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# Accelerated ultraviolet aging of structural and luminescent properties of the ureasil-polyether hybrid materials U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup>

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**KEYWORDS:** *Organic-inorganic hybrid materials, polyether, sol-gel, photooxidation, photoluminescence.*

## HIGHLIGHTS:

- Formate was the main photoproduct formed during the ultraviolet aging of ureasil-polyether hibrid (U-PEO and U-PPO).
- In Eu<sup>3+</sup> doped hybrid (U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup>) a concomitant decreases of the Eu<sup>3+</sup> complexed to urea bridges, the spatially correlated siloxane nodes, and the Judd-Ofelt parameter  $\Omega_2$  (R) were observed during the photodegradation process
- The sensitivity of structural features and PL properties arising from the photoinduced degradation of ureasil-polyether films could be exploited in the fabrication of new transparent, flexible, and biocompatible UV dosimeters and sensors.

**Abstract:** In this paper, we evaluate the role played by the polyether type and chain length on the complexation of Eu<sup>3+</sup> by the urea group of siloxane-polyether organic-inorganic hybrid materials based on polyethylene oxide (U-PEO, mw = 1900 g.mol<sup>-1</sup>) and polypropylene oxide (U-PPO, mw = 230, 2000, and 4000 g.mol<sup>-1</sup>). Investigation was made of the effect of the photooxidation of the PEO and PPO chains, induced by artificially accelerated ultraviolet aging ( $\lambda \geq 300$  nm), with the decomplexation of Eu<sup>3+</sup> from the urea-polyether array. Fourier transform infrared (FTIR) and photoluminescence (PL) spectra evidenced loss of the ligand-to-metal charge transfer (LMCT) when the Eu<sup>3+</sup> cations were unable to disrupt the strong hydrogen-bonded urea/urea aggregates formed using shorter U-PPO chains (230 g.mol<sup>-1</sup>) or semicrystalline U-PEO (1900 g.mol<sup>-1</sup>). FTIR analysis showed that formate was the main photoproduct formed continuously during the artificial ultraviolet aging of U-PEO (1900 g.mol<sup>-1</sup>) and U-PPO (2000 g.mol<sup>-1</sup>), with concomitant decreases of Eu<sup>3+</sup> complexed to urea bridges, spatially correlated siloxane nodes, and the Judd-Ofelt parameter  $\Omega_2$  (R). The kinetics of all these parameters were described by exponential laws with comparable time constants, evidencing the sensitivity of the Eu<sup>3+</sup> photoluminescence to the elimination/formation of the hydrogen-bonded urea-polyether array during the photooxidation of PPO chains. In the case of U-PEO:Eu<sup>3+</sup>, different time constant values were observed, indicating the existence of a faster process governed by the facilitated transport of the oxidative species through the amorphous moieties of the semicrystalline U-PEO. This sensitivity of structural and PL properties arising from the photoinduced degradation of ureasil-polyether films can be exploited in the fabrication of new transparent, flexible, and biocompatible UV dosimeters and sensors.

## 1. Introduction

The combination of lanthanide compounds and organic-inorganic hybrids (OIHs) provides the benefits of the intrinsic characteristics of sol-gel derived host materials (such as processability, mechanical, thermal, and chemical stability[1–3], and biocompatibility) and the outstanding luminescent properties of lanthanide compounds (such as line-like emissions, long luminescence lifetimes, large Stokes shifts, and high luminescence quantum yields)[4,5]. Research into lanthanide-doped OIH sol-gel materials emerged about thirty years ago, with the use of this soft

chemistry route to prepare homogeneous and transparent materials with tunable composition and tailored properties[6–8]. In general, the lanthanide ions (Ln) embedded in an OIH host matrix are linked to matrix moieties by specific interactions such as covalent, coordination, electrovalent, dipolar, and hydrogen bonds[9,10]. The existence of specific guest-host (Ln-OIH) coordination sites can overcome the disadvantages of concentration quenching, clustering or aggregation of emitters, inhomogeneity, or poor chemical stability[4,5,9,11].

An important example of the easy tunability of the guest-host interaction site is provided by the Ln-OIH family classified as di-ureasils. These materials present urea bridges between a silicate backbone and polyether (PE) macromers of different molecular weights, known as U-PE[12,13]. The hybrids display an intrinsic broad green-blue emission from the urea bridges, as well as emission originating from electron-hole recombination in the siliceous domains of the inorganic backbone[14–16]. The energy transfer from the di-ureasil emitter centers to incorporated  $\text{Eu}^{3+}$  ions results in white luminescence constituted of  $\text{Eu}^{3+}$  red emission in addition to the broad green-blue emission arising from the hybrid host matrix[17]. Furthermore, the emission color can be tuned across the CIE chromaticity diagram by changing the excitation wavelength, the PE chain length, or the amount of incorporated europium salt[18]. The two latter dependencies are especially related to the higher efficiency of energy transfer to  $\text{Eu}^{3+}$  coordinated to the carbonyl groups ( $\text{O}_{\text{carb}}$ ) of the urea bridges, which are located near the hybrid host emitting centers, as compared to  $\text{Eu}^{3+}$  coordination to ether oxygens of the PE chain[19]. This interesting characteristic has resulted in a large number of studies focused on the design and synthesis of U-PE: $\text{Ln}^{3+}$  for applications in different areas such as chemical sensing, bio-imaging, dye-sensitized photoelectrochemical cells, and organic light-emitting devices (OLEDs)[20–25].

A difficulty of long term application is that organic materials and polymers, such as those used in OIHs, are degraded during their life cycle. Depending on the application, these materials are subjected to environmental aging factors including UV light, adverse temperature, and the atmosphere[26]. Degradation and oxidation of the polymers have several serious effects on their chemical, physical, and mechanical properties including molecular weight, tensile strength, impact strength, and creep, among others[27]. Changes in mechanical properties result from reactions that either reduce the size of molecules or lead to crosslinking between polymer chains to form rigid three-dimensional structures[28]. In this way, chain scission and oxidation reactions can be intimately coupled in polymer degradation and can affect the stability of the OIH: $\text{Ln}^{3+}$  complex, causing loss of the  $\text{Ln}^{3+}$  optical properties[19]. The most common mechanisms of polymer degradation are photooxidation and thermooxidation[29–31]. These types of degradation are often triggered by macroradical formation from the abstraction of hydrogen atoms, leading to scission of the main chain[32]. In the case of polyethers such as polyethylene oxide (PEO) and polypropylene oxide (PPO), degradation proceeds by random chain scission, resulting in several oxidation products: esters, aldehydes, and formates[33–36]. The photooxidation and thermooxidation mechanisms are quite similar, differing in terms of the amounts of the main oxidation products. Thermooxidation produces equal amounts of esters and formates, whereas photooxidation generates formates and esters in a ratio of 5:1[37].

In recent work [20] we have demonstrated that specific coordination of  $\text{Eu}^{3+}$  to the carbonyl groups ( $\text{O}_{\text{carb}}$ ) of the urea bridges and of  $\text{Li}^+$  to the ether-type oxygen ( $\text{O}_{\text{eth}}$ ) can be exploited to prepare U-PE materials with multifunctional properties, namely photoluminescence (PL) and

ionic conduction. This emerging dual functionality was achieved by the co-incorporation of  $\text{Eu}^{3+}$  and  $\text{Li}^+$  into OIH materials based on U-PEO and U-PPO with similar molecular weights ( $1900 \text{ g mol}^{-1}$  for PEO and  $2000 \text{ g mol}^{-1}$  for PPO) using a metal molar ratio fixed at  $[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}] = 3$  and  $[\text{O}_{\text{eth}}]/[\text{Li}^+] = 15$ . Although the degradation of PEO and PPO is well documented [33–36], the photodegradation and the effect of photo ageing on the PL properties of  $\text{Eu}^{3+}$  doped materials (U-PEO:Eu and U-PPO:Eu) have not been reported. The aim of this study was therefore to investigate the effect of the ultraviolet light irradiation time ( $\lambda \geq 300 \text{ nm}$ ) on the photodegradation of U-PE film prepared with polymer of different nature and molecular weight, and doped with several  $[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}]$  molar ratios. Special attention was paid to correlate the aging on the PL properties of U-PEO:Eu $^{3+}$  and U-PPO:Eu $^{3+}$  materials with the modification of the chemical environment of  $\text{Eu}^{3+}$  induced by photo degradation of PE chains. The photodegradation studies were performed under air at atmospheric pressure using an accelerated aging device with  $\lambda > 300 \text{ nm}$  (SEPAP12/24). The effects of aging time on the photodegradation products, the OIH nanostructure, and changes of the  $\text{Eu}^{3+}$  complexation site were investigated using Fourier transform infrared (FTIR) spectroscopy, small-angle X-ray scattering (SAXS), and photoluminescence techniques.

## 2. Experimental

### 2.1 Synthesis of the materials

The OIH materials were synthesized by the sol–gel method, according to a well-known procedure[38]. In brief, the ureasil (U) cross-linking agent IsoTrEOS (3-(isocyanatopropyl)triethoxysilane, Sigma-Aldrich, 95% purity, CAS #24801-88-5) was covalently bonded to both ends of the different PE macromer by reacting the terminal amino-propyl groups, followed by the sol–gel reaction involving the  $\text{Si}(\text{OEt})_3$  hydrolysis, followed by the condensation reaction to form ureasil cross-linking nodes. This procedure was adopted to prepare U-PEO with mw  $1900 \text{ g.mol}^{-1}$  (Jeffamine<sup>®</sup> ED-2003), (CAS #65605-36-9) and UPPO with mw 230 (Jeffamine<sup>®</sup> D-230), 2000 (Jeffamine<sup>®</sup> D-2000) and 4000  $\text{g.mol}^{-1}$  (Jeffamine<sup>®</sup> D-4000 (XTJ-510)) (CAS #9046-10-0) using as reactant Jeffamine<sup>®</sup> ED-2003, Jeffamine<sup>®</sup> D-230, Jeffamine<sup>®</sup> D-2000, and Jeffamine<sup>®</sup> D-4000, respectively. The doping with different  $\text{Eu}^{3+}$  loading was achieved by suitable mass of europium(III) chloride hexahydrate (Fluka, Aldrich, 99.9% purity, CAS #13759-92-7) dissolved into the OIH precursor ethanoic solution in order to prepare U-PEO:Eu $^{3+}$  and U-PPO:Eu $^{3+}$  samples with  $[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}]$  of 0.75, 1.5, 3.0, 4.5, and 6.0. Films of  $\sim 50 \text{ }\mu\text{m}$  thick were prepared by casting OIH sol on a Teflon mould and drying at room temperature ( $25 \text{ }^\circ\text{C}$ ).

### 2.2 Accelerated ultraviolet aging

The accelerated photodegradation was performed in the presence of air, at  $60 \text{ }^\circ\text{C}$ , in a SEPAP 12/24 device. This apparatus was designed for the study of polymer photodegradation during artificial aging corresponding to moderately accelerated conditions[39]. The device was equipped with four medium-pressure mercury lamps (Mazda MA 400) in a vertical position at one axis of the chamber and the samples were placed on a rotating carousel  $13 \text{ cm}$  in diameter, which could hold 24 samples. Wavelengths below  $300 \text{ nm}$  were filtered using a glass envelope. The OIH samples were exposed in the form of rectangular films ( $2 \times 3 \text{ cm}$ ) with thickness of  $50 \text{ }\mu\text{m}$ .

## 2.3 Characterization techniques

The FTIR spectra of the OIH films were recorded with a Nicolet 760-FTIR spectrometer operating in transmission mode and controlled with OMNIC software. Spectra were obtained with accumulation of 32 scans and  $4\text{ cm}^{-1}$  resolution.

The absolute quantum yield (AQY) efficiencies and the emission spectra (upon excitation at 365 nm) were determined using a C9920-02G PL-QY system (Hamamatsu). This consisted of a 150 W monochromatized Xe lamp, an integrating sphere (with Spectralon<sup>®</sup> coating,  $\varnothing = 3.3$  inches), and a high sensitivity CCD spectrometer for detecting the entire spectral luminescence (the emission spectra were integrated from 576 to 728 nm for the europium-based compounds). The automatically controlled excitation wavelength range was from 250 to 500 nm, with a resolution bandwidth better than 5 nm. Absolute photoluminescence (PL) quantum yield excitation spectra were obtained by scanning the excitation and monitoring the overall luminescence of  $\text{Eu}^{3+}$ .

The photooxidized samples were submitted to  $\text{NH}_3$  chemical treatment in order to identify the photoproducts. The irradiated films were exposed to  $\text{NH}_3$  gas in all-Teflon vessels, at room temperature, in a simple flow system that could be sealed off to permit the reaction to proceed. In parallel to these treatments, it was observed that the unoxidized samples without  $\text{Eu}^{3+}$  loading did not react with the gas. This procedure, which has been described in a previous paper[36], was performed for the photooxidized OIH materials without and with  $\text{Eu}^{3+}$  loading.

The nanostructures of the as-prepared OIH and photooxidized samples were analyzed by small-angle X-ray scattering (SAXS) measurements performed at the SAXS1 beamline of the National Synchrotron Light Laboratory (LNLS, Campinas, Brazil). The beamline was equipped with a 2D Pilatus 300k detector located 910.9 mm from the sample, recording the image of the scattering intensity,  $I(q)$ , as a function of the modulus of the scattering vector,  $q = 4\pi/\lambda \sin(\varepsilon/2)$ , where  $\varepsilon$  is the X-ray scattering angle. The data were normalized considering the varying intensity of the direct X-ray beam, the detector sensitivity, and the sample transmission. The intensity of the parasitic scattering due to the cell windows and vacuum was subtracted from the total scattering intensity.

The Differential Scanning Calorimetry (DSC) measurements were carried out using a Q100 analyzer (TA Instruments). The samples were heated at the rate  $5\text{ }^\circ\text{C}/\text{min}$  in the range  $-90\text{ }^\circ\text{C} - 100\text{ }^\circ\text{C}$ , nitrogen was used for the gas purge, supplied at a flow rate of  $50\text{ mL}/\text{min}$ . The melting enthalpy ( $\Delta H_m$ ) was recorded during the first heating run. The degree of crystallinity ( $D_C$ ) was calculated from the relationship between  $\Delta H_m$  and the standard melting enthalpy for 100 % crystalline PEO ( $\Delta H_p$ ) ( $196.4\text{ J/g}$ ), following the equation:  $D_C (\%) = (\Delta H_m / \Delta H_p) * 100\%$ [40].

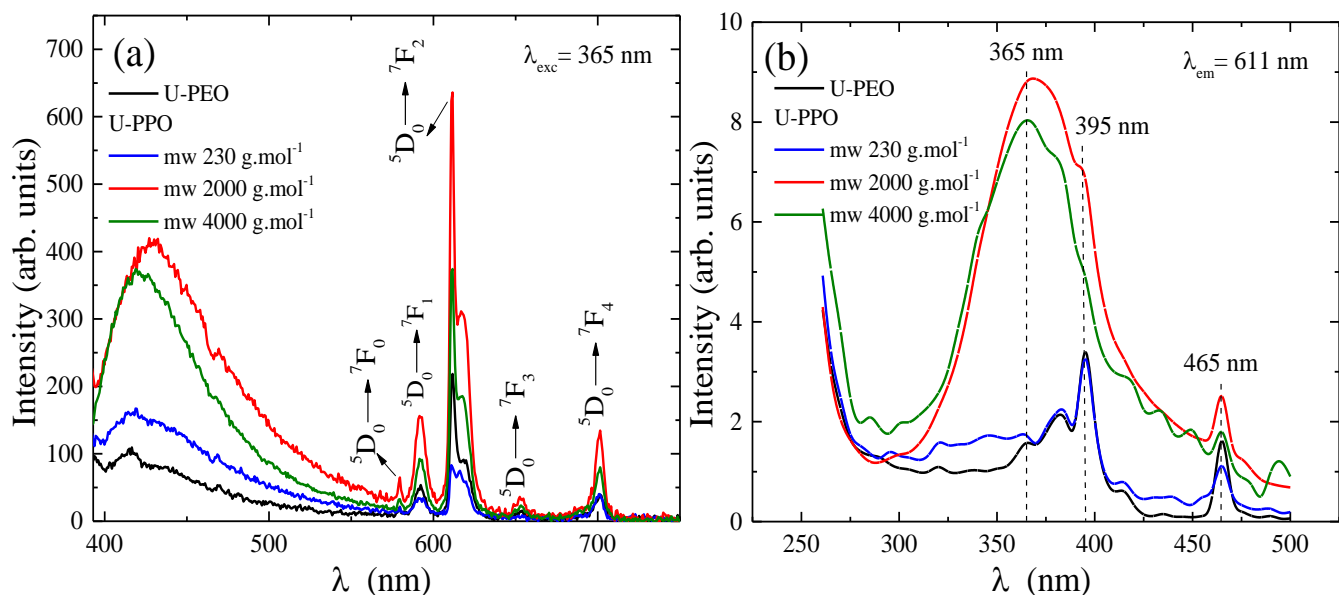
## 3. Results and Discussion

### 3.1 $\text{Eu}^{3+}$ complexation in the U-PEO and U-PPO host OIHs

In order to determine the  $[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}]$  ratios that provided the best optical efficiencies for the two polyether hybrid materials (U-PEO, with  $\text{mw} = 1900\text{ g.mol}^{-1}$ , and U-PPO, with  $\text{mw} = 2000\text{ g.mol}^{-1}$ ), the absolute quantum yields (AQY) were determined as a function of the amount of  $\text{Eu}^{3+}$ , using UV excitation at 365 nm. The results are shown in **Figure SI.1** (Supplementary Information). Regardless of the polyether type, the highest AQY values were obtained for

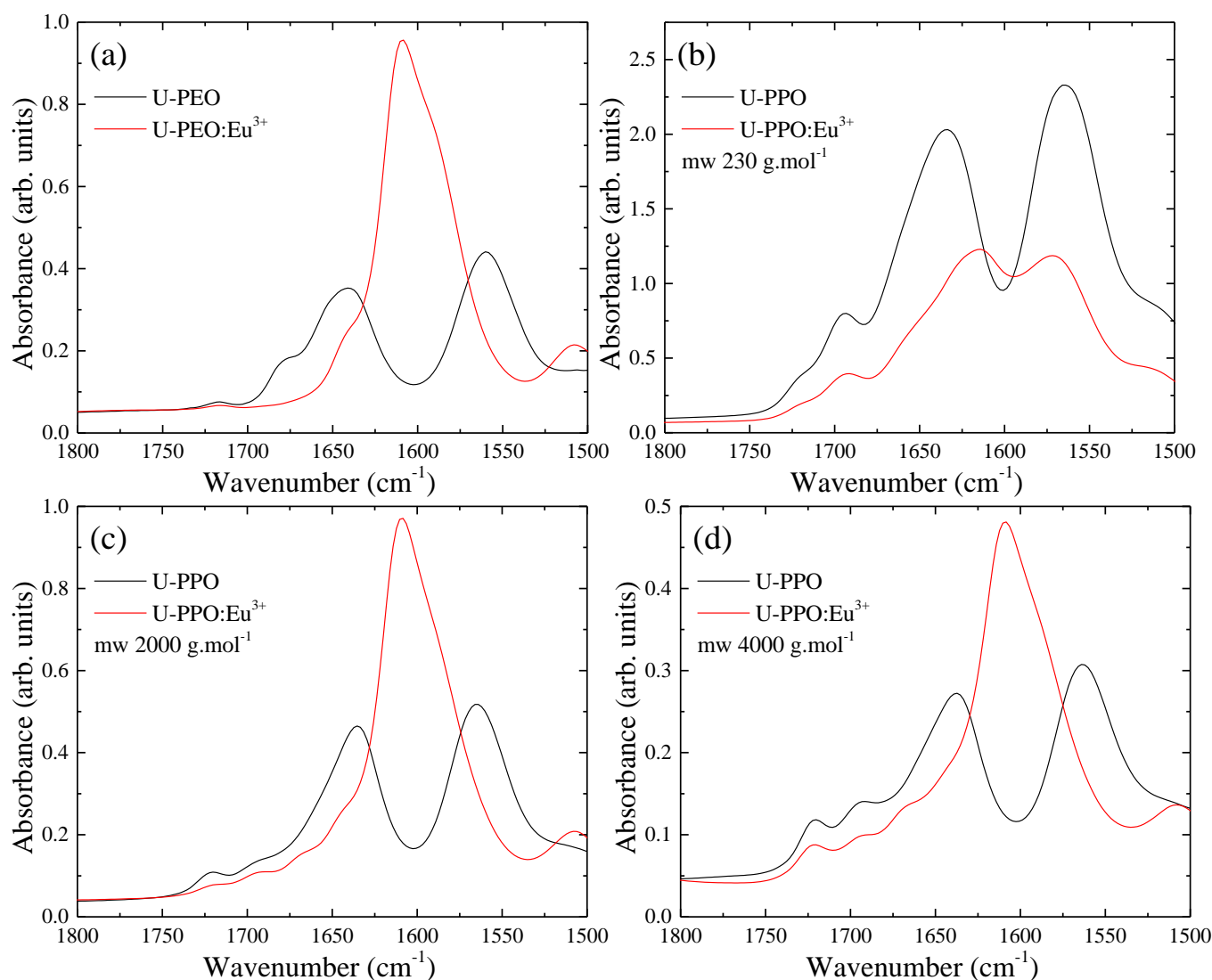
$[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}] = 3$ , with values of 2.3% and 8.9% for U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup>, respectively. The photoluminescence properties of OIH materials doped with different Eu<sup>3+</sup> concentrations were studied by Ribeiro et al.[41] and Bermudez et al.[42]. The results obtained in these studies were similar and indicated that for lower doping concentrations, the dried gels presented distributions of Eu<sup>3+</sup> typical of an amorphous environment, while a crystalline-like environment of Eu<sup>3+</sup> was observed for higher concentrations. Decreasing radiative decay rates and spectral variations suggested that the transformation of the surrounding Eu<sup>3+</sup> environment from amorphous-like to crystalline-like led to quenching of the Eu<sup>3+</sup> PL features. Considering the semicrystalline and amorphous states of U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup> [20], respectively, this explanation also supports the strong difference of AQY found for the OIHs prepared with polymers of different natures.

Photoluminescence (PL) studies were performed with the OIHs prepared with PE chains of different molecular weights (PPO of 230, 2000, and 4000 g.mol<sup>-1</sup>, and PEO of 1900 g.mol<sup>-1</sup>), containing the optimized amount of Eu<sup>3+</sup> ( $[\text{O}_{\text{carb}}]/[\text{Eu}^{3+}] = 3$ ). The PL emission spectra recorded upon UV excitation at 365 nm are displayed in **Figure 1a**. All the spectra exhibited narrow emission bands arising from *f-f* transitions of Eu<sup>3+</sup> ions from the <sup>5</sup>D<sub>0</sub> excited state to <sup>7</sup>F<sub>J</sub> (J = 0-4) ground states. The <sup>5</sup>D<sub>0</sub>→<sup>7</sup>F<sub>2</sub> transition located at 611 nm led to the strongest emission intensity, conferring a red emission feature to the samples. The broad band in the green-blue spectral range could be ascribed to the convolution of the emission from NH groups of urea bridges and the emission originating from electron-hole recombination in the siliceous domains in the inorganic backbone[43–45], which have already been observed for unloaded matrices[46] of similar hybrids[14,15]. The latter feature was much stronger for U-PPO:Eu<sup>3+</sup> of high molecular weight (2000 and 4000 g.mol<sup>-1</sup>), compared to U-PEO:Eu<sup>3+</sup> of 1900 g.mol<sup>-1</sup> and U-PPO:Eu<sup>3+</sup> of 230 g.mol<sup>-1</sup>, indicating a significant sensitization of the organic part upon excitation at 365 nm. This suggested that the Eu<sup>3+</sup> atoms in U-PPO:Eu<sup>3+</sup> of lower chain length were present in a crystalline-like environment similar to that observed for U-PEO:Eu<sup>3+</sup>. This effect of the nature and chain length of the PE was in agreement with the PL excitation spectra showing a broad band centred at 365 nm for the U-PPO:Eu<sup>3+</sup> hybrids of high molecular weight (2000 and 4000 g.mol<sup>-1</sup>), when monitoring the Eu<sup>3+</sup> <sup>5</sup>D<sub>0</sub>→<sup>7</sup>F<sub>2</sub> transition at 611 nm (**Figure 1b**). This band was associated with the ligand-to-metal charge-transfer (LMCT) and occurred in the same energy region of the <sup>7</sup>F<sub>0</sub>→<sup>5</sup>L<sub>6</sub> (at 395 nm) and <sup>7</sup>F<sub>0</sub>→<sup>5</sup>D<sub>2</sub> (at 465 nm) transitions, ascribed to the intra-configurational transitions of Eu<sup>3+</sup> [17]. This feature was consistent with an antenna effect mechanism[10], indicating efficient transfer from the high energy levels of the ligands to high excited states of the Eu<sup>3+</sup> ions. This antenna effect was lower for U-PEO:Eu<sup>3+</sup> and for U-PPO:Eu<sup>3+</sup> of low molecular weight (230 g.mol<sup>-1</sup>).



**Figure 1:** (2-column fitting image) (a) Emission and (b) excitation spectra before photooxidation ( $t$  (h) = 0) for U-PEO:Eu<sup>3+</sup> (mw = 1900 g.mol<sup>-1</sup>) and U-PPO:Eu<sup>3+</sup> of different molecular weights (mw = 230, 2000, and 4000 g.mol<sup>-1</sup>), with [Ocarb]/[Eu<sup>3+</sup>] = 3.

Valuable information concerning the dependency of the chemical environment of the Eu<sup>3+</sup> cations on the nature and chain length of the PE could be obtained from the FTIR analysis of the hydrogen bonding signature of this OIH for amide I and amide II, located in the carbonyl regions (1800-1500 cm<sup>-1</sup>). The FTIR spectra of 50  $\mu$ m thick OIH films, either undoped or doped with Eu<sup>3+</sup> at [Ocarb]/[Eu<sup>3+</sup>] = 3, are compared in **Figure 2**. Irrespective of the nature and chain length of the PE, the spectra of all the undoped OIHs presented an envelope composed of three components in the amide I region, at 1717-1721 (weak), 1678-1694 (average), and 1634-1641 cm<sup>-1</sup> (intense), together with a single intense band at 1559-1564 cm<sup>-1</sup> in the amide II region[15,19,47,48]. Comparison of the spectra evidenced that the changes induced by Eu<sup>3+</sup> doping could be classified into two families. In the first, the incorporation of Eu<sup>3+</sup> in the high molecular weight hybrids (U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup> of 2000 and 4000 g.mol<sup>-1</sup>) led to suppression of the two clearly separated bands and the appearance of a single and very intense band at 1608 cm<sup>-1</sup>, indicative of complexation of the Eu<sup>3+</sup> atom by the resonant urea linkage[19,49]. On the other hand, for the U-PPO:Eu<sup>3+</sup> with short polymer chain (230 g.mol<sup>-1</sup>), the suppression of the amide I and amide II envelope was not observed, suggesting that the Eu<sup>3+</sup> atoms were unable to disrupt the strongly hydrogen-bonded urea-urea aggregates[50], resulting in a negligible quantity of free urea groups available for Eu<sup>3+</sup> coordination. This could provide an explanation for the lower LMCT efficiency observed for U-PPO:Eu<sup>3+</sup> with shorter chains, but did not apply for U-PEO:Eu<sup>3+</sup>. In the latter case, the lower LMCT efficiency could be a consequence of the low amount of disordered U-PEO:Eu<sup>3+</sup>-urea associations resulting from the semicrystalline state of the U-PEO matrix [19]. Given these distinct classes of PL properties, an in-depth structural/optical study of the U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup> (2000 g.mol<sup>-1</sup>) was performed in order to elucidate the modifications induced by exposure to UV light, as discussed below.



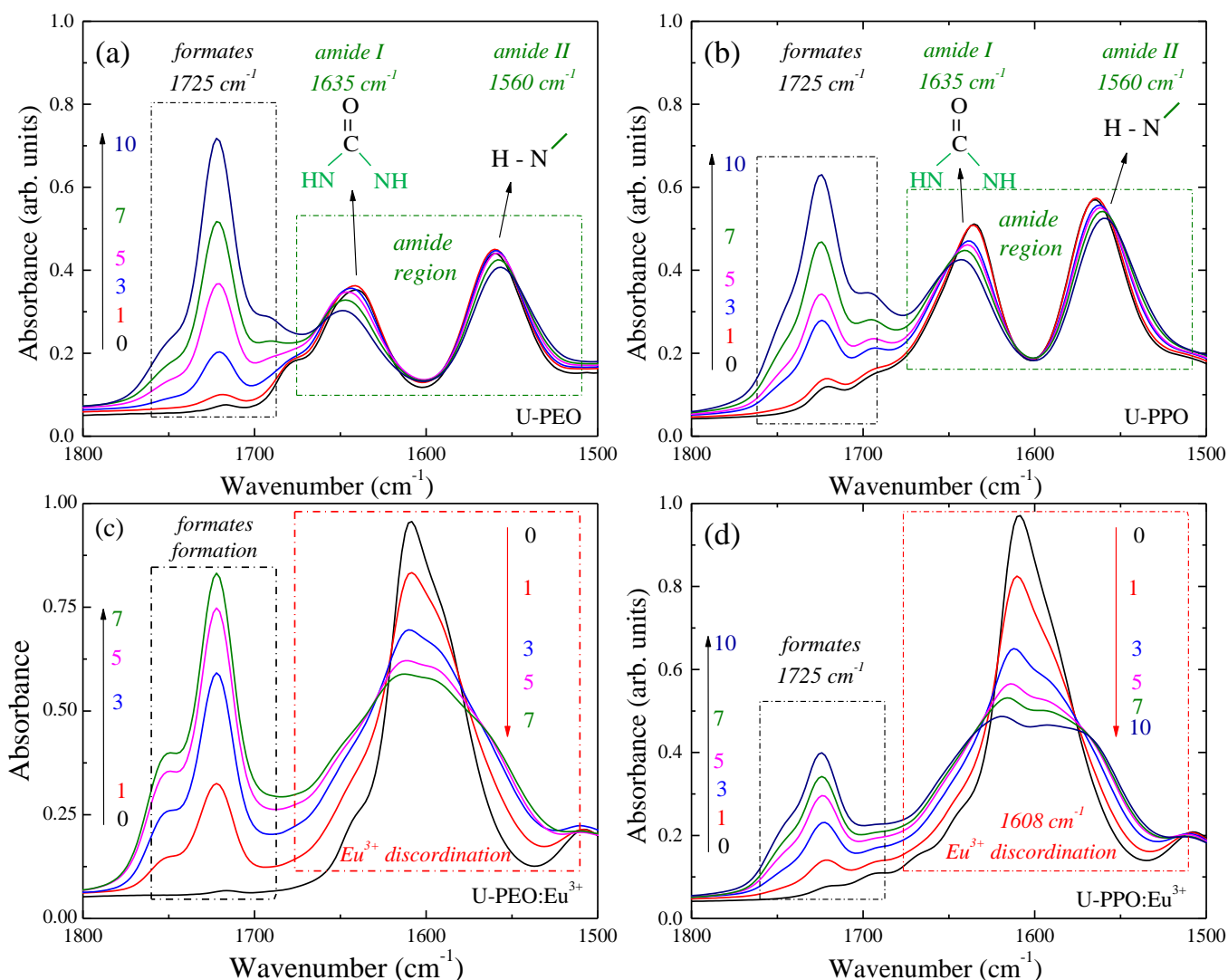
**Figure 2:** (2-column fitting image) Effect of  $\text{Eu}^{3+}$  doping on the amide I and II regions of the FTIR spectra of OIH films for (a) U-PEO (mw = 1900 g.mol<sup>-1</sup>) and for U-PPO with mw of (b) 230, (c) 2000, and (d) 4000 g.mol<sup>-1</sup>.

### 3.2 Effect of photooxidation on structural and nanostructural properties

The chemical modifications resulting from exposure of the undoped and  $\text{Eu}^{3+}$ -doped U-PEO and U-PPO hybrid samples to UV light irradiation resulted in important modifications of the FTIR spectra (**Figure 3**). The more important feature is the appearance of new bands in the carbonyl region at 1800-1500 cm<sup>-1</sup>. The spectra for the undoped U-PEO (**Figure 3a**) and U-PPO (**Figure 3b**) showed an intense band with a maximum at 1725 cm<sup>-1</sup>, together with a shoulder at 1750 cm<sup>-1</sup>. Previous studies of the photochemical behaviour of unhybridized polyethylene oxide (PEO) showed that these bands could be ascribed to the photooxidation products of the PEO chains, identified as formate (1725 cm<sup>-1</sup>) and ester (1750 cm<sup>-1</sup>)[33–36]. In the case of U-PPO (**Figure 3b**) the shoulder at 1750 cm<sup>-1</sup> was less evident, suggesting a predominance of decomposition to formate product. For OIH: $\text{Eu}^{3+}$  samples the band at around 1725 cm<sup>-1</sup> (**Figures 3c** and **3d** for U-PEO and U-PPO, respectively) was developed upon irradiation with faster rate for U-PEO: $\text{Eu}^{3+}$  and comparable rate for U-PPO, U-PEO and the U-PPO: $\text{Eu}^{3+}$ . This feature

suggests that the semicrystalline nature of the U-PEO coupled to the  $\text{Eu}^{3+}$  doping plays an important role on the photodegradation kinetics.

Moreover, slight modifications of the amide I ( $1635\text{ cm}^{-1}$ ) and amide II ( $1560\text{ cm}^{-1}$ ) bands of the urea linkage were observed during exposure of U-PEO and UPPO to UV light for 10 h. Decreased intensity and splitting of the bands indicated an increasing participation of amide groups in hydrogen-bonded aggregates[10]. In the case of U-PEO: $\text{Eu}^{3+}$  and UPPO: $\text{Eu}^{3+}$ , there was a decrease of the band at  $1608\text{ cm}^{-1}$ , which was assigned to a fingerprint of complexation of the  $\text{Eu}^{3+}$  ions by the urea bridges[20]. This suggested that photooxidation of the OIH: $\text{Eu}^{3+}$  samples led to a continuous dissociation of  $\text{Eu}^{3+}$  ions from the organic-inorganic interfaces of the U-PEO and U-PPO matrices.

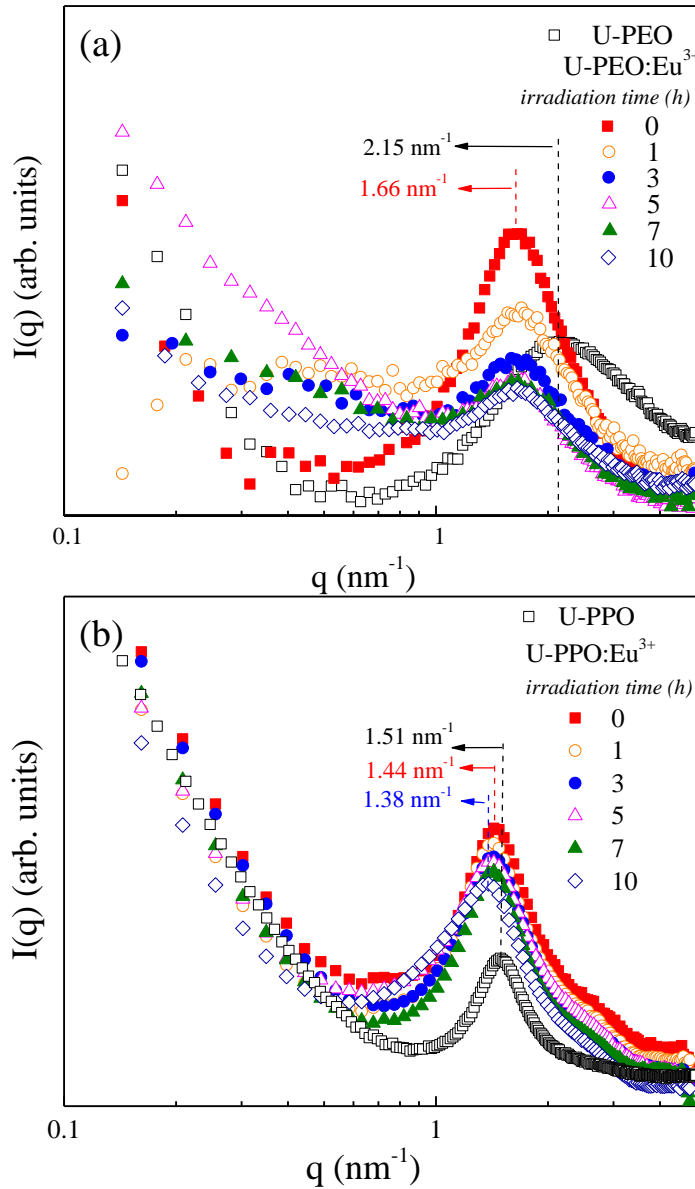


**Figure 3:** (2-column fitting image) Room temperature IR spectra from 1800 to 1500  $\text{cm}^{-1}$  of the OIH films during photooxidation: (a) U-PEO, (b) U-PPO, (c) U-PEO: $\text{Eu}^{3+}$ , and (d) U-PPO: $\text{Eu}^{3+}$ .

In order to identify the U-PE oxidation products,  $\text{NH}_3$  derivatization treatment of the photoaged films was performed according to the procedure described previously[33],[51]. The FTIR spectra of the photooxidized U-PEO: $\text{Eu}^{3+}$  and U-PPO: $\text{Eu}^{3+}$  films before and after 10 h reaction with  $\text{NH}_3$  are shown in **Figure SI.2**. After  $\text{NH}_3$  treatment, the FTIR spectrum of the photooxidized film showed a decrease of the absorption band at  $1725\text{ cm}^{-1}$ , together with the

formation of three new absorption bands at 1690, 1640, and 1565  $\text{cm}^{-1}$ , related to the formation of amide from the reaction of ester such as formate with  $\text{NH}_3$  ( $\text{HC=O-O-CH}_2\text{-} + \text{NH}_3 \rightarrow 2\text{HN-C=O-O-CH}_2\text{-}$ ) and to the dissociation of urea complexed to  $\text{Eu}^{3+}$  ions. The  $\text{NH}_3$  reaction confirmed that the both  $\text{Eu}^{3+}$ -loaded polyether hybrid materials produced similar photooxidation products.

The evolution of the OIH and OIH: $\text{Eu}^{3+}$  nanostructures, as a function of the UV light exposure time, was evidenced from the SAXS curves shown in **Figure 4**. The main effect observed following  $\text{Eu}^{3+}$  incorporation and photoaging was alteration of the peak attributed to the spatial correlation of regularly spaced siloxane cross-linked nodes[52]. After loading with  $\text{Eu}^{3+}$ , this peak was shifted from 2.15 to 1.66  $\text{nm}^{-1}$  for U-PEO and from 1.51 to 1.44  $\text{nm}^{-1}$  for U-PPO, corresponding to increases of the average correlation distance between adjacent siloxane nodes ( $D = 2\pi/q_{\text{max}}$ ). Following  $\text{Eu}^{3+}$  incorporation this distance is displaced from 2.92 to 3.78 nm for U-PEO and from 4.16 to 4.36 nm for U-PPO. The concomitant increased intensity of the correlation peak evidenced the increase of the electron density between the inorganic and the polymeric phases. These phenomena were due to the increase of the free volume of the polymeric matrix caused by the breaking of H-bonding associations, together with the complexation of  $\text{Eu}^{3+}$  by the carbonyl oxygen of the amide group[20].



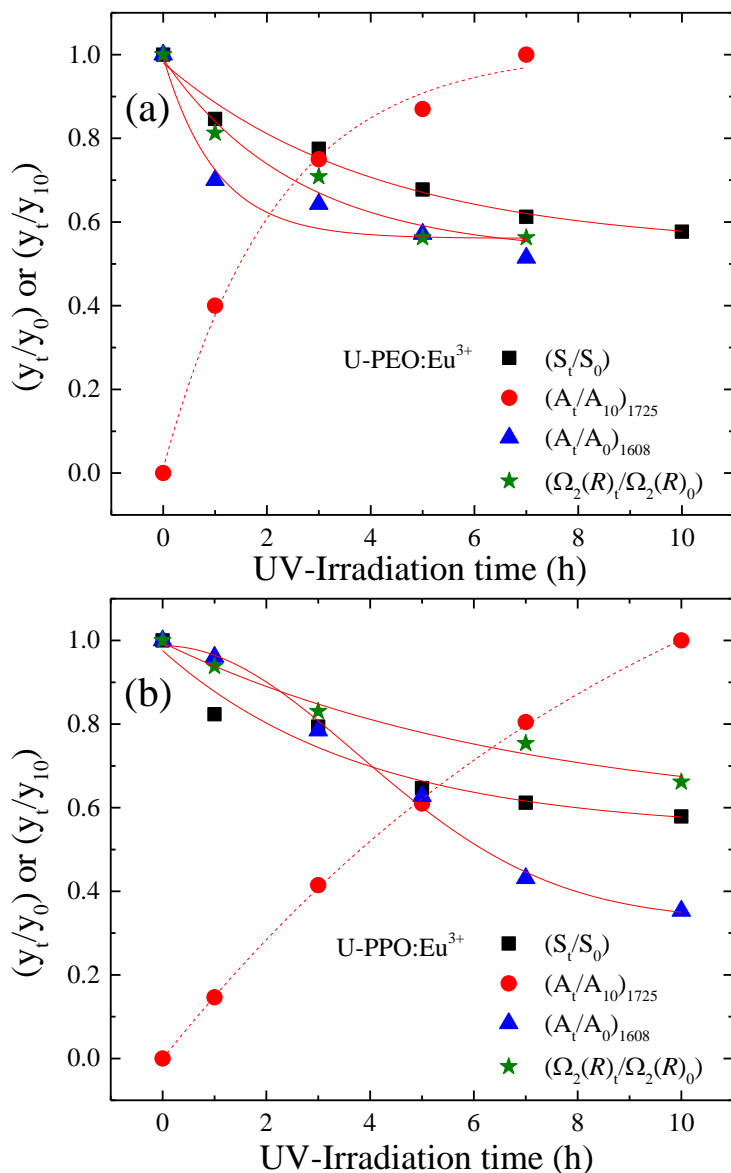
**Figure 4:** (single fitting image) SAXS curves, as a function of UV exposure time, for the  $\text{Eu}^{3+}$ -doped hybrids compared to the undoped ones: (a) U-PEO and (b) U-PPO

In the case of  $\text{U-PEO:Eu}^{3+}$ , the correlation peak intensity decreased as the UV exposure time increased. This could be ascribed to two effects, with (i)  $\text{Eu}^{3+}\text{-O}_{\text{carb}}$  dissociation decreasing the electronic density contrast between the inorganic and organic moieties, and (ii) scission of the PEO chains leading to a decrease of the number of spatially correlated siloxane nodes. It can be seen from the curves in **Figure 4a** that after 3 h of UV exposure, the  $I(q_{\text{max}})$  intensity was lower for  $\text{U-PEO:Eu}^{3+}$ , compared to the undoped U-PEO host. In contrast, even after 10 h of UV exposure, the intensity of the  $\text{U-PPO:Eu}^{3+}$  correlation peak remained higher, compared to the U-PPO matrix, as shown in **Figure 4b**. This difference revealed an important role played by the nature of the polymer (semicrystalline or amorphous; hydrophilic or hydrophobic), in terms of the sensitivity to photoaging. For semicrystalline and hydrophilic polymers (such as PEO), photoaging starts in the amorphous phase, due to the easy diffusion of oxygen through the unordered chains[53]. Accordingly, the observed decrease of the electron density contrast

between the siloxane nodes and the polymeric matrix indicated a more compact structure resulting from increase of the crystalline phase fraction[54], caused by the scission of polyether chains in the amorphous phase. However, the U-PEO:Eu<sup>3+</sup> nanostructure was preserved, as evidenced from the maintenance of the spatial correlation between the regularly spaced siloxane cross-linked nodes.

In order to understand the nanostructural modifications caused by the photoaging, it is important to note that after 10 h of UV exposure, the absorbance (A) value for the band at 1725 cm<sup>-1</sup> (see **Figure 3**), ascribed to the photoaging products, was lower for U-PPO:Eu<sup>3+</sup> (0.44) compared to U-PEO:Eu<sup>3+</sup> (0.83). This indicated that the more intense I(q<sub>max</sub>) decrease observed for U-PEO:Eu<sup>3+</sup> was due to degradation of the polyether chains. However, the effect of chain scission on the nanostructure of the amorphous U-PPO:Eu<sup>3+</sup> could be seen from the slight displacement of the q<sub>max</sub> position from 1.44 to 1.37 nm<sup>-1</sup>, due to the formation of more compact structures[54]. In contrast, no systematic changes were observed in the U-PEO:Eu<sup>3+</sup> correlation peak position following UV exposure, suggesting that the average correlation distance between two adjacent siloxane nodes of the already compact structure was less influenced by this effect, due to the semicrystalline nature of the material. Despite this complexity, the evolution of the growth of the relative area (A<sub>t</sub>/A<sub>10</sub>) of the formate band at 1725 cm<sup>-1</sup>, as well as the decreases of the amide band at 1608 cm<sup>-1</sup> and the SAXS peak (S<sub>t</sub>/S<sub>0</sub>), could be described by exponential growth ( $y_t/y_{10} = 1 + \exp(k.t)$ ) or decay ( $y_t/y_0 = 1 - \exp(k.t)$ ) laws, as shown in **Figure 5**. This approach evidenced that for U-PPO:Eu<sup>3+</sup>, all the events occurred concomitantly, with comparable time constants (k) (**Table 1**), suggesting the existence of correlations between the processes occurring at nanoscopic and molecular length scales.

In contrast, for U-PEO:Eu<sup>3+</sup>, the different time constant values indicated the existence of a faster process that was probably governed by the facilitated transport of the oxidative species through the interface and the amorphous moieties of the semicrystalline U-PEO. This hypotheses was confirmed from the fusion enthalpies values founded from DSC measurements, with evidenced a decrease from 21% to 16% of the U-PEO crystalline phase fraction upon Eu<sup>3+</sup> addition (**Figure SI.3a**). This decrease of crystallinity degree favors the diffusion of oxygen through the unordered chains, making ease the initiation of photooxidative degradation of the amorphous phase[55]. This fast degradation process leads to formation of low molecular weight products which can diffuse out the polymer matrix[56,57], decreasing the mass of amorphous polymer and increasing the crystallinity degrees during the first 3 h of UV light irradiation. The correlation between the increase of the crystallinity degree and the deaccelerated growth of formate production observed in this initial stage (**Figure SI.3a and b**) is consistent with the explanation based on the facilitated transport of oxidative species. In addition, the second increases of the format production rate coupled with a pronounced decrease of crystallinity degrees was observed in the advanced stage of photodegradation of U-PEO:Eu<sup>3+</sup>. These correlated features observed in the early and advanced stages are driven by the association of europium salt with the ether type oxygens atoms occurring at the initial addition? of europium salt and also by the continuous photo dissociation of urea-Eu<sup>3+</sup> complexes.

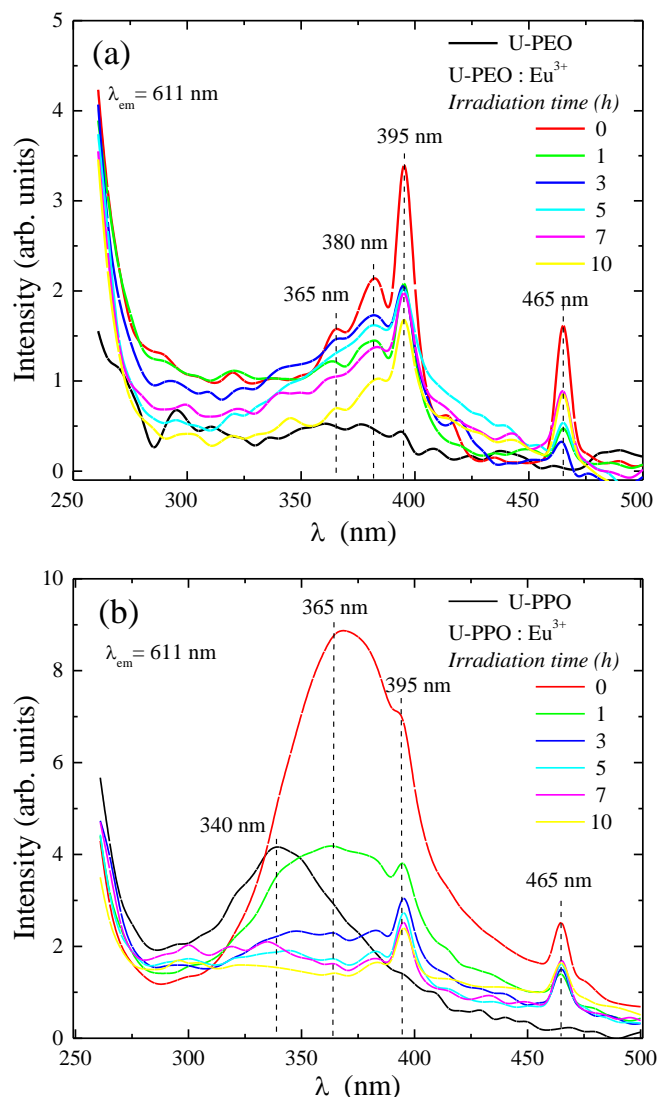


**Figure 5:** (single fitting image) Evolution of the relative areas of the SAXS correlation peak ( $S_t/S_0$ ), the formate formation band at  $1725 \text{ cm}^{-1}$  ( $A_t/A_{10})_{1725}$ , the  $\text{Eu}^{3+}$ -OIH decomplexation fingerprint band at  $1608 \text{ cm}^{-1}$  ( $A_t/A_0)_{1608}$ , and the Judd-Ofelt parameter ratio  $(\Omega_2(R)_t/(\Omega_2(R)_0))$ , as a function of the time of photooxidation, for (a) U-PEO:Eu<sup>3+</sup> and (b) U-PPO:Eu<sup>3+</sup>. The continuous and dashed lines show the best fits of the exponential decay ( $y_t/y_0 = 1 - \exp(-k.t)$ ) and growth ( $y_t/y_{10} = 1 - \exp(-k.t)$ ) laws, respectively.

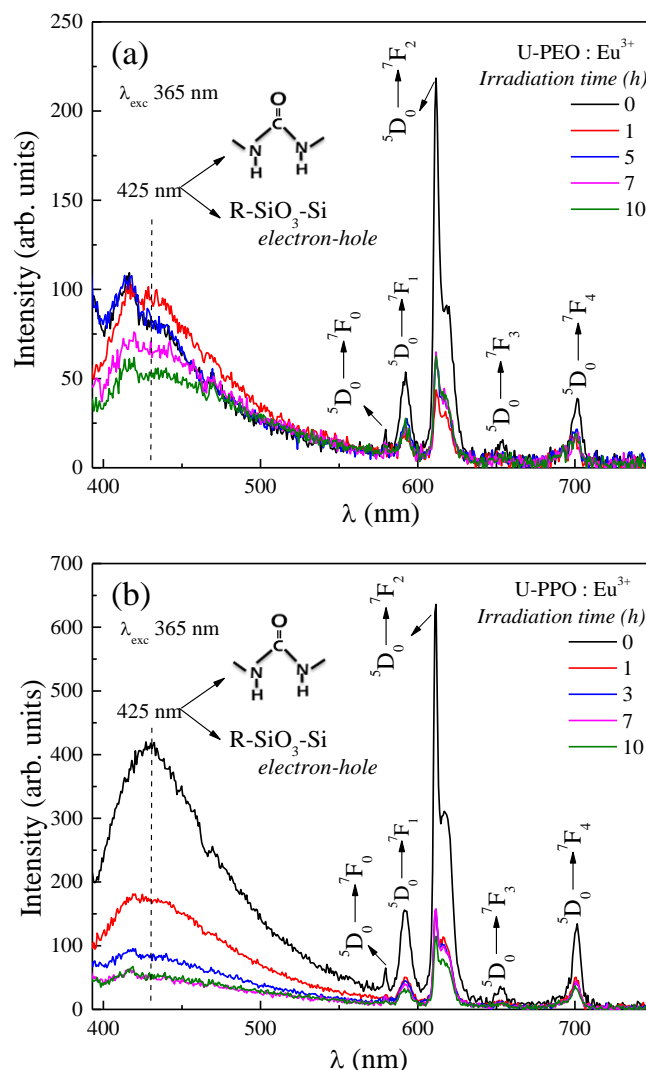
### 3.3 Effect of photooxidation on luminescence properties

The effects of the photodegradation on the photoluminescence (PL) properties of the U-PEO:Eu<sup>3+</sup> and U-PPO:Eu<sup>3+</sup> samples were investigated in the visible range. The photooxidation of the polymer and the decomplexation of  $\text{Eu}^{3+}$  ions from the hybrid host matrix could be confirmed by following the excitation (**Figure 6**) and emission (**Figure 7**) spectra recorded as a function of the UV irradiation time. As shown in **Figure 1b**, the antenna effect between the organic ligand and the  $\text{Eu}^{3+}$  ions was only efficient for the U-PPO:Eu<sup>3+</sup> sample, resulting in a broad band centred at  $365 \text{ nm}$  in the excitation spectrum. However, the excitation spectra for the photooxidized U-PPO:Eu<sup>3+</sup> (**Figure 6b**) revealed a dramatic decrease of intensity after 1 h of exposure to UV irradiation. This was accompanied by increased intensities of the bands located at  $380$ ,  $395$ , and

465 nm, attributed to the  $\text{Eu}^{3+}$  intra-configurational transitions  ${}^7\text{F}_0 \rightarrow {}^5\text{L}_7$ ,  ${}^7\text{F}_0 \rightarrow {}^5\text{L}_6$ , and  ${}^7\text{F}_0 \rightarrow {}^5\text{D}_2$ , respectively[17]. These results were consistent with degradation of the organic moiety and consequent dissociation of  $\text{Eu}^{3+}$  ions from the hybrid host matrix. The low intensity found for U-PEO: $\text{Eu}^{3+}$  could be explained by the nature of the polymer, especially the hydrophilic properties, with the water molecules that are typically found in interaction with the PEO backbone acting to quench the  $\text{Eu}^{3+}$  photoluminescence[58].



**Figure 6:** (single column fitting image) Excitation spectra of (a) U-PEO: $\text{Eu}^{3+}$  and (b) U-PPO: $\text{Eu}^{3+}$  exposed to UV irradiation for different times.



**Figure 7:** (single column fitting image) Emission spectra recorded at RT for (a) U-PEO: $\text{Eu}^{3+}$  and (b) U-PPO: $\text{Eu}^{3+}$  exposed to UV irradiation for different times.

The emission spectra recorded with excitation at 365 nm (**Figure 7**) revealed decreased intensities of both the broad band at 425 nm and the narrow lines ( ${}^5\text{D}_0 \rightarrow {}^7\text{F}_J$ ). The decreased intensity of the blue band could be attributed to photooxidation of the polyether chains, which decreased the number of spatially correlated siloxane nodes, as revealed by the reduced area of the SAXS peak shown in **Figure 5**. This hindered the complexation of  $\text{Eu}^{3+}$  to urea bridges, as observed for the shorter chain U-PPO: $\text{Eu}^{3+}$  (**Figure 2a**). The photodegradation of the organic moieties led to OIH- $\text{Eu}^{3+}$  dissociation at the organic-inorganic interface. These phenomena could

be detected from the change of the intensity of the so-called “hypersensitive” transition[59,60],  $^5D_0 \rightarrow ^7F_2$ , as a function of the UV irradiation time.

The luminescence intensity ratio,  $R = (^5D_0 \rightarrow ^7F_2)/(^5D_0 \rightarrow ^7F_1)$ , is known as the asymmetry ratio. Higher values of  $R$  indicate lower symmetry around the trivalent europium ions. The Judd-Ofelt parameter ( $\Omega_2$ ) and  $R$  provide similar information about the physical nature of the bonding between  $Eu^{3+}$  ions and the surrounding ligands, explaining the short range effects on the local structure around the  $Eu^{3+}$  ions[61–63]. Similar values of the  $\Omega_2$  ( $R$ ) parameter ( $\sim 4$ ) were observed for the non-irradiated PEO: $Eu^{3+}$  and PPO: $Eu^{3+}$  samples, indicating that the  $Eu^{3+}$  atoms occupied very similar coordination (semicrystalline or amorphous)[64]. These  $\Omega_2$  ( $R$ ) values were higher than those observed elsewhere for  $Eu^{3+}$  incorporated in several oxide glasses and polymers, including PEO and PPO[9], indicating that the symmetry of the  $Eu^{3+}$  sites was distorted[20][64]. However, after 10 h of UV exposure, the  $\Omega_2$  ( $R$ ) values decreased from  $\sim 4$  to  $\sim 3$  and  $\sim 2$  for U-PPO: $Eu^{3+}$  and U-PEO: $Eu^{3+}$ , respectively. These decreases occurred continuously during the irradiation time, following trends comparable to the decreasing intensity of the broad blue band. This behaviour suggested that the photooxidation of the polyether chains led to modification of the surrounding  $Eu^{3+}$  ligands, resulting in decreases of the asymmetry parameter.

The most noteworthy finding was the good correlation between the nanoscopic features and the local environment around the  $Eu^{3+}$  emitter, as evidenced by the comparable temporal evolution of the relative SAXS peak area values ( $S_t/S_0$ ) and  $\Omega_2(R)_t/\Omega_2(R)_0$  (**Figure 5**). These two functions were described by exponential decay laws with comparable time constants (**Table 1**), clearly evidencing the sensitivity of the photoluminescence to the elimination/formation of the hydrogen-bonded urea-polyether array during photooxidation of the PEO and PPO chains. These nanoscopic properties arising from the photoinduced degradation of ureasil-polyether can be exploited in the development of new transparent, flexible, and biocompatible UV dosimeters and sensors.

**Table 1:** Time constants ( $k$ ) of the fitted exponential growth or decay law curves ( $y_t/y_{10} = 1 \pm \exp(k.t)$ ) for the relative area of the SAXS correlation peak, the Judd-Ofelt parameter ratio ( $\Omega_2(R)_t/\Omega_2(R)_0$ ), the  $Eu^{3+}$ -OIH decomplexation fingerprint band at  $1608\text{ cm}^{-1}$ , and the formate formation band at  $1725\text{ cm}^{-1}$  (undoped in brackets), as a function of photooxidation time, for U-PEO: $Eu^{3+}$  and U-PPO: $Eu^{3+}$ .

Relative area from	Time constant $K$ ( $h^{-1}$ )	
	U-PEO: $Eu^{3+}$	U-PPO: $Eu^{3+}$
SAXS	0.25	0.26
PL ( $\Omega_2$ ( $R$ ))	0.39	0.15
IR @ $1608\text{ cm}^{-1}$	0.97	0.07
IR @ $1725\text{ cm}^{-1}$	0.45 (0.33)	0.10 (0.10)

## Conclusions

Different ureasil-polyether materials based on polyethylene oxide (U-PEO) and polypropylene oxide (U-PPO), undoped and loaded with  $Eu^{3+}$ , were successfully prepared by the sol-gel route in the forms of monoliths and thin films ( $50\text{ }\mu\text{m}$  thick). The FTIR spectra confirmed the preferential complexation of  $Eu^{3+}$  by the carbonyl group ( $O_{carb}$ ) of the urea linkage. The best photoluminescence properties were found for a doping level equivalent to  $[O_{carb}]/[Eu^{3+}] = 3$ . The PL results evidenced loss of the ligand-to-metal charge transfer when the  $Eu^{3+}$  cations were

unable to disrupt the strong hydrogen-bonded urea/urea aggregates formed using shorter chain U-PPO (230 g.mol<sup>-1</sup>) and semicrystalline U-PEO (1900 g.mol<sup>-1</sup>).

Formate was the main photoproduct formed following exposure of U-PEO and U-PPO to artificial UV light, revealing a photooxidation process quite similar to the observed for conventional PEO and PPO materials. In addition, the urea linkages present at the organic–inorganic interface are weakly dependent on the time of UV light exposure as demonstrated by slight modifications of the amide bands.

The Eu<sup>3+</sup> addition to semicrystalline U-PEO allows to an increase of the amorphous phase with accelerate initial formation of the formate, the main photooxidation product. The photoinduced dissociation of Eu<sup>3+</sup> complexed to the urea and consequent association to the ether oxygen atoms of the polymer chains favors the propagation of depolymerisation reactions in the advanced stages.

The continuous recovery of the hydrogen bond of the amide structure evidenced the Eu<sup>3+</sup> decomplexation induced by exposure to intense UV light. This induces the decreased intensity of the broad green-blue band, as revealed by PL analyses conducted as a function of the UV exposure time.

The photooxidation of the polyether chains resulted in concomitant decreases of the spatially correlated siloxane nodes and the Judd-Ofelt parameter,  $\Omega_2$  (R). The kinetic evolution of these parameters could be described by exponential laws with comparable time constants, evidencing the sensitivity of the Eu<sup>3+</sup> photoluminescence to the elimination/formation of the hydrogen-bonded urea-polyether array during the photooxidation of the PPO chains. In contrast, for U-PEO:Eu<sup>3+</sup> different time constant values were obtained, due to the existence of a faster process governed by the facilitated transport of the oxidative species through the amorphous moieties of the semicrystalline U-PEO.

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