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Conjugated Oligomers with Alternating Heterocycles from a Single Monomer: Synthesis and Demonstration of Electroluminescence

Sara Urrego-Riveros,^{a,b,f} Matthias Bremer,^c Jonas Hoffmann,^{a,b,d} Anne Heitmann,^{a,b} Thibault Reynaldo,^d Janek Buhl,^c Paul J. Gates,^e Frank D. Sönnichsen,^f Muriel Hissler,^d Martina Gerken,^c Anne Staubitz^{a,b,f,*}

Conjugated oligomers based on two different heterocycles are typically prepared by step growth polycondensation cross-coupling methods from two monomers **X-Cycle¹-X** and **M-Cycle²-M** with no control of the regioselectivity. In this work, we used a new synthetic strategy that involves an extremely chemoselective reaction of a dielectrophilic compound, **X¹-Cycle¹-X²**, with a dinucleophilic component, **M¹-Cycle²-M²**, under Stille conditions. The resulting monomers, **X¹-Cycle¹-Cycle²-M²** are di-heterocyclic push pull monomers that still contain a nucleophilic site (boronic acid) and electrophilic site (bromide) and are set up for a controlled polymerization under Suzuki conditions. In this way, two semiconducting oligomers, based on thiophene / benzene and thiophene / pyridine motifs were synthesized. Both oligomers were characterized in terms of their, thermal, electrochemical, absorption, emission and electroluminescence properties.

Introduction

Organic semiconductors are promising functional materials for applications in electronic devices such as light emitting diodes (OLED's), field effect transistors (OFET's) or photovoltaic solar cells.¹⁻⁹ For example, oligo and polythiophenes have been widely used as semiconductors due to their electron rich character, high conductivity ($> 10^5$ S/cm)¹⁰ and band gaps (E_g) ranging between 1 and 3 eV.^{11, 12} Many polythiophenes show a high aggregation in solid-state leading to high charge mobilities.^{13, 14} However this aggregation can lead to non-radiative decay pathways, such as intersystem crossing or internal conversion into the ground state, which decreases the quantum yield of oligo- and poly-thiophenes and limits their utilization in OLEDs.^{11, 15}

To overcome these limitations to be utilized in OLEDs, first of all, the aggregation in the solid state can be suppressed. This can be achieved by designing specific copolymers of thiophenes with other conjugated molecules, to prevent

interchain interactions.^{15, 16} Tuning the band gap can be achieved by combining different heterocycles in the polymer chain; a push-pull system can be generated by placing an electron rich group; e.g. a thiophene unit adjacent to an electron deficient unit; e.g. another aryl group (Figure 1). Typically, this leads to a bathochromic shift in the absorption.^{12, 15} However, the quantum yield may be decreased, depending on the strength of the dipoles. The push-pull combination decreases the mobility of the excitons and prevents them from quenching. In addition, non-radiative decay processes of excitons can be avoided.^{17, 18}

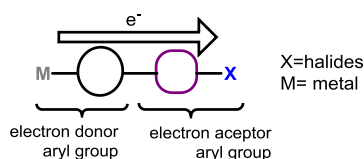


Figure 1. Design of semiconducting polymers using the push-pull principle.

One of the most common methods to produce such copolymers is based on the step-growth polymerization between two different co-monomers **A** and **B**. In such a reaction, monomer **A** would have two electrophilic sites for cross-coupling (halides, pseudohalides), whereas monomer **B** would have two metal functional groups (organotin, boronic acids, boronic esters, organozinc groups for example (Scheme 1a)).^{12, 18} The polymers **C** produced by this method can be expected to have a high dispersity (\mathcal{D}) and a low regioselectivity.

A second method is based on the chain-growth polymerization developed by McCullough and Yokozawa. The mechanism

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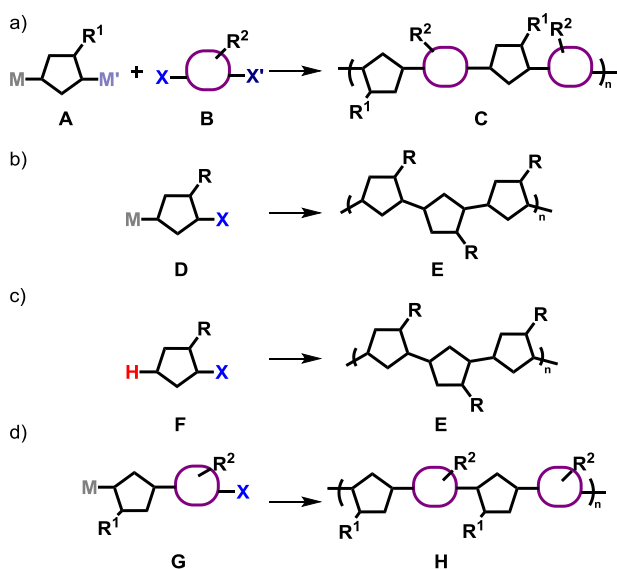
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involves a catalyst-transfer condensation using a single monomer **D**, which contains an electrophilic and nucleophilic site (Scheme 1b).¹⁹⁻²¹ With this breakthrough concept, it was possible, in the case of poly(3-alkylthiophenes), to obtain good molecular weights (M_n), a very high regioselectivity >98%, and a low dispersity; in addition, the nature of the end groups of the polymers **E** could be controlled.²² Even alternating copolymers have been obtained by this method, although in that case, only one heterocycle contained a solubilizing group, thus reducing problems of torsional defects.²³ A third, novel method is based on the C-H arylation of monomers of type **F** to produce homopolymers type **E** (Scheme 1c).²⁴⁻²⁶ The C-H bond is activated by using a catalytic system formed between a Pd, or Ni catalyst, a ligand and a base (e.g. Hermann's catalyst- Cs_2CO_3 ,^{25, 27} $\text{NiCl}_2\text{-dppe}/(\text{TMP})_2\text{Mg}\cdot 2\text{LiBr}$ ²⁸ or $\text{Pd}(\text{OAc})_2\text{-pivalic acid}$ ²⁷). This method holds much promise and can be extended to heteropolymers of the type **AB**.^{26, 29}

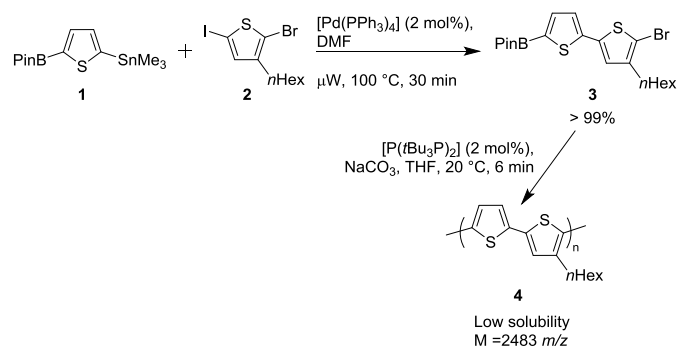
The aim of this work was to combine the concept of living polymerization with the advantage of combining two different building blocks in only one monomer. Therefore, monomers **G** of type **M-Ar-Ar-X** were designed that bear both an electrophilic and nucleophilic site allowing to obtain polymer **H** (Scheme 1d).^{19, 30}



Scheme 1. Types of polymerization a) step-growth polymerization between **A** and **B** molecules b) living polymerization from one type of heterocycle **D**. c) direct arylation polymerization from the monomer **F** d) living polymerization of copolymers from di-heterocyclic monomers **G** (this work).

Recently, we published a monomer, which was the product of a dually nucleophilic-electrophilic selective cross-coupling reaction.³¹ Therein, monomer **3** was prepared to contain two different units of thiophene; a key strategic synthetic element was that dinucleophile **1** and dielectrophile **2** reacted extremely chemo specifically (Scheme 2): Only the iodo- and the tristannyl functional groups cross-coupled. However, the polymer **4** that was obtained showed a very low solubility and could only be used as a proof-of-concept. It was impossible to analyze it by solution-state NMR and the optoelectronic

properties could not be measured. However by MALDI-MS, an oligomer with a mass of 2483 m/z was detected,³¹ which was very promising and led us to pursue the concept of monomer formation by dually electrophile and nucleophile selective reactions for monomers that would lead to *soluble* polymers or oligomers. In this context, the work of Zhang and Hong should be mentioned, who at the same time as ourselves followed a similar strategy for the synthesis of a polyarene.³²



Scheme 2. Synthetic way to obtain a polymer based on a single monomer type **M-Ar-Ar-X**.³¹

In this work, we report a dually nucleophilic-electrophilic selective cross-coupling reaction to prepare two monomers, **5** and **6** type **G** (**M-Ar-Ar-X**). The monomers **5** and **6** were designed using the concept of the push-pull principle, having an aryl or pyridine (**6**) moiety as an electron acceptor group.¹⁸ This part of the molecule was designed as the di-electrophilic part, whereas the thiophene moiety was designed as the electron-rich di-nucleophilic part. In order to improve the solubility of the monomers, a hexyl chain was installed at the thiophene unit and a hexyloxy chain on the six-membered aromatic ring (Figure 2). The oligomers **O1** and **O2** were then obtained by a Suzuki cross-coupling living polymerization.

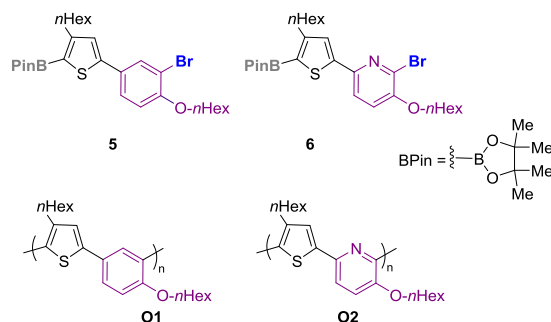


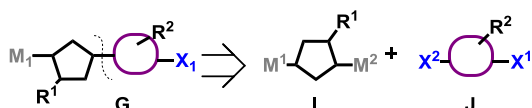
Figure 2. Newly prepared push pull monomers **5** and **6** and oligomers **O1** and **O2**.

In addition to the synthesis, we report on the structural, optical and thermal characteristics of the obtained oligomers **O1** and **O2**. First electroluminescence tests were performed to assess these materials as potential emission layers in OLEDs.

Results and discussion

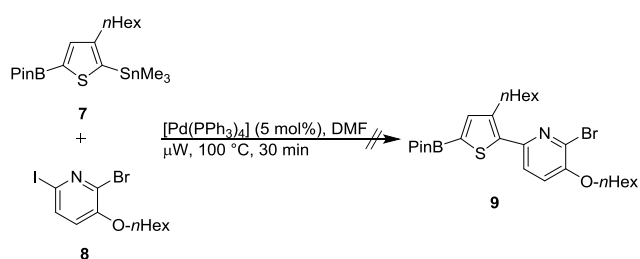
Synthesis of the monomers

The retrosynthetic analysis explains the high selectivity, which is required to obtain the monomer type **G** (Scheme 3): It is vital that in the dinucleophile **I**, only metal M^2 reacts and in dielectrophile **J**, only halogen X^2 . Any reaction of M^1 or X^1 would lead to slightly different monomers that would be inseparable from the desired ones and, moreover, there would be no regioselectivity with respect to the solubilizing side groups R^1 and R^2 .



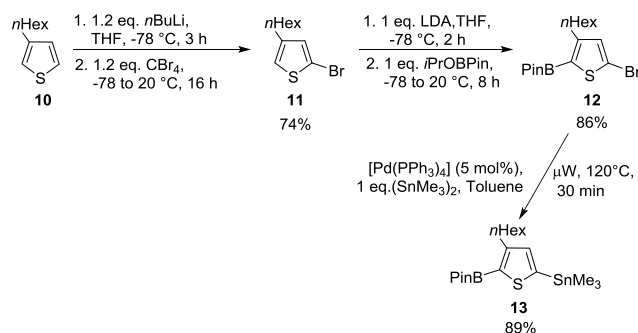
Scheme 3. Retrosynthetic pathway to obtain the desired monomers type **G**.

The dinucleophile **7** was prepared as reported in the literature.³¹ The first attempt to prepare the monomer **9** was a selective Stille cross-coupling reaction between the dinucleophile **7** and the dielectrophile **8** (Scheme 4), in a similar way to how monomer **3** was obtained (Scheme 2). However, no product was formed either with conventional or under microwave conditions. One reason might be the steric hindrance of the hexyl chain that prevents the Stille reaction between the iodo-group of **8** and the trimethylstannyl group of **7**, although electronic reasons might play a role. McCullough *et al.* showed that 3-hexylthiophene monomers react much more readily if the M functional group is in 5-position and the X functional group in 2-position. They argue that the transmetalation step is favorable if the M functional group is in a less steric hindered position relative to the catalyst center of the growing polymer.³³⁻³⁵



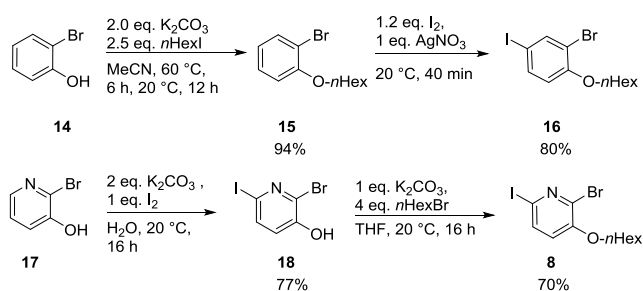
Scheme 4. Synthetic pathway to obtain monomer **9**.³¹

Therefore, the dinucleophile had to be redesigned. Thus, dinucleophile **13** was prepared (Scheme 5). The synthetic pathway involved a site-selective lithiation reaction of 3-hexylthiophene (**10**) in 5-position, followed by quenching with CBr_4 to obtain 2-bromo-4-hexylthiophene (**11**).³⁶ Site-selective deprotonation followed by a lithium-boron exchange led to mononucleophile **12**. In a final step, a selective Stille-Kelly coupling reaction of **12**,³⁷ with $(SnMe_3)_2$, Pd(0) under microwave heating led to the novel dinucleophile **13** in a yield of 89%.



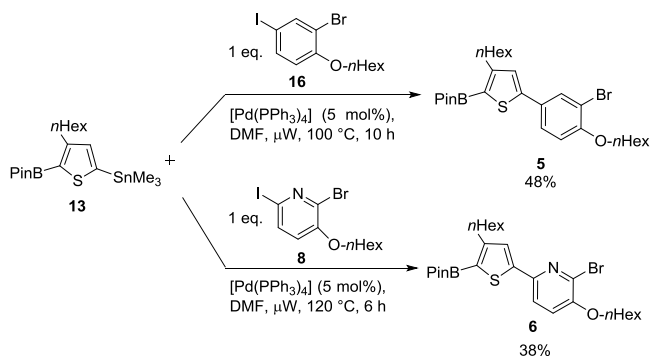
Scheme 5. Synthetic pathway to form the dinucleophile **13**.

For the Stille cross-coupling reaction, two dielectrophiles **16** and **19** were also required. The synthesis of the dielectrophile **16** started from the commercially available 2-bromophenol (**14**), which was subjected to a Williamson etherification with 1-iodohexane and led to **15** in a yield of 94%.³⁸ Then, a solvent-free iodination of 1-bromo-2-(*n*-hexyloxy)benzene (**15**) was carried out to give the 1-bromo-2-(*n*-hexyloxy)benzene-5-iodobenzene (**16**) in 80% yield (Scheme 6).³⁹ The synthesis of dielectrophile **8** started from 2-bromopyridine-3-ol (**17**), which was subjected to an iodination reaction to yield the iodo-pyridine derivate **18** in 77%. Compound **18** was then used in an alkylation reaction with 1-bromohexane to yield the dielectrophile **8** in 70% (Scheme 6).⁴⁰



Scheme 6. Synthetic path to obtain the dielectrophiles **16** and **8**.

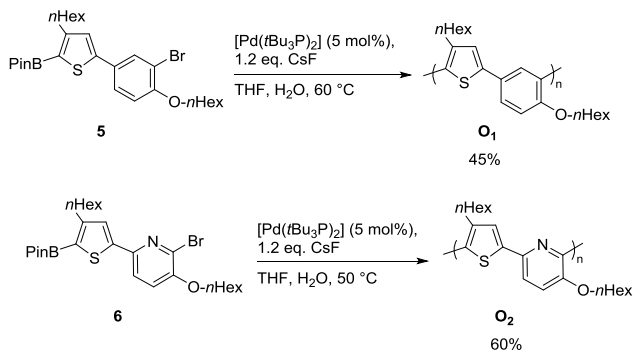
The dinucleophile **13** and dielectrophiles **16** and **8**, respectively, were subjected to a Stille cross-coupling reaction under microwave conditions leading selectively to the monomers **5** and **6** (Scheme 7).^{31, 41} The high selectivity of the reaction can be explained by the use of dinucleophile **13**, as the reaction occurred selectively between the iodo and trimethyl tin group. There is less steric hindrance for the tin functional group compared to the boronic ester, and Suzuki-reactions would require an additional base. In addition, the C-I bonds in **16** and **8** are much weaker than the C-Br bonds.



Scheme 7. Synthetic pathway to form the monomers **5** and **6** in a Stille cross-coupling reaction between dinucleophile **13** and dielectrophiles **16** and **8**, respectively.

Synthesis of oligomers **O1** and **O2**.

Polycondensation Suzuki reactions of the monomers **5** and **6** were tested with two catalysts, ($[\text{Pd}(\text{tBu}_3\text{P})_2]$ and $[\text{Pd}(\text{dppf})\text{Cl}_2]$) (see in ESI Tables SI-3 and SI-4). From these, $[\text{Pd}(\text{tBu}_3\text{P})_2]$ emerged as the best system and it was used in the polymerization reactions for both monomers, which were heated at 60 and 50 °C respectively (Scheme 8).⁴² The crude products were precipitated from the reaction mixture with a solution of HCl in methanol (1 M) and washed several times with methanol by Soxhlet extractions. The products were soluble in chloroform. Oligomer **O1** was obtained in a yield of 46% ($M_n = 1.56$ kDa, $\bar{D} = 1.12$, GPC calibrated against polystyrene) and **O2** in a yield of 60% ($M_n = 5.39$ kDa, $\bar{D} = 1.83$), respectively. It was found by MALDI-TOF mass spectrometry that for both polymers, the end groups were H and Br (see in ESI, Figure SI-2 for **O1** and SI-9 for **O2**).

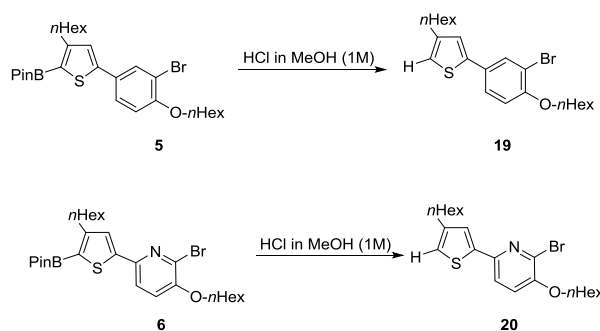


Scheme 8. Syntheses of the oligomers **O1** and **O2**.

Analysis of the oligomerization of oligomers **O1** and **O2**.

To evaluate the type of kinetics of polymer chain formation, reaction monitoring studies of the polymerization were performed. Samples from the reaction mixture were taken in the course of 24 h for **O1** and 27 h for **O2**. Each sample was quenched with a solution of HCl in methanol 1 M straight after removal from the reaction, extracted with DCM and split in two for measurements of GPC and ^1H NMR spectroscopy. The percentage of conversion for each sample, was calculated relative to the protonated products **19** and **20** that are formed by protonolysis (Scheme 9). The plots of % conversion vs. time and % conversion vs. M_n and M_n/M_w are depicted in Figure 3 for **O2**. The conversion was calculated with

respect to the consumption of product **20** observed by ^1H NMR spectra and from the UV-Vis/ elugram areas visualized in the 2D



plots from the GPC measurements (See ESI).

Scheme 9. Compounds **5** and **6** were transformed into the products **19** and **20**, respectively, by treating them with HCl/MeOH (1 M).

The oligomer appeared to grow approximately linearly in length during the first 9 h, reaching 60% conversion and a $M_n = 1.40$ kDa. Then, in the period from 9 h to 28 h, only very little growth in length took place, but a conversion of eventually 100% product was reached. This observation may be interpreted that the reaction had a living character until a conversion of 60%. A similar behavior could be observed for **O1**: after 8 h, the conversion reached 40%, growing sharply until 92% in the following 16 h (See the ESI for Figures SI-21, 22 and 23).

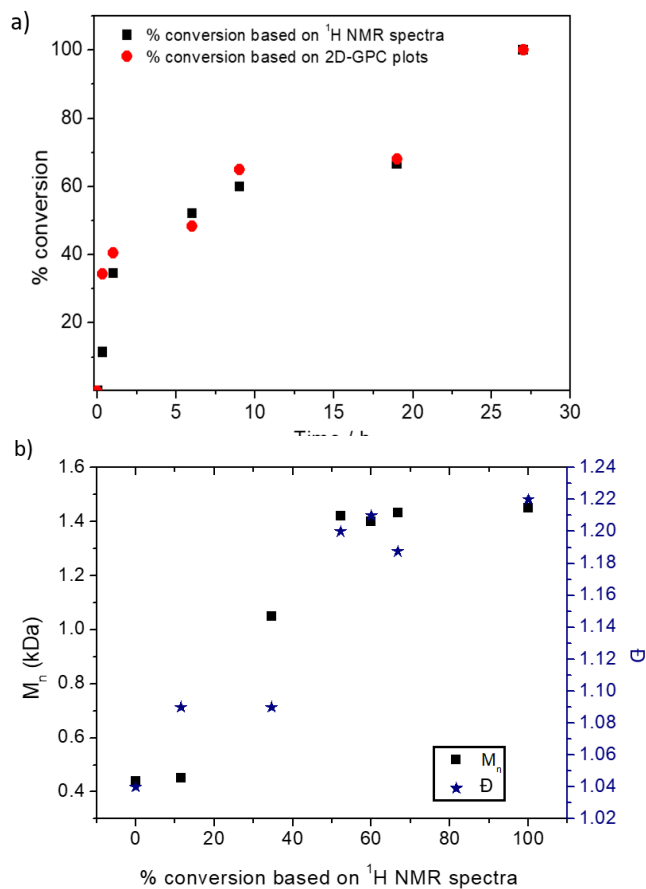


Figure 3. Plots of a) % conversion vs. time. Calculated by the two detectors (^1H and 2D-plots GPC) and b) M_n and \bar{D} vs. % conversion calculated by ^1H

NMR; for the polymerization of **6** with 5 mol% of $[\text{Pd}(\text{tBu}_3\text{P})_2]$, 1.2 eq. CsF in THF at 50 °C. The conversion was calculated relative to compound **20** (product from the quenching process with HCl/MeOH) based on ^1H NMR spectra and 2D-GPC plots.

It needs to be pointed out that although the formal classification of the polymerization kinetics is chain growth, the linear region of the quasi-living character is relatively limited, presumably because the catalyst diffuses from the growing chain more quickly than a new monomer is reacted with it.^{19, 43} This can be attributed to the sterically relatively congested site of cross-coupling (with adjacent alkyl and alkoxy groups).

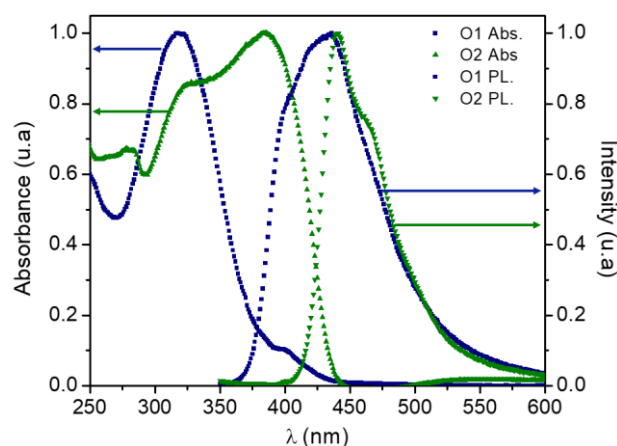
Optical and electrochemical characterization

In order to establish structure/properties relationships, the optical (UV-Vis absorption and fluorescence) and electrochemical properties of **O1** and **O2** were studied in CH_2Cl_2 . The two oligomers **O1** and **O2** presented a broadband in the absorption spectrum (Figure 4). The absorption maximum observed for **O2** ($\lambda_{\text{max}} = 387 \text{ nm}$) was bathochromically-shifted compared to the absorption maximum of **O1** ($\lambda_{\text{max}} = 318 \text{ nm}$) suggesting a higher extended π -conjugation in **O2** due to a certain charge transfer character of π - π^* transition of the thiophene-pyridine motif. A similar trend was observed for the fluorescence of the oligomers. Oligomers **O1** and **O2** showed fluorescence around 435 nm with a small red-shift for the derivative **O2** ($\Delta\lambda = 5 \text{ nm}$). The fluorescence quantum yields measured versus quinine sulfate were moderate in solution (10%, **O1** and 28%, **O2**, Table 1)). Excitation spectra also confirmed the observations made for the UV-Vis absorption of oligomers **O1** and **O2**. (see Figures SI-3 and SI-10). The optoelectronic properties from both oligomers in comparison with a 4-unit oligomer of thiophene (in chloroform solution: $\lambda_{\text{max}} = 372 \text{ nm}$,⁴⁴ $\lambda_{\text{PL}} = 494 \text{ nm}$,⁴⁴ $\Phi = 0.09$ ⁴⁴ and $E_{\text{g optical}} = 3.33 \text{ eV}$ ⁴⁵), shows that a combination with a stronger acceptor than thiophene (pyridyl group in **O2**) leads indeed to a higher bathochromic shift. However, although a phenyl group is also an acceptor with respect to thiophene,¹⁸ λ_{max} of **O1** is hypsochromically shifted, likely due to torsional defects arising from the six-membered ring that are not sufficiently counterbalanced by an acceptor property that is lower than that of a pyridyl ring. The photoluminescence is most strongly shifted in the oligomer of oligo-3-hexylthiophene.

The redox properties of oligomers **O1** and **O2** were investigated by cyclic voltammetry (CH_2Cl_2 , 0.2 M, Bu_4NPF_6 , $v = 200 \text{ mVs}^{-1}$, Table 1). All compounds showed a quasi-reversible oxidation wave at relatively low potential. Compound **O1** displayed an oxidation wave at around +0.48 V (vs. Fc^+/Fc) while **O2** was slightly easier to oxidize ($E_{\text{ox}} = +0.47 \text{ V}$ vs. Fc^+/Fc). These electrochemical characterizations indicated that the highest occupied molecular orbital (HOMO) was energetically destabilized for oligomer **O2** compared to the HOMO of **O1** (Table 1).⁴⁶ The reduction potentials, which are attributed to the energy of the lowest unoccupied molecular orbital (LUMO) and therefore to the π^* band, were not possible to detect because the potentials were out of the electrochemical window (Figure SI-4 for

O1 and Figure SI-14 for **O2** respectively). Hence, the reduction potential (LUMO) for both oligomers were estimated using the HOMO and optical energy gap values.

Figure 4. Normalized absorption spectra (abs.) and PL spectra of **O1** and **O2**



($\lambda_{\text{exc}} = 340 \text{ nm}$) in dichloromethane.

Table 1. Optical and electrochemical properties of the oligomers **O1** and **O2**.

Compound	$\Phi^a / \lambda_{\text{PL}}(\text{nm})$	E_{g} (eV) exp ^b / calc. ^f	E_{ox} (V) ^c	E_{HOMO} (eV) exp ^d / calc. ^f	E_{LUMO} (eV) Exp ^e / calc. ^f
O1	0.10 / 432	3.17 / 3.95	+0.48	-5.87 / -5.17	-2.70 / -1.22
O2	0.28 / 442	2.77 / 3.54	+0.47	-5.86 / -5.11	-3.09 / -1.57

^a The quantum yield was calculated using a solution of quinine sulfate (1 M in H_2SO_4) with a Φ of 0.54. ^b Optical band gap. ^c Obtained from the cyclic voltogram of a solution of the oligomer in DCM ($1 \times 10^{-3} \text{ M}$) with Bu_4NPF_6 (0.2 M in DCM) as electrolyte.⁴⁶ ^d The energy from the HOMO was calculated with the equation: $E_{\text{HOMO}} = -(E_{\text{onset ox. vs Fc}^+/\text{Fc}} + 5.39) \text{ eV}$.⁴⁶ ^e The energy from the LUMO was calculated with the equation $E_{\text{LUMO}} = -(E_{\text{g}} - E_{\text{HOMO}})$.⁴⁶ ^f Calculated using TD-DFT based on the trimers.

Such a difference in the optoelectronic behavior between the oligomers **O1** and **O2** can be related to the enhanced donor-acceptor effect in **O2**, because of the higher π acceptance of the pyridine compared to the benzene in **O1**.¹⁸

Computational studies

These experimental observations were further supported by time-dependent density functional theory (TD-DFT) studies in the gas phase. To simplify this computational study, the calculations were carried out on two simplified monomer structures **S1** and **S2**, and the corresponding dimers and trimers (Figure 5).

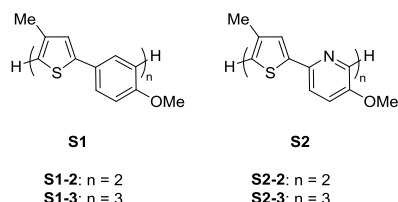


Figure 5. Monomers, dimers and trimers proposed for computational studies. The hexyl chains from the monomers **5** and **6** were replaced by methyl groups in **S1** and **S2** respectively, to reduce computational time. Further the end groups were replaced by a proton as the most neutral substituent.

TD-DFT calculations confirmed that trimer **S2-3** had a narrower energy gap (see Table 1) between the calculated HOMO and LUMO (3.54 eV) compared to **S1-3** (3.95 eV) Figure 6. Besides, the HOMO levels in both trimers were similar (-5.17 eV vs. -5.11 eV), the LUMO is more stabilized in the case of **S2-3** than in **S1-3** (-1.57 eV vs -1.22 eV). To explain this phenomenon and further estimate the effect of either the phenyl or pyridine ring on the respective thiophene unit, the bond length alternation (BLA) for both systems were calculated.

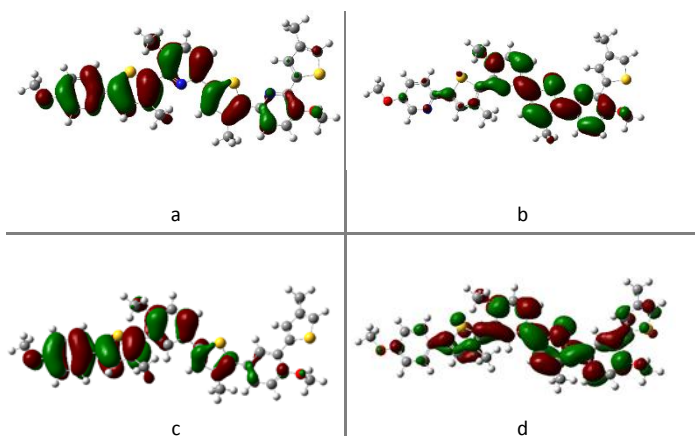


Figure 6. Molecular orbitals (displayed with an isovalue of 0.02), of trimer **S1-3**: HOMO (a) and LUMO (b) and of trimer **S2-3**: HOMO (c) and LUMO (d).

By comparing the double and single bond lengths in π -conjugated linear polymers, BLA values can indicate if the ground state is more likely to form a quinoid-like or an aromatic structure.⁴⁷ Since **S1** and **S2** are asymmetric, the BLA calculations were done by only taking the inner thiophene bonds into account. Shorter BLA values of the flanked thiophene in **S2-3** (maximum 0.045 Å) in comparison to **S1-3** (maximum 0.052 Å, Table SI-15), and a shorter central carbon bond between the thiophene and phenyl/pyridine of **S2-3** (1.467 Å vs 1.472 Å) indicate that the quinoidal structure is more likely formed in **S2-3** and explains the stabilization of its LUMO by fortified delocalization of π -electrons.

Thermal behavior of **O1** and **O2**

The thermal stability of the polymers was studied by thermogravimetric analysis under N_2 at a heating rate of 10 K/min in a range from 25-600 °C. **O1** lost 59% of weight at 406 °C (onset temperature). **O2** showed two segments of

weight loss: first 4% weight loss at 221 °C (onset temperature), probably due to traces of volatiles in the sample; and 50% weight loss at 405 °C (onset temperature), directly related to the loss of the alkoxy and alkyl chains.

The DSC thermograms indicated a glass transition temperature of -7.1 °C for **O1** and of 25 °C for **O2**.

Electroluminescence experiments

The oligomers **O1** and **O2** were tested as emissive layer between electrodes in single-layer OLED devices, to prove their electroluminescent properties. The configurations of the devices were ITO/PEDOT:PSS/**O1** or **O2**/LiF/Al. The electroluminescence spectra of both devices (Figure 7) resemble the PL spectra of compounds **O1** and **O2** but they are red shifted (Figure 4 and 7; for **O1** shifted from 432 to 533 nm and for **O2** from 442 to 593 nm). These EL spectra are obtained with high applied voltages of 7.7 V and 10.7 V (corresponding to 1 cd/m^2), respectively. At this voltage, the luminescence of **O1** started to level, indicating an instability of the compound. **O2**, albeit with a lower luminescence, did not show such a decomposition. Furthermore, a luminance of 2.5 and 0.8 cd/m^2 was achieved at ca. 10 V, respectively and the external quantum yields are quite low presumably due to unbalanced carrier injection and/or transport.

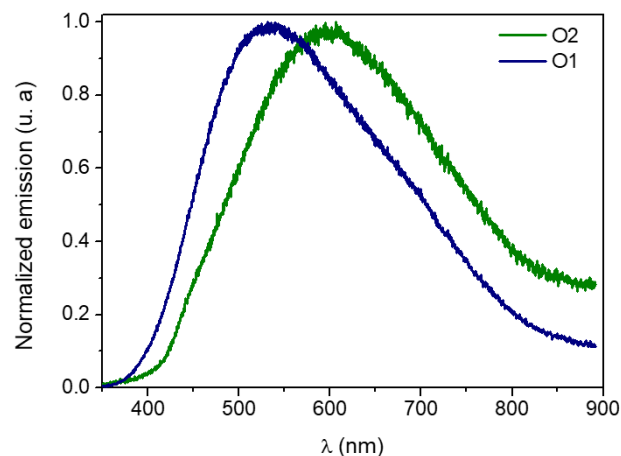


Figure 7. Electroluminescence spectra for the OLED-type devices using **O1** and **O2** as emissive layer.

Effectively, the energy level between the ITO/PEDOT:PSS-anode and the oligomers did not match well (with an energy difference of 0.66 eV for **O2** and 0.67 eV for **O1**, Figure 8). The current densities are high, suggesting that charge carriers move through the device without many emissive recombination events (see Figure SI- 28 and Table SI- 7). This may be caused by a carrier injection imbalance or a dominant non-radiate recombination in a different layer. Therefore, in order to explore the full potential of oligomers **O1** and **O2**, an optimization of the device is required, which is beyond the scope of this study and the object of further investigations.

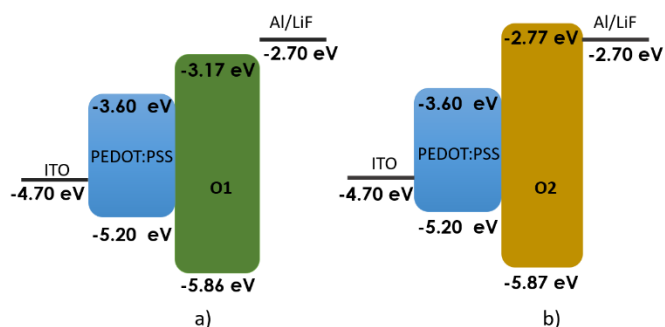


Figure 8. Schematic diagrams of the energy levels of the OLEDs using the oligomers a) **O1** and b) **O2** as emissive layers.

Experimental part

All the synthesis and characterization of the starting materials and polymers are given in the ESI.

Characterization of the OLED-type devices

The I-V characterization of the OLED-type devices was performed with a Keithley 2400 Source Measurement Unit, a calibrated photodiode and an Avantes AvaSpec-2048 USB spectrometer.

OLED-type device fabrication.

ITO coated glass substrates were cleaned in an ultrasonic bath of acetone and isopropanol for 15 min each. Subsequently, they were dried under a nitrogen stream, heated to 160 °C for 10 min and then cleaned with an oxygen plasma with 300 W for 3 min. 200 μ L of a PEDOT : PSS solution were spin-coated onto the substrates (3500 rpm for 60 s). The samples were then transferred into a glovebox and annealed at 150 °C for 5 min. Afterwards, they were cooled down to 50 °C over 2 min. A solution of the polymers either in toluene or chloroform was prepared with a concentration of 10 mg/mL. 180 μ L of the polymer solution were spin coated (500 rpm for 5 s, and 1000 rpm for 60 s) on the PEDOT : PSS layer and the samples were annealed at 70 °C. Finally, a layer of LiF (1 nm) and Al (200 nm) were deposited by thermal evaporation.

Conclusions

Monomers based on the push pull principle were synthesized using a highly chemo selective Stille reaction between a dinucleophile, containing pyridinyl **5**, or aryl groups **6**, and a dielectrophile, which contained a thiophene moiety. Based on those monomers (**5** and **6**), two oligomers **O1** and **O2** were synthesized using a Suzuki polymerization reaction. Studies analyzing the dependence of molecular weight on conversion showed that the monomers **5** and **6** may grow in a living fashion, at least until a certain point, 60% and 40% respectively, to obtain the oligomers **O1** and **O2**. After this point, the catalyst presumably dissociates from the chain. **O1** and **O2** showed optical band gaps in solution with values of 3.17 eV and 2.77 eV; and quantum yields of 10% and 28%

respectively. Due to those optoelectronic properties, simple devices to test for electroluminescence with the configuration ITO/PEDOT:PSS/**O1** or **O2**/LiF/Al were fabricated. The devices showed a turn on voltage of 4.3 and 4.6 V demonstrating the electroluminescent character from the oligomers **O1** and **O2**.

Conflicts of interest

There are no conflicts to declare.

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