U(VI)-biosorbed model  
357 compound and 40% of U(IV)-humus model compound (Figure 2; Table SI-2). This latter  
358 model compound also used by Stetten et al. (2018b) [10] corresponds to the product of  
359 reduction of the S1 sample after 20 days of incubation at high solid:liquid ratio  
360 (S1\_anox\_bio\_1ml\_t20 sample), in which U was fully reduced (Figures 2 and SI-3; Tables  
361 SI-2 and SI-3). Shell-by-shell fitting of the  $k^3$ -EXAFS spectra was performed in Stetten et al.  
362 (2018b) [10] for this humus model compound and indicated the presence of bidentate U(IV)-  
363 phosphate and monodentate U(IV)-carboxylate complexes.

364 Altogether, these results suggest that the reduction of monodentate U(VI)-phosphate and  
365 bidentate U(VI)-carboxylate complexes lead to the formation of bidentate U(IV)-phosphate  
366 and monodentate U(IV)-carboxylate complexes, without formation of U(IV) mineral, likely

367 due to the abundance of organic-bound phosphate and carboxylate ligands inhibiting uraninite  
368 formation [41,42]. Additionally, the lack of free orthophosphate may also prevent U(IV)-  
369 phosphate minerals precipitation in such organic-rich layers.

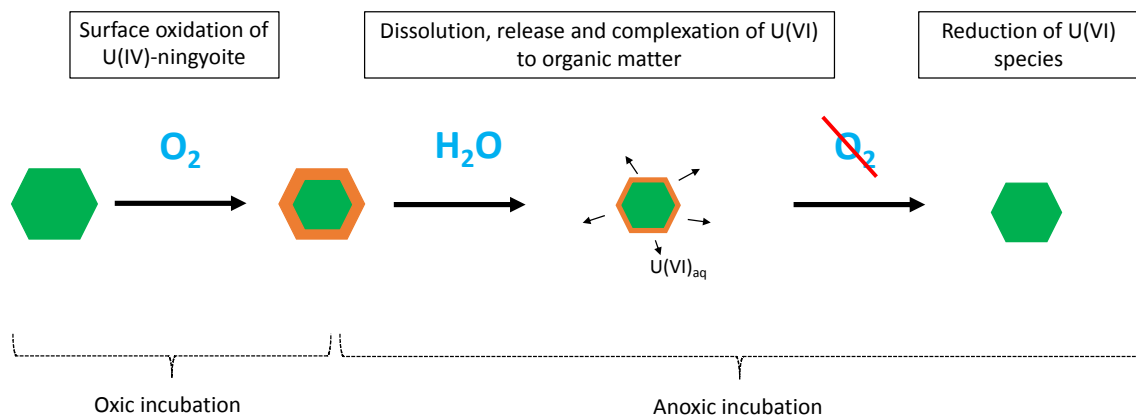
370 Sample S2 was chosen to investigate the effect of anoxic conditions on oxidized mine water  
371 deposits that contained mainly sorbed mononuclear U(VI) species (~80 %) and recalcitrant  
372 U(IV)-phosphate minerals (~20 %) (Figures 2, 3 and 4; Table SI-2). LCF-EXAFS results  
373 indicated that after 14 days of anoxic incubation (sample S2\_anox\_bio\_10ml\_t14\_A), the  
374 proportion of U(IV)-phosphate minerals increased up to 40% whereas sorbed mononuclear  
375 U(VI) species decreased by 20 % (Figures 2; Table SI-2). After 20 days of anoxic incubation  
376 (sample S2\_anox\_bio\_1ml\_t20), the proportion of U(IV) phosphate did not increase above  
377 ~40 %. In contrast, the proportion of sorbed U(VI) decreased down to 31 % while a  
378 proportion of ~30 % of U was found as U(IV)-humus (Figure 2, Table SI-2). This latter result  
379 could be interpreted as the reduction of sorbed U(VI) to sorbed U(IV), as discussed in the  
380 previous section for S1 sample.

381 The increase in the proportion of U(IV)-phosphate minerals after 14 and 20 days of anoxic  
382 incubation suggests the formation of U(IV)-phosphate minerals. Such a reaction may likely  
383 occur *via* reduction of U(VI) bound to phosphate mineral surfaces since the reduction of  
384 organically-bound U(VI) was found to lead to organically bound U(IV) as shown for S1  
385 sample. The formation of U(IV)-phosphate such as ningyoite-like minerals from the  
386 bioreduction of inorganic U(VI)-phosphate species is supported by previous laboratory assays  
387 [21,43]. Indeed, these previous studies reported slow reduction kinetics of Hydrated Uranium  
388 Phosphate (HUP), with ~20-80 % of U(VI) remaining in the solid phase after 10 days  
389 incubation of a biogenic HUP, and ~30 % of U(VI) after 20 days incubation of abiotic HUP.  
390 To this regard, the lower reduction rate observed in S2 sample compared to the organic S1

391 sample may be due to a slower bioreduction rate of inorganic U(VI)-phosphate species  
392 compared to organic ones.

393 In the incubated S2\_anox\_bio\_1ml\_t20 sample,  $\mu$ -XANES spectra for U(IV)/U(VI) ratio  
394 showed that a minor (10-35%) proportion of U(VI) remained at hot spots (Figure 3b; Table  
395 SI-4), suggesting almost complete reduction of U(VI)-bearing mineral grains. In addition, the  
396 amount of U in the soil matrix appeared to be higher after than before incubation (Figures 3a  
397 and 3b), suggesting U migration from U hot-spots to the surrounding organic-rich soil matrix  
398 (Figure 3b). One may hypothesize that such migration process could have occurred at the  
399 beginning of the incubation experiment *via* the dissolution of U(VI) initially sorbed to  
400 oxidized phosphate minerals grains (Figure 5). The presence of U(IV) in the soil matrix  
401 (Figure 3b; Table SI-4) suggests that U(VI) was then further reduced to U(IV) mononuclear  
402 species under the reducing conditions that developed in these anoxic incubations.

403



404

405

406 **Figure 5.** Schematic view of the effect of successive oxic and anoxic incubations on the redistribution  
407 of U from U(IV)-phosphate mineral such as ningyoite in mine water deposits layers in the studied  
408 contaminated wetland.

409

410 SEM-EDX observations revealed a higher amount of lermontovite-like minerals (Figures SI-  
411 5c) after anoxic incubation, *i.e.* in the S2\_anox\_bio\_1ml\_t20 sample, than in the S2 sample

412 (Figures 4d, SI-4b). Conversely, ningyoite was also found in the S2\_anox\_bio\_1ml\_t20  
413 sample (Figure SI-5ab) but in a lower amount than in the S2 sample (Figures 4d, SI-4a).  
414 Moreover, some ningyoite grains observed in this sample exhibited a weathered morphology  
415 with acicular apices and eroded surfaces (Figures 4b, SI-5b), whereas lemontovite grains  
416 exhibited massive and non-weathered morphology (Figure 4c, SI-5c). These results suggest  
417 that sorbed U(VI) previously resulting from the oxidation of ningyoite grains could have  
418 dissolved at the beginning of the anoxic incubation, whereas lemontovite-like minerals may  
419 be more recalcitrant to oxidation and dissolution. Moreover, we hypothesize that autunite-like  
420 minerals of the S2 sample could have been transformed into lemontovite-like minerals *via* a  
421 loss in Ca during the reduction process. Indeed, lemontovite is known to mainly occur as  
422 secondary mineral in reducing environments and is found in intimate intergrowth with U(VI)-  
423 phosphates from the autunite or meta-autunite group [35]. In addition, although the crystal  
424 structure of lemontovite is still unsolved, it has been reported to crystallize in the  
425 orthorhombic system and could thus have a layered structure like autunite [35], which could  
426 favor transformation of autunite phase to lemontovite *via* the loss of  $\text{Ca}^{2+}$  in the interlayer  
427 and reduction of U(VI) to U(IV). The scarcity of autunite in the wetland soil, while it is an  
428 abundant mineral of the oxidation zone of the Ty Gallen ore ([https://www.mindat.org/min-](https://www.mindat.org/min-433.html)  
429 [433.html](https://www.mindat.org/min-433.html)), and in return the abundance of lemontovite in the reduced mine water deposits  
430 layers of the wetland could be explained by a progressive transformation of autunite into  
431 lemontovite under prolonged reducing conditions. In contrast, the progressive oxidation and  
432 dissolution of ningyoite, after repeated redox cycles could explain the redistribution of U  
433 toward the organic compartment of the wetland soil. Finally, neoformation of ningyoite under  
434 reducing conditions cannot be excluded since Khijniak et al. (2005) [43] reported the  
435 formation of a ningyoite-like mineral after bioreduction of a meta-autunite mineral uramphite

436  $(\text{NH}_4)(\text{UO}_2)(\text{PO}_4).3\text{H}_2\text{O}$  at  $65^\circ\text{C}$  in a laboratory experiment with *Thermoterrabacterium*  
437 *ferrireducens*.

438

### 439 **3.4. Uranium mobility after a redox cycle: experiment vs field investigation**

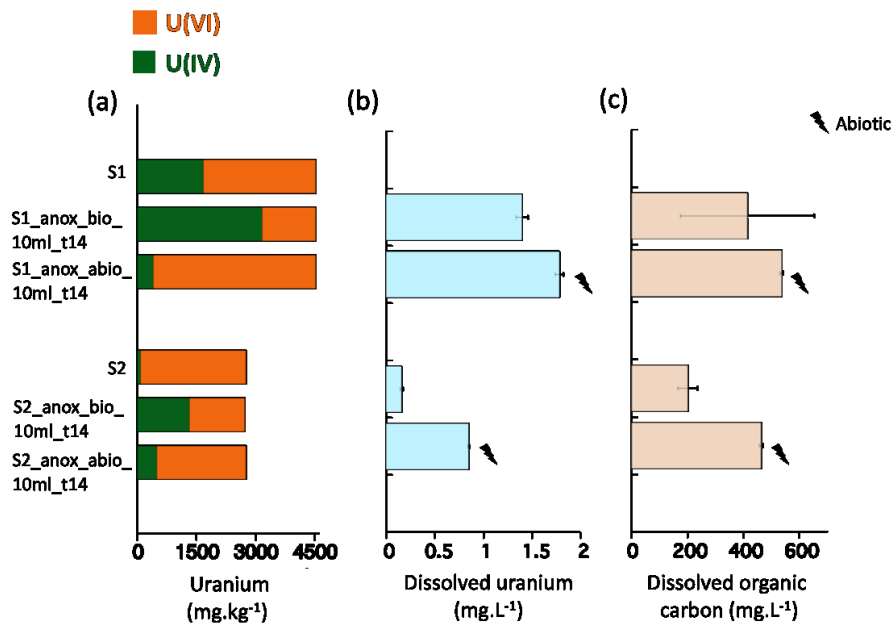
440

441 In order to investigate the release of U during a flooding event following a dry period, we  
442 have measured dissolved U concentrations in  $0.2\ \mu\text{m}$  filtered supernatant solutions collected  
443 at the end of the anoxic incubation. Figure 6 shows the dissolved uranium concentrations at  
444 the end of the 1:10 incubation performed under biotic and abiotic anoxic conditions. Uranium  
445 concentrations in the humus sample S1 experiments (S1\_anox\_biotic\_10ml ( $[\text{U}] = 1520\ \mu\text{g.L}^{-1}$ )  
446  $^1$ ) and S1\_anox\_abiotic\_10ml ( $[\text{U}] = 2156\ \mu\text{g.L}^{-1}$ ) are far higher than U concentrations of the  
447 mine water deposit layer sample S2 experiments (S2\_anox\_biotic\_10ml ( $[\text{U}] = 160\ \mu\text{g.L}^{-1}$ ) and  
448 S2\_anox\_abiotic\_10ml ( $[\text{U}] = 913\ \mu\text{g.L}^{-1}$ )). Such high concentrations represented however less  
449 than 0.05 % of both S1 and S2 solid U concentrations ( $4516\ \text{mg.kg}^{-1}$  for S1 and  $2750\ \text{mg.kg}^{-1}$   
450 for S2). These results showed that, in the conditions of our experiments, the remobilization of  
451 U in the aqueous media after a complete redox cycle is limited and did not exceed 0.05 % of  
452 the total U solid pool. Despite that the aqueous environments established in the experiments  
453 are simplified, these results were consistent with our previously developed hypothesis of the  
454 redistribution and scavenging of U in the solid phase during the incubation experiment,  
455 mostly by adsorption and subsequent reduction (Figure 3a). These results agree with a recent  
456 study showing a release of U of 0.8% from initially reoxidized mononuclear U species under  
457 reducing conditions in carbonate-free media [18].

458 In all anoxic incubations, dissolved organic carbon release was high, especially for organic S1  
459 samples (Table SI-5) and was positively correlated to dissolved U concentration, which  
460 strongly suggests solubilization of U by complexation to soluble organic ligands. Such results

461 are in line with the study of Grybos et al. (2009) [44], in which large quantities of dissolved  
462 organic matter were reported to be solubilized under reducing conditions in a wetland soil,  
463 likely because of the reductive dissolution of ferric iron oxides involved in the stabilization of  
464 organic matter. Organic species have been shown to complex U(VI) and U(IV) in aqueous or  
465 colloidal mobile species in the laboratory and the field [7,45-46]. In the study of Seder-  
466 Colomina et al. (2018) [18], U release under anoxic condition was mainly explained by  
467 inorganic carbon complexation except in experiments without added  $\text{HCO}_3^-$ , where a minor U  
468 mobilization was explained by the release of organic carbon. In addition, in the same study,  
469 geochemical modeling confirms the presence of dissolved U(VI) and U(IV) organic species as  
470 major species in the aqueous media in carbonate-free experiments. In our experiments, DIC  
471 was found negligible to the opposite of DOC. In addition, EXAFS results show an important  
472 fraction of U bound to organic moieties (Figure 2) in the solid phase. The final dissolved U in  
473 all our incubated samples could therefore be associated with soluble or colloidal forms of  
474 organic carbon. Then, for the S1 sample, the release of U could be due to the release of U-  
475 bearing organic molecules. For the S2 sample, the dissolution of U(VI)-phosphate minerals  
476 could be more limited because of the lower amount of organic carbon, but still promoted by  
477 complexation to soluble organic ligands. Accordingly, laboratory experiment have reported  
478 important remobilization of U after reoxydation of non crystalline U(IV) [15]. In contrast,  
479 U(IV)-phosphate minerals were found to be more recalcitrant to oxidative remobilisation, but  
480 significant release of U was also reported [16]. Here, our results suggested that such  
481 remobilization could be limited in organic-rich environmental samples, *via* U sorption onto  
482 particulate matter. Indeed, uranium remobilization was found to be minor (< 1% of total solid  
483 U). However, the high concentrations of U in the soil (< 4500 mg.kg<sup>-1</sup>) could contribute to  
484 significant dissolved U concentrations in pore waters under reducing condition. Moreover, the  
485 presence of U(IV)-phosphate minerals that are subject to reoxidation events and subsequent

486 dissolution appears to play an important role in the long-term release of U under redox  
 487 fluctuations, even in the absence of carbonates. In addition, mobility of U bound to particulate  
 488 and colloidal organic matter would be important to consider in order to evaluate U fluxes out  
 489 of the studied wetland.  
 490



491  
 492 **Figure 6.** (a) Uranium concentration and redox in S1, S1\_anox\_bio\_10ml\_t14,  
 493 S1\_anox\_abio\_10ml\_t14 and S2, S2\_anox\_bio\_10ml\_t14\_A and S2\_anox\_abio\_10ml\_t14 solid  
 494 samples. (b) U concentration in biotic and abiotic incubation solutions of S1\_anox and S2\_anox  
 495 incubated samples at the end of the 14 days (c) Dissolved organic carbon concentrations in biotic and  
 496 abiotic incubation solutions of S1\_anox and S2\_anox incubated samples at the end of the 14 days.  
 497 Corresponding data are reported in Table SI-5. Errors bars represent the standard deviation over the  
 498 duplicates samples for biotic incubations and the relative uncertainty of the measurement for abiotic  
 499 incubations.

500

#### 501 4. CONCLUSION

502 In our previous study of the same wetland, we showed an important redistribution of U from  
 503 U(IV)-phosphates to mononuclear U(VI)/U(IV) species, interpreted as the result of the  
 504 oxidative dissolution of U-phosphate minerals and subsequent adsorption of U by organic  
 505 matter occurring during redox oscillations [10]. Here, our results confirm that a drying period  
 506 could enhance rapid oxidation of non crystalline species and, to a lesser extent, of U(IV)-

507 phosphate minerals. Even if such oxidation has been shown to lead to an important  
508 remobilization of U, especially if it is under non-crystalline forms [15-16], our results show  
509 that in organic-rich samples, the adsorption onto soil organic ligands limits U spreading from  
510 mineral hotspots. In addition, the dissolution of U(IV)-phosphate minerals upon redox cycling  
511 raises important issues concerning the long-term fate of U in seasonally saturated  
512 environments. Our results further suggest that mononuclear U(VI) species resulting from such  
513 oxidative dissolution processes can be fully reduced after 20 days under reducing condition.  
514 Thus, in organic-rich and carbonate-free environments, U mobility appears to be mainly  
515 controlled by the fate of organic matter rather than by redox processes [7,47], which is  
516 supported by the observed positive correlation between aqueous U and DOC content in our  
517 anoxic incubation solutions. This observation raises questions concerning the long-term fate  
518 of U in highly contaminated organic-rich wetlands, since U could be slowly mobilized under  
519 both oxic and reducing conditions. As a perspective of this work, we hope that the molecular  
520 scale processes identified in this study will help to design efficient monitoring and modeling  
521 approaches to manage U migration at the landscape scale in organic-rich environments.

522

## 523 **ACKNOWLEDGEMENTS**

524 This study was conducted within the framework of the IRSN/IMPIC collaborative research  
525 program number LS 20218/C151903-P150778. The authors thank Jean-Louis Hazemann and  
526 Olivier Proux for their help on the FAME-CRG beamline (ESRF) and Samuel M. Webb and  
527 Ryan Davis for their help during measurements at the 2-3 and 11-2 SSRL beamlines  
528 respectively. Portion of this research was carried out at the Stanford Synchrotron Radiation  
529 Light Source, a national user facility operated by Stanford University, and behalf of the US  
530 DOE Office of Basic Energy Sciences. Jean-Claude Boulliard is greatly acknowledged for  
531 having supplied mineral species samples from the IMPIC mineralogy collection. The SEM-



532 EDXS and XRD platforms of IMPMC are also acknowledged for having provided time for  
533 analyses. The anonymous owner of the site is also acknowledged for giving us access to the  
534 site. Lucie Stetten's PhD has been granted by Region Ile de France, DIM R2DS PhD grant  
535 number 2015-03.

536 **REFERENCES**

537

538 [1] D. Brugge, J.L. De Lemos, B. Oldmixon, Exposure Pathways and Health Effects  
539 Associated with Chemical and Radiological Toxicity of Natural Uranium: A Review,  
540 *Reviews on environmental health* 20 (2005).

541

542 [2] S.R. Taylor, Abundance of chemical elements in the continental crust: a new table,  
543 *Geochim. Cosmochim. Acta* 28 (1964) 1273-1285.

544

545 [3] M. Cuney, The extreme diversity of uranium deposits, *Miner Deposita* 44 (2009) 3–9.

546

547 [4] K.M. Campbell, R.K. Kukkadapu, N.P. Qafoku, A.D. Peacock, E. Lesher, K. H.  
548 Williams, J.R. Bargar, M.J. Wilkins, L. Figueroa, J. Ranville, J.A. Davis, P. E. Long,  
549 Geochemical, mineralogical and microbiological characteristics of sediment from a naturally  
550 reduced zone in a uranium-contaminated aquifer, *Applied Geochemistry* 27 (2012) 1499–  
551 1511.

552

553 [5] L. Stetten, A. Mangeret, J. Brest, M. Seder-Colomina, P. Le Pape, M. Ikogou, N. Zeyen,  
554 A. Thouvenot, A. Julien, G. Alcalde, J.-L. Reyss, B. Bombled, C. Rabouille, L. Olivi, O.  
555 Proux, C. Cazala, G. Morin, Geochemical control on the reduction of U(VI) to mononuclear  
556 U(IV) species in a contaminated lake sediment, *Geochim. Cosmochim. Acta* 222 (2018a)  
557 171-186.

558

559 [6] A. Cuvier, L. Pourcelot, A. Probst, J. Prunier, G. Le Roux, Trace elements and Pb isotopes  
560 in soils and sediments impacted by uranium mining, *Science of the Total Environment* 566–  
561 567 (2016) 238–249.

562

563 [7] Y. Wang, M. Fruttschi, E. Suvorova, V. Phrommavanh, M. Descostes, A. A. A. Osman, G.  
564 Geipel, R. Bernier-Latmani, Mobile uranium(IV)-bearing colloids in a mining impacted  
565 wetland, *Nature Geosciences* 4:2942 (2013). doi: 10.1038/ncomms3942.

566

567 [8] Y. Wang, A. Bagnoud, E. Suvorova, E. McGivney, L. Chesaux, V. Phrommavanh, M.  
568 Descostes, R. Bernier-Latmani, Geochemical control on uranium(IV) mobility in a mining-  
569 impacted wetland, *Environ. Sci. Technol.* 48 (2014) 10062–10070.

570

571 [9] A. Mangeret, P. Blanchart, G. Alcalde, X. Amet, C. Cazala, M.O. Gallerand, An evidence  
572 of chemically and physically mediated migration of  $^{238}\text{U}$  and its daughter isotopes in the  
573 vicinity of a former uranium mine, *Journal of Environmental Radioactivity* 195 (2018) 67-71.  
574  
575 [10] L. Stetten, P. Banchart, A. Mangeret, P. Lefebvre, P. Le Pape, J. Brest, P. Merrot, A.  
576 Julien, O. Proux, S.M. Webb, J. R. Bargar, C. Cazala, G. Morin, Redox fluctuations and  
577 organic complexation govern uranium redistribution from U(IV)-phosphate minerals in a  
578 mining-polluted wetland soil, Brittany, France, *Environ. Sci. Technol.* 52(22) (2018b) 13099-  
579 13109.  
580  
581 [11] A. Schöner, C. Noubactep, G. Büchel, M. Sauter, Geochemistry of natural wetlands in  
582 former uranium milling sites (eastern Germany) and implications for uranium retention,  
583 *Chemie der Erde* 69 (2009) 91–107.  
584  
585 [12] C. Mikutta, P. Langner, J.R. Bargar, R. Kretzschmar, Tetra- and Hexavalent uranium  
586 forms bidentate- mononuclear complexes with particulate organic matter in a naturally  
587 uranium- enriched peatland, *Environ. Sci. Technol.* 50 (2016) 10465-10475.  
588  
589 [13] D. Li, D.I. Kaplan, H.-S. Chang, J.C. Seaman, P.R. Jaffé, P. Koster van Groos, K.G.  
590 Scheckel, C.U. Segre, N. Chen, D.-T. Jiang, Spectroscopic evidence of uranium  
591 immobilization in acidic wetlands by natural organic matter and plant roots, *Environ. Sci.*  
592 *Technol.* 49 (2015) 2823–2832.  
593  
594 [14] D.S. Alessi, B. Uster, H. Veeramani, E.I. Suvorova, J.S. Lezama-Pacheco, J.E. Stubbs,  
595 J.R. Bargar, R. Bernier-Latmani, Quantitative Separation of Monomeric U(IV) from  $\text{UO}_2$  in  
596 Products of U(VI) Reduction, *Environ. Sci. Technol.* 46 (2012) 6150–6157.  
597 <https://doi.org/10.1021/es204123z>.  
598  
599 [15] J.M. Cerrato, M.N. Ashner, D.S. Alessi, J.S. Lezama-Pacheco, R. Bernier-Latmani, J.R.  
600 Bargar, D.E. Giammar, Relative reactivity of biogenic and chemogenic uraninite and biogenic  
601 noncrystalline U(IV), *Environ. Sci. Technol.* 47 (2013) 9756-9763.  
602  
603 [16] L. Newsome, K. Morris, D. Trivedi, A. Bewsher, J.R. Lloyd, Biostimulation by  
604 Glycerol Phosphate to Precipitate Recalcitrant Uranium(IV) Phosphate, *Environ. Sci.*  
605 *Technol.* 49 (2015a) 11070-11078.

606  
607 [17] L. Newsome, K. Morris, S. Shaw, D. Trivedi, J.R. Lloyd, The stability of microbially  
608 reduced U(IV): impact of residual electron donor and sediment ageing, *Chemical Geology*  
609 409 (2015b) 125–135.

610  
611 [18] M. Seder-Colomina, A. Mangeret, L. Stetten, P. Merrot, O. Diez, A. Julien, E. Barker, A.  
612 Thouvenot, J. Bargar, C. Cazala, G. Morin, Carbonate facilitated mobilization of uranium  
613 from lacustrine sediments under anoxic conditions, *Environmental Science & Technology* 52  
614 (2018) 9615-9624.

615  
616 [19] G. Morin, A. Mangeret, G. Othmane, L. Stetten, M. Seder-Colomina, J. Brest, G. Ona-  
617 Nguema, S. Bassot, C. Courbet, J. Guillevic, A. Thouvenot, O. Mathon, O. Proux, J.R.  
618 Bargar, Mononuclear U(IV) complexes and ningyoite as major uranium species in lake  
619 sediments, *Geochemical Perspectives Letters* 2 (2016) 95-105.

620  
621 [20] R. Bernier-Latmani, H. Veeramani, E.D. Vecchia, P. Junier, J.S. Lezama-Pacheco, E.I.  
622 Suvorova, J.O. Sharp, N.S. Wigginton, J.R. Bargar, Non-uraninite products of microbial  
623 U(VI) reduction, *Environ. Sci. Technol.* 44 (2010) 9456-9462.

624  
625 [21] X. Rui, M. J. Kwon, E.J. O’Loughlin, S. Dunham-Cheatham, J.B. Fein, B. Bunker, K.M.  
626 Kemner, M.I. Boyanov, Bioreduction of Hydrogen Uranyl Phosphate: Mechanisms and U(IV)  
627 Products, *Environ. Sci. Technol.* 47 (2013) 5668-5678.

628  
629 [22] D.E. Latta, K.M. Kemner, B. Mishra, M.I. Boyanov, Effects of calcium and phosphate  
630 on uranium(IV) oxidation: Comparison between nanoparticulate uraninite and amorphous  
631 UIV– phosphate, *Geochim. Cosmochim. Acta* 174 (2016) 122-142.

632  
633 [23] O. Proux, X. Biquard, E. Lahera, J.-J. Menthonnex, A. Prat, O. Ulrich, Y. Soldo, P.  
634 Trévisson, G. Kapoujvan, G. Perroux, P. Taunier, D. Grand, P. Jeantet, M. Deleglise, J.-P.  
635 Roux, J.-L. Hazemann, FAME: a new beamline for X-ray absorption investigations of very-  
636 diluted systems of environmental, material and biological interests, *Phys. Scripta* 115 (2005)  
637 970-973.

638  
639 [24] S.M. Webb, SIXpack: a graphical user interface for XAS analysis using IFEFFIT, *Phys.*  
640 *Scr. T115* (2005) 1011–1014.

641

642 [25] B. Ravel, M. Newville, ATHENA, ARTEMIS, HEPHAESTUS: data analysis for X-ray  
643 absorption spectroscopy using IFEFFIT, *J. Synchrotron Radiat.* 12 (2005) 537–541.

644  
645 [26] G. Morin, V. Noël, N. Menguy, J. Brest, B. Baptiste, M. Tharaud, G. Ona-Nguema, M.  
646 Ikogou, E. Viollier, F. Juillot, Nickel accelerates pyrite nucleation at ambient temperature,  
647 *Geochemical Perspectives Letters* 5 (2017) 6-11.

648  
649 [27] D. Marshall, Ternplot: An excel spreadsheet for ternary diagrams, *Computers*  
650 *Geosciences* 22 (1996) 697–699. [https://doi.org/10.1016/0098-3004\(96\)00012-X](https://doi.org/10.1016/0098-3004(96)00012-X).

651  
652 [28] M. Seder-Colomina, G. Morin, J. Brest, G. Ona-Nguema, N. Gordien, J.-J. Pernelle, D.  
653 Banerjee, O. Mathon, G. Esposito, E.D. Van Hullebusch, Uranium(VI) Scavenging by  
654 Amorphous Iron Phosphate Encrusting *Sphaerotilus natans* Filaments, *Environ. Sci. Technol.*  
655 49 (2015) 14065–14075.

656  
657 [29] S. Kelly, K. Kemner, J. Fein, D. Fowle, M. Boyanov, B. Bunker, N. Yee, X-ray  
658 absorption fine structure determination of pH-dependent U-bacterial cell wall interactions,  
659 *Geochimica et Cosmochimica Acta* 66 (2002) 3855–3871. [https://doi.org/10.1016/S0016-](https://doi.org/10.1016/S0016-7037(02)00947-X)  
660 [7037\(02\)00947-X](https://doi.org/10.1016/S0016-7037(02)00947-X).

661  
662 [30] J.O. Sharp, J.S. Lezama-Pacheco, E.J. Schofield, P. Junier, K.-U. Ulrich, S. Chinni, H.  
663 Veeramani, C. Margot-Roquier, S.M. Webb, B.M. Tebo, D.E. Giammar, J.R. Bargar, R.  
664 Bernier- Latmani, Uranium speciation and stability after reductive immobilization in aquifer  
665 sediments, *Geochim. Cosmochim. Acta.* 75 (2011) 6497–6510.

666  
667 [31] H.S. Moon, J. Komlos, P.R. Jaffé, Uranium reoxidation in previously bioreduced sed-  
668 iment by dissolved oxygen and nitrate, *Environ. Sci. Technol.* 41 (2007) 4587–4592.

669  
670 [32] S.D. Kelly, Uranium chemistry in soils and sediments, in: B. Singh, M. Grafe (Eds.),  
671 *Dev. Psychiatry*, 2010, Vol. 34, pp. 411.

672  
673 [33] A. Singh, J.G. Catalano, K.U. Ulrich, D.E. Giammar, Molecular-Scale Structure of  
674 Uranium(VI) Immobilized with Goethite and Phosphate, *Environ. Sci. Technol.* 46 (2012)  
675 6594–6603.

676

677 [34] T. Muto, R. Meyrowitz, A. M. Pommer, T. Murano, Ningyoite, a new uranous phosphate  
678 mineral from Japan, *Am. Mineral.* 44 (1959) 633–650.

679 [35] S.V. Krivovichev, J. Plášil, *Mineralogy and Crystallography of Uranium*, in: P.C. Burns,  
680 G.E. Sigmon (Eds.), *Uranium: From Cradle to Grave*, Mineralogical Association of Canada  
681 Short Courses, 2013, pp 15-119

682  
683 [36] G.T.W. Law, A. Geissler, I.T. Burke, F.R. Livens, J.R. Lloyd, J.M. McBeth, K. Morris,  
684 Uranium Redox Cycling in Sediment and Biomineral Systems, *Geomicrobiology Journal* 28  
685 (2011) 497–506, <https://doi.org/10.1080/01490451.2010.512033>.

686  
687 [37] B. Gu, J. Chen, Enhanced microbial reduction of Cr(VI) and U(VI) by different natural  
688 organic matter fractions, *Geochimica et Cosmochimica Acta* 67 (2003) 3575–3582,  
689 [https://doi.org/10.1016/S0016-7037\(03\)00162-5](https://doi.org/10.1016/S0016-7037(03)00162-5).

690  
691 [38] D.R. Lovley, E.J.P. Phillips, Y.A. Gorby, E.R. Landa, Microbial reduction of uranium,  
692 *Nature* 350 (1991) 413–416.

693  
694 [39] Y. Suzuki, S.D. Kelly, K.A. Kemner, J.F. Banfield, Direct microbial reduction and  
695 subsequent preservation of uranium in natural near-surface sediment, *Appl. Environ.*  
696 *Microbiol.* 71 (2005) 1790–1797.

697  
698 [40] T. Behrends, P. Van Cappellen, Competition between enzymatic and abiotic reduction of  
699 uranium(VI) under iron reducing conditions. *Chem Geol* 220 (2005) 315–327.

700  
701 [41] M.I. Boyanov, K.E. Fletcher, M.J. Kwon, X. Rui, E.J. O’Loughlin, F.E. Löffler, K.M.  
702 Kemner, Solution and microbial controls on the formation of reduced U(IV) phases, *Environ.*  
703 *Sci. Technol.* 45 (2011) 8336–8344.

704  
705 [42] D.S. Alessi, J.S. Lezama-Pacheco, J.E. Stubbs, M. Janousch, J.R. Bargar, P. Persson, R.  
706 Bernier-Latmani, The product of microbial uranium reduction includes multiple species with  
707 U(IV)–phosphate coordination, *Geochim. Cosmochim. Acta* 131 (2014) 115-127.

708  
709 [43] T.V. Khijniak, A.I. Slobodkin, V. Coker, J.C. Renshaw, F.R. Livens, E.A. Bonch-  
710 Osmolovskaya, N.-K. Birkeland, N.N. Medvedeva-Lyalikova, J. R. Lloyd, Reduction of  
711 Uranium(VI) Phosphate during Growth of the Thermophilic Bacterium *Thermo-*

712 *terrabacterium ferrireducens*, Applied and environmental microbiology 71 (2005)  
713 6423–6426.

714  
715 [44] M. Grybos, M. Davranche, G. Gruau, P. Petitjean, M. Pédrot, Increasing pH drives  
716 organic matter solubilization from wetland soils under reducing conditions, Geoderma 154  
717 (2009) 13-19.

718  
719 [45] P. Warwick, N. Evans, A. Hall, G. Walker, E. Steigleder, Stability constants of U(VI)  
720 and U(IV)-humic acid complexes, J. Radioanal. Nucl. Chem. 266 (2005) 179-190.

721  
722 [46] S.E. Bone, J.J. Dynes, J. Cliff, J.R. Bargar, Uranium(IV) adsorption by natural organic  
723 matter in anoxic sediments, Proc. Natl. Acad. Sci. U. S. A. 114 (2017) 711–716.

724  
725 [47] M.Grybos, M.Davranche, G. Gruau, P. Petitjean, Is trace metal release in wetland soils  
726 controlled by organic matter mobility or Fe-oxyhydroxides reduction?, Journal of Colloid and  
727 Interface Science 314, (2007) 490–501. <https://doi.org/10.1016/j.jcis.2007.04.062>.

728