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Investigating laminar mixing in high pressure microfluidic systems

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Abstract

In this study, the hydrodynamic behavior of coflowing fluids CO₂ and ethanol has been investigated in a high-pressure microfluidic reactor working at supercritical conditions, in which the two fluids are completely miscible. The velocity field has been measured by Micro Particle Image Velocimetry (μ PIV) for different temperatures between 20 and 50 °C at a fixed pressure of 100 bar. Meanwhile, we have developed a model to investigate numerically the mixing. By comparing the experimental results to a three-dimensional numerical simulation, the mixing model has been validated for the laminar coflow in the micromixer. In order to understand the mixing condition effects, several parameters have been investigated, namely: the Reynolds number, the temperature and the CO₂/ethanol ratio. A mixing time constant is defined by using the segregation intensity curve and later used to characterize the mixing quality. The characteristic mixing time has been related to the laminar energy dissipation rate ϵ , similarly to the stretching efficiency model in previous studies. The mixing quality is eventually analyzed in term of segregation index and mixing time.

Highlights

Velocity field of CO₂-ethanol fluid mixture is measured in a micromixer under SAS conditions by μ PIV.

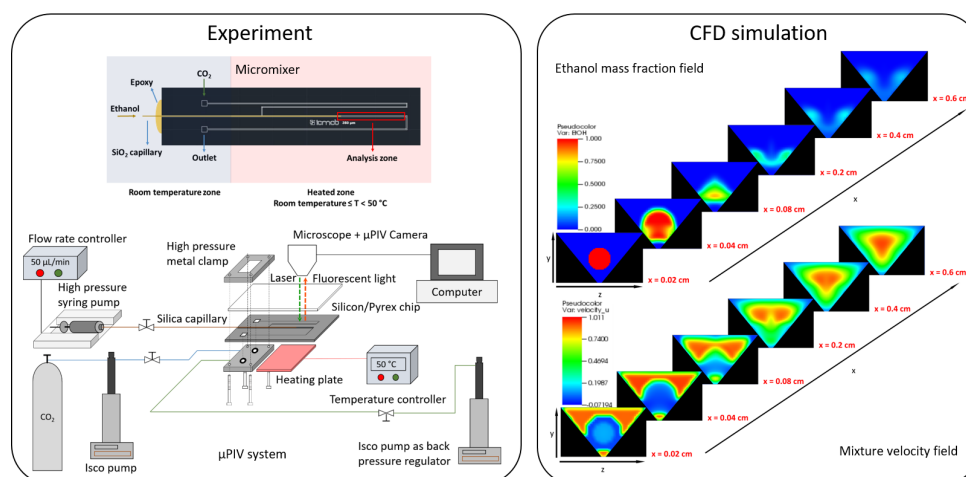
3D CFD simulation is validated by the experiments.

The effects of the operating conditions on the laminar mixing quality are examined in the micromixer.

Characteristic mixing time is plotted as a function of the energy dissipation rate.

Keywords— Supercritical antisolvent process (SAS); Micro particle image velocimetry (μ PIV); High pressure microfluidic mixing; Computational fluid dynamics (CFD); Mixing time constant

Graphical abstract



Nomenclature

D	diffusion coefficient	$m^2 \cdot s^{-1}$
D_h	hydraulic diameter	m
I	mixing index	
L	characteristic length	m
N	Total number	
Q	flow rate	$m^3 \cdot s^{-1}$
R	ideal gas constant	$J \cdot kg^{-1} \cdot K^{-1}$
Re	Reynolds number	
T	temperature	K
V	molar volume	$m^3 \cdot mol^{-1}$
g	gravitational acceleration	$m \cdot s^{-2}$
\mathbf{u}	velocity vector	$m \cdot s^{-1}$
a	first coefficient in Peng–Robinson equation of state	
b	second coefficient in Peng–Robinson equation of state	
k, l	interaction parameters in Peng–Robinson equation of state	
l_s	mixing channel length	m
p	pressure	bar
t	time	s
u	absolute velocity	$m \cdot s^{-1}$
x	mass fraction	

Greek symbols

α	α function in Peng–Robinson equation of state	
ϵ	energy dissipation	$W \cdot kg^{-1}$ or $m^2 \cdot s^{-3}$
μ	dynamic viscosity	$Pa \cdot s$
ν	kinematic viscosity	$m^2 \cdot s^{-1}$
ω	acentric factor	
ρ	density	$kg \cdot m^{-3}$

Subscripts

0 initial temperature

c critical point

$EtOH$ ethanol

m mixture

1 Introduction

As a promising bottom up method to process nano or microparticles with narrow particle size distribution, supercritical antisolvent processes (SAS) have been largely studied for decades. The SAS is an environmentally friendly process, mostly using supercritical CO₂ as an antisolvent for generating the precipitation of materials. The main interest is that no cleaning steps are required to remove the solvent residues at the end of the process. Indeed, the depressurization at the end of precipitation results in a removal of the CO₂ as a gas. The general principle consists in dissolving a solute in a dedicated solvent and to mix it with an antisolvent around its critical point, in which the solute has a very low solubility. Because of the created supersaturation due to the mixing of these two fluids, the solute precipitates in a burst nucleation generating nuclei, which later grow to form eventually particles. The supersaturation, as the precipitation driving force, is the key point for the particle precipitation and is mainly influenced by the mixing of the solution (the solute and the solvent) and the antisolvent. A quick mixing of these two fluids promotes a higher supersaturation and a higher nucleation rate. Consequently, the particle growth is relatively restrained so that the obtained particles have a narrow size distribution and a small size. By applying a supercritical antisolvent, one improves the mixing process with other fluids because compared to liquid antisolvents, the supercritical ones are less viscous and dense as fluid, possessing higher diffusivity.

In the literature, many experimental studies have been conducted to capture process condition effects. However, some controversial results are found: some authors demonstrated that an increase of the initial solute concentration may lead to a growth of mean particle size (De Marco and Reverchon, 2011; Rossmann *et al.*, 2014). Whereas, based on the experimental results, other authors found some opposite effect on the particle size (Boutin, 2012). Concerning the temperature, again, several teams made opposite claims on the temperature dependency of the particle size (De Marco and Reverchon, 2011; Campardelli *et al.*, 2017; Chang *et al.*, 2008; Miguel *et al.*, 2015). These debatable results reveal a lack of fundamental understanding on the SAS process. As a matter of fact, the preliminary and essential work is to better understand the fluid behaviors during mixing process and to optimize conditions to meet an efficient mixing.

As a powerful tool to understand fluid flow mechanisms, microfluidics provide means towards controlled hydrodynamics and temperature inside microchannels. Furthermore, due to the small dimensions (typically between 10 and 500 μm), micromixers have been proven to be more efficient than the conventional industrial ones (Falk and Commenge, 2010), so fluid mixing has been largely studied and characterized in microchannels at low pressure (Aubin *et al.*, 2010), because many microreactors are not designed for high pressure applications. Whereas, SAS processes could also be applied on chip in a successful manner to attain the processing of small organic particles, as previously demonstrated in our laboratory for P_3HT nanoparticles (Couto *et al.*, 2015). The originality of this paper is due to a combination of an experimental microfluidic approach with a CFD numerical modeling to investigate the mixing efficiency at small scales for a representative SAS case. The model system, which was chosen, consists of coflowing ethanol and CO₂ at a constant pressure of 100 bar. Thanks to *in-situ* micro Particle Image Velocimetry (μPIV) measurements within a coflow geometry microchannel, we have accessed the thermo-hydrodynamic behavior of the fluid mixing in the laminar regime and validated the numerical simulation. Then, the CFD simulation was used to explore the mixing quality in such micro device geometries by varying several operating parameters such as the Reynolds number, the temperature and the

CO₂/ethanol ratio. This was achieved thanks to the definition of a mixing index I_m , which will be described. Eventually, a characteristic time constant t_m is introduced and correlated to the energy dissipation rate ϵ .

2 Experimental Section

2.1 Chemical materials

The 100% CO₂ used in this work is supplied by Messer company. The 96% ethanol is provided by Xilab. The purchased fluorescent particles are polystyrene beads from Thermo Fisher Scientific and have a refractive index of 1.59, a density of 1.06 g/cm³ and an average diameter of 1 μ m. They are doped with red fluorescent dyes and packaged in deionized water as aqueous suspension at 1 wt.%, with an excitation maxima at 542 nm and an emission maxima at 612 nm.

2.2 Microreactor design and Fabrication

In this study, we have considered the design of a microfluidic reactor reproducing SAS experiment by contacting the solvent (ethanol) and the antisolvent (CO₂) in a coflow geometry.

The microreactor was made in Silicon-Pyrex, which is a micro fabrication technology largely utilized for high pressure microfluidics (Couto *et al.*, 2015; Marre *et al.*, 2010), combining the mechanical properties of silicon and Pyrex with the good thermal conductivity of silicon and the visible transparency of Pyrex, thus providing an easy optical access.

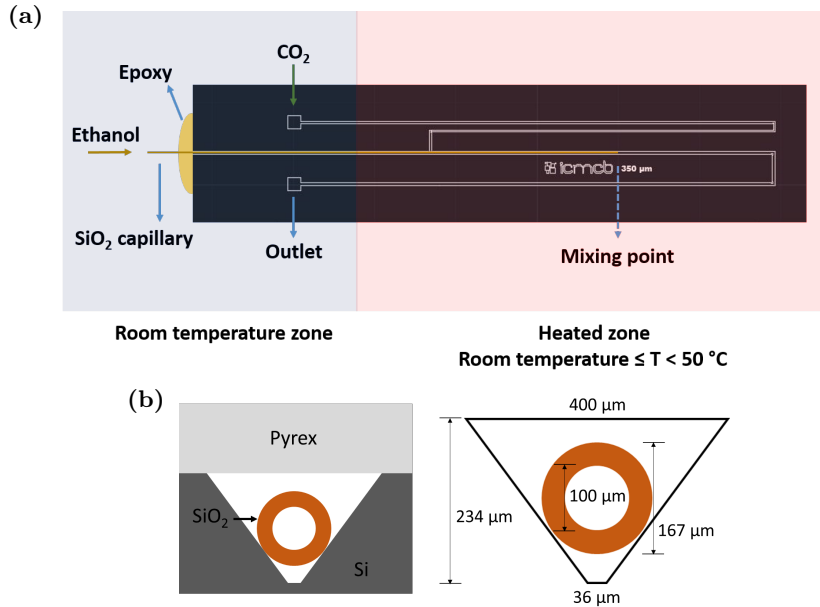


Figure 1: Illustration of the developed microfluidic device used for the experimental mixing studies: (a). Top view of the general design and (b). cross-section views with additional details of dimensions.

The microchannel was first wet etched in silicon wafer following a photolithography step involving the UV exposure of a photoresist through a photomask on which the design was printed. Ultimately, the etched silicon wafer was anodically bonded to a top Pyrex cover. The microchannel cross-section is trapezoidal due to the

wet etching process. The general design of the microreactor is presented in Figure 1a. A main channel (top and bottom width of 400, 36 μm respectively and height of 234 μm), where the mixing occurs, is fed by a side channel aiming at preheating the CO_2 flow. Then, a silica capillary (100 μm ID and 167 μm OD), used to introduce the solvent ethanol, is inserted inside the main channel and further epoxy glued to provide a full 3D injection as shown in the cross-section view of Figure 1b. This configuration allows working up to 150 bar in pressure. The temperature of the heated part of the chip is maintained thanks to a heating element controlled with an Eurotherm $\text{\textcircled{R}}$ temperature controller.

2.3 Experimental set-up and μPIV measurements

The microreactor is implemented in a general experimental set-up detailed in Figure 2. CO_2 and ethanol are injected with an ISCO 100 DM pump equipped with a cooling jacket, and a Harvard PhD 2000 high pressure syringe pump, respectively. The pumps are connected to the microreactor and the silica capillary thanks to a house-made compression part and Valco/Vici commercial fittings, respectively, as seen in Figure 2. The overall pressure is controlled using another ISCO pump working downstream in constant pressure mode. A heating plate is placed on the silicon surface of the microchip to control the temperature. In order to acquire locally the fluid velocities during the injection and mixing process, the hydrodynamic inside the microreactor was characterized by an *in-situ* μPIV system. The μPIV set-up includes (i) a laser diode emitting at $\lambda = 532 \text{ nm}$, which frequency is set to 4 Hz, with the pulse duration controlled at 15 μs with a time delay at 6 μs , (ii) a ZEISS Axiovert 200M microscope with a 20 \times magnification objective, (iii) a CCD camera of the Vieworks company displaying a resolution of 3296×2472 pixels, with a frame rate of 10 fps. The PIV experimental data are processed by using the software Hiris, developed by R&D Vision company.

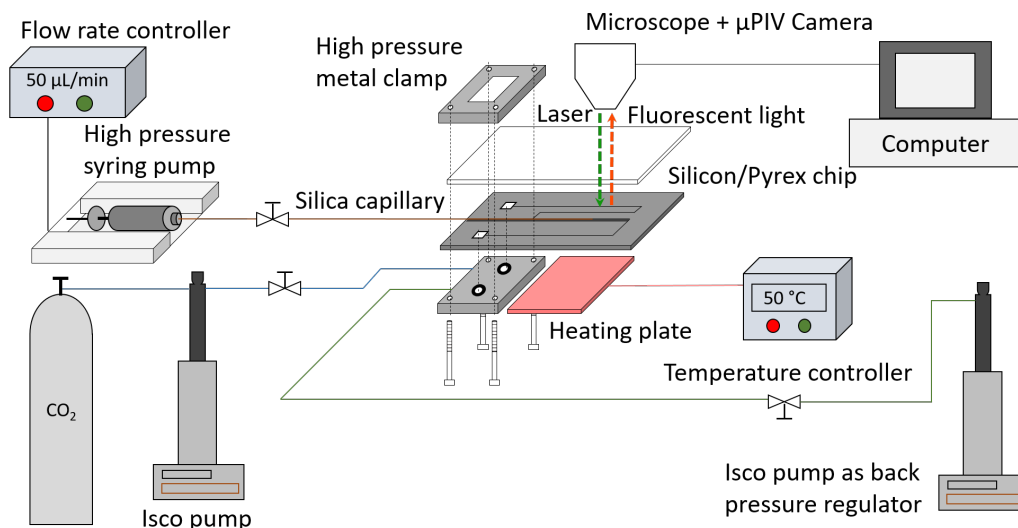


Figure 2: Microfluidic set-up developed for the μPIV analysis of the high pressure and controlled temperature mixing of CO_2 and ethanol inside a microchannel.

The general principle of this characterization consists in tracking the displacement of fluorescent particles, which follow the streamlines of the fluids. The particles, in suspension in the fluid, are excited by a laser inducing the fluorescence of the particles, which can be captured by a camera. In a typical measurement procedure, by comparing the particles' positions between two pictures taken at a very short interval of time Δt , one can

estimate the instantaneous velocity field and then the mean velocities of these particles, which are assumed to be the local velocity of the mixing fluids because the particle diameters are much smaller than the width of the microchannel, so they are supposed to follow the current streamlines in the fluid mixture.

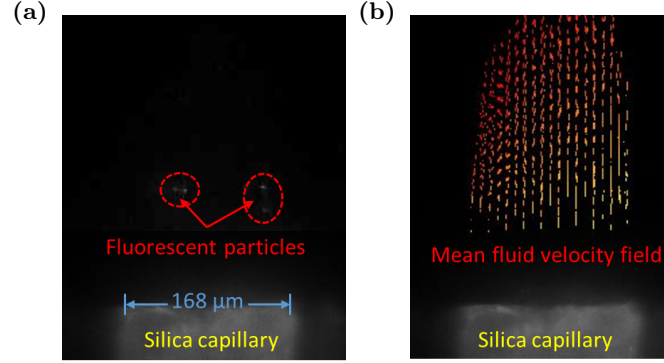


Figure 3: Illustration of velocity field reconstruction by μ PIV (a). A typical image of the acquisition mode of μ PIV; (b). Mean velocity field of mixing fluids processed by Hiris software.

The tracer particles are preliminarily dried and diluted about 100 times in ethanol (0.01% of volume). The set-up is first pressurized gradually up to 100 bar from downstream with an ISCO pump filled with pure ethanol as a back pressure regulator. The liquefied CO_2 is then injected by another ISCO pump in which the temperature is set originally at -5°C and the pressure at 100 bar. Once the pressure and temperature are stabilized, the particle suspension in ethanol is injected inside the capillary thanks to a high pressure syringe pump. The two fluids encounter at the outlet of the capillary, forming a coflow mixing. While the laser arrives through the focused mixing zone, the particles in the microchannel are excited and return red fluorescent light at a wavelength of about 612 nm through a dichroic filter and captured by the camera. In a recorded acquisition of 1000 images (500 pairs), the tracer particle displacements can be observed in a black background to determine the average particle velocity field with a fixed interframe time Δt of 100 μs . In the recorded images, the measured area is an adjustable rectangle of 0.5 mm \times 0.3 mm (Figure 3). The images are integrated with a field depth of 10 μm . One pixel corresponds to 0.263 μm and the window size is chosen to be 64 pixels \times 32 pixels, according to estimated fluid velocity. The reproducibility has been tested and the measurements are reproduced 3 times under fixed conditions (Figure 5 and Figure 6). During the experiment, no agglomeration of particles has been observed. However, some sedimentation is sometimes detected and some particles are trapped and attached to the wall. Nevertheless, according to the good reproducibility of the results, this effect could be neglected.

3 Numerical modeling

In the numerical modeling, the fluid velocity and ethanol concentration field are obtained by the resolution of the continuity equation, Navier-Stokes (NS) equation and the species transport equation. The effects of the presence of the fluorescent particles on the fluid properties are neglected so the mixture contains only CO_2 and ethanol, which are supposed to be completely miscible in the investigated conditions according to their mixture phase diagram. Furthermore, the fluid flow is assumed to be isothermal.

3.1 Hydrodynamics

Since the mixture density relies only on the mixture composition and the conditions investigated, which always ensure the pressure above the critical point of the CO₂-ethanol mixture, the equation of continuity (Equation 1) and the NS equation (Equation 2) are solved for a completely miscible and incompressible fluid mixture in one phase, as previously reported (Erriguible *et al.*, 2013).

$$\nabla \cdot \mathbf{u} = 0 \quad (1)$$

$$\rho \left(\frac{\partial \mathbf{u}}{\partial t} + \mathbf{u} \cdot \nabla \mathbf{u} \right) = -\nabla p + \nabla \cdot (\mu(\nabla \mathbf{u} + \nabla^T \mathbf{u})) \quad (2)$$

The ethanol mass fraction in the fluid mixture is calculated by the conservation equation of the species, according to the Fick's law (Equation 3). The mass fraction of CO₂ can then be directly deduced.

$$\frac{\partial \rho x_j}{\partial t} + \nabla \cdot (\rho x_j \mathbf{u} - \rho D_j \nabla x_j) = 0 \quad (3)$$

3.2 Mixture thermophysical properties

In order to solve the equations mentioned in the model, it is necessary to have access or to calculate the thermophysical properties of the CO₂-ethanol mixture, such as density, viscosity and diffusion coefficient.

The density of the fluid mixture is considered to be a function of its composition as well as the temperature and the pressure. It is calculated by the Peng Robinson equation of state (PREOS) (Equation 4) with Van der Waals mixing rules (Equation 5). The parameters a and b are solved first for pure components and then for the mixture. The interaction parameters k_{ij} and l_{ij} depend on the temperature (Maeta *et al.*, 2015).

$$p = \frac{RT}{V_m - b_m} - \frac{a_m}{V_m(V_m + b_m) + b_m(V_m - b_m)} \quad (4)$$

$$a_m = \sum_i^n \sum_j^n x_i x_j a_{ij} \quad b_m = \sum_i^n \sum_j^n x_i x_j b_{ij} \quad (5)$$

$$a_{ij} = (1 - k_{ij})(a_{ii}a_{jj})^{0.5} \quad b_{ij} = (1 - l_{ij})\frac{b_{ii} + b_{jj}}{2}$$

$$a = \frac{0.45724\alpha R^2 T_c^2}{P_c} \quad b = \frac{0.0778RT_c}{P_c}$$

$$\alpha = \left(1 + (0.37464 + 1.5422\omega - 0.26992\omega^2)(1 - \sqrt{\frac{T}{T_c}}) \right)^2$$

The viscosities of the pure fluids are obtained from the NIST database for the considered experimental conditions and the mixture viscosity is calculated by a logarithmic mixing rule according to Equation 6:

$$\ln \mu_m = x_{CO_2} \ln \mu_{CO_2} + x_{EtOH} \ln \mu_{EtOH} \quad (6)$$

The diffusion coefficient of ethanol in CO₂ is estimated by applying the Hayduk-Minhas correlation (Equation 7)

(Hayduk and Minhas, 1982; Fadli *et al.*, 2010):

$$D_m = 1.33 \cdot 10^{-7} \cdot T^{1.47} \cdot V_{EtOH}^{-0.71} \cdot \mu_{CO_2}^{\frac{10.2}{V_{EtOH}^{-0.791}}} \quad (7)$$

4 Results and discussion

4.1 Numerical procedure

The partial differential equations (1-3) are solved numerically to obtain the velocities and the species mass fractions for each time step, thanks to the CFD open source code “Notus” developed at the I2M/TREFLE Department. Meanwhile, the thermophysical properties are calculated at each time step according to local species composition. The approximation of the model described in Section 3 has been realized by the finite volume method on a fixed staggered grid. An implicit centered scheme is applied for all the terms in the Navier-Stokes equation while the transport equation is solved by a weighted essentially non-oscillatory (WENO) explicit scheme (Jiang and Shu, 1996). The gravity is neglected in the confined microchip because of a small value of the Bond number. The velocity–pressure coupling is solved by the solver BiCGSTAB II, preconditioned with a Jacobi method for the prediction step and the MCG approach for the correction step (Falgout and Yang *et al.*, 2002).

Due to the trapezoidal asymmetrical geometry, a three-dimensional simulation is required and the configuration is exhibited in Figure 4. The whole shape of the simulated area is a cuboid of $0.6 \text{ cm} \times 0.023 \text{ cm} \times 0.04 \text{ cm}$. The circular injector has a length of 0.03 cm with an inner and an outer radius of $51 \text{ }\mu\text{m}$ and $84 \text{ }\mu\text{m}$. Solid walls are penalized to shape the trapezoidal channel. The mean velocities for both CO_2 and ethanol are provided at $x = 0$ based on the fluid flow rates and the fluid velocity is then developed by the simulation at the mixing point $x = 0.03 \text{ cm}$. The ethanol mass fraction is defined to be 1 in the injector and 0 at the outside, leading to pure ethanol and pure CO_2 at $x = 0$ for the boundary conditions of species transport. The grid size is chosen to be $10 \text{ }\mu\text{m}$, according to a convergence study, which is comparable to the grid size of μPIV measurement ($16.8 \text{ }\mu\text{m} \times 8.4 \text{ }\mu\text{m}$) so the total number of nodes is 552 000 ($600 \times 23 \times 40$). Given the high number of nodes required for describing the process with accuracy, the simulations are performed by a Message Passing Interface (MPI) parallel programming on 16 processors.

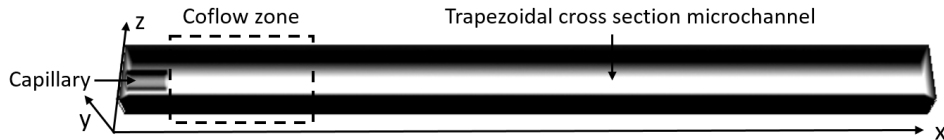


Figure 4: The 3D geometry of the microchannel in the CFD simulation.

4.2 Model validation

In order to validate the numerical model, the experimental and simulated data obtained for the fluid velocities were compared. The μPIV data were recorded on the mid cross-section of the channel corresponding

to the plane $y = 0$. First, the fluid velocities were measured in the microchannel far from the capillary tip, where the velocity profile is well developed and assumed to have a Poiseuille parabolic profile (Figure 5). Then, another area near the capillary tip was also investigated, in the mixing zone (Figure 6). The particles taken into account for the μ PIV fluid velocity measurements can be actually at different height in the channel in a three-dimensional area with a fine thickness. However, as mentioned before, the thickness integrated with our equipment is $10 \mu\text{m}$ and one can consider that all the recorded particle movements are in a two-dimensional plane. For the case far from the injector, a comparison is shown in Figure 5b and both μ PIV and numerical model provide similar results. The slight difference is probably due to the microfabrication procedure (wet etching step) as the microchannel can not be perfectly trapezoidal with smooth walls everywhere.

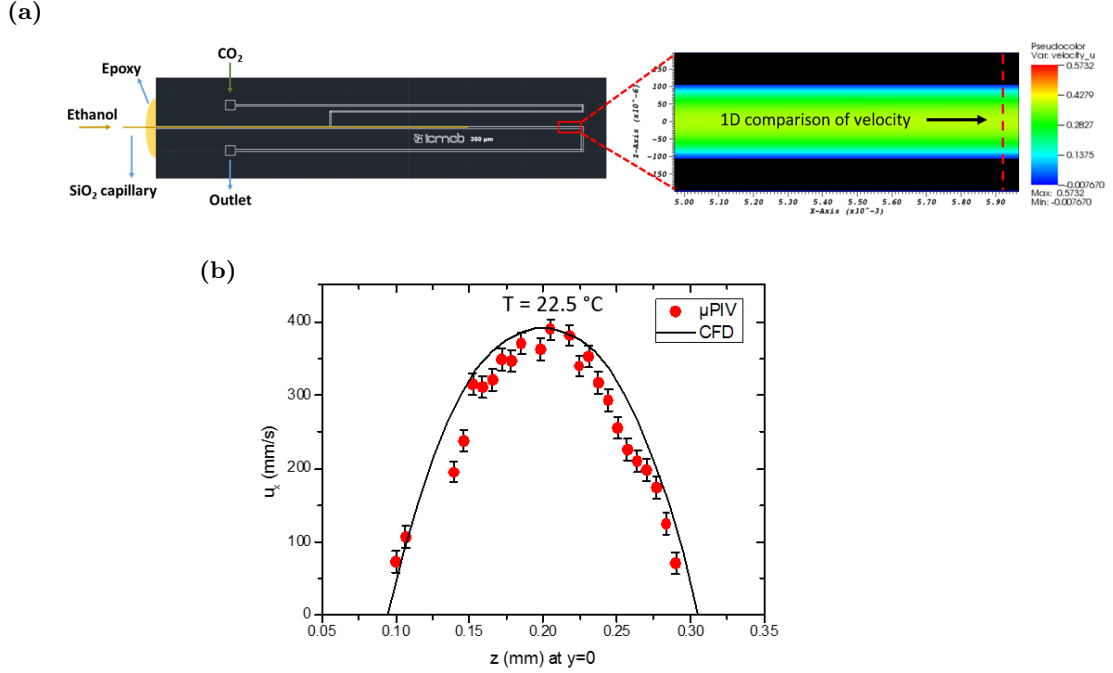


Figure 5: CFD validation for a fully developed flow (far from the injector): (a). 1D velocity comparison on the red dashed line far from the capillary outlet in the plane $y = 0$ (b). comparison at the mid-line for $T = 22.5 \text{ }^\circ\text{C}$, $\text{CO}_2 \text{ wt.\%} = 93.6$, with error bars according to replicated experiments (test No. 11 in Table 1).

Secondly, the fluid velocity at the capillary tip has been examined for different temperatures and fluid flow rates. Based on the experimental results, the velocity of the inner fluid (ethanol) drops first at the injector outlet. Actually the ethanol velocity in the capillary cannot be correctly measured by μ PIV because the silica capillary reflects light and the particle movement inside it cannot be captured. However, the velocity of ethanol at the tip of the injector can be simply deduced from its flow rate. The fluid velocity increases then gradually in the x direction along with the outer fluid CO_2 . A one dimensional comparison is proposed in Figure 6 and data are retrieved from the mid-line of the 2D plane (red dashed line in Figure 6a).

In general, according to the two comparisons above, both experiments and simulations present similar hydrodynamic behaviors. The numerical model has been validated with simulated velocities in agreement with measured velocity profiles obtained by μ PIV. We then used the CFD code to investigate the influence of various parameters on the mixing quality in the microfluidic chip.

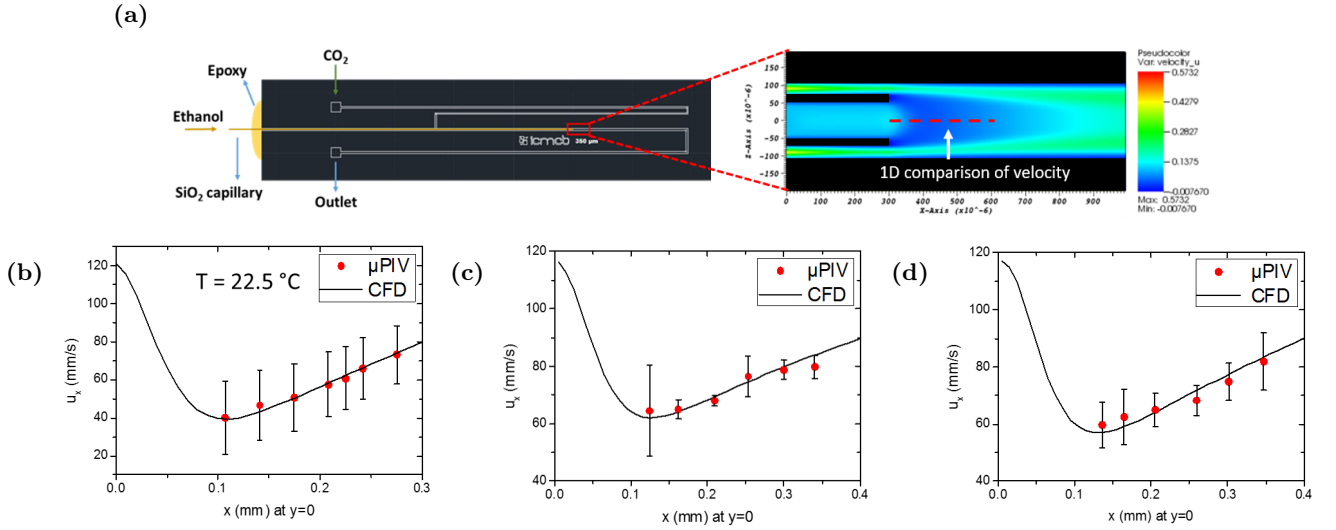


Figure 6: CFD validation in the vicinity of the injector: (a). 1D velocity comparison at the capillary outlet on the red dashed line in the plane $y = 0$ (b). comparison at the mid-line with error bars according to replicated experiments for $T = 22.5 \text{ }^\circ\text{C}$, $\text{CO}_2 \text{ wt.\%} = 93.6$ (test No. 11) (c). $T = 38 \text{ }^\circ\text{C}$, $\text{CO}_2 \text{ wt.\%} = 83.1$ (test No. 4) (d). $T = 47.5 \text{ }^\circ\text{C}$, $\text{CO}_2 \text{ wt.\%} = 83.1$ (test No. 5) in Table 1.

4.3 Fluid velocity and ethanol mass fraction field

A three-dimensional simulation case is presented in Figure 7a for the velocity component u_x in the x direction of the 2D plane $y = 0$. Similarly, the ethanol mass fraction field is shown in the 2D plane $y = 0$ (Figure 7b) and $z = 0$ (Figure 7c) for a Reynolds number of 1109, the temperature at $47.5 \text{ }^\circ\text{C}$ and the total mass fraction of CO_2 at 93.7% (No.12 in Table 1). The black color symbolizes the silicon wall of the microchannel or the silica capillary tube. For the ethanol mass distribution, the red and blue colors represent the pure ethanol and the CO_2 , respectively. As expected, the simulation exhibits a mean field flow as no turbulent fluctuation is present.

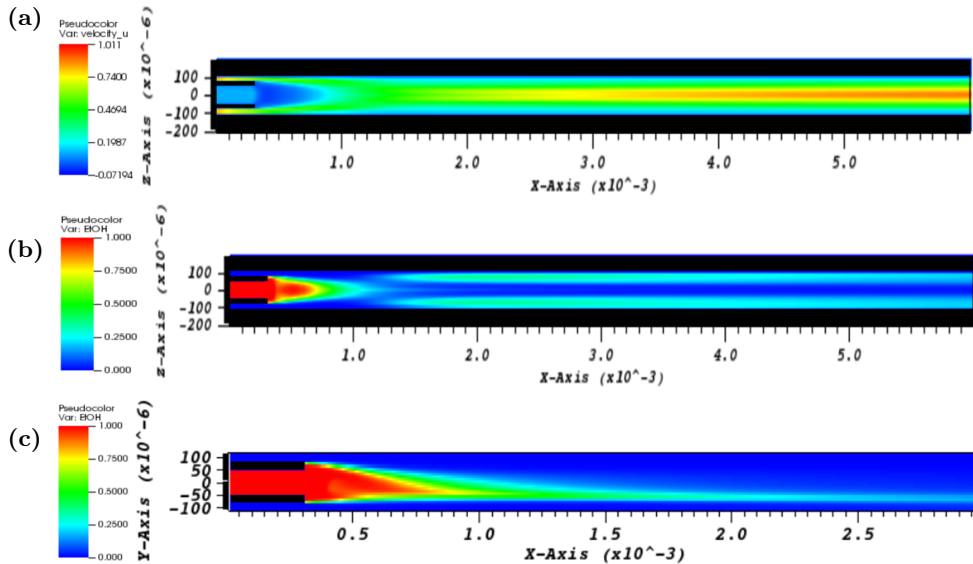


Figure 7: $y = 0$, $z = 0$ views of the 3D simulation of the CO_2 -ethanol mixing in the microchip for $T = 47.5 \text{ }^\circ\text{C}$ and $\text{CO}_2 \text{ wt.\%} = 93.7$ (No.12 in Table 1): (a) Velocity component u_x field ($y = 0$ plane); (b) Ethanol mass fraction field ($y = 0$ plane); (c) Ethanol mass fraction field ($z = 0$ plane).

Figure 7c shows the asymmetric ethanol mass fraction field in the plane $z = 0$. This result is related to the asymmetric trapezoidal shape of the channel and to the CO_2 velocity, which is higher than the ethanol

at the injector outlet. The flow of CO₂ first pushes the ethanol towards the bottom of the channel and the ethanol later goes up near the wall when the overall fluid velocity is higher in the center of channel, as indicated in Figure 8.

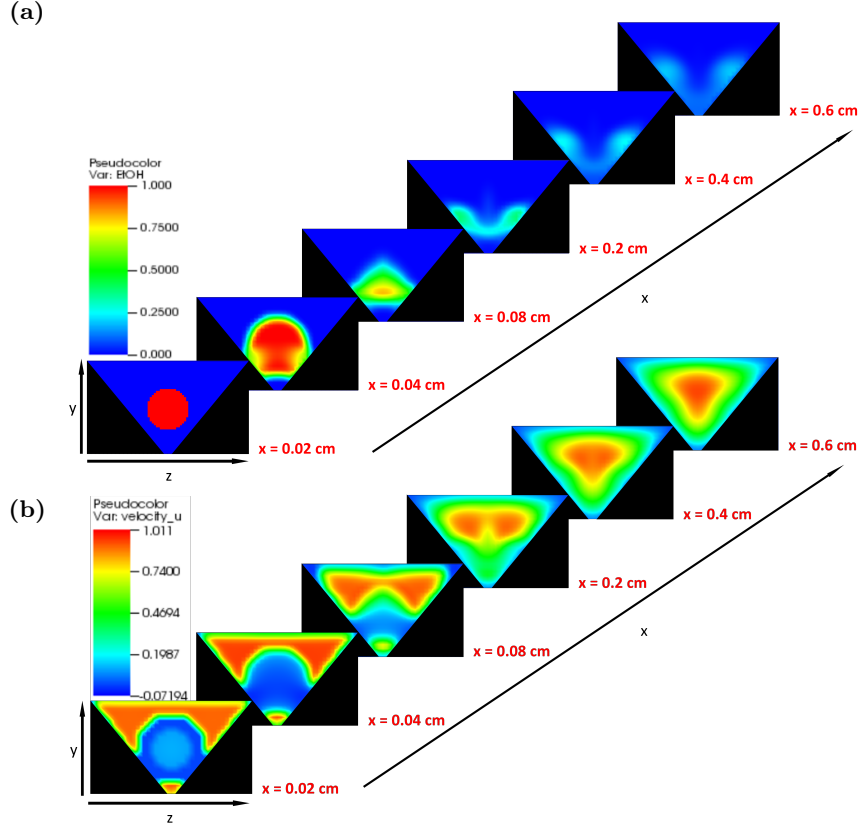


Figure 8: Ethanol mass fraction field and fluid velocity field variation in the asymmetrical 3D microchannel at steady state for $T = 47.5$ °C and CO_2 wt.% = 93.7 (No.12 in Table 1): (a). Ethanol field evolution in the x direction; (b). Velocity component u_x evolution in the x direction.

4.4 Reynolds number, temperature and CO₂/ethanol ratio effects on the mixing quality

In this part, various simulation conditions have been performed to examine their effects on the mixing efficiency. All parameters are listed in Table 1, with average values calculated for the fluid velocity component $\overline{u_x}$, the density $\overline{\rho}$, the viscosity $\overline{\mu}$ as well as the Reynolds number Re . These average values are described later.

(i) Mixing quality criteria

In order to investigate the effects of the process parameters on the mixing quality with simulated results of the validated numerical model, we have first defined a mixing index. The mixing quality is related to the homogeneity of the fluid mixture, which may be expressed classically by a relation known as the intensity of segregation or the mixing index (Equation 8) (Danckwerts, 1958):

$$I_m = \frac{\sum (x_i - \bar{x})^2}{N \cdot \bar{x} \cdot (1 - \bar{x})} \quad (8)$$

with I_m , the mixing index, x_i the mass fraction of component "i", \bar{x} the mean fraction of component "i" in

Table 1: Mixing conditions for the different simulation cases, along with the average fluid mixture properties and the characteristic time constant of mixing (* initial velocity ratio of CO₂ to ethanol, based on the fluid flow rates).

No.	T (°C)	CO ₂ (wt.%)	Q_{EtOH} ($\mu\text{L}\cdot\text{min}^{-1}$)	Q_{CO_2}	\bar{u}_x ($\text{m}\cdot\text{s}^{-1}$)	$\frac{u_{CO_2}}{u_{EtOH}}$ *	$\bar{\rho}$ ($\text{kg}\cdot\text{m}^{-3}$)	$\bar{\mu}$ ($\mu\text{Pa}\cdot\text{s}$)	$D\cdot 10^8$ ($\text{m}^2\cdot\text{s}^{-1}$)	Re	t_m (ms)
1	20.0	82.8	25	100	0.0386	1.22	935	118	1.58	62	28.92
2	28.0	82.8	50	100	0.0802	1.37	896	96	1.86	152	18.65
3	22.5	82.9	50	200	0.0827	1.39	923	111	1.65	136	18.64
4	38.0	83.1	50	200	0.0903	1.85	843	72	2.34	208	14.28
5	47.5	83.1	50	200	0.0959	2.66	786	51	3.12	295	11.21
6	22.5	90.6	50	400	0.1559	2.77	896	90	1.65	308	7.74
7	38.0	90.8	50	400	0.1767	3.70	789	58	2.34	478	6.45
8	24.0	93.5	50	600	0.2326	4.20	864	80	1.70	499	4.74
9	38.0	93.6	50	600	0.2816	5.50	713	53	2.34	749	3.18
10	47.5	90.9	50	400	0.1972	5.37	706	40	3.12	697	4.46
11	22.5	93.6	50	600	0.2318	4.16	875	83	1.65	486	4.77
12	47.5	93.7	50	600	0.3143	8.06	644	36	3.12	1109	3.18
13	47.5	93.7	22	266	0.1391	8.07	644	36	3.12	491	3.88
14	47.5	90.7	23	180	0.0889	5.29	704	40	3.12	312	6.77
15	38.0	94.8	25	371	0.1690	6.86	732	51	2.34	478	3.66
16	38.0	97.7	10	351	0.1693	16.23	671	47	2.34	478	3.86
17	38.0	84.6	100	447	0.1998	2.07	836	69	2.34	478	11.59

the trapezoidal plane, N the total number of numerical grid elements in the cross-section. This index has a value ranging between 0 and 1, implying respectively a homogeneous mixture and a total segregation. Based on the simulation results, we have calculated this index for each trapezoidal cross-section as a function of the corresponding time:

$$t = \frac{l_s}{\bar{u}_x} \quad (9)$$

with l_s the distance to the injector outlet and \bar{u}_x the mean fluid velocity component in the x direction far from the injector for a fully developed and homogeneous flow, calculated assuming the mass flow rate conservation:

$$\bar{u}_x = \frac{\rho_{CO_2} Q_{CO_2} + \rho_{EtOH} Q_{EtOH}}{\bar{\rho} A} \quad (10)$$

with ρ_{CO_2} , ρ_{EtOH} the densities of pure CO₂ and ethanol under the test conditions, Q_{CO_2} , Q_{EtOH} the volumetric flow rates in the microchip before mixing, $\bar{\rho}$ the density of complete mixture calculated by the global composition as well as A the trapezoidal area of the microchannel.

(ii) Influence of the total Reynolds number

We first have investigated the effect of the global Reynolds number on the mixing efficiency. Figure 9a and Figure 9b represent the evolution of the intensity of segregation calculated from the simulation for different average Reynolds numbers. Based on the simulation results, the I_m values never reached 0 within the channel whose the length is 6 mm, because a part of ethanol is pushed and trapped near the wall by CO₂ with a high velocity due to the channel geometry, as shown in Figure 8. The average Reynolds number and the hydraulic

diameter of microchannel are defined as:

$$Re = \frac{\overline{u_x} D_h \bar{\rho}}{\bar{\mu}} \quad (11)$$

where

$$D_h = \frac{4A}{P} \quad (12)$$

here $\bar{\mu}$ the viscosity of complete mixture calculated by the global composition, D_h the hydraulic diameter, estimated to be 200 μm for the microchannel and P the perimeter of the trapezoidal microchannel. It is not clear to determine the effect related to the Re number based on Figure 9a because of different mean velocities and the interest of introducing the time axis is to be able to compare with the characteristics mixing time published in the literature, which is discussed in the next part, so the length has been replaced by time for the axis, also resulting in the mixing index more comparable among different tests (Figure 9b). As expected, an improved mixing quality involves a higher global Reynolds number.

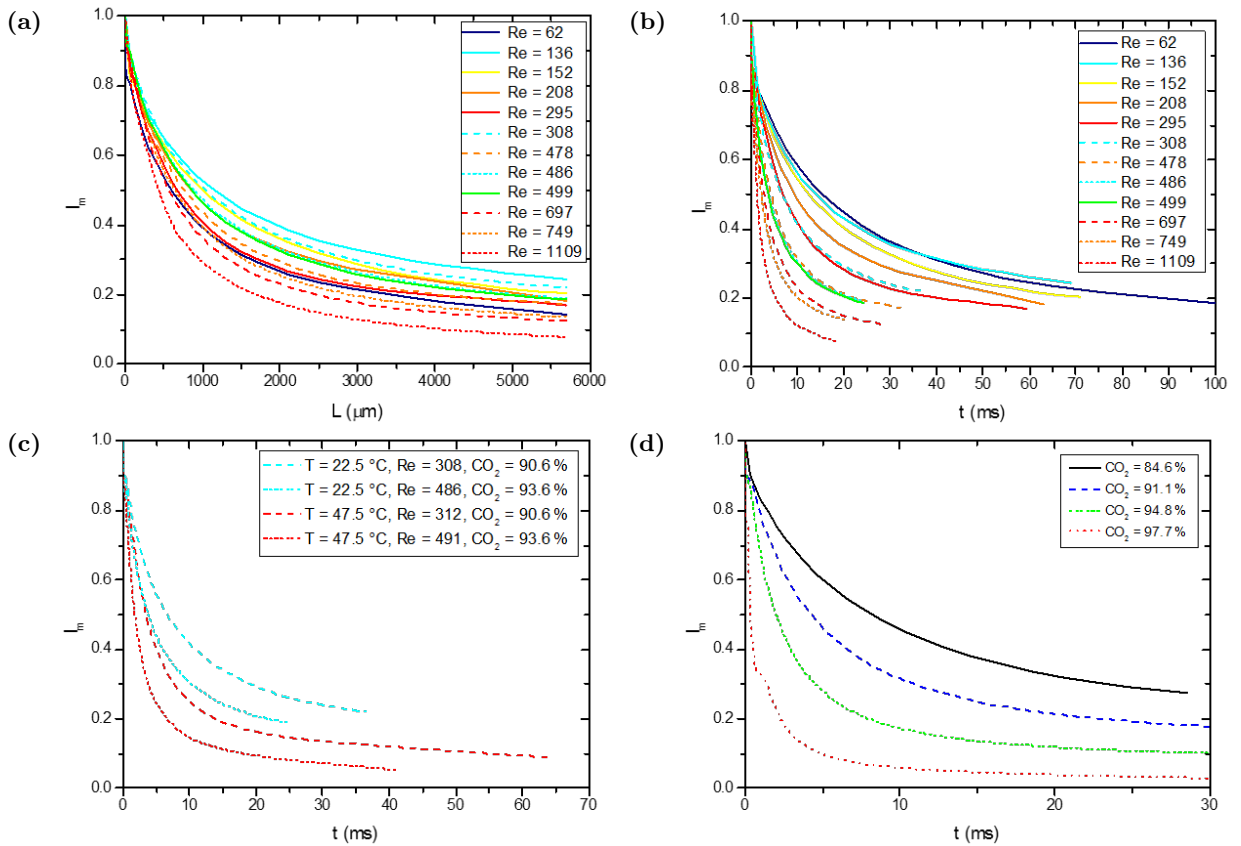


Figure 9: Mixing index / intensity of segregation I_m derived from simulation results: (a). I_m as a function of the distance from the injector tip in the microchannel (b). Reynolds number effect on I_m (c). Temperature effect on I_m with equivalent CO_2 /ethanol ratio and Reynolds number (d). CO_2 /ethanol ratio effect on I_m with fixed temperature at 38 °C and Reynolds number at 478.

(iii) Influence of the temperature

Then, the influences of the temperature have been also studied numerically and presented in Figure 9c. Near the critical temperature of the mixture CO_2 /ethanol, the temperature change has a strong effect on the mixture properties, such as density and viscosity and also on the molecular diffusion coefficient of ethanol in CO_2 . For a fixed CO_2 flow rate, a higher temperature in the channel leads to a much lower CO_2 density therefore a higher

velocity, according to the conservation of mass (Equation 13):

$$\rho_0 \cdot Q_0 = \rho_T \cdot Q \quad (13)$$

where ρ_0 is the fluid density in the cooling pump, Q_0 is the initial volumetric flow rate injected into the system by the pump, ρ_T is the fluid density in the microchannel at temperature T and Q is the real volumetric flow rate in the chip. Simulations have been performed to compare the evolution of the segregation intensity curve due to a temperature change, keeping an equivalent Reynolds number and fixed CO₂/ethanol ratio in the mixture. As a matter of fact, it is impossible to examine the effect of each parameter independently, because some parameters are related. For example, when the temperature varies, in order to keep the same Reynolds number and the same CO₂ fraction in the system, the injected flow rates must be changed so the velocity ratio changes. Indeed, at laminar flow conditions, the mass transfer is related not only to the shear stress, but also to the diffusion in the perpendicular plane to the flow direction. As observed in Figure 9c, a higher temperature results in a quicker decrease of the mixing index value so the mixing quality is improved. This can be attributed to two main effects: first, the augmentation of the diffusion coefficient, and then, the change of the CO₂ velocity due to a density decrease, leading to higher velocity ratio and consequently higher shear stress. For example, for the tests No.6 and No.14 in Table 1, the temperature increases from 22.5 °C to 47.5 °C, resulting in a rise of fluid velocity ratio (CO₂ initial velocity over ethanol initial velocity) from 2.77 to 5.29.

(iv) Influence of the CO₂/ethanol ratio

Eventually, we have investigated the effect of the CO₂/ethanol ratio on the mixing efficiency. Figure 9d illustrates that a higher ratio promotes the mixing quality. When this parameter increases, it results in a tremendous difference of two fluids' velocities. The strong shear stress at the capillary outlet due to an important CO₂/ethanol ratio actually leads the CO₂ flow towards the ethanol in the middle of the channel and improves mixing. The mean velocity of CO₂ is 15 times higher than the ethanol velocity in the injector for the case of CO₂ mass fraction of 97.7%, even creating recirculation vortices that improve significantly the mixing, shown in Figure 10 in which the velocity vectors are superimposed with the ethanol mass fraction field in the plane $y = 0$. The arrows in the figure represent only the velocity vector directions without magnitude. The same behavior has been captured in a numerical simulation for a two phase coflow in microfluidics (Zhang *et al.*, 2018). However, for other cases of CO₂ lower than 97.7%, no recirculations were detected.

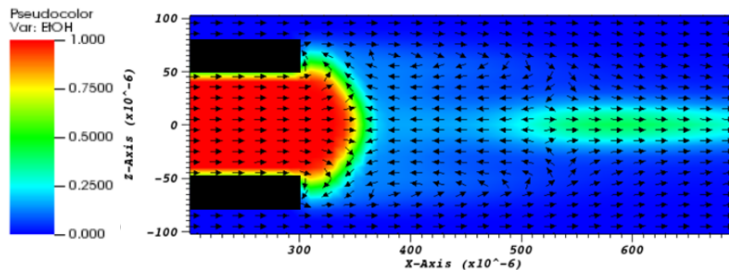


Figure 10: Recirculation detected in the mixing zone in the plane $y = 0$ for test case of which the CO₂ mass fraction is at 97.7%, with temperature at 38 °C and mean Reynolds number at 478.

We conclude that even if it is complicated to extract the influence of a single parameter over the mixing efficiency, we could identify some general trends based on the considered simulated cases. As expected, a higher

global Reynolds number results in a better mixing quality. An increase of the CO₂/ethanol ratio enhances mass transfer in the mixture by increasing the shear stress generated by an important difference between the inner and the outer velocities at the injector outlet. This shear stress results in some cases in the creation of vortices near the tip of the nozzle, which largely contribute in decreasing the mixing time. According to the simulation results, a temperature increase improves the mixing because of both increased diffusion and shear stress. In general, a higher Reynolds number, an increased temperature and a strong CO₂/ethanol ratio are recommended to accelerate fluid mixing.

4.5 Characteristic mixing time

Another important parameter for characterizing the mixing efficiency is the mixing time. A theoretical way to define the mixing time was proposed by Baldyga and Bourne (1986). It is defined as the time required to obtain a homogeneous mixture in a slab, considering both diffusion and advection mixing, also known as the stretching efficiency model (Equation 14). This equation has been largely applied to determine micromixer efficiency (Baldyga and Bourne, 1984; Falk and Commenge, 2010; Ghanem *et al.*, 2014):

$$t_m = \frac{1}{\sqrt{2}} \sqrt{\frac{\nu}{\epsilon}} \ln \frac{1.52 \cdot L \cdot u}{D} \quad (14)$$

with ν the kinematic viscosity, ϵ the energy dissipation rate, L the characteristic length, u the fluid velocity and D the diffusion coefficient.

In this study, we have chosen to define the characteristic mixing time as the time constant of a first order system, which in here is the evolution of the segregation index as a function of the time. As illustrated in Figure 11, this time constant is defined as the intersection point of the time axis and the tangent of the origin of the mixing index curve.

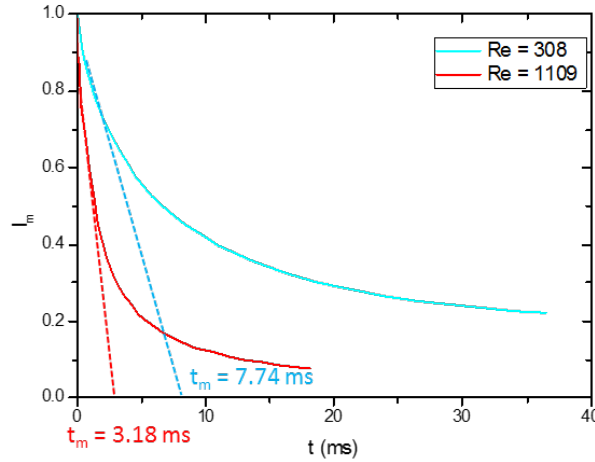


Figure 11: Illustration of the mixing time constant determination for 2 different cases: the red one corresponds to the temperature of 47.5 °C and the CO₂ fraction of 93.7 wt.% (No.12); the blue one corresponds to the temperature of 22.5 °C and the CO₂ fraction of 90.6 wt.% (No.6).

Similar method of tangent construction has been adopted to define a characteristic mixing length from supersaturation curve (Metzger, 2017). This mixing time definition reveals the decline rate of index value at the beginning of the mixing process, which is critical in SAS. Indeed, a mixing as quick as possible is desired

once the two fluids contact each other, in order to create high supersaturation and to promote nucleation.

Since the theoretical mixing times in microchannels of different diameters published previously have been collected and correlated as a function of the energy dissipation rate ϵ (Falk and Commenge, 2010), the correlation was plotted and compared to our results in Figure 12. The energy dissipation rate for the laminar flow conditions employed in the present study has been determined by Equation 15:

$$\epsilon = \frac{26.12 \cdot \nu \cdot \overline{u_x}^2}{D_h^2} \quad (15)$$

The coefficient 26.12 is recalculated, with details in appendix, according to the trapezoidal cross-section area of the microchannel (Bahrami *et al.*, 2005).

The characteristic mixing times in this study (red points framed by two dotted lines illustrating a 50% relative error) are in agreement with the theoretical behavior shown as the blue line in Figure 12 (Falk and Commenge, 2010). Therefore, it seems that the determination of the mixing time in this study is an appropriate method to evaluate the mixing performance under laminar flow conditions for microfluidic coflow.

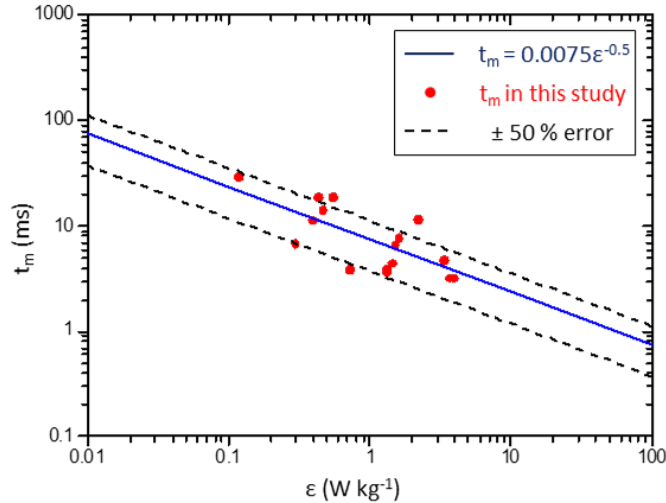


Figure 12: Comparison of characteristic time constant obtained in this study (red points) with the data published (blue line) in the literature (Falk and Commenge, 2010).

5 Conclusion

In this work, the CO₂ and ethanol mixture has been studied by both experimental and numerical approaches at microfluidic scale in laminar flow conditions. The mixing was studied under SAS process conditions, which means the two fluids are completely miscible in a single phase at 100 bar in a controlled temperature isothermal environment. By applying a μ PIV characterization, the fluid velocity was experimentally measured in a microchip considering a coflow configuration. A 3D CFD simulation has been validated by comparing the experimental and numerical results. Different average Reynolds numbers, the temperature between 20 and 50 °C, as well as the ratio of CO₂ to ethanol, have been investigated to capture their effects on mixing, which is characterized by the intensity of segregation or the mixing index I_m . It has been discussed that these mixing parameters are not independent so that the effect of only one variable cannot be determined without any other

condition changes. We have obtained some tendencies of the conditions to achieve a fast mixing in the microchannel. Finally, the mixing performance has been related to the characteristic mixing time constant, which is a simple criterion to examine the mixing efficiency in the micromixer. A first order system has been applied to estimate this mixing time constant according to the evolution of the mixing index curve as a function of time and this characteristic time has been well related to the energy dissipation rate ϵ of laminar flows. After comparing the mixing time constant of this study to the theoretical ones previously published, the results are reasonable and reliable. Therefore the method used in this paper is a robust way to study the laminar mixing of two different fluids under microfluidic SAS conditions.

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Appendix

Determination of the energy dissipation rate coefficient in a trapezoidal channel

For a laminar flow, the energy dissipation rate can be expressed based on Equation 16:

$$\epsilon = \frac{Q\Delta p}{\rho V} = \frac{Q}{\rho A} \cdot \frac{\Delta p}{L} \quad (16)$$

with p the pressure, Q the flow rate, ρ the fluid density, V the container volume, A the cross-section surface and L the length of channel.

According to the Darcy-Weisbach equation (Equation 17) and the Darcy friction factor f_D in a circular tube Equation 18, one can easily obtain Equation 19:

$$\frac{\Delta p}{L} = f_D \cdot \frac{\rho}{2} \cdot \frac{\bar{u}^2}{D_h} \quad (17)$$

$$f_D = \frac{64}{Re} \quad (18)$$

$$\frac{\Delta p}{L} = \frac{128\mu Q}{\pi D_h^4} \quad (19)$$

with \bar{u} the mean velocity, D_h the hydraulic diameter, Re the Reynolds number, μ the fluid viscosity.

By combining Equation 16 and Equation 19, the energy dissipation rate ϵ is finally presented in Equation 20 for laminar flow in a circular tube:

$$\epsilon = \frac{32\nu(\bar{u})^2}{D_h^2} \quad (20)$$

Above ϵ estimation is deducted only for fluid flow in a round cross-section conduit. Whereas, the coefficient 32 should be different for a different channel geometry. A similar correlation of Equation 18 is proposed for a trapezoidal shape channel, given Equation 21 (Bahrami *et al.*, 2005):

$$f_F Re_{\sqrt{A}} = \frac{8\pi^2(3\omega^2 + 1) + \beta(1 - 3\omega^2)}{9\sqrt{\omega}(\omega + \sqrt{\omega^2 - \beta\omega^2 + 1})} \quad (21)$$

in which f_F is the Fanning friction factor and the $Re_{\sqrt{A}}$ is the Reynolds number with the square root of cross-sectional channel area as the characteristic length instead of the hydraulic diameter.

The parameters ω and β are determined by the trapeze shape, shown in Figure 13.

$$\omega = \frac{a + b}{2h} \quad (22)$$

$$\beta = \frac{4ab}{(a + b)^2} \quad (23)$$

As known that the Darcy friction factor is 4 times of the Fanning friction factor, the former one can be

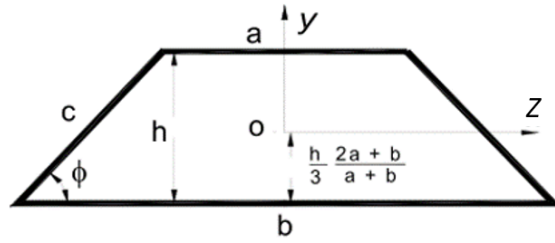


Figure 13: Schematic of trapeze cross-section geometry, extracted and modified from a published work (Bahrami *et al.*, 2005).

calculated for a trapezoidal channel from Equation 24. By comparing Equation 18 and Equation 20 for the circular channel, the coefficient 26.12 is eventually derived for Equation 15.

$$f_D = \frac{52.24}{Re} \quad (24)$$