



HAL
open science

Tuning electronic structure in layered hybrid perovskites with organic spacer substitution

Joshua Leveillee, Claudine Katan, Jacky Even, Dibyajyoti Ghosh, Wanyi Nie,
Aditya D Mohite, Sergei Tretiak, André Schleife, Amanda J Neukirch

► **To cite this version:**

Joshua Leveillee, Claudine Katan, Jacky Even, Dibyajyoti Ghosh, Wanyi Nie, et al.. Tuning electronic structure in layered hybrid perovskites with organic spacer substitution. *Nano Letters*, 2019, 19 (12), pp.8732-8740. 10.1021/acs.nanolett.9b03427 . hal-02343044

HAL Id: hal-02343044

<https://hal.science/hal-02343044v1>

Submitted on 12 Feb 2021

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Tuning electronic structure in layered hybrid perovskites with organic spacer substitution

Joshua Leveillee,^{*,†,‡} Claudine Katan,[¶] Jacky Even,[§] Dibyajyoti Ghosh,[‡] Wanyi Nie,[‡] Aditya D. Mohite,^{||} Sergei Tretiak,[‡] André Schleife,^{†,⊥,#} and Amanda J. Neukirch^{*,‡}

[†]*Department of Materials Science and Engineering, University of Illinois at Urbana-Champaign, Urbana, IL 61801, USA*

[‡]*Los Alamos National Laboratory, Los Alamos, NM 87545, USA*

[¶]*Univ Rennes, ENSCR, INSA Rennes, CNRS, ISCR - UMR 6226, F-35000 Rennes, France*

[§]*Univ Rennes, INSA Rennes, CNRS, Institut FOTON - UMR 6082, F-35000 Rennes, France*

^{||}*Department of Chemical and Biomolecular Engineering, Rice University, Houston, TX 77005, US*

[⊥]*Materials Research Laboratory, University of Illinois at Urbana-Champaign, Urbana, IL 61801, USA*

[#]*National Center for Supercomputing Applications, University of Illinois at Urbana-Champaign, Urbana, IL 61801, USA*

E-mail: leveill2@illinois.edu; ajneukirch@lanl.gov

Abstract

Two-dimensional layered halide organic perovskites (LHOPs) are promising candidates for many optoelectronic applications due to their interesting tunable properties.

They provide a unique opportunity to control energy and charge dynamics via the independent tunability of the energy levels within the perovskite and the organic spacer for various optoelectronic applications. In the perovskite layer alone, one can replace the Pb (Sn), the halide (X=I,Br,Cl), the organic component, and the number of layers between the organic spacer layers. In addition, there are many possibilities for organic spacer layers between the perovskite layers, making it difficult for experimental methods to comprehensively explore such extensive combinatorial space. Of particular technological interest is alignment of electronic levels between the perovskite layer and the organic spacer layer, leading to desired transfer of energy or charge carriers between perovskite and organic components. For example, as band edge absorption is almost entirely attributed to the perovskite layer, one way to demonstrate energy transfer is to observe triplet emission from organic spacers. State-of-the-art computational chemistry tools can be used to predict the properties of many stoichiometries in search for LHOPs that have the most promising electronic structure features. In this first principles study, we survey a group of π -conjugated organic spacer candidates for use in triplet light emitting LHOPs. Utilizing density functional theory (DFT) and time-dependent DFT, we calculate the first singlet (S_1) and triplet (T_1) excitation energy in the ground-state geometry and the first triplet excitation energy in the excited triplet state relaxed geometry (T_1^*). By comparing these energies to the known lowest exciton energy level of $\text{Pb}_n\text{X}_{3n+1}$ perovskite layers (X=I, Br, Cl), organic spacer and perovskite layer pairings can be identified for possible transfer of Wannier excitons from the inorganic perovskite lattice to spin-triplet Frenkel excitons located on the organic cation. We successfully identify ten organic spacer candidates for possible pairing with perovskite layers of specific halide composition to achieve triplet light emission across the visible energy range. Molecular dynamics simulations predict that finite temperatures and perovskite environment have little influence on the average excitation energies of the two common organic spacers naphthylethylammonium (NEA) and phenylethylammonium (PEA). We find significant thermal broadening up to 0.5 eV of the optical excitation energies appearing due to finite temperature effects. The findings

herein provide insights into alignment of electronic levels of conjugated organic spacer with perovskite layer.

Keywords: **Layered hybrid organic-inorganic perovskites, halide perovskite, organic spacer substitution, triplet light emission, exciton energy alignment, first principles simulations.**

Two-dimensional (2D) layered hybrid metal-halide-organic perovskites (LHOPs) have come to the forefront of optoelectronic materials research due to their diverse and promising properties and applications. Most LHOPs assume a Ruddelston-Popper crystalline structure of $A_2(\text{CH}_3\text{NH}_3)_{n-1}\text{M}_n\text{X}_{3n+1}$, where A^{1+} is a large organic spacer, M^{2+} is a metal cation (usually Pb or Sn), X^{1-} is a halide anion, and n is the number of octahedra along the stacking axis in a single perovskite layer in stacking direction.¹ Stoichiometric design flexibility arises in the choice of perovskite layer thickness, metal-halide combination, and organic cation spacer.² Successes of LHOP materials designs include highly efficient solar cells,³ tuned light emitting diodes (LEDs),⁴⁻⁷ and possible spintronics.⁸ Recent studies have discovered that electronic and optical properties of the perovskite layer are heavily influenced by the choice of the organic spacer. The geometries of organic spacers strongly affect octahedra tilting in the perovskite layer and determine the band gap of the bulk material² as well as energies of surface states.⁹ Furthermore, introducing more complex conjugated organic spacers has diversified the light-emitting properties of LHOPs,¹⁰ leading to increased interest in band alignment and charge transfer between the perovskite and organic layers.¹¹⁻¹⁴ It is clear that the choice of organic cation is critical to the optical and electronic functionality of LHOP devices.

The electronic structures of 2D LHOPs are broadly tunable by the versatile alignment between the electronic levels of the perovskite and the organic spacer components. Depending on the degree of conjugation and polarity of the spacer moiety, one can precisely align desired electronic states including the perovskite levels and the singlet and triplet manifolds of the organic spacer layer. Knowledge of this provides materials scientists with a rich tool to control

energy and charge dynamics via independent tunability of energy levels in the perovskite and organic components. Different applications require distinct electronic dynamics and transport, from separating p and n charge carrier channels between the perovskite and spacer layers, to storing excitons in the triplet states of the organic spacer layer and then re-using them as in thermally activated delayed fluorescence (TADF) emitters,¹⁵ to attaining white-like broad-band dual emission characteristics.¹⁶ Of particular technological interest for many applications is engineering of the directed energy transfer between the perovskite layer and the organic spacer layer. Typically in these LHOP systems, the perovskite layer is the dominant component of band edge absorption.^{2,14,17-19} Therefore, a way to demonstrate energy transfer would be to observe triplet emission from the organic spacer layer, which itself could be harnessed for several applications.

Markedly, over the past two decades, select LHOPs have demonstrated the ability to induce energy transfer from perovskite layer exciton states to low energy spin-triplet exciton states in the organic layer.²⁰⁻²² The fundamental differences of the singlet-triplet exciton energy split between the perovskite and organic layers is critical to inducing such energy transfer processes. Within the perovskite layer, strong spin-orbit coupling and a high dielectric constant cause the differences between the singlet and triplet exchange integrals to be small and thus cause a little difference between singlet and triplet states. Consequently, these weakly bound and delocalized S_1 and T_1 Wannier-Mott excitons are hardly distinguishable in optical emission spectra.^{19,23,24} In contrast, the collinear spin nature and small dielectric screening in the organic layer leads to a large difference in singlet and triplet exchange integrals, resulting in a drastic lowering of the T_1 excitation energy compared to the S_1 excitation energy.²⁵ The excitons in the organic layer are strongly bound Frenkel excitons. When the energy of the weakly bound exciton in the perovskite layer aligns with the energy T_1 of the strongly bound exciton in the organic layer, charge transfer from the perovskite to the organic layer may occur.²⁰⁻²² After transfer, the T_1 excitation energy in the organic relaxes to a lower T_1^* energy due to enhanced short-range atomic deformation, thus

reaching optimal triplet molecular geometry. Phosphorescent photo-emission is subsequently measured at the T_1^* excitation energy.²⁶ The choice and alignment of the exciton energy in the perovskite and T_1 in the organic, and the photo-emission energy T_1^* from the organic, are key design parameters leading to different options for LHOP photo-emission materials. Subsequent light emission due to recombination from this triplet exciton is referred to as triplet emission. Electronic structure methods provide a tool to better understand how to choose the various components within LHOPs to achieve desired material properties and optical response.

In this letter, we detail our first-principles survey of organic spacer triplet emission candidates for LHOPs. Each candidate consists of a conjugated chromophore attached to an ethyl-ammonia tail ($C_2H_4-NH_2$). The lowest spin-triplet excitation levels are calculated on the ground state (S_0) geometries to obtain the triplet energy T_1 and the excited triplet geometry to obtain the triplet energy T_1^* in the Δ SCF framework. The Δ SCF approach has been shown to be a robust framework for predicting the lowest energy triplet state.²⁷⁻³⁰ We also calculate the singlet and triplet excitation energies of several organic spacer candidates in clustered geometries and in the perovskite environment, and predict similar excitation energies compared to the vacuum-relaxed molecular geometries. By means of molecular dynamics simulations for a few LHOP systems, we verify that our initial conclusions hold also under dynamic conditions but with significant thermally induced broadening of the excitation energies up to 0.5 eV. We are able to successfully identify promising triplet emission candidates for pairing with known layered perovskite exciton levels to obtain possible triplet emission devices. We consider pairings with perovskite layers of varying layer thickness and halide stoichiometry, and the results herein provide the fundamental concepts for crafting materials with the desired response to various stimuli (light, field, charge injection).

Four classes of organic cation spacer molecules have been explored in this work (see molecular structures in Fig. 1). The choices are motivated by organic spacers successfully used in LHOPs experimentally^{12,18,20-22} and span diverse forms of conjugation (polycarbo-

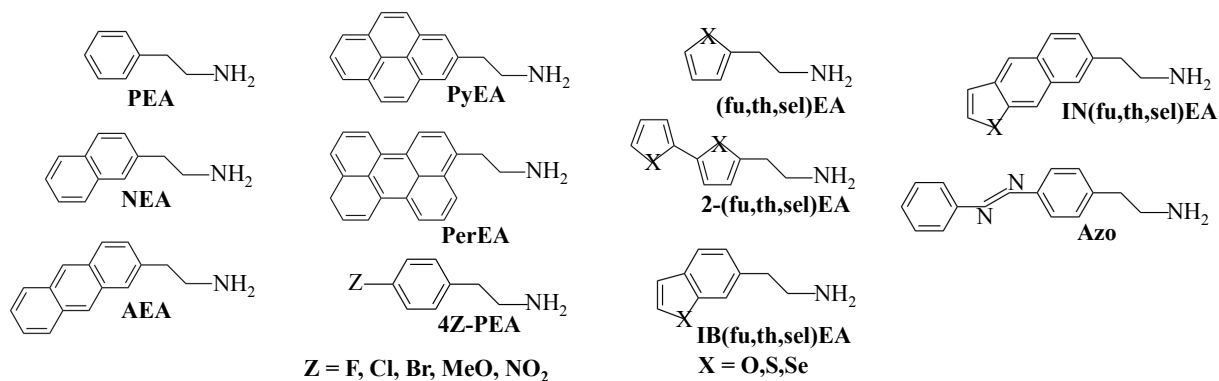


Figure 1: Organic spacer candidates considered in this work. Abbreviations include furan (fu), thiophene (th), selenophene (sel), isobenzol (IB), isonaphthyl (IN), 2-2'-bis (2), and Azobenzene ethylammonia (Azo).

cyclic, heterocyclic, and electron rich/poor functionalization): First, five polycarbocyclic compounds with ethylammonium (EA) tails are studied. Amongst them are the basic phenylethylammonium (PEA), used extensively in LHOP synthesis, and the two-ring naphthylethylammonium (NEA), a different isomer of which was investigated by Ema *et al.* and has shown promising application as a triplet emitter.²² Three-ring anthrylethylammonium (AEA) were explored more seldom, but have promising low singlet and triplet energy states.²⁰ Four-ring pyrene-ethylammonium (PyEA) has only briefly been studied for its triplet emission properties,²¹ whereas five-ring perylene-ethylammonium (PerEA) was not explored for its photo-emission properties as spacers in LHOP systems.¹² The second class is functionalized PEA with the goal of adding electron-rich or -poor groups and additional linearly conjugated compounds. For electron-rich additions, halides F, Cl, and Br and NO₂ are added to the para position, opposite from the EA bonding site, on the phenyl group. Electron-poor methoxy (MeO) is attached to the same site. The third class are the heterocyclic compounds with carbon and chalcogen based rings. These include the pentacyclic furan-, thiophene-, and selenophene-ethylammonium (XEA, X=fu,th,sel) and their 2,2' linked pentacyclic counterparts (2-XEA, X=fu,th,sel). Further, there are the fused ring compounds isobenzyl(furan,thiophene,selenophene)-ethylammonium (IBXEA, X=fu,th,sel) and isonaphthyl(furan,thiophene,selenophene)-ethylammonium (INXEA, X=fu,th,sel). As

a fourth class, we included azobenzene-ethylammonium to examine the effects of extended conjugation featuring the N–N double bond. In this work we refer to the atomic positions of the organic cation layers without the inorganic perovskite lattice, taken from a single unit cell of the experimental LHOP structures, as organic cluster geometries. PEA and NEA,¹⁸ and PyOPA¹² clusters are investigated and initialized from atomic structures determined experimentally.

The predicted T_1 and T_1^* energies of organic spacer candidates calculated by the Δ SCF approach (see section Methods) span the visible and near-UV energy range (see Table S1 in the supplemental information). Our calculations are in agreement with available experimental data. For the most simple conjugated spacer, PEA, we compute T_1 of 4.43 eV and T_1^* of 3.77 eV using B3LYP. Since benzene is identical to the phenyl group in PEA we use it for comparison here, in the absence of PEA experimental data. This is further justified by the good agreement of Δ SCF results for benzene, $T_1^*=3.88$ eV, and PEA as well as the good agreement between Δ SCF and phosphorescence emission data of 3.40–3.60 eV.³¹ The lowest phosphorescence emission energy T_1^* is predicted for NEA and PyEA, within 0.2 eV of the experimental NEA-PbBr₄²² isomer and PyEA-PbI₄.²¹ Furthermore, the Δ SCF model also predicts the T_1 excitation energies of fuEA and thEA to within 100 meV of experimental results.³² Given such good agreement, the Δ SCF model can be used to make predictions on a number of other organic spacer candidates. We note that while this agreement is encouraging, these data were computed for vacuum-optimized isolated molecules. It is, thus, important to check what influence geometric and thermal effects have on these results, which will be discussed later.

We further analyze spatial localization of triplet excitations using spin densities (defined as the difference between up and down spin densities) and complementary natural orbitals corresponding to unpaired electrons forming the triplet state. These orbitals are calculated as eigenfunctions of an open-shell ground-state density matrix of the molecule using the NBO (Natural Bond Orbital) package³³ built in the Gaussian suite. We further refer the reader

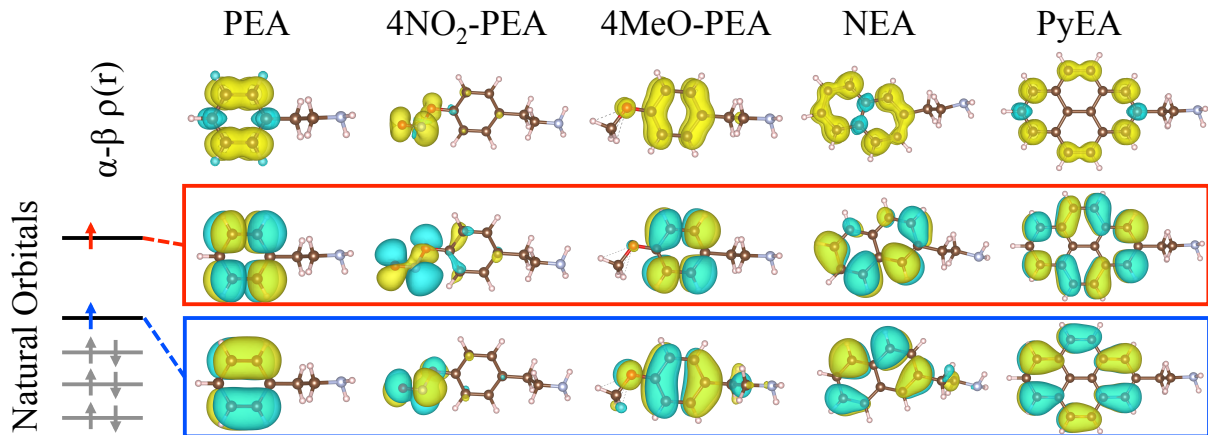


Figure 2: Spin density of the triplet state (top row) and natural orbitals of the open shell states in selected organic spacer candidates.

to a review³⁴ discussing various flavors of the natural orbital analysis in molecular systems. Both spin densities and natural orbitals are visualized for selected organic spacer candidates in Fig. 2. It is found for all candidate spacers that the spin density is dispersed around the π -conjugated bonds of organic rings or on the non-tail functional group. None of the cases studied here exhibits significant spin density localized on the EA tail. The natural orbitals for the two spins of the triplet configuration are generally localized in the same fashion. This indicates that orbitals delocalized over the conjugated rings or localized on the non-tail functional group are the main participants in triplet excitations, while the non-conjugated tails provide little contributions.

Unlike other conjugated ring spacer candidates considered, the triplet excitation on 4NO₂-PEA is localized near the bonds of the NO₂ as opposed to the conjugated ring. This removal of the excitation from the PEA ring causes 4NO₂-PEA to exhibit a much lower T_1 energy of 2.29 eV compared to the other functionalized PEA candidates (4.05–4.59 eV). This predicted T_1 energy of 4NO₂-PEA is in good agreement with the experimental T_1 energy of NO₂ of 2.35 eV.³⁵ The origin of the lowered excitation energy is the conjugated bond series of O–N=O. The localization of the spin density, predicted at the density functional level of theory, is validated for this case and confirmed by performing second order Møller-

Plesset perturbation theory (MP2) calculations.^{36,37} 4NO₂-PEA highlights the possibility of conjugated compounds as alternative to ring structures for lowered triplet energy level.

After calculating the excitation energies of organic organic spacer candidates in the gas phase, we move on to predict the excitation energies of clusters of organic spacers without the perovskite layer, in the geometries determined by X-ray diffraction experiments.^{12,18} The geometries and orientation of individual organic molecules in LHOP crystals differ from that of isolated molecules due to interactions with each other and the perovskite layers. Du *et al.* examined PEA, PMA, NEA, and NMA organic spacers in Pb(I,Br,Cl)₄ systems and found varying behavior of tail geometries and penetration depth of the ammonium group into the perovskite octahedra layer.¹⁸ One uniform characteristic of the conjugated ring organic spacers in the $\langle 100 \rangle$ oriented class of LHOPs is the curling of the tail relative to the geometries of isolated molecules in vacuum. For example, the primary carbon of the naphthyl ring (C₁) in NEA of the EA in the vacuum relaxed geometry is 3.8 Å away from the nitrogen of the EA tail. In the perovskite environment, the tail curls back resulting in the nitrogen moving closer to the carbon rings with a C₁-N distance of 3.02 Å in NEA-PbI₄.¹⁸ Smaller conjugated groups like PEA tend to arrange parallel to the perovskite layer while larger chromophores begin to tilt to a perpendicular position. Direct face-to-face π -stacking is not present in the organic layers considered in this study. Instead, the conjugated rings tend to arrange edge-to-face, a formation that is favorable to template the perovskite network.¹² Given these fundamental differences between perovskite and isolated geometries, it is important to understand how excitation spectra of organic spacer candidates change in the perovskite environment.

To this end, we examine how excitation energies of clusters of organic molecules PEA, NEA, and PyOPA in their experimental geometries differ compared to their isolated counterparts calculated in the gas-phase. The clusters are taken as isolated layers from the unit cell, are non-periodic, and consist of 4, 8, and 4 molecular species of PEA, NEA, and PyOPA, respectively. Fig. S1 (c) and (d) in the supplemental information show the simulated

atomic structures for the NEA and PyOPA cluster, respectively. These organic compounds are chosen by available experimental atomic structures of layered perovskites in which they are the spacer. While PEA is unlikely to be a triplet emission candidate, NEA and PyEA (or PyOPA) show great potential due to their low T_1 energies. PEA and NEA layer geometries of four and eight molecules, respectively, are taken from Ref.¹⁸ and pyrene-based organic-layer geometries with OPA tails are taken from Ref.¹² First, no relaxations are performed on C, N, and O atoms, but since H atoms are not present in the XRD data, their positions are relaxed around the experimentally known atomic positions. Below we also test the effects of tail geometry by relaxing O, N, and C atoms, while keeping the tails fixed. We also note that one H atom is omitted from each NH_3^+ group to simulate neutral layers. This does not introduce detrimental effects on the geometries, since the tails are kept fixed.

The T_1 energies calculated via the ΔSCF method using the B3LYP XC functional for the organic-spacer layer geometry without the perovskite layer, taken from experimental data of PEA ($T_1 = 4.0$ eV) and PyOPA ($T_1 = 2.35$ eV), are predicted within 0.5 eV of the vacuum relaxed molecules of PEA ($T_1 = 4.43$ eV) and PyEA ($T_1 = 2.43$ eV). In contrast, the NEA experimental organic-layer geometry ($T_1 = 2.26$ eV) is predicted to have a ≈ 0.96 eV decrease in excitation energies compared to the isolated NEA molecule relaxed in vacuum ($T_1 = 3.14$ eV). The source of this discrepancy lies with the molecular geometry differences where the experimental C–C bond lengths on the conjugated ring vary between 1.29 and 1.51 Å, whereas the isolated molecule has a smaller variation of 1.36–1.42 Å. It is worth noting that X-ray diffraction measurements do not pinpoint well the positions of light atoms such as carbon,³⁸ and therefore some variation in experimentally measured C–C bond lengths and atomic positions is to be expected. Finite temperature can induce additional variation in the molecular structure of spacers and, consequently, spacer excitation energies. Other crystallographic measurements of NEA and NMA (naphthal-methylammonium) structures have observed a C–C bond length range in the naphthal ring between 1.34–1.43 Å in NEA- CuCl_4 ³⁹ and NMA- SnI_4 .⁴⁰ This is very similar to the range found here for the isolated

structures. Nonetheless it needs to be clarified (i) how the curled tail geometry influences the conjugated rings and (ii) how the periodic potential of the perovskite and other conjugated rings affects the variance of the chromophore geometry.

To further test the effects of tail geometry on the predicted T_1 energies of organic candidates, a second set of calculations are performed where tail geometries are kept fixed at experimental geometries^{12,18} while the conjugated rings are optimized. This test serves to determine if the bent tail significantly changes the orbital nature on the conjugated ring, thus changing the excitation spectrum. The resulting ring-relaxed B3LYP Δ SCF T_1 energies for the PEA cluster ($T_1 = 4.42$ eV), NEA cluster ($T_1 = 3.07$ eV), and PyOPA cluster ($T_1 = 2.35$ eV) are predicted within 0.15 eV compared to the respective energies computed for isolated geometries. We find that tail curling alone does not strongly affect the excitation spectrum of these molecules and does not change the spin density of the molecule, which remains localized on the ring (see Fig. 2). Therefore, differences in the excitation energies calculated for isolated geometries and experimental structures, must originate from distinct conjugated ring geometries. This might be explained by the difficulties of pinpointing the precise position of light atoms, such as carbon, in XRD when heavy atoms, such as lead, are present in the system.

Using ab-initio molecular dynamics (AIMD) simulations we further explore how the excitation energies of the organic spacer embedded in the perovskite layers, at finite temperature, differ from those relaxed in a vacuum. To this end, atomic geometries for PEA-PbI₄ and NEA-PbI₄ are initialized from atomic positions determined by XRD measurements by Du *et al.*¹⁸ Prior to performing AIMD, we calculate the S_1 and T_1 energies of each unique crystalline organic PEA and NEA in these experimental geometries using TDDFT for S_1 and Δ SCF for T_1 excitation energies. We find that the calculated T_1 energy of PEA remains in a narrow range between 4.43 eV and 4.46 eV for the individual spacer molecule geometries in the unit cell taken from the experimental XRD data. In contrast, the calculated T_1 energies of the NEA molecules vary between 2.23–3.26 eV, warranting further investigation by AIMD

as discussed next.

AIMD results demonstrate that the excitation energies of PEA and NEA in the PbI_4 LHOP environment experience a broadening as a consequence of thermal fluctuations in the molecular geometries. Table S2 in the supplemental information lists the explicit results and Fig. 3 shows the distributions of S_1 and T_1 excitation energies calculated across post-equilibrium molecular conformations from AIMD. The average excitation energies for the different post-equilibration MD geometries are noticeably lower, compared to their isolated vacuum-relaxed counterparts. This results from a combination of longer average bond lengths in MD compared to XRD as well as complex torsion and bending in the NEA and PEA. The standard deviation in excitation energies lay between 0.11 and 0.20 eV, with the highest attributed to the T_1 excitation in PEA. The C–C bond lengths show high variability in PEA and NEA. For PEA, the simulated C–C variation exceeds the C–C bond length variations observed in XRD (1.39–1.41 Å). Alternatively, the variation in NEA bond length in the AIMD simulations remains lower compared to the XRD results. This explains why the T_1 energies calculated in XRD PEA conformation fall within a narrow range whereas geometries sampled from AIMD show a higher variation. Overall, the findings herein demonstrate the importance of finite temperature dynamics on the broadening of S_1 and T_1 excitation energies of organic spacer candidates in LHOPs. As good agreement is observed between \bar{T}_1 obtained from AIMD and T_1 calculated from vacuum-relaxed simulations, the latter results can be used for first-principles materials design of LHOPs for applications requiring specific alignments of electronic features.

Thus, from our analysis of the calculated excitation energies we identify multiple organic spacer candidates with low-lying triplet states in Pb-based LHOPs with a variety of halides and number of perovskite layers (see Table 1 and Fig. 4). The first major design consideration is the relative T_1 excitation energy of the organic spacer to the lowest excitation energy of the perovskite (with number of perovskite octahedra layers $n = 1, 2, 3$, etc.) E_x . As previously discussed, the lowest triplet and singlet excitation energies in the perovskite layer

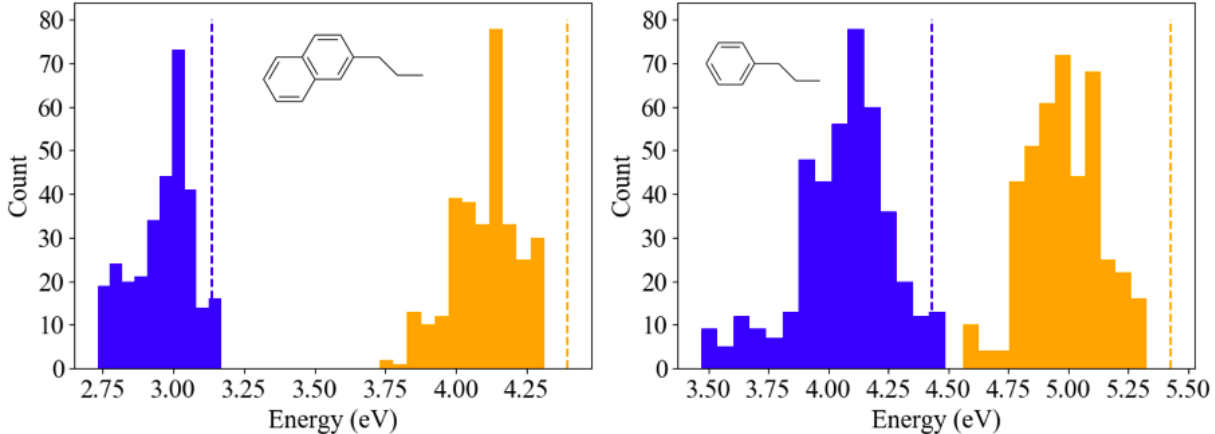


Figure 3: Histograms of Δ SCF T_1 (blue) and TDDFT S_1 (orange) energies calculated on MD molecular conformations of NEA in NEA-PbI₄ (left) and PEA in PEA-PbI₄ (right). The vertical dashed line indicates the vacuum-relaxed result for S_1 (orange) and T_1 (blue).

are nearly indistinguishable,^{23,24} and thus we generate energy alignments with the lowest optical excitation peak energy in the perovskite E_x from experimental data. When T_1 is less than but close to E_x , strongly-bound exciton energy transfer can be initiated from the perovskite to the organic. Once this transfer occurs, the organic spacer relaxes in the excited state, leading to a new T_1^* emission energy which is the lowest energy phosphorescence peak. In the following, energy pairings will be given in $E_x:T_1:T_1^*$ notation. Starting with single-layer PbCl₄, the two optimal pairings are thEA-PbCl₄ and selEA-PbCl₄ with energy alignments of 3.7:3.69:3.02 and 3.7:3.57:2.93 eV, respectively. These LHOPs could produce a violet or indigo phosphorescence spectrum with a lowest emission wave length of ≈ 400 nm. As the functionalized PEA T_1 tends to be higher than 3.7 eV, thEA and selEA seem to be the only viable options for pairing with $n = 1$ PbCl₄ found in our work. The lowest exciton energy level of single-layer $n = 1$ PbCl₄ are referenced from Du *et al.*¹⁸

For $n = 1$ PbBr₄ perovskite layers, we also identify two promising organic spacer pairings: The optimal pairing is NEA-PbBr₄, with an energy alignment of 3.20:3.14:2.67 eV. This first-principles prediction is supported by experimental evidence of highly efficient energy transfer from PbBr₄ to the triplet state of a NEA isomer by Ema *et al.*²² They observe this process in a variety of naphthyl-based PbI₄ systems, encouraging further studies of these. Another

possible pairing is 2-fuEA-PbBr₄, which has an energy alignment of 3.20:2.91:2.40 eV and could exhibit triplet emission at 516 nm, corresponding to a green light. The lowest exciton energy level of single-layer $n = 1$ PbBr₄ are referenced from Du *et al.*¹⁸

The highest number of promising pairings of organic spacers for triplet emission occurs with $n = 1$ PbI₄ and $n = 2$ Pb₂I₇. The most obvious pairing for $n = 1$ PbI₄ is PyEA-PbI₄ with an energy alignment of 2.5:2.43:2.10 eV. This material would show triplet emission of yellow/orange light, which has been experimentally verified by Braun *et al.*²¹ Another possibility is AEA-PbI₄ for an alignment of 2.5:2.17:1.75 eV, leading to red light triplet emission. Interestingly, AEA spacers in layered perovskites have rarely been explored for their light emitting properties, despite promising exciton energy level alignment. A similar energy alignment of 2.5:2.17:1.77 eV and potential red light triplet emission can be achieved by the pairing IBselEA-PbI₄. Finally, 4NO₂-PEA is predicted to have an alignment with $n = 1$ PbI₄ of 2.5:2.29:0.68 eV, indicating potential infrared triplet emission in 4NO₂-PEA-PbI₄. As previously discussed, a low T_1^* in 4NO₂-PEA is attributed to the localization of the triplet excitation on the NO₂ functional group (Fig. 2).

Two-layer $n = 2$ Pb₂I₇ is also predicted to have spacer candidates for deep red triplet emission: The first two are PerEA and AzEA, with alignments of 2.15:1.80:1.50 eV and 2.15:2.08:1.75 eV, respectively. The pairing with PerEA should be given careful attention, given the recent inclusion of PerEA in layered perovskite by Passarelli *et al.*¹² Finally, Azo can be paired with $n = 2$ Pb₂I₇ for an alignment of 2:15:1.95:1.25 eV, giving yet another spacer for possible infrared triplet emission. The lowest exciton energy levels of one, two, three, four, and infinite (3D) layered Pb_nI_{4n-1}, as discussed here and shown in Fig. 4, are references from Blancon *et al.*⁴¹

In addition to these, multiple of the considered compounds have non-ideal alignment with the known perovskite layer absorption edges, but should not be entirely discounted. Compounds with T_1 energies between 2.5 and 2.8 eV are mismatched with $n = 1$ PbI₄ and are unlikely to induce energy transfer from $n = 1$ PbBr₄, including 2-(th,sel)EA and

Table 1: Suggested pairings of organic spacers with perovskite layers of varying halide and layer compositions. The energy alignment shows the perovskite exciton level E_x , the organic triplet level in the ground state geometry T_1 (charge transfer energy), and the organic triplet level in the excited state geometry T_1^* (triplet emission energy), i.e. $E_x:T_1:T_1^*$ (in eV).

Perovskite	Org. Spacer	Energy Alignment (eV)
PbI ₄ ($n=1$)	PyEA	2.50 : 2.43 : 2.10
	AEA	2.50 : 2.17 : 1.75
	IBselEA	2.50 : 2.17 : 1.77
	4NO ₂ -PEA	2.50 : 2.29 : 0.68
Pb ₂ I ₇ ($n=2$)	PerEA	2.15 : 1.80 : 1.50
	Azo	2.15 : 1.95 : 1.25
PbBr ₄ ($n=1$)	NEA	3.20 : 3.14 : 2.67
	2fuEA	3.20 : 3.14 : 2.40
PbCl ₄ ($n=1$)	thEA	3.70 : 3.69 : 3.02
	selEA	3.70 : 3.57 : 2.93

IN(fu,th,sel)EA. However, little is currently known about the perovskite exciton levels of $n = 2$ Pb₂Br₇ and it is possible that alignment is better in this case. Given experimental evidence from $n = 2$ Pb₂I₇,⁶ it is likely that $n = 2$ Pb₂Br₇ would demonstrate a E_x level between 3.2 eV for $n = 1$ PbBr₄ and 2.4 eV for $n = \infty$ PbBr₃. Thus, some of the aforementioned non-ideal compounds for pairing with PbI₄ or PbBr₄ could provide an ideal pairing with $n = 2$ Pb₂Br₇. Experimental observation of E_x from $n = 2$ Pb₂Br₇ would be required to make a definite claim. In contrast, the single ring compounds and {Br, Cl, F, MeO}-functionalized PEA all have high T_1 compared to E_x of $n = 1$ PbCl₄, and, therefore, are highly unlikely to induce excitation transfer.

A brief discussion on possible exciton transfer mechanisms, though not simulated in this work, is useful. It is important to note that the S_1 and T_1 states in both 3D three-dimensional HOPs and two-dimensional LHOPs are nearly indistinguishable, differing by up to only 25 meV.^{23,42} This can be attributed to the strong spin-orbit coupling, where $J = S + L$ is a good quantum number for the Wannier-type exciton generated in the perovskite upon optical absorption or charge injection. The influence of spin-orbit coupling is much weaker in the organic layer, where the enhanced Coulomb exchange strongly splits the S_1 and T_1 energies, resulting in strongly bound Frenkel excitons. Thus, the energy transfer involves

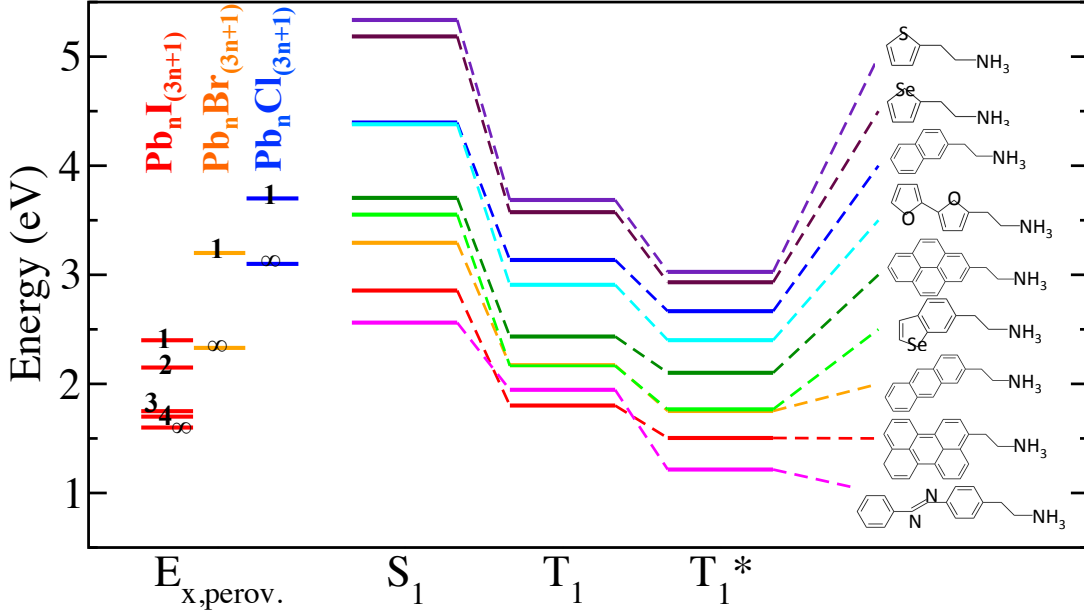


Figure 4: Alignment between perovskite E_x exciton levels, organic triplet T_1 excitation energies, and organic T_1^* excited state geometry emission energies for a select subset of examined organic spacers. Perovskite exciton energy levels are referenced from previous experimental data.^{18,41}

a transfer of the mixed singlet and triplet Wannier-Mott exciton in the strongly spin-orbit coupled perovskite layer to the triplet Frenkel exciton in the organic layer. The first, and most likely, possible charge transfer mechanism is the Dexter mechanism,⁴³ hypothesized by Ema *et al.* as the source of triplet emission in an isomer of $\text{NEA}_2\text{-PbBr}_4$.²² Dexter charge transfer requires a significant orbital overlap between the organic and perovskite species and is highly dependent on individual atomic structures of LHOP systems. Ema *et al.* noted along these lines that the length of the organic cation tail strongly influences the triplet exciton transfer rate, suggesting more strongly that the Dexter mechanism is at play.²² Another possible mechanism would be a charge transfer through a series of excimer and exciplex intermediate states,⁴⁴ though no studies have been performed on this concerning LHOPs to the knowledge of the authors. Finally, Förster resonance energy transfer (FRET) could be induced by dipole-dipole interaction, leading to excitation in the organic spacer layer via non-radiative de-excitation in the perovskite layer.⁴⁵ Notably, while FRET is less sensitive

to the distance between donor and acceptor species compared, for example, to the Dexter mechanism, vanishing transition dipole moments of T_1 states of organic molecules effectively rule out this scenario. However, the spin-orbit coupling in organics may be enhanced by the presence of perovskites. This may potentially enable a weak FRET channel and improve triplet emission compared to that in the isolated molecules. Further investigation on the energy and charge transfer mechanisms between the perovskite layer and the organic spacer will thus be an important next step to designing functional LHOPs.

In conclusion, we use electronic-structure calculations to predict multiple viable pairings of organic spacer candidates with single perovskite layers, to facilitate design of LHOPs with desired energy level alignments. These pairings may achieve potential triplet emission across the visible spectral range, offering potential solutions for tuned light emission applications. We carefully verified that our predictions remain consistent even when tail geometries are altered and organic candidates are placed into their clustered LHOP geometries. Finally, finite temperature dynamics of organic spacers in the perovskite crystal, as simulated by ab-initio molecular dynamics, result in a broadening of excitation energies slightly below the results obtained for isolated geometries. This thermal broadening of the excitation energies provides an important consideration for future theoretical and experimental work in designing LHOP optical emission devices. Altogether our computational predictions demonstrate that electronic structure of LHOPs paired with conjugated organic spacers is extremely tunable, which suggests broad variability and control over dynamics of energy and charge carriers. In particular, light emitting applications involving LHOPs is a promising field of research and should be further investigated to discover potential next generation, wave length tuned, highly efficient light sources.

Computational Methods

First-principles simulations for the singlet excitation, triplet excitation, and the energy gap between the highest occupied and lowest unoccupied molecular orbitals (HOMO-LUMO gap) based on Δ SCF and time-dependent density functional theory (TDDFT) are performed in this work.

Ground state closed- and open-shell DFT calculations of spin singlet and triplet excitations are conducted using the Gaussian code.⁴⁶ These use occupation number and spin orientation constraints to enforce the total spin S to 0 or 1, respectively. As such, the pure spin states are eigenstates of the Kohn-Sham Hamiltonian. Charge densities and single-particle wave functions are calculated in a Gaussian 6-31G* basis set on pruned ultra-fine real space grids with 99,590 real space integration points around the atoms. For heavy Br ions, a LANL2DZ Gaussian basis set is tested and found to be comparable in accuracy to 6-31G* (see Table S9 in the supplemental material). For neutral molecules in vacuum, structural relaxations are performed in the singlet configuration with the B3LYP exchange-correlation functional.⁴⁷

Triplet excitations are calculated using the Δ SCF approach, due to its predictive power for the S_0 geometry triplet and T_1^* geometry triplet energies. In addition, Δ SCF avoids triplet instability errors that occur in TDDFT:^{29,30,48,49} In TDDFT, the difference between singlet and triplet can become larger than the HOMO-LUMO gap,³⁰ forcing the triplet excitation energies to become excessively small or even negative in some cases. Sears *et al.* point to the difference of the exchange contribution between singlet and triplet within a given exchange-correlation functional as the source of this triplet instability. The Δ SCF approach avoids this issue, since it relies on the difference between total Kohn-Sham energies of singlet and triplet states. Lowest-energy singlet excitations are calculated within the TDDFT framework in the frequency domain, from the solution of the Casida equation.⁵⁰ TDDFT has been shown to provide reliable singlet excitation energies in organic molecules for charge transfer and optical excitations.⁵¹ While T_1 excitation energies are also calculated by TDDFT (see Tables

S3 through S12 in the supplemental information), they underestimate the excitation energies due to the aforementioned triplet instability⁴⁸ and are thus not discussed in the main text. Excitation energies are separately calculated using HSE,⁵² B3LYP,⁴⁷ and CAM-B3LYP⁵³ exchange correlation functionals within Δ SCF and TDDFT. It is found that the introduction of solvent effects of water using the polarization continuum model⁵⁴ does not change the predicted S_1 and T_1 energies significantly (see Table S13 of the supplemental material). CAM-B3LYP is used to compare the excitation energies of charged (NH_3^+ -terminated) and uncharged (NH_2 -terminated) species, due to its improved handling of long-range exchange under charged conditions. The results are comparable within tens of meV (see Tables S8 through S12 in the SI). As tail charge is inconsequential to the chromophore excitation energies and since B3LYP better handles the triplet instability compared to CAM-B3LYP and HSE in extensively conjugated systems, B3LYP results are presented in this letter. However, for natural bonding orbital analysis and to predict changes in natural atomic charges when going from tails with NH_2 to NH_3^+ , we use CAM-B3LYP due to its ability to better handle excitations in the charged state. All these simulations are done on geometries relaxed in the vacuum and detailed results are tabulated in Tables S3 through S12 of the supplemental material.

To understand how finite temperature in the perovskite environment effects molecular geometry and singlet and triplet excitation energies, we run ab-initio molecular dynamics simulations. These are performed using the VASP code,⁵⁵⁻⁵⁸ the PAW method, and the PBE exchange-correlation functional⁵⁹ with Grimme-D2 Van der Waals corrections⁶⁰ in the NVT ensemble with a Nose-Hoover thermostat at 300 K. The simulations are initialized with experimental geometries¹⁸ and after thermal equilibrium is reached, molecular geometries at each time step are used for simulating the singlet and triplet excitation energies using the Gaussian code.

Acknowledgement

The work at Los Alamos National Laboratory (LANL) was supported by the LANL Directed Research and Development Funds (LDRD). This work was conducted, in part, at the Center for Nonlinear Studies and the Center for Integrated Nanotechnologies, U.S. Department of Energy (DOE), Office of Basic Energy Sciences user facility at LANL. This research used resources provided by the LANL Institutional Computing (IC) Program. LANL is operated by Triad National Security, LLC, for the National Nuclear Security Administration of the U.S. Department of Energy (Contract No. 89233218NCA000001). Work performed at UIUC was supported by the National Science Foundation under Grant No. DMR-1555153. This research is partially supported by the Blue Waters sustained-petascale computing project, which is supported by the National Science Foundation (awards OCI-0725070 and ACI-1238993) and the state of Illinois. Blue Waters is a joint effort of the University of Illinois at Urbana-Champaign and its National Center for Super Computing Applications. J.E. acknowledges the Institut Universitaire de France. The work in France was supported by Agence Nationale pour la Recherche (TRANSHYPERO projects).

References

- (1) Ruddlesden, S. N.; Popper, P. New compounds of the K_2NiF_4 type. *Acta Crystallogr* **1957**, *10*, 538–539.
- (2) Pedesseau, L.; Saponi, D.; Traore, B.; Robles, R.; Fang, H.-H.; Loi, M. A.; Tsai, H.; Nie, W.; Blancon, J.-C.; et al., Advances and Promises of Layered Halide Hybrid Perovskite Semiconductors. *ACS Nano* **2016**, *10*, 9776–9786.
- (3) Cho, K. T.; Grancini, G.; Lee, Y.; Oveisi, E.; Ryu, J.; Almora, O.; Tschumi, M.; Schouwink, P. A.; Seo, G.; Heo, S.; et al., Selective growth of layered perovskites for stable and efficient photovoltaics. *Energy Environ Sci* **2018**, *11*, 952–959.

- (4) Dohner, E. R.; Jaffe, A.; Bradshaw, L. R.; Karunadasa, H. I. Intrinsic White-Light Emission from Layered Hybrid Perovskites. *J Am Chem Soc* **2014**, *136*, 13154–13157.
- (5) Smith, M. D.; Karunadasa, H. I. White-Light Emission from Layered Halide Perovskites. *Acc Chem Res* **2018**, *51*, 619–627.
- (6) Mao, L.; Ke, W.; Pedesseau, L.; Wu, Y.; Katan, C.; Even, J.; Wasielewski, M. R.; Stoumpos, C. C.; Kanatzidis, M. G. Hybrid Dion–Jacobson 2D Lead Iodide Perovskites. *J Am Chem Soc* **2018**, *140*, 3775–3783.
- (7) Tsai, H.; Nie, W.; Blancon, J.-C.; Stoumpos, C. C.; Soe, C. M. M.; Yoo, J.; Crochet, J.; Tretiak, S.; Even, J.; Sadhanala, A.; et al., Stable Light-Emitting Diodes Using Phase-Pure Ruddlesden–Popper Layered Perovskites. *Adv Mater* **2018**, *30*, 1704217.
- (8) Wang, J.; Zhang, C.; Liu, H.; McLaughlin, R.; Zhai, Y.; Vardeny, S. R.; Liu, X.; McGill, S.; Semenov, D.; Guo, H.; et al., Spin-optoelectronic devices based on hybrid organic-inorganic trihalide perovskites. *Nat Commun* **2019**, *10*, 129.
- (9) Kepenekian, M.; Traore, B.; Blancon, J.-C.; Pedesseau, L.; Tsai, H.; Nie, W.; Stoumpos, C. C.; Kanatzidis, M. G.; Even, J.; Mohite, A. D.; et al., Concept of Lattice Mismatch and Emergence of Surface States in Two-dimensional Hybrid Perovskite Quantum Wells. *Nano Lett* **2018**, *18*, 5603–5609.
- (10) Smith, M. D.; Connor, B. A.; Karunadasa, H. I. Tuning the Luminescence of Layered Halide Perovskites. *Chemical Rev* **2019**, *119*, 3104–3139.
- (11) Mitzi, D. B.; Chondroudis, K.; Kagan, C. R. Design, Structure, and Optical Properties of Organic-Inorganic Perovskites Containing an Oligothiophene Chromophore. *Inorg Chem* **1999**, *38*, 6246–6256.
- (12) Passarelli, J. V.; Fairfield, D. J.; Sather, N. A.; Hendricks, M. P.; Sai, H.; Stern, C. L.; Stupp, S. I. Enhanced Out-of-Plane Conductivity and Photovoltaic Performance in n

- = 1 Layered Perovskites through Organic Cation Design. *J Am Chem Soc* **2018**, *140*, 7313–7323.
- (13) Traore, B.; Pedesseau, L.; Assam, L.; Che, X.; Blancon, J.-C.; Tsai, H.; Nie, W.; Stoumpos, C. C.; Kanatzidis, M. G.; Tretiak, S.; et al., Composite Nature of Layered Hybrid Perovskites: Assessment on Quantum and Dielectric Confinements and Band Alignment. *ACS Nano* **2018**, *12*, 3321–3332.
- (14) Katan, C.; Mercier, N.; Even, J. Quantum and Dielectric Confinement Effects in Lower-Dimensional Hybrid Perovskite Semiconductors. *Chem Rev* **2019**, *119*, 3140–3192.
- (15) Uoyama, H.; Goushi, K.; Shizu, K.; Nomura, H.; Adachi, C. Highly efficient organic light-emitting diodes from delayed fluorescence. *Nature* **2012**, *492*, 234–238.
- (16) Sheng, C. X.; Singh, S.; Gambetta, A.; Drori, T.; Tong, M.; Tretiak, S.; Vardeny, Z. V. Ultrafast intersystem-crossing in platinum containing π -conjugated polymers with tunable spin-orbit coupling. *Sci Rep* **2013**, *3*, 2653.
- (17) Gebhardt, J.; Kim, Y.; Rappe, A. M. Influence of the Dimensionality and Organic Cation on Crystal and Electronic Structure of Organometallic Halide Perovskites. *J Phys Chem C* **2017**, *121*, 6569–6574.
- (18) Du, K.-z.; Tu, Q.; Zhang, X.; Han, Q.; Liu, J.; Zauscher, S.; Mitzi, D. B. Two-Dimensional Lead(II) Halide-Based Hybrid Perovskites Templated by Acene Alkylamines: Crystal Structures, Optical Properties, and Piezoelectricity. *Inorg Chem* **2017**, *56*, 9291–9302.
- (19) Leveille, J.; Katan, C.; Zhou, L.; Mohite, A. D.; Even, J.; Tretiak, S.; Schleife, A.; Neukirch, A. J. Influence of π -conjugated cations and halogen substitution on the optoelectronic and excitonic properties of layered hybrid perovskites. *Phys Rev Materials* **2018**, *2*, 105406.

- (20) Braun, M.; Tuffentsammer, W.; Wachtel, H.; Wolf, H. Tailoring of energy levels in lead chloride based layered perovskites and energy transfer between the organic and inorganic planes. *Chem Phys Lett* **1999**, *303*, 157 – 164.
- (21) Braun, M.; Tuffentsammer, W.; Wachtel, H.; Wolf, H. Pyrene as emitting chromophore in organic–inorganic lead halide-based layered perovskites with different halides. *Chem Phys Lett* **1999**, *307*, 373 – 378.
- (22) Ema, K.; Inomata, M.; Kato, Y.; Kunugita, H.; Era, M. Nearly Perfect Triplet-Triplet Energy Transfer from Wannier Excitons to Naphthalene in Organic-Inorganic Hybrid Quantum-Well Materials. *Phys Rev Lett* **2008**, *100*, 257401.
- (23) Tamarat, P.; Bodnarchuk, M. I.; Trebbia, J.-B.; Erni, R.; Kovalenko, M. V.; Even, J.; Lounis, B. The ground exciton state of formamidinium lead bromide perovskite nanocrystals is a singlet dark state. *Nat Mater* **2019**, *18*, 717–724.
- (24) Thouin, F.; Valverde-Chávez, D. A.; Quarti, C.; Cortecchia, D.; Bargigia, I.; Beljonne, D.; Petrozza, A.; Silva, C.; Srimath Kandada, A. R. Phonon coherences reveal the polaronic character of excitons in two-dimensional lead halide perovskites. *Nat Mater* **2019**, *18*, 349–356.
- (25) Tanaka, K.; Takahashi, T.; Kondo, T.; Umeda, K.; Ema, K.; Umebayashi, T.; Asai, K.; Uchida, K.; Miura, N. Electronic and Excitonic Structures of Inorganic–Organic Perovskite-Type Quantum-Well Crystal (C₄H₉NH₃)₂PbBr₄. *Jp J Appl Phys* **2005**, *44*, 5923–5932.
- (26) Jüstel, T.; Möller, S.; Winkler, H.; Adam, W. *Ullmann’s Encyclopedia of Industrial Chemistry*; American Cancer Society, 2012.
- (27) Frank, I.; Hutter, J.; Marx, D.; Parrinello, M. Molecular dynamics in low-spin excited states. *J Chem Phys* **1998**, 4060–4069.

- (28) Grimm, S.; Nonnenberg, C.; Frank, I. Restricted open-shell Kohn–Sham theory for π - π^* transitions. I. Polyenes, cyanines, and protonated imines. *J Chem Phys* **2003**, *119*, 11574–11584.
- (29) Gavnholt, J.; Olsen, T.; Englund, M.; Schiøtz, J. Δ self-consistent field method to obtain potential energy surfaces of excited molecules on surfaces. *Phys Rev B* **2008**, *78*, 075441.
- (30) Sears, J. S.; Koerzdoerfer, T.; Zhang, C.-R.; Brédas, J.-L. Communication: Orbital instabilities and triplet states from time-dependent density functional theory and long-range corrected functionals. *J Chem Phys* **2011**, *135*, 151103.
- (31) Ishikawa, H.; Noyes, W. A. The Triplet State of Benzene. *J Am Chem Soc* **1962**, *84*, 1502–1503.
- (32) Flicker, W. M.; Mosher, O. A.; Kuppermann, A. Triplet states of furan, thiophene, and pyrrole. *Chem Phys Lett* **1976**, *38*, 489 – 492.
- (33) Glendening, E. D.; Badenhop, J. K.; Reed, A. E.; Carpenter, J. E.; Bohmann, J. A.; Morales, C. M.; Landis, C. R.; Weinhold, F. NBO 6.0. 2012.
- (34) Reed, A. E.; Curtiss, L. A.; Weinhold, F. Intermolecular interactions from a natural bond orbital, donor-acceptor viewpoint. *Chem Rev* **1988**, *88*, 899–926.
- (35) Hochstrasser, R. M.; Marchetti, A. P. Electronic, Vibrational, and Zeeman Spectra of Triplet NO₂. *J Chem Phys* **1969**, *50*, 1727–1736.
- (36) Møller, C.; Plesset, M. S. Note on an Approximation Treatment for Many-Electron Systems. *Phys Rev* **1934**, *46*, 618–622.
- (37) Head-Gordon, M.; Pople, J. A.; Frisch, M. J. MP2 energy evaluation by direct methods. *Chem Phys Lett* **1988**, *153*, 503 – 506.

- (38) Cullity, B.; Stock, S. *Elements of X-ray Diffraction, Third Edition*; Prentice-Hall, 2001.
- (39) Mande, H. M.; Ghalsasi, P. S.; Navamoney, A. Synthesis, structural and spectroscopic characterization of the thermochromic compounds A_2CuCl_4 : [(Naphthyl ethylammonium) $_2CuCl_4$]. *Polyhedron* **2015**, *91*, 141 – 149.
- (40) Lemmerer, A. Phase transitions and structural motifs of inorganic-organic lead halide hybrids. Ph.D. thesis, University of the Witwatersrand, Johannesburg, 2008.
- (41) Blancon, J.-C.; Tsai, H.; Nie, W.; Stoumpos, C. C.; Pedesseau, L.; Katan, C.; Kepenekian, M.; Soe, C. M. M.; Appavoo, K.; Sfeir, M. Y.; et al., Extremely efficient internal exciton dissociation through edge states in layered 2D perovskites. *Science* **2017**, *355*, 1288–1292.
- (42) Ema, K.; Umeda, K.; Toda, M.; Yajima, C.; Arai, Y.; Kunugita, H.; Wolverson, D.; Davies, J. J. Huge exchange energy and fine structure of excitons in an organic-inorganic quantum well material. *Phys Rev B* **2006**, *73*, 241310.
- (43) Dexter, D. L. A Theory of Sensitized Luminescence in Solids. *J Chem Phys* **1953**, *21*, 836–850.
- (44) Yip, W. T.; Levy, D. H. Excimer/Exciplex Formation in van der Waals Dimers of Aromatic Molecules. *J Phys Chem* **1996**, *100*, 11539–11545.
- (45) Förster, T. Zwischenmolekulare Energiewanderung und Fluoreszenz. *Ann Phys-Berlin* **1948**, *437*, 55–75.
- (46) Frisch, M. J.; Trucks, G. W.; Schlegel, H. B.; Scuseria, G. E.; Robb, M. A.; Cheeseman, J. R.; Scalmani, G.; Barone, V.; Petersson, G. A.; Nakatsuji, H.; et al., Gaussian~16 Revision B.01. 2016; Gaussian Inc. Wallingford CT.
- (47) Becke, A. D. Density-functional thermochemistry. III. The role of exact exchange. *J Chem Phys* **1993**, *98*, 5648–5652.

- (48) Peach, M. J. G.; Williamson, M. J.; Tozer, D. J. Influence of Triplet Instabilities in TDDFT. *J Chem Theory Comput* **2011**, *7*, 3578–3585.
- (49) Tozer, D. J.; Handy, N. C. On the determination of excitation energies using density functional theory. *Phys Chem Chem Phys* **2000**, *2*, 2117–2121.
- (50) Casida, M.; Ipatov, A.; Cordova, F. In *Time-Dependent Density Functional Theory*; Marques, M. A., Ullrich, C. A., Nogueira, F., Rubio, A., Burke, K., Gross, E. K. U., Eds.; Springer Berlin Heidelberg: Berlin, Heidelberg, 2006; pp 243–257.
- (51) Guido, C. A.; Knecht, S.; Kongsted, J.; Mennucci, B. Benchmarking Time-Dependent Density Functional Theory for Excited State Geometries of Organic Molecules in Gas-Phase and in Solution. *J Chemical Theory Comput* **2013**, *9*, 2209–2220.
- (52) Heyd, J.; Scuseria, G. E.; Ernzerhof, M. Hybrid functionals based on a screened Coulomb potential. *J Chem Phys* **2003**, *118*, 8207–8215.
- (53) Yanai, T.; Tew, D. P.; Handy, N. C. A new hybrid exchange–correlation functional using the Coulomb-attenuating method (CAM-B3LYP). *Chem Phys Lett* **2004**, *393*, 51 – 57.
- (54) Tomasi, J.; Mennucci, B.; Cammi, R. Quantum Mechanical Continuum Solvation Models. *Chem Revi* **2005**, *105*, 2999–3094.
- (55) Kresse, G.; Hafner, J. Ab Initio Molecular Dynamics for Liquid Metals. *Phys Rev B* **1993**, *47*, 558–561.
- (56) Kresse, G.; Furthmüller, J. Efficiency of Ab-Initio Total Energy Calculations for Metals and Semiconductors Using a Plane-Wave Basis Set. *Comp Mater Sci* **1996**, *6*, 15–50.
- (57) Kresse, G.; Furthmüller, J. Efficient Iterative Schemes for Ab Initio Total-Energy Calculations Using a Plane-Wave Basis Set. *Phys Rev B* **1996**, *54*, 11169–11186.

- (58) Kresse, G.; Joubert, D. From Ultrasoft Pseudopotentials to the Projector Augmented-Wave Method. *Phys Rev B* **1999**, *59*, 1758–1775.
- (59) Perdew, J. P.; Burke, K.; Ernzerhof, M. Generalized Gradient Approximation Made Simple. *Phys Rev Lett* **1996**, *77*, 3865–3868.
- (60) Grimme, S. Semiempirical GGA-type density functional constructed with a long-range dispersion correction. *J Comput Chem* **2006**, *27*, 1787–1799.

Supporting Information Available

Explicit calculated data for singlet S_1 and triplet T_1 and triplet emission T_1^* energies are tabulated in the supporting information for all investigated compounds along with a supplemental figure showing the simulated organic clusters of NEA and PyOPA.

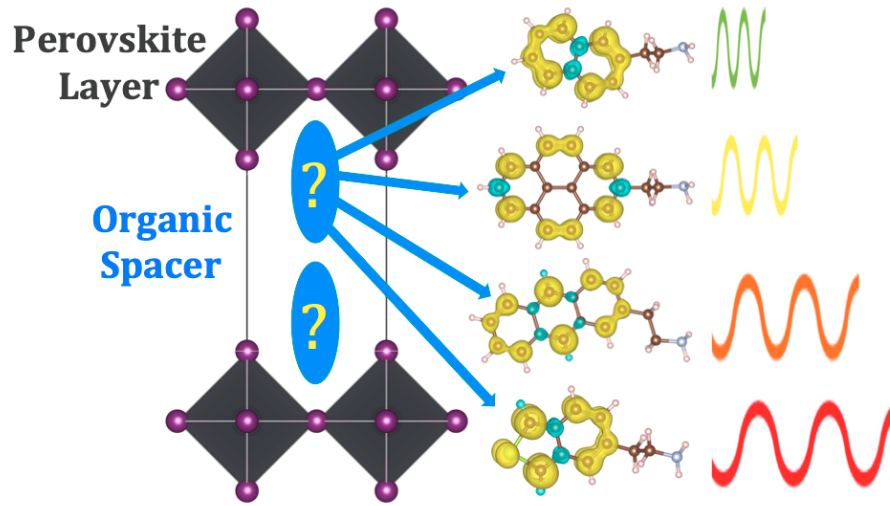


Figure 1: Table of contents graphic