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Effect of Anti-Symmetric Anchoring Groups on Electronic Transport in Hybrid Metal/Molecule/Graphene Single Molecule Junctions

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Abstract

A combined experimental and theoretical study on molecular junctions with asymmetry in both the electrode type and asymmetry in the anchoring group type is presented here. A scanning tunneling microscope is used to create the “anti-symmetric” Au-S-(CH₂)_n-COOH-graphene molecular junctions and determine their conductance. These measurements are combined with electron transport calculations based on Density Functional Theory (DFT) to analyze the electrical conductance and its length attenuation factor from a series of junctions of different molecular length (n). Our results show an unexpected trend with a rather high conductance and a smaller attenuation factor for the Au-S-(CH₂)_n-COOH-graphene configuration compared to the equivalent junction with the “symmetrical” COOH contacting using the HOOC-(CH₂)_n-COOH series. Due to the effect of the graphene electrode, the attenuation factor is also smaller than the one obtained for Au/Au electrodes. These results are interpreted through the relative molecule/electrode couplings and molecular level alignments as determined with DFT. We show that in an anti-symmetric junction, the electrical current flows through the less resistive conductance channel, similarly to what is observed in the macroscopic regime.

Introduction

Aviram and Ratner first proposed the creative idea of using molecules as rectifiers by connecting molecules within a sandwich structure.¹ Since then, molecular electronics has been put forward as a notion for extending the physical scale limits of conventional silicon integrated circuits. In terms of technical breakthrough in the measurement of electrical current through “molecular circuits” a key breakthrough was made in 1988 by Aviram *et al.*, who deployed a scanning tunneling microscope (STM) to observe the *I-V* characteristics of an “anti-symmetrical” semiquinone containing molecular bridges (by “anti-symmetrical” we mean here molecular bridges which do not have a center of symmetry, either as a result of their intrinsic molecular symmetry or as a result of different contacting to the electrode contacts at either end). This was a seminal attempt

to measure the electrical properties of single molecules using a STM to record the I - V response for molecular junctions (MJs).¹ Today, mechanically controlled break junctions (MCBJ)², scanning tunneling microscopy-break junctions (STM-BJ)³, conductive probe atomic scanning force microscopy (CP-AFM)⁴ and STM $I(s)$ ⁵ (I = current, s = distance) have all been developed as effective methods to characterize the essential properties of MJs. Examples of the use of these techniques to study the electrical characteristics of single molecule can be found in review.⁶⁻¹²

So far, well-studied molecular junction systems most often have symmetric configurations, for instance using molecules terminated with similar anchoring groups at each extremity of the junction, including dithiol¹³⁻¹⁵, diamine^{14, 16} and dicarboxylic acid^{14, 17} etc. Likewise, the majority of single-molecule electrical studies have used the same metal contact types on either respective side, with gold being the most common⁶⁻¹², although other electrodes have been used such as Cu^{18, 19} and Ag¹⁸ etc. There have also been a selection of studies where different contact electrodes have been used for example gold as one contact, with the second contact being graphite or graphene²⁰ or semiconductors such as gallium arsenide²¹ or silicon.²² Nevertheless, the majority of single molecule electrical characterizations have been carried out with both symmetrical molecules and identical metal contact types at either end of the molecular bridge. Such symmetrical MJ configurations have been welcome since they are generally most readily constructed particularly with gold contacts and have formed the testbed for most fundamental studies of charge flow through single molecule junctions.⁶⁻¹² However, anti-symmetrical configurations occur opportunities for new ways of generating electrical functionality as such rectification, switching, molecular transistor behavior and charge storage. Indeed, some of the founding studies in molecular electronics demonstrated rectification through the use of directionally orientated and non-symmetric molecular monolayers, formed for example by Langmuir-Blodgett methods.^{23, 24} A number of studies²⁵ have attempted to describe how introducing asymmetry into single molecule electrical junctions can impact on the

current-voltage or electrical conductance response. For example, Martin *et al.*²⁶ found that a series of molecules (SH-(CH₂)_n-COOH) with anti-symmetric anchoring groups exhibits a lower conductance than their comparable symmetric molecules. On the other hand, Wang *et al.*²⁷ presented an investigation of 1,4-disubstituted benzene molecular junctions and they analyzed molecular rectification and the impact arising from anti-symmetric anchoring groups. Diez-Perez *et al.*²⁸ have achieved current rectification at the single molecule level by controlling the orientation of a non-symmetric diblock dipyrimidinyl diphenyl molecule bound between a pair of gold electrode contacts which responded as a single molecule diode. Other examples of the use of anti-symmetric MJs to create the future single-molecule information processing devices (rectification, switching and negative differential resistance) can be found in review.²⁹

In our previous works, we have investigated anti-symmetric hybrid Au-molecule-graphene junctions, where the anti-symmetry arises from the different contacts at either ends of the molecular bridge. In these studies a number of different anchoring groups have been employed (dithiol^{30, 31}, diamine³², and dicarboxylic acid³³). We found that the anti-symmetric MJs with a graphene bottom electrode contact lead to a lower current attenuation factor (decay constant, β -factor) than the corresponding symmetric gold-molecule-gold junctions. These results bring in new perspectives for enhanced conductance at longer molecular length. It is also clear that the mechanism of transport properties for gold-molecule-graphene junctions is remarkably different from the gold-molecule-gold counterpart. Anti-symmetric graphene/gold contacts open new routes for controlling the electrical properties of single-molecule junctions. We now extend this work by combining two anti-symmetric elements (anti-symmetric anchoring groups and anti-symmetric electrodes) simultaneously into one molecular junction. The molecular bridges deployed in this study are polymethylene (alkane) chains with different lengths terminated on one side by a carboxylic acid anchoring group and on the other end by a thiol moiety. The electrodes are a STM gold tip and a graphene bottom electrode. The conductance of these junctions is measured using a so-called

STM $I(s)$ method, which can be considered as a variant of the *in situ* break junction method (*in situ* BJ). In this technique the STM tip is used to form molecular junctions by approaching very close to the graphene substrate, but direct physical contacting between the gold tip and graphene is avoided to prevent damage to either. In this study, we found that the conductance of Au-S-(CH₂)_n-COOH-graphene junctions decays exponentially with the number of methylene groups with an attenuation factor of 0.37 per molecular unit length, similar to the alkanedithiol molecular bridges tethered between gold and graphene electrodes. These anti-symmetric junctions with Au/graphene contacts present a higher conductance than those formed with symmetric molecular bridges with carboxylic acid anchoring groups. However, the conductance of Au-S-(CH₂)_n-COOH-graphene junctions is found to be lower than the equivalent junctions with thiol moieties on both ends of the molecular junction (*i.e.* Au-S-(CH₂)_n-SH-graphene electrodes). In this respect it is noted that the attenuation factor is driven by the low attenuation channel conductance associated with the thiol group. Density Functional Theory (DFT) calculations confirm this feature and help in interpreting the obtained attenuation factors and relative conductance values.

Experimental methods

In this experiment, gold-molecule-graphene junctions were formed using the STM- $I(s)$ method⁶ (Bruker Multimode 8 microscope) and all the conductance measurements were performed under a stable liquid (mesitylene) environment. Gold wires (99.99%) bought from Tianjing Lucheng Metal were made into STM tips by electrochemical etching as has been reported previously.³⁴ The 1 × 1 cm size graphene substrates (Graphene Supermarket, US) were used as the bottom electrode. During the conductance measurement, the graphene substrate was fixed onto the sample plate using silver gel to make the electrical contact between the graphene and the metal sample plate. The target molecules were prepared as 10 mM solution in mesitylene (99%, Aladdin) and the solution was placed in a liquid cell and first allowed to stabilize before the formation of MJs. The gold STM tip was set at an initial vertical distance (4 nm) and then toward

to the graphene substrate. The gold tip was then retracted to its initial distance, and the process of approach and retraction cycle was repeated continuously to obtain a large number of $I(s)$ curves. The tip bias in each experiment was set at + 0.3 V unless otherwise stated. Over 10000 $I(s)$ curves were collected from the conductance measurements of each MJ. The $I(s)$ curves which present big noise and no plateau were removed. The remaining $I(s)$ curves were divided into a few equal bins according to the position of the plateaus. More than 500 $I(s)$ curves from the bins with the most $I(s)$ curves were selected for data analysis. The selected current data was converted into conductance values to obtain a one-dimensional (1D) conductance histogram. A Gaussian fit was applied to the most prominent peak in the 1D histogram. The conductance data was also used to plot the two-dimensional (2D) frequency histograms to locate the most probable range for the conductance value of molecular junctions. The color of specific area in 2D histogram ranges from blue (low count) to red (high count) as an indicator of an increasing number of data point within this area. To verify the configuration of MJs, control experiments were conducted. The gold STM tips were dipped into either SH-(CH₂)₇-COOH or SH-(CH₂)₁₁-COOH containing solutions for 24 hours in order to promote close-packed surface coverage of the molecular target on the gold STM tip. This procedure will favor of the formation of Au-S-(CH₂)_n-COOH-graphene molecular junctions (the H atom is expected to be removed through the formation of Au-S, as has been observed previously²¹). The measurement was carried out punctually on the day following the 24 hour adsorption period to form the molecular film on the gold STM tips. The collected $I(s)$ curves were closely monitored and the new gold STM tip was generally changed if after a few hours no further plateaus were observed in the $I(s)$ traces.

Results and discussions

The conductance of a series of hybrid graphene/molecular junctions has been measured. Figure 1 shows typical $I(s)$ traces of molecular junctions during the opening process of the STM- $I(s)$ method. Two types of curves are presented. The black exponential decay

curve which shows no current steps indicates the absence of molecular junction formation. When a molecular junction is captured between the two electrodes, a pronounced plateau appears as can be observed in other color curves. The plateau is due to the enhanced current through the molecule, with the position of the plateaus reflecting the MJ conductance. For example, the conductance of Au-S-(CH₂)₇-COOH-Graphene junctions is around 4-6 nS, while the conductance of the shorter Au-S-(CH₂)₅-COOH-Graphene junctions is located at around 10-15 nS. To find the dominant location of conductance from a large amount of raw data, selected plateau curves were used to plot 1D and 2D histograms as shown in Figure 2a and 2b. In Figure 2a, a significant peak is observed for the Au-S-(CH₂)₇-COOH-Graphene molecular junctions, giving a peak conductance value of around 5 nS after a Gaussian curve fit is applied. The corresponding 2D histogram is presented in Figure 2b. In this 2D plot a high count feature is apparent, as seen in the red area located at around 4 to 6 nS; this is in good agreement with the peak in Figure 2a and the plateaus for this MJ (Figure 1).

Although these MJs are anti-symmetric, only a single conductance peak is observed indicating that the two different orientations, with the thiol bound to either graphene or gold, respectively, do not give discernable conductance values. Here it is noted that thiols (thiolates) bind much more strongly with gold than carboxylic acid end groups do, so from this fact it would appear likely that Au-S-(CH₂)₇-COOH-graphene junctions would be favored over Au-COOH-(CH₂)₇-S-graphene junctions. Differences in MJs formed with either thiol (thiolate) or carboxylic (carboxylate) anchoring groups to gold has been previously studied. It was found that when changing from a thiol to a carboxylic acid anchor the contact resistance increased, and this increase was correlated with the decrease in bond strength with the Au surface.¹⁴ Control experiments have also been carried out in which the self-assembly of the target molecules was done in two different ways. The comparison here was to either (a) form the SAM on the graphene substrate or (b) to form the SAM directly onto the gold STM tip. Figure 3 shows a comparison of MJ conductance using these two different self-assembly modes. For SH-

(CH₂)₇-COOH (Figure 3a) and SH-(CH₂)₁₁-COOH (Figure 3b), it is seen that the conductance value of SAM formed on the gold tip and graphene substrate are essentially similar to each other.

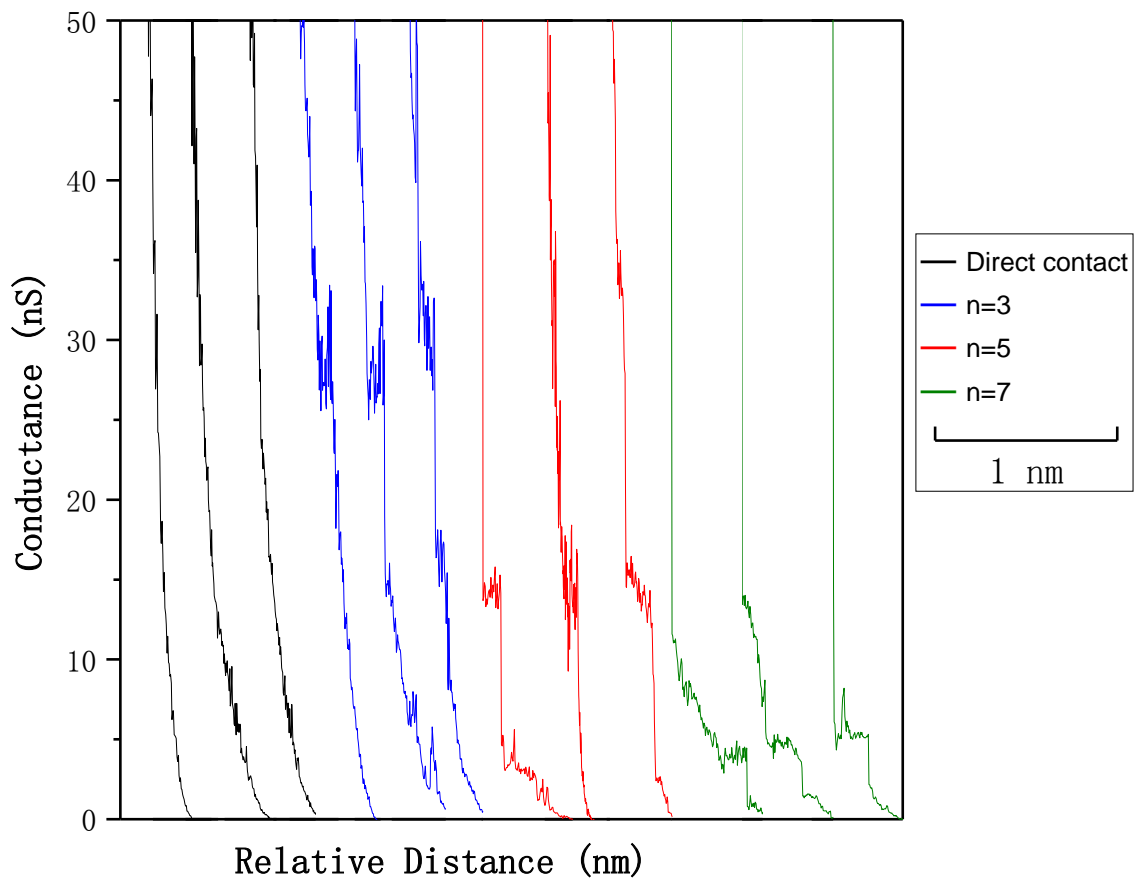


Figure 1: Typical $I(s)$ curves of $\text{SH}-(\text{CH}_2)_n-\text{COOH}$, with $n = 3$ (blue), 5 (red) and 7 (green). The black curve is recorded in absence of molecules on graphene (without forming molecular junctions).

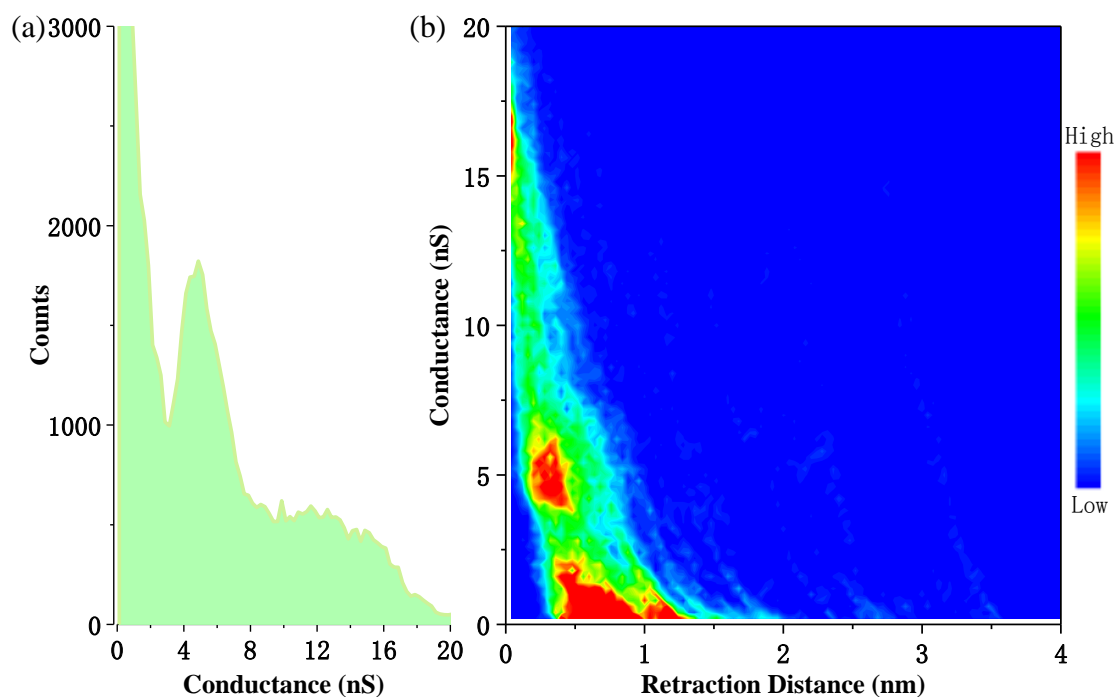


Figure 2: (a) A 1D histogram for the Au-S-(CH₂)₇-COOH-graphene hybrid junction (from selected 534 I(s) curves). (b) The corresponding 2D histogram, where the red colour represents high point count. The distance is the retraction distance from the set point conditions.

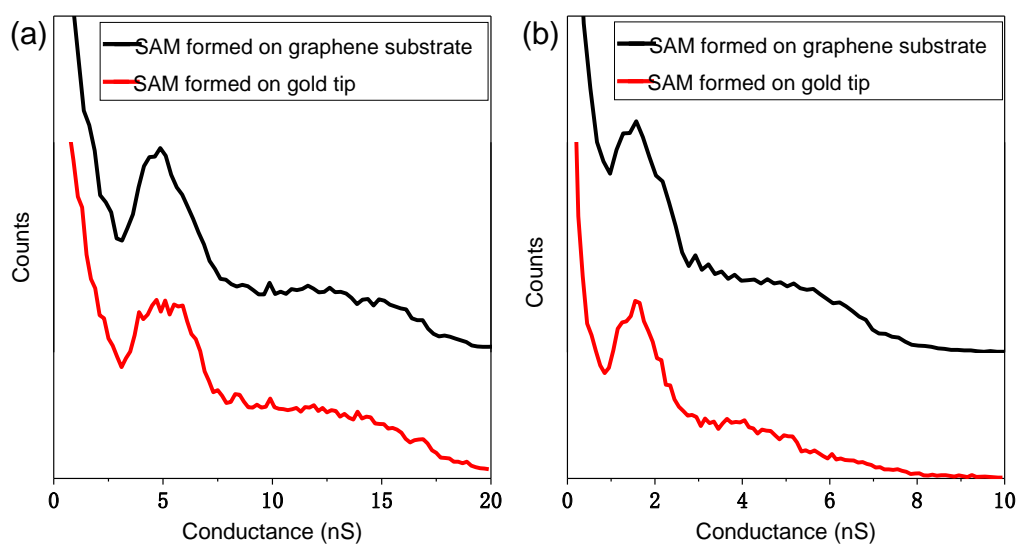


Figure 3 : Comparison of molecular conductances between SAMs formed on graphene substrate (black) and SAMs (red) formed on gold tip (a) SH-(CH₂)₇-COOH (b) SH-(CH₂)₁₁-COOH.

Figure 4 presents the five conductance histograms with the same scaling in order to compare the conductance peaks for different molecular lengths. Only a single main

peak was found for each histogram, and the conductance values of SH-(CH₂)₁-COOH (grey), SH-(CH₂)₃-COOH (blue), SH-(CH₂)₅-COOH (pink), SH-(CH₂)₇-COOH (green) and SH-(CH₂)₁₁-COOH (purple) are determined to be 63, 30, 15, 4.9, and 1.7 nS, respectively. As expected, the conductance decreases as the molecular length increases; clearly the molecular backbone lengths play a key important role in determining the magnitude of the molecular conduction.³⁵ It has been demonstrated that for a wide variety of different anchoring groups (-SH^{14, 30, 36}, -NH₂^{14, 32, 36}, -COOH^{14, 33, 37}, -I³⁸ ...) in the case of polymethylene (alkane) backbones conductance decreases sharply with molecular length, generally satisfying the well-known equation:

$$G = A \exp(-\beta_n N)$$

where G is the conductance, β_n is the attenuation factor per number of -CH₂ units, N is the number of phenyl groups and A is a constant related to the contact resistance between the molecular target and the electrode. The slope of the plot of $\ln(G)$ versus N gives the β_n value. Figure 5 shows the natural logarithmic plots of the conductance as a function of the number of methylene groups for Au-S-(CH₂)_n-SH-graphene (blue), Au-COO-(CH₂)_n-COOH-graphene (grey), Au-S-(CH₂)_n-COOH-graphene (red) and Au-S-(CH₂)_n-COOH-graphene (theory, green). The measured conductance decreases with the molecular length and the linear fit yields a tunneling decay constant of around 0.37 per methylene unit for Au-S-(CH₂)_n-COOH-graphene junctions.

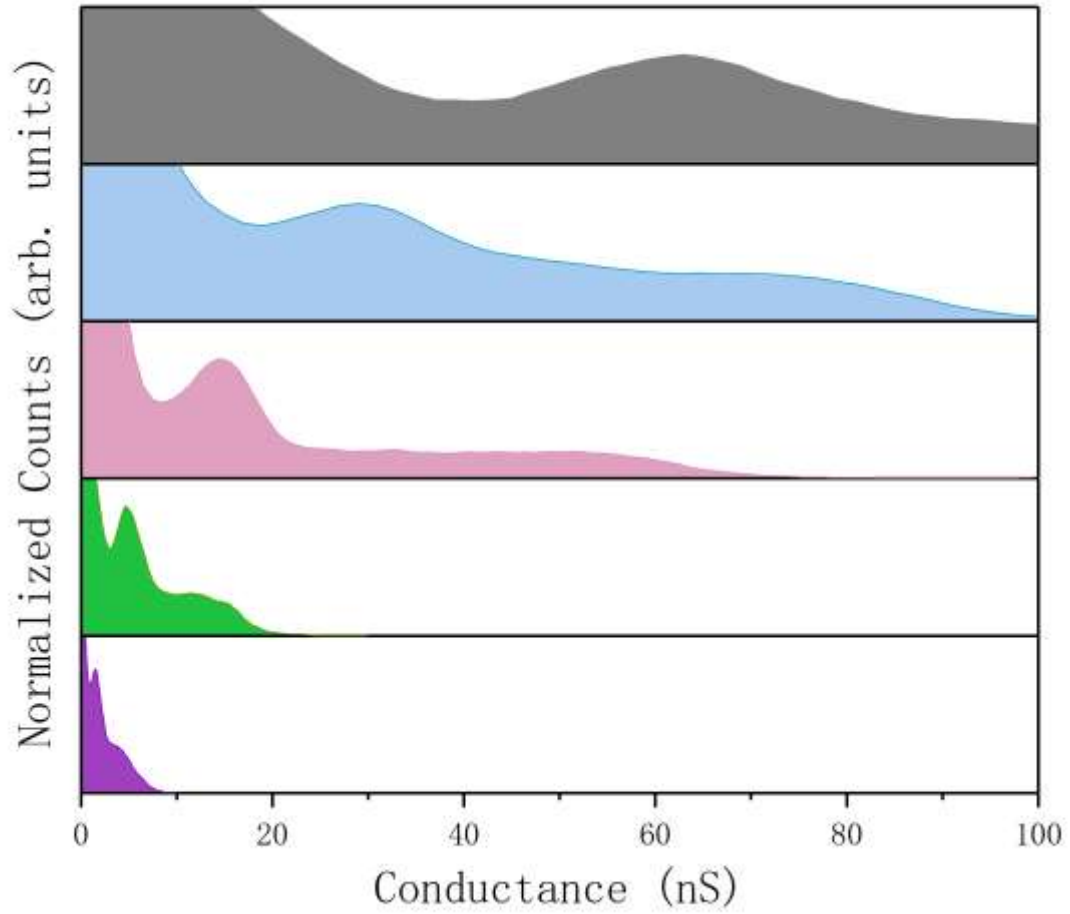


Figure 4 : Conductance histograms for $\text{SH}-(\text{CH}_2)_n-\text{COOH}$ junctions in which $n = 1$ (grey), 3 (blue), 5 (red), 7 (green), and 11 (purple).

Table 1 summarizes the conductance values and tunnelling decay constants (β_n) of the anti-symmetric MJs ($\text{SH}-(\text{CH}_2)_n-\text{COOH}$) and symmetric MJs ($\text{SH}-(\text{CH}_2)_n-\text{SH}$, $\text{COOH}-(\text{CH}_2)_n-\text{COOH}$) with Au/Au electrodes and Au/graphene electrodes. We found that the decay constant of Au/S- $(\text{CH}_2)_n$ -COOH/graphene junctions is lower than the Au/S- $(\text{CH}_2)_n$ -COOH/Au junctions.³⁷

Table 1 Conductance values and tunneling decay constant (β) for $\text{SH}-(\text{CH}_2)_n-\text{COOH}$, $\text{SH}-(\text{CH}_2)_n-\text{SH}$ and $\text{COOH}-(\text{CH}_2)_n-\text{COOH}$

Molecular junctions	Number of CH_2	Conductance (nS)		Decay constant (β_n)	
		Experiment	Theory	Experiment	Theory

	1	63	43.6		
Au/S-(CH ₂) _n -COOH	3	30	26.4		
/graphene	5	15	14.4	0.37	0.30
	7	4.9	4.81		
	11	1.7	2.62		
Au/ S-(CH ₂) _n -SH	2	98	152		
/graphene ^{30, 32}	4	53	46		
	6	24	25	0.40	0.32
	8	9	18		
	10	4	6		
Au/ COO-(CH ₂) _n -	2	15.6	38.5		
COOH /graphene ³³	3	10.3	13		
	4	5.1	5.3	0.69	0.69
	5	2.4	4		
	6	1.08	2.2		
Au/ S-(CH ₂) _n -COO	5	2.5			
/Au ²⁶	7	0.88		0.87	
Au/ S-(CH ₂) _n -S	6	28.2			
/Au ^{13, 14}	8	3.9		1.08	
	10	0.2			
Au/ COO-(CH ₂) _n -	6	2.48			
COO /Au ²⁶	8	0.80		0.78	

Similarly to our previous results^{30, 32, 33}, junctions with Au/graphene electrodes yield a lower decay constant compared to the similar MJs with symmetric gold electrodes resulting in a higher conductance for MJs with alkane backbone at longer lengths. Our experimental results show that the decay constant (0.37) of Au-S-(CH₂)_n-COOH-graphene junctions is smaller than that obtained for Au-S-(CH₂)_n-SH-graphene junctions (0.40), which is in good agreement with the results reported in those

molecules measured between Au/Au electrodes.²⁶ In Figure 5, we present the evolution of the logarithm of the conductance vs. the molecular length for the different Au/graphene MJs considered in this work, as well as theoretical results. We can now classify the different MJs according to their respective conductance values from the highest to the lowest as: Au-S-(CH₂)_n-SH-graphene junctions, Au-S-(CH₂)_n-COOH-graphene junctions, Au-COO-(CH₂)_n-COOH-graphene junctions.

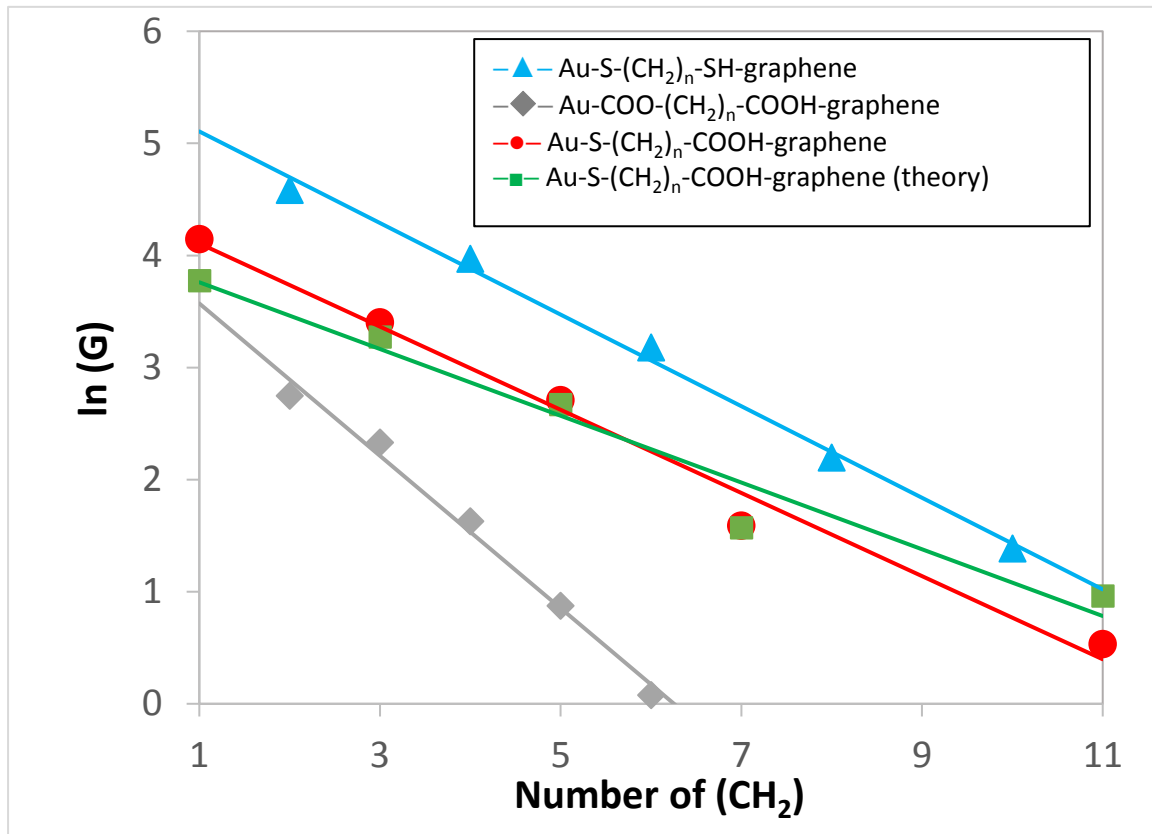


Figure 5 : Natural logarithmic plots of the conductance as a function of the number of methylene groups for Au-S-(CH₂)_n-SH-graphene (blue), Au-COO-(CH₂)_n-COOH-graphene (grey), Au-S-(CH₂)_n-COOH-graphene (red) and Au-S-(CH₂)_n-COOH-graphene (theory, green)

Extending the linear fits in Figure 5 to the origin, we can get the value of the contact resistance. The contact resistance of Au-S-(CH₂)_n-SH-graphene junctions, Au-S-(CH₂)_n-COOH-graphene junctions and Au-COO-(CH₂)_n-COOH-graphene junctions are 3.9 MΩ³⁰, 11.3 MΩ and 14.2 MΩ³³, respectively. This difference indicates that the coupling is rather weak at the carboxylate group-electrode interface.

In order to understand the underlying physical mechanism at the origin of this phenomenon, we have performed Density Functional Theory (DFT) calculations on these junctions. Following a now well-established procedure^{21,23}, we have considered Au-S-(CH₂)_n-COOH-graphene junctions, for n = 1,3,5,7,11, as represented in Figure 6 (a). After structural optimization, electronic structure has been determined as well as electronic transport properties using a Keldysh-Green formalism.²¹ The evolution of the electronic transmission is represented in Figure 6 (b).

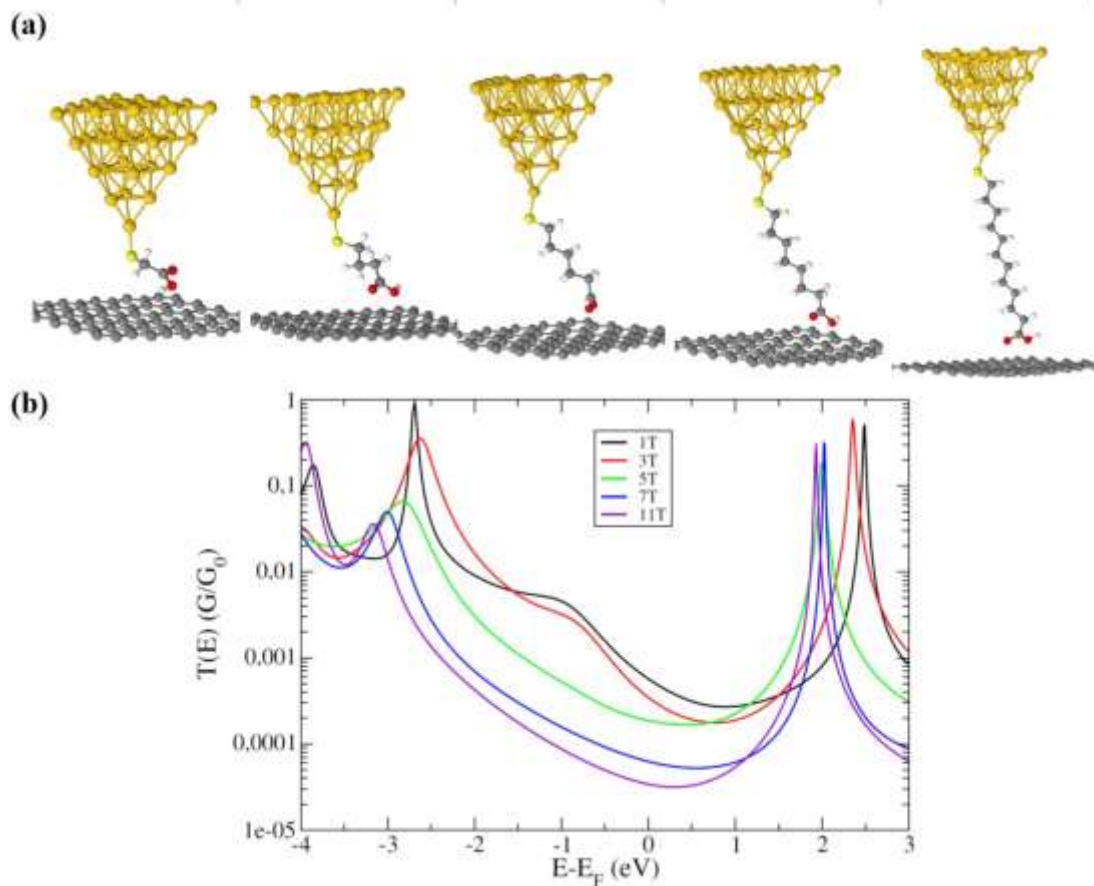


Figure 6 : (a) Atomic representation of the Au-S-(CH₂)_n-COOH-graphene junctions calculated in DFT. (b) Corresponding electronic transmissions.

From those transmissions, we can observe that the electronic transport in the Au-S-(CH₂)_n-COOH-graphene junctions lies in the non-resonant tunnelling regime, since the Fermi level is in the gap, slightly closer to the LUMO level. Hence, the transport properties are driven by the LUMO level and its position with respect to the Fermi level represents a key parameter for the determination of the attenuation factor. Indeed, from

a simple barrier tunnelling model, β_n can be expressed as: $\beta_n = 2d_0\sqrt{(2m\phi)/\hbar}$ where d_0 is the unit length between the monomers in the molecule, m is the mass of the electron, and ϕ is the barrier height, corresponding here to the energy difference between the LUMO and the Fermi level. In order to understand the obtained value for the attenuation factor, we have also calculated the electronic transmission of the Au-COO-(CH₂)₇-COOH-graphene junction. This transmission is presented in Figure 7, in comparison with the transmission of the Au-S-(CH₂)₇-COOH-graphene junction. As we can observe, the two transmission curves are very similar in shape, with a gap of about 5 eV and an electronic conductance slightly dominated in both cases by the LUMO level. Although the shapes and gap of these two transmission curves are similar, the conductance at the Fermi level clearly differs.

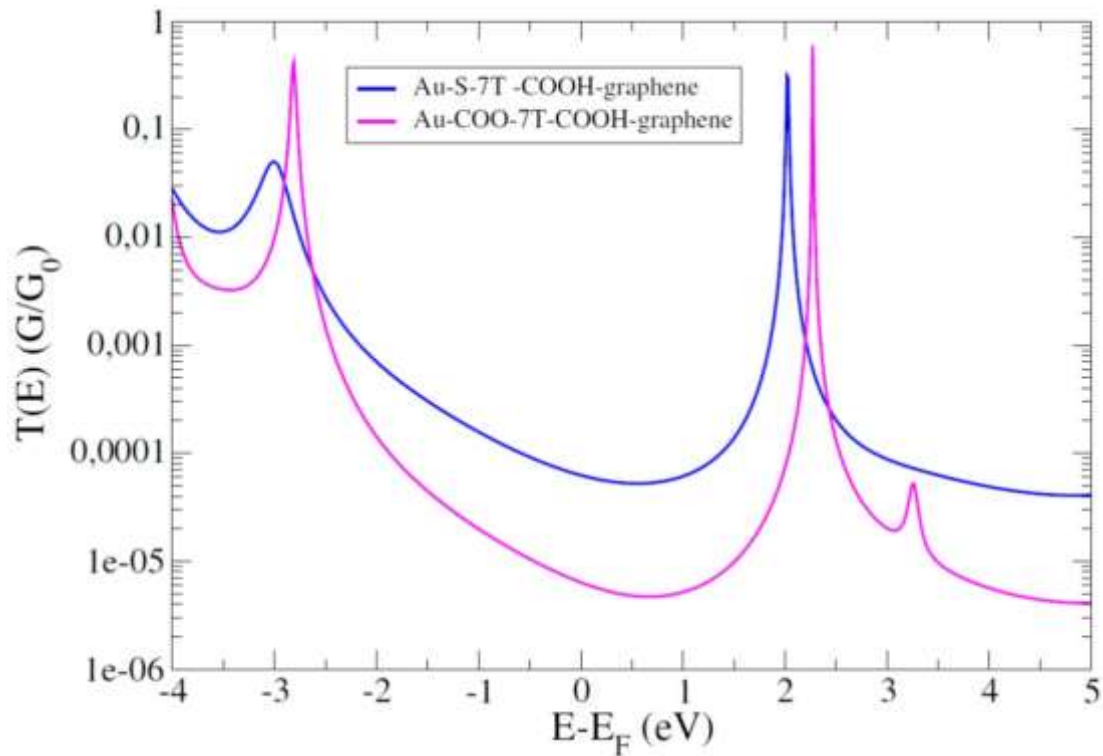


Figure 7 : evolution of the calculated electronic transmissions for Au-S-7T-COOH-graphene and Au-COO-6T-COOH-graphene molecular junctions.

Referring to Figure 7, the LUMO level in the junction with -COO(H) contacting groups at both molecular ends (“full COOH junction”) is located at about 2.4 eV from the Fermi level, which is a bit farther than in the case of the S-COOH junction, where the

LUMO is located at about 2 eV. This latter value is due to the stronger coupling of the sulphur atom to the gold electrode²⁶. This results in a slightly smaller electronic barrier for the anti-symmetric junction. Hence, by referring to the expression for β_n , we can deduce that the attenuation factor will be reduced for the Au-S-(CH₂)_n-COOH-graphene junctions with respect to the full COOH junctions, in a similar manner to the observation for the Au-S-(CH₂)_n-SH-graphene junctions. Another way of interpreting this value is to consider that the electrons choose the less resistive conductance channel offered by the coupling of the sulphur group to the electrode, as observed in macroscopic electronic circuits. In the case of Au/Au electrodes, Martín *et al.*²⁶ also measured the conductance of X-bridge-X, X-bridge-Y and Y-bridge-Y (X=thiol terminus and Y=COOH) molecular bridges and their results showed that the anti-symmetric configuration (Au-S-(CH₂)_n-COO-Au) produces the lowest conductance. In this respect, our present results show an unexpected trend in Figure 5 with a rather high conductance for the Au-S-(CH₂)_n-COOH-graphene configuration with respect to the Au-COO-(CH₂)_n-COOH-graphene configuration. To elucidate this peculiar behaviour, we analyse these two junctions within the frame of a single level model³⁹. Here, the LUMO level determines the conductance and therefore, the electronic transmission at the Fermi energy can be expressed as:

$$T(E_F) = 4\Gamma_L\Gamma_R / [(E_L - E_F)^2 + (\Gamma_L + \Gamma_R)^2]$$

with $\Gamma_{L/R}$ the left (right) coupling to the electrodes (namely Γ_{S-Au} or Γ_{COO-Au} for Γ_L and $\Gamma_{COOH-graphene}$ for Γ_R), $E_L - E_F$ the energy difference between the LUMO and the Fermi level, namely, the electronic barrier potential of the system. Comparing the Au-S-(CH₂)_n-COOH-graphene to the Au-COO-(CH₂)_n-COOH-graphene, it is known from Ref [19] that $\Gamma_{S-Au} > \Gamma_{COO-Au}$ (respectively 0.6 and 0.4 eV). Moreover, from Figure 7, $E_L - E_F$ is smaller for the anti-symmetric junction than for the full COOH junction (respectively 2 and 2.4 eV). Consequently, considering these values and $\Gamma_{COOH-graphene}$ necessarily smaller than Γ_{S-Au} or Γ_{COO-Au} (due to the weak

coupling at the molecule/graphene interface), we find that $T(E_F)$ is higher for the anti-symmetric junction than for the full COOH junction, which confirms the experimental and theoretical results presented in Figure 5. From this analysis, we can deduce that the anti-symmetric Au-S-(CH₂)_n-COOH-graphene junction is more conductive than the corresponding Au-COO-(CH₂)_n-COOH-graphene junction due to the stronger coupling at the S-Au interface, which has an effect both on the coupling strength Γ_{S-Au} and on the reduced potential barrier $E_L - E_F$. Hence, the electronic and transport properties of the anti-symmetric junction are driven by the most favourable conductance channel related to the thiol anchoring group.

Conclusions

In summary, we have studied the conductance of the anti-symmetric SH-(CH₂)_n-COOH molecular junctions ($n = 1, 3, 5, 7$ and 11), using graphene bottom and gold top electrodes. The conductance decays exponentially with the number of methylene groups with an attenuation factor of 0.37 , much lower than the value obtained with Au/Au electrodes. The obtained value is very similar to the one found for alkanedithiols, and smaller than for alkanedicarboxylic acids in the same configuration. Moreover, the conductance is also close to the one of the alkanedithiol, contrary to what is observed in Au/Au junctions. DFT-based electronic transport calculations and analysis of the relative molecule/electrode couplings and molecular level alignments confirm these features. As a result, the electronic transport properties of the anti-symmetric junctions are enhanced by the introduction of the graphene electrode, and driven by the most coupled anchoring group, namely the thiol group. In other words, the current in anti-symmetric junctions flows along the most favourable conductance channel associated to the sulphur atom, similarly to what happens in macroscopic circuits.

Author Contributions

Chunhui He carried out the main experiments and wrote the first draft of the manuscript, Li Yang, Yannick J. Dappe and Richard J. Nichols designed experiments and wrote the

manuscript; Qian Zhang, Yinqi Fan and Jingyao Ye conducted other complementary experiments; Cezhou Zhao and Chun Zhao analyzed experimental results. Yannick J. Dappe developed the simulation work and analyzed sequencing simulation data. All authors have given approval to the final version of the manuscript.

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References

1. A. Aviram and M. A. Ratner, *Chem. Phys. Lett.*, 1974, **29**, 277–283.
2. M. A. Reed, *Science*, 1997, **278**, 252–254.
3. B. Q. Xu and N. J. J. Tao, *Science*, 2003, **301**, 1221–1223.
4. X. D. Cui, A. Primak, X. Zarate, J. Tomfohr, O. F. Sankey, A. L. Moore, T. A. Moore, D. Gust, G. Harris and S. M. Lindsay, *Science*, 2001, **294**, 571–574.
5. W. Haiss, H. van Zalinge, S. J. Higgins, D. Bethell, H. Hobenreich, D. J. Schiffrin and R. J. Nichols, *J. Am. Chem. Soc.*, 2003, **125**, 15294–15295.
6. S. V. Aradhya and L. Venkataraman, *Nat Nanotechnol*, 2013, **8**, 399–410.
7. F. Chen and N. J. Tao, *Accounts Chem. Res.*, 2009, **42**, 429–438.
8. S. Marques-Gonzalez and P. J. Low, *Aust J Chem*, 2016, **69**, 244–253.
9. R. J. Nichols and S. J. Higgins, *Annu Rev Anal Chem*, 2015, **8**, 389–417.
10. R. J. Nichols and S. J. Higgins, *Curr. Opin. Electrochem.*, 2017, **4**, 98–104.
11. K. Wang and B. Q. Xu, *Topics Curr Chem*, 2017, **375**, 1–43.
12. D. Xiang, X. L. Wang, C. C. Jia, T. Lee and X. F. Guo, *Chem Rev*, 2016, **116**, 4318–4440.
13. S. Y. Jang, P. Reddy, A. Majumdar and R. A. Segalman, *Nano Lett.*, 2006, **6**, 2362–2367.
14. F. Chen, X. L. Li, J. Hihath, Z. F. Huang and N. J. Tao, *J. Am. Chem. Soc.*, 2006, **128**, 15874–15881.
15. N. B. Zhitenev, H. Meng and Z. Bao, *Phys Rev Lett*, 2002, **88**, 226801.
16. L. Venkataraman, J. E. Klare, C. Nuckolls, M. S. Hybertsen and M. L. Steigerwald, *Nature*, 2006, **442**, 904–907.
17. Y.-H. Wang, X.-Y. Zhou, Y.-Y. Sun, D. Han, J.-F. Zheng, Z.-J. Niu and X.-S. Zhou, *Electrochimica Acta*, 2014, **123**, 205–210.
18. Z. L. Peng, Z. B. Chen, X. Y. Zhou, Y. Y. Sun, J. H. Liang, Z. J. Niu, X. S.

- Zhou and B. W. Mao, *J. Phys. Chem. C*, 2012, **116**, 21699–21705.
19. Y. H. Wang, D. F. Li, Z. W. Hong, J. H. Liang, D. Han, J. F. Zheng, Z. J. Niu, B. W. Mao and X. S. Zhou, *Electrochemistry Communications*, 2014, **45**, 83–86.
 20. F. Chen, Z. F. Huang and N. J. Tao, *Applied Physics Letters*, 2007, **91**, 162106.
 21. A. Vezzoli, R. J. Brooke, N. Ferri, S. J. Higgins, W. Schwarzacher and R. J. Nichols, *Nano Lett.*, 2017, **17**, 1109–1115.
 22. A. C. Aragonés, N. Darwish, S. Ciampi, F. Sanz, J. J. Gooding and I. Diez-Perez, *Nature Communications*, 2017, **8**, 15056.
 23. R. M. Metzger, *Chem Rev*, 2015, **115**, 5056–5115.
 24. R. M. Metzger, B. Chen, U. Hopfner, M. V. Lakshmikantham, D. Vuillaume, T. Kawai, X. L. Wu, H. Tachibana, T. V. Hughes, H. Sakurai, J. W. Baldwin, C. Hosch, M. P. Cava, L. Brehmer and G. J. Ashwell, *J. Am. Chem. Soc.*, 1997, **119**, 10455–10466.
 25. E. Lortscher, *Chemphyschem*, 2011, **12**, 2887–2889.
 26. S. Martin, D. Z. Manrique, V. M. Garcia-Suarez, W. Haiss, S. J. Higgins, C. J. Lambert and R. J. Nichols, *Nanotechnology*, 2009, **20**, 125203.
 27. K. Wang, J. Zhou, J. M. Hamill and B. Xu, *J Chem Phys*, 2014, **141**, 054712.
 28. I. Diez-Perez, J. Hihath, Y. Lee, L. P. Yu, L. Adamska, M. A. Kozhushner, I. I. Oleynik and N. J. Tao, *Nature Chemistry*, 2009, **1**, 635–641.
 29. T. Tamaki and T. Ogawa, *Topics Curr Chem*, 2017, **375**.
 30. Q. Zhang, L. Liu, S. Tao, C. Wang, C. Zhao, C. Gonzalez, Y. J. Dappe, R. J. Nichols and L. Yang, *Nano Lett*, 2016, **16**, 6534–6540.
 31. Q. Zhang, S. Tao, Y. Fan, C. Zhao, C. Zhao, W. Su, Y. J. Dappe, R. J. Nichols and L. Yang, *The Journal of Physical Chemistry C*, 2018, **122**, 23200–23207.
 32. Q. Zhang, S. Tao, R. Yi, C. He, C. Zhao, W. Su, A. Smogunov, Y. J. Dappe, R. J. Nichols and L. Yang, *J Phys Chem Lett*, 2017, **8**, 5987–5992.
 33. L. Liu, Q. Zhang, S. Tao, C. Zhao, E. Almutib, Q. Al-Galiby, S. W. Bailey, I. Grace, C. J. Lambert, J. Du and L. Yang, *Nanoscale*, 2016, **8**, 14507–14513.
 34. B. Ren, G. Picardi and B. Pettinger, *Review of Scientific Instruments*, 2004, **75**, 837–841.
 35. T. A. Su, M. Neupane, M. L. Steigerwald, L. Venkataraman and C. Nuckolls, *Nature Reviews Materials*, 2016, **1**, 16002.
 36. C. H. He, Q. Zhang, S. H. Tao, C. Z. Zhao, C. Zhao, W. T. Su, Y. J. Dappe, R. J. Nichols and L. Yang, *Phys. Chem. Chem. Phys.*, 2018, **20**, 24553–24560.
 37. S. Martin, W. Haiss, S. Higgins, P. Cea, M. C. Lopez and R. J. Nichols, *J. Phys. Chem. C*, 2008, **112**, 3941–3948.
 38. L. L. Peng, B. Huang, Q. Zou, Z. W. Hong, J. F. Zheng, Y. Shao, Z. J. Niu, X. S. Zhou, H. J. Xie and W. Chen, *Nanoscale Res Lett*, 2018, **13**, 121.
 39. B. Larade and A. M. Bratkovsky, *Phys. Rev. B*, 2003, **68**, 235305.