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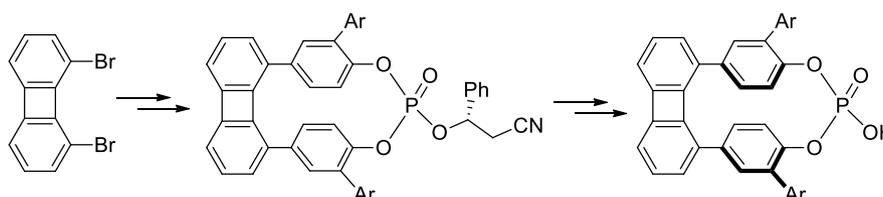
# Planar Chiral Phosphoric Acids With Biphenylene-tethered Paracyclophane Scaffolds: Synthesis, Characterization and Catalytic Screening

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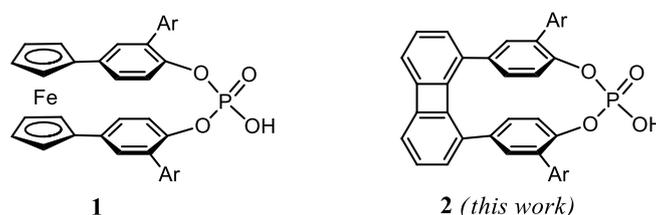
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**ABSTRACT:** Phosphoric acids with planar chiral paracyclophane scaffolds have been prepared in optically pure form starting from 1,8-dibromobiphenylene, by means of a chiral phosphorodiamidate as the phosphorylating agent. Structural characterization and configurational assignment have been performed by X-ray diffraction studies. The acids promote the organocatalytic enantioselective H-transfer reduction of  $\alpha$ -arylquinolines with up to 90% enantiomeric excess.

## Introduction

The widespread uses of phosphoric acids and phosphates as chiral acids,<sup>1</sup> chiral anions and ligands<sup>2</sup> are among the major achievements of modern enantioselective catalysis. In this field, the atropochiral Binol derived phosphoric acids are privileged auxiliaries, but alternative scaffolds have also been developed, including Spinol<sup>3</sup> and Taddol<sup>4</sup> derivatives, which display axially chiral structures and stereogenic carbons respectively. (For recent references, see:<sup>5</sup>) The above series of chiral phosphoric acids might be suitably complemented by analogous derivatives exploiting planar chiral scaffolds. In this context, having noticed that a single example of planar chiral phosphoric acids had been reported,<sup>6</sup> we have initiated a research project oriented toward the development of phosphoric acids of this class. In our previous papers<sup>7</sup> we have disclosed the first series of planar chiral phosphoric acids represented hereafter by the general formula **1** (**Figure 1**).



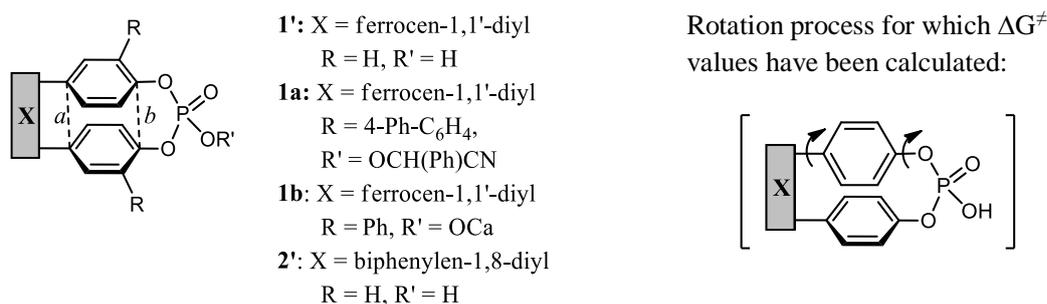
**Figure 1.** Planar chiral phosphoric acids based on paracyclophane scaffolds.

They are paracyclophanes with a 1,1'-ferrocendiyl and an O-P-O units tethering the aromatic rings in their *para*-positions. Two aryl substituents (Ar) generate the desired planar chirality while giving a formally C<sub>2</sub>-symmetric scaffold. The 1,1'-ferrocendiyl unit had been targeted initially as a suitable linker, based on DFT studies showing that it will give moderate ring strain, combined with satisfying

configurational stability of the paracyclophane. Acids **1** proved to be good catalysts for the enantioselective H-transfer reduction of 2-substituted quinolines (e.e. up to 92%). They displayed however an only moderate thermal stability, giving an estimated 10% decomposition rate after 18h heating at 60°C. Their uses will be therefore reasonably restricted to catalytic reactions taking place at room temperature or lower. To hopefully overcome this limitation, we have turned then our attention to analogous paracyclophane-type derivatives bearing different tethering chains. This paper reports on our studies oriented toward this goal and especially toward the use of a biphenylen-1,8-diyl unit as the tethering chain, as typified by compound **2** in **Figure 1**.

## Results and Discussions

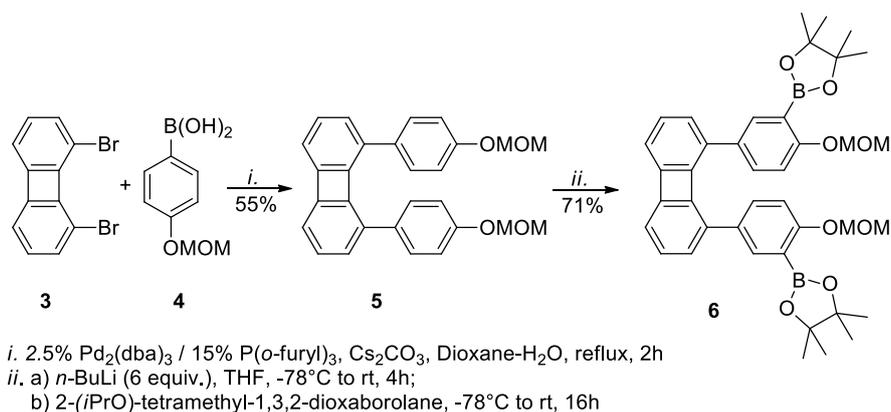
In this work, with the aim of accessing new phosphoric acids with paracyclophane structures, we have envisioned the biphenylen-1,8-diyl unit as a possible paracyclophane tethering chain. Compared to the previously used 1,1'-ferrocendiyl motif of **1**, the four atom biphenylene tether is anticipated to give slightly longer C-C distances between the aromatic rings of the paracyclophane: *a* and *b* distances have been calculated indeed at 3.76 and 3.13 Å respectively for **2'** vs. 3.51 and 3.09 Å for **1'**, at the M06/6-31G(d,p) DFT level (Figure 2). The ring strain energy has been calculated<sup>8</sup> however to be slightly higher for **2'** with respect to **1'** (24 kJ/mol vs 17 kJ/mol), which is likely related to the rigidity of the biphenylene structure, compared to the flexible ferrocene scaffold. The activation energy required for the rotation of the aromatic ring around the C- -O axis ( $\Delta G^\ddagger$  in **Figure 2**) has been calculated at 180 kJ/mol for **2'**, i.e. slightly lower than for the ferrocenic derivative **1'** (216 kJ/mol), but still high enough to likely ensure the configurational stability of the molecule.



	<i>a</i> (Å)	<i>b</i> (Å)	$\Delta G^\ddagger$
<b>Ferrocen-1,1'-diyl derived paracyclophanes, 1</b>			
Calculated values for <b>1'</b>	3.51	3.09	216 kJ/mol
Experimental values for <b>1a</b> <sup>7b</sup>	3.495	3.099	
Experimental values for <b>1b</b> <sup>7b</sup>	3.404	3.022	
<b>Biphenylen-1,8-diyl derived paracyclophanes, 2</b>			
Calculated values for <b>2'</b>	3.76	3.13	180 kJ/mol

**Figure 2.** Calculated/experimental structural parameters and rotation barriers for paracyclophane-type phosphoric acids **1** and **2**

Based on these rather encouraging calculation results, we have investigated then synthetic approaches to acids **2**. To this end, we have adapted the strategy previously used for the synthesis of **1**, which made use of two sequential Suzuki couplings to create both the ferrocene-aryl and the aryl-aryl bonds.<sup>7b</sup> When applied to the synthesis of **2**, the method involves 1,8-dibromobiphenylene **3** as the starting material, and the two Suzuki coupling steps shown in **Scheme 1** and 2. The bis-boronate **6**, diol **8** and the chiral phosphates **10** in Scheme 3 are the key intermediates.

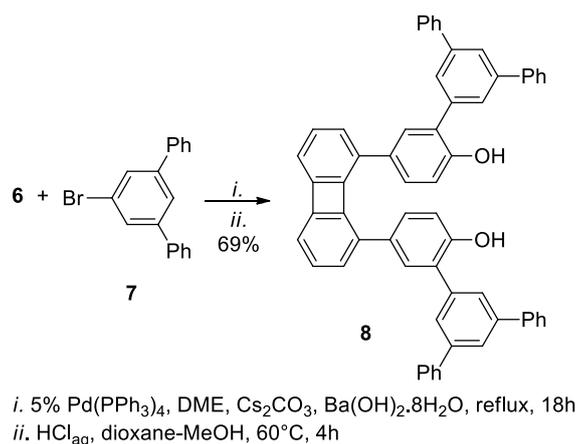


**Scheme 1.** Synthesis of the bis-boronate **6**

The synthetic procedure starts with a Suzuki coupling between 1,8-dibromobiphenylene, **3**,<sup>9</sup> and the aryl-boronic acid **4**<sup>10</sup> which displays a MOM-protected hydroxyl function in its *para*-position. The reaction takes place in a dioxane-water mixture, in the presence of a 2.5 mol% amount of Pd<sub>2</sub>(dba)<sub>3</sub> and 15 mol% of tri(*o*-furyl)phosphine. The desired product **5** has been isolated in 55% yield at a 17 g scale. Other catalysts and conditions previously reported for the double arylation of **3** (e.g. Pd(PPh<sub>3</sub>)<sub>4</sub>)<sup>11</sup> did not afford improved yields.

Compound **5** has been converted then into the corresponding bis-boronate **6** by a metalation-borylation reaction<sup>12</sup> taking advantage from the *ortho*-directing effect of the O-MOM functions: addition of an excess *n*-BuLi to **5** in THF, afforded the dilithiated derivative which was trapped then with isopropyl pinacolborate. The bis-boronate **6** has been isolated in 71% yield at a 19 g scale.

As demonstrated in our previous work on the analogous series of ferrocene-derived paracyclophanes **1**, this compound should represent a useful platform for the introduction of various substituents on the aryl rings by coupling reactions. This step has been illustrated here by the synthesis of the highly hindered *meta*-terphenyl-substituted diol **8** by palladium-catalyzed Suzuki coupling, under Pd(PPh<sub>3</sub>)<sub>4</sub> catalysis (**Scheme 2**).

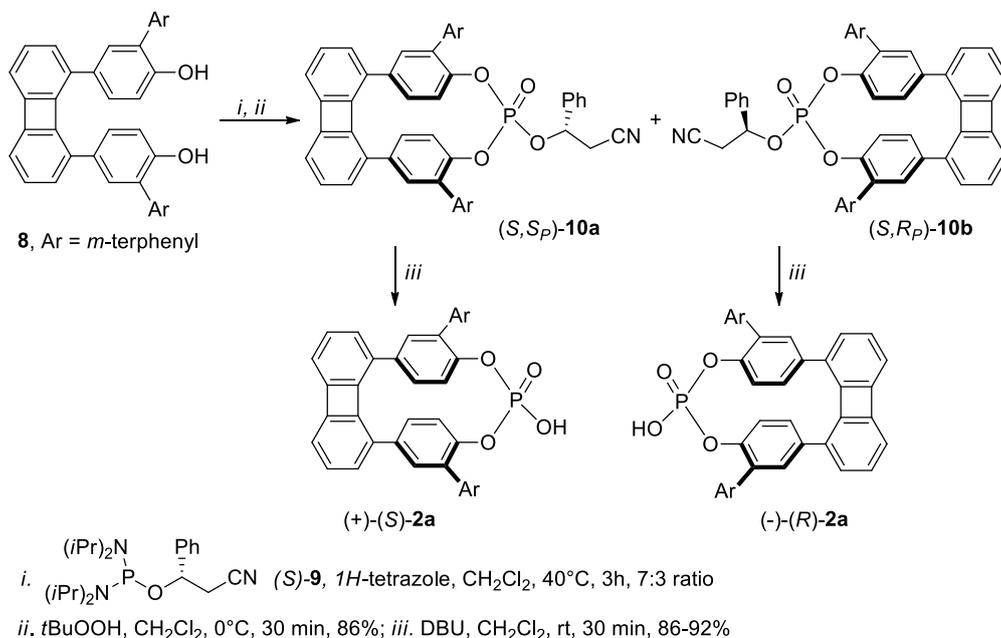


**Scheme 2.** Synthesis of diol **8**

After *in situ* removal of the MOM protecting groups from the intermediate coupling product, the desired diol **8** has been isolated in 69% yield. The <sup>1</sup>H and <sup>13</sup>C NMR spectra of **8** show single sets of signals that reveal unhindered rotation of the aryl substituents of the biphenylene unit.

The synthesis of the desired phosphoric acid in optically pure form has been carried out then by reacting diol **8** with the chiral phosphinating agent **9**, according to our recently established procedure<sup>7b,13</sup> (**Scheme 3**). The cyclization reaction of diol **8** with phosphoramidite **9** in the presence

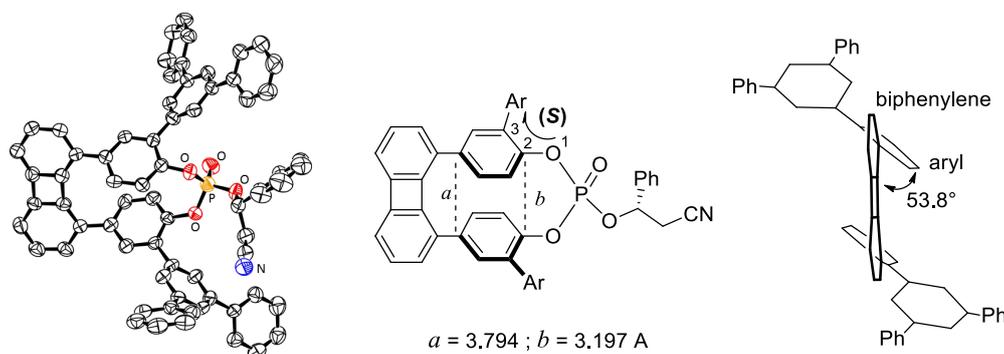
of *1H*-tetrazole, followed by oxidation of the resulting phosphites with *tert*-butylhydroperoxide, afforded a 7:3 mixture of the corresponding cyclic phosphates **10a** and **10b**. In principle, the cyclization reaction might afford also the third isomer with non-chiral configuration of the paracyclophane scaffold, in which the two aryl substituents would be located on the same face of the paracyclophane. This isomer has not been detected in the reaction mixture. Phosphates **10a,b** have been obtained in high, 86% yield over two steps. The cyclization step takes place here in much better yield than the analogous cyclization leading to the ferrocenic acids **1** (86% vs 40%). This undoubtedly relates to the rigid biphenylene unit, which forces the aromatic rings of **8** in a suitable, almost parallel arrangement.



### Scheme 3. Synthesis of the planar chiral phosphoric acids (+)-**2a** and (-)-**2a**

The two epimers of **10** could be separated easily by semi-preparative HPLC (NW50 silica gel column, heptane/toluene/THF 4/95/1, retention times: 10.0 min for (*S,S,P*)-**10a** ( $[\alpha]_D = +68$  ( $c=1$ ,  $CHCl_3$ )) and 13.0 min for (*S,R,P*)-**10b** ( $[\alpha]_D = -123$  ( $c=1$ ,  $CHCl_3$ )) and fully characterized. Thus, although the chiral auxiliary (*S*)-**9** induced an only moderate diastereoselectivity in this cyclization reaction, HPLC separation of the epimers afforded synthetically useful amounts of the pure compounds **10a** and **10b**.

The configuration of the epimeric paracyclophanes has been assigned from the X-ray crystal structure of **10a** shown in **Figure 3** (CCDC 1014151). This epimer displays an  $S_P$  configuration (relative configuration with respect to the known (*S*)-configuration of the chiral auxiliary).<sup>14</sup> The structural parameters of (*S,S,P*)-**10a** from X-ray data are in good agreement with the calculated parameters in **Figure 2**. Notably the key non-bonding *a* and *b* distances which define the paracyclophane geometry and modulate the ring strain, have been measured at 3.794 Å and 3.197 Å respectively, vs 3.76 Å and 3.13 Å for calculated values. With respect to the analogous ferrocene tethered paracyclophane **1a**, the larger macrocyclic ring of **10a** generates a slightly larger O-P-O intracyclic bond angle (106.89° for **10a**, vs 105.76° for phosphate **1a**<sup>7b</sup>), while the P-O distances measure about 1.60 Å in both series.



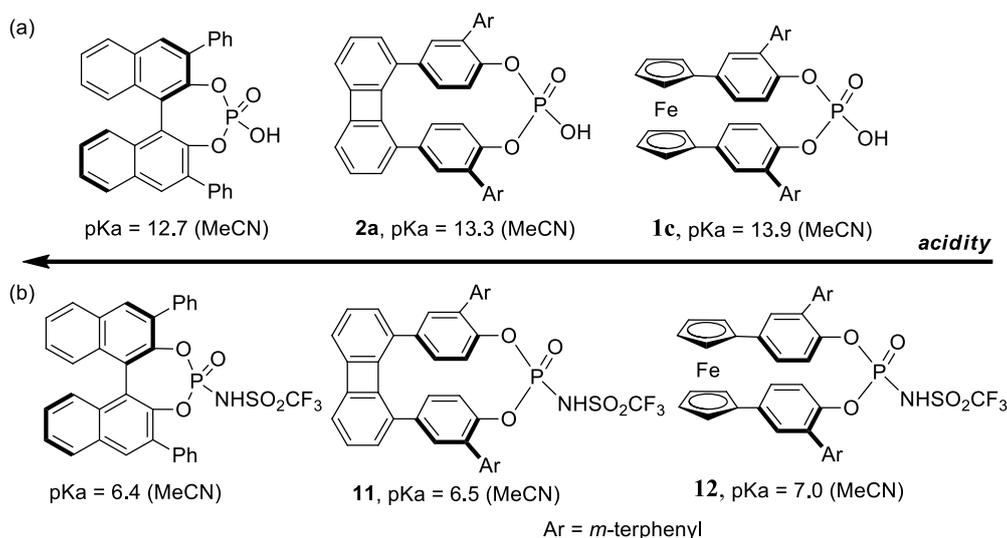
**Figure 3.** X-ray crystal structure of  $(S,SP)$ -**10a**, stereochemical descriptor and key geometrical features.

X-ray data highlight significantly different geometrical features for the ferrocene based paracyclophane **1a** and the biphenylene analogue **10a**, with respect to their three-dimensional structure, in the solid state at least. Indeed, in phosphate  $(S,SP)$ -**10a** the dihedral angle between the parallel aryl planes and the biphenylene-tether plane measures  $53.8^\circ$ , while in the ferrocene-based paracyclophane **1a** the parallel aryl planes are almost orthogonal to the median plane of the ferrocene tether ( $83.03^\circ$  dihedral angle). The observed differences in the geometrical features are anticipated to cause slightly different catalytic behaviors for the two series of paracyclophanic acids, **1** and **2**.

As shown in **Scheme 3**, phosphates  $(S,SP)$ -**10a** and  $(S,RP)$ -**10b** have been converted into the desired phosphoric acids (+)- $(S)$ -**2a** ( $[\alpha]_D = +90$  ( $c = 0.5$ ,  $\text{CHCl}_3$ )) and (-)- $(R)$ -**2a** respectively, by removal of the 1-phenyl-2-cyanoethyl substituent in the presence of DBU and subsequent washing of the resulting DBU-phosphate with an HCl 6N solution. The acids have been obtained in 86-92% yields at a 1g to 2 g scale, as colorless solids. The enantiomeric purity of these acids has been established by HPLC on an ID column with heptane/THF/TFA/ $\text{Et}_3\text{N}$  (40:60:0.5:0.3) as the eluent: retention times 3.9 and 5.9 min ( $ee > 98\%$ ).

In order to ascertain the configurational stability of the biphenylene-based [4.3]-paracyclophane scaffold of **2a**, we have heated a sample of acid  $(S)$ -**2a** ( $>98\%$  ee) at  $110^\circ\text{C}$  in toluene. Chiral HPLC analysis didn't show any epimerization after heating for 16h. The sample did not show significant decomposition either, so that the experiment also demonstrates that acid **2a** is chemically more stable than the ferrocene-based acids **1** which start to decompose at about  $60^\circ\text{C}$ .<sup>7b</sup> Thus, the biphenylene based acid **2a** has the potential of overcoming the limitations of the ferrocene-based acids in terms of thermal stability, while overall retaining the same key structural features related to planar chirality.

For comparison purposes, pKa values have been measured for the biphenylene derived acid **2a** and the ferrocene-derived acid **1c**.



**Figure 4.** (a)  $\text{pK}_a$  values for the phosphoric acids **2a**, **1c** and a representative binol-derived acid;<sup>15</sup> (b)  $\text{pK}_a$  values for the corresponding sulfonimides

The  $\text{pK}_a$  values have been obtained in MeCN, according to the published procedure.<sup>15</sup> With a  $\text{pK}_a$  value of 13.3, the biphenylene derived acid **2a** is slightly more acidic than the ferrocenic acid **1c** ( $\text{pK}_a = 13.9$ ), both of them being less acidic than the 3,3'-diphenyl-1,1'-binol derived phosphoric acid ( $\text{pK}_a = 12.7$ ) (**Figure 4**). As expected, the  $\text{pK}_a$  values remain however in the same range.

A decrease of the acidic character is observed also, although at a lower degree, in the mixed phosphoric-sulfonic imides, when going from Binol derived imides ( $\text{pK}_a = 6.4$ ) to paracyclophanic imides with biphenylene (**11**,  $\text{pK}_a = 6.5$ ) and ferrocene tethers (**12**,  $\text{pK}_a = 7.0$ ).<sup>16</sup> The lower sensitivity of  $\text{pK}_a$  values towards structural changes in the imides series is reasonably expected, since most of the negative charge in the corresponding anions is delocalized on the strongly electron-withdrawing triflyl group. Structural changes in the remaining part of the molecule should have only minor effects on  $\text{pK}_a$  values.

The last step of this study has been to validate the catalytic efficiency of the newly designed acid **2a** by preliminary screenings in a model reaction, i.e. the organocatalytic H-transfer hydrogenation of the  $\alpha$ -arylquinolines **13** with Hantsch esters.<sup>17</sup> Results are shown in **Table 1**.

**Table 1.** Enantioselective H-transfer hydrogenation of  $\alpha$ -arylquinolines

entry	Substrate <sup>a</sup>	Ar	Cat.	R (Hantsch ester)	Product	ee
1	<b>13a</b>	Ph	( <i>S</i> )- <b>2a</b>	Et	<b>14a</b>	77%
2	<b>13a</b>	Ph	( <i>R</i> )- <b>1c</b>	Et	<b>14a</b>	60% <sup>7b</sup>
3	<b>13a</b>	Ph	( <i>S</i> )- <b>2a</b>	CH <sub>2</sub> (4-BnO-C <sub>6</sub> H <sub>4</sub> )	<b>14a</b>	78%
4	<b>13a</b>	Ph	( <i>R</i> )- <b>1c</b>	CH <sub>2</sub> (4-BnO-C <sub>6</sub> H <sub>4</sub> )	<b>14a</b>	87% <sup>7b</sup>
5	<b>13a</b>	Ph	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14a</b>	<b>82%</b>
6	<b>13a</b>	Ph	( <i>R</i> )- <b>1c</b>	<i>t</i> -Bu	<b>14a</b> <sup>b</sup>	70% <sup>7b</sup>
7	<b>13b</b>	2-naphthyl	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14b</b>	<b>90%</b>
8	<b>13c</b>	<i>p</i> -biphenyl	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14c</b>	<b>87%</b>
9	<b>13d</b>	4-MeO-C <sub>6</sub> H <sub>4</sub>	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14d</b>	<b>77%</b>
10	<b>13e</b>	3,5-(CF <sub>3</sub> ) <sub>2</sub> -C <sub>6</sub> H <sub>3</sub>	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14e</b>	60%
11	<b>13f</b>	1-naphthyl	( <i>S</i> )- <b>2a</b>	<i>t</i> -Bu	<b>14f</b>	47%

12	<b>13b</b>	2-naphthyl	( <i>S</i> )- <b>11</b> <sup>c</sup>	<i>t</i> -Bu	<b>14b</b>	69%
13	<b>13b</b>	2-naphthyl	( <i>S</i> )- <b>12</b> <sup>c</sup>	CH <sub>2</sub> (4-BnO-C <sub>6</sub> H <sub>4</sub> )	<b>14b</b>	60%

<sup>a</sup>Reactions performed on 0.05 mmol. <sup>b</sup>Reaction time 20h. <sup>c</sup>reaction time = 3h

For the hydrogenation of **13a** (Ar = Ph) in the presence of acid (*S*)-**2a**, three different Hantsch esters have been tested as the H-transfer reagents (R = Et, CH<sub>2</sub>(4-BnO-C<sub>6</sub>H<sub>3</sub>), *t*-Bu; entries 1, 3, 5 in **Table 1**). In all these experiments, the starting material has been converted quantitatively into the expected tetrahydroquinoline **14a** after 2h at room temperature. Enantiomeric excess of 77-78% were obtained by using the ethyl or the 4-benzyloxybenzyl esters, while the *t*-butyl ester afforded the desired product in slightly improved, 82% ee. For comparison purposes, entries 2,4,6 in **Table 1** recall the results obtained previously in the same reactions by using the ferrocene-derivative **1c** as the acid catalyst.

The *t*-Bu Hantsch ester and catalyst (*S*)-**2a** were used then in the reduction of a larger set of substrates. Good enantiomeric excesses were obtained for Ar = 2-naphthyl, *p*-biphenyl and 4-MeO-phenyl, with ee between 77% and 90% (entries 7-9). The mixed amides (*S*)-**11** and (*S*)-**12** have been tested also in the reduction of the  $\alpha$ -2-naphthylquinoline **13b** (entries 12, 13). In spite of their lower pK<sub>a</sub> values, these catalysts gave lower catalytic activity and lower enantiomeric excesses than the corresponding phosphoric acids in entry 7.

## Conclusions

This work complements our previous studies by highlighting the biphenylene-based paracyclophanes **2** as a new kind of planar chiral phosphoric acids giving promising levels of enantioselectivity in organocatalytic H-transfer hydrogenations. With respect to the analogous ferrocene framework of **1**, the biphenylene unit rigidifies the cyclophane conformation and gives advantages in terms of both synthetic availability and thermal stability of the corresponding phosphoric acids. This work corroborates the good potential of these paracyclophane scaffolds in enantioselective acid catalysis. Further studies on their catalytic applications are currently in progress.

## Experimental Section

Anhydrous solvents were obtained by filtration through drying columns (THF, toluene, CH<sub>2</sub>Cl<sub>2</sub>). All reagents and solvents were of commercial quality and were used without further purification. Analytical thin-layer chromatography (TLC) was performed on plates precoated with silica gel layers. Flash column chromatography was performed using 40-63 mesh silica. NMR spectra (<sup>1</sup>H, <sup>13</sup>C) were recorded at 500 or 300 MHz spectrometers. High resolution mass spectrometry (HRMS) was performed using electrospray ionization (ESI) and time-of-flight (TOF) analyzer, in positive-ion or negative-ion detection mode.

*1,8-Bis(4-(methoxymethoxy)phenyl)biphenylene (5)*. 1,8-Dibromobiphenylene **3** (22.55 g, 72.7 mmol), [4-(methoxymethoxy)phenyl]boronic acid **4** (29.12 g, 160.0 mmol) and tri-*o*-furyl phosphine (2.53 g, 10.9 mmol, 15 mol%) were dissolved in 1,4-dioxane (700 mL) then a 1M solution of Cs<sub>2</sub>CO<sub>3</sub> (290 mL) was added. The flask was evacuated and refilled with argon and Pd<sub>2</sub>dba<sub>3</sub> (1.67 g, 1.82 mmol, 2.5 mol%) was added under a flow of argon. The mixture was refluxed for 2 h, then allowed to cool to room temperature, quenched with a saturated solution of NH<sub>4</sub>Cl. The organic phase was extracted twice with EtOAc and dried over MgSO<sub>4</sub> and the solvents were removed under reduced pressure. The black residue was purified by column chromatography (heptane/EtOAc = 95:5 to 9:1). The product **5** (R<sub>f</sub> = 0,15 in heptane/EtOAc 7:3) was obtained as a pale yellow solid (17.0 g, 55%). <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) :  $\delta$  6.86-6.81 (m, 8H), 6.67-6.64 (m, 2H), 6.60 (d, *J* = 8.3 Hz, 4H), 5.11 (s, 4H), 3.48 (s, 6H); <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>):  $\delta$  156.8 (C), 151.4 (C), 148.1 (C), 132.5 (C), 131.9 (C), 129.3 (CH), 128.8 (CH), 128.7 (CH), 115.5 (CH), 115.4 (CH), 94.7 (CH<sub>2</sub>), 56.1 (CH<sub>3</sub>) ppm. HRMS (ESI) calcd. for C<sub>28</sub>H<sub>24</sub>NaO<sub>4</sub> [M+Na]<sup>+</sup> : 447.1572, found : 447.1563.

*1,8-Bis-[4-(methoxymethoxy)-3-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)phenyl]-biphenylene* (**6**). In a flame dried round-bottomed flask is dissolved 1,1'-bis-[4-(methoxymethoxy)phenyl]-biphenylene **5** (8.4 g, 19.8 mmol) in anhydrous THF (200 mL). After cooling to -78 °C, *n*-BuLi (2.5 M solution in hexane, 47.5 mL, 118 mmol, 6.0 eq.) was added dropwise to the solution. After 5 minutes at -78 °C the mixture is allowed to warm to room temperature for 4 h (after few minutes at r.t. the pale yellow solution turns into a dark brown solution and after ca 20 min the dilithiated species appears as a white precipitate). Then the suspension is cooled to -78 °C and 2-isopropoxy-4,4,5,5-tetramethyl-1,3,2-dioxaborolane (32.0 ml, 158.3 mmol) is added dropwise and the suspension is stirred overnight at room temperature. To the resulting suspension was added a saturated aqueous solution of NH<sub>4</sub>Cl and EtOAc. The phases were separated and the aqueous phase is extracted once with EtOAc. The organic phases are dried over MgSO<sub>4</sub> and the solvents were removed under reduced pressure. The procedure was repeated twice. The mixture was separated by flash chromatography on silica gel (eluent: Toluene/EtOAc = 9:1) to afford **6** (*R<sub>f</sub>* = 0.42 in heptane/EtOAc 7:3) as a pale yellow solid (19.0 g, 71% for two batches). mp 176 °C; <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) : δ 7.46 (d, *J* = 2.4 Hz, 2H), 6.84-6.81 (m, 4H), 6.75 (dd, *J* = 8.5, 2.4 Hz, 2H), 6.63 (m, 2H), 6.42 (d, *J* = 8.5 Hz, 2H), 5.12 (s, 4H), 3.51 (s, 6H), 1.30 (s, 24H); <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>) : δ 161.4 (C), 151.3 (C), 148.1 (C), 135.5(CH), 132.4 (C), 132.1 (CH), 131.3 (C), 129.4 (CH), 128.4 (CH), 115.2 (CH), 114.2 (CH), 95.3 (CH<sub>2</sub>), 83.3 (C), 56.1 (CH<sub>3</sub>), 25.0 (CH<sub>3</sub>), 24.9 (CH<sub>3</sub>), 24.6 (CH<sub>3</sub>) ppm; HRMS (ESI) calcd. for C<sub>40</sub>H<sub>46</sub>O<sub>8</sub>B<sub>2</sub>Na [M+Na]<sup>+</sup> : 699.3276, found : 699.3293.

*1,8-Bis-[4-(hydroxy)-3-[(3',5'-diphenyl)phenyl]phenyl]biphenylene* (**8**). Bis-boronate **6** (17.1 g, 25.2 mmol), 1-bromo-3,5-diphenylbenzene **7** (16.0 g, 51.7 mmol) and Ba(OH)<sub>2</sub>·8H<sub>2</sub>O (31.8 g, 101 mmol), DME (200 mL) and water (55 mL) were introduced in a round bottomed flask. The flask was evacuated and refilled with argon (3 cycles), Pd(PPh<sub>3</sub>)<sub>4</sub> (1.46 g, 1.26 mmol, 5 mol%) was added under a flow of argon. The mixture was refluxed for 18 h. The resulting white suspension was allowed to cool to room temperature, diluted with CH<sub>2</sub>Cl<sub>2</sub> and washed with water. The organic phase was dried over MgSO<sub>4</sub> and concentrated *in vacuo*. The residue was taken up in 1,4-dioxane (500 mL) and MeOH (500 mL). Then, HCl 37% was added dropwise until pH = 1 (15 mL) and the mixture was heated at 60 °C for 4 h. A saturated aqueous solution of NaHCO<sub>3</sub> was added, the layers were separated and the aqueous layer was extracted with EtOAc. After drying over MgSO<sub>4</sub> the solvents were removed under reduced pressure. The mixture was separated by flash chromatography on silica gel (Toluene/EtOAc = 97:3) to afford the desired product **8** (*R<sub>f</sub>* = 0,35 in heptane/EtOAc 8:2) as a pale yellow solid (13.7 g, 69%). mp 224 °C; <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>) : δ 7.59 (bs, 2H), 7.46-7.45 (m, 8H), 7.36-7.32 (m, 16H), 6.98 (bs, 2H), 6.93 (d, *J* = 8.1 Hz, 2H), 6.87-6.82 (m, 4H), 6.68 (d, *J* = 6.5 Hz, 2H), 6.60 (d, *J* = 8.0 Hz, 2H), 5.09 (bs, 2H); <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>) : δ 152.2 (C), 151.4 (C), 147.8 (C), 142.4 (C), 140.5 (C), 137.8 (C), 132.3 (C), 131.1 (C), 130.0 (CH), 129.6 (CH), 128.9 (CH), 128.8 (CH), 128.6 (CH), 127.7 (CH), 127.3 (CH), 126.6 (CH), 125.2 (CH), 115.5 (CH) ppm; HRMS (ESI) calcd. for C<sub>60</sub>H<sub>41</sub>O<sub>2</sub> [M+H]<sup>+</sup> : 793.3107, found: 793.3104.

**Synthesis of the epimeric phosphates 10.** To a solution of bis-phenol **8** (6.5 g, 8.2 mmol) and solid 1*H*-tetrazole (2.41 g, 34.4 mmol) in anhydrous DCM (410 mL) was added dropwise *via* cannula a solution of (*S*)-2-cyanoethyl-1-phenyl-*N,N,N',N'*-tetra-isopropylphosphorodiamidite (*S*)-**9** (3.71 g, 9.8 mmol) in anhydrous DCM (10 mL). The mixture was heated at 40 °C for 3 h, then quenched by addition of saturated aqueous NaHCO<sub>3</sub> and extracted with DCM. The organic phase was dried over MgSO<sub>4</sub> and the solvents were removed under reduced pressure. The crude macrocyclic phosphites were obtained as a 7:3 mixture of diastereoisomers and engaged in the next step without purification. <sup>31</sup>P NMR (200 Hz, CDCl<sub>3</sub>): δ 130.0 (minor), 129.4 ppm (major). HRMS (ESI) calcd. for C<sub>69</sub>H<sub>46</sub>NO<sub>3</sub>NaP [M+Na]<sup>+</sup>: 990.3113, found : 990.3148.

The mixture of crude phosphites (8.2 mmol, 1 eq.) was dissolved in DCM (20 mL). TBHP (5.5M in decane, 4.5 mL, 3 eq.) was added to the solution at 0 °C. After 15 min, the reaction mixture was allowed to warm to room temperature and stirred for 45 minutes. The reaction was treated with saturated aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, the layers were separated, dried over MgSO<sub>4</sub> and concentrated *in vacuo*.

The crude phosphates were purified by chromatography on silica gel (eluent: toluene/EtOAc = 7:3) to afford the mixture of phosphates **10a** and **10b** (6.9 g, 86% over 2 steps) as a pale yellow solid. Phosphates (*S,S*)-**10a** and (*S,R*)-**10b** were separated by semi-preparative HPLC on a SiO<sub>2</sub> column (250 x 10 mm, 5 mic) with toluene/heptane:THF (95:4:1) as the eluent; flow rate, 100 mL/min. Retention times: 10.0 min for (*S,S*)-**10a**, 13.0 min for (*S,R*)-**10b**.

(*S,S*)-**10a** (major diastereoisomer), white solid (3.6 g, 45% yield). mp 251 °C. <sup>31</sup>P NMR (200 MHz, CDCl<sub>3</sub>) : δ -17.2 ppm; <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) : δ 7.91 (bs, 1H), 7.78 (bs, 3H), 7.70 (d, *J* = 7.4 Hz, 4H), 7.65-7.62 (m, 6H), 7.45-7.34 (m, 12H), 7.29 (d, *J* = 7.5 Hz, 2H), 7.14 (t, *J* = 7.5 Hz, 1H), 7.03 (t, *J* = 7.6 Hz, 2H), 6.94-6.86 (m, 3H), 6.81-6.74 (m, 4H), 6.68-6.63 (m, 2H), 6.21 (dd, *J* = 8.5, 2.2 Hz, 1H), 6.14 (d, *J* = 8.5 Hz, 1H), 6.08 (dd, *J* = 8.5, 2.2 Hz, 1H), 5.17 (ddd, *J*<sub>H-P</sub> = 7.7, *J* = 6.5, 5.5 Hz, 1H), 2.44 (dd, *J* = 16.5, 6.5 Hz, 1H), 2.34 (dd, *J* = 16.5, 5.5 Hz, 1H); <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>) : δ 151.6 (C), 151.5 (C), 149.6 (C), 149.2 (C), 149.1 (d, *J*<sub>C-P</sub> = 5.9 Hz, C-O), 146.8 (d, *J*<sub>C-P</sub> = 8.3 Hz, C-O), 142.5 (C), 142.0 (C), 141.0 (C), 140.5 (C), 138.1 (C), 137.9 (C), 136.9 (C), 136.5 (C), 136.1 (d, *J*<sub>C-P</sub> = 3.9 Hz, C<sub>Ph</sub>), 132.9 (C), 131.8 (C), 131.5 (C), 131.1 (C), 131.0 (C), 129.5 (CH), 129.2 (CH), 129.0 (CH), 128.9 (CH), 128.2 (CH), 128.1 (CH), 128.0 (CH), 127.7 (CH), 127.6 (CH), 127.5 (CH), 127.4 (CH), 127.2 (CH), 127.0 (CH), 125.9 (CH), 125.8 (CH), 125.6 (CH), 121.9 (CH), 118.9 (CH), 116.8 (CH), 114.9 (CN), 76.7 (OCH), 26.9 (d, *J*<sub>C-P</sub> = 7.7 Hz, CH<sub>2</sub>) ppm; HRMS (ESI) calcd. for C<sub>69</sub>H<sub>47</sub>NO<sub>4</sub>P [M+H]<sup>+</sup> : 984.3243, found : 984.3210. [α]<sub>D</sub><sup>20</sup> : +68 (*c* = 1.0, CHCl<sub>3</sub>).

(*S,R*)-**10b** (minor diastereoisomer), white solid (1.8 g, 22% yield). mp 253 °C. <sup>31</sup>P NMR (200 MHz, CDCl<sub>3</sub>) : δ -17.7 ppm. <sup>1</sup>H NMR (300 MHz, CDCl<sub>3</sub>) : 7.90 (bs, 1H), 7.81 (d, *J* = 1.4 Hz, 2H), 7.77 (bs, 1H), 7.71 (d, *J* = 7.6 Hz, 4H), 7.65 (d, *J* = 7.6 Hz, 6H), 7.48-7.30 (m, 14H), 7.16-7.10 (m, 1H), 7.02-6.96 (m, 4H), 6.81-6.77 (m, 4H), 6.68-6.64 (m, 2H), 6.60 (d, *J* = 8.4 Hz, 1H), 6.15-6.10 (m, 2H), 6.02 (dd, *J* = 8.4, 2.1 Hz, 1H), 5.40 (ddd, *J*<sub>H-P</sub> = 8.2, *J* = 8.2, 4.6 Hz, 1H), 2.57 (dd, *J* = 16.6, 4.6 Hz, 1H), 2.37 (dd, *J* = 16.6, 8.2 Hz, 1H). <sup>13</sup>C NMR (75 MHz, CDCl<sub>3</sub>) : δ 151.5 (C), 149.4 (C), 148.2 (d, *J*<sub>C-P</sub> = 8.3 Hz, C-O), 147.8 (d, *J*<sub>C-P</sub> = 8.9 Hz, C-O), 142.6 (C), 142.2 (C), 141.1 (C), 140.5 (C), 138.0 (C), 137.6 (C), 136.9 (C), 136.7 (C), 135.6 (d, *J*<sub>C-P</sub> = 6.0 Hz, C<sub>Ph</sub>), 132.2 (C), 132.1 (C), 131.6 (C), 131.1 (C), 131.0 (C), 129.7 (CH), 129.2 (CH), 129.0 (CH), 128.9 (CH), 128.4 (CH), 128.2 (CH), 128.0 (CH), 127.6 (CH), 127.4 (CH), 127.1 (CH), 126.3 (CH), 126.0 (CH), 125.7 (CH), 120.7 (CH), 120.0 (CH), 116.8 (CH), 115.0 (CN), 76.1 (d, *J*<sub>C-P</sub> = 5.6 Hz, OCH), 26.0 (d, *J*<sub>C-P</sub> = 4.3 Hz, CH<sub>2</sub>) ppm. HRMS (ESI) calcd. for C<sub>69</sub>H<sub>47</sub>NO<sub>4</sub>P: 984.3243, found: 984.3212. [α]<sub>D</sub><sup>20</sup> : -123 (*c* = 1.0, CHCl<sub>3</sub>).

*Phosphoric acid (S)*-**2a**. DBU (920 μL, 6.1 mmol) was added dropwise to a solution of phosphate (*S,S*)-**10a** (3.0 g, 3.0 mmol) in DCM (15 mL). After 30 min at room temperature, the crude DBU phosphate salt was purified by column chromatography (eluent: MeOH/DCM = 0:100 to 5:95) to remove the side-product cinnamitrile. The fractions were collected, concentrated *in vacuo* and diluted in DCM (15 mL). The solution was treated with HCl 6N (3\*15 mL). The layers were separated and the organic layer was concentrated *in vacuo* to afford the chiral phosphoric acid (*S*)-**2a** (2.40 g, 92% yield) as a white solid. mp 306 °C; <sup>31</sup>P (125 MHz, CDCl<sub>3</sub>) : δ -11.4; <sup>1</sup>H NMR (500 MHz, CDCl<sub>3</sub>) : δ 7.66 (bs, 2H), 7.65 (bs, 4H), 7.61 (d, *J* = 7.7 Hz, 8H), 7.44 (d, *J* = 1.7 Hz, 2H), 7.34 (t, *J* = 7.7 Hz, 8H), 7.24 (t, *J* = 7.7 Hz, 4H), 6.93-6.89 (m, 4H), 6.77 (dd, *J* = 5.1, 2.4 Hz, 2H), 6.39 (d, *J* = 8.5 Hz, 2H), 6.02 (1H), 6.00 (dd, *J* = 8.5, 1.7 Hz, 2H); <sup>13</sup>C NMR (125 MHz, CDCl<sub>3</sub>) : δ 151.9 (C), 149.7 (C), 148.2 (d, *J*<sub>C-P</sub> = 7.6 Hz, C-O), 142.5 (C), 141.3 (C), 137.8 (C), 136.5 (C), 132.0 (C), 131.9 (d, *J*<sub>C-P</sub> = 5.9 Hz, C<sub>Ph</sub>), 129.0 (CH), 128.2 (CH), 128.1 (CH), 127.9 (CH), 127.6 (CH), 127.1 (CH), 126.0 (CH), 120.8 (CH), 116.8 (CH) ppm; HRMS (ESI) calcd. for C<sub>60</sub>H<sub>38</sub>O<sub>4</sub>P: 853.2513, found: 853.2493; [α]<sub>D</sub><sup>20</sup> : +90 (*c* = 0.5, CHCl<sub>3</sub>).

**Representative procedure for the synthesis of the mixed imides 11 and 12.** (*a*) *Conversion of (S)*-**2a** into the corresponding chlorophosphate. A solution of triphenylphosphine (105 mg, 0.40 mmol) in anhydrous dichloromethane (1.5 mL) was added dropwise to a solution of phosphoric acid (*S*)-**2a** (171 mg, 0.20 mmol) and 2,2,2-trichloroacetonitrile (40 μL, 0.40 mmol) in anhydrous dichloromethane (1.5

mL). The mixture was stirred for 1 h at r.t., then the solvent was evaporated under reduced pressure and the residue was purified on silica gel (heptane/ethyl acetate = 9:1 to 8:2) to afford the desired chlorophosphate as a white solid (150 mg, 86 %).  $R_f = 0,5$  (heptanes/ethyl acetate 8:2); mp 282 °C;  $^{31}\text{P}$  NMR (200 MHz,  $\text{CDCl}_3$ ):  $\delta$  -14.0 ppm;  $^1\text{H}$  NMR (500.1 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.87 (t,  $J = 1.6$  Hz, 1H), 7.85 (t,  $J = 1.6$  Hz, 1H), 7.81 (d,  $J = 1.6$  Hz, 2H), 7.79 (d,  $J = 1.6$  Hz, 2H), 7.75 (d,  $J = 8.0$  Hz, 4H), 7.73 (d,  $J = 8.0$  Hz, 4H), 7.54 (bs, 2H), 7.52-7.48 (m, 8H), 7.41 (t,  $J = 8.0$  Hz, 2H), 7.39 (t,  $J = 8.0$  Hz, 2H), 6.92-6.86 (m, 4H), 6.77 (d,  $J = 7.0$  Hz, 2H), 6.63 (d,  $J = 8.5$  Hz, 1H), 6.61 (d,  $J = 8.5$  Hz, 1H), 6.29 (dd,  $J = 8.5, 2.1$  Hz, 1H), 6.26 (dd,  $J = 8.5, 2.1$  Hz, 1H).  $^{13}\text{C}$  NMR (75.5 MHz,  $\text{CDCl}_3$ ):  $\delta$  151.6 (C), 149.5 (C), 149.4 (C), 148.6 (d,  $J_{\text{C-P}} = 12.2$  Hz, C-O), 148.2 (d,  $J_{\text{C-P}} = 12.2$  Hz, C-O), 142.7 (C), 142.6 (C), 141.1 (C), 140.9 (C), 137.6 (C), 137.5 (C), 137.0 (C), 132.2 (d,  $J_{\text{C-P}} = 6.4$  Hz,  $\text{C}_{\text{Ph}}$ ), 131.5 (C), 131.4 (C), 129.1 (CH), 129.0 (CH), 128.8 (CH), 128.5 (CH), 128.3 (CH), 128.1 (CH), 127.9 (CH), 127.7 (CH), 127.5 (CH), 127.3 (CH), 127.2 (CH), 126.5 (CH), 126.3 (CH), 121.2 (CH), 119.2 (CH), 117.0 (CH) ppm; HRMS (ESI) calcd.for  $\text{C}_{60}\text{H}_{39}\text{O}_3\text{ClP}$   $[\text{M}+\text{H}]^+$ : 873.2325, found: 873.2340.  $[\alpha]_{\text{D}}^{20}$ : +89 ( $c = 1$ ,  $\text{CHCl}_3$ ).

(b) *Synthesis of the mixed imide 11*. To a solution of potassium carbonate (33 mg, 0.24 mmol) and trifluoromethanesulfonamide (36 mg, 0.24 mmol) in anhydrous acetonitrile (2.5 mL) at r.t. was added a solution of chlorophosphate (105 mg, 0.12 mmol) in anhydrous dichloromethane/acetonitrile (0.5 mL / 2.5 mL). The mixture was heated at 60°C for 16h. The solvent was evaporated under reduced pressure and the residue was purified on silica gel (eluent: heptane/ethyl acetate gradient, from 8:2 to 5:5). The fractions were collected, concentrated *in vacuo* and diluted in DCM (5 mL). This solution was treated with HCl 6N (3\*5 mL). The layers were separated and the organic layer was concentrated *in vacuo* to afford **11** as a white solid (105 mg, 89% yield).  $R_f = 0,6$  (heptane/ethyl acetate 1:1); mp 235 °C;  $^{31}\text{P}$  NMR (202.5 Hz,  $\text{CDCl}_3$ ):  $\delta$  -18.8 ppm;  $^{19}\text{F}$  NMR (282.4 Hz,  $\text{CDCl}_3$ ):  $\delta$  -77.6 ppm;  $^1\text{H}$  NMR (300.2 MHz,  $\text{CDCl}_3$ ):  $\delta$  7.91 (bs, 1H), 7.77 (bs, 4H), 7.73 (d,  $J = 7.7$  Hz, 4H), 7.65 (bs, 1H), 7.62 (d,  $J = 7.7$  Hz, 4H), 7.55 (bs, 1H), 7.52 (bs, 1H), 7.45-7.31 (m, 12H), 7.04-6.83 (m, 6H), 6.74 (d,  $J = 8.5$  Hz, 1H), 6.51 (d,  $J = 8.5$  Hz, 1H), 6.29 (dd,  $J = 8.5, 2.1$  Hz, 1H), 6.25 (dd,  $J = 8.5, 2.1$  Hz, 1H);  $^{13}\text{C}$  NMR (75.5 MHz,  $\text{CDCl}_3$ ):  $\delta$  151.7 (C), 149.8 (C), 149.4 (C), 149.3 (d,  $J_{\text{C-P}} = 8.8$  Hz, C-O), 144.6 (d,  $J_{\text{C-P}} = 8.8$  Hz, C-O), 142.5 (C), 142.4 (C), 141.3 (C), 141.0 (C), 137.9 (C), 137.5 (C), 137.1 (C), 136.8 (C), 132.9 (C), 132.0 (C), 131.8 (d,  $J_{\text{C-P}} = 8.2$  Hz,  $\text{C}_{\text{Ph}}$ ), 131.6 (C), 129.2 (CH), 128.9 (CH), 128.7 (CH), 128.4 (CH), 128.2 (CH), 128.1 (CH), 127.7 (CH), 127.5 (CH), 127.0 (CH), 126.3 (CH), 125.9 (CH), 122.2 (CH), 119.1 (q,  $J_{\text{C-F}} = 321.2$  Hz,  $\text{CF}_3$ ), 118.5 (CH), 117.0 (CH); HRMS (ESI) calcd.for  $\text{C}_{61}\text{H}_{40}\text{NO}_3\text{F}_3\text{PS}$   $[\text{M}+\text{H}]^+$ : 986.2317, found: 986.2324;  $[\alpha]_{\text{D}}^{20}$ : +99 ( $c = 1$ ,  $\text{CHCl}_3$ ).

Mixed imide **12**:  $^{31}\text{P}$  NMR (202.5 MHz  $\text{CDCl}_3$ ):  $\delta$  -10.5 ppm;  $^{19}\text{F}$  NMR (282.4 MHz,  $\text{CDCl}_3$ ): -78.7 ppm;  $^1\text{H}$  NMR (500.1 MHz,  $\text{DMSO-}d_6$ - 90 °C): 7.91-7.87 (m, 10H), 7.81-7.80 (m, 2H), 7.71-7.64 (m, 2H), 7.53-7.48 (m, 8H), 7.39 (t,  $J = 7.1$  Hz, 4H), 7.25-7.22 (m, 3H), 6.52 (d,  $J = 8.2$  Hz, 1H), 6.43 (d,  $J = 8.2$  Hz, 1H), 6.36 (d,  $J = 8.2$  Hz, 1H), 4.90 (s, 1H), 4.87 (s, 1H), 4.68 (s, 1H), 4.62 (s, 1H), 4.40 (s, 2H), 4.37 (s, 2H);  $^{13}\text{C}$  NMR (125.7MHz,  $\text{DMSO-}d_6$ - 90 °C):  $\delta$  148.0 (d,  $J_{\text{C-P}} = 7.2$  Hz, C), 146.4 (d,  $J_{\text{C-P}} = 9.2$  Hz, C), 140.0 (C), 139.7 (C), 138.2 (C), 137.5 (C), 131.8 (C), 131.4 (C), 129.3 (d,  $J_{\text{C-P}} = 9.0$  Hz, C), 128.6 (C), 127.6 (CH), 127.5 (2CH), 127.4 (CH), 127.1 (C), 126.2 (CH), 126.1 (CH), 126.0 (CH), 125.8 (CH), 125.7 (CH), 125.6 (2CH), 124.7 (CH), 124.2 (CH), 123.1 (CH), 122.8 (CH), 122.5 (CH), 118.7 (CH), 85.4 (C), 85.2 (C), 67.8 (2CH), 67.1 (2CH), 65.4 (CH), 65.0 (CH), 64.4 (2CH); HRMS (ESI) calcd.for  $\text{C}_{59}\text{H}_{40}\text{NF}_3\text{FeO}_5\text{PS}$   $[\text{M}-\text{H}]^-$ : 1018.1666, found: 1018.1647;  $[\alpha]_{\text{D}}^{20} = +747$  ( $c = 1$ ,  $\text{CHCl}_3$ ).

*Representative procedure for the catalytic reductions of  $\alpha$ -arylquinolines*. A solution of 2-phenylquinoline **13a** (11 mg, 0.05 mmol), Hantzsch dihydropyridine ( $R = \text{Et}$ , 0.12 mmol) and the acid catalyst (*S*)-**2a** (0.005 mmol) in toluene (1 ml) was stirred at r.t. for 2h. The solvent was removed under reduced pressure and the residue was purified on silica gel with toluene as the eluent. 2-Phenyl-1,2,3,4-tetrahydroquinoline **14a** was obtained as a colorless oil. The enantiomeric excess was determined by HPLC using a CHIRALPAK® IB column. Eluent: *i*-PrOH/*n*-heptane 5:95, at a flow

rate of 1 mL/min [detection at 275 nm]. Retention times: 6.7 min for (*S*)-**14a** and 8.2 min for (*R*)-**14a**. [Litt:<sup>18</sup> (*S*)-**14a**:  $[\alpha]_{\text{D}} = -35.7$  ( $c = 0.8$ ,  $\text{CHCl}_3$ )].

The same procedure was applied to the reduction of quinolines **13b-f**. Spectral data for the corresponding tetrahydroquinolines **14b-f** and conditions for ee measurements by HPLC have been reported previously.<sup>7b</sup>

### Computational Method

Calculations have been carried out with the Gaussian09 package of programs.<sup>19</sup> All the structures were optimized at the DFT level by means of the M06 functional.<sup>20</sup> The 6-31G(d,p) basis set was applied for all atoms. This level of calculation has been previously shown to provide reliable structures for phosphoric acids derivatives.<sup>7b</sup> To get accurate geometries and energies, the SCF convergence criterion has been systematically tightened to  $10^{-8}$  a.u., and the force minimizations were carried out until the rms force becomes smaller than (at least)  $1 \times 10^{-5}$  a.u. The optimized geometries were characterized by harmonic analysis, and the nature of the stationary points was determined according to the number of negative eigenvalues of the Hessian matrix

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**Associated content.** Copies of the <sup>1</sup>H, <sup>13</sup>C and <sup>31</sup>P NMR spectra, X-ray crystal data for **10a**, details for pKa measurements. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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