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Molecular Reactivity of Busulfan Through Its Experimental Electrostatic Properties in the Solid State

Nour Eddine Ghermani,^{1,2,4} Anne Spasojević-de Bire,²

Nouzha Bouhmaida,^{2,3} Souad Ouharzoune,^{1,2} Jérôme Bouligand,¹ Anne Layre,¹ Ruxandra Gref,¹ and Patrick Couvreur¹

Purpose. In the route of developing novel liquid phase formulations based on the encapsulation of busulfan into liposomes in nontoxic solvents, drug crystallization inevitably occurs. In order to better understand the reactivity of busulfan, the characterization of its molecular properties was therefore considered as a key point. Also, preliminary attempts to prevent crystallization using cyclodextrins were explored.

Methods. An accurate single-crystal high-resolution X-ray diffraction experiment at 100 K has been carried out. The experimental electron density of busulfan was refined using a multipole model. Busulfan/ cyclodextrin coprecipitates were analyzed by powder X-ray diffraction and ¹H-NMR spectroscopy.

Results. The electrostatic properties of busulfan and the methylsulfonate fragment dipole moment (3.2 D) were determined. The polar moieties play a key role in the crystallization of busulfan, which presents a nucleophilic region surrounding the sulfonate part, whereas the carbon chain displays an electrophilic character. This highlights the subtle busulfan/cyclodextrin association.

Conclusions. Busulfan electrostatic properties were used to quantify its chemical reactivity. This explains the difficulty to formulate busulfan into liposomes due to a strong polar character of the methylsulfonate terminal groups. The complexation with cyclodextrins deserves to be further investigated to allow the formulation of busulfan in nontoxic solvents.

KEY WORDS: busulfan; busulfan–cyclodextrin complex; electron density; high-resolution crystallography.

INTRODUCTION

The formulation of molecules with a trend to crystallize is a major problem in pharmacy, not only because the shape and overall dimensions of the obtained crystals have an important impact in the mechanical manufacturing of solid dosage forms (storage, flow, compression, tableting, and so forth), but also because it may impact the release profile of

¹ Laboratoire Physico-Chimie, Pharmaceutique, Biopharmacie (PPB) UMR CNRS 8612, Faculté de Pharmacie, Université Paris XI, 92296 Châtenay-Malabry Cedex, France. ²

Laboratoire Structures, Propriété et Modélisation des Solides (SPMS) UMR CNRS 8580, Ecole Centrale Paris, 92295 Châtenay-Malabry Cedex, France. ³

Laboratoire des Sciences des Matériaux (LSM), Faculté des Sciences Sémalialia, Université Cadi Ayyad, BP 2390, 40000 Marrakech, Morocco. ⁴

To whom correspondence should be addressed. (e-mail: noureddine.ghermani@cep.u-psud.fr)

the drug. Additionally, the design of liquid pharmaceutical formulations with crystallizable molecules is still more problematic, especially when the drug concentration is close to the limit of solubility. Another concern is the risk of the crystallization of drugs *in vivo* after administration, as it is the case with busulfan (1,4-butanediol-dimethylsulfonate), a powerful antitumoral agent in leukemias (1). Typically, one of the main side effects observed clinically for therapies with a high dose rate of busulfan is the hepatic veno-occlusive complication due to the microcrystallization of this drug in the microvenous system of the liver (2–4). These considerations have promoted the search for more appropriate and effective formulations of busulfan among which liposomes have been proposed and tested in phase I/II clinical trials (5). However, although these formulations were not satisfyingly characterized, it is likely that in the conditions used for the preparation of the liposomes, busulfan should crystallize outside the vectors. In general, there is a lack of understanding the drug properties at the atomic level to guide the formulation in order to find new strategies to prevent the crystallization. In the current study, using busulfan as a drug model, we have carried out a high-resolution X-ray diffraction experiment at 100 K to characterize the electron density and electrostatic properties of this drug in the solid state. Crystal structures of busulfan and related molecules [CH₃–SO₂–O(CH₂)_n–SO₂–CH₃, n = 2, 3, 4] were first reported by MacKenna *et al.* (6) from room-temperature experiments. The authors focused on the molecular conformations with respect to the DNA crosslinking potential. In our work, we try to go further in the investigation of busulfan electronic properties. First, the characterization of the molecular bond electron density through topological analysis using the “Atoms in Molecules” theory (7) clearly demonstrated the alkylating power of busulfan in relation to the cross-linking of DNA strands. Afterwards, the electrostatic potential and the methylsulfonate fragment dipole moment were quantified experimentally. These two properties are fundamental to highlight the reactivity of the busulfan and to better understand the self-affinity of molecules giving rise to the observed strong crystallization.

Finally, in the search of an alternative way to stop the crystallization of busulfan in nontoxic solvents for further formulations, we considered the use of cyclodextrins. These molecules are able to form inclusion complexes with drugs, enhancing their stability and

solubility in water (8,9). The second part of this study concerns the characterization of the busulfan/-cyclodextrin-associated systems by powder X-ray diffraction and ¹H-NMR spectroscopy. Correlations with the molecular electrostatic properties were established.

MATERIALS AND METHODS

Materials

Crystallization of Busulfan

Crystalline white powder of busulfan was purchased from Sigma-Aldrich (Schnelldorf, Germany). Busulfan is soluble in numerous common organic solvents but weakly in pure water. Crystals are easily obtained in few hours by slow evaporation at room temperature. The addition of water, polymers, or other molecules in different v/v ratios with busulfan in cosolvent solutions did not affect the crystallization of the drug (10). However, no crystals of busulfan were obtained in more polar solvents like dimethylsulfoxide (DMSO) or dimethylacetamide (DMAC) in normal conditions of supersaturated solution evaporation. The crystals of busulfan are of very good quality, colorless, and have a shape of prisms or elongated blocks. The crystal of busulfan used in this study was obtained from a solution

of ethanol.

Preparation of Busulfan–Cyclodextrin Systems

Crystalline powder of native -cyclodextrin hydrate was purchased from Roquette Frères (Lestrem, France). Thirty milliliters of acetone unsaturated solution of busulfan (0.16 and 0.08 g for 1:1 and 1:2 molar ratios) was added to 50 ml of aqueous solution of -cyclodextrin (0.75 g). The limpid solutions were stirred for 5 days at room temperature. In a slow kinetic process, a precipitation progressively appeared in the glass beakers. The solutions were then filtered, and the solid coprecipitates (white powders) were air-dried during 24 h and first analyzed by X-ray powder diffractometry. Afterwards, the busulfan–cyclodextrin coprecipitates were dissolved in deuterated water (D₂O) for a ¹H-NMR spectroscopy analysis.

Methods

Single-Crystal X-Ray Data Collection

The single-crystal X-ray high-resolution data collection was performed at 100 K on a Bruker-Smart 3-circle (, ,) diffractometer (Karsrhue, Germany) equipped with a CCDship area detector. The X-ray beam corresponds to the Mo K radiation (0.71073 Å). The high resolution reached was $d = 0.408 \text{ \AA}$ ($1 \text{ \AA} = 10^{-10} \text{ m}$) corresponding to $(\sin/\lambda)_{\text{max}} = 1.225 \text{ \AA}^{-1}$. The collected frames were recorded with a scan interval of 0.15°. The SAINT program (11) of the Smart software package was used to correct data from the Lorentz and polarization phenomena and to integrate the peak diffraction intensities. In the experiment, 46,878 reflections were collected, yielding 5421 unique data having intensities $I > 3(I)$ used as observations in the electron density refinements. The standard deviations (I) were estimated from the variance analysis performed during the averaging of the equivalent reflection intensities (12). Table I summarizes the experimental details.

X-Ray Powder Diffractometry and ¹H-NMR Spectroscopy

X-ray diffraction lines of the busulfan–cyclodextrin hydrate and the coprecipitates of busulfan–cyclodextrin (molar ratios 1:1 and 1:2) were recorded at room temperature on a Siemens D5000 diffractometer (Karsrhue, Germany) using Cu K radiation (1.54051 Å, 40 kV, and 35 mA) in

Bragg–Brentano reflection geometry. The powder diffraction patterns correspond to a 2 angle scan range from 3° to 70° with a 0.02° step size and 1 s/step. The diffractograms were visualized using the WINPLOTR program of the FULLPROF package (13). ¹H-NMR resonances of the busulfan, -cyclodextrin, and busulfan–cyclodextrin systems were recorded on a Bruker DRX 400-MHz spectrometer (Wissenbourg, France).

Table I. Single-Crystal Experiment and Refinement Details

Chemical formula	C ₆ H ₁₄ O ₆ S ₂
Formula weight (g mol ⁻¹)	246.3
Temperature (K)	100.0(1)
Wavelength (Å)	0.71073
Crystal system	Triclinic
Space group	$\bar{1}$
a (Å)	5.5246(1)
b (Å)	6.7386(1)
c (Å)	7.4728(1)

(°)	68.788(1)
(°)	89.297(1)
(°)	75.856(1)
Volume (Å ³)	250.96(1)
Z	1
Density (calculated) (Mg/m ³)	1.63
Absorption coefficient (mm ⁻¹)	0.550
Crystal size (mm ³)	0.70 × 0.55 × 0.40
Color	Colorless
[sin/] _{max} (Å ⁻¹)	1.225
Total number of reflections	46,878
Unique reflections [I > 3 (I)]	5421
R _{int}	0.0244
Multipole refinement	
R ₁ (F) (%)	1.58
wR ₁ (F) (%)	1.85
gof(F)	0.87
Kappa refinement	
R ₁ (F) (%)	2.16
wR ₁ (F) (%)	2.36
gof(F)	1.09

$$R_{\text{int}} = \frac{1}{N_H} \sum_{H=1}^{N_H} \sqrt{\frac{1}{N_H} \sum_{i=1}^{N_H} |I_i(H) - \bar{I}_i(H)|^2}$$

I(H) is the observed reflection intensity at the Bragg vector H. N_H is the number of equivalent and redundant reflections.

$$R_1 = \frac{\sum_w |F_{\text{obs}} - F_{\text{calc}}|}{\sum_w F_{\text{obs}}} = \frac{\sum_w |F_{\text{obs}} - F_{\text{calc}}|^2}{\sum_w F_{\text{obs}}^2} \cdot \frac{1}{2} \cdot \frac{1}{n_p}$$

F_{obs} and F_{calc} are the observed and calculated structure factors, respectively. The statistical weight is w = 1/(F_{obs}² g of (F)), where ² is the variance and m_{obs} and n_p are the number of observations and refined parameters, respectively.

Structure and Electron Density Refinements

The crystal structure was solved by direct methods (14) and first refined with the SHELXL-97 program of the WINGX software package (15). The obtained parameters were then transferred to the input files of the MOLLY program (16) based on the Hansen–Coppens electron density model, where the total electron density ρ_{model}(r) of the molecule is calculated as the sum of atomic contributions

$$\rho_{\text{at}}(\mathbf{r}) = \rho_{\text{core}}(\mathbf{r}) + P_{\text{val}} \rho_{\text{val}}(\mathbf{r}) + \sum_{l=0}^3 R_l r^{n_l} \sum_{m=0}^l P_{lm\pm} y_{lm\pm}, \quad (1)$$

In Eq. (1), ρ_{core} and ρ_{val} are Hartree–Fock free atomic core and valence spherical electron densities (17), respectively. The third part in Eq. (1) is the aspherical atomic electron density described by the real spherical harmonic angular functions y_{lm±}(,) and modulated by a radial Slater-type function

$$R_l r = \frac{1}{n_l + 2!} r^{n_l} e^{-r} \quad (2)$$

The least-squares fitted parameters are the valence population P_{val} yielding the atomic charge (q = P_{val} - N_{val}; N_{val} is the valence population of a free atom) and the multipole populations P_{lm±} characterizing the electron deformation density. The contraction (, > 1) or expansion (, < 1) coefficients of the atomic electron density are also refined.

In this study, the chosen radial function arguments in Eq. (2) were 3.95, 4.50, 3.00, and 2.26 bohr⁻¹ (1 bohr = 0.52918 Å) for S, O, C, and H, respectively (18). The best refinement results were obtained for exponents n₁ 3, 3, 4, 4 (up to hexadecapole level) for S, n₁ 2, 2, 3 (up to octupole level)

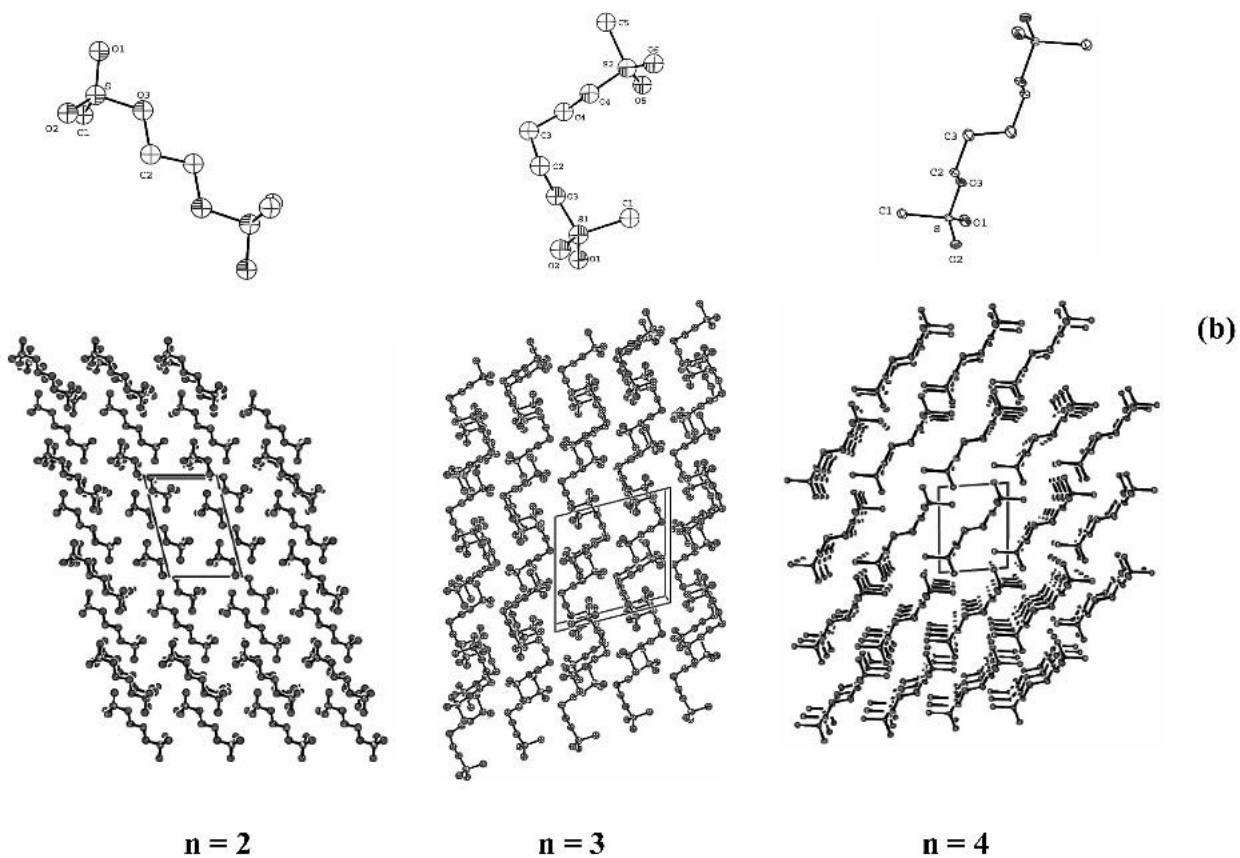
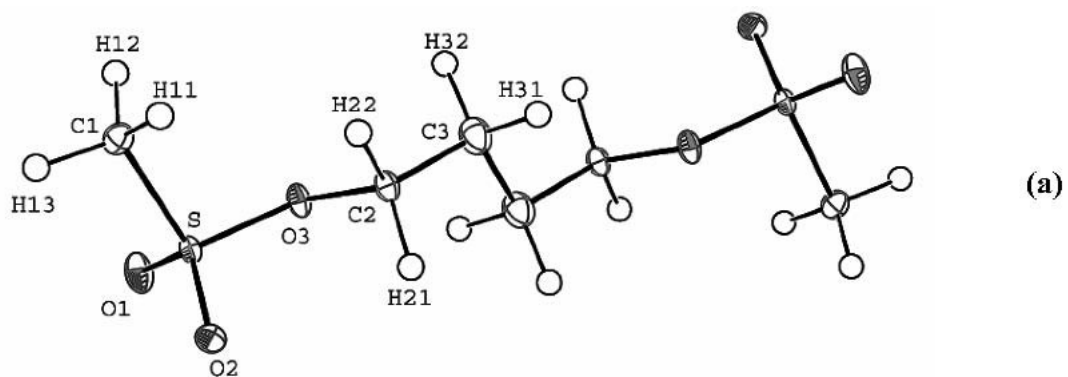


Fig. 1. (a) ORTEP diagram of the busulfan molecule with 50% probability thermal ellipsoids. (b) View of the busulfan $\text{CH}_3\text{-SO}_2\text{-O}(\text{CH}_2)_n\text{-SO}_2\text{-CH}_3$ analogs [$n = 2, 3$ from MacKenna *et al.* (6) and $n = 4$ from this study] crystal packing perpendicular to **b** crystallographic axis. Hydrogen atoms are omitted for clarity.

for C and O, and $n = 1$ (up to dipole level) for H. Charge and dipole parameter constraints were imposed to chemically equivalent hydrogen atoms. After the multipole refinements, the atomic net charges were estimated through a kappa refinement (19) based on the spherical part in Eq. (1). Figure 1a shows the ORTEP diagram (20) displaying the atomic thermal vibrations of the busulfan molecule. The refinement statistical factors are given in Table I.

STATDENS program (21) was used to calculate and plot the electron deformation density as $\rho_{\text{def}} = \rho_{\text{model}} - \rho_{\text{free}}$ where ρ_{free} is the superposition of the free atom model electron densities. To gain further insight into the atomic bond properties, the ‘‘Atoms in Molecules’’ theory developed by Bader (7) establishes a direct relation between the total electron density distribution of a chemical system and the energetic properties of the bonds. The topological analysis was performed through the gradient $\nabla \rho(\mathbf{r})$ and the Laplacian $\nabla^2 \rho(\mathbf{r})$ of the electron density. The zeroes of $\nabla \rho(\mathbf{r})$ localize the extrema at \mathbf{r}_c called bond critical points (BCPs). $\nabla^2 \rho(\mathbf{r})$ is also related to the potential $V(\mathbf{r})$ (negative) and kinetic $G(\mathbf{r})$ (positive) energy densities

$$\nabla^2 \rho(\mathbf{r}) = -2G(\mathbf{r}) + 2V(\mathbf{r}) \quad (3)$$

Therefore, $\nabla^2 \rho(\mathbf{r}_c) < 0$ indicates a concentration of electrons at \mathbf{r}_c [$\rho(\mathbf{r}_c)$ dominates] and $\nabla^2 \rho(\mathbf{r}_c) > 0$ corresponds to a depletion of electrons at \mathbf{r}_c [$G(\mathbf{r}_c)$ dominates]. The NEWPROP (22) program was used to analyze the topological features of the experimental electron density of the busulfan. The bond energy density $H(\mathbf{r}_c) = G(\mathbf{r}_c) + V(\mathbf{r}_c)$ can also be quantified from the values of $\nabla^2 \rho(\mathbf{r}_c)$ [Eq. (3)] at \mathbf{r}_c using the empirical formula of the kinetic energy density (23)

$$G(\mathbf{r}_c) = \frac{3}{10} \sqrt{\frac{1}{\rho(\mathbf{r}_c)}} \left[\nabla^2 \rho(\mathbf{r}_c) + 1 \right] \quad (4)$$

6

Electrostatic Properties

The molecular electrostatic potential $V(\mathbf{r})$ was calculated using the ELECTROS program (21) as

$$V(\mathbf{r}) = \sum_{\text{at}} \frac{Z_{\text{at}}}{|\mathbf{r} - \mathbf{R}_{\text{at}}|} - \int \frac{\rho(\mathbf{r}')}{|\mathbf{r} - \mathbf{r}'|} d\mathbf{r}' \quad (5)$$

the sum of the contributions of the positive nuclear charge Z_{at} and the electron density $\rho(\mathbf{r})$ of the pseudoatom at \mathbf{R}_{at} [Eq. (1)]. We also have used the electrostatic potential to derive the dipole moment of the methylsulfonate fragment of the busulfan molecule (24).

RESULTS AND DISCUSSION

Molecular Conformation and Crystal Packing

The busulfan crystallizes in P1 triclinic space group with one molecule in the unit cell. Figure 1a shows the crystal conformation of the busulfan molecule where the inversion center is in the middle of the C₃–C₃ bond. The bond lengths and angles obtained after the multipole refinement are in good agreement with those reported in the previous study (6). The sulfur atom of the terminal methylsulfonate (CH₃SO₃) fragment is in a distorted tetrahedral coordination with two short SO bond distances for SO₁ and SO₂ [1.4276(3) and 1.4097(3) Å, respectively]. Consequently, the oxygen electron repulsion in the SO₂ part of the methylsulfonate fragment widens the O₁–S–O₂ bond angle [119.39(3)°]. In contrast, O₃ is singly bonded to S [S–O₃ 1.5847(2) Å] and to C₂ [C₂–O₃ 1.4671(4) Å]. C₁ methyl carbon is at 1.7354(3) Å from S. All these SO, S–O, S–C, and C–O bond distance values are in relatively good agreement with the average estimates retrieved from the Cambridge Structural Data Center (CSDC) (25) for similar fragments: <SO> 1.412, <S–O> 1.571, <S–C> 1.732, and <C–O> 1.462 Å, respectively. In the carbon chain of busulfan, the C₂–C₃ bond length (1.5253(4) Å) is similar to that of the central C₃–C₃ link [1.5291(5) Å]. The *gauche*-, *trans*-, *gauche*- crystal conformation (6) maintains the two O₃ ester oxygen atoms at a distance of 5.1956(5) Å. On the other hand, Fig. 1b compares the crystal packing of busulfan analogs

[CH₃–SO₂–O(CH₂)_n–O–SO₂–CH₃, n = 2, 3, 4]. It is worth noting that the crystal packing looks very similar for n = 2 and 4 for which the molecules are stretched, whereas for n = 3, the central carbon permits the molecule to be bent. Given the nature of these molecules, only weak intermolecular O···H(C) interactions are observed (>2.37 Å). Therefore, the crystal packing of these three molecules mainly originates from electrostatic and van der Waals interactions, particularly those involving the methylsulfonate fragment with a pronounced polar character. Indeed, the inspection of S···S distances reveals short contacts: two S···S 4.671 (n = 2), S···S 4.237 and 4.546 (n = 3), S···S 4.119 (n = 4 at room temperature) and 4.0729(1) Å (n = 4 at 100 K), respectively. This latter emphasizes that the methylsulfonate fragment interactions are enhanced in the busulfan crystal lattice. Furthermore, the intermolecular torsion angles for the closest pairs of methylsulfonate fragments are C–S···S–C 48.2° (for the two S···S contacts for n = 2), 180.0° (n = 3), and 180.0° for the busulfan (n = 4). Except for n = 2, the C–S axes of the methylsulfonate fragments are antiparallel in order to minimize the dipole–dipole interactions.

Electron Deformation Density

The static electron deformation density in the chemical bonds of busulfan is depicted in Fig. 2. Figure 2a shows the electron density distribution in O₁SO₂ part of the methylsulfonate fragment. Electron density peak heights of 0.70 eÅ^{−3} and 0.55 eÅ^{−3} are found for SO₁ and SO₂ bonds, respectively. Figure 2a also shows that the oxygen lone pairs are not symmetric. On the contrary, theoretical calculations on the isolated H₃NSO₃ molecule show symmetrical density in the SO bond and the lone pairs (26). In the C₁–S–O₃ plane of Fig. 2b, the

electron density peak heights are $0.50 \text{ e}\text{\AA}^{-3}$ in both $\text{S}-\text{O}_3$ and $\text{S}-\text{C}_1$ bonds. Figure 2c gives a global view of the electron deformation density in the busulfan. The inversion center connecting C_3 to C_3' carbon atoms (with electron density peak height of $0.35 \text{ e}\text{\AA}^{-3}$) is in the middle of the map in Fig. 2c.

Bond Topological Properties

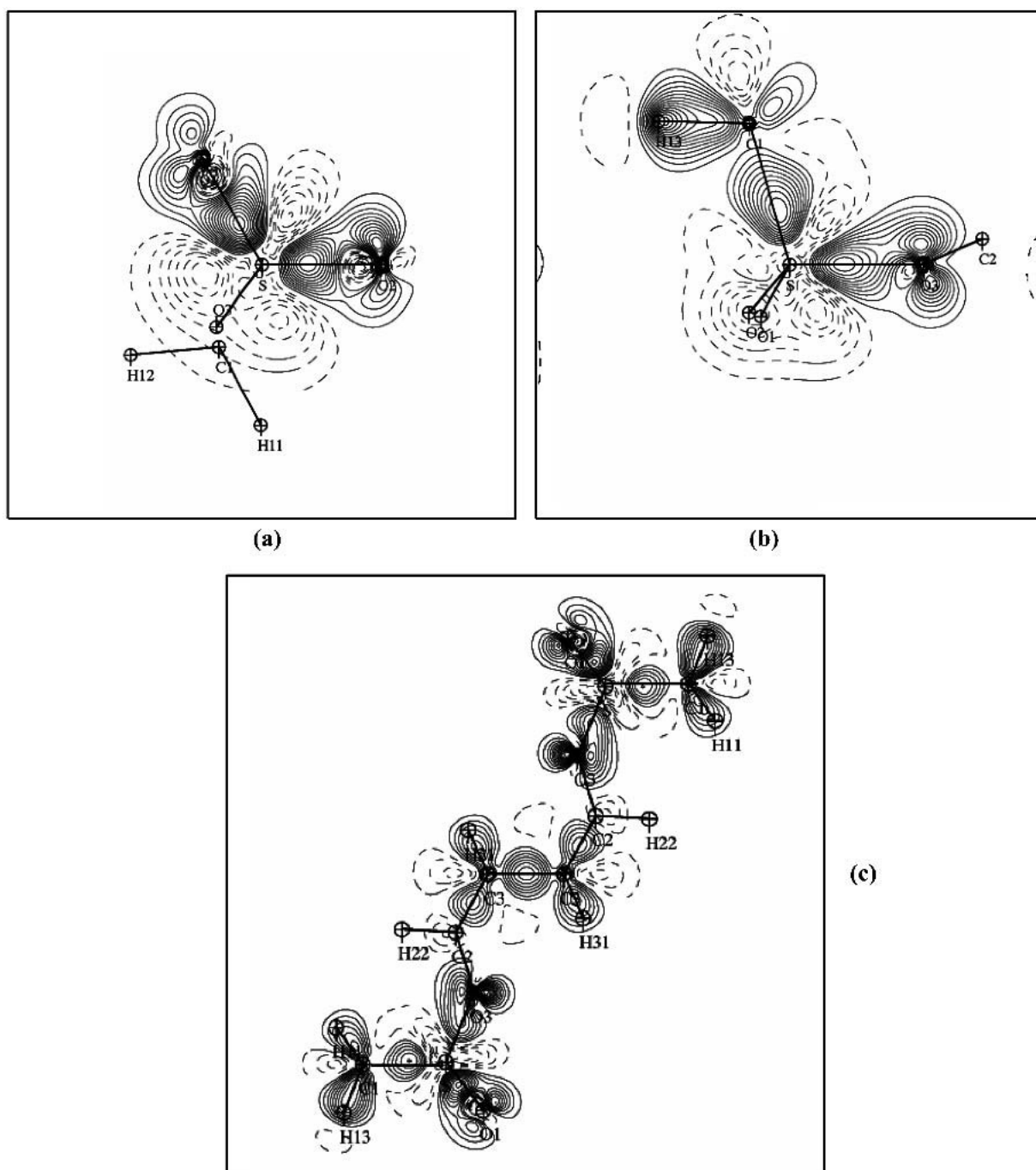


Fig. 2. Static electron deformation density maps: (a) O₁–S–O₂; (b) C₁–S–O₃; (c) global view centered on C₃–C₃ bond. Contours interval is ± 0.05 e \AA^{-3} ; negative contours are dashed; zero contour omitted.

Table II lists the topological descriptors of the busulfan atomic bonds. All ρ^2 values have a negative sign at the bond critical points (BCPs). This means from Eq. (3) that the bonds in busulfan have a covalent character (shared shell interactions) where the potential energy dominates. In the methylsulfonate fragment of busulfan, the double bond character of SO₁ and SO₂ is clearly revealed by large values of ρ^2 at BCP (2.52 and 2.43 e \AA^{-3}), ρ^2 (–21.97 and –16.53 e \AA^{-5}) compared to those of the single bond S–O₃ (ρ^2 1.90 e \AA^{-3} , ρ^2 –13.72 e \AA^{-5}). A much lower electron density is found in S–C₁ bond (ρ^2 1.43 e \AA^{-3}) with an intermediate magnitude of Laplacian (ρ^2 –8.18 e \AA^{-5}). The C₂–O₃ bond has the lowest value of the Laplacian (ρ^2 –1.78 e \AA^{-5}) and a BCP electron density 1.47 e \AA^{-3} . In the carbon chain, the topological parameters given in Table II are in good agreement with those reported for organic molecules like peptides (27). On the other hand, the highest moduli of H [Eq. (4)] is found for the SO double bonds ($|H|$ 1656.47 for SO₁ and 1521.45 kJ·mol^{–1} per atomic unit volume for SO₂). This is followed by the C–H bond energy density modulus in the range of 909.50 to 1032.80 kJ·mol^{–1} per atomic unit volume.

However, half of these values are found for S–C₁ and C₂–O₃. This latter has the lowest modulus ($|H|$ 609.63 kJ·mol^{–1} per atomic unit volume) and then quantifies the alkylating property of the busulfan showing that the C₂–O₃ bond can easily be cleaved.

Net Atomic Charges

The kappa-refinement (19) was carried out in order to estimate the net atomic charges (Table III) in e units (1 e 1.602 10^{–19} C). Except for hydrogen atoms that are highly

Table II. Topological Properties at the Bond Critical Points (BCPs) of Busulfan*

Bond	d ₁	d ₂	ρ^2	ρ	H
SO ₂	0.662	0.768	–21.97	2.52	–1656.47
SO ₂	0.641	0.770	–16.53	2.43	–1521.45
S–O ₃	0.764	0.821	–13.72	1.90	–1034.58
C ₁ –S	0.731	1.005	–8.18	1.43	–640.97
C ₂ –O ₃	0.620	0.848	–1.78	1.47	–609.63
C ₃ –C ₂	0.739	0.787	–11.82	1.67	–841.23
C ₃ –C ₃	0.765	0.765	–9.94	1.53	–724.52
H ₁₁ –C ₁	0.377	0.671	–15.71	1.85	–1013.05
H ₁₂ –C ₁	0.384	0.667	–15.28	1.81	–978.00
H ₁₃ –C ₁	0.383	0.689	–14.38	1.79	–954.44
H ₂₁ –C ₂	0.363	0.678	–17.89	1.85	–1032.80
H ₂₂ –C ₂	0.372	0.708	–14.98	1.75	–929.41
H ₃₁ –C ₃	0.359	0.716	–15.27	1.72	–909.50
H ₃₂ –C ₃	0.360	0.701	–15.65	1.73	–920.43

*Laplacian ρ^2 (in e \AA^{-3}); total electron density ρ (in e \AA^{-3}); d₁ and d₂ are the BCP-atom distances (in Å). The last column gives the total energy density H (in kJ mol^{–1} per atomic unit volume).

contracted (ρ 1.3 to 1.4), the kappa parameters for other atoms are very close to 1.0. The sulfur [+0.93(8) e] and the methyl carbon C₁ [–0.8(1) e] atoms carry the highest charges in modulus but with opposite signs. C₁ negative charge is compensated by 3 times +0.28(4) e corresponding to the three attached protons H₁₁, H₁₂, and H₁₃. The double SO and single S–O bonds are characterized by different charges of the oxygen atoms O₁, O₂, and O₃, which are –0.54(5),

–0.44(5), and –0.27(5) e, respectively. In the carbon chain, C₂ is almost neutral [ρ –0.1(1) e], whereas the central C₃ carries –0.35(8) e. The agreement with Mulliken charges is still confirmed here as the theoretical values, reported for the methylsulfonate –1 anion, are +1.16, –0.65, –0.67, and +0.15 e for S, O, C, and H, respectively (28). Comparatively, in the busulfan molecule, the methylsulfonate fragment has a global charge of –0.28 e. In consequence, the carbon chain in busulfan is positively charged (+0.56 e). It is then expected that the removal of the methylsulfonate groups during the alkylating process will liberate a high electrophilic radical.

Electrostatic Potential

Experimentally fitted multipole parameters and net atomic charges were used to calculate the electrostatic poten-

Table III. Kappa Parameters (ρ) and Net Atomic Charges (q) of Busulfan (in e Unit)

	ρ	q
S	1.043(6)	+0.93(8)
O ₁	0.973(4)	–0.54(5)
O ₂	0.985(4)	–0.44(5)
O ₃	0.994(4)	–0.27(5)
C ₁	0.988(8)	–0.8(1)
C ₂	1.009(9)	–0.1(1)

C ₃	0.999(8)	-0.35(8)
H ₁₁ H ₁₂ H ₁₃	1.41(5)	+0.28(4)
H ₂₁ H ₂₂	1.29(4)	+0.14(4)
H ₃₁ H ₃₂	1.30(4)	+0.21(4)

Standard deviations are given in parentheses.

tial generated around an isolated busulfan molecule. Figure 3 depicts the electrostatic potential isosurfaces corresponding to $\pm 0.05 \text{ e}\text{\AA}^{-1}$ ($1 \text{ e}\text{\AA}^{-1} = 14.4 \text{ V}$ and energetically $1 \text{ e}^2\text{\AA}^{-1} = 332.4 \text{ kcal}\cdot\text{mol}^{-1} = 1389.4 \text{ kJ}\cdot\text{mol}^{-1}$). The minimum electrostatic potential value of $-0.14 \text{ e}\text{\AA}^{-1}$ ($-0.18 \text{ e}\text{\AA}^{-1}$ from kappa refinement) is found in the vicinity of O₁ oxygen atom compared to $-0.10 \text{ e}\text{\AA}^{-1}$ (from the two refinements) for O₂. These holes of attractive electrostatic potential for positive species like protons are at a distance shorter than 1.5 Å from the O₁ and O₂ nuclei. The nucleophilic (negative electrostatic potential) regions, similar from the two refinements, have a kidney shape and surround the O₁ and O₂ oxygen sites of the methylsulfonate moiety. The negative potential does not extend to O₃, which is completely inhibited by the strong electrophilicity (positive electrostatic potential) of the carbon chain and methyl groups of busulfan.

Methylsulfonate Fragment Dipole Moment and Crystallization Impacts

Figure 3 also shows the polar character of the terminal methylsulfonate fragments of busulfan. Positive and negative barycenters are located on the methyl and SO₃ groups, respectively. We used the electrostatic potential to calculate the dipole moment of the methylsulfonate fragment using the method we developed earlier (24). The obtained values are 3.23(3) and 4.69(2) D (1 Debye = $3.336 \cdot 10^{-30} \text{ C}\cdot\text{m}$) from the multipole and kappa refinements, respectively. The kappa model slightly overestimates the value of the fragment dipole moment, which is very sensitive to the multipole parameters as pointed out by Abramov *et al.* (29). Comparatively, the estimated dipole moments of the H₂O and the DMSO molecules are 2.2 D and 3.96 D, respectively (24,30). The high value of the methylsulfonate fragment dipole moment explains the antiparallel interactions and the crystal stacking of busulfan molecules. Also, the noncrystallization in polar DMSO or DMAC solutions can be explained by the competition of the dipole interactions between busulfan–busulfan and busulfan–solvent couples.

Busulfan–Cyclodextrin Systems

The use of cyclodextrins has been considered to try to obtain an inclusion complex with busulfan. The interior of the cyclodextrin cavity is hydrophobic, whereas the external surface is hydrophilic (9). In the inclusion complexes, both steric and electrostatic considerations must be taken into account. The inspection of the large extent of the negative electrostatic potential (Fig. 3) suggests the use of the -cyclodextrin with an intermediate cavity diameter (6 Å) in our preparation of coprecipitates with busulfan. The cyclodextrins have cyclic conformations resembling, to some extent, the crownmolecules. The Structural Cambridge Data Bank contains some crystallographic structures of clathrate complexes in presence of CH₃–SO₂ or methylsulfonate fragments (25). The reported structures reveal that the methyl group is more involved in the interaction with the host molecule than the sulfonate part. The complex is stabilized by weak interactions between the oxygen atoms of the host molecule and the methyl groups pointing their protons (electrophilic part) toward the molecular cavity.

Figure 4 shows the powder X-ray diffraction patterns

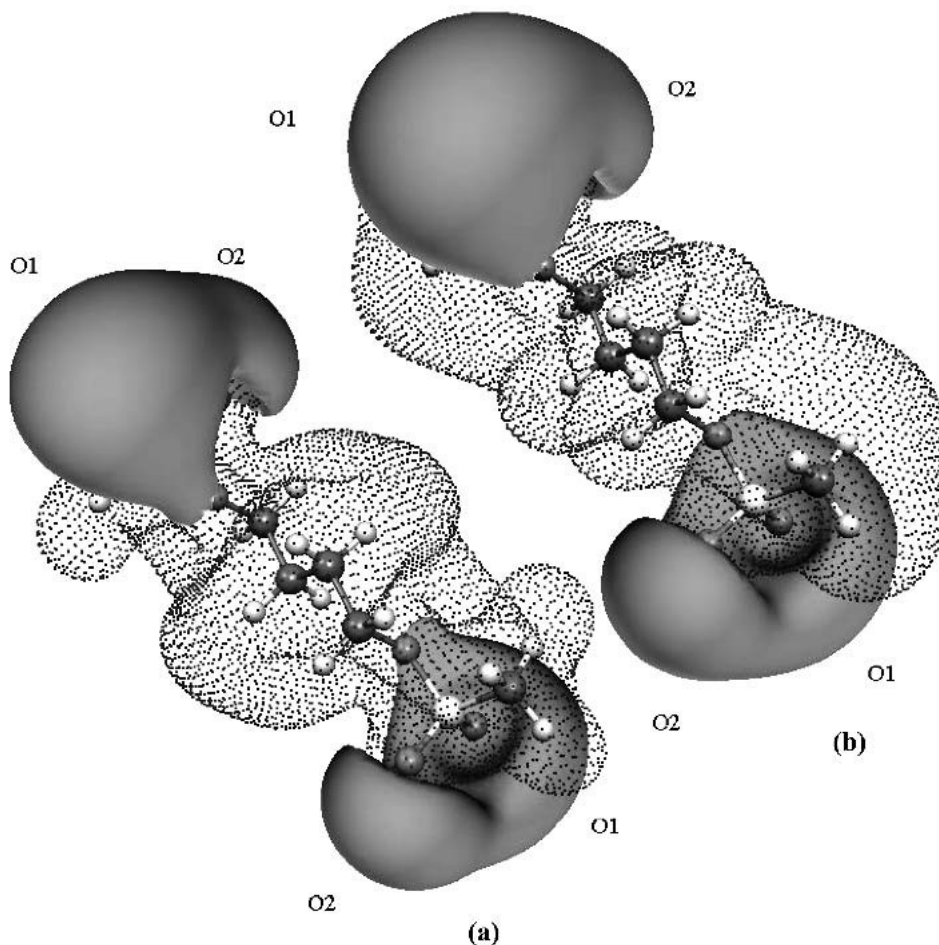


Fig. 3. Electrostatic potential isosurfaces around the busulfan molecule: (a) from multipole refinement; (b) from kappa refinement. The positive electrostatic potential isosurface ($+0.05 \text{ e}\text{\AA}^{-1}$) is in dots and the negative one ($-0.05 \text{ e}\text{\AA}^{-1}$) is smoothed.

obtained for the different compounds used in this study. Figures 4a and 4b correspond to the starting products of the pure busulfan and -cyclodextrin hydrate, respectively. The association in the solid state of the busulfan and -cyclodextrin hydrate with initial 1:1 and 1:2 molar ratios yield crystalline phase compounds with a small part of amorphization characterized by a humped background as shown in Figs. 4c and 4d. The view of these patterns suggests two new phases with numerous peaks indicating a quite different behavior of the busulfan–cyclodextrin interactions with respect to the chosen stoichiometries. These new phases are clearly dissimilar from the single physical mixture (not shown here), corresponding to an exact superposition of the patterns in Fig. 4a and 4b. 1:1 and 1:2 busulfan/-cyclodextrin diffractograms also differ from that of pure -cyclodextrin hydrate depicted in Fig. 4b. The d-spacing abscissa (\AA) is preferentially used in order to well separate the peaks corresponding to the long-range crystal packing parameters. In the -cyclodextrin hydrate diffractogram, the peaks at d 10.0, 14.3, and 20.4 \AA can be chosen as reference. The latter value is in good agreement with the highest cell parameter [20.857(6) \AA] determined from a single-crystal study of the -cyclodextrin hydrate (31). Following the three reference peaks, Figs. 4e and 4f clearly show differences in the crystal lattice of -cyclodextrin hydrate and those of 1:1 and 1:2 compounds, which suggests that a true association of busulfan and -cyclodextrin exists. This is much more visible for 1:1 compound presenting the most distinguished peak position shifts. However, it is premature at this stage of investigation to identify the type of crystal packing of such solid complexes and the respective arrangements of the host and guest molecules. 1

H-NMR was used to achieve the analysis of the busulfan–cyclodextrin coprecipitates in terms of complexation but also in terms of stability in aqueous medium. The chemical shifts of free -cyclodextrin, pure busulfan, and 1:1 and 1:2 busulfan–cyclodextrin systems are given in Table IV. Owing to the insolubility of the drug in water, we cannot consider the induced chemical shifts for busulfan. However, in the presence of busulfan in the 1:1 molar ratio system, the H_5 (-0.015 ppm) and H_6 (-0.019 ppm) protons of the -cyclodextrin display the most important induced shifts whereas of H_3 is found equal to -0.005 ppm . The upfield shifts of H_5 and H_6 protons, respectively located on the inner surface and the rim of the cyclodextrin, suggest a shallow position of the busulfan that interacts with the hydroxyl groups of the narrow side of the cavity. This hypothesis can also be supported in light of the observed positive electrostatic potential region surrounding the methyl group and the carbon chain of the busulfan molecule (Fig. 3). The busulfan molecule likely exposes this electrophilic region toward the

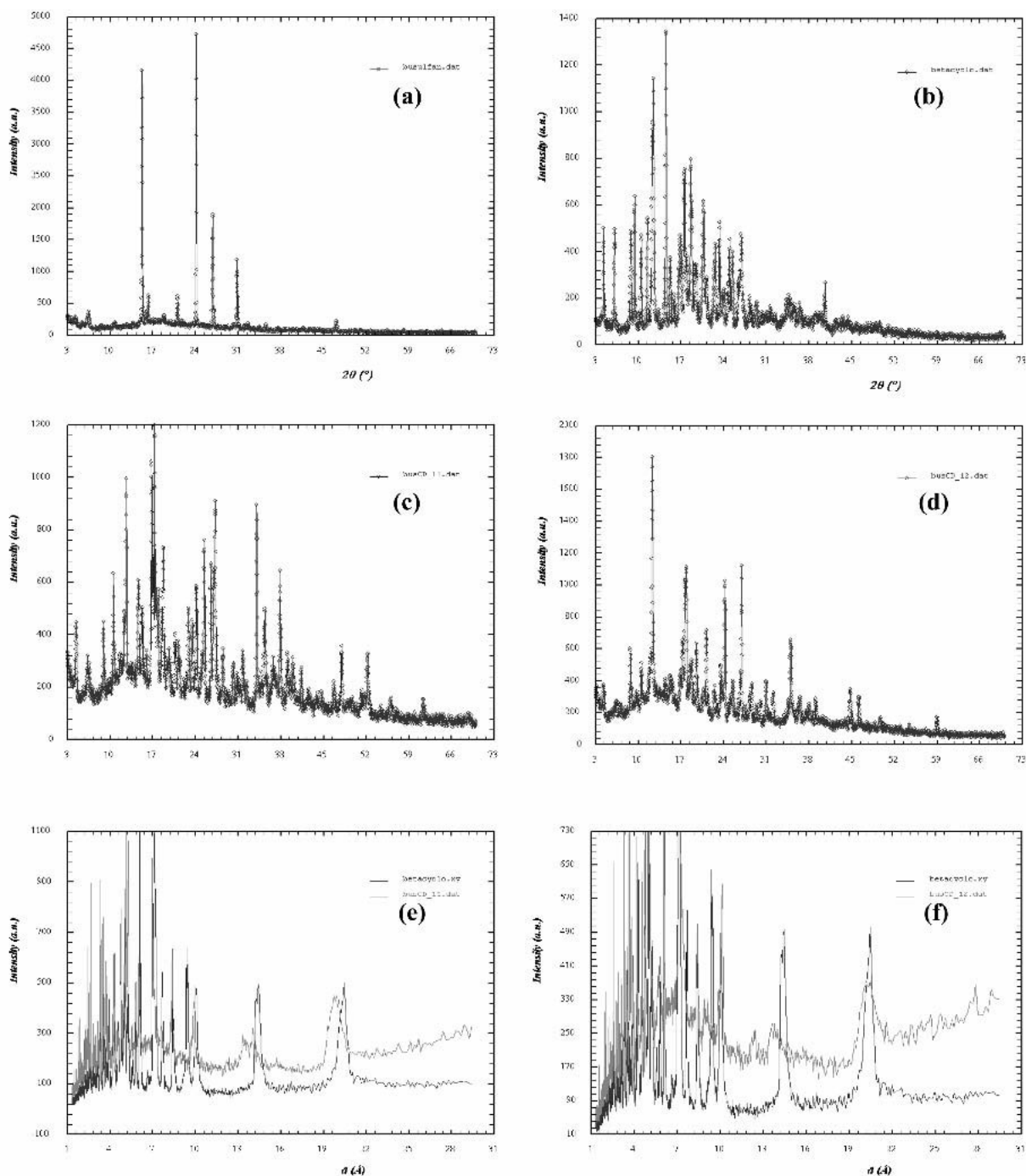


Fig. 4. Powder diffraction patterns: (a) busulfan, (b) -cyclodextrin hydrate, (c) busulfan–cyclodextrin 1:1, (d) busulfan–cyclodextrin 1:2, (e) busulfan–cyclodextrin 1:1 (upper curve) and -cyclodextrin hydrate (lower curve) with d-spacing abscissa (Å), and (f) busulfan–cyclodextrin 1:2 (upper curve) and -cyclodextrin hydrate (lower curve) with d-spacing abscissa (Å). Diffraction intensities are in arbitrary units.

cavity aperture of the cyclodextrin. We note that the crystal conformation of the busulfan depicted in Fig. 1a and the real one in the complex are probably different; the exposed electrophilic part can be extended as consequence of a possible torsion of the carbon chain. On the other hand, no significant shifts are observed for the cyclodextrin protons for the 1:2 coprecipitate in solution. The slight chemical shifts of the cyclodextrin protons in presence of busulfan in the 1:2 system cannot confirm any association in this case. The $^1\text{H-NMR}$ results are thus in good agreement with those observed in the solid state through X-ray powder patterns showing the difference between the 1:1 and 1:2 systems.

CONCLUSIONS

This work provides a characterization of the busulfan electrostatic properties derived from an accurate X-ray high-resolution diffraction experiment. The chemical bond strengths and the alkylating property of busulfan are quantified through the topological electron density descriptors. The

Table IV. ¹H-NMR Chemical Shifts of Busulfan (in CDCl₃), -Cyclodextrin (in D₂O), and Busulfan–Cyclodextrin 1:1 and 1:2 Compounds (in D₂O)

	Free	(ppm)																																																
		1:1	1:2																																															
Busulfan																																																		
H	4.295	—	—	–OCH ₂	Triplet																																													
H	3.048	2.807	2.809	–CH ₃	Singlet																																													
H	1.922	1.910	1.885	–CH ₂ CH ₂	Multiplet																																													
complex – free																																																		
<table border="1" style="margin-left: auto; margin-right: auto;"> <thead> <tr> <th></th> <th>1:1</th> <th>1:2</th> </tr> </thead> <tbody> <tr> <td colspan="6">-Cyclodextrin</td> </tr> <tr> <td>H₁</td> <td>5.060</td> <td>5.059</td> <td>5.062</td> <td>–0.001</td> <td>0.002</td> </tr> <tr> <td>H₂</td> <td>3.585</td> <td>3.590</td> <td>3.590</td> <td>0.005</td> <td>0.005</td> </tr> <tr> <td>H₃</td> <td>3.920</td> <td>3.915</td> <td>3.919</td> <td>–0.005</td> <td>–0.001</td> </tr> <tr> <td>H₄</td> <td>3.562</td> <td>3.566</td> <td>3.566</td> <td>0.004</td> <td>0.004</td> </tr> <tr> <td>H₅</td> <td>3.794</td> <td>3.779</td> <td>3.791</td> <td>–0.015</td> <td>–0.003</td> </tr> <tr> <td>H₆</td> <td>3.823</td> <td>3.804</td> <td>3.822</td> <td>–0.019</td> <td>–0.001</td> </tr> </tbody> </table>							1:1	1:2	-Cyclodextrin						H ₁	5.060	5.059	5.062	–0.001	0.002	H ₂	3.585	3.590	3.590	0.005	0.005	H ₃	3.920	3.915	3.919	–0.005	–0.001	H ₄	3.562	3.566	3.566	0.004	0.004	H ₅	3.794	3.779	3.791	–0.015	–0.003	H ₆	3.823	3.804	3.822	–0.019	–0.001
	1:1	1:2																																																
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refined atomic multipole parameters and net charges are used to generate the electrostatic potential of busulfan emphasizing the polar character of the terminal methylsulfonate groups. This latter property governs the arrangement of molecules in the solid state and is likely responsible for the strong crystallization of busulfan. In this study, we suggest overcoming this crystallization obstacle for liquid formulations into nontoxic solvents by trying to prepare inclusion complexes of busulfan and cyclodextrins. Powder diffraction and ¹H-NMR analyses show that native -cyclodextrins can associate with busulfan in the coprecipitation process described in this paper.

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