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Improved single-chain-magnet behavior in a biradical-based nitronyl nitroxide-Cu-Dy chain

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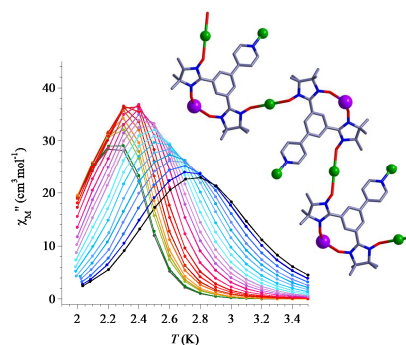
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Abstract :

A biradical with ferromagnetic intramolecular interaction was found to improve the SCM behavior of a nitronyl nitroxide-Cu-Dy chain, resulting in an energy barrier for magnetization reversal of 40 K.



One-dimensional (1D) heterometallic coordination polymers represent an intriguing family of molecular magnetic compounds. A seminal example was reported by Kahn et al.¹ and since then this design strategy has remained a flourishing research topic.² A current focus concerns 1D compounds exhibiting slow magnetic relaxation of molecular origin known as single-chain magnets (SCM),³ and numerous heterometallic SCMs involving various metal combinations (for instance 3d-3d',⁴ 3d-4d,⁵ 3d-5d,⁶ 3d-4f,⁷ and even heterotrimetallic⁸) have been reported.⁹ For SCMs, the correlation energy contributes to the

energy barrier for magnetization reversal,¹¹ therefore strong exchange interactions within the chains are desirable for increasing the blocking temperature. This latter represents a crucial synthetic challenge for rising the energy barrier of SCMs.¹² A promising strategy consists in using organic radicals as bridging ligands in order to shorten the distance between the magnetic centers and achieve largest possible magnetic communication along the 1D array.^{2c,13} This is obtained when an organic radical unit directly coordinates two consecutive metal centers. Several remarkable SCMs based on this approach have been reported, involving either 3d or 4f metal ions,^{14,15} and such metal-radical systems have been demonstrated to be photoswitchable and compatible with observation of magneto-optical properties.¹⁰ Related 2p-3d-4f SCMs (i.e. hetero-tri-spin compounds) were hardly investigated,^{16,17} however they provide new opportunities for fine-tuning the magnetic couplings between spins, the ligand field, and magnetic anisotropy of the metal centers.

Our contribution to this topic concerns nitronyl nitroxide radical-3d-4f derivatives,¹⁷ in particular ferromagnetic chains with this radical unit acting as a bridge between the metal ions.¹⁸ Slow magnetic relaxation have been observed when anisotropic Ln ions were involved but SCM characteristics remained modest. A likely cause is the antiferromagnetic next-neighbour radical-radical interaction via the Ln centers,^{18b,19} a contribution of similar amplitude but opposite to the ferromagnetic radical-Ln interactions resulting in a situation similar to spin frustration. We conjectured that the effect of this competitive contribution could be reduced by an additional ferromagnetic exchange pathway between radicals, a feature easily implemented in biradicals.²⁰ This is confirmed herein by a hetero-tri-spin SCM constructed using the nitronyl nitroxide (NIT hereafter) biradical, bisNITPhPy (i.e. [5-(4-pyridyl)-1,3-bis(1'-oxyl-3'-oxido-4',4',5',5'-tetramethyl-4,5-hydro-1*H*-imidazol-2-yl)]benzene), with *meta*-positioned radical moieties.

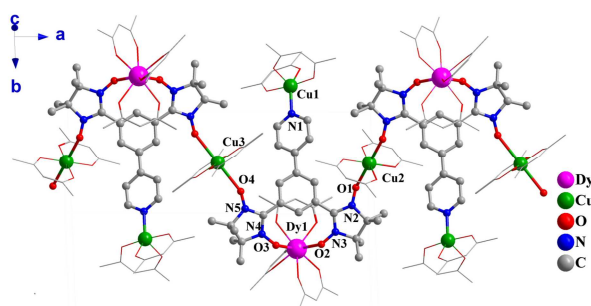


Figure 1. One-dimensional structure of complex 1. All hydrogen and fluorine atoms are omitted for clarity.

The 1D coordination polymer $[\text{DyCu}_2(\text{hfac})_7(\text{bisNITPhPy})]$ **1** was obtained as dark-green single crystals from the one-pot reaction between $\text{Dy}(\text{hfac})_3 \cdot 2\text{H}_2\text{O}$, $\text{Cu}(\text{hfac})_2$ and bisNITPhPy with a mole ratio of 1:2:1 in a *n*-hexane- CHCl_3 mixture (see ESI). Single crystal X-ray analysis

revealed that **1** has one-dimensional chain structure consisting in an alternation of Cu(II) and Dy(III) ions bridged by NIT units (Fig. 1). This arrangement is made up by one bisNITPhPy, one Dy(hfac)₃, and two Cu(hfac)₂ units. The organic ligand and the Dy form a chelate resulting from coordination of one aminoxyl (NO) moiety of each NIT to a central Dy, while the second NOs are coordinated to one Cu(II) each. The ligand is further linked to one independent and terminal Cu(hfac)₂ by the means of the pyridyl group. The Dy ion is ligated by six oxygen atoms from three bidentate hfac anions and two nitroxide groups forming a biaugmented trigonal prism (C_{2v})²¹ coordination polyhedron (Table S3). The Dy–O_{rad} distances are 2.368(5) Å and 2.385(5) Å, respectively, and the Dy–O_{hfac} bond lengths fall within the range of 2.298(5)–2.375(5) Å (Table S2 in ESI). The O(rad)–Dy–O(rad) angle is 82.4(7)° and the dihedral angles between the ON–C–NO fragment and the phenyl ring of biradical unit are 29.75(9)° and 29.02(1)°, respectively. The Cu units linked to the NIT moieties are crystallographically independent and appear as Cu2 and Cu3 in Fig. 1. They exhibit an elongated octahedral coordination environment formed by four O atoms from hfac ligands located in the equatorial positions, and two O atoms from two bisNITPhPy ligands in the apical positions. The Cu–O(hfac) bond lengths vary from 1.922(5) Å to 1.938(6) Å and the Cu–O_{rad} bond distances are 2.388(6) Å for Cu2 and 2.604(6) Å for Cu3. The pyridyl-coordinated Cu1 has a distorted pyramidal geometry consisting of four O atoms from two hfac[–] and one N atom from pyridine, in which the axial position is occupied by one O atom (Cu1–O5: 1.995(7) Å). In the chain, the Dy···Cu distances are 8.619(5) Å for Dy1···Cu2 and 8.486(1) Å for Dy1···Cu3. The crystal packing diagram is shown in Fig. S1. The shortest interchain Dy···Dy separation is 10.013(8) Å and the nearest Dy···Cu distance between the chains is 10.310(8) Å.

The temperature dependence of the magnetic susceptibility, χ_M (given for a bisNITPhPyDyCu₂ moiety) for **1** has been investigated between 2 and 300 K, and the results are plotted as $\chi_M T = f(T)$ in Fig. 2 (see also Fig. S4 in ESI). The phase purity of the sample was confirmed by PXRD (Fig. S3). The value found at 300 K for $\chi_M T$ is 16.15 cm³ K mol^{–1}, slightly above the Curie contribution of 15.67 cm³ K mol^{–1} expected for one Dy^{III} ion (⁶H_{15/2}, 14.17 cm³ mol^{–1} K), two Cu^{II} ($S = \frac{1}{2}$, $g = 2.0$) and two organic radicals ($S = 1/2$) in the absence of exchange interactions. The $\chi_M T$ product increases smoothly as T is reduced and below 20 K it rises rapidly and reaches 313 cm³ K mol^{–1} for 2 K. Such a behaviour is characteristic for long-range magnetic correlation without magnetic ordering, in agreement with the expected behaviour for a spin chain. The field dependence of the magnetization recorded at 2 K is characterized by a very sharp increase for small fields followed by a smoother but steady increase to reach 8.5 μB with 50 kOe. A similar behaviour was found for 3, 4, and 5 K (Fig. 2 and S5). The magnetisation reached for the largest investigated field suggests dominant ferromagnetic interactions within the chain, in agreement with the $\chi_M T = f(T)$ behaviour.

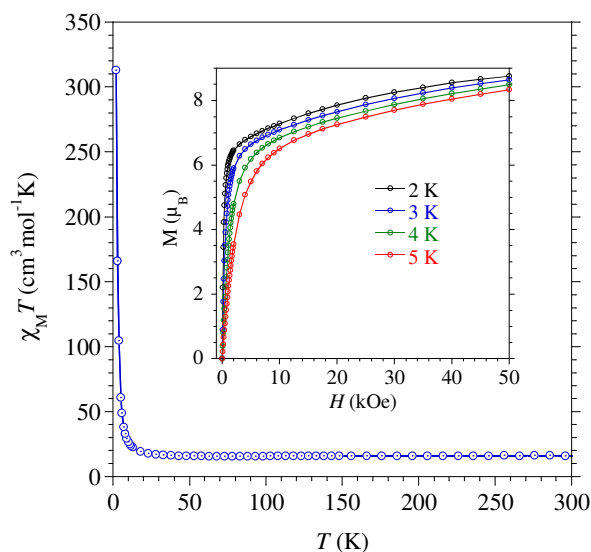


Figure 2. $\chi_M T$ versus T behavior for **1** (data from 300 to 100 K have been collected in an applied field of 1 kOe and those below with $H = 50$ Oe to avoid saturation (see Fig. S14); the inset shows the field dependence of the magnetization at 2, 3, 4, and 5 K.

This is consistent with the anticipated behaviour for the present system. Indeed, the axial coordination of aminoxyl groups to the Cu(II) ions are expected to lead to ferromagnetic Cu-radical interactions due to orthogonality of the magnetic orbital (dx^2-y^2) on copper and π^* SOMO on the radical,²² and the Dy-nitronyl nitroxide exchange interaction is also anticipated to be ferromagnetic;²³ moreover ferromagnetic intra-ligand radical-radical interaction is also expected.²⁴

For **1**, $\ln(\chi_M' T) = f(T^{-1})$ (χ_M' , the in-phase AC susceptibility in the absence of dc field) shows a linear region between 3.0 and 25 K (Fig. S4) that leads to a correlation energy of $\Delta_E/k_B = 7.7$ K (5.3 cm^{-1}) and $C_{eff} = 12.6 \text{ cm}^3 \text{ mol}^{-1} \text{ K}$, thus confirming the 1D Ising-like character of **1**.

AC susceptibilities, measured in the absence of a static field, showed frequency dependent χ_M' and χ_M'' behaviors (Fig. 3), a typical feature associated with slow magnetic relaxation. The analysis of the Cole–Cole plots, $\chi_M'' = f(\chi_M')$, gave a distribution width for the relaxation times of $\alpha = 0.11$ to 0.37 between 2 and 3 K. Such a distribution might be indicative for a population of chains of different lengths, however this was not detected when analysing the relaxation times (*vide infra*).

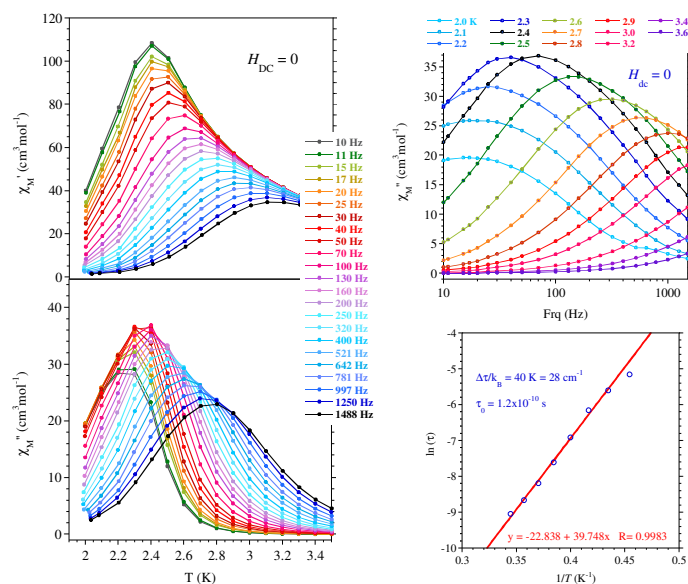


Figure 3. (left) χ_M' and $\chi_M'' = f(T)$ for frequencies between 10 and 1488 Hz obtained in zero field; (right) $\chi_M'' = f(\text{Freq})$ for temperatures between 2 and 3.6 K, and $\ln(\tau) = f(T^{-1})$ with best fit (red line) to Arrhenius equation.

The relaxation times (τ) between 2.2 and 3.0 K were deduced from $\chi_M'' = f(\nu)$ behaviors using the Debye model,²⁵ and subsequently fitted by the Arrhenius law [$\tau = \tau_0 \exp(\Delta_\tau/k_B T)$] to afford an energy barrier of 40.0 K with a pre-exponential factor (τ_0) of 1.2×10^{-10} s. In SCM, the variation of τ with T often exhibits a crossover between two thermally activated modes related to the occurrence of finite- and infinite-size regimes in the relaxation process.¹¹ This is not observed here, likely due to the small Ln-radical exchange interactions in **1** and the narrow temperature window for which τ could be assessed. The above energy barrier, Δ_τ , comprises two contributions which are the correlation energy Δ_ξ , and the blocking energy due to the magnetic anisotropy of the elementary magnetic unit, Δ_A .¹¹ For **1**, supposing finite-size chains, this latter contribution is in the order of $\Delta_A = 23 \text{ cm}^{-1}$ (32 K) from the difference between Δ_τ and Δ_ξ .

The magnetic characteristics exhibited by **1** are clearly improved compared to related 2p-3d-4f chain compounds¹⁸ and this supports our assumptions and strategy disclosed herein. The effective role of the biradical species promoting a ferromagnetic NIT-NIT interaction and thus limiting the next-neighbour contribution needs to be confirmed but the present results show that such biradicals provides promising opportunities to develop SCM with appealing properties.

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References

- 1 O. Kahn, Y. Pei, M. Verdaguer, J. P. Renard, J. Sletten, *J. Am. Chem. Soc.*, 1988, **110**, 782.
- 2 (a) M. Verdaguer, A. Bleuzen, V. Marvaud, J. Vaissermann, M. Seuleiman, C. Desplanches, A. Scullier, C. Train, R. Garde, G. Gelly, C. Lomenech, I. Rosenman, P. Veillet, C. Cartier and F. Villain, *Coord. Chem. Rev.*, 1999, **190-192**, 1023; (b) J. S. Miller, *Chem. Soc. Rev.*, 2011, **40**, 3266; (c) L. Bogani, A. Vindigni, R. Sessoli, D. Gatteschi, *J. Mater. Chem.*, 2008, **18**, 4750; (d) H. L. Sun, Z. M. Wang, S. Gao, *Coord. Chem. Rev.*, 2010, **254**, 1081; (e) S. Turner, O. Kahn, L. Rabardel, *J. Am. Chem. Soc.*, 1996, **118**, 6428; (f) K. S. Pedersen, J. Bendix, R. Clérac, *Chem. Commun.*, 2014, **50**, 4396. (g) M. Andruh, *Chem. Commun.*, 2018, **54**, 3559.
- 3 A. Caneschi, D. Gatteschi, N. Lalioti, C. Sangregorio, R. Sessoli, G. Venturi, A. Vindigni, A. Rettori, M. G. Pini, M. A. Novak, *Angew. Chem. Int. Ed.*, 2001, **40**, 1760.
- 4 (a) R. Clérac, H. Miyasaka, M. Yamashita and C. Coulon, *J. Am. Chem. Soc.*, 2002, **124**, 12837; (b) S. Wang, J.-L. Zuo, S. Gao, Y. Song, H.-C. Zhou, Y. Z. Zhang, X. Z. You, *J. Am. Chem. Soc.*, 2004, **126**, 8900; (c) E. Coronado, J. R. Galán-Mascarós and C. Martí-Gastaldo, *J. Am. Chem. Soc.*, 2008, **130**, 14987; (d) H. R. Wen, Y. Z. Tang, C. M. Liu, J. L. Chen, C. L. Yu, *Inorg. Chem.*, 2009, **48**, 10177; (e) E. Pardo, C. Train, R. Lescouezec, Y. Journaux, J. Pasán, C. R. Pérez, F. S. Delgado, R. R. Garcia, F. Lloret, C. Paulsen, *Chem. Commun.*, 2010, **46**, 2322; (f) Y. Q. Wang, Q. Yue, E. Q. Gao, *Chem. Eur. J.*, 2017, **23**, 896; (g) X. B. Li, Y. Ma, E. Q. Gao, *Inorg. Chem.*, 2018, **57**, 7446. (h) C. Pichon, N. Suaud, C. Duhayon, N. Guihéry, J.-P. Sutter, *J. Am. Chem. Soc.* 2018, **140**, 7698.
- 5 (a) J. F. Guo, X. T. Wang, B. W. Wang, G. C. Xu, S. Gao, L. Szeto, W. T. Wong, W. Y. Wong, T. C. Lau, *Chem. Eur. J.*, 2010, **16**, 3524; (b) D. P. Zhang, L. F. Zhang, Y. T. Chen, H. L. Wang, Z. H. Ni, W. Wernsdorfer, J. Z. Jiang, *Chem. Commun.*, 2010, **46**, 3550. (c) T. S. Venkatakrishnan, S. Sahoo, N. Bréfuel, C. Duhayon, C. Paulsen, A. L. Barra, S. Ramasesha, J. P. Sutter, *J. Am. Chem. Soc.*, 2010, **132**, 6047.
- 6 (a) S. W. Choi, H. Y. Kwak, J. H. Yoon, H. C. Kim, E. K. Koh, C. S. Hong, *Inorg. Chem.*, 2008, **47**, 10214; (b) X. W. Feng, T. D. Harris, J. R. Long, *Chem. Sci.*, 2011, **2**, 1688; (c) X. W. Feng, J. J. Liu, T. D. Harris, S. Hill, J. R. Long, *J. Am. Chem. Soc.*, 2012, **134**, 7521; (d) Y. Q. Zhang, C. L. Luo, X. B. Wu, B. W. Wang, S. Gao, *Inorg. Chem.*, 2014, **53**, 3503; (e) Y. Z. Zhang, B. S. Dolinar, S. H. Liu, A. J. Brown, X. Zhang, Z. X. Wanga, K. R. Dunbar, *Chem. Sci.*, 2018, **9**, 119.
- 7 (a) A. Figuerola, J. Ribas, D. Casanova, M. Maestro, S. Alvarez, C. Diaz, *Inorg. Chem.*, 2005, **44**, 6949; (c) M. X. Yao, Q. Zheng, K. Qian, Y. Song, S. Gao, J. L. Zuo, *Chem. Eur. J.*, 2013, **19**, 294; (d) A. Dey, S. Das, S. Kundu, A. Mondal, M. Rouzières, C. Mathonière, R. Clérac, R. S. Narayanan, V. Chandrasekhar, *Inorg. Chem.*, 2017, **56**, 14612.
- 8 D. Visinescu, A. M. Madalan, M. Andruh, C. Duhayon, J.-P. Sutter, L. Ungur, W. Van den Heuvel, L. F. Chibotaru, *Chem. Eur. J.* 2009, **15**, 11808.
- 9 (a) S. Dhers, H. L. C. Feltham, S. Brooker, *Coord. Chem. Rev.*, 2015, **296**, 24; (b) R. Lescouezec, L. M. Toma, J. Vaissermann, M. Verdaguer, F. S. Delgado, C. Ruiz-Pérez, F. Lloret, M. Julve, *Coord. Chem. Rev.* 2005, **249**, 2691.
- 10 (a) E. Heintze, F. El Hallak, C. Clauß, A. Rettori, M. G. Pini, F. Totti, M. Dressel, and L. Bogani, *Nat. Mater.*, 2013, **12**, 202; (b) R. Sessoli, M.-E. Boulon, A. Caneschi, M. Mannini, L. Poggini, F. Wilhelm, A. Rogalev, *Nat Phys* 2015, **11**, 69-74.
- 11 C. Coulon, V. Pianet, M. Urdampilleta, R. Clérac, *Struct. Bond.*, Vol. 164 (Ed.: S. Gao), Springer Berlin Heidelberg, 2015, pp. 143.

- 12 (a) H. Miyasaka, M. Julve, M. Yamashita, R. Clérac, *Inorg. Chem.*, 2009, **48**, 3420; (b) W. X. Zhang, R. Ishikawa, B. Breedlove, M. Yamashita, *RSC Adv.*, 2013, **3**, 3772.
- 13 X. X. Meng, W. Shi, P. Cheng, *Coord. Chem. Rev.*, 2019, **378**, 134.
- 14 (a) H. Miyasaka, T. Madanbashi, K. Sugimoto, Y. Nakazawa, W. Wernsdorfer, K. Sugiura, M. Yamashita, C. Coulon, R. Clérac, *Chem. Eur. J.*, 2006, **12**, 7028; (b) N. Ishii, Y. Okamura, S. Chiba, T. Nogami, T. Ishida, *J. Am. Chem. Soc.* 2008, **130**, 24; (c) M. Bałanda, Z. Tomkowicz, W. Haase, M. Rams, *J. Phys. Conf. Ser.*, 2011, **303**, 012036; (d) M. G. F. Vaz, R. A. A. Cassaro, H. Akpinar, J. A. Schlueter, P. M. Lahti, M. A. Novak, *Chem. Eur. J.*, 2014, **20**, 5460 ; (e) R. Ishikawa, K. Katoh, B. K. Breedlove, M. Yamashita, *Inorg. Chem.*, 2012, **51**, 9123; (f) R. A. A. Cassaro, S. G. Reis, T. S. Araujo, P. M. Lahti, M. A. Novak, M. G. F. Vaz, *Inorg. Chem.*, 2015, **54**, 9381.
- 15 (a) L. Bogani, C. Sangregorio, R. Sessoli, D. Gatteschi, *Angew. Chem., Int. Ed.*, 2005, **44**, 5817; (b) K. Bernot, L. Bogani, A. Caneschi, D. Gatteschi, R. Sessoli, *J. Am. Chem. Soc.*, 2006, **128**, 7947 ; (c) X. F. Wang, Y. G. Li, P. Hu, J. J. Wang, L. C. Li, *Dalton Trans.*, 2015, **44**, 4560. (d) R. N. Liu, L. C. Li, X. L. Wang, P. P. Yang, C. Wang, D. Z. Liao, J. P. Sutter, *Chem. Commun.*, 2010, **46**, 2566. (e) P. Hu, X. F. Wang, Y. Ma, Q. L. Wang, L. C. Li, D. Z. Liao, *Dalton Trans.*, 2014, **43**, 2234 ; (d) T. Han, W. Shi, Z. Niu, B. Na, P. Cheng, *Chem. Eur. J.*, 2013, **19**, 994.
- 16 Z. X. Wang, X. Zhang, Y. Z. Zhang, M. X. Li, H. H. Zhao, M. Andruh, K. R. Dunbar, *Angew. Chem. Int. Ed.*, 2014, **53**, 11567 .
- 17 M. Zhu, L. Li, J.-P. Sutter, *Inorg. Chem. Front.*, 2016, **3**, 994.
- 18 (a) M. Yang, J. Xie, Z. Sun, L. C. Li, J. P. Sutter, *Inorg. Chem.*, 2017, **56**, 13482; (b) M. Zhu, P. Hu, Y. G. Li, X. F. Wang, L. C. Li, D. Z. Liao, V. M. L. Durga Prasad Goli, S. Ramasesha, J. P. Sutter, *Chem. Eur. J.*, 2014, **20**, 13356; (c) X. F. Wang, P. Hu, L. C. Li, J. P. Sutter, *Inorg. Chem.*, 2015, **54**, 9664.
- 19 J. P. Sutter, M. L. Kahn, S. Golhen, L. Ouahab, O. Kahn, *Chem. Eur. J.* 1998, **4**, 571.
- 20 (a) H. Iwamura, *Pur. Appl. Chem.* 1986, **58**, 187. (b) L. Catala, J. L. Moigne, N. Kyritsakas, P. Rey, J. J. Novoa, P. Turek, *Chem. Eur. J.*, 2001, **7**, 2466.
- 21 (a) D. Casanova, M. Lluell, P. Alemany, S. Alvarez, *Chem. Eur. J.* 2005, **11**, 1479. (b) Lluell, M. Casanova, D. Cirera, J. Alemany, P. Alvarez, S. SHAPE, 2.1 ed.; University of Barcelona: Barcelona, Spain, 2013.
- 22 (a) D. Gatteschi, J. Laugier, P. Rey, C. Zanchini, *Inorg. Chem.*, 1987, **26**, 938; (b) D. Luneau, P. Rey, J. Laugier, P. Fries, A. Caneschi, D. Gatteschi, R. Sessoli, *J. Am. Chem. Soc.*, 1991, **113**, 1245.
- 23 (a) M. L. Kahn, J. P. Sutter, S. Golhen, P. Guionneau, L. Ouahab, O. Kahn and D. Chasseau, *J. Am. Chem. Soc.*, 2000, **122**, 3413; (b) K. Bernot, F. Pointillart, P. Rosa, M. Etienne, R. Sessoli, D. Gatteschi, *Chem. Commun.*, 2010, **46**, 6458.
- 24 D. Shiomi, M. Tamura, H. Sawa, R. Kato, M. Kinoshita, *J. Phys. Soc. Jpn.*, 1993, **62**, 289.
- 25 (a) S. M. J. Aubin, Z. Sun, L. Pardi, J. Krzystek, K. Folting, L. C. Brunel, A. L. Rheingold, G. Christou, D. N. Hendrickson, *Inorg. Chem.*, 1999, **38**, 5329; (b) C. Dekker, A. F. M. Arts, H. W. de Wijn, A. J. van Duynveldt, J. A. Mydosh, *Phys. Rev. B*, 1989, **40**, 11243.

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5. References

Experimental section

1. Materials and Physical Measurements

All reagents and solvents were purchased from commercial sources and used as received. The biradical ligand NITPhPybis was prepared according to literature method.^[S1] Elemental analysis (for C, H, and N) was implemented on a PerkinElmer 240 elemental analyzer. Infrared spectra were measured in the 4000–400 cm⁻¹ range on KBr pellets using a Bruker TENOR 27 spectrometer. Magnetic measurements were performed on a Quantum Design MPMS 5 SQUID magnetometer using crystalline powder samples mixed to grease. Magnetic susceptibility data were corrected for the diamagnetic contribution of all the constituent atoms with Pascal's constants and the sample holder.

2. Synthesis

[DyCu₂(hfac)₃](NITPhPybis) **1**

A mixture of Dy(hfac)₃·2H₂O (0.01 mmol, 0.0082 g) and Cu(hfac)₂ (0.02 mmol, 0.009 g) was dissolved in 15 mL dry boiling *n*-hexane. The solution was maintained to reflux for 6 h, then cooled to 50 °C and NITPhPybis ligand (0.01 mmol, 0.0047 g) in a CHCl₃ solution (6 mL) was added to the above solution. The resulting solution was heated for 15 min and then was cooled to room temperature, filtered, and stored at room temperature for slow evaporation. After a week, dark-green strip crystals were obtained for single crystal X-ray analysis. Yield: 65%. Calcd for C₆₀H₃₈Cu₂F₄₂DyN₅O₁₈ (2204.56 g mol⁻¹): C, 32.69; H, 1.72; N, 3.18. Found: C, 32.67; H, 2.05; N, 3.25. FT-IR (KBr): 3417(m), 2877(m), 2432(m), 1789(m), 1653(s), 1533(s), 1477(s), 1360(s), 1154(s), 1074(s), 949(s), 860(s), 662(s), 588(m), 547(s) cm⁻¹.

3. X-ray crystallography

The crystal data were collected at 113 K on a Rigaku Saturn CCD diffractometer (Mo/K α radiation, λ = 0.71073 Å). The multi-scan absorption corrections were conducted using the SADABS program.^[S2] These structures were solved by direct methods and then refined by full-matrix least-squares on F^2 using crystallographic software package program SHELXS-2014.^[S3] Non-H atoms were refined anisotropically. All hydrogen atoms were set in calculated positions and refined as riding on the corresponding non-hydrogen atoms. Information concerning detailed crystallographic data collections and structure refinements parameters are provided in Table S1. Selected bond distances (Å) and angles (°) are shown in Table S2. Crystallographic data for the structural analyses have been deposited with the Cambridge Crystallographic Data Centre, CCDC 1879743 contain the supplementary crystallographic data for complex **1**. These information can be obtained free of charge via www.ccdc.cam.ac.uk/data_request/cif.

Table S1. Crystallographic data and structure refinement parameters for **1**.

Complexes	1Dy
Empirical formula	C ₆₀ H ₃₈ Cu ₂ F ₄₂ DyN ₅ O ₁₈
<i>Mr</i>	2204.53
<i>T</i> (K)	113
Crystal system	monoclinic
Space group	<i>P</i> 2 ₁ / <i>c</i>
<i>a</i> /Å	20.944(4)
<i>b</i> /Å	20.520(4)
<i>c</i> /Å	19.008(4)
α /°	90.000
β /°	95.45(3)
γ /°	90.000
<i>V</i> /Å ³	8132(3)
<i>Z</i>	4
<i>D</i> _{calcd} /g cm ⁻³	1.801
μ /mm ⁻¹	1.590
ϑ /°	1.816to 25.009
<i>F</i> (000)	4316
Collected reflections	64162
Independent reflections	14313
<i>R</i> _{int}	0.0749
GOF(<i>F</i> ²)	1.069
<i>R</i> 1/ <i>wR</i> 2 (<i>I</i> > 2σ(<i>I</i>))	0.0806/0.2053
<i>R</i> 1/ <i>wR</i> 2 (all data)	0.0997/0.2218

Table S2. Selected bond lengths [Å] and angles [°] for **1**.

Bond distances			
Dy(1)-O(2)	2.367(5)	Cu(2)-O(1)	2.388(6)
Dy (1)-O(3)	2.385(5)	Cu(2)-O(1)#1	2.388(6)
Dy (1)-O(13)	2.347(6)	Cu(2)-O(9)	1.932(6)
Dy (1)-O(14)	2.368(6)	Cu(2)-O(10)	1.938(6)
Dy (1)-O(15)	2.298(5)	Cu(3)-O(4)	2.604(6)
Dy (1)-O(16)	2.374(5)	Cu(3)-O(4)#1	2.604(6)
Dy (1)-O(17)	2.306(6)	Cu(3)-O(11)	1.939(6)
Dy (1)-O(18)	2.374(5)	Cu(3)-O(12)	1.922(6)
Cu(1)-N(1)	2.038(7)	O(1)-N(2)	1.259(9)
Cu(1)-O(5)	1.995(7)	O(2)-N(3)	1.304(8)
Cu(1)-O(6)	1.935(6)	O(3)-N(4)	1.293(8)
Cu(1)-O(7)	2.129(6)	O(4)-N(5)	1.268(9)
Cu(1)-O(8)	1.943(6)		
Angles			
O(2)-Dy(1)-O(3)	82.4(7)	O(1)-Cu(2)-O(1)#1	180.0
O(13)-Dy(1)-O(2)	115.9(2)	O(9)#1-Cu(2)-O(9)	180.0
O(13)-Dy(1)-O(3)	71.3(1)	O(10)#1-Cu(2)-O(10)	180.0
O(15)-Dy(1)-O(2)	146.3(2)	O(10)-Cu(2)-O(1)	98.3(3)
O(15)-Dy(1)-O(3)	83.6(2)	O(10)-Cu(2)-O(1)#1	81.7(3)
O(17)-Dy(1)-O(2)	84.0(2)	O(4)-Cu(3)-O(4)#1	180.0
O(17)-Dy(1)-O(3)	142.8(3)	O(11)-Cu(3)-O(11)#2	180.0
O(5)-Cu(1)-O(6)	90.9(3)	O(12)-Cu(3)-O(12)#2	180.0
O(7)-Cu(1)-O(8)	89.7(2)	N(3)-O(2)-Dy(1)	136.6(5)
O(5)-Cu(1)-N(1)	144.2(3)	N(4)-O(3)-Dy(1)	138.7(4)
O(6)-Cu(1)-N(1)	93.1(3)	N(2)-O(1)-Cu(2)	147.3(7)
O(7)-Cu(1)-N(1)	108.4(3)	N(5)-O(4)-Cu(3)	144.4(5)
O(8)-Cu(1)-N(1)	92.6(3)		

Symmetry transformations used to generate equivalent atoms: #1 -x,y,-z+1/2 ;#2 -x+1, y, -z+ 1/2

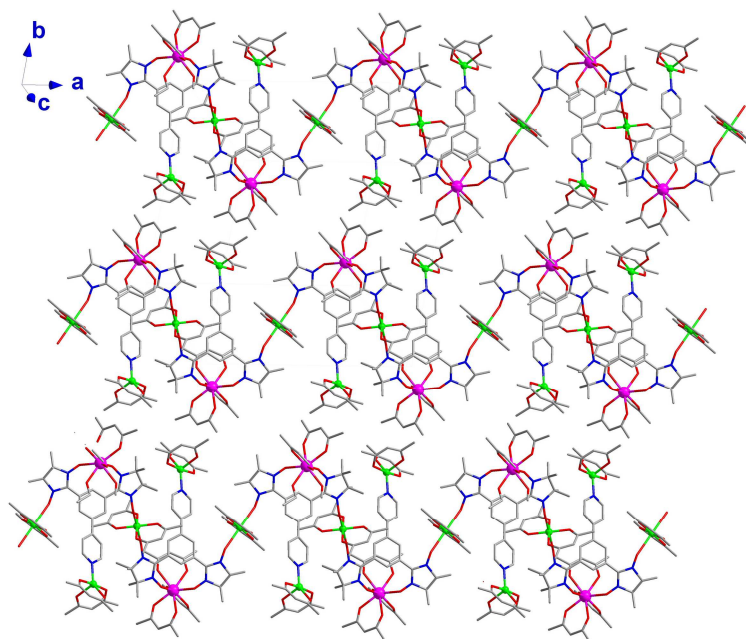


Fig S1. Packing diagram of complex **1**. All H, F atoms are omitted for clarity.

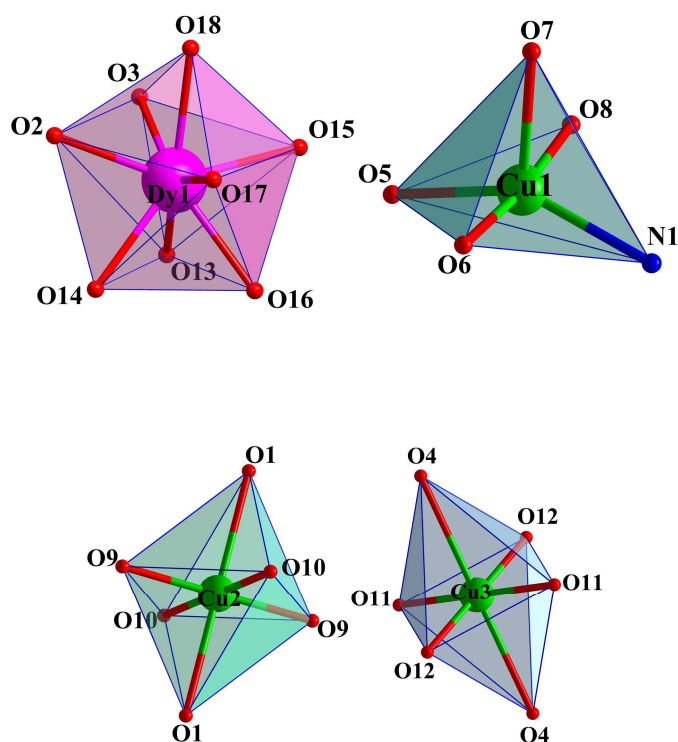


Fig S2. Coordination polyhedron of the Dy^{III}, Cu^{II} 1, Cu^{II}2 and Cu^{II}3 ions.

Table S3. Detailed geometry analysis results by Continuous Shape Measures

Compound	SAPR-8	TDD-8	JBTPR-8	BTPR-8	JSD-8
1Dy	2.074	1.879	1.489	0.743	3.658

SAPR-8: square antiprism. TDD-8: triangular dodecahedron. JBTPR-8 : Biaugmented trigonal prism J50.

BTPR-8: biaugmented trigonal prism. JSD-8 : Snub diphenoid J84

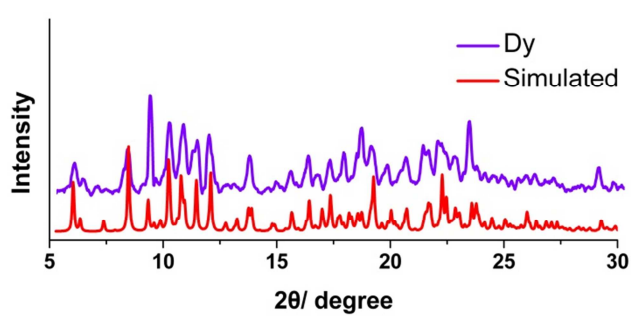


Fig. S3 Powder X-ray diffraction pattern of Dy complex at 293 K, together with the calculated pattern from the crystal structure data.

4. Magnetic measurements

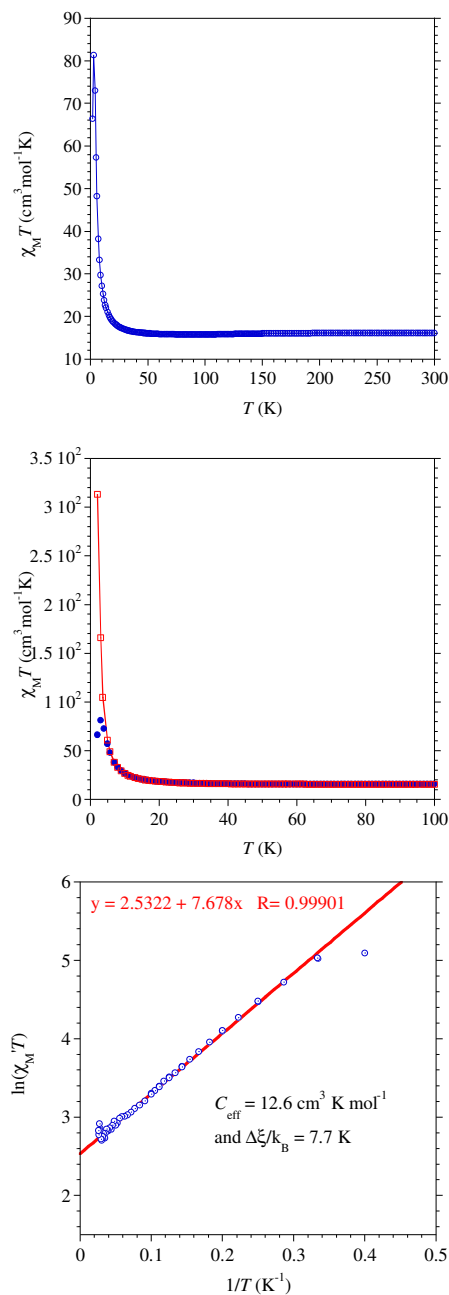


Figure SI4. Temperature dependence of $\chi_M T$ with χ_M (*top*) recorded in an applied field of 1 kOe, (*middle*) variation of $\chi_M T$ below 100 K in an applied field of 50 Oe (in red) and 1 kOe (blue) revealing a saturation effect in larger field; and , (*bottom*) $\ln(\chi_M' T) = f(T^{-1})$ (χ_M' , the in-phase susceptibility measured with an AC field of 3 Oe ($\nu = 100$ Hz) in the absence of a dc field), the solid red line is the best fit of the Glauber's expression to the linear part between 3 and 25 K.

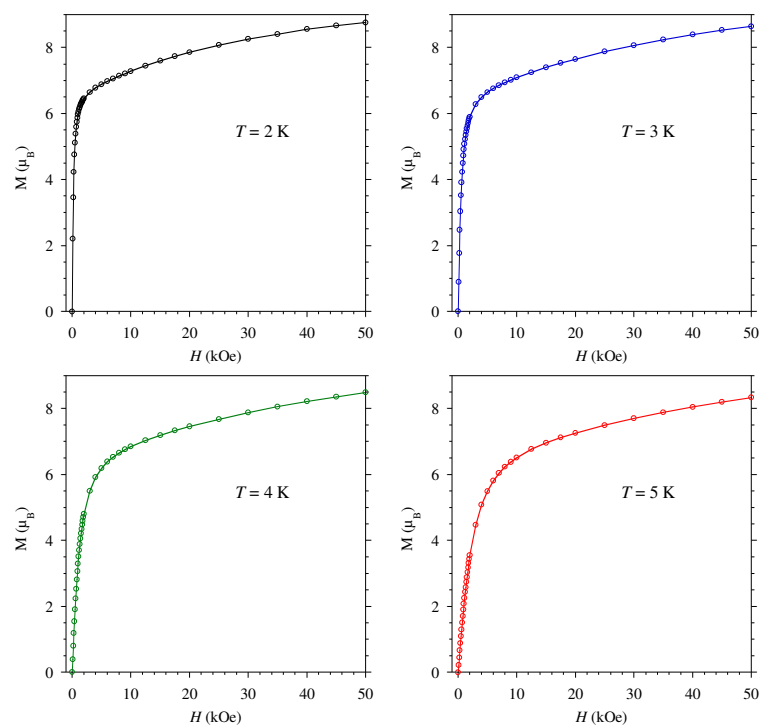


Figure S5. Field dependence of the magnetization at different temperatures for **1**.

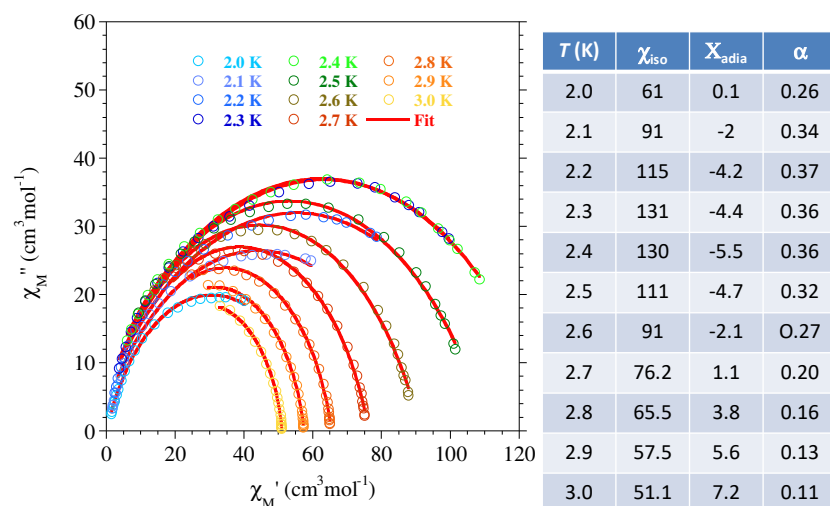


Figure S6. $\chi_M' = f(\chi_M'')$, i.e. Cole-Cole plots and best fits (solid lines) with a generalized Debye model, the fit parameters are gathered in the table.

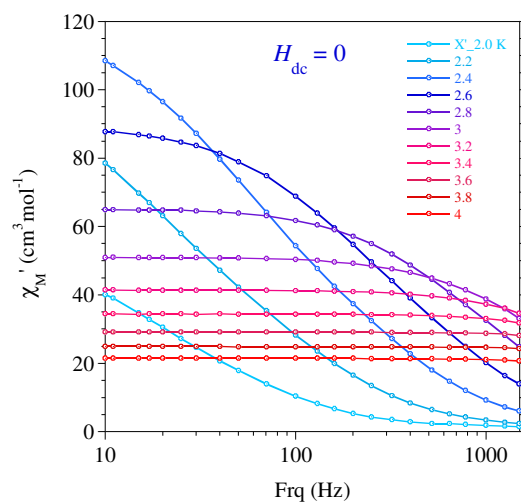


Figure S7. Frequency-dependent out-of-phase signals (χ'') for **1** under zero dc field with $H_{AC} = 3$ Oe

5. References

- [S1]. H. D. Li, J. Sun, M. Yang, Z. Sun, J. K. Tang, Y. Ma, L. C. Li., *Inorg. Chem.* 2018, 57, 9757-9765.
- [S2].Sheldrick, G. M. SADABS, Program for Absorption Correction of Area Detector Frames, BRUKER AXS Inc., Madison, WI.
- [S3].Sheldrick, G. M. SHELXS-2014, Program for structure solution; Universitat Gottingen: Gottingen, Germany, 2014.