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Corrosion Behavior in a Neutral Artificial Saliva of Several Binary Co-Cr Alloys with Various Chromium Contents

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Abstract

Substitution of too expensive alloys rich in noble metals by alloys based on nickel or chromium is often considered in prosthetic dentistry. When containing chromium in quantities high enough the corrosion performances of the “predominantly base” alloys very interesting, as well as their mechanical properties. Such alloys are generally rather complex because of the presence of other elements such as molybdenum, tungsten, silicon and other, which may be active during corrosion and mask the importance of the chromium content. In this work it was chosen to look for a critical chromium content over which the corrosion behaviour is acceptable or very good, by working with simple binary Co-xCr alloys (x varying from 0 to 30 wt.% by slices of 5 wt.%) especially prepared by casting, by avoiding any interference of other elements classically present (Mo, W, Si...). Electrochemical investigations were carried out in a neutral 37°C-heated 9 g/L NaCl solution, classically used in such studies. Follow-up of open circuit potential, application of the Stern-Geary and Tafel methods, and cyclic polarizations allowed identifying 25 wt.% as being a threshold to respect for being certain to benefit from the highest corrosion resistance.

Keywords: Cast dental alloys; Cobalt-chromium; Neutral artificial saliva; Corrosion; Electrochemical experiments

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Introduction

Prostheses for dental restoration such as crowns or fixed partial dentures contain metallic pieces to allow the prostheses resisting intense stresses. In the case of fixed partial dentures the reinforcing metallic internal parts are named frameworks. These ones are constituted of parts made of a “parent alloy”, which are joined together by soldering [1]. The parent alloys can be of different categories, which are named in function of the presence or not of noble metals in their chemical composition (Au, Pt, Pd, Ag...). The alloys containing noble metals are the “High Noble” ones (>60 wt.% of noble metals) or the “Noble” ones (>25 wt.%). Other parent alloys, must cheaper, do not contain noble metals: the “Predominantly Base” alloys. Many of them are based on nickel and contain significant quantities of chromium for resisting corrosion in clinic situation. Unfortunately they may potentially induce allergic phenomena [2]. Other “Predominantly Base” alloys also exist with not so serious problems: some alloys based on cobalt [3].

Beside gold-cobalt [4] and titanium-cobalt [5] dental alloys, cobalt-chromium alloys are the most popular cobalt-based dental alloys. Molybdenum is often also present in their chemical composition [6], to exploit its strengthening effect by solid solution hardening. The roles of Cr (resistance to aqueous corrosion) and Mo (mechanical resistance) can be compared to the ones of these same elements at elevated temperatures ($\approx 1000^\circ\text{C}$) in some cobalt-based super alloys [7] which present similar chemical compositions: resistance to high temperature oxidation and hot corrosion [8] and resistance to high temperature creep deformation [9].

Co-Cr-Mo dental alloys can be synthesized by casting [10] or by powder metallurgy [11] for example. Significant contents of other elements such as W [12] or Ti [13] can be also present. The (Co, Cr)-based alloys are perhaps not so used in prosthetic dentistry as the (Ni, Cr)-based ones but their microstructures [14], mechanical properties [15] and corrosion behaviours [16] were subjected to characterizations and studies in several works.

Many pended patents concern cobalt-based alloys for medical applications of various types, not only for dental but also surgical (orthopedic prostheses, stents...) [17]. Several elaboration and shaping routes are concerned by these patents, not only foundry [18]: subtracting machining [19], forging [20], soldering alloys [21], and coating alloys [22]. Various types of composition are covered: Co-Cr-Mo [23], Co-Cr-Mo-Ti [24], Co-Cr-Mo-W [25], Co-Cr-Ni-(Mo, Ti, Ta) [26], Co-Cr-W-Nb [27] and even Co-Ru-Cr [28]. Other alloys protected by patents contain additionally noble metals [29] or simply carbon for obtaining high strength [30]. These examples illustrate the high number of patents which were pended concerning cobalt-chromium alloys for various medical uses. Curiously published articles about these cobalt alloys are rather rare, notably concerning corrosion studies.

All these dental alloys based on cobalt are logically rich in this later element, and the second most present element is chromium. Indeed, the Cr content is the most often close to 30 wt.%, which can induce some problems when foundry is the chosen way for synthesizing the Co-based dental alloys. High content in this rather refractory element (melting point equal to around 1890°C, much higher the Co one: 1490°C [31]) may induce some difficulties to totally melt the composition and not molten particles of Cr can be still present inside the alloys. Despite much higher melting points tungsten and molybdenum generally cause fewer problems since they are present in lower quantities in cobalt-based dental alloys. Some rare compositions found in literature involve chromium contents lower than these 30 wt.%, such as 20 wt.% and even 15 wt.%. Melting such alloys ought to be easier than for Cr-rich alloys but one can be afraid that the very good resistance against corrosion of such alloys in the buccal milieu will be threatened and even lost. For efficiently monitoring the chromium content in the composition of dental alloys it can be useful to study the corrosion of cobalt-based alloys in artificial saliva versus their chromium content. This is what was undertaken here, in a very simple case, binary Co-Cr cast alloys. Their chromium content was successively lowered by slice of 5 wt.%, from 30 wt.% down to zero. These seven alloys were elaborated by high frequency induction melting under inert atmosphere, then tested in corrosion for conditions close to the buccal ones. This was done in a simple 9 g/L NaCl aqueous solution (classically used artificial saliva, e.g. [32], maintained at 37°C)- using classical electrochemical techniques of corrosion characterization: three-electrode cell and potentiostat driven by a software allowing performing the usual stationary electrochemical methods (free potential, Stern-Geary, Tafel, cyclic polarization).

Materials and Methods

The experimental details of this study may be resumed as follows:

Elaboration of the alloys

Weighing of the pure cobalt and nickel (99.9, Alfa Aesar), for around 10 grams of alloy for each composition; targeted Cr contents: 30, 25, 20, 15, 10, 5 and 0 wt.% (**Figure 1**). Placing in the cold crucible of the high frequency induction furnace (CELES

and positioning of a silica tube for isolating the melting chamber from laboratory air. Generation of an internal atmosphere of 0.3 bars of pure argon (three cycles of vacuuming using a primary pump followed by filling by pure argon). Heating, melting, chemical homogenization, cooling and solidification (110 kHz, 5 kV, isothermal stage of three minutes in liquid state).

Metallographic preparation; chemical and microstructure control

Cutting of the obtained spherical ingots in two halves, using a metallographic saw (Delta Abrasimet Cutter, BUEHLER). Embedding of one of an ingot half in a cold-reacting mixture (resin and hardener: Mecaprex MA2+, PRESI). Grinding using SiC papers from 400-grit to 4000-grit, under water flow, ultrasonic cleaning, final polishing with textile enriched by 1 µm diamond particles until mirror-like state, short cleaning by water then by ethanol and drying. Microstructure control using a Scanning Electron Microscope (JSM-6010LA, JEOL) in Back Scattered Electrons mode (SEM, BSE). Chemical composition control using the Energy Dispersion Spectrometer (EDS) attached to the SEM.

Making of the electrodes

Sawing of the upper part of the second ingot half and insertion of the denuded copper part of a plastic gained electrical wire. Embedding of the half ingot with the same resin and hardener mixture as previously used for the metallographic samples, with total immersion of the spherical part of the alloy surface and of the denuded part of the electrical wire, the circular plane



Figure 1 Melting of one of the alloys.

surface of the half ingot remaining free. Grinding and polishing following the same procedure as previously described for the metallographic samples (mirror-like state too).

Preparation of the artificial saliva

Weighing of NaCl (AnalaR Normapur) and dissolution in distilled water, homogenization for 10 minutes using a magnetic agitator. Heating to $37^{\circ}\text{C} \pm 2^{\circ}\text{C}$ and oxygenation by air bubbling for 15 minutes. Rating of the pH to either 7.4 or to 2.3 using a 0.01 Mol/L NaOH solution, followed by readjustment of temperature to 37°C .

Preparing the apparatus allowing the corrosion characterization

Positioning of the double-envelope three-electrode cell, connection to a F32 heater/cooler (JULABO) and heating of the cell by internal circulation of warm water (heated and maintained at about 39°C by the F32 device) (**Figure 2**). Connection of the double-envelope three-electrode cell to a potentiostat (263A, EG&G Instruments Princeton Applied Research) driven by a computer. Filling of the cell with the $\{37^{\circ}\text{C}$ -heated, 15 min aerated, pH 7.4)-solution. Positioning of the working electrode (the Co-Cr alloy), the counter electrode (platinum) and the potential reference electrode (saturated Calomel). Running of the electrochemical experiments (driven by the PowerSuite software) Follow up of the free potential (open circuit potential, E_{ocp1}) of the working electrode for one hour. Polarization from $E_{ocp} - 20 \text{ mV}$ to $E_{ocp} + 20 \text{ mV}$ at 10 mV/min and measurement of the polarization resistance R_{p1} (leading to the assessment of the current density of corrosion I_{corr} according to the Stern-Geary method).

Follow up of the free potential (open circuit potential, E_{ocp1}) of the working electrode for a second hour. Second polarization from $E_{ocp} - 20 \text{ mV}$ to $E_{ocp} + 20 \text{ mV}$ at 10 mV/min and measurement of the polarization resistance R_{p2} (new I_{corr}). Then:

1. Either a Tafel run (polarization from $E_{ocp} - 250 \text{ mV}$ to $E_{ocp} + 250 \text{ mV}$ at 10 mV/min , to determine the I_{corr} , the corrosion potential I_{corr} and eventually the cathodic and anodic Tafel coefficients β_c and β_a
2. Or a cyclic polarization between $E_{ocp} - 150 \text{ mV}$ to $E_{ocp} + 1.23 \text{ V}$ at $+1 \text{ mV/s}$ followed by return to $E_{ocp} - 150 \text{ mV}$ at -1 mV/s to examine the behavior of the alloy at high potential (e.g. passivation, stability of the passive scale).

Results

Chemical and microstructural state of the alloys

SEM observations of the mounted and polished samples were carried out in BSE mode to examine the microstructure for several brightness and contrast ratings, and to control the obtained global chemical compositions. The SEM/BSE micrographs displayed in **Figure 3** (only four of the seven alloys) illustrate the metallurgical state of all alloys: they are all single-phased and chemically homogeneous. The Cr contents are well respected for all alloys.



Figure 2 Photograph of the double-envelope cell with the three positioned electrodes.

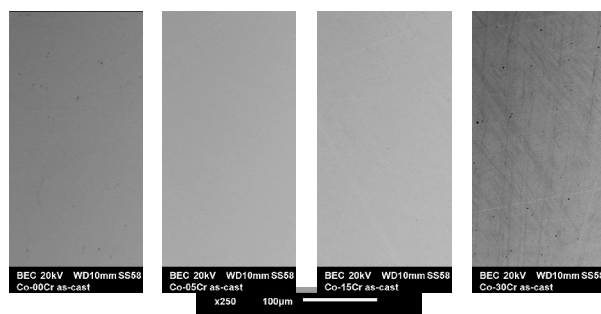


Figure 3 Illustration of the single-phased states of the alloys (here only 4 out of the 7 ones, SEM/BSE, 250X).

Free-potential evolution

These alloys were also mounted as electrodes and subjected to electrochemical experiments to characterize their corrosion behavior in the artificial saliva preliminarily aerated and heated/maintained at the body temperature (37°C). The first electrochemical experiments which were carried out were the follow-up of the free potential or open circuit potential during two hours. This was performed two times per alloy, one time preceding a Tafel experiment (TF) and one time preceding a cyclic polarization one (CP). The obtained evolution curves are plotted versus time all together in **Figure 4**. Except one for the Co-0Cr alloy and one for the Co-30Cr alloy, all these curves are rather close to one another without any real hierarchy according to the chromium content in alloy. A common point for a majority of these curves is an initial increase in potential during the first

half an hour of immersion. In two or three cases this increase may take more time. The values of these potentials after 1 hour and after 2 hours are given in **Table 1**. The potential domains of stability of the O, Co and Cr species are added, on the left for oxygen and on the right for cobalt and chromium. This allows noting that the potential stays in all cases in the domains of stability of water (H_2O), of the double hydroxide of cobalt ($Co(OH)_2$) and of the oxide of chromium (chromia, Cr_2O_3).

Low potential range linear polarization runs

After one hour as well as after two hours of immersion with E_{ocp} follow-up, a potential-increasing linear polarization of low amplitude (± 20 mV) centered on the current E_{ocp} was applied to each working electrode in order to issue values of polarization resistance R_p which may lead to an estimation of the corrosion current according to the Stern-Geary method. The obtained R_p values are given in **Table 2**. The polarization resistances are of only $3\text{ k}\Omega \times \text{cm}^2$ for the Co-0Cr alloy while adding 5 wt.% Cr leads to multiply it by ten (20 to $60\text{ k}\Omega \times \text{cm}^2$ for the Co-5Cr alloy). With 10 wt.% Cr the R_p values reach about $100\text{ k}\Omega \times \text{cm}^2$ and

enriching more in Cr leads to $200\text{ k}\Omega \times \text{cm}^2$. Even some values are approaching $1\text{ M}\Omega \times \text{cm}^2$ (seen one time in the case of the Co-25Cr alloy) it is more convenient to say that they are of several hundreds of $\text{k}\Omega \times \text{cm}^2$ as soon as the chromium content is equal to or higher than 15 wt.%.

Tafel runs

After two hours of immersion a Tafel experiment was run, between -250 mV and $+250$ mV around the current open circuit potential. This led to the $E=f(\log [(I/A)/(S/\text{cm}^2)])$ curves plotted together in **Figure 5**. The $I=0$ peaks (i.e. E_{corr}) are all globally centered on about -200 mV/NHE (except one of the two Tafel curves obtained for the Co-30Cr alloy). In all cases this corresponds to the H_2O , $Co(OH)_2$ and Cr_2O_3 domains. The Tafel curve which is placed the most on the right of the graph is the one corresponding to the Co-0Cr alloy. Its abscissa (i.e. I_{corr}) is thus the highest among all the tests. The Tafel curves obtained for the Co-5Cr and Co-10Cr alloys are in intermediate position (medium corrosion current densities) while the curves corresponding to the Co-15Cr, Co-20Cr, Co-25Cr and Co-30Cr alloys are all situated on the left of the graph (low corrosion current densities). The values of E_{corr} and I_{corr} resulting of the Tafel treatment of these curves are displayed in **Table 3** as well as the cathodic and anodic Tafel coefficients β_c and β_a . The E_{corr} values are rather fluctuating, with no relation with the chromium content. In contrast the I_{corr} values steadily decreases when more and more chromium is present, from the Co-0Cr alloy to the Co-15Cr/Co-20Cr alloys, before becoming very low when the Cr content reaches 25 and 30 wt.% Cr (only several tens of Nano-Amperes per cm^2).

Cyclic polarization runs

After two hours of immersion cyclic polarizations were performed between a potential a little slower than the open circuit potential and a rather high anodic potential (>1 Volt by regards to the Normal Hydrogen Electrode). The potential-increasing parts of the curves obtained for the seven alloys are displayed in **Figure 6**. In all cases the open circuit potentials are located in the H_2O , $Co(OH)_2$ and Cr_2O_3 domains, this suggesting that the alloys are corroded through the passivation layer made of the cobalt hydroxide or of

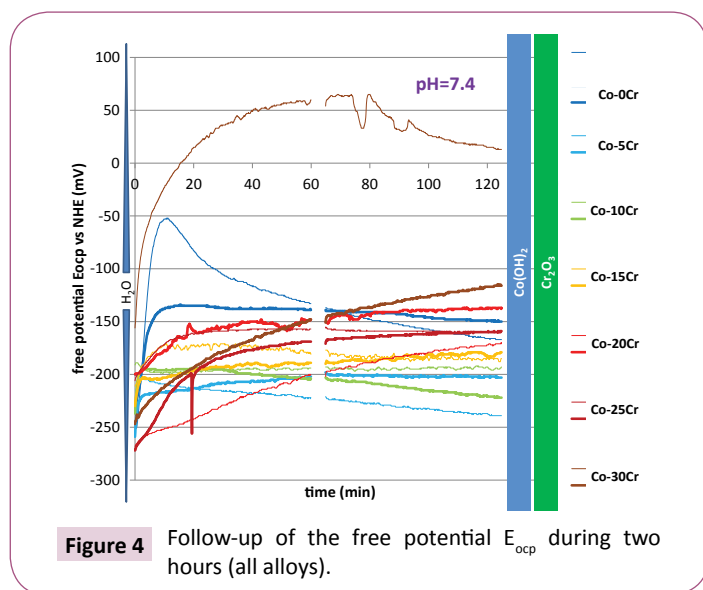


Figure 4 Follow-up of the free potential E_{ocp} during two hours (all alloys).

Table 1 Values of the free potential after 1 hour and after 2 hours of immersion preceding the Tafel plot (TF) or the cyclic polarization (CP) for all alloys.

		Eocp/HNE in mV							
Immersion time	TF or PP	wt.% Cr	0	5	10	15	20	25	30
1 h	TF		-248	-201	-191	-211	-268	-206	-159
	CP		-232	-259	-237	-230	-200	-272	-247
2 h	TF		-137	-221	-193	-177	-202	-159	+54
	CP		-139	-198	-203	-190	-150	-171	-151

Table 2 Values of the polarization resistances after 1 hour and after 2 hours of immersion preceding the Tafel plot (TF) or the cyclic polarization (CP) for all alloys.

		Rp values in $\Omega \times \text{cm}^2$							
Immersion time	TF or PP	wt.% Cr	0	5	10	15	20	25	30
1 h	TF		2995	33881	84366	343763	163638	767212	250878
	CP		3224	37822	41883	139152	201776	322118	349065
2 h	TF		4167	51979	130051	356338	287846	926971	95183
	CP		3169	56005	66884	215847	366515	451024	539688

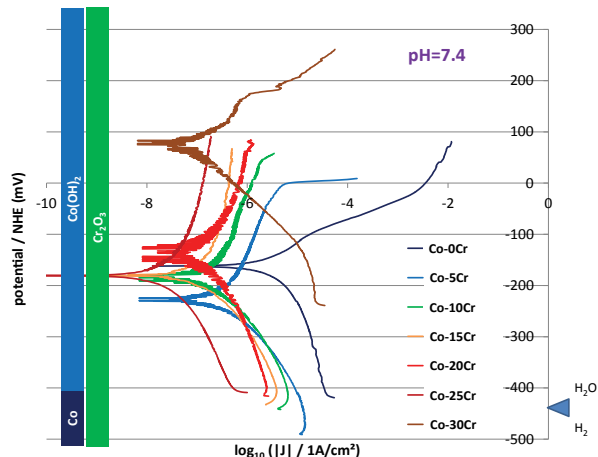


Figure 5 Tafel plot after two hours of immersion (all alloys).

Table 3 Values of the corrosion potential and the corrosion density of current after 2 hours of immersion, according to the Tafel method; values of the cathodic and anodic Tafel coefficients.

Tafel calculations results							
wt.% Cr	0	5	10	15	20	25	30
E _{corr} (mV/NHE)	-162.27	-226.57	-182.75	-178.66	-126.79	-181.95	+82.45
I _{corr} (μA/cm²)	1.711	0.412	0.216	0.201	0.261	0.031	0.067
β _c (mV/dec)	74.938	120.67	127.16	156.55	277.16	164.22	90.436
β _a (mV/dec)	60.345	236.30	238.26	575.67	311.31	260.64	45.927

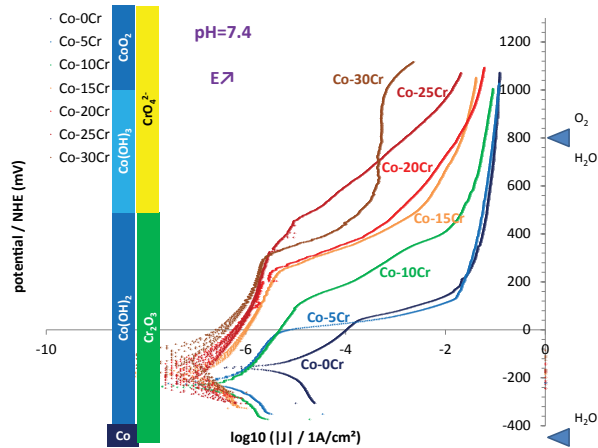


Figure 6 Increasing-potential part of the cyclic polarization curves after two hours of immersion (all alloys).

the chromium hydroxide ($\text{Cr}(\text{OH})_3$) or oxide (Cr_2O_3), by dissolved oxygen (O_2). When the applied anodic potential raises to higher and higher values jumps in anodic current occur, for rather low anodic potentials for the Co-0Cr and Co-5Cr alloys (between 0 and +50 mV/NHE), at around +100 mV/NHE for the Co-10Cr alloy, and around +350 mV/NHE for the other chromium-rich

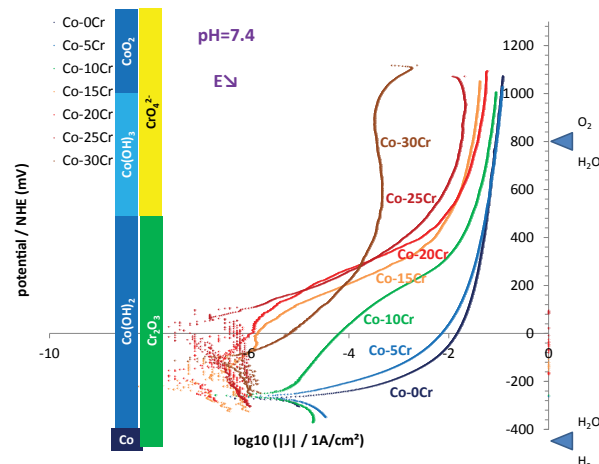


Figure 7 Decreasing-potential part of the cyclic polarization curves after two hours of immersion (all alloys).

alloys. For potentials high enough (more than +500 mV/NHE) transpassivation takes place (CrO_4^{2-} domain). Chromia is then re-oxidized in chromate ions dissolving in the solution, leading to an additional anodic current. The potential-decreasing parts of the cyclic polarization curves are shown in **Figure 7**. Their shapes are similar to the potential-increasing parts. Furthermore the new corrosion potentials and corrosion current densities are similar to the ones observed per alloy in the potential-increasing parts.

Discussion

These binary cobalt-chromium alloys were thus successfully elaborated, despite the high melting points of both elements. However, the fusion apparatus which was used is very powerful by comparison by the devices used by the dental technician for melting and shaping the metallic parts of prostheses. In this way having less chromium (the most refractory one among the two) in the composition keeps great interest, and knowing how the corrosion behaviour of a Co-xCr alloy evolves versus x is useful to optimize the Cr content.

The study of the microstructures reveals that all alloys are single-phased, and moreover chemically homogeneous. This facilitates the investigation of the influence of the chromium content on the corrosion behaviour of these alloys since all potentials and current intensity data are not composed of several contributions the decorrelating of which is generally difficult and speculative. Despite this great simplification for the interpretation of the electrochemical results in the field of the microstructures of the alloys, some scattered results remained, notably the curiously great value of the open circuit potential of the Co-30Cr alloy in **Figure 4** and the corresponding particularly high Tafel curve in **Figure 5**. One can guess that this is the concentration in dissolved oxygen – the main oxidant specie involved in these experiments – which was difficult to master, despite that a precise protocol was defined and respected the best possible.

Nevertheless some very interesting observations were done concerning the dependence on the chromium content of

the average level of open circuit potential, the polarization resistance, the potential and current densities of corrosion as assessed according to the Tafel method. These results are consistent and they all show that the alloys get all passivated after less than one hour of free immersion. The passivation scale is either $\text{Co}(\text{OH})_2$ (Co-0Cr alloy, maybe also Co-5Cr and Co-10Cr by parts) and $\text{Cr}(\text{OH})_3/\text{Cr}_2\text{O}_3$ for all alloys containing chromium. The polarization resistances are rather high (Co-0Cr alloy), high (Co-5Cr and Co-10Cr) or very high (Co-15Cr, Co-2Cr, Co-25Cr and Co-30Cr) and the corrosion currents vary in the opposite way.

All the alloys underwent passivation which allowed them resisting corrosion. But the cyclic polarization tests, notably their potential-increasing parts, seemed revealing a lack of stability of the passivation scale by the sudden increases in anodic current when the applied anodic potential exceeded a potential step. Before reaching this step (e.g. +100 mV/NHE for the Co-10Cr alloy) one thinks that the passivation layer was well covering and stable, with a corrosion current (low) corresponding to the equilibrium between the anodic current in the passivation plateau and the reduction of the dissolved oxygen. When exceeding the same step the passivation layer became unstable and get fractured, with a part of the alloy surface, now no longer protected, falling again in an active state. For the highest chromium contents considered in this work, 25 and 30 wt.%, the passivation layer maybe thicker and more resistant, thanks to more chromium available: the anodic current in passive state is thus very low and the deterioration of the passivation layer is delayed to higher potentials, and even does not occur, the increase in anodic

current being simply related to the transpassivation $\text{CrIII}(\text{OH})_3$ or $\text{CrIII}_2\text{O}_3 \rightarrow \text{CrVIO}_4^{2-}$.

Conclusion

The complementary electrochemical techniques used to characterize the corrosion behaviour of these Co-xCr alloys with different Cr contents led thus to consistent results in term of corrosion current. The chromium-free pure cobalt material logically displayed the worst behaviour with polarization of several $\text{k}\Omega \times \text{cm}^2$ and corrosion currents of $2 \mu\text{A}/\text{cm}^2$ (no higher thanks to a neutral pH allowing passivation). With 5 or 10 wt.% Cr the behaviour is better, without dispute (R_p higher of an order of magnitude and corrosion current of several hundreds of nA/cm^2). But this is for 15 wt.% Cr and higher that the behaviour became very good according to Stern-Geary (very high R_p). However, this is for 25 wt.% Cr and more that the alloys are the best corrosion resistant, with corrosion currents of only several tens of nA/cm^2 and the best stability of the passive state.

The aqueous solution prepared from distilled water and NaCl (9 g/L) with pH rated to 7.4 is artificial saliva classically used for such corrosion studies on dental alloys, but the pH may vary in mouth in real situations, as after absorption of acidic drinks. To anticipate any variation in corrosion behaviour due to pH fluctuation, as well as to any possible detrimental effect of elements added to the Co-xCr binary base to improve the mechanical resistance, weldability and other properties (dental alloys may also contain molybdenum, silicon...) it is may be preferable to retain 25 wt.% and not 15 wt.% for the chromium content.

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