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# The nature and partitioning of invisible gold in the pyrite-fluid system

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31 **Keywords:**

32 Invisible gold; pyrite; sulfur radical ions; high-energy resolution fluorescence-detected x-ray  
 33 absorption spectroscopy (HERFD-XAS); solubility; gold deposits

34

35 **Highlights:**

- 36 • Au in sulfide-sulfate solutions is complexed with HS<sup>-</sup>, S<sub>3</sub><sup>2-</sup> and S<sub>2</sub><sup>-</sup> ions
- 37 • In As-poor systems, invisible gold in pyrite is present as Au(0) and Au(I)
- 38 • Au(I) in pyrite occurs as S-Au-S<sub>n</sub> moieties similar to those in the fluid
- 39 • Pyrite-fluid Au(I) partition coefficient ( $D_{py/fl}$ ) is  $0.15 \pm 0.07$  at 450°C and 700 bar
- 40 • Au(I) uptake from fluid by pyrite is enhanced with decreasing temperature
- 41 •  $D_{py/fl}$  values allow predictions of tenors of chemically bound gold in pyrite ore

42

43 **Abstract:**

44 The most characteristic feature of hydrothermal deposits of gold is its intimate association with  
 45 pyrite. Microscopically visible gold occurs in pyrite ore as metal particles of >0.1 micron in size,  
 46 together with so called “invisible” gold, undetectable by conventional microscopic methods. The  
 47 chemical, redox and structural state of this invisible gold and the mechanisms of its incorporation  
 48 into pyrite remain both inconsistent and controversial since the dawn of economic geology. To  
 49 clarify these issues, we performed laboratory experiments to simulate interactions of gold-bearing  
 50 sulfur-rich hydrothermal fluids with arsenic-free pyrite at temperatures from 350 to 450°C and  
 51 pressures from 400 to 700 bar, typical of the formation conditions of many types of gold deposits.  
 52 Gold solubility was measured in these fluids as a function of sulfur speciation and acidity. Gold  
 53 redox and structural state in pyrite was characterized by high-energy resolution fluorescence-  
 54 detected x-ray absorption spectroscopy (HERFD-XAS), together with more traditional analytical  
 55 techniques such as scanning electron microscopy (SEM), x-ray diffraction (XRD), electron probe  
 56 micro analysis (EPMA), laser ablation inductively coupled plasma mass spectrometry (LA-  
 57 ICPMS), and inductively coupled plasma atomic emission spectrometry (ICP-AES). Results show  
 58 that dissolved Au in sulfide-sulfate solutions forms complexes with hydrogen sulfide, and tri- and  
 59 di-sulfur radical ions whose amounts depend mostly on the fluid pH and total sulfur concentration.  
 60 Invisible gold in pyrite occurs as Au metal submicron- to nano-sized particles and chemically  
 61 bound Au(I) in the form of (poly)sulfide clusters composed on S-Au-S linear units, similar to  
 62 those in aqueous complexes. Our findings contest the common belief that Au(I) substitutes for Fe  
 63 and/or S in the structure of As-poor pyrite. The partition coefficient of Au(I) between pyrite and  
 64 the fluid,  $D_{py/fl}$ , is determined to be  $0.15 \pm 0.07$  at 450°C in a wide range of Au fluid phase  
 65 concentrations (10-1000 ppm), but much higher  $D_{py/fl}$  values, between 10 and 50, are found at  
 66 350°C. These Au partitioning trends coupled with the new data on Au molecular environment in  
 67 pyrite suggest a control of Au(I) incorporation in the mineral by a chemisorption step.  
 68 Extrapolated to Au contents of hydrothermal fluids of the Earth’s crust which are typically below  
 69 1 ppm, our  $D_{py/fl}$  values reproduce fairly well the natural Au tenors in As-poor pyrites (~0.1-1 ppm  
 70 Au), which are 100 to 1000 times lower than those typically observed in arsenian pyrites and  
 71 arsenopyrites (10-1000 ppm Au at As tenors of 0.01-10 wt%). Our results thus indirectly highlight  
 72 a key role played by arsenic in gold enrichment in As-bearing iron sulfide ore, a role that yet  
 73 remains to be fully understood and quantified.

74

## 1. Introduction

75

76 The most characteristic feature of hydrothermal deposits of gold is its intimate  
77 association with iron sulfide minerals, pyrite and arsenopyrite, in which Au concentrations may  
78 reach 100 to 1000 ppm, i.e. almost million times the Au average crustal abundance. Gold affinity  
79 for these minerals has been recognized since the first half of the 20<sup>th</sup> century (e.g., Goldschmidt,  
80 1937). Both ‘visible’ native gold (i.e., optically detectable particles larger than 0.1-1  $\mu\text{m}$ ) in  
81 physical association with pyrite, and ‘invisible’ (i.e., optically undetectable) gold, possibly  
82 incorporated into the mineral structure, occur (e.g., Boyle, 1969; Cathelineau et al., 1989; Cook  
83 and Chryssoulis, 1990; references therein). Invisible gold, also termed refractory gold, is not only  
84 difficult to identify, but also hard to recover from ore (e.g., Yang et al., 1998; Adams, 2005;  
85 Marsden and House, 2009). Such refractory gold poses problems both for accurate estimates of the  
86 gold distribution and endowment of a deposit and for efficient extraction of the metal from iron  
87 sulfide ore. Resolving these issues requires knowledge of the nature, distribution and mechanisms  
88 of incorporation of invisible gold in pyrite, which all yet remain enigmatic. The present study was  
89 designed in an attempt to better understand the gold-pyrite relationships, using a novel integrated  
90 experimental, spectroscopic and modeling approach.

91

92 The literature on Au-bearing pyrite over the four last decades is full of discussions about  
93 the chemical and redox state of invisible gold. Based on a variety of analytical techniques such as  
94 scanning electron microscopy (SEM), high resolution transmission electron microscopy  
95 (HRTEM), secondary ion mass spectrometry (SIMS), laser ablation inductively coupled plasma  
96 mass spectrometry (LA-ICPMS), electron probe microanalyzer (EPMA), Mössbauer spectroscopy  
97 (MBS), proton induced x-ray emission (PIXE), x-ray photoelectron spectrometry (XPS), x-ray  
98 absorption spectroscopy (XAS), atom probe microscopy (APM), atomic force microscopy (AFM),  
99 tunneling spectroscopy (STS), coupled with analyses of elemental distribution and inter-element  
100 correlations, most studies agree that invisible gold occurs in two major chemical states: an  
101 elemental form as nanometer-size particles of Au(0) or alloys with some other trace metals (e.g.,  
102 Bi, Te, Ag), and a chemically bound form (e.g., Marion et al., 1986; Arehart et al., 1993; Scaini et  
103 al., 1998; Friedl et al., 1995; Fleet and Mumin, 1997; Genkin et al., 1998; den Besten et al., 1999;  
104 Simon et al., 1999; Tauson, 1999; Cabri et al., 2000; Palenik et al., 2004; Reich et al., 2005, 2006,  
105 2010; Mikhlín et al., 2011; Deditius et al., 2011; Fougèrouse et al., 2016; Trigub et al., 2017).  
106 However, despite significant recent advances in analytical and microscopic methods listed above,  
107 the nature of this latter type of gold and factors controlling its incorporation in pyrite remain both  
inconsistent and controversial. For example, the redox state of chemically bound gold varies from

108 Au<sup>3+</sup> to Au<sup>1+</sup>, and its coordination number from 2 to 6, according to different authors (e.g., see  
109 Cabri et al., 2000 for review); such large variations make it difficult to attribute to Au an exact  
110 structural position in pyrite and to identify the nature of coordinating atoms (e.g., Au, Fe, S). The  
111 situation is further complicated by the presence of other minor and trace elements (Cu, Sb, Ni, Co,  
112 Ag, Bi, Se, Te) and, in particular, of arsenic with which gold is systematically correlated in  
113 arsenian pyrite (e.g., Reich et al., 2005). These elements, often present in far greater  
114 concentrations (>0.1-1.0 wt%) than Au (typically <0.01-0.1wt%), greatly complicate the  
115 assessment, using inter-element correlations, of the Au oxidation state and atomic substitutions in  
116 pyrite. The use of direct redox-sensitive spectroscopic methods such as X-ray photoelectron, X-  
117 ray absorption or Mössbauer spectroscopy is often limited by low concentrations, poor selectivity  
118 for Au in the presence of As and other elements, and a lack of adequate standards.

119 As a result of these limitations, the mechanisms of Au incorporation in hydrothermal  
120 pyrite and arsenopyrite and the role of arsenic and other elements played in this process still  
121 remain hypothetical. For example, some earlier studies invoked gold transport by the fluid phase  
122 in the form of Au-As-S (Sb, Se, Te) aqueous complexes and their co-precipitation with ferrous  
123 iron (Boyle, 1969; Boiron et al., 1989). Other studies suggested chemisorption of dissolved Au as  
124 Au-S(As) complexes at As-rich, Fe-deficient sites (Mao, 1991; Fleet and Mumin, 1997; Cepedal  
125 et al., 2008), electrochemically driven adsorption of negatively charged Au(HS)<sub>2</sub><sup>-</sup> complexes on  
126 semiconducting As-pyrite and arsenopyrite surfaces (Mironov et al., 1981; Moller and Kersten,  
127 1994; Widler and Seward, 2002), or Au<sup>3+</sup> and Au<sup>+</sup> precipitation on sulfide surfaces due to Au  
128 reduction (Bancroft and Hyland, 1990; Scaini et al., 1998, Maddox et al., 1998; Mikhlin et al.,  
129 2011). However, these hypotheses are mostly based on low-temperature experiments (<100°C),  
130 some of them involving Au aqueous complexes (AuCl<sub>4</sub><sup>-</sup>, Au(CN)<sub>2</sub>Br<sub>2</sub><sup>3-</sup>), which are unlikely at  
131 hydrothermal conditions (e.g., see Pokrovski et al., 2014 for a review of Au speciation in  
132 hydrothermal fluids). These limitations complicate the extrapolation of the results of such  
133 experiments to the far greater temperatures of Au-bearing pyrite formation in nature (>150-  
134 200°C). Different isomorphic substitution models of Au in the pyrite structure were also proposed,  
135 including coupled substitutions of As<sup>3+</sup> + Au<sup>+</sup> for 2 Fe<sup>2+</sup>, As<sup>3+</sup> + Au<sup>3+</sup> for 3 Fe<sup>2+</sup> (Deditius et al.,  
136 2008), Cu<sup>+</sup> + Au<sup>3+</sup> for 2 Fe<sup>2+</sup> (Chouinard et al., 2005), or Au<sup>+</sup> + S<sup>2-</sup> for Fe<sup>2+</sup> + HS<sup>-</sup> (Tauson, 1999),  
137 but analytical methods capable of unambiguously supporting these models are lacking. Note that  
138 the geometry of the available structural sites in pyrite composed of Fe(S<sub>6</sub>) octahedrons and  
139 S(Fe<sub>3</sub>S) distorted tetrahedrons (Le Pape et al., 2018) is not favorable for an energetically stable  
140 incorporation of both Au<sup>+</sup> and Au<sup>3+</sup> ions that have linear and square planar coordination  
141 geometries in most inorganic and organometallic compounds (e.g., Cotton and Wilkinson, 1988).

142 Furthermore, common precipitation mechanisms of As and Au due to fluid reduction and/or  
143 cooling, or coupled As-Au redox reactions on the pre-existing (arseno)pyrite have also been  
144 considered (e.g., Heinrich and Eadington, 1986; Pokrovski et al., 2002), but they remain  
145 qualitative owing to the complexity of the fluid composition and evolution in hydrothermal  
146 systems and the interplay of multiple factors controlling mineral solubility such as temperature  
147 ( $T$ ), pressure ( $P$ ), redox, acidity (pH), and fluid composition (Pokrovski et al., 2014).

148 It follows from this brief overview that a deeper understanding of the gold-pyrite(-  
149 arsenic) relationships in hydrothermal systems awaits experimental studies at pertinent laboratory-  
150 controlled conditions, coupled with Au-selective in situ spectroscopic techniques. In the recent  
151 years, advances in synchrotron-based high-resolution X-ray absorption/fluorescence spectroscopy  
152 (e.g., Proux et al., 2017; Manceau et al., 2018; Le Pape et al., 2018; references therein) have  
153 provided an unprecedented gain both in sensitivity (metal detection limits <1 ppm) and selectivity  
154 (ability to isolate the trace metal fluorescence signal in a complex mineral/organic matrix), now  
155 enabling to directly “see” invisible gold in iron sulfide minerals. In this study, we designed  
156 hydrothermal experiments aimed at investigating the Au-pyrite interactions in model S-bearing  
157 fluid-pyrite systems at temperatures (350-450°C) and pressures (400-700 bar) typical of those of  
158 various types of gold deposits, as a function of gold and sulfur speciation in the fluid phase. The  
159 Au-bearing pyrites resulting from such experiments were investigated by high-resolution X-ray  
160 absorption spectroscopy (HR-XAS) coupled with more “traditional” methods (EPMA, SEM,  
161 XRD, LA-ICPMS, ICP-AES) to reveal the state of invisible gold in pyrite and to derive pyrite-  
162 fluid partition coefficients for chemically bound Au. The results allow predictions of invisible Au  
163 contents in ore, and more reliable identification of the factors leading to Au enrichment in iron  
164 sulfide minerals. Although the present work deals with As-free systems, the results also help  
165 indirectly assess the role of arsenic, thus providing the foundation for future experimental and  
166 analytical studies on far more complex As-bearing systems.

167

168

## 2. Materials and methods

169

### 2.1. *Experimental strategy and conditions*

171

172 The primary goal of this exploratory experimental study was to investigate the state of gold  
173 incorporated into pyrite under conditions pertinent to natural hydrothermal systems forming gold  
174 deposits. Gold incorporation in pyrite is expected to occur via *i*) exchange of an Au-bearing

175 aqueous fluid with already existing sulfide minerals, and/or *ii*) simultaneous precipitation of gold  
176 and pyrite (and other metals) from an aqueous fluid, driven by changes in  $T$ ,  $P$  or fluid  
177 composition. Such fluid-mineral interactions are very difficult to track from the observation of  
178 complex natural mineral assemblages; alternatively, they may be simulated experimentally in  
179 simplified but well-controlled systems. In this study, we designed two types of experiments to  
180 reproduce the two major types of expected Au-pyrite interactions, fluid-mineral exchange and co-  
181 precipitation.

182 Most previous laboratory syntheses of Au-bearing (arseno)pyrite were performed either in  
183 dry (salt-flux method) or water-bearing systems at temperatures above 450-500°C from mixtures  
184 of native elements with no control of redox potential, pH, and sulfur speciation, or from pre-  
185 existing pyrite and fluid with excess of native sulfur leading to extremely acidic solutions (e.g.,  
186 Fleet and Mumin, 1997; Tauson, 1999; Trigub et al., 2017). Such conditions are rather far from  
187 those encountered in natural Au-forming hydrothermal systems, which are characterized by  
188 moderate temperatures ( $\leq 450^\circ\text{C}$ ), and near-neutral pH and redox environments allowing  
189 coexistence of sulfate and sulfide in the fluid phase (Pokrovski et al., 2014; 2015). Thus, in  
190 contrast to those works, our experiments were conducted in S-bearing gold-saturated fluids at  
191 temperatures 350-450°C and pressures 400-700 bar, and in the presence of sulfide and sulfate over  
192 a pH range 4 to 8. Our experiments allow accurate control on the concentration and speciation of  
193 Au and S in the fluid phase, enabling to link, for the first time, both Au status and degree of  
194 partitioning in pyrite to the fluid composition and Au aqueous speciation. A scheme summarizing  
195 the experimental conditions and major steps is given in Fig. 1.

196 The run temperatures were chosen to be in the upper range of the common temperature  
197 window of most Au-bearing pyrite deposits (150-500°C) to facilitate mineral-fluid exchange and  
198 attainment of equilibrium within the short-time laboratory scale. Aqueous solutions of potassium  
199 thiosulfate ( $\text{K}_2\text{S}_2\text{O}_3$ ) with HCl or KOH to adjust the pH were used in both types of experiments. In  
200 aqueous solution above 200°C, thiosulfate breaks down within hours to sulfate ( $\text{HSO}_4^-/\text{SO}_4^{2-}$ ) and  
201 sulfide ( $\text{H}_2\text{S}/\text{HS}^-$ ) as the dominant sulfur species with variable amounts of the radical sulfur ions  
202 ( $\text{S}_3^{\cdot-}$  and  $\text{S}_2^{\cdot-}$ ) and molecular sulfur ( $\text{S}_n^0$ ), whose fractions depend on pH, temperature and S total  
203 amount. Their concentrations can be predicted using the thermodynamic properties derived from  
204 in-situ Raman spectroscopy data on similar thiosulfate solutions (Pokrovski and Dubessy, 2015).  
205 Importantly, our S-bearing systems enable both  $f_{\text{O}_2}$  and pH buffering via sulfide-sulfate and  
206 sulfate-hydrogen sulfate equilibria (Jacquemet et al., 2014), which is essential for interpretation of  
207 gold solubility and speciation and can be quantified using thermodynamic modeling (section 2.8).  
208 The choice of elevated S concentrations in this study (1-6 wt% total S), as compared to most

209 natural systems, was dictated by the necessity to solubilize sufficient amount of gold into the fluid  
210 to favor fluid-mineral exchange and Au incorporation into pyrite in concentrations sufficient to be  
211 analyzable by the methods chosen, and in particular by HR-XAS. Our experimental conditions  
212 enable an accurate control on equilibrium and robust extrapolations of the obtained structural and  
213 partitioning data to natural systems. Metallic gold was used as the Au source in all experiments.  
214 Gold solubility in such S-rich solutions was shown to be controlled by hydrogensulfide ( $\text{HS}^-$ ) and  
215 trisulfur ion ( $\text{S}_3^{2-}$ ) complexes, depending on pH and S concentration (Pokrovski et al., 2015). Thus,  
216 the present experiments not only provide an opportunity to verify these and other alternative Au  
217 speciation models in a wide pH range, but also they enable, for the first time, the exploration of  
218 the impact of changes in aqueous speciation of sulfur and gold on Au incorporation in pyrite.

219

## 220 *2.2. Gold bearing fluid-pyrite exchange experiments*

221

222 Pyrite-fluid exchange experiments were conducted at 450°C and 700 bar in batch Ti-  
223 based alloy reactors of ~20 mL volume (Fig. 1). Details about the reactor design and handling are  
224 given in Pokrovski et al. (2002). The reactor was loaded with a weighed amount of aqueous ~0.5  
225 mol/kg (m)  $\text{K}_2\text{S}_2\text{O}_3$ (-HCl-NaOH) solution. A weighed piece of gold foil (Heraeus, 99.99% purity)  
226 was attached to a titanium container suspended in the upper part of the reactor. A weighed amount  
227 of iron(II) disulfide ( $\text{FeS}_2$ , pyrite, powder, 325 mesh, 99.8% trace metals basis, <10 ppm As,  
228 Sigma-Aldrich) or iron(II) monosulfide ( $\text{FeS}$ , powder, 100 mesh, 99.9% trace metals basis, <0.1  
229 ppm As, Sigma-Aldrich) was placed into the container, avoiding direct contact with the gold foil.  
230 The runs starting with  $\text{FeS}_2$  and  $\text{FeS}$  are named 'Py' (for pyrite) and 'Po' (for pyrrhotite),  
231 respectively (Table 1). Before experiments, the solids were characterized by XRD and SEM and,  
232 after total digestion in aqua-regia, by ICP-AES (see below); they and were found to be free of gold  
233 (<0.01 ppm) and contain only minor (<0.1 wt%) impurities of  $\text{BaSO}_4$  (in  $\text{FeS}_2$ ), and of  $\text{FeS}_2$  and  
234 native Fe and S (in  $\text{FeS}$ ). The reactors were placed in a temperature-controlled oven ( $\pm 1^\circ\text{C}$ ); the  
235 pressure at the run temperature was estimated from the degree of filling of the reactor (which is  
236 equivalent of the fluid density at experimental  $P$ - $T$ ), and using the  $P$ - $T$ - $X$ -density properties of the  
237 well-known NaCl- $\text{H}_2\text{O}$  system (Driesner and Heinrich, 2007) and assuming that the density of a  
238  $\text{K}_2\text{S}_2\text{O}_3$ (-HCl-KOH)-bearing supercritical fluid is similar to that of an NaCl fluid of the same mass  
239 concentration (wt%); this approximation induces an error of less than  $\pm 100$  bar at 450°C and 700  
240 bar (e.g., Pokrovski et al., 2009a).

241 The run duration was 24 days (except 19 days for Py3 run), which was considered long  
242 enough to attain the solid-fluid chemical equilibrium as shown by numerous previous experiments



243 on Au solubility in S-rich systems over a wide  $T$  range (250-450°C, Pokrovski et al., 2009a, 2015)  
244 as well as by studies of Fe/Ni sulfide recrystallization in the presence of aqueous fluids at  
245 temperatures as low as 120°C (e.g., Xia et al., 2009) that demonstrated attainment of equilibrium  
246 within a period of a few hours to a few days. At the end of the experiment, the reactor was placed  
247 in water for cooling, resulting in a fast separation of the container with the pyrite phase from the  
248 fluid that condenses into liquid in the lower part of the reactor thus avoiding fluid-solid back  
249 reactions, then cooled for 20 min, unloaded, and washed 3 times with hot aqua regia to remove all  
250 gold that was dissolved in the fluid at the run temperature but (partly) precipitated on reactor walls  
251 upon cooling. The resulting washout solution was treated and analyzed for Au (section 2.4). The  
252 solid phase was rinsed with water to remove traces of salt, dried at 50°C, and analyzed as  
253 described below (sections 2.5 and 2.6). The gold foil was also rinsed with water, dried, and  
254 weighed to determine weight loss due to gold dissolution in the fluid and incorporation in pyrite,  
255 thus providing an additional control on Au total dissolved concentrations.

256

### 257 2.3. Gold-pyrite co-precipitation experiments

258

259 Gold-pyrite co-precipitation was simulated at 350°C/400 bar (runs Copy1 and 2) and at  
260 450°C/700 bar (run Copy3, Table 2) using a hydrothermal reactor (Fig. 1) equipped with an ultra-  
261 fast sampling device and a flexible inner gold cell (150-180 cm<sup>3</sup>), adapted to multicomponent S-  
262 rich systems similar to those of the present study (Pokrovski et al., 2008; Kokh et al., 2017). The  
263 reactor is placed in a rocking furnace and the temperature ( $\pm 1^\circ\text{C}$ ) and pressure ( $\pm 5$  bar) are  
264 controlled independently during the run. The reactor is equipped with a sampling tube, two high-  
265 pressure titanium valves and a titanium vial tightly assembled to the second valve; this design  
266 allows for almost instantaneous (within  $<1$  s) sampling of a portion of the fluid phase pushed by  
267 the internal pressure from the reactor, thus avoiding any solute precipitation or degassing, which  
268 would have inevitably occurred in S-rich systems using traditional sampling procedures. The cell  
269 was loaded with an aqueous thiosulfate solution and the reactor was brought to the desired  $P$ - $T$ ;  
270 the fluid was periodically sampled to monitor attainment of gold solubility equilibrium, and  
271 processed and analyzed as described below (section 2.4). The composition of the fluid may be  
272 modified in situ at high  $P$ - $T$  during the run via injection, using a calibrated capstan pump, of a  
273 given amount of another solution (e.g., HCl in CoPy3 experiment) inducing controlled changes in  
274 pH and Au solubility. Furthermore, in the final step of the three CoPy experiments, an FeCl<sub>2</sub>  
275 solution was injected into the cell, inducing pyrite precipitation followed by re-equilibration  
276 between the precipitated pyrite and the fluid of modified composition. At the run termination, the

277 reactor was cooled down within a few hours and the pyrite precipitate was recovered. Compared  
278 to the batch reactor technique (section 2.2), the flexible-cell reactor enables far more accurate,  
279 time- and composition-resolved measurement of gold solubility and sulfur speciation in the fluid  
280 at high  $P$ - $T$ , but does not allow fast recovery of the solid phase that might potentially be subject to  
281 some modifications (e.g., Au precipitation or re-dissolution) during the inevitably slow cooling of  
282 the 80-kg reactor.

283

#### 284 *2.4. Chemical analyses of experimental solutions*

285

286 Chemical treatment and analyses of gold, other metals and sulfur in the experimental  
287 solutions were performed using protocols developed in Kokh et al. (2016, 2017). Briefly, Au  
288 concentrations in aqua-regia washout solutions from experiments in batch reactors were analyzed  
289 by ICP-MS and/or ICP-AES. The sample preparation recipe consists of gentle evaporation of the  
290 aqueous sample on hot plate (at 60-80°C) in a clean Teflon vial (Savilex©) until a wet salt residue  
291 is obtained; then reacting the residue with 2 g of hot aqua regia in the closed vial (at 120°C for ~2  
292 h), followed by gentle evaporation of the solution to 0.2-0.3 g (at 60-70°C), and finally dilution of  
293 the rest by ultra-pure doubly-distilled 0.5wt% HCl – 1.5wt% HNO<sub>3</sub> in water. Gold, K and Fe in  
294 the fluids from experiments in the flexible-cell reactor were treated and analyzed using the same  
295 protocol for fluids sampled into an NH<sub>3</sub> (28 wt% NH<sub>3</sub> in water) solution (Kokh et al., 2016). The  
296 overall analytical uncertainties (which include sample treatment and ICP-MS or ICP-AES  
297 analysis) on the final concentration do not exceed 10% of the value.

298 Chlorine and sulfur were also analyzed in fluids sampled from the flexible-cell reactor.  
299 Total chloride was quantified (as the Cl<sup>-</sup> ion) by high-performance liquid chromatography  
300 (HPLC). Total dissolved sulfur (S<sub>tot</sub>) was analyzed by ICP-AES and/or HPLC (as the SO<sub>4</sub><sup>2-</sup> ion)  
301 after complete oxidation to sulfate in NH<sub>3</sub>-H<sub>2</sub>O<sub>2</sub> solutions. Both methods for S showed an  
302 agreement within 10% of the total concentration value. Total reduced sulfur (S<sub>red</sub>, dominantly  
303 hydrogen sulfide, and eventually sulfite and polysulfide) was quantified by iodometric titration.  
304 Hydrogen sulfide (H<sub>2</sub>S/HS<sup>-</sup>) was separated from the other sulfur species by precipitation as  
305 cadmium sulfide (CdS) and centrifugation, followed by iodometric titration of CdS in acid  
306 solution (10 wt% HCl). Sulfate was analyzed gravimetrically from the fluid sampled directly in a  
307 1 M Ba acetate solution resulting in precipitation of BaSO<sub>4</sub>, which was subsequently washed,  
308 centrifuged, and dried before weighing. Uncertainties of the Cl and S analyses are less than 10%  
309 of the value. The concentrations of Au and other solutes in batch and flexible-cell reactor  
310 experiments are reported in Table 1 and 2, respectively.

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312 *2.5. Conventional analyses of experimental pyrites*

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The solids recovered from both types of experiment were analyzed for phase composition and major element contents by XRD, SEM and EPMA (Table 3). The analyses showed pyrite as the major phase with variable amounts of visible gold particles (0.1-10  $\mu\text{m}$ ) in almost all experiments, plus minor pyrrhotite (<5%) in the two batch-reactor runs started with FeS (Po1 and Po2, Table 1, Fig. 2). Total gold was quantified by dissolution of a weighed amount of the bulk pyrite solid in hot aqua-regia, followed by the same treatment as that for the aqueous solutions (Section 2.4) and ICP-AES analysis. Visible gold contents were also independently estimated on polished sections by SEM-EDS mapping of a given surface area (typically  $\sim 1 \text{ mm}^2$ ). The surface of gold particles, being much brighter in back scattered electron (BSE) mode than pyrite (e.g., Fig. 2), was measured and compared to the surface of all pyrite grains using the ImageJ software. The visible Au concentration (in ppm) was then estimated as:

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$$C_{Au(0)}(\text{ppm}) = (S_{Au(0)} \times \rho_{Au} \times 10^6) / (S_{Pyrite} \times \rho_{Pyrite}) \quad (1)$$

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where  $S_{Au(0)}$  and  $S_{Pyrite}$  are the surfaces of gold and pyrite grains respectively, and  $\rho_{Au(0)}$  and  $\rho_{Pyrite}$  are the densities of pure metallic gold (19.3  $\text{g/cm}^3$ ) and pyrite (5.0  $\text{g/cm}^3$ ), respectively. Although these analyses are subject to large uncertainties due to the potentially limited representativeness of the examined polished surfaces and SEM resolution, they nevertheless provide a useful estimate of visible gold content for Au-rich samples (>0.1 wt% total Au).

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An attempt was also made to analyze invisible Au in large pyrite crystals (>10  $\mu\text{m}$ ), by carefully avoiding visible Au particles, using EPMA (Cameca SX Five microprobe, electron beam spot  $\sim 1 \mu\text{m}$ , accelerating voltage 25 kV, current 100 nA, counting time 240 s). All analyses for the fluid-pyrite exchange runs (Py and Po series) were, however, only a factor of  $\sim 2$  above the EPMA detection limit of  $\sim 120 \text{ ppm Au}$ , in agreement with low invisible gold concentrations found by X-ray absorption spectroscopy in these samples, whereas the co-precipitation runs (CoPy series) yielded much higher Au contents (1000s ppm), which match the XAS-determined concentrations within a factor of 2 (section 2.6).

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Selected samples were also analyzed for Au by LA-ICPMS using a Finnigan MAT Element equipped with a femtosecond UV laser (laser spot diameter  $\sim 20 \mu\text{m}$  yielding an ablation crater of 25-30  $\mu\text{m}$ , energy 0.003 mJ, fluency 0.94  $\text{J/cm}^2$ , power 12%, frequency 5 Hz, background time 45 s, ablation time 60 s). Transient ablation signals were converted to concentration using the SILLS software (Guillong et al., 2008), and an external Au-bearing FeS

344 standard (LaFlamme Po726 standard reference material,  $46\pm 2$  ppm Au, Sylvester et al., 2005) and  
345 Fe content in pyrite as the internal standard. Although the LA-ICPMS technique offers much  
346 higher sensitivity compared to EPMA, with typical detection limits of 0.1 ppm Au for our  
347 samples, the major challenge of such analyses is the small pyrite grain size being typically <10-40  
348  $\mu\text{m}$  with rare exceptions (Fig. 2), which is smaller than or comparable to the ablation spot size. In  
349 addition, the LA-ICPMS analysis is much more penetrative than EPMA, with ablation depth  
350 superior to the beam spot diameter ( $>20 \mu\text{m}$ ). As a result, these analyses of individual pyrite grains  
351 (by avoiding “visible”  $>1 \mu\text{m}$  gold particles, where possible), allow only a rough estimate of total  
352 invisible gold content, and provide information about Au distribution across the grain depth. This  
353 analysis cannot exclude the presence of gold nuggets inside or underneath a pyrite grain. An  
354 attempt to reduce the laser spot to  $5\mu\text{m}$  allowed elimination of some of the largest spikes, but  
355 yielded too unstable and weak ablation signals to obtain accurate analytical results for our  
356 particular samples. Recent developments of LA-ICPMS time-of-flight set-ups (e.g., Gundlach-  
357 Graham et al., 2018) would enable, in near future, high-resolution quantitative elemental imaging  
358 of gold and other trace metals in analytically challenging pyrite samples like those of the present  
359 study.

360

## 361 *2.6. High-resolution X-ray absorption spectroscopy (HR-XAS) analyses of experimental pyrites*

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363 HR-XAS analyses at the Au  $L_3$  edge ( $\sim 11.920$  keV) of experimental pyrite samples were  
364 performed at FAME (BM30b) beamline of the European Synchrotron Radiation Facility,  
365 Grenoble, France (Proux et al., 2005), using a recently developed 5-crystal spectrometer operating  
366 in high-energy resolution fluorescence detected (HERFD) mode (Llorens et al., 2012). The two  
367 main advantages of using HERFD compared to traditional XAS spectroscopy are *i*) a significant  
368 gain in spectral resolution compared to nominal resolution defined by the core hole width of the  
369 absorption edge, allowing the detection of different features in X-ray absorption near-edge  
370 structure (XANES) spectral region, which are indicative of Au redox state and coordination  
371 geometry, and *ii*) the ability to efficiently filter out all unwanted contributions from elastic  
372 scattering and fluorescence from other elements (e.g., As whose absorption edge is within 50 eV  
373 of that of Au) in the sample and thus significantly improve both the limit of detection for Au and  
374 signal-to-noise ratio of XAS spectra. In addition,  $L_3$ -edge XANES spectra at nominal resolution of  
375 most Au(I) compounds and complexes with common ligands (S, Cl, P) are quite similar and rather  
376 featureless (Fig. 3; Pokrovski et al., 2009a,b, 2015), making it difficult to unambiguously

377 distinguish between the different Au coordination environments and the identity of neighboring  
378 atoms.

379         The beamline X-ray optics incorporated a Si(220) double crystal monochromator with  
380 sagittal focusing providing a beam spot (full width at half maximum) of  $\sim 200 \times 300 \mu\text{m}^2$  at the  
381 sample, Rh-coated mirrors for harmonic rejection, and a crystal analyzer spectrometer (Llorens et  
382 al., 2012) with five Si(660) crystals placed in a Roland circle geometry (Bragg angle =  $85.71^\circ$ ).  
383 This device allows a resolution of  $1.3 \pm 0.1$  eV at the  $L\alpha$  Au fluorescence line measured as full  
384 width middle height (FWMH) of the elastic peak, providing a gain in resolution by a factor of  $\sim 5$   
385 compared to nominal spectral resolution (natural core-hole width is 5.4 eV at the Au  $L_3$ -edge).  
386 High-resolution fluorescence XANES (hereafter HR-XANES) spectra were recorded using a  
387 Vortex EX-90 mono-element detector. In addition, nominal-resolution transmission spectra were  
388 acquired at the same time as HERFD spectra for all samples and reference compounds. The pyrite  
389 samples from hydrothermal experiments were homogenized by mild grinding, which was  
390 necessary for obtaining XAS signals devoid of harmonics, diffraction peaks, spectral distortions,  
391 and self-absorption artifacts which are all common for inhomogeneous samples (e.g., Curis et al.,  
392 2005; references therein). The produced powders were pressed in 5-mm diameter pellets, affixed  
393 with silicon grease to a copper holder and placed in a liquid He cryostat ( $\sim 10$  K) to reduce X-ray  
394 beam-induced damage common for redox-sensitive elements like Au (Pokrovski et al., 2009b). A  
395 helium-filled bag was placed to fill the space between the crystal analyzers, cryostat window  
396 where the pyrite sample is placed, and detector, to reduce the absorption of the Au fluorescence  
397 signal by air. Different reference Au-bearing compounds (Au,  $\text{Au}_2\text{S}$ ,  $\text{Na}_3\text{Au}(\text{S}_2\text{O}_3) \times 1-2\text{H}_2\text{O}$ ,  
398  $\text{AuAgS}$ , Au-thiomalate,  $\text{K}_2\text{AuCl}_4 \times 1-2\text{H}_2\text{O}$ ,  $\text{AuTe}_2$ , Au-doped CuS,  $\text{AuSb}_2$ ,  $\text{AuClP}(\text{Phen})_3$  and  
399  $\text{AuCl}(\text{P}(\text{Phen})_3)_2$  where Phen is the phenol cycle  $\text{C}_6\text{H}_5\text{OH}$ ) were diluted by mixing with boron  
400 nitride powder to obtain an Au concentration of 1 to 2 wt%, comfortable for fluorescence  
401 measurements but low enough to avoid self-absorption effects that lead to damping of XANES  
402 amplitudes. The prepared powders were pressed in pellets and recorded similarly to the pyrite  
403 samples, both in transmission (nominal resolution) and fluorescence (high-resolution) modes.

404         Since our pyrites contain both visible and invisible gold (Fig. 2), care was taken to avoid  
405 gold particles larger than  $\sim 1 \mu\text{m}$  and to assure relative homogeneity of Au concentrations over the  
406 whole beam size during the acquisition, which is essential for obtaining accurate and reproducible  
407 XAS spectra. The acquisition spots were thus carefully chosen by scanning the pellets in two  
408 directions (e.g., Fig. A1). The X-ray beam was positioned sufficiently far from large fluorescence  
409 peaks arising from gold particles to allow us to probe essentially invisible and relatively  
410 homogeneously distributed gold in the sample. For most pyrite samples of low Au concentration

411 (<100 ppm) multiple scans (5 to 20) were recorded and, if found similar, merged together to  
 412 increase the signal-to-noise ratio. The concentration of invisible gold in pyrite in the acquisition  
 413 spots may directly be determined from the edge height of the fluorescence spectrum using a  
 414 standard of known Au concentration with similar matrix major elements (Fe, S), measured at  
 415 identical acquisition conditions. The standard used in this study is a mixture of natural Au-bearing  
 416 pyrite and arsenopyrite from the Villeranges deposit, France (sample MCB-7358) in which gold is  
 417 present exclusively in a chemically bound state as demonstrated by Mössbauer spectroscopy  
 418 (Boiron et al., 1989). Total concentrations of Fe, S, As and Au in the standard were determined by  
 419 ICP-AES after complete aqua-regia digestion of the solid. The obtained solid composition  
 420 corresponds to the formula  $\text{FeAs}_{0.28}\text{S}_{1.72}$  with  $450\pm 50$  ppm Au. The solid was ground, pressed in a  
 421 pellet, and measured by XAS similarly to the experimental pyrite samples (detailed account of Au  
 422 XAS spectra of this and other natural and synthetic arsenian pyrites and arsenopyrites will be  
 423 given elsewhere). Gold concentrations in our pyrite samples were thus determined using  
 424 proportionality between absorption edge height in fluorescence mode ( $\Delta\mu_{\text{fl}}$ ) and Au concentration  
 425 for a pyrite sample (*Py*) and the standard (*St*):

$$426 \quad C_{\text{Au}}(\text{Py}) = \Delta\mu_{\text{fl}}(\text{Py}) \times C_{\text{Au}}(\text{St}) / \Delta\mu_{\text{fl}}(\text{St}) \quad (2)$$

427 HR-XANES spectra were normalized to the absorption edge height with the Athena  
 428 software (Ravel and Newville, 2005) to compare with spectra of reference compounds and  
 429 perform linear combination fits (LCF) to quantify the contributions from different Au structural  
 430 environments and to directly distinguish between native and chemically bound gold. Furthermore,  
 431 the experimental XANES spectra were compared with quantum-chemically simulated ones for  
 432 different Au cluster geometries and substitution models in pyrite. Due to relatively low Au  
 433 concentrations, the insufficient signal-to-noise ratio of most pyrite samples did not allow  
 434 acquisition of extended x-ray absorption fine structure (EXAFS) spectra good enough to extract  
 435 accurate structural information. Most EXAFS spectra had low signal-to-noise ratios and short  
 436 exploitable energy ranges ( $3 < k < 8 \text{ \AA}^{-1}$ , where *k* is photoelectron wavenumber), which resulted in  
 437 large errors on Au nearest coordination numbers ( $2 < N < 6$  atoms) and interatomic distances  
 438 ( $2.2 < R < 2.4 \text{ \AA}$ ). As a result, in this study we rely on the analyses of the HR-XANES experimental  
 439 and quantum-chemistry simulated spectra.

440

## 441 2.7. Quantum-chemistry ab-initio modeling of XANES spectra using FDMNES

442

443 Direct and unambiguous interpretation of XANES spectra, particularly in high-resolution  
 444 mode, requires reference compounds with Au redox and structural environments closely matching

445 those of the samples. Since no such stable compounds with known structures exist in the Au-Fe-  
446 S(-As) system, a theoretical simulation of XANES spectra for expected Au geometries and  
447 structures in pyrite is the only alternative. Such simulations become now possible due to  
448 significant progress in quantum-chemical modelling of full electron potential in the near-edge  
449 absorption region, coupled with increasing computer power enabling the use of Fine Difference  
450 Methods (FDM) for solving the Schrödinger equation on the node points of a three-dimensional  
451 grid (Amestoy et al., 2006; Guda et al., 2015). We used the FDMNES code (Joly, 2001; Bunau and  
452 Joly, 2009) to calculate theoretical Au L<sub>3</sub>-edge XANES spectra of different possible Au local  
453 structures and geometries in pyrite (e.g., Au substitution in the Fe or S structural sites, different  
454 Au-S<sub>n</sub> clusters similar to those in aqueous solution adsorbed on pyrite surface) and to compare  
455 them with the experimental spectra. All calculations were performed in FDM mode (note that in  
456 case of low symmetry, the muffin-tin approximation is not sufficient in the near-edge energy  
457 range; Joly, 2001; Testemale et al., 2004), and accounting for relativistic effects and spin-orbit  
458 interactions intrinsic to heavy atoms such as gold (program's keywords 'Relativism' and  
459 'Spinorbit'; Joly, 2018). The obtained raw calculations represent the evolution of the photo-  
460 absorption cross-section of Au as a function of X-ray photon energy and correspond to the  
461 transition amplitude between the initial and final states with an energy resolution of the FD method  
462 (<0.1 eV). These spectra, which display almost all possible electronic transitions, are further  
463 convoluted with a Lorentzian function (keyword 'Arc') with a width of 1.3 eV (keyword  
464 'Gamma\_hole'), which corresponds to the energy resolution of our HERFD setting (section 2.6).

465

## 466 *2.8. Thermodynamic calculations in experimental fluid-mineral systems*

467

468 Gold and sulfur speciation and solubility in the fluid phase of the experimental systems  
469 were modelled using available robust thermodynamic data, and the results were compared with the  
470 measured Au solubility. Calculations were performed using the HCh software package and  
471 associated Unitherm database, allowing chemical equilibrium simulations in multicomponent  
472 fluid-mineral systems based on the minimization of the Gibbs energy of the system (Shvarov,  
473 2008; 2015), and accounting for non-ideality of the fluid using the extended Debye-Hückel  
474 equation (Helgeson et al., 1981). The selection of thermodynamic data sources was discussed in  
475 detail elsewhere (Pokrovski et al., 2015; Kokh et al., 2017), and will only briefly be summarized  
476 here.

477 The thermodynamic properties of the minerals, major fluid components, and most sulfur  
478 aqueous species were taken from the updated SUPCRT (Johnson et al., 1992), JANAF (Chase,

479 1998), and Robie and Hemingway (1995) databases, complemented by recent data for ionic sulfur  
480 forms including  $S_3^-$  (Pokrovski and Dubessy, 2015; references therein) using the revised and  
481 extended HKF (Helgeson-Kirkham-Flowers) equation of state (Oelkers et al., 2009; Sverjensky et  
482 al., 2014; references therein). The thermodynamic properties of the molecular sulfur aqueous  
483 forms,  $H_2S$ ,  $SO_2$  and  $H_2$  and  $O_2$ , were adopted according to a recent model for aqueous non-  
484 electrolytes (Akinfiyev and Diamond, 2003), which allows a more accurate description over the  $P$ - $T$   
485 range relevant to our study (350-450°C; <1 kbar) than the SUPCRT database, which used a more  
486 limited experimental dataset in derivation of the HKF model coefficients for these species.  
487 Noteworthy, the thermodynamic data from Akinfiyev and Diamond (2003) were also used in  
488 derivation of the thermodynamic properties of Au species (see below), and thus were chosen here  
489 to maintain thermodynamic consistency.

490 The thermodynamic properties of the  $Au^+$  cation, and its traditional hydrogen sulfide  
491 ( $AuHS$ ,  $Au(HS)_2^-$ ) and hydroxide ( $AuOH$ ) species were taken from a recent compilation of  
492 Pokrovski et al. (2014), consistent with most recent experimental data. The Gibbs energy value of  
493 the di-chloride species ( $AuCl_2^-$ ) for which there is a more substantial disagreement among the  
494 major data sources, was taken as an average between those reported in Pokrovski et al. (2014) and  
495 Stefansson and Seward (2003). Species such as  $AuCl$  and  $Au(OH)_2^-$ , tentatively suggested in some  
496 previous compilations, were ignored in the present modeling both because their existence in  
497 hydrothermal solutions is not explicitly proven by direct experimental data and their parameters  
498 are based on estimations at ambient conditions and thus are subject to large errors at elevated  
499 temperatures. Among the traditional species, only  $Au(HS)_2^-$  was found to significantly contribute  
500 to the measured Au solubility in our S-rich systems at near neutral pH. The recently discovered  
501 gold complex with the trisulfur ion,  $Au(HS)S_3^-$  (Pokrovski et al., 2015), accounting for Au  
502 solubility in acidic-to-neutral solutions, was also included in the model. In addition, a tentative Au  
503 complex with the disulfur radical ion,  $S_2^-$ , which has been detected by in-situ Raman spectroscopy  
504 in similar thiosulfate aqueous solutions (Pokrovski and Dubessy, 2015), was also tested to account  
505 for the measured solubility (section 4.1). It should be emphasized that our choice of  
506 thermodynamic data sources for major sulfur forms and its complexes with gold is self-consistent,  
507 which is a key condition for accurate interpretation of Au speciation and solubility in S-rich  
508 systems at elevated temperatures, similar to those of the present study (e.g., Pokrovski et al.,  
509 2015).

510

511

### 3. Results



### 512 *3.1. Gold concentration in the fluid phase*

513

#### 514 *3.1.1. Batch-reactor experiments*

515 Gold concentrations measured in quenched solutions from batch-reactor experiments at  
516 450°C and 700 bar are reported in Table 1. The dissolved Au concentrations vary from ~100 ppm  
517 (in acidic solutions) to almost 3000 ppm (in alkaline solutions). A reasonable quantification limit  
518 for aqueous Au in our experimental systems is about 100-150 ppm, which is due to intrinsic  
519 uncertainties of the batch-reactor technique such as incomplete recovery of Au precipitated on the  
520 reactor walls after the quench, possible recrystallization of Au in the reactor during the run due to  
521 minor temperature gradients, mechanical detachment of gold metal particles from the foil  
522 suspended in the upper part of the reactor, and uncontrolled gold dissolution during the reactor  
523 heating or cooling. These artifacts may result either in overestimation or underestimation of the  
524 true gold solubility (see section 4.1). Total gold dissolved concentrations can independently be  
525 estimated from weight loss of the gold foil after the experiment (with a reasonable detection limit  
526 of ~50 ppm Au and an uncertainty of  $\pm 50$  ppm Au corresponding to weight loss of <0.5 mg). This  
527 estimation takes into account Au incorporation in pyrite, but is obviously subject to the same  
528 artifacts. It can be seen in Table 1 that for the runs with low Au solubility in the fluid producing  
529 Au-poor pyrites (Py1, 3 and 6) the gold weight-loss derived concentrations are similar within  
530 errors to those derived from quench-solution analysis. For runs in the presence of initial pyrite and  
531 with higher Au aqueous concentrations and less acidic pH (>400 ppm Au, Py2, 4 and 5), the  
532 weight-loss derived concentrations are also in agreement with those from the quenched solution,  
533 demonstrating almost complete recovery of gold and a negligibly small (within the associated  
534 errors) fraction of the Au incorporated in pyrite. By contrast, runs starting with FeS (Po1, 2)  
535 yielded significantly higher loss-weight Au concentrations compared to those derived from  
536 quenched solution, demonstrating that a more important fraction of Au was precipitated on and/or  
537 incorporated into the solid phase during FeS transformation into pyrite, as also confirmed by direct  
538 analyses of the solid phase (section 3.2).

539

#### 540 *3.1.2. Flexible-cell reactor experiments*

541 Gold and other solute (S, Cl, K, Fe) concentrations, measured in sampled portions of fluid  
542 as a function of time in co-precipitation experiments conducted at 350 and 450°C using the  
543 flexible-cell reactor (Copy1, 2 and 3), are reported in Table 2. It should be emphasized that these  
544 direct measurements are not subjected to the artifacts common for the batch-reactor method  
545 enumerated above and thus may be considered to represent the true gold solubility values. The

546 uncertainties of Au solubility are less than 10% of the total value; they stem mostly from the S-  
547 rich solutions treatment and analysis by independent methods (ICP-MS, ICP-AES). It can be seen  
548 that Au dissolved concentrations before pyrite precipitation were constant within errors for run  
549 durations from 1 to 10 days, demonstrating attainment of saturation with metallic gold. After  
550 pyrite precipitation was induced by injection of FeCl<sub>2</sub> solutions, gold dissolved concentrations  
551 dropped by a factor of 5 to 10 (depending on the experiment), in response to the large decrease of  
552 aqueous S concentration upon scavenging by pyrite and the resulting breakdown of gold-sulfur  
553 complexes. The Au aqueous concentrations after pyrite precipitation either increase by a small  
554 factor (2 to 3) with time (e.g., CoPy1), remain stable (CoPy2), or slightly decrease (CoPy3), likely  
555 reflecting the longer times for the system to recover after a large and rapid change in S  
556 concentration, and the different reaction kinetics of Au-pyrite-fluid exchange and pyrite  
557 recrystallization, which may depend on fluid composition, acidity, and temperature.

558 The concentrations of other fluid constituents (total S, K, Cl) are in excellent agreement  
559 with the initial and injected quantities. Measured total dissolved sulfur contents ( $S_{\text{tot}}$ ) are identical  
560 within errors to the sum of aqueous sulfide and sulfate concentrations in CoPy2 and CoPy3,  
561 suggesting that other S species concentrations (e.g., radical ions, polysulfides) are small. By  
562 contrast, in more acidic CoPy1 experiment,  $S_{\text{tot}}$  concentrations are by 20 to 30% higher than the  
563 sum of sulfate and sulfide, likely due to the presence of molecular sulfur in the experiment, as  
564 demonstrated by in situ Raman spectroscopy on similar solutions (Pokrovski and Dubessy, 2015);  
565 this form of sulfur is not analyzable in sampled fluids by the I<sub>2</sub>, CdS and BaSO<sub>4</sub> protocols  
566 employed in this study (section 2.4). A more detailed account of sulfur speciation and isotope  
567 fractionation in these experiments will be given elsewhere.

568

### 569 *3.1.3. Comparison of Au solubility between the two methods*

570 The dissolved Au concentrations in the flexible-cell reactor experiment at 450°C and 700  
571 bar (CoPy3, Table 2) before and after HCl injection can directly be compared with those from  
572 batch-reactor runs of similar *P-T* and fluid composition (Py4 and 5; Table 1), in particular to  
573 evaluate the accuracy of the batch-reactor solubility measurements. The Au concentrations  
574 measured in batch reactors both in alkaline and neutral solutions (~2800 ppm Au in Py5, 0.5m  
575 K<sub>2</sub>S<sub>2</sub>O<sub>3</sub>-0.3m KOH; ~1600 ppm Au in Py4, 0.5m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub>; Table 1) are 2 to 3 times higher than the  
576 flexible-cell reactor Au concentrations at similar solution compositions (~1300 ppm Au in CoPy3,  
577 initial fluid with 0.5m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub>-0.3m KOH; ~600 ppm Au in fluid after HCl injection resulting in  
578 almost complete neutralization of KOH, ~0.41m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub>-0.2m KCl). Such discrepancies cannot be  
579 caused by minor differences in solution composition between the two types of dissolution

580 experiments or by the presence of pyrite in the batch-reactor experiments since pyrite solubility is  
581 negligibly small at these conditions as shown by thermodynamic modeling ( $<10^{-4}$  m Fe).  
582 Therefore, the overestimated Au concentrations in the S-rich solutions in the batch-reactor  
583 compared to the flexible-cell reactor experiments at 450°C are likely due to enhanced dissolution  
584 of Au during the temperature rise in batch-reactor runs that largely exceeds the Au saturation  
585 concentrations at 450°C. Indeed, thermodynamic predictions of gold solubility, using the stability  
586 constant for the dominant  $\text{Au}(\text{HS})_2^-$  complex at this pH range (see Section 2.8), show a solubility  
587 maximum at 200-250°C in neutral-to-alkaline S-rich solutions, followed by a pronounced  
588 solubility decrease at higher  $T$ , with Au saturation concentrations  $\sim 10$  times lower at 450°C (Fig.  
589 A2a,b). Such Au solubility pattern versus temperature in sulfide solutions is known for a long time  
590 (Pokrovski et al., 2014; references therein). Considering the very fast dissolution kinetics of  
591 metallic Au in such solutions over the  $T$  range 150-450°C, as was shown by time-dependent in-  
592 situ XAS measurements (Pokrovski et al., 2009a), it is thus very likely that some additional Au  
593 was added to the solution by dissolution between  $\sim 200$  and  $\sim 300^\circ\text{C}$  during heating of the batch  
594 reactors. When reaching 450°C, such Au-rich solution became oversaturated and gold re-  
595 precipitated on the reactor walls, contributing both to the washout solution and weight loss  
596 analyses reported in Table 1. This artifact of the batch-reactor method, which has not been  
597 recognized in previous studies that employed the same technique in S-rich ( $>0.5$  m) neutral to  
598 alkaline solutions (e.g., Trigub et al., 2017), requires corrections when interpreting Au speciation  
599 and solubility.

600 Consequently, in the following discussion (section 4.1), we tentatively corrected by a  
601 factor of 0.5 the measured Au concentrations in batch-reactor runs Py4 and Py5. By contrast, in  
602 more acidic S-rich solutions, in which Au speciation is dominated by  $\text{Au}(\text{HS})\text{S}_3^-$ , gold solubility  
603 only weakly depends on temperature, and thus is not expected to be overestimated in batch-reactor  
604 experiments (Fig. A2c). Similarly, at lower sulfide concentrations ( $<0.2$  m, such as in Po1 and Po2  
605 runs, in which  $>$ half of initial S was consumed by pyrite formation), the amplitude of Au solubility  
606 decrease from 250 to 450°C is much smaller and thus may be neglected. The Au aqueous  
607 concentrations adopted in this study are reported in Table 4 and used for interpreting Au aqueous  
608 speciation and fluid-pyrite partitioning (sections 4.1 and 4.3).

609

### 610 *3.2. Gold state and concentration in pyrite from conventional analyses*

611

612 The complementary analytical methods employed in this study help distinguish the  
613 different states of gold associated with pyrite and estimate their approximate concentrations in the

614 solid phase (Table 3). Visible gold contents in pyrites from exchange experiments determined by  
615 SEM mapping are highly variable, ranging from ~100 ppm for runs starting with pyrite (Py1 to 5)  
616 to ~20,000 ppm in runs starting with FeS (Po1 and 2). The Au enrichment in the Po runs is due to  
617 precipitation of native gold from the solution upon sulfidation of FeS to FeS<sub>2</sub> and scavenging of  
618 part of sulfur from the fluid and the resulting decrease in gold solubility, as a result of fast gold-  
619 fluid re-equilibration (section 3.1). Native gold contents estimated by SEM, despite their large  
620 uncertainties (>50% of the value, section 2.2), compare reasonably well with total Au contents in  
621 the solid phase determined by aqua-regia digestion followed by ICP-AES. These contents are 100s  
622 ppm for Py runs and 1,000s ppm for Po runs. Furthermore, an independent estimate of the total  
623 gold amount associated with pyrite in batch-reactor experiments may be made from the difference  
624 between Au concentrations measured in washout solutions and those calculated from weight loss  
625 of the Au foil (Table 1). For the Py runs this difference is zero within errors. In contrast, in the two  
626 Po runs Au weight-loss significantly exceeds Au washout content; this difference corresponds to  
627 pyrite-scavenged Au concentrations of ~10,000 ppm, fairly comparable (within 50%) with those  
628 derived from SEM and total digestion analyses (Table 3).

629 An attempt to determine invisible gold contents in pyrites was made using EPMA by  
630 carefully avoiding microscopic grains of native gold. These contents were found to be only  
631 slightly higher (factor of <2) than the detection limit of the technique (<120 ppm Au) for all  
632 pyrites from exchange experiments, and thus EPMA analyses may provide only qualitative  
633 information. Nevertheless, the variation of Au concentrations among different data points acquired  
634 at different positions at the grain and for different grain sizes (from 5 to ~50  $\mu\text{m}$ ) is rather small  
635 (relative standard deviation, RSD <20%, Table 3), suggesting no significant invisible gold zoning  
636 or noticeable enrichment of small grains. In contrast, pyrites from co-precipitation runs (Copy1  
637 and 2) exhibit quantifiable invisible Au concentrations of 1,000s ppm on average but with  
638 significant data scatter (RSD > 50%, Table 3) likely due to presence of heterogeneously  
639 distributed submicron to nano particulate gold, which could be observed at highest magnification  
640 in EDS images of Au-bearing precipitated pyrite (Fig. 2e, f).

641 LA-ICPMS analyses of selected samples also show a rather heterogeneous gold  
642 distribution across an individual pyrite grain and between different grains in the same sample  
643 (Table 3, Fig. A3). Coexistence of spiky signals and regions of relatively constant signal intensity  
644 as a function of time in the same ablation measurement confirms, in agreement with the other  
645 conventional methods above, the presence of different types of gold in pyrite at micron scale.  
646 These include relatively large ( $\mu\text{m}$  size) particles of native gold, together with clusters of smaller  
647 sub-micron particles, and crystal domains with more homogeneously distributed invisible gold.

648 The few spatially resolved data points that could be acquired across rare large grains (40-50  $\mu\text{m}$ )  
649 or on individual grains of different size (from 10 to 50  $\mu\text{m}$ ) do not show systematic correlations  
650 with Au concentration (Fig. A3). Because of the relatively low spatial resolution of LA-ICMPS,  
651 compared to SEM and EPMA, and small data statistics, no further quantification was possible for  
652 our fine grain samples. Nevertheless, despite these limitations of LA-ICPMS, the mean Au  
653 concentrations for each analyzed sample (Py2, Py5, Po1, and CoPy3; Table 3) match within errors  
654 those of invisible gold from EPMA and XAS analyses. A more accurate assessment of the  
655 different Au chemical forms and their fractions is provided by HR-XAS spectroscopy in the  
656 following section.

657

### 658 *3.3. Gold speciation and concentration in pyrite from HR-XANES spectroscopic analyses*

659

660 HR-XANES scans were recorded on relatively “flat” regions of the pyrite pellets (i.e.,  
661 those showing a constant fluorescence signal versus beam spot position, e.g., Fig. A1) by avoiding  
662 large Au particles ( $>1 \mu\text{m}$ ) clearly identifiable by their intense fluorescence signals and XANES  
663 spectra similar to those of native gold references. The total invisible Au concentration, which is  
664 determined using the edge step height of the fluorescence spectrum at each acquisition point  
665 (equation 2), varied less than 20% of the value; its average for each sample is reported in Table 3.  
666 These fluorescence-derived invisible gold concentration values are far more accurate than those  
667 from the EPMA and LA-ICPMS analyses above, since the technique is sensitive enough  
668 (detection limit is  $<1 \text{ ppm Au}$ ) and more representative of the whole sample because of the much  
669 larger X-ray beam size ( $\sim 300 \mu\text{m}$ ) than the EPMA (1-3  $\mu\text{m}$ ) or LA-ICPMS (20-30  $\mu\text{m}$ ). Thus, the  
670 fluorescence-derived values of invisible gold concentration were retained in the subsequent  
671 interpretation (section 4).

672 Absorption edge-height normalized XANES spectra of selected pyrite samples are shown  
673 in Fig. 4. There are three groups of samples differing by the dominant state of invisible Au, as  
674 inferred by comparisons with the spectra of metallic gold and Au(I) and Au(III) reference  
675 compounds. Some pyrites (Po1, 2; CoPy2, 3) have spectra very similar to those of elemental gold  
676 demonstrating that Au(0) largely dominates in those samples. The two major distinguishing  
677 features of the HR-XANES Au(0) spectrum are *i*) a relatively low-amplitude white-line with a  
678 wide uneven top over 11,920-11,923 eV, and *ii*) an intense resonance centered around 11,945 eV.  
679 These main features are little sensitive to gold particle size, ranging from  $>1 \mu\text{m}$  down to 5 nm, as  
680 shown by XANES spectra of micron-particle Au metal and a nano-particle Au suspension (Fig. 4).

681 The spectra directly confirm the large abundance of metallic gold in the first group of samples as  
682 also qualitatively indicated by SEM, EMPA and LA-ICPMS (Table 3). Another group of samples  
683 (Py1, 2, 3, 5) exhibits XANES spectra very different from those of Au(0) and Au(III) compounds  
684 but strikingly similar to those of Au(I) S-bearing compounds such as Au<sub>2</sub>S, Au-thiomalate or Au-  
685 thiosulfate (not shown), as to the energy position, shape and amplitude of the white line at ~11,921  
686 eV and the post-edge resonance at ~11,930 eV. These spectral features are characteristic of Au(I)  
687 coordinated by two sulfur atoms in a linear geometry, typical for the majority of Au(I)-S  
688 compounds (e.g., Pokrovski et al., 2009a). In addition, Py1 sample exhibits a very minor, yet  
689 detectable, Au(0) feature at 11,945 eV (which can be quantified using LCF analyses below). The  
690 third group of samples (Py3, 4; CoPy1) displays spectra with both Au(I) and Au(0) features. Note  
691 also that the spectra of all pyrite samples are clearly distinct from those of Au(III)-bearing  
692 compounds (e.g., KAuCl<sub>4</sub>×2H<sub>2</sub>O), which show a very intense feature at 11,920 eV due to the 2p-  
693 5d transition, allowed in Au<sup>3+</sup> but forbidden in Au<sup>+</sup> (Berrodier et al., 2004; Pokrovski et al.,  
694 2009b), together with other smaller post-edge resonances, which are all absent in the pyrite  
695 spectra, demonstrating that no Au(III) occurs in our samples.

696 To further quantify the fractions of Au(0) and Au(I), linear combination fit (LCF)  
697 analyses were performed on the pyrite spectra using as reference compounds metallic Au for  
698 Au(0) and Au(I) sulfide (Au<sub>2</sub>S), thiosulfate or thiomalate for Au(I) (Au<sub>2</sub>S giving the best fit  
699 quality). An example of LCF is given in Fig. A4. Owing to contrasting spectral features of Au(I)  
700 and Au(0) in high-resolution mode, our method has a detection limit for Au(0) of ~3 mol% and an  
701 uncertainty of ±3% on the value of Au(0) mole fraction, as evaluated using variable fit energy  
702 ranges and different Au(I) compounds as standards. The fraction of chemically bound Au in pyrite  
703 samples, Au(I), relative to total Au, Au(I)+Au(0), derived from LCF is reported in Table 3. These  
704 data allow direct quantification of the absolute Au(I) concentrations, which are the product of  
705 Au(I) mole fraction and the total invisible Au concentration from fluorescence spectra (as  
706 determined using equation 2, section 2.6).

707

#### 708 *3.4. Gold structural position in pyrite from FDMNES-simulated XANES spectra*

709

710 Comparisons of the experimental XANES spectra of Au-bearing pyrites with those of  
711 reference compounds indicate that invisible chemically bound Au associated with pyrite is likely  
712 present as S-Au(I)-S moieties (i.e., part of a molecule), probably associated with more distant  
713 atoms (Fig. 4). However, the absence of stable Au-Fe-S-bearing solids with well-known  
714 stoichiometry and structure whose spectra could be included in the comparison does not allow

715 unambiguous assessment of other possible Au coordination environments and their contribution to  
716 the spectral signal. For example, some previous studies (e.g., Trigub et al., 2017) suggested that  
717 chemically bound Au might also substitute for Fe or S in the pyrite structure; however, there is no  
718 unambiguous proof for such models in the absence of adequate reference compounds.

719 To verify the impact of such potential substitutions on the XANES signal, we simulated  
720 XANES spectra of Au in the Fe and S crystallographic sites of pyrite using the FDMNES code.  
721 The Au-Fe/S distances and site geometries ranged from those of a regular pyrite structure, in  
722 which Fe is octahedrally coordinated by 6 S atoms at 2.26 Å, and S is surrounded by 1 S and 3 Fe  
723 atoms in a distorted tetrahedron, to those of relaxed structures, in which the Fe substitution site is  
724 expanded to accommodate an AuS<sub>6</sub> octahedron with Au-S distances between 2.40 and 2.50 Å as  
725 predicted by DFT simulations (Appendix A). The resulting calculated XANES spectra were found  
726 to be only weakly sensitive to variations of Au-S and Au-Fe distances within the same  
727 coordination geometry. The main factor affecting the XANES spectral shape, white line intensity,  
728 and energy position is no doubt the Au nearest coordination sphere itself (i.e., number and cluster  
729 geometry of the first-shell neighbors). It can be seen in Fig. 5a that calculated spectra for Au in Fe  
730 and S sites in the pyrite structure are characterized by a very intense white line and significant  
731 energy shifts as compared with the experimental spectra. The differences are far too large to be  
732 attributed to imperfections of the FDMNES simulations or chosen structural models and  
733 calculation parameters.

734 The FDMNES parameters have been validated by simulations of spectra of reference  
735 compounds with known structures, which were found to adequately match the experimental  
736 spectra (Fig. A5) and, in particular, the white-line intensity and position which is primarily  
737 function of absorber first-shell coordination (Au site geometry and identity of neighboring atoms).  
738 The effect of coordination environment is further demonstrated by simulations of XANES spectra  
739 of isolated AuS<sub>n</sub> clusters (2 < n < 6), with geometries ranging from linear AuS<sub>2</sub> to square-plane or  
740 tetrahedral AuS<sub>4</sub> to octahedral AuS<sub>6</sub> with Au-S distances from 2.3 to 2.5 Å. It can be seen in Fig.  
741 5b that there is a net increase in the white-line intensity with increasing Au coordination by S from  
742 2 to 6, with the best match of the experimental pyrite spectra by a linear S-Au-S cluster (Au-S  
743 distance 2.3 Å), which is the structural unit for Au(I) in most of its S-bearing solids and aqueous  
744 complexes (Pokrovski et al., 2009a). The simple AuS<sub>2</sub> cluster has a narrower white line and a first  
745 post-edge resonance maximum slightly shifted to lower energies compared to the experimental  
746 pyrite spectrum, likely due to the presence of more distant atoms in the latter. The effect of such  
747 atoms (H, S) on the XANES spectra of Au clusters may further be simulated by using the  
748 optimized geometries of the Au(HS)<sub>2</sub><sup>-</sup> and Au(HS)S<sub>3</sub><sup>-</sup> complexes obtained using first principles

749 molecular dynamics in aqueous solution (Pokrovski et al., 2015). It can be seen in Fig. 5c that the  
750 FDMNES simulated XANES spectra of these species do provide a better match of the  
751 experimental spectrum than a simple AuS<sub>2</sub> cluster. Interestingly, there are small but distinctive  
752 differences in the spectra of the two aqueous Au complexes (white-line energy shift and energy  
753 and amplitude differences in the post-edge feature) suggesting that they might directly be  
754 distinguished in aqueous solution in high-resolution mode. Finally, even a better match with the  
755 experimental spectrum is obtained for an H-free S-Au-S<sub>3</sub> cluster that exhibits a white line  
756 amplitude almost identical to the experimental spectrum and an energy shift of the maximum of  
757 the first post-edge resonance toward that of the experiment (Fig. 5c). These similarities suggest  
758 that Au(I) might be adsorbed on pyrite surface as such or similar complexes, and thus S<sub>3</sub><sup>2-</sup> (and S<sub>2</sub><sup>2-</sup>  
759 ) ions might favor gold scavenging by pyrite. More in-depth insight into the nature and structural  
760 environment of gold species on the pyrite surface requires systematic molecular dynamics  
761 simulations of the structures of adsorbed complexes, which is beyond the scope of the present  
762 study.

763 Note that the effect of Au coordination on the XANES spectrum is neatly emphasized in  
764 the high-resolution mode, which is necessary for unambiguously distinguishing the different Au  
765 coordination environments in sulfur-bearing systems. In summary, both experimental and  
766 FDMNES-simulated XANES spectra demonstrate that invisible Au in pyrite forms di-sulfide  
767 clusters, likely bound to more distant S atoms, and no detectable incorporation of Au occurs in Fe  
768 or S sites of the pyrite structure. The nature of this gold and its fluid-mineral partitioning are  
769 discussed in the following sections.

770

771

## 4. Discussion

772

### 4.1. Interpretation of gold solubility and speciation in sulfur-bearing fluids

773

774  
775 The interpretation of Au solubility measured in our experimental sulfate-sulfide systems in  
776 terms of Au aqueous complexes is primarily based on knowledge of sulfur speciation itself, which  
777 has been a subject of recent systematic research in similar systems at elevated temperatures and  
778 pressures using in situ spectroscopic methods (Pokrovski and Dubrovinsky, 2011; Jacquemet et  
779 al., 2014; Barré et al., 2017; Pokrovski and Dubessy, 2015; Pokrovski et al., 2015). These direct  
780 spectroscopic data, combined with available thermodynamic properties of the major traditional S  
781 species such as sulfate and sulfide, generally agree that the H<sub>2</sub>S/HS<sup>-</sup>, S<sub>3</sub><sup>2-</sup> and, potentially, S<sub>2</sub><sup>2-</sup> ions



782 are the major ligands for gold. Figure 6 shows the distribution of S species calculated using the  
783 available thermodynamic data, as a function of pH (equivalent to addition of HCl or KOH) at  
784 450°C and 700 bar in a 0.5 m thiosulfate solution used in this study. The  $S_2^{2-}$  ion was also included  
785 in these calculations based on the few available data points from in situ Raman spectroscopy  
786 measurements of Pokrovski and Dubessy (2015) at 450°C in similar solutions (see Table A1 for  
787 derivation of  $S_2^{2-}$  stability constant). It can be seen that the S species concentrations are strongly  
788 pH dependent, resulting in a strong pH control on both Au complexes distribution and overall Au  
789 solubility. In this discussion, we focus mostly on Au solubility data points obtained in this study at  
790 450°C that represent a dataset large enough to allow a systematic examination of the effect of pH  
791 and testing different Au speciation models. Our other experiments conducted at lower temperature  
792 and lower S dissolved contents (Po1 and Po2, and CoPy1 and CoPy2 runs, Table 1 and 2) were  
793 not considered in detail here because these few data points provide insufficient pH and ligand  
794 concentration contrast in such dilute solutions to allow unambiguous discrimination among the  
795 different Au speciation models.

796 Figure 7 compares the measured gold solubility at 450°C from this study with different  
797 thermodynamic models. Model 1 includes both the  $Au(HS)_2^-$  (dominant at neutral to basic pH) and  
798  $Au(HS)S_3^-$  (dominant at acidic to neutral pH) complexes, whose thermodynamic properties have  
799 been reported in recent studies based on Au solubility and in-situ XAS measurements in  
800 hydrothermal  $S_3^{2-}$ -bearing solutions (Pokrovski et al., 2015). This model predicts Au solubility in  
801 acidic conditions (pH<6, the neutral pH at 450°C, 700 bar being 6.05, see Fig. 6) within 50% of  
802 the experimental value, but underestimates measured solubility at neutral-to-alkaline pH by a  
803 factor of 5 to 10. For comparison, a traditional model involving  $Au(HS)_2^-$  alone (model 2),  
804 established on a large literature dataset obtained in  $H_2S/HS^-$ -dominated reducing solutions in  
805 which  $S_3^{2-}$  is negligible (Pokrovski et al., 2014), largely underestimates, by a factor of 10 to 100,  
806 the measured solubility over the whole pH range; this disagreement further confirms the  
807 significance of Au- $S_3^-$  complexes, particularly at pH<6, but points to the existence of some  
808 additional Au species at more alkaline pH.

809 Because the  $S_2^{2-}$  ion attains its maximum concentration at pH 7-8, becoming comparable  
810 with that of  $S_3^{2-}$  (Fig. 6), an  $S_2^{2-}$ -bearing complex,  $Au(HS)S_2^-$  analogous to  $Au(HS)S_3^-$ , was  
811 considered in an attempt to account for the measured elevated solubility in that pH range. The  
812 stability constant of  $Au(HS)S_2^-$  was adjusted to match the measured solubility at pH~8 at which  
813 the disagreement of experiment with both model 1 and 2 is the largest. This new model involving  
814 the three species, shown as model 3 in Fig. 7, largely overestimates the solubility at more acidic  
815 pH, clearly demonstrating that  $Au(HS)S_2^-$  cannot be a dominant Au species at pH >7. For

816 comparison, model 3bis shows the concentration of the same  $\text{Au}(\text{HS})\text{S}_2^-$  complex calculated,  
 817 following the analogy with  $\text{Au}(\text{HS})\text{S}_3^-$  (Pokrovski et al., 2015) assuming an equilibrium constant  
 818 of 1 for the ligand exchange reaction



820 This approximation is expected to be more reasonable than model 3 above since the chemical  
 821 affinities for Au(I) of the  $\text{HS}^-$ ,  $\text{S}_3^{2-}$  and  $\text{S}_2^-$  ions of same electric charge and close formal redox  
 822 state of S (-1, -0.33, and -0.5, respectively) should be similar (Pokrovski et al., 2015). Thus,  
 823  $\text{Au}(\text{HS})\text{S}_2^-$  appears too weak to account for the elevated measured Au solubility at  $\text{pH} > 6$ . It  
 824 cannot be excluded, however, that complexes with  $\text{S}_2^-$  ion of different stoichiometry and charge  
 825 than those considered here may contribute to Au solubility the neutral-to-alkaline region, but our  
 826 dataset is too small to evaluate their impact and would require more systematic work.

827 Finally, model 4 considers  $\text{Au}(\text{HS})_2^-$  alone, according to recent measurements of Trigub et  
 828 al. (2017) in concentrated S-NaOH solutions using a similar batch-reactor quench technique as  
 829 that of our study. These authors, however, used a simplistic S speciation model compared to that  
 830 in our study, by ignoring  $\text{S}_3^{2-}$  and  $\text{S}_2^-$  ions, as well as NaHS and Na(H)SO<sub>4</sub> ion pairs and  
 831 Na<sub>2</sub>SO<sub>4</sub>(s) precipitation in their high-temperature alkaline solutions, which all may strongly affect  
 832 predictions of both pH and redox potential and the resulting Au solubility interpretation.  
 833 Combined with the artifact of the batch-reactor technique leading to overestimation of Au  
 834 solubility as shown in this study (section 3.1; Fig. A1), their Au speciation model should be  
 835 regarded with care. It can be seen in Fig. 7 that Trigub et al.'s model severely underestimates Au  
 836 solubility at acidic pH, by at least a factor of 10, and overestimates Au solubility at neutral-to-  
 837 alkaline pH by a factor of 2.

838 In summary, none of the models considered above can fully account for the measured Au  
 839 solubility over the whole pH range at 450°C, and in particular in the alkaline region ( $\text{pH} > 6$ ),  
 840 which have not been covered before by direct systematic experiments. If the actual models  
 841 qualitatively agree on the existence of Au complexes with  $\text{HS}^-$ ,  $\text{S}_3^{2-}$  and, possibly,  $\text{S}_2^-$  ligands, the  
 842 exact stoichiometry and charge of these complexes, in particular those with  $\text{S}_2^-$ , will require more  
 843 systematic studies by both solubility and spectroscopic approaches.

844

## 845 *4.2. The speciation of invisible gold in pyrite and comparison with literature data*

846

### 847 *4.2.1. Invisible metallic gold*

848 Our findings of two major forms of invisible gold in pyrite, metallic gold submicron- to  
 849 nanoparticles, Au(0), and chemically bound gold, Au(I), are in general agreement with the large

850 amount of literature published over almost 50 years (e.g., Boyle, 1969; Cabri et al., 1989; Reich et  
851 al., 2005; Tauson et al., 2018; references therein). Both visible and invisible metallic gold  
852 systematically found in our experimental pyrites is likely to have precipitated *i)* on cooling and/or  
853 during the run due to minor thermal gradients favoring gold redistribution in the pyrite-fluid  
854 system (Py runs), *ii)* upon FeS<sub>2</sub> formation by scavenging S from solution in experiments started  
855 with FeS (Po runs), and *iii)* upon FeCl<sub>2</sub> injection inducing FeS<sub>2</sub> precipitation (CoPy runs), all these  
856 phenomena leading to gold-supersaturated solutions favoring metal precipitation. The degree of  
857 supersaturation, reaction time, and temperature may be the major factors controlling the size and  
858 distribution of gold particles. As it has been recognized for a long time (Kouzmanov and  
859 Pokrovski, 2012, for an overview and references therein), similar processes of fluid cooling, Fe-  
860 rich rock sulfidation, and pyrite precipitation in nature lead to formation of metal gold. Another  
861 possible process leading to nano-particle gold formation in pyrite and other sulfo-arsenides both in  
862 nature and experiment may be fluid-assisted recrystallization and breakdown of chemically bound  
863 Au that was initially incorporated into the mineral lattice or adsorbed on the surface. For example,  
864 this process likely operated in orogenic gold deposits during pyrite and arsenopyrite metamorphic  
865 recrystallization leading to expulsion of chemically bound metastable gold (e.g., Cook et al., 2009;  
866 Large et al., 2009; Velasquez et al., 2014), and in mantle chromitites hosting platinum group  
867 elements (PGE) that were subjected to fluid-assisted metamorphism leading to formation of  
868 nanoparticles of Ru-Os-Ir alloys through thermal breakdown of PGE-bearing sulfide minerals  
869 (e.g., González-Jiménez et al., 2015). In addition, slow crystal growth favoring migration and  
870 clustering of adsorbed Au complexes on the surface of arsenopyrite might be a controlling factor  
871 of nanoscale Au distribution, as recently evidenced by atom probe microscopy on arsenopyrite  
872 from the Obiasi orogenic gold deposit (Fougerouse et al., 2016). Furthermore, because solid  
873 solution miscibility is generally favored at elevated temperatures (e.g., Barton and Tolmin, 1966;  
874 Barton, 1969), cooling itself may destabilize Au dissolved as a solid solution in the mineral lattice  
875 leading to the formation of gold nanoparticles. Reheating such pyrites leads to dissolution of Au  
876 nanoparticles as was shown by in situ experiments coupled with TEM observations on Au-bearing  
877 pyrites from Carlin deposits (Reich et al., 2006). However, an opposite behavior was recently  
878 reported by Trigub et al. (2017) who found formation of metallic gold while heating, in an  
879 anhydrous FeS<sub>2</sub>-S system, synthetic pyrite originally containing dominantly chemically bound  
880 gold. Such behavior is at odds both with general knowledge of the thermal stability of most solid  
881 solutions and the existing in-situ observations, thus indicating that chemically bound Au in their  
882 pyrites is unlikely to be dissolved in the mineral lattice as an isomorphic substitution.

883

884 *4.2.2. Redox state of chemically bound gold*

885 If the mechanisms leading to occurrence of invisible *metallic gold* in pyrite are widely  
886 recognized and explored in existing work as outlined above, the major debate remains on the exact  
887 redox state, coordination, and structural position of *chemically bound gold*. Our data clearly  
888 demonstrate that this type of invisible Au in pyrite occurs in nominally +1 redox state (Au(I)),  
889 bound to sulfide and/or polysulfide ions, very similar to Au(I) in the fluid (such as Au(HS)S<sub>3</sub><sup>-</sup>) and  
890 Au solids with reduced sulfur (Au<sub>2</sub>S, Au-thiosulfate, Au-thiomalate). This finding is in agreement  
891 with some previous work using Mössbauer spectroscopy (Marion et al., 1986; Friedl et al., 1995;  
892 Palenik et al., 2004) and X-ray absorption spectroscopy on both natural (Simon et al., 1999) and  
893 synthetic (Trigub et al., 2017) pyrites. Our results and previous data convincingly rule out the  
894 significant presence of trivalent gold, Au(III), suggested in some studies on the basis of theoretical  
895 considerations of cationic substitutions in the pyrite structure involving Fe<sup>2+</sup>, Au<sup>3+</sup>, As<sup>3+</sup> and  
896 vacancies (e.g., Arehart et al., 1993; Chouinard et al., 2005; Deditius et al., 2008). Note that the  
897 nominally +3 redox state, Au(III), is known to only be stable in surficial oxygenated environments  
898 (e.g., Berrodier et al., 2004) or chemical compounds with strongly electronegative ligands such as  
899 F (Cotton and Wilkinson, 1988). Furthermore, trivalent gold association with reduced sulfur in  
900 pyrite would be highly unlikely because no stable Au(III) sulfide compounds are known in  
901 chemistry. Note, however, that the notion of redox state of a metal in covalent-bond dominated  
902 compounds such as those in the Au-Fe-S(-As) system should be taken with care (e.g., Cabri et al.,  
903 2000). Furthermore, in the presence of arsenic (arsenopyrite or lollingite), Au chemical state might  
904 be different and a similar debate exists on those minerals (e.g., Cathelineau et al., 1989; Genkin et  
905 al., 1998; Cabri et al., 2000), which is beyond the scope of our study that deals with As-poor  
906 systems.

907

908 *4.2.3. Coordination and structural position of chemically bound gold*

909 Another major topic of debate is Au coordination and the corresponding position in the  
910 pyrite structure. It should be emphasized that the only method enabling unambiguous  
911 determination of Au site geometry and local atomic environment is XAS. Our HR-XANES  
912 experimental spectra and FDMNES simulations collectively demonstrate that Au is coordinated in  
913 its nearest shell by two S atoms, forming [-S-Au(I)-S-] moieties, similar to those known in most  
914 Au sulfide and thiol-bearing compounds and Au-S species in hydrothermal solutions (e.g.,  
915 Pokrovski et al., 2009a). Our findings agree with the pioneering XAS study of Simon et al. (1999)  
916 on pyrites from Carlin-type gold deposits. Comparisons of their XANES data with reference  
917 compounds would suggest a linear or square-planar coordination of Au with surrounding atoms,

918 but the insufficient resolution resulting in rather featureless spectra did not allow distinguishing  
 919 among the possible neighboring atoms (Fe/S/As) in that early work. The Au L<sub>3</sub>-edge XANES  
 920 spectra of experimental pyrites of our study are also similar to those of natural pyrites from  
 921 orogenic gold deposits, recently investigated using high-resolution synchrotron XRF and XANES  
 922 spectroscopy, showing the presence of both Au(I) and Au(0) (Stromberg et al., 2019); however, no  
 923 further development concerning Au(I) structural and coordination state was given in that study.  
 924 Our interpretation disagrees with a recent XAS study of Trigub et al. (2017) who suggested, on the  
 925 basis on XANES modeling and EXAFS fits of hydrothermally synthesized Au-bearing pyrites,  
 926 that Au<sup>+</sup> substitutes for Fe<sup>2+</sup> in the pyrite structure. As in our study, however, their XANES  
 927 spectra astonishingly resemble those of Au<sub>2</sub>S, and the poor signal-to-noise ratio of their pyrite  
 928 EXAFS spectra renders quantification of Au coordination number and Au-ligand interatomic  
 929 distances uncertain (see Appendix A for detailed comparisons). Our data clearly rule out  
 930 significant Au substitution in pyrite (at least As-free), which is also in agreement with studies of  
 931 Tauson's group (Tauson, 2018, for overview) whose early experiments showed that only a few  
 932 ppm of Au could be accommodated as a solid solution in the pyrite structure in Au metal-saturated  
 933 systems. These authors further suggested, based on statistical analyses of bulk Au concentration  
 934 and distribution as a function of grain size of synthetic and natural pyrites, that most Au is  
 935 concentrated in so called 'non-autonomous' phase forming a layer of less than 1 μm thick at the  
 936 pyrite surface. However, the exact composition of such a phase and Au structural and redox state  
 937 remain unknown in the absence of direct spatially-resolved spectroscopic data. Further insight into  
 938 the nature of chemically bound gold in pyrite may be gained from Au pyrite/fluid partition  
 939 coefficients as discussed below.

940

#### 941 *4.3. Gold partitioning between pyrite and hydrothermal fluid*

942

943 The direct data obtained in this study on Au concentrations ( $C_{\text{Au(I)}}$ ) in the fluid phase and in  
 944 coexisting pyrite allow derivation of the partition coefficient of chemically bound gold between  
 945 pyrite and fluid defined as

$$946 D_{\text{py/fl}} = C_{\text{Au(I)}} \text{ (in pyrite)} / C_{\text{Au(I)}} \text{ (in fluid)} \quad (4).$$

947 A good linear correlation, apparent in Fig. 8 between Au(I) concentrations in the fluid phase and  
 948 in pyrite from the exchange experiments at 450°C and 700 bar (Py and Po runs) in a wide range of  
 949 aqueous Au contents (from 10s to 1,000s ppm), suggests an equilibrium exchange reaction. Such  
 950 an exchange is little (if at all) dependent of the nature of Au(I) dominant species in the fluid  
 951 (Au(HS)S<sub>3</sub><sup>-</sup> vs Au(HS)<sub>2</sub><sup>-</sup> vs Au-S<sub>2</sub><sup>-</sup> complexes) and the initial solid phase (pyrite vs pyrrhotite).

952 The slope of the linear regression through zero of these data points corresponds to a partition  
953 coefficient  $D_{\text{py/fl}}$  of  $0.15 \pm 0.07$  (1 SD). The co-precipitation experiment at the same  $P$ - $T$  conditions  
954 (CoPy3) yielded an uncertain  $D_{\text{py/fl}}$  value ranging from  $<0$  to  $\sim 4$ , since Au(I) fraction in this  
955 sample dominated by metallic Au was close to the detection limit ( $\sim 3$  mol% Au(I), Tables 3 and  
956 4). Considering this uncertainty, these co-precipitation data overlap, within error, with the  $D$   
957 values from exchange experiments above. It seems thus likely that similar physical chemical  
958 processes control Au(I) partitioning into pyrite in both types of experiment. The partition  
959 coefficients of chemically bound Au for the lower-temperature co-precipitation experiments  
960 (CoPy 1 and 2 at  $350^\circ\text{C}$ ) are significantly higher, being of the order of 10 to 50 (Fig. 8, Table 4),  
961 indicating that incorporation of Au(I) in pyrite is favored at lower temperatures. This experimental  
962 finding is also in agreement with the higher abundance of chemically bound Au in pyrites from  
963 low-to-moderate temperature hydrothermal deposits (e.g., Carlin-type, epithermal) compared to  
964 pyrites from high-temperature magmatic or metamorphic deposits (porphyry, orogenic) that  
965 contain dominantly metallic gold in ore (e.g., Deditius et al., 2014; references therein).

966 Our Au partitioning values from pyrite-fluid exchange experiments at  $450^\circ\text{C}$  are in  
967 excellent agreement with those reported by Tauson et al. (2018) from experiments at the same  $T$   
968 and similar  $P$  ( $D_{\text{py/fl}} = 0.3 \pm 0.2$ ). Interestingly, no effect of arsenic on Au incorporation in pyrite  
969 was found in their study at similar conditions, in agreement with the paucity of Au-rich arsenian  
970 pyrites in high-temperature hydrothermal settings. These authors interpreted such Au partitioning  
971 as an isomorphic substitution in a non-autonomous submicron-thick solid phase at the pyrite  
972 surface, but the structural and chemical position of Au in such a phase yet remain hypothetical.  
973 The methods employed in our study do not have sufficient spatial resolution to verify the existence  
974 of such a phase, which would require nano-scale in situ techniques, such as nano-beam XRD,  
975 XRF, and XAS that will potentially be offered by the next generation of synchrotrons (e.g.,  
976 Dimper et al., 2014). Furthermore, the robust linear correlation between dissolved and  
977 incorporated Au(I) displayed in Fig. 8 suggests that the Au(I) activity (equivalent to  
978 concentration) in pyrite increases with increasing the activity (or concentration) of aqueous Au(I)  
979 in the fluid. Such a dependence rules out formation of Au proper solid phases (e.g., such as  $\text{Au}_2\text{S}$ )  
980 on the pyrite surface, since Au activity in such a gold-dominated phase is constant, and thus must  
981 be independent of Au activity in the fluid, which is not the case of our data shown in Fig. 8  
982 displaying a clear positive correlation between Au in solution and pyrite. Even though our data do  
983 not rule out Au substitution as an isomorphic impurity in a non-autonomous phase suggested in  
984 Tauson's studies, a more plausible and simple explanation of the Au partitioning trends would be

985 formation of adsorbed S-Au-S complexes on the pyrite surface, similar to those in aqueous  
986 solution as discussed below.

987

#### 988 *4.4. Chemisorption control on Au(I) incorporation in As-poor pyrite*

989

990 Chemisorption has indeed been the most commonly invoked phenomenon to explain Au  
991 scavenging by (arseno)pyrite (e.g., Fleet and Mumin, 1997), but detailed physical-chemical  
992 mechanisms and driving forces of such sorption remain elusive. Only few experimental studies  
993 devoted to Au sorption on sulfide mineral surfaces are available, none of them being conducted  
994 above 90°C, to the best of our knowledge. Some studies (Mironov et al. 1982; Bancroft and  
995 Hyland, 1990; Mycroft et al. 1995; Maddox et al. 1998; Mikhlin et al., 2011) employed acidic  
996 Au<sup>3+</sup> and Au<sup>+</sup> chloride, cyanide, thiosulfate or citrate solutions, of little relevance to hydrothermal  
997 sulfide-bearing fluids. They found that part of Au was rapidly reduced to the metallic state on the  
998 surface of pyrite and arsenopyrite (depending of the relative stability of the Au aqueous  
999 complexes). Some other studies (Scaini et al., 1998; Widler and Seward, 2002) examined sorption  
1000 by iron sulfide minerals from Au(I) sulfide solutions, but reached rather different conclusions. For  
1001 example, Scaini et al. (1998) found that Au sorption on FeS<sub>2</sub> increased with increasing pH  
1002 accompanied by Au(I) reduction to metal, whereas Widler and Seward (2002) observed a  
1003 moderate decrease of Au(HS)<sup>0</sup> and Au(HS)<sub>2</sub><sup>-</sup> adsorption on pyrite with increasing pH above 5,  
1004 with no changes in redox state of adsorbed Au(I). Furthermore, these authors found that Au(I)  
1005 adsorption on marcasite and pyrrhotite was significant only in acidic solutions (pH<4). Even  
1006 though the temperature-composition parameters of our study are quite different from their work,  
1007 there is a surprising similarity in acidity trends in our experiments, with the most acidic run (Py 1,  
1008 pH~4.4) displaying a  $D_{py/fl}$  value (~0.3) twice as high as the less acidic experiments (Table 4, Fig.  
1009 8). Our values at acid pH are also close to those recently reported from fluid-pyrite Au exchange  
1010 experiments in highly acidic S-H<sub>2</sub>O solutions at 450°C ( $D_{py/fl}$  ~1; Trigub et al., 2017) in line with a  
1011 sorption control on Au incorporation in pyrite. Similar tendencies with pH were observed for  
1012 Au(HS)<sup>0</sup> and Au(HS)<sub>2</sub><sup>-</sup> sorption on As<sub>2</sub>S<sub>3</sub> and Sb<sub>2</sub>S<sub>3</sub> at  $T<90^{\circ}\text{C}$  (Renders and Seward, 1989).  
1013 Because the isoelectric point of pyrite surface in H<sub>2</sub>S-bearing aqueous solution is around pH~2 at  
1014 ambient conditions (i.e., pyrite surface is positively charged at pH<2, and negatively charged at  
1015 pH>2; Bebie et al., 1998; Widler and Seward, 2002), the strong adsorption of Au-sulfide  
1016 negatively charged complexes dominant in S-bearing hydrothermal fluids investigated in this  
1017 study (Au(HS)<sub>2</sub><sup>-</sup>, Au(HS)S<sub>3</sub><sup>-</sup>) should be different from simple electrostatically controlled outer-  
1018 sphere sorption of cations or anions on oxide surfaces (i.e. loose electrostatic attraction of

1019 H<sub>2</sub>O/H<sup>+</sup>/OH<sup>-</sup> covered surface with the absorbing ions; e.g., Dzombak and Morel, 1990). Thus,  
1020 Au(I) is expected to sorb via an inner-sphere mechanism, by forming direct Au-S-Fe or Au-S<sub>n</sub>-Fe  
1021 bonds with the pyrite surface, probably similar to that of arsenite, As(III) (As(OH)<sub>3</sub> and  
1022 AsO(OH)<sub>2</sub><sup>-</sup>), sorption on FeS<sub>2</sub>. Assuming the same surface site density for Au(I) as for As(III) (3.7  
1023 sites/nm<sup>2</sup>; Bostick and Fendorf, 2002), it can be estimated that the maximal concentration of  
1024 surface bound Au-S complexes for an ideal cubic-shape pyrite grain of 10 μm in size (having a  
1025 surface area of 6×10×10 μm<sup>2</sup>), representative of our experiments, would reach as much as 20,000  
1026 ppm Au if scaled to the mass of the grain. This sorption capacity is significantly superior to the  
1027 Au(I) concentrations in pyrite measured in this study (<200 ppm Au(I) at 450°C, Table 3), thus  
1028 supporting the linear trend with the concentration of Au(I) in the fluid shown in Fig. 8. Similar  
1029 sorption capacities of pyrite *vis-à-vis* As(III) in aqueous solution were recently reported in  
1030 siderite-to-pyrite replacement experiments at 200°C with As contents in pyrite reaching 10,000s  
1031 ppm (Kusebauch et al., 2018), and independently by thermodynamic models of arsenian pyrite -  
1032 fluid interactions at hydrothermal conditions (Xing et al., 2019). Furthermore, the plausibility of  
1033 our adsorption mechanism is independently supported by natural observations of a general  
1034 tendency to Au enrichment in small-grain pyrite from Carlin and MVT deposits (e.g., Reich et al.,  
1035 2005; Tauson et al., 2018), since smaller grains offer larger surface areas favorable for adsorption.

1036 Gold hydrogen sulfide and radical ions may attach to pyrite surface by forming -S-S-  
1037 bonds with sulfur atoms belonging to the corners of FeS<sub>6</sub> octahedrons (Fig. 9). Polysulfide sulfur  
1038 is ubiquitous on pyrite surfaces as systematically demonstrated by direct XPS measurements in a  
1039 wide range of solution compositions at near-ambient *T* (e.g., Nesbitt et al., 1998; Widler and  
1040 Seward, 2002; references therein) and may be a key factor for scavenging Au(I) at higher *T* as  
1041 well, further favored by existence of direct Au complexes with poly-sulfur ions (S<sub>3</sub><sup>-</sup> and S<sub>2</sub><sup>-</sup>) in  
1042 hydrothermal solution. The close resemblance of the XANES spectra of our Au-bearing pyrites to  
1043 the simulated spectra of Au-S<sub>3</sub><sup>-</sup> complexes (Fig. 5c) provides further support to this mechanism.  
1044 The uptake of polysulfide Au species should also be favored during early stages of pyrite  
1045 precipitation from hydrothermal fluids since pyrite formation occurs via polysulfide ion precursors  
1046 (e.g., Schoonen and Barnes, 1991).

1047 Our laboratory experiments probe the key early stages of Au incorporation in pyrite ore  
1048 such as Au scavenging by preexisting pyrite and simultaneous gold-pyrite precipitation that are  
1049 fast phenomena compared to natural timescales of ore deposit formation and evolution. The fate of  
1050 the sorbed Au-S species during the much slower processes of further pyrite crystal growth and  
1051 ageing in natural systems remains an open question in pyrite geochemistry. It is expected that  
1052 some part of these sorbed Au-S clusters are incorporated in defects and dislocations of the



1053 growing crystal owing to the lack of energetically favorable crystallographic position for Au(I) in  
1054 the As-free pyrite structure. Other, likely major, part of sorbed Au would be reduced into Au(0)  
1055 nanoparticles that may further migrate, grow, melt or get remobilized following cooling, thermal  
1056 overprint, pressure changes, and fluid-assisted metamorphism (e.g., Reich et al., 2006; Velazquez  
1057 et al., 2014; González-Jiménes et al., 2015). Future improvements in sensitivity and spatial  
1058 resolution of in-situ spectroscopic techniques for ppm-level metal concentrations will enable more  
1059 quantitative insight into the fate of invisible gold in pyrite during such processes.

1060 In the presence of significant amounts of arsenic in pyrite, however, the mechanisms and  
1061 driving forces of Au scavenging and long-term storage may be different, since arsenic may  
1062 considerably affect both surface electrical and chemical properties (e.g., Mironov et al., 1981;  
1063 Moeller and Kersten, 1994) and the crystal structure, for example, facilitating Au incorporation as  
1064 a solid solution. Although As-bearing systems are beyond the scope of this study, our new data on  
1065 fluid-pyrite partition coefficients of Au in As-poor systems indirectly support the key role of  
1066 arsenic in gold enrichment in iron sulfide minerals in nature, as will be shown below.

1067

#### 1068 *4.5. Role of arsenic in gold incorporation in pyrite*

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1070 The partition coefficients of Au(I) obtained in this study allow prediction of Au dissolved  
1071 contents in As-poor pyrite in nature. Detailed Au-As systematics of pyrites from various gold  
1072 deposits (Carlin, orogenic, porphyry, epithermal, OICG, VHMS, and Witwatersrand) has recently  
1073 been provided by Deditius et al. (2014). Their analysis demonstrates that invisible Au and As are  
1074 generally positively correlated in pyrite at As contents above 0.01-0.1 wt% and that the solubility  
1075 limit of chemically bound Au in arsenian pyrite is defined by an Au/As atomic ratio of 0.02  
1076 independently of the geochemical environment. Both Au and As tenors in natural pyrite generally  
1077 increase with decreasing  $T$  in the range 500-200°C.

1078 These natural observations are in excellent agreement with our experimental data showing  
1079 Au(I) pyrite-fluid partition coefficients at 350°C ( $D_{py/fl} \sim 10-50$ ) significantly greater than at 450°C  
1080 ( $D_{py/fl} \sim 0.2$ ). Gold concentrations in hydrothermal fluids forming porphyry-epithermal and  
1081 orogenic deposits, as inferred from both direct fluid inclusion analyses and thermodynamic  
1082 calculations of gold solubility, are typically less than 1 ppm, even though some high-temperature  
1083 sulfur-rich magmatic fluids ( $\geq 450^\circ\text{C}$ ) may exceptionally contain up to 30 ppm Au (e.g.,  
1084 Kouzmanov and Pokrovski, 2012; Garofalo et al., 2014; Pokrovski et al., 2014). Combining these  
1085 natural concentrations with the  $D_{py/fl}$  values derived in this study at 450°C yields Au(I) tenors in  
1086 pyrite of less than 3 ppm Au in high-temperature porphyry and orogenic environments. This is in

1087 agreement with the low Au(I) tenors in As-poor pyrites from such settings, which are largely  
1088 dominated by metallic gold in pyrite ore. For lower temperature epithermal, Carlin, volcanic  
1089 massive sulfide (VMS), and some orogenic deposits ( $\leq 350^\circ\text{C}$ ) that often lack analyzable fluid  
1090 inclusions, there is virtually no direct data on Au fluid contents; nevertheless, Au fluid-phase  
1091 concentrations are expected to be even smaller ( $< 0.1$  ppm Au) due to the much lower sulfide  
1092 mineral solubility (pyrite and pyrrhotite) and the resulting low concentrations of reduced sulfur  
1093 aqueous species ( $\text{H}_2\text{S}/\text{HS}^-$ ,  $\text{S}_3^{2-}$ ), which are the key ligands for Au transport as shown in this study  
1094 and previous work (e.g., Kouzmanov and Pokrovski, 2012; Pokrovski et al., 2015; references  
1095 therein). Such sub-ppm Au(I) fluid concentrations, combined with our  $D_{\text{py/fl}}$  values obtained at  
1096  $350^\circ\text{C}$  ( $\sim 10$ - $50$ ), would correspond to pyrite Au(I) contents of less than a few ppm. These  
1097 predicted Au concentrations are in excellent agreement with the solubility limit of chemically  
1098 bound gold in *arsenic-free pyrite* of 2 ppm Au (Reich et al., 2005). By contrast, our predicted  
1099 values are 1 to 3 orders of magnitude lower, on average, than the Au concentrations analyzed in  
1100 natural *arsenian pyrites* ( $> 0.01$ - $0.1$  wt% As) from these environments (Deditius et al., 2014;  
1101 references therein).

1102         Therefore, our experimental data obtained in As-free systems indirectly support the key  
1103 role of arsenic in gold enrichment in arsenian pyrites and arsenopyrites from such deposits. Our  
1104 results provide a sound physical-chemical interpretation of the famous wedge-shaped pattern of  
1105 Au solubility in arsenian pyrite, established in the seminal work of Reich et al. (2005) on the basis  
1106 of pyrite analyses from Carlin-type deposits, and more recently shown to be applicable to other  
1107 types of Au deposits. According to Reich et al.'s gold solubility model in pyrite, at low As  
1108 contents relevant to our study, Au(I) solubility is low ( $< 2$  ppm) and independent of As pyrite  
1109 content up to  $\sim 0.01$  wt% As, but increases proportionally to As at higher As contents (Fig. 9). Our  
1110 direct partitioning and spectroscopic data are in quantitative agreement with the lower-As part of  
1111 the solubility curve where Au(I) incorporation in pyrite is controlled by chemisorption as Au-  
1112 (poly)sulfide complexes, partly incorporated in defects and dislocations by the growing pyrite  
1113 crystal, and partly rejected from the structure by reduction to Au(0) during further pyrite life. In  
1114 contrast, for As-richer pyrites our data imply a fundamental change in the mechanism of Au  
1115 scavenging by the mineral. Arsenic would presumably favor both chemisorption and incorporation  
1116 of gold into the crystal structure as isomorphic constituent, but the physical-chemical mechanisms  
1117 and driving forces of this 'arsenic effect' and Au exact structural position yet remain to be fully  
1118 understood and quantified via integrated experimental, spectroscopic, and modeling studies.

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## 5. Concluding remarks

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1123 The key points of this study are the following:

1124 A combination of experimental, analytical and modeling approaches used in this work  
1125 enables new insight into one of the oldest enigmas of Ore Deposit Research, the nature of invisible  
1126 gold in pyrite. We show that gold is incorporated from hydrothermal fluids into As-poor pyrite  
1127 both as native metal, Au(0) in the form of nano- and submicron particles, and as chemically bound  
1128 gold of a nominal redox state of +1, Au(I).

1129 High-resolution XANES experimental and simulated spectra of Au-bearing pyrites  
1130 demonstrate that Au(I) is coordinated with two nearest sulfur atoms in a linear geometry, similar  
1131 to Au(I) complexes with the HS<sup>-</sup> and S<sub>3</sub><sup>2-</sup> ligands dominant in hydrothermal ore-forming fluids.  
1132 The Au(I) incorporation in pyrite at molecular level is likely to proceed via specific sorption of S-  
1133 Au-S aqueous species on the pyrite surface with formation of polysulfide moieties, [-S-Au-S-(S-  
1134 )<sub>n</sub>]. Our results rule out, at least in As-poor systems, the commonly accepted paradigm postulating  
1135 Au<sup>+</sup> substitution for Fe<sup>2+</sup> in the pyrite structure with formation of a solid solution as the mode of  
1136 chemically bound Au incorporation in pyrite.

1137 Combining the derived Au(I) content in pyrite with accurate Au concentration  
1138 measurements in the coexisting hydrothermal fluid from our experiments allows generation of a  
1139 first set of Au(I) partition coefficients between fluid and pyrite at *P-T* conditions typical of many  
1140 types of gold deposits. These coefficients imply relatively low dissolved gold contents, of less  
1141 than a few ppm Au, in As-poor pyrites (<0.01 wt% As), in full agreement with numerous natural  
1142 observations. Our estimations contrast with the elevated invisible Au contents, of 100 to 1,000s  
1143 ppm Au, commonly recorded in As-rich pyrites (>1 wt% As) and arsenopyrites (10s wt% As) in  
1144 many ore deposits; they thus highlight the fundamental control of As on Au incorporation in these  
1145 minerals. Quantitative assessment of the true role of arsenic in the gold fate in fluid-mineral  
1146 systems could only be achieved using a combination of cutting-edge experimental, spectroscopic  
1147 and modeling approaches.

1148 Furthermore, the results of this study highlight limitations of the existing thermodynamic  
1149 and speciation models of Au in S-bearing hydrothermal fluids that still persist despite more than  
1150 50 years of extensive research. In addition to 'traditional' Au complexes with HS<sup>-</sup> ligands, our  
1151 data confirm the existence, in high-temperature fluids (≥450°C), of Au species with other ligands  
1152 such as S<sub>3</sub><sup>2-</sup> and S<sub>2</sub><sup>2-</sup> ions whose nature, stability, and role in Au transport and scavenging by  
1153 sulfide minerals will require future integrated research.

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- 1407

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1409 **Table 1.** Batch reactor experiments: initial system compositions and gold dissolved concentrations determined  
 1410 from gold weight-loss and analyses of quenched fluid from runs conducted in this study at 450°C and 700 bar.

Run number	Mass of aqueous solution loaded, g	Solids loaded	Mass of Py/Po loaded, g	Initial system composition, mol/kg fluid			$\Delta m_{Au}$ , mg weight loss	Au in fluid and solid, ppm weight loss	Au in quenched fluid, ppm	Au in solid <sup>c</sup> , ppm
				<i>K<sub>2</sub>S<sub>2</sub>O<sub>3</sub></i>	<i>HCl</i>	<i>KOH</i>				
Py1	13.639	Py, Au	1.120	0.50	0.65	-	0.7	47	<150	ND
Py2	13.460	Py, Au	1.121	0.50	0.20	-	7.3	545	440	1260
Py3	16.843	Py, Au	0.659	0.50	0.20	-	<0.5	<30	237	ND
Py4	14.264	Py, Au	1.023	0.50	-	-	22.3	1631	1664	ND
Py5	15.246	Py, Au	0.730	0.50	-	0.31	39.0	2559	2832	ND
Py6 <sup>a</sup>	8.786	Py	0.935	0.50	0.20	-	<0.5	<60	2	ND
Py7 <sup>b</sup>	15.266	Py, Au	1.004	-	-	-	0.5	41	135	ND
Po1	15.324	Po, Au	0.308	0.50	0.20	-	10.0	652	426	6980
Po2	12.984	Po, Au	0.309	0.49	0	0.31	24.6	1608	946	13800
<i>error</i>	$\pm 0.010$ g		$\pm 0.005$ g				$\pm 0.5$ mg	$\pm 50$ ppm	$\pm 15\%$	$\pm 20\%$

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1412 <sup>a</sup> Au-free blank run for evaluating Au contamination from the autoclave and washing solutions.

1413 <sup>b</sup> S-poor Au-bearing blank run in water for evaluating detection limits for Au concentration determinations.

1414 <sup>c</sup> Derived from the difference of Au concentrations from weight loss and quenched fluid analyses and using the masses of the loaded solution and solid phase. ND here means not detected; it is applicable to cases where this difference is negative.

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**Table 2.** Flexible-cell reactor experiments: fluid composition and gold and other element dissolved concentrations measured in sampled fluids of K<sub>2</sub>S<sub>2</sub>O<sub>3</sub>(-HCl-KOH-FeCl<sub>2</sub>) composition at the indicated *P-T*-time conditions. Also reported are the initial aqueous solution compositions and concentrations of the compounds injected into the experimental fluid. Concentrations of major solute components are in mole per 1 kg of fluid (m); those of Au in ppm.

Run name, Sample number	Days from start	S tot	sulfide	sulfate	K	Cl	Fe	Au, ppm
<b>CoPy1 350°C, 370 bar; 0.49 m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub> + 0.14 m HCl</b>								
1	1	1.00	0.31	0.40	1.00	0.13	1e-4	653
2	3	1.07	0.31	0.42	1.02	0.13	2e-5	709
3	7	0.95	0.30	0.41	0.92	0.13	2e-5	625
4	10	0.97	0.30	0.39	0.97	0.13	1e-5	556
<i>Injection of KCl (0.11m)-BaCl<sub>2</sub>(0.090m)-FeCl<sub>2</sub> (0.18m)</i>								
5	14	0.26	0.0043	0.21	0.86	0.74	0.027	6
6	17	0.25	0.0050	0.20	0.88	0.74	0.023	16
7	24	0.24	<0.003	0.19	0.86	0.75	0.014	37
<b>CoPy2 350°C, 400 bar; 0.19 m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub> + 0.091 m KOH</b>								
1	1	0.37	0.16	0.14	0.58	<0.02	4×10 <sup>-5</sup>	474
2	3	0.33	0.11	0.13	ND	<0.001	5×10 <sup>-5</sup>	403
<i>Injection of FeCl<sub>2</sub> (0.056 m)</i>								
3	14	0.12	<0.001	0.12	0.49	0.20	0.0022	51
4	17	0.11	<0.001	0.084	0.44	0.19	0.0014	51
<b>CoPy3 450°C, 700 bar; 0.50 m K<sub>2</sub>S<sub>2</sub>O<sub>3</sub> + 0.30 m KOH</b>								
4	16	0.71	0.40	0.32	0.79	ND	ND	1552
5	20	0.69	0.35	0.24	0.78	ND	ND	1358
6	23	0.75	ND	ND	0.93	ND	ND	1365
7	27	0.59	0.37	0.28	0.60	ND	ND	1299
<i>Injection of HCl (0.172 m)</i>								
8	29	0.62	0.20	0.35	0.74	0.20	ND	561
9	32	0.57	0.19	0.31	0.68	0.18	ND	638
10	35	0.54	0.25	0.32	0.61	0.22	ND	563
<i>Injection of FeCl<sub>2</sub> (0.051 m)</i>								
11	37	0.43	0.13	0.33	0.68	0.27	ND	148
12	41	0.47	0.11	0.28	0.76	0.29	ND	49
13	44	0.48	0.084	0.30	0.75	0.28	ND	75
<i>Error</i>		±10%	±10%	±10%	±10%	±10%	±50%	±10%

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ND = not determined.

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**Table 3.** Total gold, visible and invisible native gold, and invisible chemically bound gold concentrations in pyrite from hydrothermal batch-reactor and flexible-cell reactor experiments of this study obtained using different analytical methods.

Run	Acid digestion and ICP-AES; Au total, ppm	SEM; Au visible			EPMA <sup>a</sup> ;	LA-ICPMS <sup>b</sup> ;	X-ray Fluorescence; Au invisible		% Au(I) LCF XANES of Au invisible	Au(I) chemically bound in pyrite, ppm <sup>c</sup>
		S <sub>Au</sub> , μm <sup>2</sup>	S <sub>Py</sub> , μm <sup>2</sup>	Au <sup>0</sup> , ppm	Au invisible, ppm (Nb pts)	Au invisible, ppm (Nb pts)	Δμ <sub>n</sub> ×10 <sup>4</sup>	Au invisible, ppm		
Py1	100	2.7	68017	150	180 (30)	ND	0.531	52	95	50
Py2	117	3.2	57915	210	180 (21)	81 (9)	0.598	59	>97	58
Py3	ND	ND	ND	ND	ND	ND	0.552	48	88	42
Py4	605	3.2	60696	201	180 (29)	ND	4.81	471	30	141
Py5	486	51	307653	639	190 (27)	180 (8)	2.08	204	>97	204
Po1	5591	470	82194	23941	ND	320 (9)	13.73	1190	4.5	54
Po2	20418	410	89092	19286	ND	ND	16.25	1593	7.3	116
CoPy1	ND	ND	ND	ND	4000 (20)	ND	36.57	3586	46	1664
CoPy2	ND	ND	ND	ND	5000 (45)	ND	129.7	12716	5	636
CoPy3	ND	ND	ND	ND	ND	6047 (5)	132.2	12957	<3	<390
Error	±5%			±50%	±50%	±50%	±15%	±20%	±3%	±25%

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<sup>a</sup> EPMA detection limit is 120 ppm Au; <sup>b</sup> LA-ICPMS typical detection limit is 0.1 ppm Au

<sup>c</sup> obtained as the product of Au(I) mole fraction and the total Au concentration from fluorescence spectra from equation (2); XRF detection limit is 1 ppm Au.

ND = not determined

1439

1440 **Table 4.** Adopted gold solubility and speciation in equilibrium with metallic gold and pyrite in S-rich  
 1441 hydrothermal fluid, and pyrite/fluid partition coefficient for chemically bound Au(I), obtained in this  
 1442 study from pyrite-fluid exchange and Au-pyrite co-precipitation experiments.

Run	Adopted Au solubility in fluid in equilibrium with gold metal and pyrite, ppm	Calculated fluid pH	Dominant Au(I) aqueous species in fluid	Au(I) chemically bound in pyrite, ppm	Partition coefficient $D_{py/fl}$ (Au(I)) <sup>a</sup>
<b>Pyrite-gold exchange experiments, 450°C, 700 bar</b>					
Py1	150	4.4	Au(HS)S <sub>3</sub> <sup>-</sup>	50	0.33
Py2	440	5.4	Au(HS)S <sub>3</sub> <sup>-</sup>	58	0.13
Py3	237	5.4	Au(HS)S <sub>3</sub> <sup>-</sup>	42	0.18
Py4	832	6.6	Au(HS)S <sub>3</sub> <sup>-</sup> + Au(HS) <sub>2</sub> <sup>-</sup>	141	0.17
Py5	1394	8.0	Au(HS) <sub>2</sub> <sup>-</sup> + Au-S <sub>2</sub> <sup>-</sup> + other complexes	204	0.15
Po1	426	5.8	Au(HS)S <sub>3</sub> <sup>-</sup>	54	0.13
Po2	946	8.1	Au(HS) <sub>2</sub> <sup>-</sup> + Au-S <sub>2</sub> <sup>-</sup> + other complexes	116	0.12
<b>Pyrite-gold co-precipitation experiment, 450°C, 700 bar</b>					
CoPy3	91	6.6	Au(HS) <sub>2</sub> <sup>-</sup> + Au(HS)S <sub>3</sub> <sup>-</sup>	<390 <sup>a</sup>	<4.2 <sup>b</sup>
<b>Pyrite-gold co-precipitation experiments, 350°C, 400 bar</b>					
CoPy1	37	4.5	Au(HS)S <sub>3</sub> <sup>-</sup>	1664	45.0
CoPy2	51	7.4	Au(HS) <sub>2</sub> <sup>-</sup>	636	12.5
<i>Error</i>	±20%	±0.2 unit		±25%	±30%

1443 <sup>a</sup> equation (4)

1444 <sup>b</sup> maximum value, since Au(I) in pyrite was below detection limit in this experiment

1445

## Figure captions

1446

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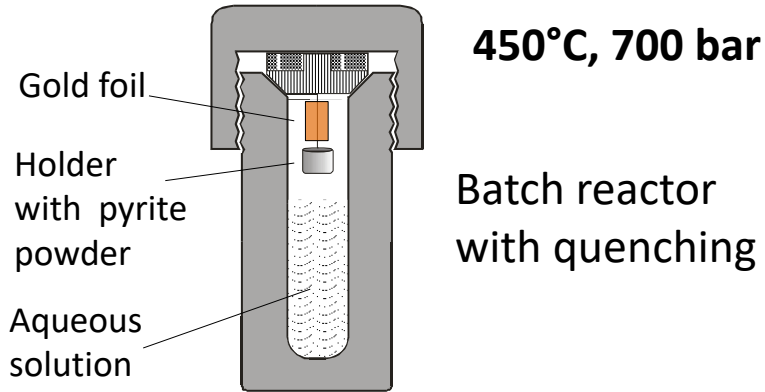
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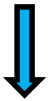
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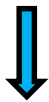
## Fluid-pyrite exchange of gold



Reactor quenching upon run termination



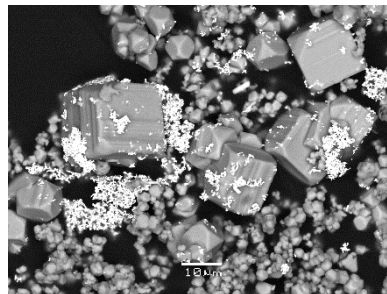
Au analysis of washout solution



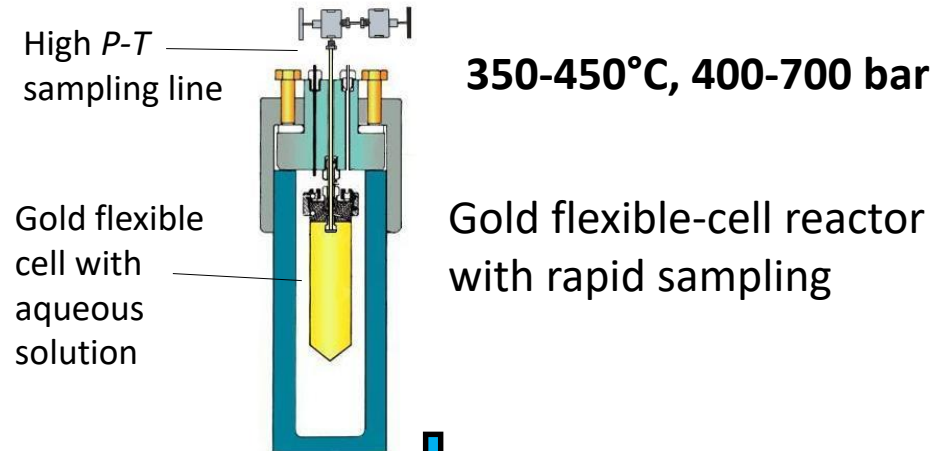
Gold foil weight loss



Pyrite analyses



## Gold-pyrite co-precipitation



Fluid sampling, metal and sulfur analyses



FeCl<sub>2</sub> injection to induce pyrite precipitation



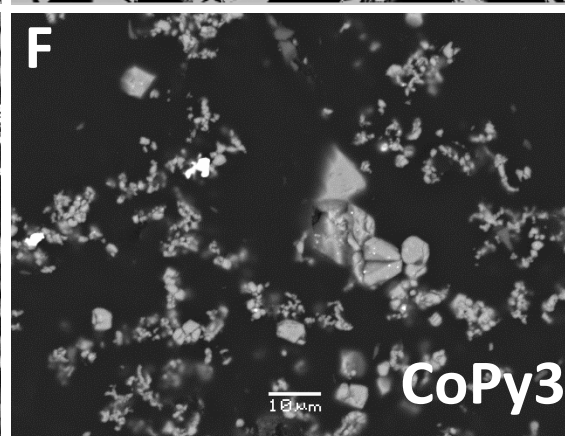
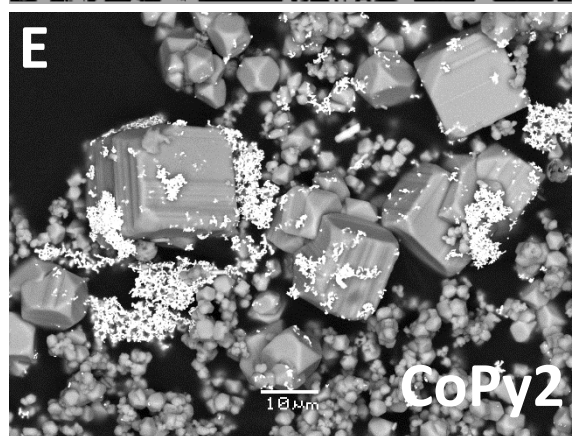
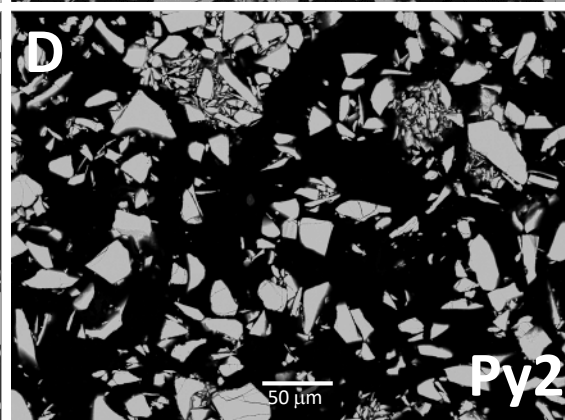
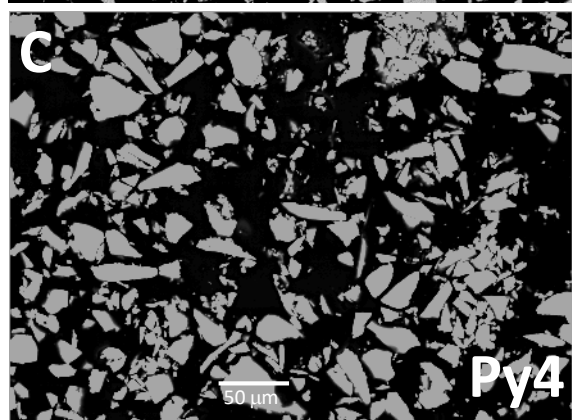
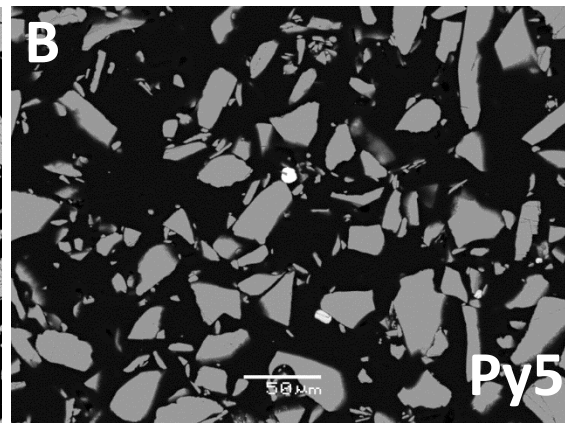
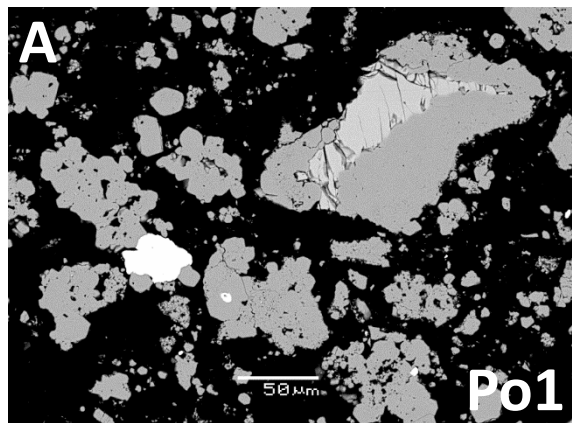
Fluid sampling, metal and sulfur analyses

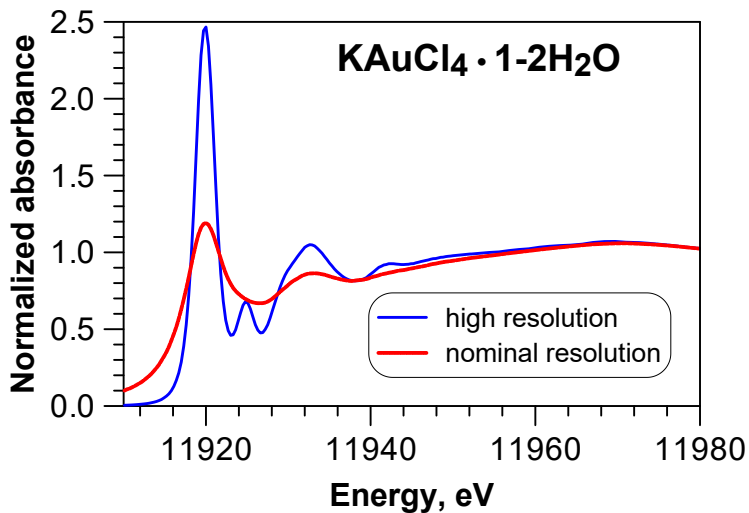
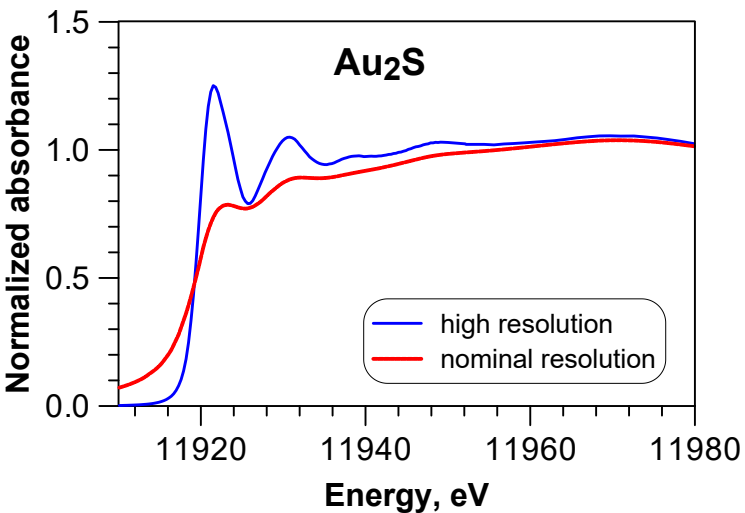
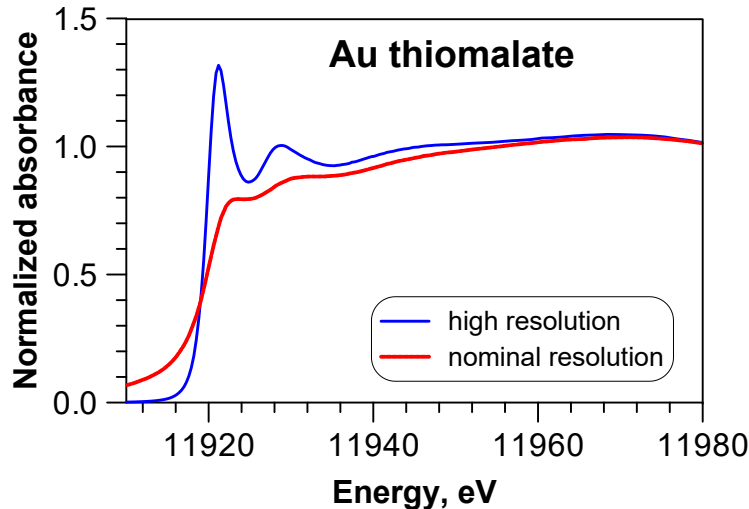
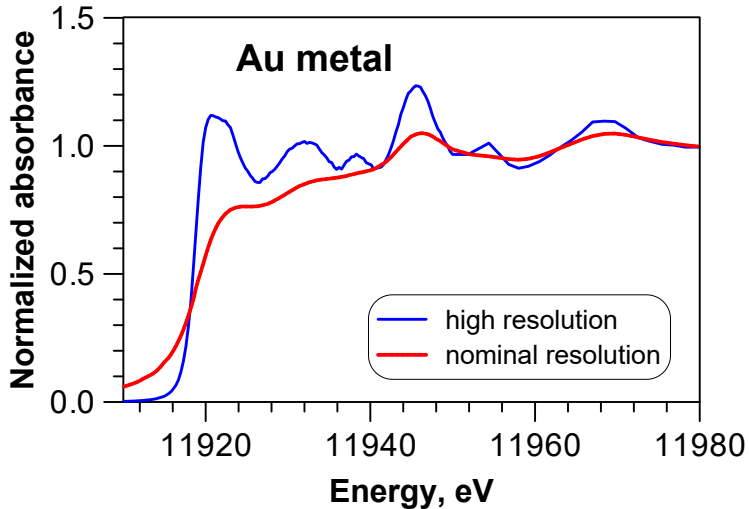


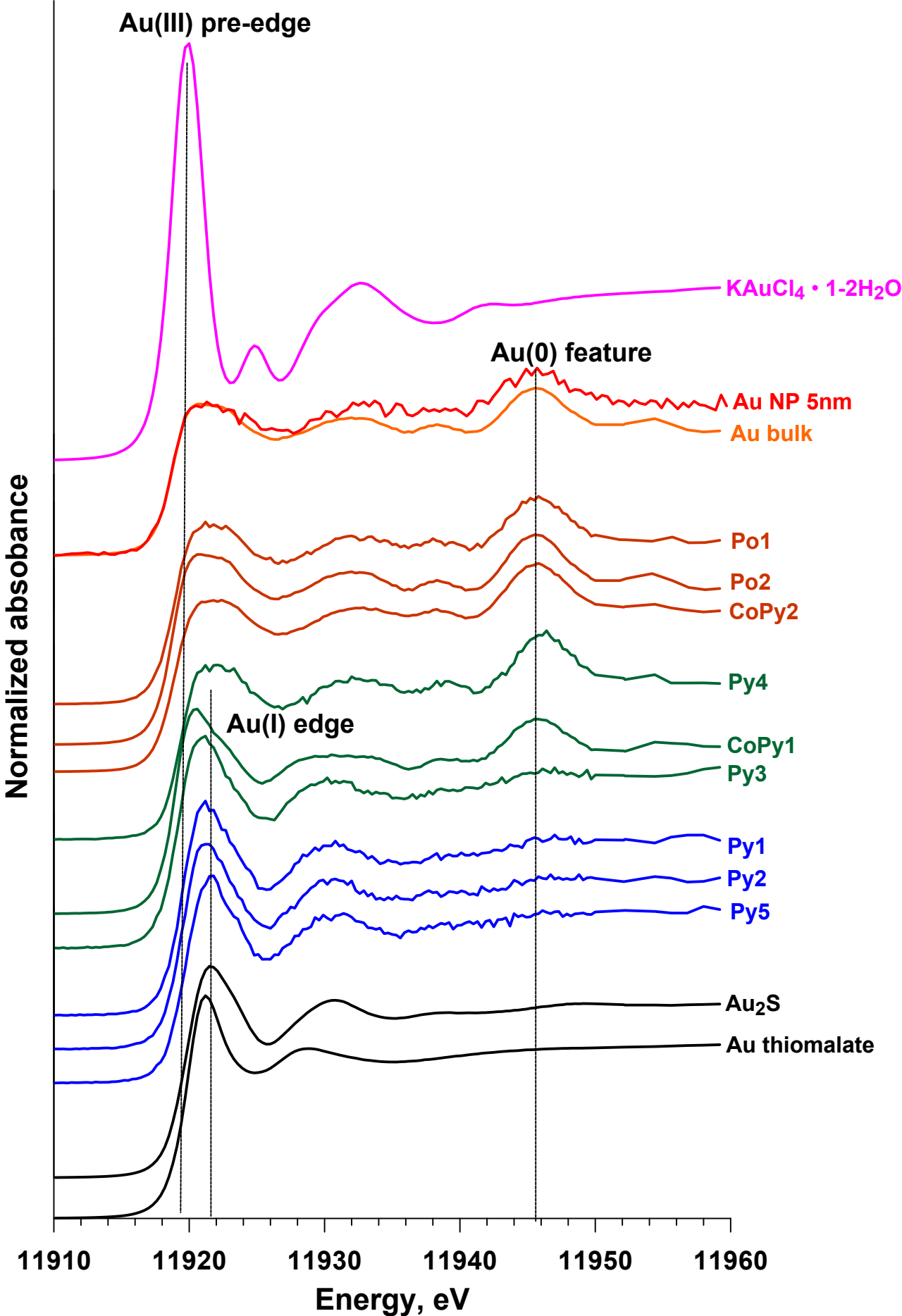
Reactor cooling

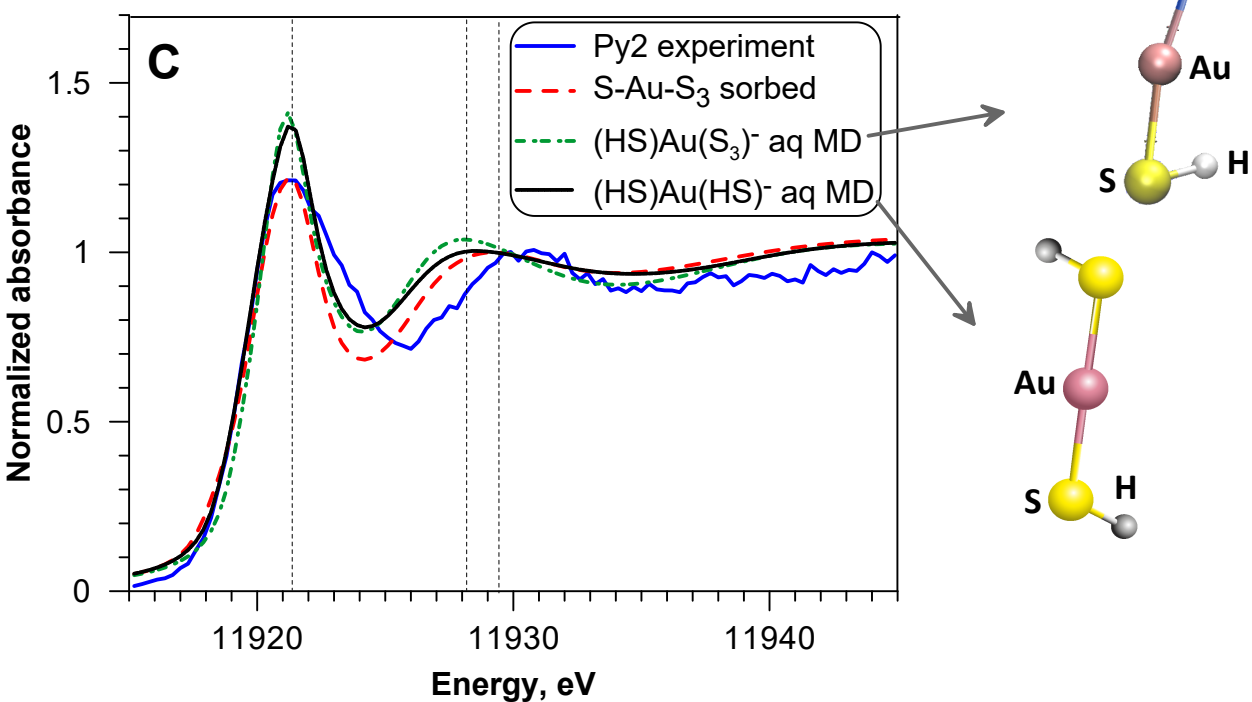
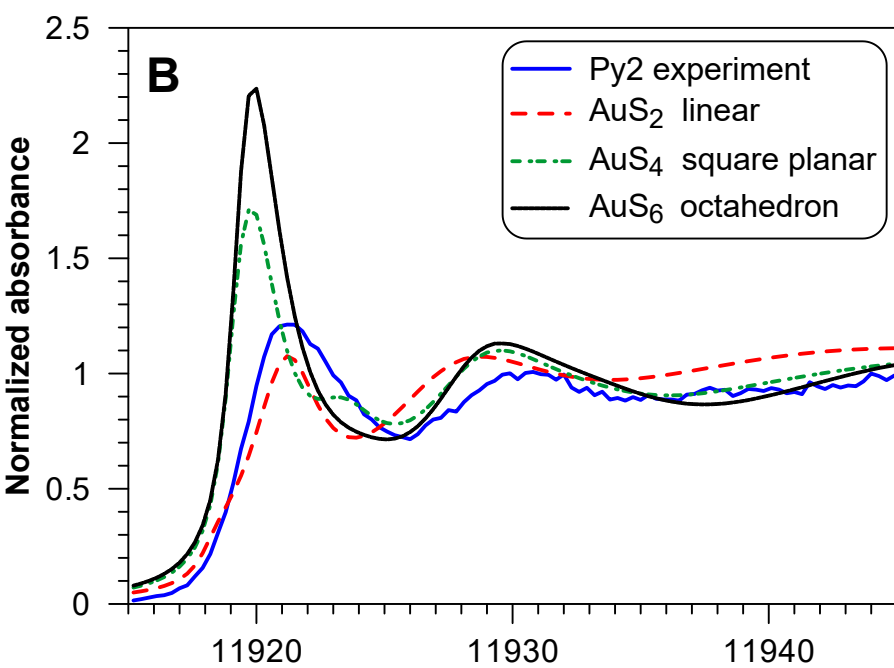
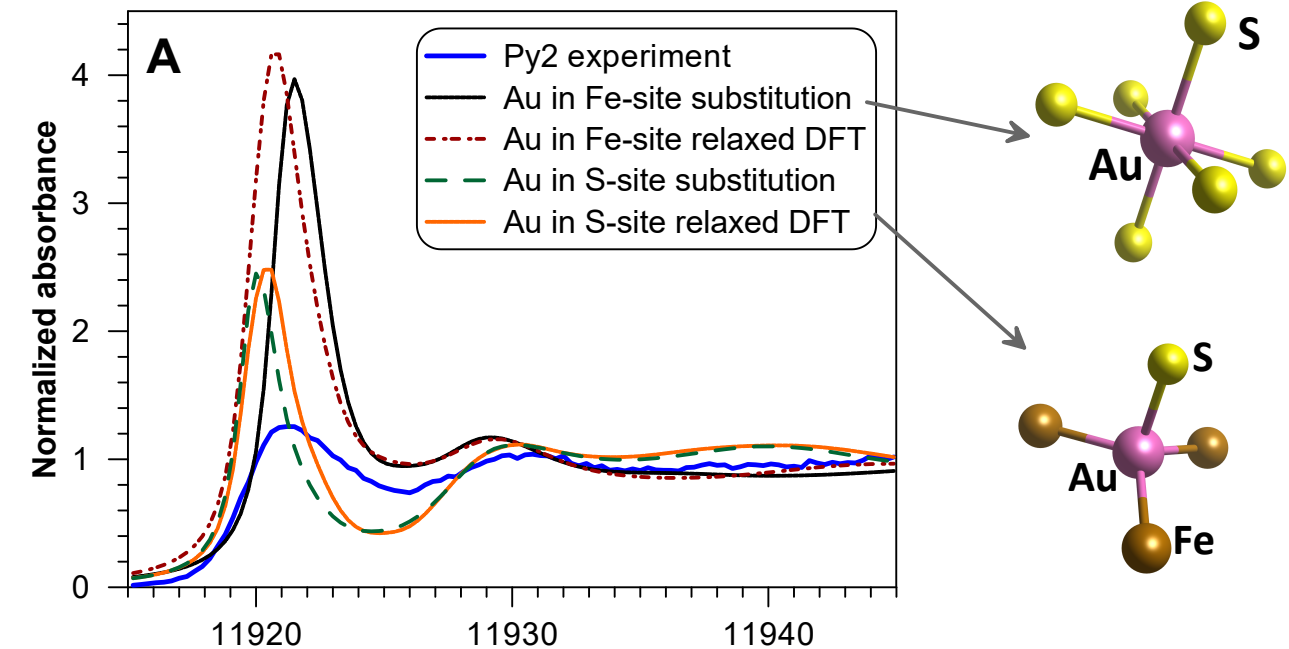


Pyrite analyses

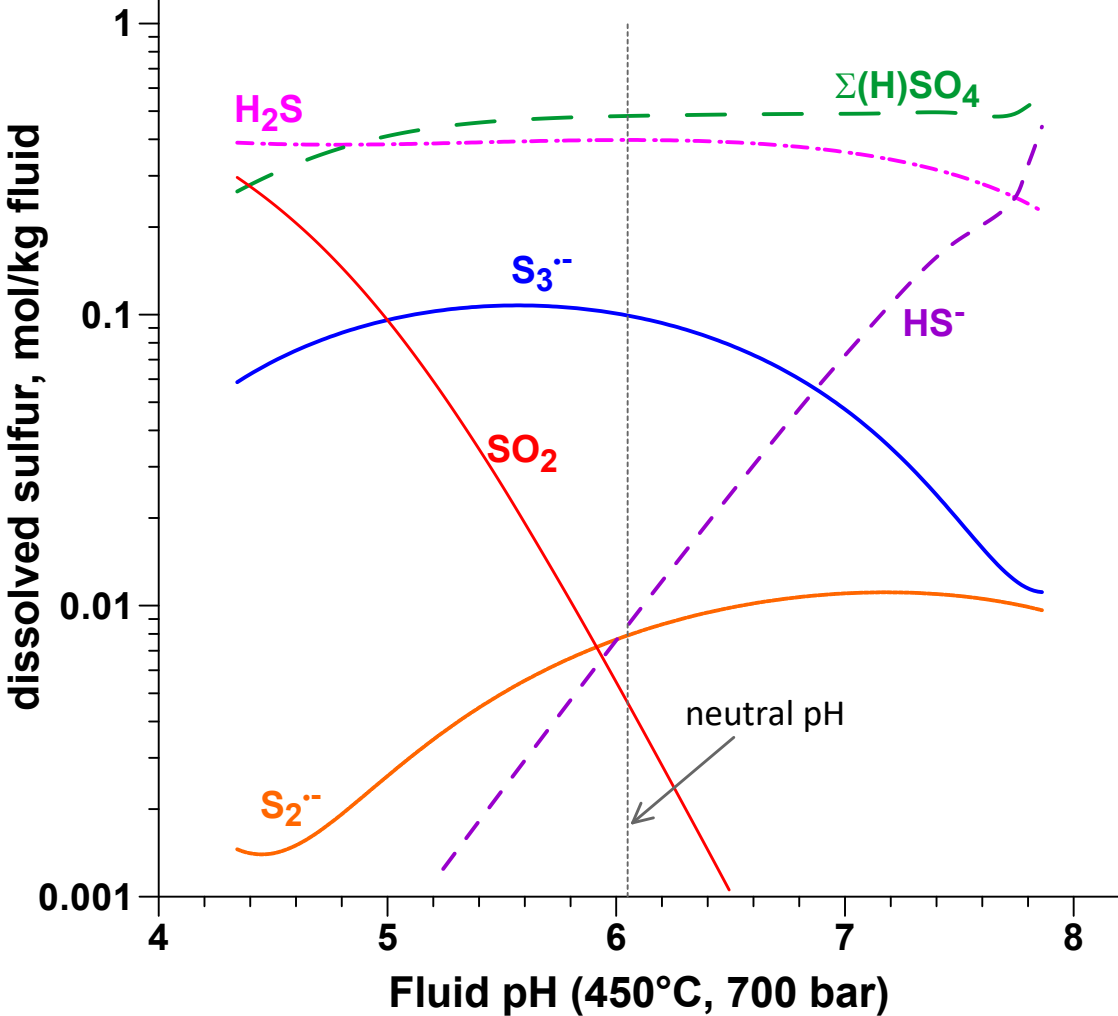




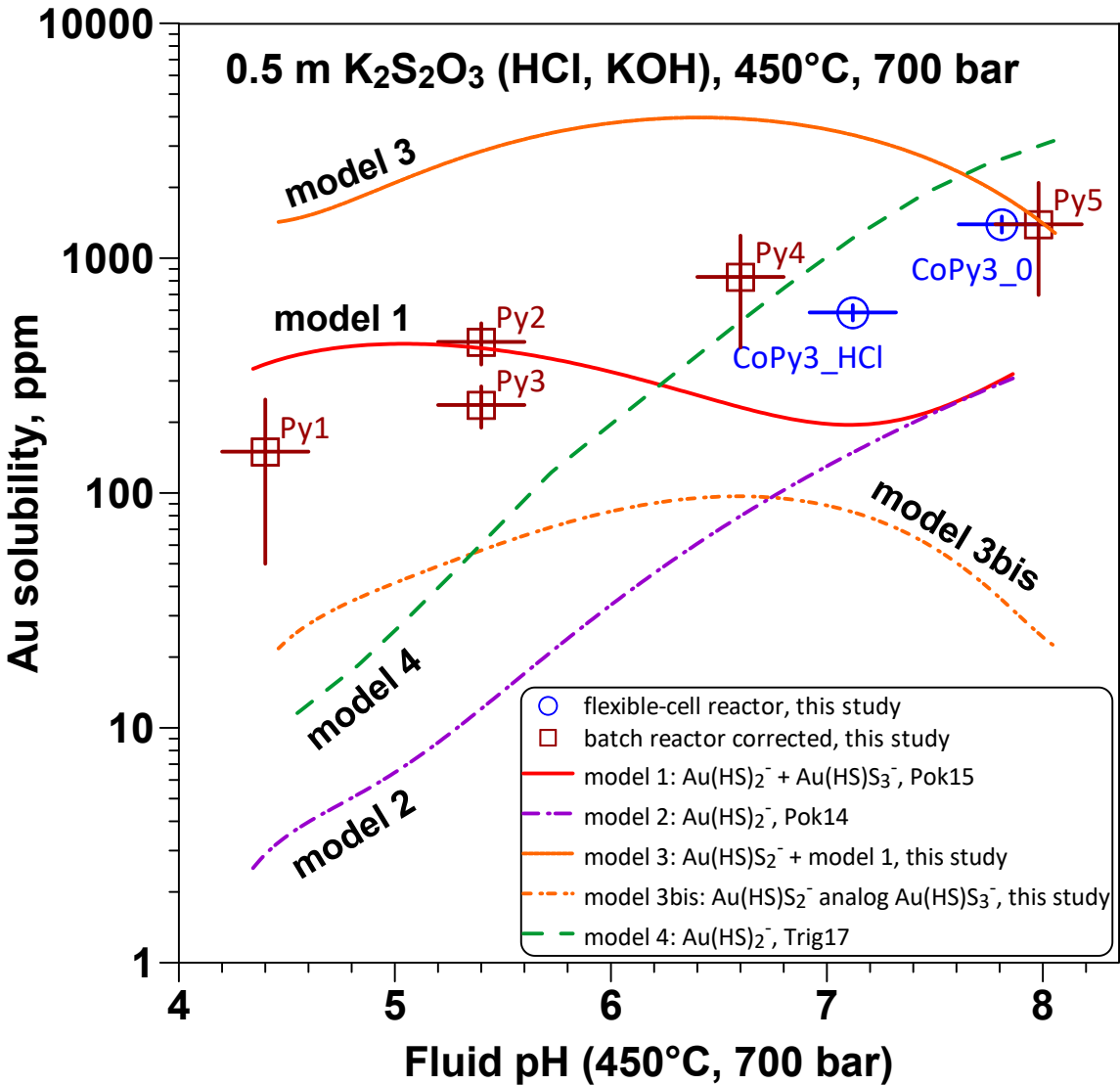


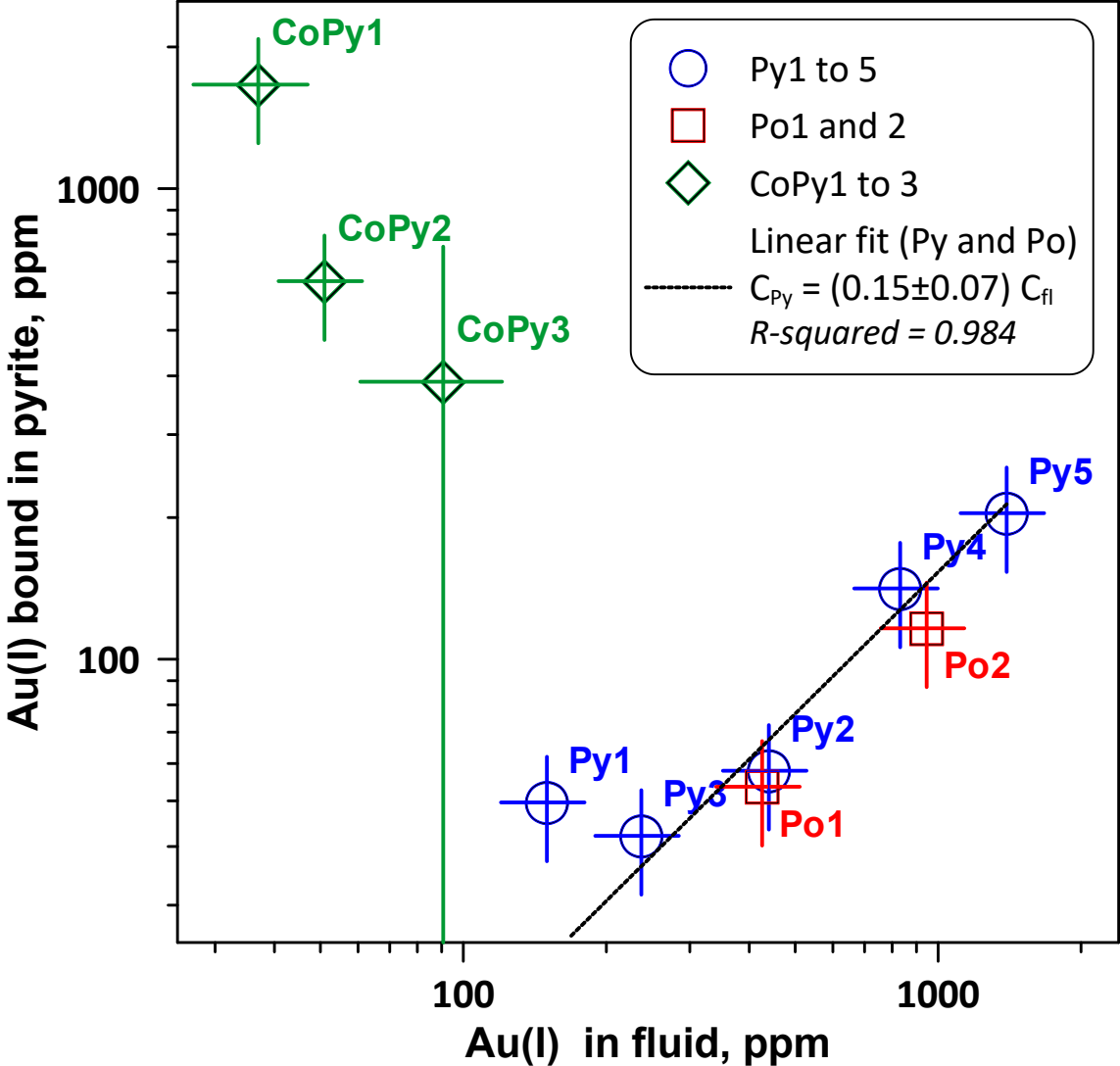


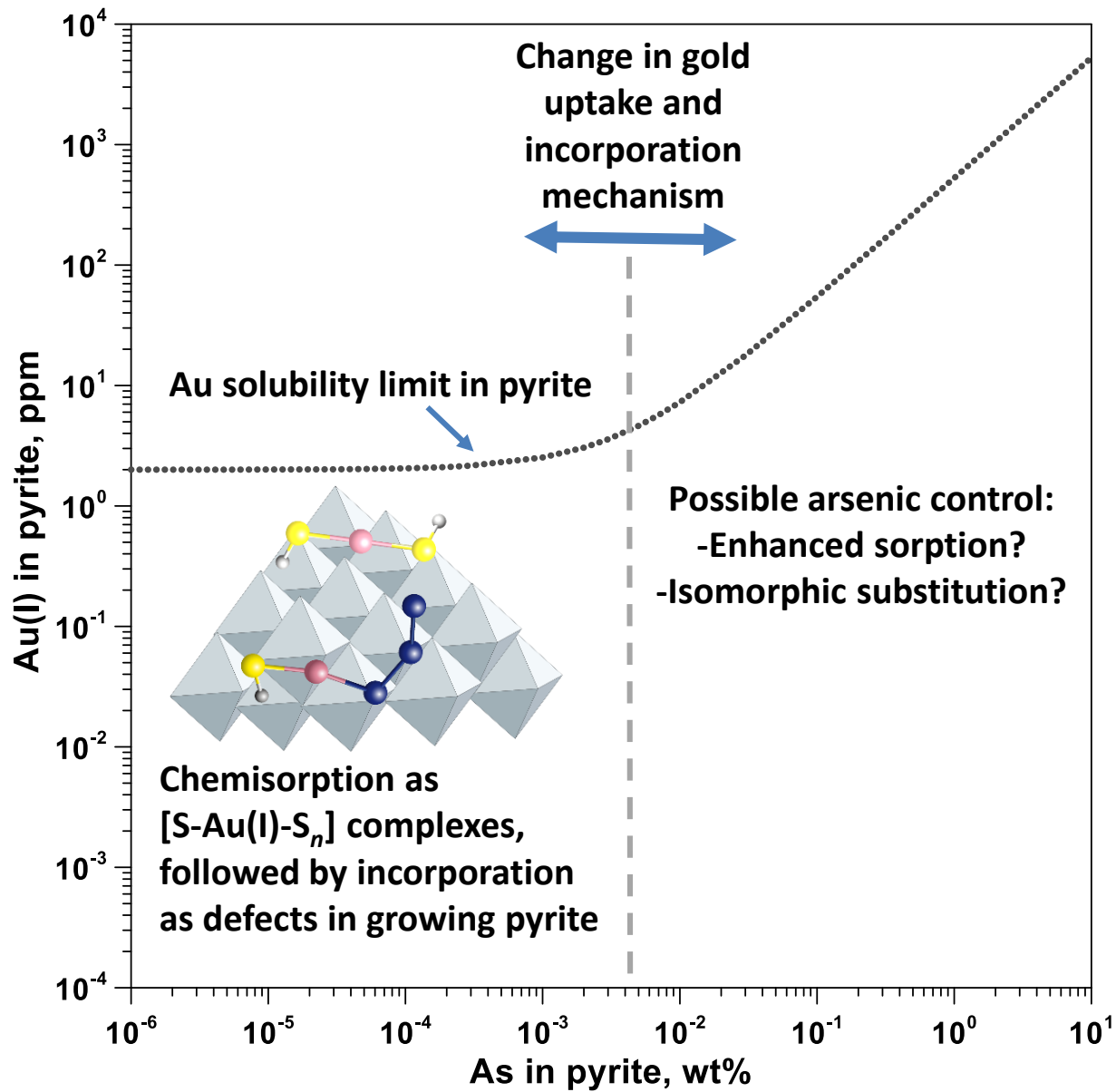
0.5 m  $\text{K}_2\text{S}_2\text{O}_3$  (HCl, KOH),  $450^\circ\text{C}$ , 700 bar



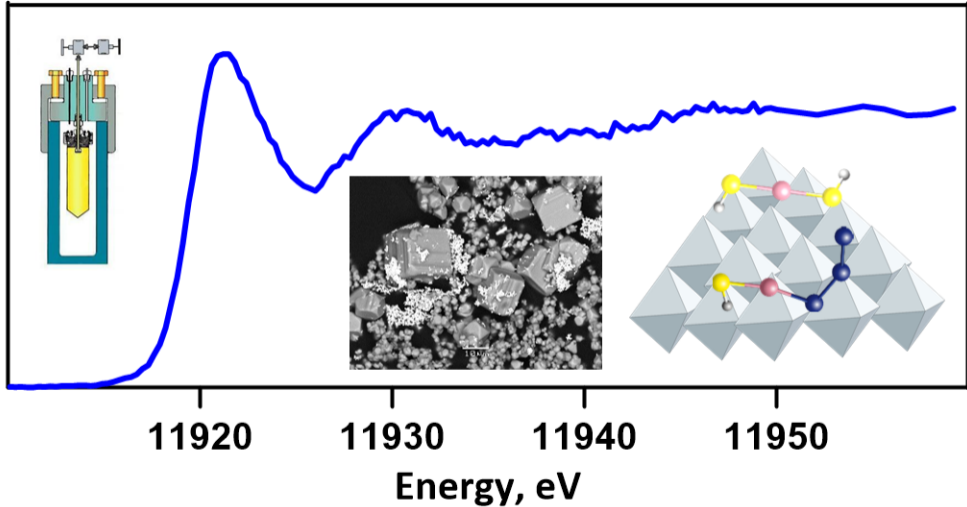








Absorbance



11920

11930

11940

11950

Energy, eV