



Thermal expansion of α -boron and some boron-rich pnictides

Kirill Cherednichenko, Vladimir Solozhenko

► To cite this version:

Kirill Cherednichenko, Vladimir Solozhenko. Thermal expansion of α -boron and some boron-rich pnictides. Solid State Communications, 2019, 303-304, pp.113735. 10.1016/j.ssc.2019.113735 . hal-02331172

HAL Id: hal-02331172

<https://hal.science/hal-02331172>

Submitted on 24 Oct 2019

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Thermal expansion of α -boron and some boron-rich pnictides

Kirill A. Cherednichenko and Vladimir L. Solozhenko *

LSPM-CNRS, Université Paris Nord, 93430 Villetaneuse, France

Abstract

Thermal expansion of α -rhombohedral boron (α -B₁₂) and two isostructural boron-rich pnictides (B₁₂P₂ and B₁₂As₂) has been studied between 298 and 1280 K by high-temperature synchrotron X-ray diffraction. For all studied phases no temperature-induced phase transitions have been observed. The observed temperature dependencies of the lattice parameters and unit cell volumes were found to be quasi-linear. Variation of the thermal expansion coefficients in the group of boron-rich pnictides (B₁₃N₂ – B₁₂P₂ – B₁₂As₂) was analyzed.

Keywords : A. Boron; A. Boron pnictides; C. X-ray diffraction; D. Thermal expansion

I. Introduction

During last few decades elemental boron and boron-rich compounds have been attracting a considerable attention due to their unusual crystal structure and properties. The common features of these compounds are B₁₂ icosahedra and other *closo*-clusters with polycentric metal-like bonding system [1]. At the same time, boron atoms within *closo*-clusters form strong covalent bonds with neighbouring atoms of the adjacent clusters and various interstitial atoms (e.g. C, N, O, P, Zr, Y), giving a big variety of boron-rich compounds: B₄C, B₁₂O₂, B₁₃N₂, ZrB₁₂, YB₁₂, etc. A combination of the metal-like intra-cluster bonds and strong covalent inter-cluster bonding makes boron-rich compounds extremely stable, eventually leading to high melting temperatures, chemical inertness, outstanding mechanical properties and interesting electronic properties (e.g. self-healing resistance to radiation damage, superconductivity, etc.) [1-9]. It should be noted that properties of boron-rich solids are significantly influenced by the interstitial atoms [3,7]. Thus, study of the impact of interstitial atoms on various properties (hardness, compressibility, thermal expansion, band gap, etc.) is of great interest from both fundamental and application points of view.

Very recently thermal expansion of two boron-rich nitrides synthesized under extreme conditions has been studied by synchrotron X-ray diffraction [10]. In the present paper we report the results on thermal expansion of α -rhombohedral boron (α -B₁₂) and two boron-rich pnictides, B₁₂P₂ and B₁₂As₂. The retrieved thermal expansion coefficients compared with those reported earlier [10,11] allowed us to analyze the thermal expansion variation in the family of boron-rich pnictides.

* Corresponding author: vladimir.solozhenko@univ-paris13.fr

II. Experimental

Highly crystalline α -B₁₂ (99.98%) was provided by Dr. Igor N. Goncharenko, Laboratoire Léon Brillouin (CEA-Saclay, France). Boron subphosphide B₁₂P₂ was produced by self-propagating high-temperature reaction in the BPO₄–MgB₂–Mg system according to the method described elsewhere [12]. Boron subarsenide B₁₂As₂ was synthesized by reaction of amorphous boron with arsenic melt at 5 GPa and 2100 K in a toroid-type apparatus (for details, see [13]). According to X-ray diffraction study (TEXT 3000 Inel diffractometer, CuK α 1 radiation) all samples are single-phase, and their lattice parameters determined using Maud [14] and PowderCell [15] software are in good agreement with literature data [1,3].

Thermal expansion of boron-rich pnictides and α -B₁₂ has been studied at MCX beamline, Elettra Synchrotron (Trieste, Italy) and B2 beamline, DORIS III storage ring, HASYLAB-DESY (Hamburg, Germany). The powdered samples were loaded into quartz-glass capillaries under Ar atmosphere. Debye-Scherrer geometry with rotating capillary was used at both beamlines. The wavelength (1.0352 Å at MCX beamline and 0.6841 Å at B2 beamline) was calibrated using Si and LaB₆ as external standards; temperature calibration has been done by measuring thermal expansion of platinum [16] under the same experimental conditions. X-ray diffraction patterns were collected in the 2-60 θ -range using image plate detectors upon stepwise heating with 100-K steps from room temperature up to 1280 K (see Fig. 1). The ramp time was 10 min with 5 min of temperature stabilization; the acquisition time was 60-120 s. FIT2D software [17] was used to integrate the collected X-ray diffraction data into one-dimensional patterns. The lattice parameters of the studied phases at different temperatures were determined by Le Bail refinement procedure using PowderCell software [15].

III. Results and Discussion

α -rhombohedral boron is the low-temperature allotrope, and its thermodynamic stability domain in the equilibrium phase diagram of boron [18] is still not precisely defined. It has rhombohedral crystal structure ($R\bar{3}m$) containing B₁₂ icosahedra located at the vertices of the rhombohedral cell. Due to the lack of the valence electrons necessary for the formation of inter-icosahedral covalent bonds, the equatorial boron atoms of the neighboring icosahedra lying in one layer forming relatively weak "two electrons-three centers" bonds, $2e3c$ [1,2] (Fig. 2a). The boron-rich pnictides, B₁₂X₂Y (where X = N, P, As; Y = B), have structures related to α -rhombohedral boron with two or three interstitial atoms in $1b$ and $2c$ Wyckoff positions (in hexagonal setting, $3b$ and $6c$) placed along the body diagonal of the rhombohedral unit cell (Fig. 2b). The interstitial atoms in $2c$ positions form the "two electrons-two centers" covalent bonds, $2e2c$, with equatorial boron atoms of the neighboring icosahedra.

The unit cell parameters of α -B₁₂, B₁₃N₂ [10], B₁₂P₂ and B₁₂As₂ in hexagonal (a^{hex} , c^{hex}) and rhombohedral (a^{rh} , α^{rh}) settings are presented in Table I. Further we will consider the thermal expansion in hexagonal unit cell only, so the "hex" indexes of lattice parameters will not be used. As one can see, a parameter increases in the α -B₁₂ – B₁₃N₂ – B₁₂P₂ – B₁₂As₂ row, which happens

due to insertion of the interstitial atoms in α -boron unit cell and subsequent "squeezing" of the B_{12} units outside [3]. On the other hand, X-X bonds ($X = N, P, As$) tend to pull closer the adjacent planes of icosahedra resulting in decrease of c parameter [19]. Different "compression" of boron-rich pnictides along c -axis might be explained by different nature of X-X bonds i.e. P-P and As-As are $2e2c$ covalent bonds (however, they are considerably longer than the corresponding diatomic equilibrium bonds [20][20], see Table II), whereas B-N-B chain in the $B_{13}N_2$ structure can be considered as $3e3c$ bond (for details, see [21]), and is significantly longer. Nevertheless, as it follows from the formula for the hexagonal unit cell volume:

$$V_0 = \frac{\sqrt{3}}{2} a^2 c \quad (1)$$

the enlargement of hexagonal unit cell in the ab -plane has significantly greater impact on increase of the unit cell volume than the "compression" along c -axis.

The results of high-temperature X-ray diffraction study of α - B_{12} , $B_{12}P_2$ and $B_{12}As_2$ (Fig. 3) revealed quasi-linear temperature dependency of the lattice parameters, hence, the linear thermal expansion coefficients (TEC) do not change over the whole temperature range under study. The linear TECs (α_l) can be estimated by Eq. 2 [22]:

$$\alpha_l = \frac{l - l_{298\text{ K}}}{l_{298\text{ K}} \cdot (T - 298\text{ K})} \quad (2)$$

where l is a unit cell parameter and $l_{298\text{ K}}$ is a unit cell parameter at room temperature. The α_l values are presented in Table I. The thermal expansion of different interatomic distances was not analyzed due to insufficient quality of powder X-ray diffraction data (for that the single-crystal diffraction data are highly required); here and further the thermal expansion of the inter- and intra-icosahedral bonds is assumed to be the same for all studied solids.

As it has been expected, thermal expansion of α -rhombohedral boron and isostructural boron-rich pnictides in different crystallographic directions was found to be anisotropic. The anisotropy was estimated by α_c/α_a ratio (see Table I): in contrast to α - B_{12} , thermal expansion of all boron-rich pnictides along c -axis (α_c) is larger than that in the ab -plane (α_a) (see Fig. 4a). Moreover, the liner TEC values of α -boron and boron-rich pnictides vary considerably.

Analysis of the α_c variation of boron-rich pnictides requires special consideration of B-X and X-X bonds. According to Fig. 2, the interstitial atoms ($X = N, P, As$) in $3c$ Wyckoff positions form the bonds with four neighboring atoms and, hence, are tetrahedrally coordinated as well as X atoms in cubic boron pnictides BN [23], BP [24] and BAs [25]. All four B-X bonds in the cubic BX phase have equal lengths, thus, forming an ideal tetrahedron. Taking into account thermodynamic stability of all cubic boron pnictides, the lengths of the B-X bonds in these compounds were assumed to be the optimal for the tetrahedral geometry. Unlike cubic boron pnictides, the tetrahedra (with central X atom) in $B_{13}N_2$, $B_{12}P_2$ and $B_{12}As_2$ are significantly distorted (see Table II). Employing the explanation that has been already used in the case of $B_{50}N_2$ thermal expansion [10] we assume that thermal vibrations lead to the reduction of the distorted tetrahedra towards the ideal ones i.e. at high temperatures the X-X bonds in $B_{12}P_2$ and $B_{12}As_2$ tend to shrink (see Table II for corresponding diatomic equilibrium bonds), whereas N-B-N chains in $B_{13}N_2$ tend to expand. In the case of boron subnitride, nitrogen atom is in the center of tetrahedron formed by boron atoms only, that is why the B-N distances in $B_{13}N_2$ are comparable with the corresponding bond lengths in cubic BN, but not

with diatomic equilibrium B–N bond. Such a different thermal expansion of boron-rich pnictides along c -axis results in α_c decrease in the row: $B_{13}N_2 - B_{12}P_2 - B_{12}As_2$ (see Fig. 4a). Despite one could expect the maximal thermal expansion in c -direction for α - B_{12} , its α_c value was found to be close to that of $B_{12}As_2$.

Following the same logic, one should expect that at high temperatures B–P and B–As bonds in the ab -plane will tend to expand (see the difference between B–X bond lengths in $B_{12}X_2$ and cubic BX, Table II), while B–N bonds will tend to shrink, which should result in the higher α_a values for $B_{12}P_2$ and $B_{12}As_2$. Nevertheless, the hexagonal unit cell initially enlarged in the ab -plane leads to the smaller increase of the inter-icosahedral distances and B–X bonds at high temperatures, and, thus, to lower α_a values; α_a decrease in the row $B_{13}N_2 - B_{12}P_2 - B_{12}As_2$. The largest thermal expansion of α - B_{12} in the ab -plane might be explained by the presence of weak $2e3c$ inter-icosahedral bonds instead of $2e2c$ covalent bonds in boron-rich pnictides i.e. stronger $2e2c$ bonds prevent significant thermal expansion in this crystallographic direction.

Fig. 3 presents the variations of the normalized unit cell volumes $V(T)/V_0$ (V_0 is the unit cell volume at 298 K) of α - B_{12} , $B_{12}P_2$ and $B_{12}As_2$ *versus* temperature. The observed temperature dependencies of unit cell volume are quasi-linear and, thus, can be approximated by Eq. 3 [26]:

$$V(T) = V_0(1 + \alpha_v(T - 298 \text{ K})), \quad (3)$$

where α_v is volume TEC.

The retrieved volume TECs of α - B_{12} , $B_{12}P_2$ and $B_{12}As_2$ (see Table I) were found to be of the same order of magnitude as those of other boron-rich solids [10,11,22,26-30]. The α_v value of $B_{12}As_2$ ($15.3(1) \times 10^{-6} \text{ K}^{-1}$) is in excellent agreement with the literature data ($15(2) \times 10^{-6} \text{ K}^{-1}$) [11]. According to Table I, boron subnitride has the highest volume TEC among all studied phases. The main reasons for that are: not too much expanded unit cell at room temperature (thus, allowing the further increase of the interatomic distances) and the presence of $3e3c$ N–B–N chains with its tendency to expand at high temperatures. α - B_{12} has the second largest volume TEC owing to strong impact of the rather weak $2e3c$ bonds resulting in significant expansion in the ab -plane. The volume thermal expansion of boron subphosphide and subsarsenide is considerably influenced by the "negative" tendency for the P–P and As–As bonds at high temperatures and the unit cells initially enlarged in a and b directions.

Interestingly, the volume TECs and bulk moduli (B_0) of $B_{13}N_2$, $B_{12}P_2$ and $B_{12}As_2$ [31-33] vary similar (see Table I and Fig. 4b). As one can see, the α_v vs B_0 dependence for boron-rich pnictides can be approximated by the following exponential function:

$$\alpha_v = 14.94 + 1.37 \cdot 10^{-4} \cdot e^{0.05 \cdot B_0}, \quad (4)$$

Data for α -rhombohedral boron is also presented in Fig. 3 (we used B_0 value of α - B_{12} reported in [34]), however it does not follow the exponential $\alpha_v(B_0)$ dependency. It should be noted that for all studied phases we did not observe temperature-induced phase transitions and/or decomposition up to the highest experimental temperatures.

IV. Conclusions

Thermal expansion of α -rhombohedral boron and isostructural boron-rich pnictides ($B_{12}P_2$, $B_{12}As_2$) was studied *in situ* by synchrotron X-ray diffraction up to 1280 K. The precise measurements of

lattice parameters at different temperatures allowed us to retrieve the corresponding linear (α_l) and volume (α_v) thermal expansion coefficients of α -B₁₂, B₁₂P₂ and B₁₂As₂. The obtained values were compared with the literature data for boron subnitride B₁₃N₂, and variation of α_l and α_v of three boron-rich pnictides was analyzed in terms of their crystal structure.

Acknowledgements

The authors thank Dr. Vladimir Mukhanov for the samples synthesis and Drs. Jasper Rikkert Plaisier, Lara Gigli and Giulio Zeraushek for assistance in measurements at Elettra. Synchrotron X-ray diffraction experiments were carried out during beam time allocated for Proposal 20160086 at Elettra Sincrotrone Trieste and Proposal I-20090172 EC at HASYLAB-DESY. This work was financially supported by the European Union's Horizon 2020 Research and Innovation Programme under Flintstone2020 project (grant agreement No 689279).

ORCID IDs

Vladimir L. Solozhenko  <https://orcid.org/0000-0002-0881-9761>

References

- [1]. B. Albert, H. Hillebrecht, *Angew. Chem. Int. Ed.*, **48**, 8640 (2009).
- [2]. D. Emin, *Phys. Today*, **40**, 55 (1987).
- [3]. G.A. Slack, K.E. Morgan, *J. Phys. Chem. Solids*, **75** 1054 (2014).
- [4]. S. Veprek, R.F. Zhang, A.S. Argon, *J. Superhard Mater.*, **33**, 409 (2011).
- [5]. D. Emin, *J. Solid State Chem.*, **179**, 2791 (2006).
- [6]. M. Carrard, D. Emin, L. Zuppiroli, *Phys. Rev. B*, **51**, 11270 (1995).
- [7]. G.A. Slack, K.E. Morgan, *Solid State Sci.*, **47**, 43 (2015).
- [8]. G. Akopov, M.T. Yeung, Z.C. Sobell, C.L. Turner, C.W. Lin, R.B. Kaner, *Chem. Mater.*, **28**, 6605 (2016).
- [9]. J.Y. Ge, V.N. Gladilin, N.E. Sluchanko, A. Lyashenko, V.B. Filipov, J.O. Indekeu, V.V. Moshchalkov, *New J. Phys.*, **19**, 093020, (2017).
- [10]. K.A. Cherednichenko, L. Gigli, V.L. Solozhenko, *Solid State Comm.*, **275**, 21 (2018).
- [11]. C.E. Whiteley, M.J. Kirkham, J.H. Edgar, *J. Phys. Chem. Solids*, **74**, 673 (2013).
- [12]. V.A. Mukhanov, P.S. Sokolov, O. Brinza, D. Vrel, V.L. Solozhenko, *J. Superhard Mater.*, **36**, 18 (2014).
- [13]. K.A. Cherednichenko, Y. Le Godec, V.L. Solozhenko, *High Pressure Res.*, **38**, 224 (2018).
- [14]. L. Lutterotti, M. Bortolotti, G. Ischia, I. Lonardelli, H.-R. Wenk, *Z. Kristallogr. Suppl.* **26**, 125 (2007).
- [15]. W. Kraus, G. Nolze, *J. Appl. Cryst.*, **29**, 301 (1996).
- [16]. P. Riello, A. Lausi, J. Macleod, J. Rikkert Plaisier, G. Zeraushek, P. Fornasiero, *J. Synchrot. Rad.*, **20** 194 (2013).
- [17]. A.P. Hammersley, S.O. Svensson, M. Hanfland, A.N. Fitch, D. Hausermann, *High Pressure Res.*, **14**, 235 (1996).
- [18]. V.L. Solozhenko, O.O. Kurakevych, *Sci. Rep.*, **3**, 2351 (2013).
- [19]. C.L. Beckel, N. Lu, B. Abbott, M. Yousaf, *Inorg. Chim. Acta*, **289**, 198 (1999).
- [20]. K.P. Huber, G. Herzberg, "Constants of Diatomic Molecules", Van Nostard Reinhold, New York (1979).
- [21]. S. Mondal, E. Bykova, S. Dey, Sk. Imran Ali, N. Dubrovinskaia, L. Dubrovinsky, G. Parakhonskiy, S. van Smaalen, *Sci. Rep.*, **6**, 19330 (2016).
- [22]. G.V. Tsagareishvili, T.G. Nakashidze, J.Sh. Jobava, G.P. Lomidze, D.E. Khulelidze, D.Sh. Tsagareishvili, O.A. Tsagareishvili, *J. Less-Common Met.*, **117**, 159 (1986).
- [23]. V.L. Solozhenko, V.V. Chernyshev, G.V. Fetisov, V.B. Rybakov, I.A. Petrusha, *J. Phys. Chem. Solids*, **51**, 1011 (1990).
- [24]. P. Popper, T.A. Ingles, *Nature*, **179**, 1075 (1957).

- [25]. J.A. Perri, S. La Placa, B. Post, *Acta Cryst.*, **11**, 310(1958).
- [26]. O.O. Kurakevych, V. L. Solozhenko, *J. Superhard Mater.*, **36**, 270 (2014).
- [27]. T.R. Pilladi, G. Panneerselvam, S. Anthonysamy, V. Ganesan, *Ceram. Int.*, **38**, 3723 (2012).
- [28]. H.L. Yakel, *J. Appl. Cryst.*, **6** (1973) 471.
- [29]. C.E. Holcombe Jr., D.D. Smith, J.D. Lore, W.K. Duerksen, D.A. Carpenter, *High Temp. Sci.*, **5**, 349 (1973).
- [30]. T. Lundstrom, B. Lonnberg, J. Bauer, *J. Alloys Compd.*, **267**, 54 (1998).
- [31]. K.A. Cherednichenko, Y. Le Godec, V.L. Solozhenko, *High Pressure. Res.*, **38**, 224 (2018).
- [32]. V.L. Solozhenko, K.A. Cherednichenko, O.O. Kurakevych, *J. Superhard Mater.*, **39**, 71 (2017).
- [33]. K.A. Cherednichenko, V.L. Solozhenko, *J. Appl. Phys.*, **122**, 155901 (2017).
- [34]. I. Chuvashova, E. Bykova, M. Bykov, V. Svitlyk, L. Dubrovinsky, N. Dubrovinskaia, *Sci. Rep.*, **7**, 8969 (2017).

Table I. Lattice parameters, unit cell volumes, mean linear thermal expansion coefficients (α_l), linear thermal expansion anisotropy parameters (α_a/α_c), volume thermal expansion coefficients (α_v) and bulk moduli (B_0) of α -B₁₂, B₁₃N₂, B₁₂P₂ and B₁₂As₂.

	α -B ₁₂	B ₁₃ N ₂ [10]	B ₁₂ P ₂	B ₁₂ As ₂
$a, b^{hex}, \text{\AA}$	4.9161(3)	5.4537(3)	5.9894(3)	6.1353(2)
$c^{hex}, \text{\AA}$	12.5752(5)	12.2537(7)	11.8594(8)	11.8940(7)
$V^{hex}, \text{\AA}^3$	263.20(4)	315.62(5)	368.43(6)	387.72(5)
$a^{rh}, \text{\AA}$	5.0623(3)	5.1573(3)	5.2521(1)	5.3166(2)
$\alpha^{rh}, ^\circ$	58.09	63.84	69.53	70.48
$V^{rh}, \text{\AA}^3$	87.73(1)	105.21(2)	122.81(2)	129.24(2)
$\alpha_a \times 10^6, \text{K}^{-1}$	6.4(3)	6.1(1)	5.7(1)	5.0(1)
$\alpha_c \times 10^6, \text{K}^{-1}$	5.4(2)	8.9(1)	6.5(1)	5.3(1)
α_c/α_a	0.84	1.46	1.14	1.06
$\alpha_v \times 10^6, \text{K}^{-1}$	18.3(6)	21.3(2)	17.9(3)	15.3(1)
B_0, GPa	224(7) [34]	205(2) [33]	192(11) [32]	150(4) [31]

Table II. Bond lengths in boron-rich pnictides and cubic boron pnictides compared with diatomic equilibrium B–B and B–X (X = N, P, As) distances.

	BN [23]	BP [24]	BA _s [25]	α -B ₁₂	B ₁₃ N ₂ [10]	B ₁₂ P ₂	B ₁₂ As ₂
B–X, Å	1.566	1.967	2.069	–	1.6302	1.9074	1.9914
X–X, Å	–	–	–	–	1.5390*	2.2428	2.3833
Diatomic equilibrium bonds [20], Å				B–B	B–N	P–P	As–As
				1.590	1.281	1.893	2.103

* X–X distance in B₁₃N₂ is given for the N–B bond of the N–B–N chain.

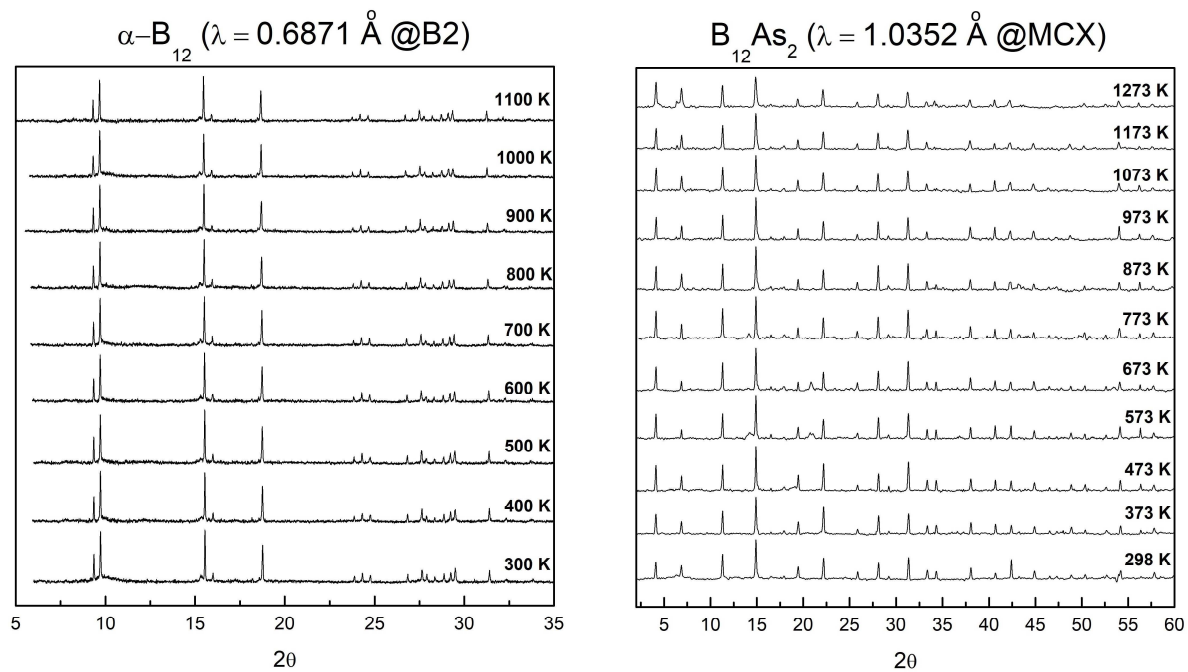


Fig. 1 X-ray diffraction patterns of α -rhombohedral boron (*left*) and B₁₂As₂ boron subarsenide (*right*) acquired at B2 (DORIS III) and MCX (Elettra), respectively.

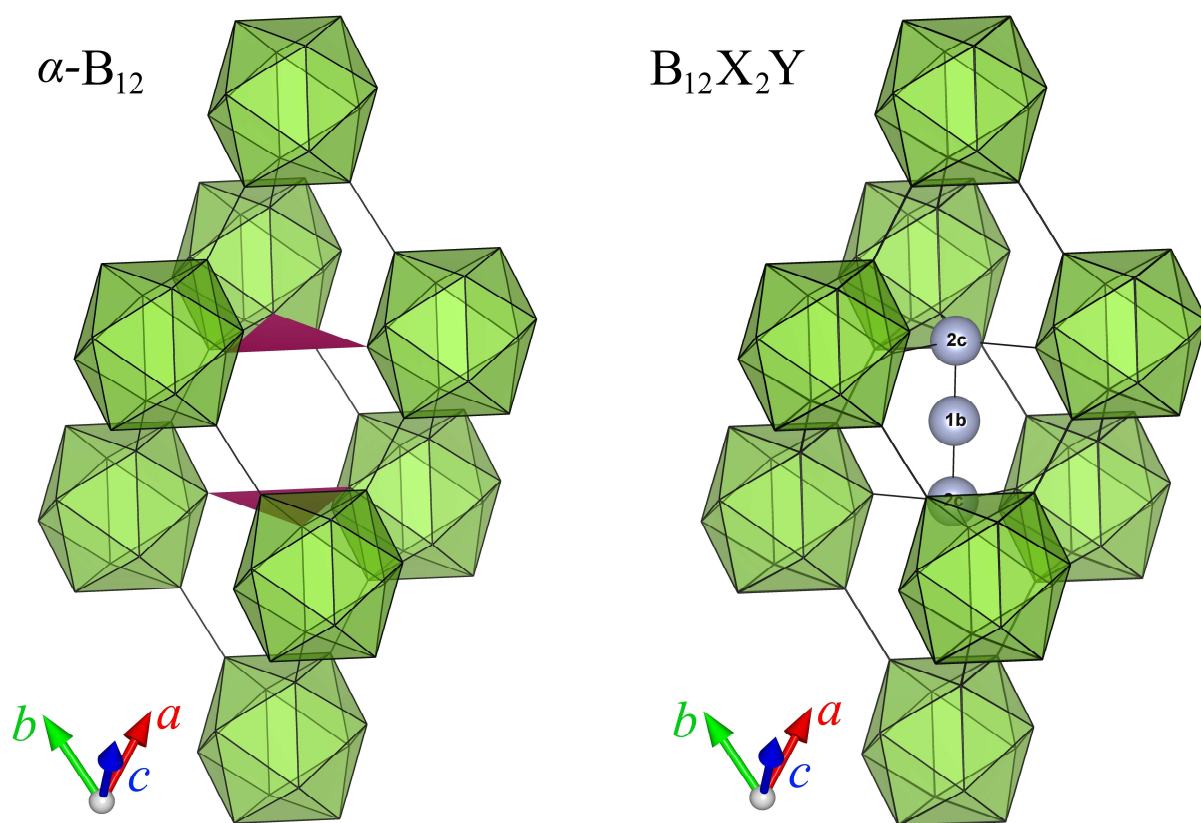


Fig. 2 Crystal structures of α -rhombohedral boron (*left*) and $\alpha\text{-B}_{12}$ -related boron-rich pnictides, $\text{B}_{12}\text{X}_2\text{Y}$, (*right*) in rhombohedral setting. The B_{12} clusters are presented by green icosahedra, the $2e3c$ bonds in $\alpha\text{-B}_{12}$ are shown by rose triangles, the interstitial X (X = N, P, As) and Y atoms (Y = B) in $2c$ and $1b$ Wyckoff positions, respectively, are shown by grey balls.

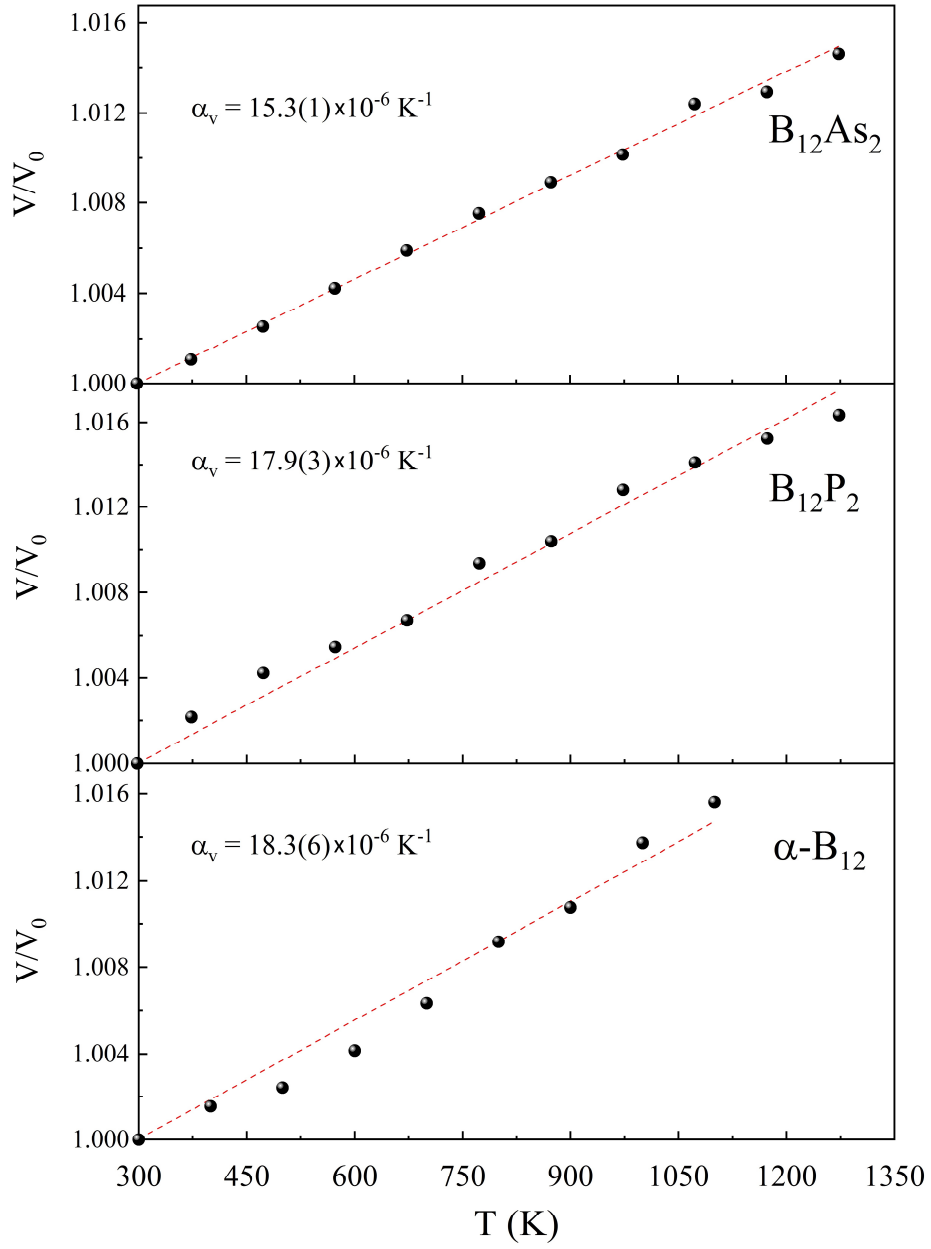


Fig. 3 Variation of the normalized unit cell volumes of α -B₁₂, B₁₂P₂ and B₁₂As₂ as function of temperature. The dashed lines represent the linear fits to the experimental data. The corresponding values of volume thermal expansion coefficients are indicated.

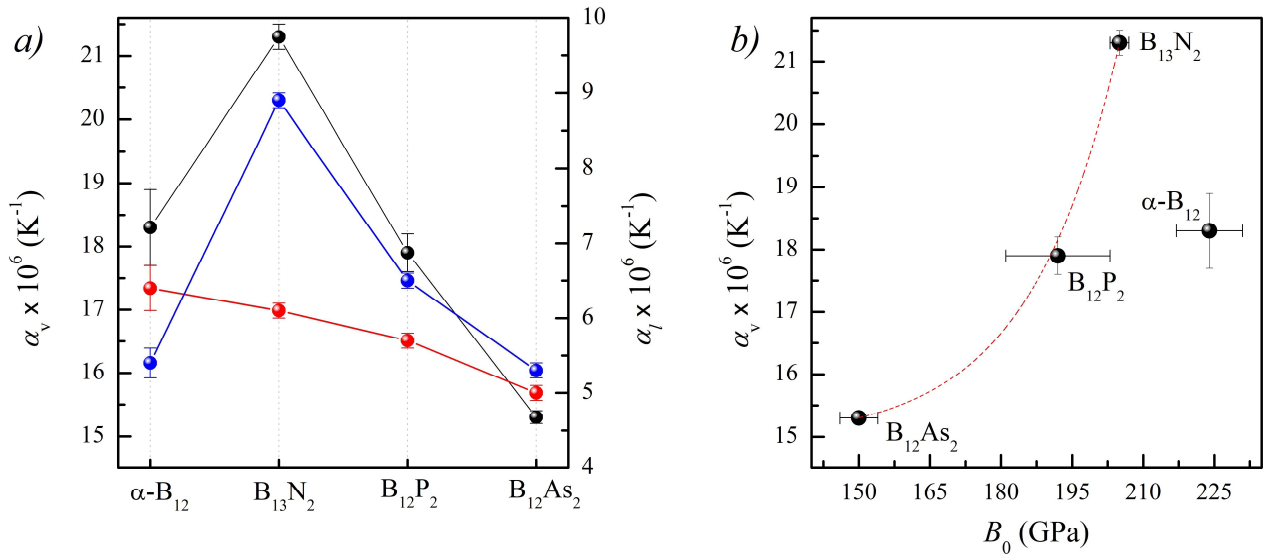


Fig. 4 **a)** Volume (α_v , black) and linear (α_a , red, and α_c , blue) thermal expansion coefficients of $\alpha\text{-B}_{12}$, B_{13}N_2 [8], B_{12}P_2 and B_{12}As_2 ; **b)** Volume (α_v) thermal expansion coefficients of $\alpha\text{-B}_{12}$, B_{13}N_2 , B_{12}P_2 , B_{12}As_2 versus bulk modulus (B_0) [29-32]; the dashed line represents exponential fit.