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Enhanced stability and bangap tuning of α-$[HC(NH_2)_2]PbI_3$ Hybrid Perovskite by Large Cation Integration.

Antonin Leblanc,[a] Nicolas Mercier,*[a] Magali Allain,[a] Jens Dittmer,[b] Thierry Pauporté,[c] Vincent Fernandez,[d] Florent Boucher,[d] Mikael Kepenekian,[e] Claudine Katan*[e]

[a] MOLTECH-ANJOU, UMR-CNRS 6200, Université d'Angers, 2 Bd Lavoisier, 49045 Angers, France.

[b] Institut des Molécules et Matériaux du Mans, CNRS UMR 6283, Le Mans Université, Avenue Olivier Messiaen, 72085 Le Mans cedex 9, France

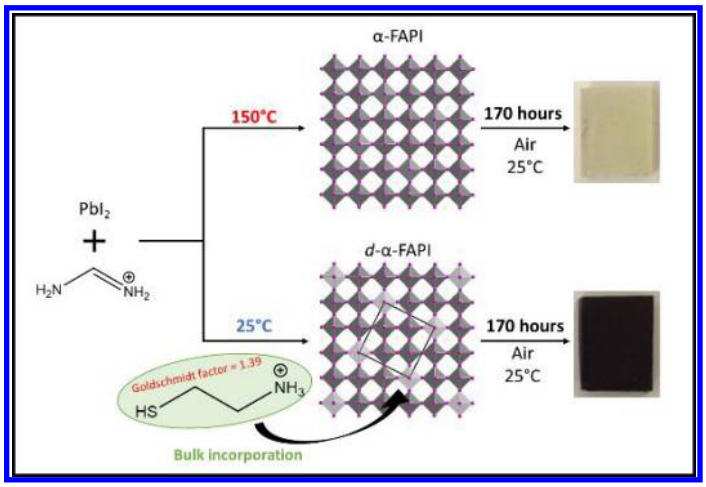
[c] Chimie ParisTech, PSL Research University, CNRS, Institut de Recherche de Chimie Paris (IRCP), 11 rue P. et M. Curie, F-75005 Paris, France.

[d] Institut des Matériaux Jean Rouxel, UMR-CNRS 6502, Université de Nantes, 2 rue de la Houssinière, BP 32229, 44322 Nantes Cedex 3, France

[e] Univ Rennes, ENSCR, INSA Rennes, CNRS, ISCR (Institut des Sciences Chimiques de Rennes) - UMR 6226, F-35000 Rennes, France

Keywords: halogenated perovskite, Goldschmidt factor, iodoplumbate, formamidinium, alpha-FAPI, lead deficient perovskite, 3D hybrid perovskite

ABSTRACT



We report room temperature synthesis of lead and iodide deficient α-$[HC(NH_2)_2]PbI_3$ perovskites (abbreviated *d*-α-FAPI, FA^+ = formamidinium), owning the general formula $(A',FA)_{1+x}[Pb_{1-x}I_{3-x}]$ (with A' = hydroxyethylammonium (HEA^+) or thioethylammonium

(TEA^+) cations, $0.04 \leq x \leq 0.15$). These materials retain a 3D character of their perovskite network despite incorporation of large HEA^+ or TEA^+ cations, demonstrating that the Goldschmidt tolerance factor can be bypassed. We found that thin films of $(TEA,FA)_{1+x}[Pb_{1-x}I_{3-x}]$ (x= 0.04 and 0.13) show exceptional α-phase stability under ambient conditions, one order of magnitude higher compared to α-FAPI and α-$(Cs,FA)PbI_3$ thin films. *d*-α-FAPI phases are shown to maintain a direct bandgap, which increases monotonously for *x* ranging from 0 up to 0.20, with characteristics of a p-type semiconductor for low concentrations of vacancies ($x \leq 0.13$) and n-type for larger ones. They offer an alternative to reach methylammonium- and bromine-free stable α-FAPI type phase, and open new avenues in the field of perovskite solar cells, up to bandgap tuning desirable for tandem solar cells.

INTRODUCTION

Since 2012, perovskite based solar cells (PSCs) have shown an impressive potential, combining advantages of low-temperature thin film processing (cost process less than half of the price of c-Si solar cell technology),[1,2] a high power conversion efficiency (PCE) certified up to 23.7%,[3,4] and the possibility of recycling the solar cells.[5,6] However, the well-known $(CH_3NH_3)PbI_3$ hybrid perovskite (abbreviated MAPI, with MA^+ = methylammonium) firstly used in PSC has shown a fast degradation under air exposure that hinders market development.[7] Consequently, a chemical approach has been applied consisting in substituting the MA^+ cation with a less hydrophilic one, presenting a Goldschmidt tolerance factor in the 0.8-1.0 range suitable for the perovskite phase.[8] Formamidinium (FA^+) has been considered as a good candidate since the Goldschmidt factor of $FAPbI_3$ (called FAPI), is close to 1.0.[8] Unfortunately, the perovskite α-FAPI

phase, which possesses a better thermal stability and a more suitable bandgap than MAPI,[9] is only synthesized and is thermodynamically stable beyond 150°C. When cooling down to room temperature under ambient atmosphere, α-FAPI rapidly transforms into δ-FAPI, an undesirable polymorph phase that does not present a perovskite architecture nor photoactivity.[10] In order to slow down this α-to-δ phase transformation, a double cation strategy was first developed. This chemical approach consists of the α-phase stabilisation by mixing two materials possessing a large and a small Goldschmidt tolerance factors, leading to an alloy which owns a balanced factor comprised between the suitable range [0.8 - 1.0]. The resulting two archetype compounds are $Cs_{0.15}FA_{0.85}PbI_3$ and $MA_{0.20}FA_{0.80}PbI_3$.[11,12] This strategy was successful as the α-to-δ phase transformation was observed after a longer time (air condition) than with α-FAPI, and improved efficiencies for solar cells were also achieved. Subsequently, other combinations of multiple cations and mixed halide were employed, leading to the most stable and efficient alloys α-$(Rb/Cs/MA)_{0.20\text{-}0.25}FA_{0.75\text{-}0.80}Pb(I_{0.85}Br_{0.15})_3$, regularly exceeding 20% solar cells.[13-20] However, despite the beneficial effects of the presence of MA^+ and Br^- ions on the α-phase purity, crystallinity, stability and on solar cell efficiencies, both volatile MA^+ cation and Br^- halide induce thermal degradation and an unsuitable blue-shift of the bandgap.[21] Moreover, from a chemical point of view, this multi-cations alloying strategy is still sorely limited by the choice of the cations that possess a suitable size to keep a 3D perovskite architecture.

Recently, we have reported unprecedented lead and iodide deficient hybrid perovskite materials (*d*-HP family), owning a general formulation $(A')_{3.48x}(A)_{1\text{-}2.48x}[Pb_{1\text{-}x}I_{3\text{-}x}]$, where A= MA^+ and A'= HEA^+ (hydroxyethylammonium $HO\text{-}(CH_2)_2\text{-}NH_3^+$).[22] The resulting

phases, named *d*-MAPI, present a 3D perovskite corner-shared $Pb_{1-x}I_{3-x}$ (x = 0 to 0.20) network, which can be described as a lead and iodide deficient PbI_3 MAPI network. A slightly improved air stability has been demonstrated for this *d*-MAPI phases as compared to MAPI.[22] More recently, Diau *et al*. have prepared the tin analogue (*d*-(MA)SnI_3) which also shows good air stability.[23] This superior air stability was also observed for the *hollow* hybrid perovskites named (A',A)MI_3 (A'=en^{2+} = ethylenediammonium, A = MA^+/FA^+, M = Pb^{2+}/Sn^{2+}) reported by the Kanatzidis'group.[24] When A= FA^+, a series of *hollow* perovskites are obtained: $(FA)_{1-x}(en)_xPb_{0.7x}I_{3-0.4x}$. However, they are synthesized at temperatures over 100 °C, as the α-FAPI phase, and no thin film preparation (and so no thin film stability test) of these *hollow*-α-FAPI phases has been reported. [24]

In this study, we report the room temperature synthesis of a family of lead and iodide deficient α-FAPI (*d*-α-FAPI) hybrid perovskites. These 3D perovskites incorporate the formamidinium cation and a larger size cation, either the hydroxyethylammonium (HEA^+, *d*-α-FAPI-$\mathbf{H_x}$ compounds) or the thioethylammonium (TEA^+, *d*-α-FAPI-$\mathbf{T_x}$ compounds) monocation. The general formula of these materials is $(FA,A')_{1+x}[Pb_{1-x}I_{3-x}]$ with A'= HEA^+ or TEA^+. Crystals of *d*-α-FAPI have been grown, allowing the selection and structure determination of one composition, $(HEA)_{0.406}(FA)_{0.698}[Pb_{0.896}I_{2.896}]$ (x = 0.104, written *d*-α-FAPI-$\mathbf{H_{0.104}}$), which revealed the 3D nature of the perovskite network. Several materials have been also prepared as pure crystallized powder phases, two *d*-α-FAPI-$\mathbf{H_x}$ (with *x* = 0.13 and *x* = 0.15, here abbreviated $\mathbf{H_{0.13}}$, and $\mathbf{H_{0.15}}$), and one *d*-α-FAPI-$\mathbf{T_x}$ (x = 0.04, abbreviated $\mathbf{T_{0.04}}$), as well as thin films ($\mathbf{H_{0.13}}$, $\mathbf{H_{0.15}}$, $\mathbf{T_{0.04}}$ and also $\mathbf{T_{0.13}}$ (x = 0.13)). Powders and thin films have been fully characterized by X-ray diffraction, NMR (liquid and solid state), energy dispersive X-ray spectroscopy (EDX), elemental analysis, UV-

visible absorption and X-ray photoelectron spectroscopy (XPS) measurements. Moreover, the analysis has been completed by computational DFT investigations. Besides the fact that these α-FAPI type phases can be synthesized at room temperature, we show that the stability of the α phase at room temperature and under ambient condition (thin films, case of *d*-α-FAPI-$\mathbf{T_x}$ materials) is exceptionally improved compared to both α-FAPI and α-$Cs_{0.15}FA_{0.85}PbI_3$ reference materials.

RESULTS AND DISCUSSION

Crystals of *d*-α-FAPI have been obtained at room temperature by slow liquid-gas diffusion process using aqueous hydroiodic acid and ethanol as solvent and anti-solvent, respectively (more details in Supporting Information -I.a-). These crystals have a cross shape (Figure **1a**), and depending on experimental conditions they can be black or dark red. It is also worth noting that no yellow crystals of δ-FAPI were co-crystallized. For all *x* values , single-crystal X-ray diffraction analysis of *d*-α-FAPI-$\mathbf{H_x}$ has revealed room temperature tetragonal unit cells (a = b = 14.471 Å, c = 6.448 Å, α = β = γ = 90°, V = 1350.2 Å^3), slightly larger than the previously reported *d*-MAPI unit cell (also obtained at room temperature: a = b = 14.312 Å, c = 6.366 Å, α = β = γ = 90°, V = 1303.9 Å^3).[22] This can be correlated to the effective radius of FA^+ (253 pm) which is larger than the one of MA^+ (217 pm).[25] The structure of the *d*-α-FAPI-$\mathbf{H_x}$ was studied from one single crystal selected among many other. The *d*-α-FAPI-$\mathbf{H_x}$ phase crystallises at room temperature into the P4/m space group (see S.I. IV, Table S7). Despite several attempts at different temperatures (100-293K range), the X-ray study has only allowed the determination of the inorganic network. This can be partly explained at room temperature by a dynamical disorder of both HEA^+

and FA^+ organic cations, as revealed by solid state NMR experiments (see below). At lower temperature, this may be explained by a statistical disorder. Structural refinements finally led to a $Pb_{4.48}I_{14.48}$ formula per unit cell for the inorganic part. The structure can be described from a main 3D Pb_4I_{14} perovskite network of corner-sharing PbI_6 octahedra with a 3D arrangement which generates channels along the c axis (Figure **1b**). These channels are half filled with extra Pb^{2+} and I^- ions, located in (0, 0, 0) and (0, 0, 0.5), respectively. To highlight the lead and iodide deficiency of this 3D perovskite network, the $Pb_{4.48}I_{14.48}$ formula should be written as $Pb_{0.896}I_{2.896}$ or $Pb_{1-x}I_{3-x}$ with $x = 0.104$ (*d*-α-FAPI-$\mathbf{H_{0.104}}$). The general formulation of the *d*-α-FAPI-$\mathbf{H_x}$ phases is $(HEA)_{3.9x}(FA)_{1-2.9x}[Pb_{1-x}I_{3-x}]$ or $(HEA,FA)_{1+x}[Pb_{1-x}I_{3-x}]$ according to different chemical analyses of crystals and powders (see below and S.I.) and therefore the formula of this studied crystal is $(HEA)_{0.406}(FA)_{0.698}[Pb_{0.896}I_{2.896}]$. Thus, the lead and iodide deficient *d*-α-FAPI phase can be described from the α-FAPI phase through a substitution of x $(PbI)^+$ units by x organic cations, HEA^+ or FA^+. In the case of $x = 0.2$ ($Pb_{0.8}I_{2.8}$ or Pb_4I_{14} network), all Pb^{2+} and I^- ions along the *c* direction have been substituted by organic cations leading to empty channels. When $0 \leq x \leq 0.2$ (case of the studied crystal *d*-α-FAPI-$\mathbf{H_{0.104}}$), the X-ray analysis shows that channels contain Pb^{2+} and I^- anions as well as organic cations. Consequently, one hypothesis is a partial filling of channels by Pb^{2+} and I^-. Another hypothesis is that intermediate values of x may correspond to a limited number of empty channels in the structure (see below the model structures used for calculations). Finally, it must be noted that the poor quality of crystals of *d*-α-FAPI-$\mathbf{T_x}$ precluded their X-ray single crystal analysis.

From the crystal structure analysis of the anhydrous iodide salts of hydroxyethylammonium (HEAI) and thioethylammonium (TEAI) which have revealed that HEA^+ and TEA^+ cations adopt a curving shape (S.I. Tables S8 and S9), we estimated the effective radius of HEA^+ (3.44 Å) and TEA^+ (4.14 Å) according to the model proposed by Kieslich *et al*.[25] Thus, the Goldschmidt factors of the hypothetical $(HEA)PbI_3$ and $(TEA)PbI_3$ perovskites were calculated as 1.24 and 1.39, respectively. These values are far from the expected 0.8-1.0 range usually necessary to stabilize the 3D ABX_3 perovskite.[8] However, the 3D perovskite structures of α-FAPI type prepared in this work (*d*-α-FAPI phases) have been stabilized despite the incorporation of these HEA^+ / TEA^+ cations (see below). This is explained by the presence of $(PbI)^+$ vacancies in the 3D network. This family of compounds (*d*-HP) shows that it is possible to circumvent the geometrical Goldschmidt factor and to obtain 3D hybrid perovskites even using larger $R\text{-}NH_3^+$ organic cations. Moreover, the integration of HEA^+/TEA^+ cation allows the synthesis at room temperature of the α-FAPI type network (*d*-α-FAPI) while α-FAPI reference is prepared only at temperature over 150 °C.

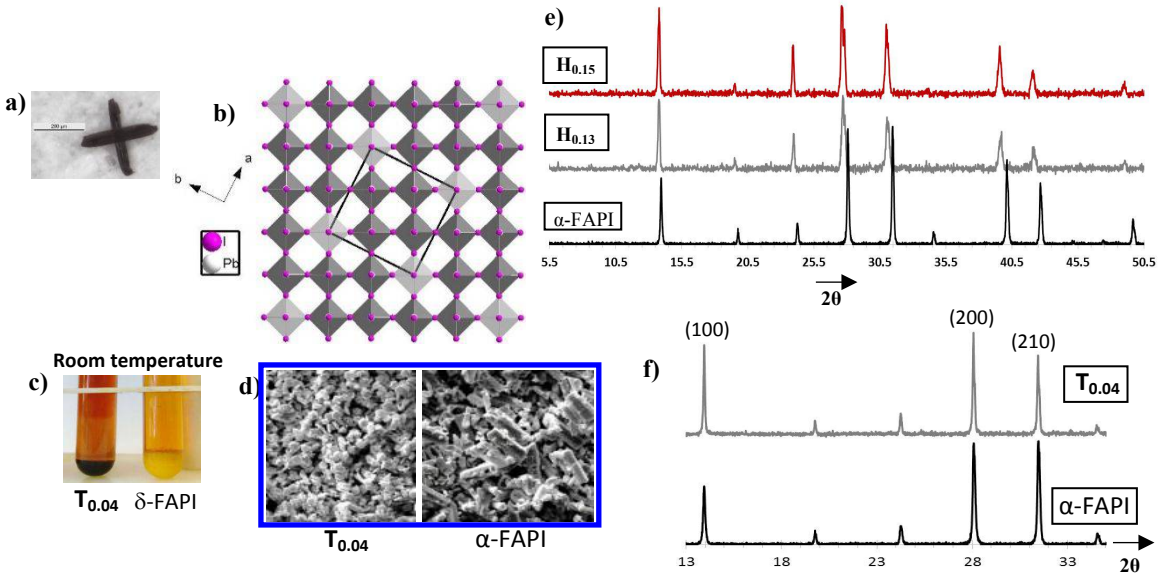


Figure 1. a) *Cross-shaped crystals of d-α-FAPI.* ***b)*** *View along the c axis of the tetragonal unit cell (black square) of the structure of* $(HEA,FA)_{1.104}[Pb_{0.896}\,I_{2.896}]$ *showing the lead and iodide deficient perovskite network (dark grey octahedra : full* Pb^{2+} *and* I^{-} *sites occupancy, light grey octahedra: half* Pb^{2+} *and* I^{-} *sites occupancy).* ***c)*** *Photography of* $\boldsymbol{T_{0.04}}$ *and δ-FAPI powders synthesized at room temperature;* ***d)*** *SEM images of crystallized powders of* $\boldsymbol{T_{0.04}}$ *and α-FAPI;* ***e)*** *XRD patterns of* $\boldsymbol{H_{0.13}}$, $\boldsymbol{H_{0.15}}$ *and α-FAPI powders showing the shift of lines towards low angles for* $\boldsymbol{H_{0.13}}$ *and* $\boldsymbol{H_{0.15}}$; ***f)*** *XRD patterns of* $\boldsymbol{T_{0.04}}$ *and α-FAPI powders highlighting the (100), (200) and (210) diffraction lines.*

Powders of *d*-α-FAPI-$\mathbf{H_x}$ and *d*-α-FAPI-$\mathbf{T_x}$ were prepared by a fast precipitation method at room temperature. Increasing amounts of hydroxyethylamine/FAI/PbI_2 or thioethylamine/FAI/PbI_2 were dissolved at room temperature into aqueous hydroiodic acid under ultrasonication until the saturation was reached. Then, the saturated solution was

poured into ethyl acetate, immediately leading to the precipitation of a powder (more details in S.I. I.b). Pure *d*-α-FAPI-$\mathbf{H_x}$ phases with x = 0.13 ($\mathbf{H_{0.13}}$) and x = 0.15 ($\mathbf{H_{0.15}}$), as well as one pure phase of *d*-α-FAPI-$\mathbf{T_x}$ with x = 0.04 ($\mathbf{T_{0.04}}$), were obtained and fully characterized. We also prepared the δ-FAPI by starting from PbI_2 and FAI as reagents, and also synthesized α-FAPI compound by heating the δ-FAPI powder at 155°C during 30 min. A color change then occurs from yellow to black characterizing the δ-FAPI to α-FAPI phase transition. Despite our best efforts to precipitate pure phases with compositions in the x = 0-0.20 range, only one composition could be obtained for *d*-α-FAPI-$\mathbf{T_x}$ ($\mathbf{T_{0.04}}$) while two distinct compositions ($\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$) where obtained with HEA^+. In all other experiments, the yellow δ-FAPI was crystallized as impurity. Figure **1c** (zoomed Figure S7) shows a photography of δ-FAPI and $\mathbf{T_{0.04}}$ precipitated powders. δ-FAPI powder has a typical yellow color while the $\mathbf{T_{0.04}}$ powder is black, showing that another phase is obtained when a small amount of TEA^+ is incorporated in the system, which will be proved to be a α-FAPI type phase (*d*-α-FAPI). Furthermore, it is worth noting that the morphology of α-FAPI and $\mathbf{T_{0.04}}$ micrometer-size crystals (obtained by the precipitation method) are very different, as shown by scanning electron microscopy images (Figure **1d**, zoomed Figure S8). Indeed, the elongated shape of α-FAPI crystals is the consequence of the preferential growth along the PbI_3^- chain direction of the initial δ-FAPI compound,[10] while a cube shape is observed for $\mathbf{T_{0.04}}$ crystals proving that the cubic α-FAPI type network was directly obtained. $\mathbf{H_{0.13}}$, $\mathbf{H_{0.15}}$ and $\mathbf{T_{0.04}}$ samples have been fully characterized, first by energy dispersive X-ray spectroscopy (EDX) allowing the determination of the I/Pb atomic ratios. Considering the inorganic network $[Pb_{1-x}I_{3-x}]$ as found in the *d*-MAPI phases $(HEA,MA)_{1+x}[Pb_{1-x}I_{3-x}]$,[22] and in this studied crystal $\mathbf{H_{0.104}}$ ($Pb_{0.896}I_{2.896}$ network, x= 0.104),

the measured I/Pb atomic ratios gave $x = 0.13$ for $\mathbf{H_{0.13}}$, $x = 0.15$ for $\mathbf{H_{0.15}}$, and $x = 0.04$ for $\mathbf{T_{0.04}}$ (Figures S1, S8, and Tables S3, S6). Moreover, ^{1}H nuclear magnetic resonance spectroscopy (NMR) in solution led to determine the following HEA^+/FA^+ and TEA^+/FA^+ molecular ratios: $HEA^+/FA^+ = 0.81$ for $\mathbf{H_{0.13}}$ and 1.13 for $\mathbf{H_{0.15}}$, and $TEA^+/FA^+ = 0.13$ for $\mathbf{T_{0.04}}$ (Figures S2-S6, S9-11). The following formulations were thus deduced: $(HEA)_{0.51}(FA)_{0.62}[Pb_{0.87}I_{2.87}]$ ($\mathbf{H_{0.13}}$) and $(HEA)_{0.60}(FA)_{0.55}[Pb_{0.85}I_{2.85}]$ ($\mathbf{H_{0.15}}$), consistent with the general formulation of $(HEA)_{3.9x}(FA)_{1-2.9x}[Pb_{1-x}I_{3-x}]$, with $x = 0.13$ and $x = 0.15$; and $(TEA)_{0.12}(FA)_{0.92}[Pb_{0.96}I_{2.96}]$ ($\mathbf{T_{0.04}}$, $x = 0.04$). All these compositions were finally confirmed by elemental analysis (Tables S1, S2, S4, and S5). The purity of all these crystallized powders was also verified by powder X-ray diffraction (XRD) analysis. The XRD patterns of the yellow δ-FAPI and the black α-FAPI powders (Figures S12 and S13) perfectly fit with their theoretical XRD patterns, calculated from both known δ-FAPI and α-FAPI structures (δ-FAPI: hexagonal $P6_3mc$, a = b = 8.7 Å, c = 7.9 Å; α-FAPI: cubic Pm-3m, a = 6.362 Å for).[10,26]

As regards $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ powders, their XRD patterns are in good agreement with the calculated XRD pattern from the single crystal data of $(HEA,FA)_{1.104}[Pb_{0.896}I_{2.896}]$ (Figure S14). In Figure **1e**, diffractograms of $\mathbf{H_{0.13}}$, $\mathbf{H_{0.15}}$ and α-FAPI powders are provided. Compared to α-FAPI, a slight shift of most peaks of $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ towards lower angles is observed (see also zoom in Figure S15). This phenomenon, revealing a slight increase of unit cell parameters through the substitution process of $(PbI)^+$ units by organic cations, has been already observed in the case of *d*-MAPI phases $(HEA,MA)_{1+x}[Pb_{1-x}I_{3-x}]$.[22] Three peaks of low intensity which were expected at low 2θ angles (6.2°, 8.7°, 12.3°) are absent in the XRD patterns of $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ (Figure **1e** and Figure S14 -theoretical XRD pattern-

). These lines are assigned to (1 0 0), (1 1 0) and (2 0 0) planes of the tetragonal unit cell (a = 14.312 Å, c = 6.366 Å) typical of the *d*-HP phase[22] (see Figure **1b**). The fact that these lines are not observed could be explained by a preferential orientation phenomenon of (0 0 l) type, but most probably the smaller unit cell observed is resulting of a non-ordered distribution of defaults (lead and iodide vacancies) as observed in *hollow* structures based on ethylenediammonium cations.[24] Unlike the *d*-α-FAPI-$\mathbf{H_x}$ phases, the XRD pattern of $\mathbf{T_{0.04}}$ does not present significant shifts compared to the α-FAPI XRD one (Figures **1f**, S16). This may be explained by the low incorporation of TEA^+ cations into this $\mathbf{T_{0.04}}$ phase (x= 0.04, TEA/FA= 0.13). Nevertheless, we can notice differences for several diffraction peak intensities. Thus, the intensities I(100)/I(200) ratio ((100) line at 13.9° and (200) line at 28.0°) appears higher in $\mathbf{T_{0.04}}$ than in α-FAPI, and this cannot be explained by a preferential orientation phenomenon. We also notice that I(210) ((210) line at 31.4°) is lowered in $\mathbf{T_{0.04}}$ compared to α-FAPI. This feature is probably the consequence of a slight structural modification involved by the incorporation of TEA^+ in $\mathbf{T_{0.04}}$. Moreover, this phenomenon is also observed when materials ($\mathbf{T_{0.04}}$, $\mathbf{T_{0.13}}$, and α-FAPI) are prepared as thin films (see below).

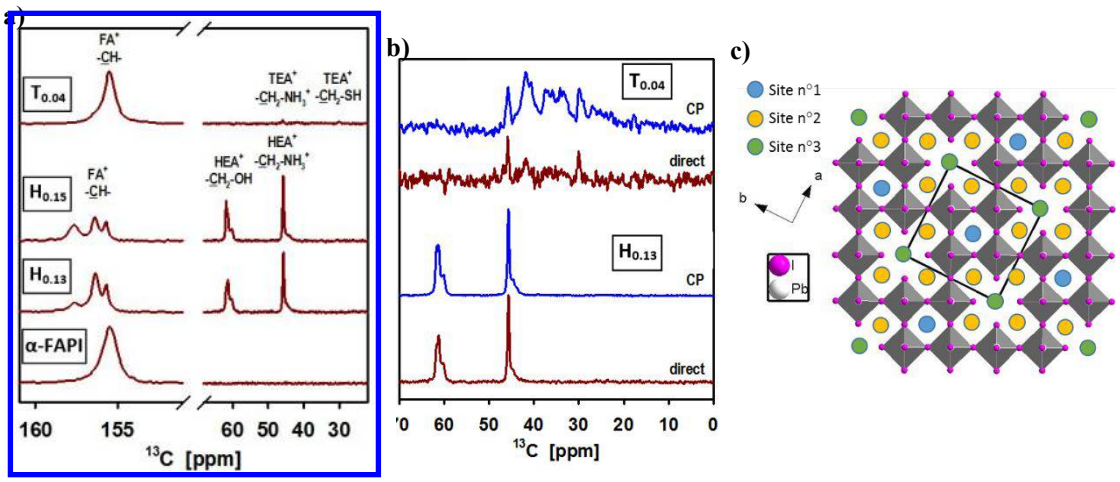


Figure 2. a) *Direct excitation solid state NMR spectra of α-FAPI,* $\boldsymbol{H_{0.13}}$*,* $\boldsymbol{H_{0.15}}$*, and* $\boldsymbol{T_{0.04}}$*, scaled according to the number of scans.* ***b)*** *The upfield region of the direct excitation spectra (red) of* $\boldsymbol{H_{0.13}}$ *and* $\boldsymbol{T_{0.04}}$ *compared to the corresponding cross polarization spectra (blue). Arbitrary scale.* ***c)*** *View along the c axis of the tetragonal unit cell (black square) of a hypothetical structure of d-α-FAPI with x = 0.20 (resulting formula of (HEA or TEA, FA)*$_{1.2}$*[Pb*$_{0.8}$ *I*$_{2.8}$*]). The organic cations are located on the three different crystallographic site n°1, n°2 and n°3 (blue, yellow and green circles, respectively).*

^{13}C solid state NMR was used to characterize the organic cations, invisible to XRD, verify their incorporation into the perovskite framework and gain information on their dynamics. While the direct excitation ^{13}C spectrum of α-FAPI shows the expected single signal of the FA^+ carbon (at 155.5 ppm), the materials $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ show two additional signals around 61 and 45 ppm, corresponding to $–CH_2–OH$ and of $–CH_2–NH_3^+$ of the HEA^+ cation, respectively (Figure **2a**). Concerning the FA^+ signal, it shifts upon the insertion of the other cations, similar to MA^+ in *d*-MAPI phases,[22] but also splits into three signals, at identical

positions for $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$. These observations are consistent with the possible occupation of different sites in this lead and iodide deficient perovskite network: three sites for the FA^+ ions, and probably only the sites n°2 and n°3 for HEA^+ ions (see below Figure **2c**). The least shifted (0.2 compared to 0.9 and 2.2 ppm) signal is here most likely from FA^+ on the site corresponding to the unmodified FAPI structure (site n°1). There is also dispersion in the HEA^+ signals: both signals in both samples consist of a dominant peak on the downfield side, interpreted as stemming from HEA^+ on the more abundant site n°2 , and a broader foot upfield, attributed to HEA^+ on site n°3. The chemical shift range of the latter contains the shift of HEAI, because the local environment of the carbons on site n°3 is closer to those in HEAI (Figure S22). The HEA^+ signals show substructures, in particular the foot of $–CH_2–NH_3^+$ on site n°3, indicating that these cations are not all incorporated in the same way or that there are preferential orientations. In fact, ^{13}C cross-polarization (CP) spectra of $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ exhibit intense signals compared to α-FAPI showing that the free rotation of the cations is restrained by the structure (Figure S22). This counts in particular for the bigger HEA^+ cations. It is thus imaginable that there are preferred orientations that have lifetimes being long enough that the NMR signals are not fully averaged.

For $\mathbf{T_{0.04}}$, the CH_2 ^{13}C signals of TEA^+ are hardly detectable by a direct excitation experiment. Their shifts are 45.9 and 30.3 ppm, respectively, different from those of TEAI (Figure S22). The ^{13}C CP experiment reveals a broad distribution of additional signals in the CH_2 region (Figure **2b**). Note that this experiment is not quantitative and enhances signals stemming from less mobile molecules or molecule groups. In the direct excitation experiment their intensity is so low that they are largely covered by the noise. Nevertheless, as they are spread over a broader range, they probably constitute the major part of TEA^+.

(This finding is in contrast to $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$, where the shapes of the signals acquired with CP and direct excitation are essentially identical, indicating that the dynamics of all HEA^+ is similar (Figure **2b**, $\mathbf{H_{0.13}}$).) Incorporated TEA^+ thus presents in the majority much less rotational mobility compared to the other organic cations (FA^+, HEA^+ in $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$).

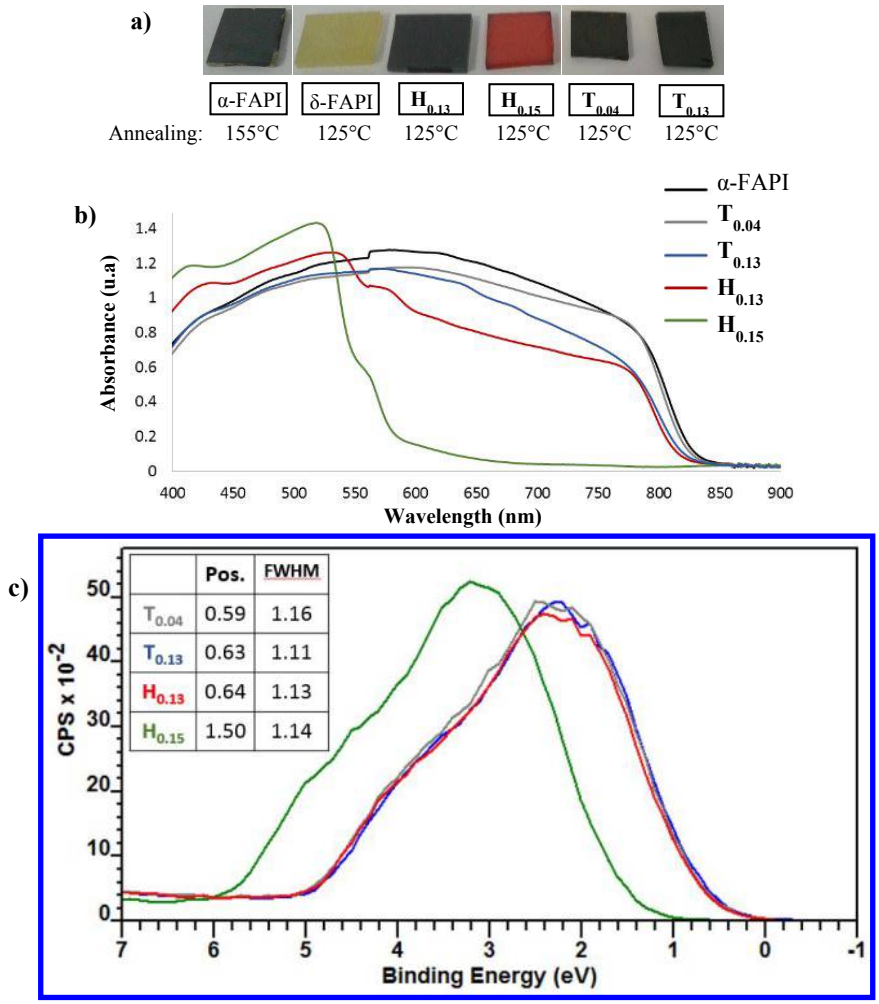


Figure 3. a) *Photography of α-FAPI, δ-FAPI,* $\boldsymbol{H_{0.13}}$, $\boldsymbol{H_{0.15}}$, $\boldsymbol{T_{0.04}}$ *and* $\boldsymbol{T_{0.13}}$ *thin films and their associated temperatures of annealing.* ***b)*** *UV-visible absorbance spectra of α-FAPI (black),* $\boldsymbol{H_{0.13}}$ *(red),* $\boldsymbol{H_{0.15}}$ *(green),* $\boldsymbol{T_{0.04}}$ *(grey), and* $\boldsymbol{T_{0.13}}$ *(blue) thin films.* ***c)*** *XPS spectra of* $\boldsymbol{H_{0.13}}$ *(red),* $\boldsymbol{H_{0.15}}$ *(green),* $\boldsymbol{T_{0.04}}$ *(grey) and* $\boldsymbol{T_{0.13}}$ *(blue) thin films, representing the valence band (VB) energies relative to the Fermi energy* (E_F-E_{VB})*. Inset: respective both energies (eV) of* E_F-E_{VB} *(pos.) and FWHM.*

Thin films of *d*-α-FAPI-$\mathbf{H_x}$ and *d*-α-FAPI-$\mathbf{T_x}$ phases as well as α-FAPI were prepared by spin coating at room conditions (see S.I. -I.e- for details). The precursors were added into a solution of dimethylformamide (DMF), and depending on stoichiometries of the reagents, we prepared α-FAPI (PbI_2/FAI: 1/1), $\mathbf{H_{0.13}}$ (PbI_2/FAI/HEAI: 3/3/1), $\mathbf{H_{0.15}}$ (PbI_2/FAI/HEAI: 3/3/2), $\mathbf{T_{0.04}}$ (PbI_2/FAI/TEAI: 3/3/0.5), and contrary to the precipitated powders, we were able to prepare another successfully pure composition $\mathbf{T_{0.13}}$ (x = 0.13; PbI_2/FAI/TEAI: 3/3/1.0). The spinning step was followed by an annealing treatment of thin films, at 155°C for α-FAPI and at 125°C for $\mathbf{H_{0.13}}$, $\mathbf{H_{0.15}}$, $\mathbf{T_{0.04}}$, and $\mathbf{T_{0.13}}$. We must notice that the annealing of the FAPI thin film at 125 °C results in the formation of δ-FAPI as a pure phase (yellow thin film). After the annealing step, thin films of α-FAPI, $\mathbf{H_{0.13}}$, $\mathbf{T_{0.04}}$, and $\mathbf{T_{0.13}}$ are black, and red for $\mathbf{H_{0.15}}$ (see Figure **3a**).

As for powders, the I/Pb ratio of this films were determined from EDX analysis, leading to $x = 0.13$ ($\mathbf{H_{0.13}}$), $x = 0.15$ ($\mathbf{H_{0.15}}$) (Figure S1, Table S3), $x = 0.04$ ($\mathbf{T_{0.04}}$), and $x = 0.13$ ($\mathbf{T_{0.13}}$) (Figure S8, Table S6). XRD measurements revealed that $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ thin films were obtained as pure phases without the presence of δ-FAPI, α-FAPI or PbI_2. The XRD patterns of $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ exhibit a shift of most of peaks towards lower angles compared to the lines of α-FAPI, as observed for powders. We also notice that the three very weak peaks at low 2θ angle (6.2°, 8.7°, 12.3°), typical of the tetragonal super-cell of such *d*-HP (see single crystal study) are observed in the thin film pattern of $\mathbf{H_{0.15}}$ (Figure S17). As in the case of *d*-α-FAPI-$\mathbf{H_x}$ phases, XRD patterns of $\mathbf{T_{0.04}}$ and $\mathbf{T_{0.13}}$ thin films show single phase layers without the presence of PbI_2 (Figure S20). This is in contrast with the XRD of *α*-FAPI thin film (PbI_2 is known to be easily generated during the annealing treatment at 155°C -room conditions-) (Figures S18, S19).[27,28] As observed in the XRD of $\mathbf{T_{0.04}}$ powder and the XRD

of ethylene diammonium (en^{2+}) based *hollow* perovskites $(FA)_{1-x}(en)_x(Pb)_{1-0.7x}(I)_{3-0.4x}$,[24] there is no significant shift of diffraction lines in the XRD patterns of $\mathbf{T_{0.04}}$ and $\mathbf{T_{0.13}}$ thin films compared to the one of α-FAPI phase. However, the specific features observed in the XRD pattern of $\mathbf{T_{0.04}}$ powder compared to the one of α-FAPI, which have a higher I(100)/I(200) ratio ((100) line at 13.9°, (200) line at 28.0°) and a lower I(210) intensity ((210) line at 31.4°), are also observed for $\mathbf{T_{0.04}}$ and $\mathbf{T_{0.13}}$ thin films (Figure S20).

The UV-visible absorption spectra of α-FAPI and $\mathbf{H_{0.13}}$, $\mathbf{H_{0.15}}$, $\mathbf{T_{0.04}}$, $\mathbf{T_{0.13}}$ thin films are compared in Figure **3b**, and the optical bandgaps were determined from Tauc plot graphs (S.I., Figure S21) leading to 1.50 eV for α-FAPI, 1.51 eV for $\mathbf{T_{0.04}}$, 1.53 eV for $\mathbf{H_{0.13}}$ and $\mathbf{T_{0.13}}$, 1.90 eV for $\mathbf{H_{0.15}}$. For $x \leq 0.13$, both the observed blue shift of the bandgap and the decreasing absorption intensity in the 600-800 nm range are proportional to the x value of the perovskite frameworks $[Pb_{1-x}I_{3-x}]$, consistently with earlier findings on *d*-MAPI and *hollow* perovskites.[22,24] Interestingly, from $x = 0.13$ to $x = 0.15$, the blue shift of the optical bandgap becomes huge. A similar distinctive behaviour is observed using XPS measurements (Figure **3c**): the valence band (VB) energies relative to the Fermi energy (E_F-E_{VB}) is found almost unchanged for $\mathbf{T_{0.04}}$, $\mathbf{T_{0.13}}$ and $\mathbf{H_{0.13}}$, (0.59 eV, 0.63 eV and 0.64 eV, respectively), whereas it largely increases for $\mathbf{H_{0.15}}$ (1.50 eV). Meanwhile, by using low-lying electronic levels of iodine (4d) or lead (4f) as a reference, it is shown that the position of the VB is independent of the composition (S.I., Table S10). Besides, the XPS revealed also two different types of N atom for our *d*-α-FAPI compounds: NH_3^+ group belonging to HEA^+/TEA^+ and NH_2^+ to FA^+. Note that, as expected, only the N atom of NH_2^+ was observed for α-FAPI (as for examples $\mathbf{T_{0.13}}$ and α-FAPI, Figure S23). Moreover, as reported with *d*-MAPI,[22] the incorporation of HEA^+/TEA^+ cations has the consequence

to eliminate the formation of Pb(0) under continuous X-ray irradiation compared to the reference α-FAPI (Figure S24).

Furthermore insight into the electronic structure of *d*-α-FAPI compounds has been obtained by computational investigations based on density functional theory (DFT) including spin-orbit coupling. The starting point for the calculations is the X-ray resolved structure of *d*-α-FAPI-$\mathbf{H_{0.104}}$ crystal in the P4/m group with partial occupation of (0, 0, 0) and (0, 0, 0.5) lead and iodide sites (see above). As a result of the dynamical motion of the organic cation evidenced by solid state NMR spectroscopy (see above), no structural information is available for those cations. Thus, in order to compensate the negative charge of the inorganic part, Cs^+ cations have been placed at the center of the vacancies, where FA^+ are expected to be located (Figures **4a** & **4b**, and zoomed view Figure S25). As such, the distinctive nature between the three sites defined in Figure **2c** relies solely on the inorganic framework. Such a substitution has proved very useful for 3D and layered perovskites,[30,31] and holds whenever the molecular frontier orbitals are far above and far below the conduction (CB) and valence bands, respectively. The first structural model corresponds to an ideal 3D network with no deficiency of lead and iodide (x = 0.00). As expected, it shows a direct bandgap at A with a computed bandgap energy E_G = 0.28 eV (Figure **4a**). It is well-known that Kohn-Sham DFT dramatically underestimates bandgaps in semiconductors, although it correctly describes trends. The second structural model considers the other extreme of the experimentally determined crystal structure where all the partially occupied sites are taken empty (x = 0.20). To retain charge neutrality, additional Cs^+ cations are placed at the position of missing Pb atoms, where HEA^+ or TEA^+ cations are expected to be inserted. As a result, channels appear with deficiency of lead and

iodide. Despite the extent of the transformation, the resulting band structure (Figure **4b**) shows remarkably similar features to the pristine system, including a direct bandgap at A. Meanwhile, the deficiency in lead and iodide atoms leads to two notable differences. Firstly, the bandgap E_G undergoes a sizeable increase to 0.92 eV, which is in qualitative good agreement with UV-visible spectra (Figure **3b**). Concomitantly, the bandwidth of both the valence and conduction bands decreases, which is indicative of increased effective masses with variations of about +50% and +40% for holes and electrons, respectively. Naturally, the larger effective masses of charge carriers hinder the optoelectronic properties of x = 0.20 material by lowering the probability to extract charges in the case of solar cells.

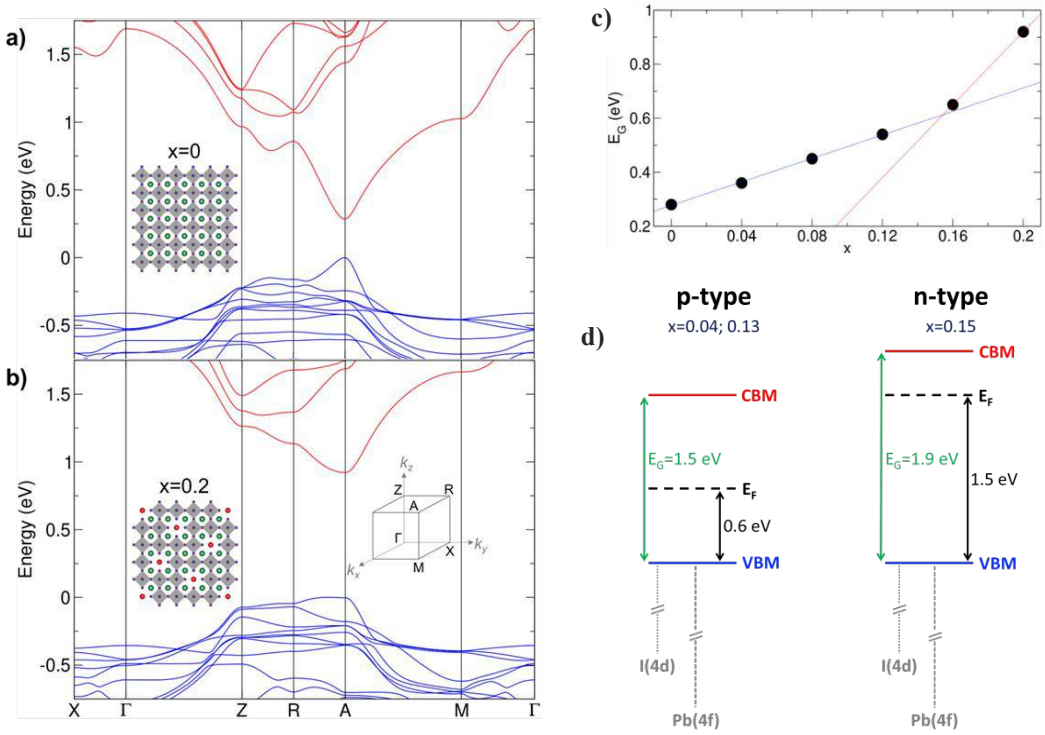


Figure 4. a) *Electronic band structure for the structure deprived of lead vacancy.* ***b)*** *Same for a concentration of vacancies x = 0.2. Insets present the structure used for calculations with Pb and I atoms in grey and violet, respectively. Green and red balls show the positions of* Cs^+ *cations. Red balls mark the position of PbI vacancies.* ***c)*** *Computed bandgap* E_G *(eV) with respect to the vacancy concentration x. Lines in blue and red correspond to linear fits plotted as guide for the eye.* ***d)*** *Schematic evolution of band edges and Fermi level position derived from XPS and DFT results.*

To further investigate the impact of composition, we designed a supercell (see S.I., Figure S25). This supercell allows to consider structural defective compositions with x = 0.00, 0.04, 0.08, 0.12, 0.16 and 0.20 and monitor the evolution of E_G with respect to x (Figure **4c**). As expected, a monotonous increase of E_G is obtained between the two end compositions but, two regions can be identified marked by a steep opening of the bandgap for x > 0.12 much greater than the one observed for x ≤ 0.12. (blue and red lines on Figure **4c**). Indeed, the linear opening of the bandgap seen for x >0.12 is three times steeper than the one observed for x ≤0.12. This computationally acceleration of bandgap opening for x >0.12 are qualitatively in line with the large variation of E_G and E_F-E_{VB} experimentally observed by UV-visible spectrometry and XPS, respectively, when increasing x from 0.13 to 0.15 (Figure **3b** and **3c**). In addition, considering the Pb 5d orbitals of as a reference, DFT calculations also confirm that the valence band energy is hardly affected in this range of composition, as the systematic downshift with increasing x value remains below 100meV.

Thus, despite the limitations of the here-presented computational model, in which the x value is monitored by considering an increasing number of infinite channels of defects starting from the experimental structure recorded on the single crystal $(HEA)_{0.406}(FA)_{0.698}[Pb_{0.896}I_{2.896}]$ (**$H_{0.104}$**), DFT results recover qualitatively both experimental trends obtained on thin films. These experimental and computed results suggest different band diagrams for x ≤0.13 and x >0.13 that are schematically sketched in Figure **4d**. It shows that the Fermi level is closer to the VB for the less-deficient (x ≤ 0.13) compositions, which is characteristic of a p-type semiconductor with positive charges as the majority charge carriers. On the contrary, for x = 0.15, the Fermi level is found closer

to the conduction band edge, indicative of a n-type semiconductor. Recently, G. Paul *et al.* inferred a correlation between electronic conductivity from p-type to n-type with PbI_2-deficient and PbI_2 rich precursors both for MAPI and FAPI.[32] Unfortunately, for the *d*-FAPI investigated in our work, any attempt to correlate composition to stoichiometry of precursors failed and the overall trend is found in the opposite direction as the films with $x = 0.15$ have been made using a stoichiometric ratio more PbI_2-deficient than the one used for $x \leq 0.13$ (Table S10). This suggests that other factors can influence the semiconductor doping type and level. Amongst others, it has been shown that thermal annealing can convert an initially p-type halide perovskite to n-type.[33]

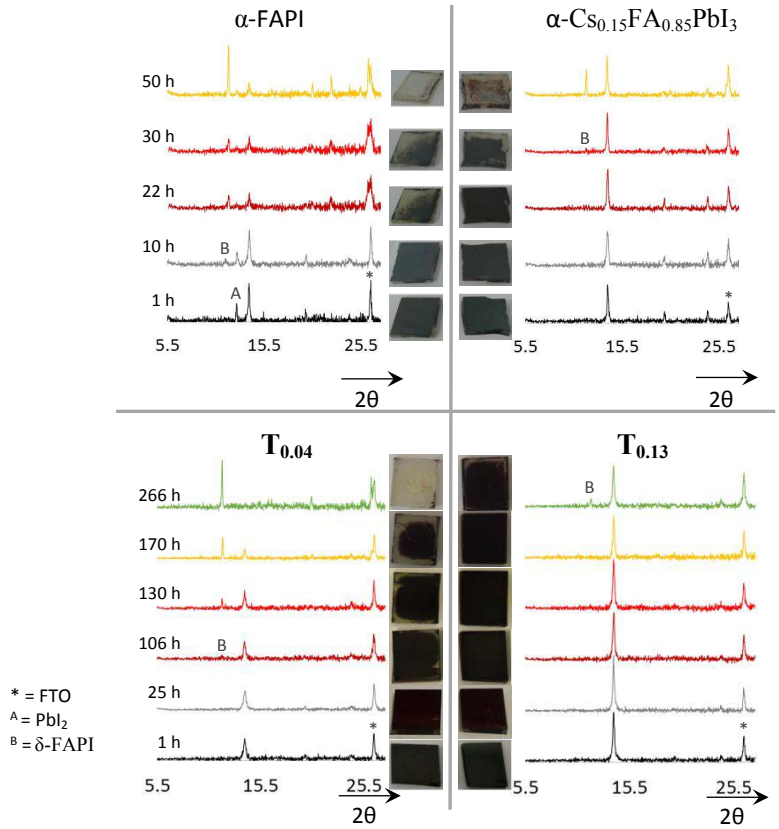


Figure 5. *Stability test (ambient conditions: 25 °C, 75-90% humidity) of α-FAPI, α-$Cs_{0.15}FA_{0.85}PbI_3$,* ***$T_{0.04}$, $T_{0.13}$*** *thin films. The XRD patterns and the pictures were realized in the 1-50 hours range and 1-266 hours range for α-FAPI, α-$Cs_{0.15}FA_{0.85}PbI_3$ and* ***$T_{0.04}$, $T_{0.13}$,*** *respectively.*

The α-phase air stability of **$H_{0.13}$**, **$H_{0.15}$**, **$T_{0.04}$** and **$T_{0.13}$** thin films has been compared to both α-FAPI and α-$Cs_{0.15}FA_{0.85}PbI_3$ samples. This mixed cation α-FAPI type perovskite has been selected as a reference material due to its higher α-phase stability and performances in solar cells.[11,34,35] The corresponding thin films were prepared by spin-coating with a final annealing step at 125 °C for α-$Cs_{0.15}FA_{0.85}PbI_3$, **$H_{0.13}$**, **$H_{0.15}$**, **$T_{0.04}$**, **$T_{0.13}$**, and at 155°C for α-FAPI (more details in S.I. -X-). All these films were aged under ambient conditions (25

°C, air with 75-90% humidity, ambient sunlight), without encapsulation and their aspect and structural changes upon aging were followed (Figure **5** and Figures S26, S27). After 1 hour, only the XRD pattern of α-FAPI thin film presented the typical peak of PbI_2 (12.8° 2θ), a compound which is in fact generated during the annealing step at 155°C. After 10 hours, the typical line of δ-FAPI at 11.7° 2θ appeared in the XRD pattern of α-FAPI, then the α-to-δ phase transformation was completed after around 50 hours of aging. α-$Cs_{0.15}FA_{0.85}PbI_3$ thin film owns a higher α-phase stability under ambient conditions than α-FAPI since the α-to-δ transformation starts after 30 hours and is completed between 50 and 106 hours of aging (S.I. Figure S27). The $\mathbf{H_{0.13}}$ and $\mathbf{H_{0.15}}$ thin films exhibited a similar stability to the α-FAPI one (total transformation into δ-FAPI between 22 and 50 hours of aging, Figure S26). Importantly, the $\mathbf{T_{0.04}}$ and $\mathbf{T_{0.13}}$ thin films were significantly more stable. Indeed, the first sign of δ-phase formation occurred after 106 hours and 266 hours for $\mathbf{T_{0.04}}$ and $\mathbf{T_{0.13}}$, respectively. This means that under ambient conditions, $\mathbf{T_{0.13}}$ thin films are approximately twenty times more stable than α-FAPI and ten times more stable than α-$Cs_{0.15}FA_{0.85}PbI_3$. This exceptional α-phase stability is presumably the consequence of the presence of the TEA^+ cation inside the 3D perovskite network. Interestingly, taking into account that HEA^+ and TEA^+ cations own the same alkyl chain, and that their size and amounts of incorporation are similar, such a higher stability of $\mathbf{T_{0.13}}$ compared to $\mathbf{H_{0.13}}$, is due to added benefit thanks to the thiol function as compared to the alcohol one. The thiol function is less hydrophilic and may make stronger interactions with the inorganic network, thus allowing a better stabilisation of the α-phase in air condition. This hypothesis of stronger interactions with the inorganic network is reinforced by the solid NMR data,

which evidenced that TEA^+ presents much less rotational mobility as compared to the two other organic cations considered in this work (FA^+, HEA^+).

CONCLUSION

In summary, we have shown that both HEA^+ and TEA^+ large cations (effective radius of 3.44 Å and 4.14 Å, respectively) can be incorporated into a 3D $[Pb_{1-x}I_{3-x}]$ α-FAPI type perovskite network, showing that it is possible to circumvent the geometrical Goldschmidt factor. The resulting phases *d*-α-FAPI-$\mathbf{H_x}$ and *d*-α-FAPI-$\mathbf{T_x}$ can be obtained by an easy precipitation method at room temperature, while the well-known α-FAPI parent phase can only be obtained above 150 °C. Incorporation of such larger size cations, particulary the thiol-cation TEA^+, leads to thin films demonstrating exceptional α-phase stability under ambient conditions, surpassing the one of α-FAPI and α-$Cs_{0.15}FA_{0.85}PbI_3$ by more than an order of magnitude. The higher α-phase stability of *d*-α-FAPI-$\mathbf{T_x}$ materials compared to *d*-α-FAPI-$\mathbf{H_x}$ ones, points out the positive impact of the thiol function compared to alcohol function. This certainly originates from the less hydrophilic nature of –SH compared to –OH, and it may also indicate greater abilities of the thiol function to make quite strong interactions with the inorganic perovskite network, which leads to a slower motion of the TEA^+ cations as evidenced by solid-state NMR. Besides, computational investigations correlated with experimental observations reveal the progressive opening of the direct band gap and increase of the effective masses when *x* increase from 0 to 0.13. For greater values of *x*, those trends remain but the evolution dramatically accelerates. Moreover, UV-visible, XPS and computational results suggest that *d*-α-FAPI compounds turn from p-type to n-type semiconductors for *x* values above 0.15. These observations prove that it is possible

to accurately tune the electronic properties of this *d*-α-FAPI phases lead and iodide deficient hybrid perovskite family (*d*-HP) by controlling *x*. Thus, *d*-α-FAPI materials appears as serious MA^+ and Br^--free candidates as an alternative to the widespread alloys $(Cs,MA,FA)Pb(I,Br)_3$ for single junction but also tandem PSCs, which recently achieved certifief power conversion efficiencies of 28.0%.[3] We believed that these results pave a path to the discovery of new perovskite materials for energy applications thanks to the greater versatility of defective sites.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the

ACS Publications website

AUTHOR INFORMATIONS

Corresponding Authors
nicolas.mercier@univ-angers.fr ; claudine.katan@univ-rennes1.fr

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REFERENCES

(1) Brenner, T. M.; Egger, D. A.; Kronik, L.; Hodes, G.; Cahen, D. Hybrid Organic-Inorganic Perovskites: Low-Cost Semiconductors with Intriguing Charge-Transport Properties. *Nature Reviews Materials.* **2016**, *1*, 15007.

(2) Song, Z.; McElvany, C. L.; Phillips, A. B.; Celik, I.; Krantz, P. W.; Watthage, S. C.; Liyanage, G. K.; Apul, D.; Heben, M. J. A Technoeconomic Analysis of Perovskite Solar Module Manufacturing with Low-Cost Materials and Techniques. *Energy Environ. Sci*. **2017**, *10*, 1297-1305.

(3) National Renewable Energy Laboratory, N.R.E.L. Accessed 17 December 2018. https://www.nrel.gov/pv/assets/pdfs/pv-efficiency-chart.20181214.pdf.

(4) Lee, M. M.; Teuscher, J.; Miyasaka, T.; Murakami, T. N.; Snaith, H. J. Efficient Hybrid Solar Cells Based on Meso-Superstructured Organometal Halide Perovskites. *Science*. **2012**, *338*, 643-647.

(5) Kadro, J. M.; Pellet, N.; Giordano, F.; Ulianov, A.; Müntener, O.; Maier, J.; Grätzel, M.; Hagfeldt, A. Proof-of-Concept for Facile Perovskite Solar Cell Recycling. *Energy Environ. Sci.* **2016**, *9*, 3172-3179.

(6) Kim, B. J.; Kim, D. H.; Kwon, S. L.; Park, S. Y.; Li, Z.; Zhu, K.; Jung, H. S. Selective Dissolution of Halide Perovskites as a Step Towards Recycling Solar Cells. *Nature Communications*. **2016**, *7*, 11735.

(7) Aristidou, N.; Eames, C.; Islam, M. S. ; Haque, S. A. Insights into the Increased Degradation Rate of $CH_3NH_3PbI_3$ Solar Cells in Combined Water and O_2 Environments. *J.Mater.Chem.A*. **2017**, *5*, 25469-25475.

(8) Correa-Baena, J.-P.; Saliba, M.; Buonassisi, T.; Grätzel, M. ; Abate, A.; Tress, W.; Hagfeldt, A. Promises and Challenges of Perovskite Solar Cells. *Science.* **2017**, *358*, 739-744.

(9) Yang, W. S.; Noh, J. H.; Jeon, N. J.; Kim, Y. C.; Ryu, S.; Seo, J.; Seok, S. I. High-Performance Photovoltaic Perovskite Layers Fabricated through Intramolecular Exchange. *Science*. **2015**, *348* (6240), 1234-1237.

(10) Prathapani, S.; Choudhary, D.; Mallick, S.; Bhargava, P.; Yella, A. Experimental Evaluation of Room Temperature Crystallization and Phase Evolution of Hybrid Perovskite Materials. *CrystEngComm.* **2017**, *19*, 3834-3843.

(11) Li, Z.; Yang, M.; Park, J.-S.; Wei, S.-H.; Berry, J. J.; Zhu, K. Stabilizing Perovskite Structures by Tuning Tolerance Factor: Formation of Formamidinium and Cesium Lead Iodide Solid-State Alloys. *Chem. Mater.* **2016**, *28*, 284-292.

(12) Pellet, N.; Gao, P.; Gregori, G.; Yang, T.-Y. ; Nazeeruddin, M. K.; Maier, J.; Gratzel, M. Mixed-Organic-Cation Perovskite Photovoltaics for Enhanced Solar-Light Harvesting. *Angew. Chem. Int. Ed.* **2014**, *53*, 3151-3157.

(13) Abdi-Jalebi, M.; Andaji-Garmaroudi, Z.; Cacovich, S.; Stavrakas, C.; Philippe, B.; Richter, J. M.; Alsari, M.; Booker, E. P.; Hutter, E. M.; Pearson, A. J.; Lilliu, S.; Savenije, T. J.; Rensmo, H.; Divitini, G.; Ducati, C.; Friend, R. H.; Stranks, S. D. Maximizing and Stabilizing Luminescence From Halide Perovskites with Potassium Passivation. *Nature*. **2018**, *555*, 497-501.

(14) Cao, J.; Wu, B.; Chen, R.; Wu, Y.; Hui, Y.; Mao, B.-W. ; Zheng, N. Efficient, Hysteresis-Free, and Stable Perovskite Solar Cells with ZnO as Electron-Transport Layer: Effect of Surface Passivation. *Adv. Mater.* **2018**, *30*, 1705596.

(15) Cho, K. T.; Grancini, G.; Lee, Y.; Oveisi, E.; Ryu, J.; Almora, O.; Tschumi, M.; Schouwink, P. A.; Seo, G.; Heo, S.; Park, J.; Jang, J.; Paek, S.; Garcia-Belmonte, G.; Nazeeruddin, M. K. Selective Growth of Layered Perovskites for Stable and Efficient Photovoltaics. *Energy Environ. Sci.* **2018**, *11*, 952-959.

(16) Zhao, Y.; Tan, H.; Yuan, H.; Yang, Z.; Fan, J. Z.; Kim, J.; Voznyy, O.; Gong, X.; Quan, L. N.; Tan, C. S.; Hofkens, J.; Yu, D.; Zhao, Q.; Sargent, E. H. Perovskite Seeding Growth of Formamidinium-Lead-Iodide-based Perovskites for Efficient and Stable Solar Cells. *Nature Communication*. **2018**, *9*, 1607.

(17) Gratzel, M. The Rise of Highly Efficient and Stable Perovskite Solar Cells. *Acc. Chem. Res.* **2017**, *50*, 487-491.

(18) Saliba, M.; Matsui, T.; Domanski, K. ; Seo, J.-Y. ; Ummadisingu, A. ; Zakeeruddin, S. M.; Correa-Baena, J.-P.; Tress, W. R.; Abate, A.; Hagfeldt, A.; Grätzel, M. Incorporation of Rubidium Cations into Perovskite Solar Cells Improves Photovoltaic Performance. *Science*. **2016**, *354* (6309), 206-209.

(19) Saliba, M.; Matsui, T.; Seo, J.-Y. ; Domanski, K. ; Correa-Baena, J.-P.; Nazeeruddin, M. K.; Zakeeruddin, S. M.; Tress, W.; Abate, A.; Hagfeldt, A.; Gratzel, M. Cesium-ContainingTriple Cation Perovskite Solar Cells: Improved Stability, Reproducibility and High Efficiency. *Energy Environ. Sci.* **2016**, *9*, 1989-1997.

(20) Wang, P.; Ulfa, M.; Pauporté, T. Effects of Perovskite Monovalent Cation Composition on the High and Low Frequency Impedance Response of Efficient Solar Cells. *J. Phys. Chem. C*. **2018**, *122*, 1973-1981.

(21) Turren-Cruz, S.-H.; Hagfeldt, A.; Saliba, M. Methylammonium-Free, High-Performance and Stable Perovskite Solar Cells on a Planar Architecture. *Science*. **2018**, *362* (6413), 449-453.

(22) (a) Leblanc, A.; Mercier, N. ; Allain, M. ; Dittmer, J. ; Fernandez, V. ; Pauporté, T. Lead- and Iodide-Deficient $(CH_3NH_3)PbI_3$ (*d*-MAPI): The Bridge between 2D and 3D Hybrid Perovskites. *Angew. Chem.* **2017**, *129*, 16283-16288. (b) Patent CRO BFF 17P0180, Project (1), 14th March **2017**, Depositor SATT OUEST VALORISATION – Client reference: DV3293.

(23) Tsai, C.-M.; Lin, Y.-P. ; Pola, M. K. ; Narra, S. ; Jokar, E. ; Yang, Y.-W. ; Diau, E. W.-G. Control of Crystal Structures and Optical Properties with Hybrid Formamidinium and 2-Hydroxyethylammonium Cations for Mesoscopic Carbon-Electrode Tin-Based Perovskite Solar Cells. *ACS Energy Lett.* **2018**, *3*, 2077-2085.

(24) Spanopoulos, I.; Ke, W.; Stoumpos, C. C.; Schueller, E. C. ; Kontsevoi, O. Y. ; Seshadri, R. ; Kanatzidis, M. G. Unraveling the Chemical Nature of the 3D "Hollow" Hybrid Halide Perovskites. *J. Am. Chem. Soc.* **2018**, *140*, 17, 5728-5742.

(25) Kieslich, G.; Sun, S.; Cheetham, A. K. Solid-State Principles Applied to Organic-Inorganic Perovskites: New Tricks for an Old Dog. *Chem. Sci.* **2014**, *5*, 4712-4715.

(26) Weller, M. T.; Weber, O. J.; Frost, J. M.; Walsh, A. Cubic Perovskite Structure of Black Formamidinium Lead Iodide, α-$[HC(NH_2)_2]PbI_3$, at 298 K. *J. Phys. Chem. Lett.* **2015**, *6*, 3209-3212.

(27) Taylor, V. C. A.; Tiwari, D.; Duchi, M.; Donaldson, P. M.; Clark, I. P.; Fermin, D. J.; Oliver, T. A. A. Investigating the Role of the Organic Cation in Formamidinium Lead Iodide Perovskite Using Ultrafast Spectroscopy. *J. Phys. Chem. Lett.* **2018**, *9*, 895-901.

(28) Eperon, G. E.; Stranks, S. D.; Menelaou, C.; Johnston, M. B.; Herz, L. M.; Snaith, H. J. Formamidinium Lead Trihalide: a Broadly Tunable Perovskite for Efficient Planar Heterojunction Solar Cells. *Energy Environ. Sci.* **2014**, *7*, 982-988.

(29) Kohrt, C.; Spannenberg, A.; Werner, T. 2-Hydroxyethylammonium Iodide. *Acta Cryst*. **2014**, *E70*, o628. DOI : 10.1107/S1600536814009581.

(30) Even, J.; Pedesseau, L.; Katan, C. Proc. SPIE, 2014, 9140, 91400Y.

(31) Even, J.; Pedesseau, L.; Jancu, J.-M.; Katan, C. Importance of Spin-Orbit Coupling in Hybrid Organic/Inorganic Perovskites for Photovoltaic Applications. *J. Phys. Chem. Lett*. **2013**, *4*, 2999-3005.

(32) Paul, G.; Chatterjee, S.; Bhunia, H.; Pal, A. J. Self-Doping in Hybrid Halide Perovskites via Precursor Stoichiometry: To Probe the Type of Conductivity through Scanning Tunneling Spectroscopy. *J. Phys. Chem. C*. **2018**, *122*, 35, 20194-20199.

(33) Wang, Q.; Shao, Y.; Xie, H.; Lyu, L.; Liu, X.; Gao, Y.; Huang, J. Qualifying Composition Dependent p and n Self-Doping in $CH_3NH_3PbI_3$. *Applied Physics Letters*. **2014**, *105*, 163508.

(34) Song, Z.; Wang, C.; Phillips, A. B.; Grice, C. R.; Zhao, D.; Yu, Y.; Chen, C.; Li, C.; Yin, X.; Ellingson, R. J.; Heben, M. J.; Yan, Y. Probing the Origins of Photodegradation in Organic–Inorganic Metal Halide Perovskites with Time-Resolved Mass Spectrometry. *Sustainable Energy Fuels*. **2018**, *2*, 2460-2467.

(35) Lee, J.-W.; Kim, D.-H.; Kim, H.-S.; Seo, S.-W.; Cho, S. M.; Park, N.-G. Formamidinium and Cesium Hybridization for Photo- and Moisture-Stable Perovskite Solar Cell. *Adv. Energy Mater.* **2015**, *5*, 1501310.

(36) Mercier, N.; Poiroux, S.; Riou, A.; Batail, P. Unique Hydrogen Bonding Correlating with a Reduced Band Gap and Phase Transition in the Hybrid Perovskites $(HO(CH_2)_2NH_3)_2PbX_4$ (X = I, Br). *Inorganic Chemistry*. **2004**, *43*, 26, 8361-8366.

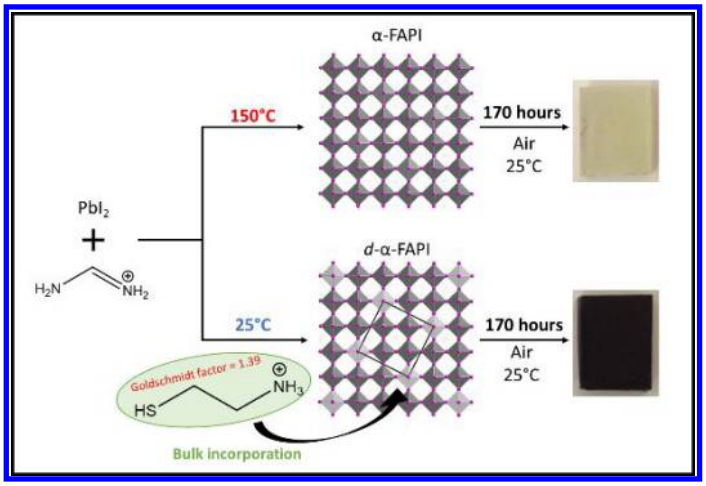
(37) Soler, J. M.; Artacho, E.; Gale, J. D.; García, A.; Junquera, J.; Ordejón, P.; Sánchez-Portal, D. The SIESTA Method for ab initio Order-N Materials Simulation. *J. Phys.: Condens. Matter.* **2002**, *14*, 2745-2779.

(38) Artacho, E.; Anglada, E.; Diéguez, O.; Gale, J. D.; García, A.; Junquera, J.; Martin, R. M.; Ordejón, P.; Pruneda, J. M.; Sánchez-Portal, D. ; Soler, J. M. The SIESTA Method; Developments and Applicability. *J. Phys.: Condens. Matter.* **2008**, *20*, 064208.

(39) Zhang, Y.; Yang, W. Generalized Gradient Approximation Made Simple. *Phys. Rev. Lett.* **1998**, *80*, 890.

(40) Troullier, N.; Martins, J. L. Efficient Pseudopotentials for Plane-Wave Calculations. *Phys. Rev. B*. **1991**, *43*, 1993-2006.

(41) Artacho, E.; Sánchez-Portal, D.; Ordejón, P.; García, A.; Soler, J. M. Linear-Scaling ab-initio Calculations for Large and Complex Systems. *Phys. Stat. Sol. (b)* **1999**, *215*, 809-817.

(42) Fernández-Seivane, L.; Oliveira, M. A.; Sanvito, S.; Ferrer, J. On-Site Approximation for Spin-Orbit Coupling in Linear Combination of Atomic Orbitals Density Functional Methods. *J. Phys.: Condens. Matter*. **2006**, *18*, 7999-8013.

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