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1 **Application of a mineral binder to reduce VOC emissions from indoor**
2 **photocatalytic paints**

3
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14
15 **ABSTRACT**

16 Given toxic nature of many volatile organic compounds (VOCs) present within indoor
17 environments, it is necessary to measure and quantify indoor VOC emissions to better inform
18 and protect the public from possible adverse health effects of indoor air pollution. To better
19 understand and quantify this problem, a horizontal flow tube reactor was used to study VOC
20 emissions from selected paint. Studied paints include mineral silicate binders, acrylic binders
21 and acrylic/siloxane binders with and without incorporated titanium dioxide (nano-TiO₂)
22 nanoparticles. Surface emission fluxes of selected VOCs from the tested paints were detected
23 and quantified using a High Sensitivity-Proton Transfer Reaction-Mass Spectrometry (HS-
24 PTR-MS). Low VOC emissions were observed for reference paints (absence of nano-TiO₂) in

1 the presence of UV irradiation. On the other hand, important formation of formaldehyde,
2 acetaldehyde and pentanal were observed for photocatalytic paints (impregnated with nano-
3 TiO₂). VOC emission fluxes from reference paints and photocatalytic paints were compared
4 to determine the formation of VOCs due to a reaction between the binder and radical reactive
5 species created on photocatalytic surfaces. Different matrix impacts for each paint were
6 studied and an important difference in VOC emission between acrylic binder paints and
7 mineral binder paints was observed under UV irradiation. A 66%, 29% and 88% decreases in
8 formaldehyde, acetaldehyde and pentanal, respectively, emission were observed for mineral
9 binder compared to acrylic binder VOC emission. In majority of the experiments, mineral
10 binder emitted less VOCs compared to acrylic binder. This mineral binder seems to be an
11 important factor in improving indoor air quality.

12

13 **Keywords:** Indoor Air; Formaldehyde; Photocatalytic Paints; VOC emissions; Surface
14 emission fluxes

15

1. INTRODUCTION

It is now well recognized that private residences, schools and commercial buildings among others are important environments of human exposure to air pollutants¹. In the modern world, an average human spends over 90% of her/his lifetime within indoor settings². Numerous compounds that include nitrous acid (HONO)³, particulate matter (PM), carbon monoxide (CO), nitrogen oxides (NOx) and volatile organic compounds (VOCs) are known to be ubiquitous indoor air pollutants^{4,5}. Some VOCs are known to cause adverse human health effects because they are easily absorbed through skin and/or mucous membranes^{6,7}. Based on a study carried out by the United States Environmental Protection Agency (U.S. EPA)², the VOC levels within indoor environments in US are typically 5 to 10 times higher than those of outdoor settings. Further, it has been observed that within Chinese homes, the indoor VOC levels are 100 to 1000 times higher compare to European home environment⁸. Indoor VOC sources are numerous and include emissions from building materials, combustion processes, furnishing or household products. Recently, Geiss *et al.*⁹ studied indoor (offices, schools, private homes) and outdoor environment emissions of selected VOCs across Europe to calculate their indoor-to-outdoor environment concentration ratios. These investigators calculated the work to school environment ratios of 21 VOCs that included BTEX compounds (namely, benzene, toluene, ethylbenzene, xylenes) and aldehydes (namely, formaldehyde and acetaldehyde among others) to range from 1.2 to 14.3. The results of Geiss *et al.*⁹ imply point sources of VOCs within indoor settings since the ratio indoor-to-outdoor environment ratio > 1 indicates higher VOC concentrations within indoor environments than outdoors. In France¹⁰, similar work has been performed in low-energy buildings, representative of future building designs and constructions. In a simulation similar to the one listed above, the calculated indoor-to-outdoor ratios were > 1 meaning that health risks for indoor environment occupants will remain a major occupational concern well into the future.

1 To better improve indoor and outdoor air quality, different nano-TiO₂ particle based
2 photocatalytic paint materials continue to be developed. In the environmental remediation of
3 atmospheric pollutants, nano-sized TiO₂ photocatalyst was chosen because of its superior
4 photocatalytic oxidation potentials, non-toxic properties, low price and anti-photocorrosive
5 properties¹¹ compared to other heterogeneous photocatalytic methods or photoelectrochemical
6 phenomena used in the building and construction industry.

7 To date, a majority of small scale laboratory and larger scale chamber¹² experimental studies
8 have concentrated their research effort to study the photocatalytic consumption and/or
9 degradation of NO_x¹³⁻¹⁷ (NO_x = NO + NO₂) and VOC¹⁸⁻²¹ pollutants on selected
10 photocatalytic materials. Fewer studies have been reported in the literature that researched
11 VOC emissions from the photocatalytic paint materials themselves^{9,21-23}. However, it has
12 been proposed that the binder molecules present within the paint may react with reactive
13 radical species, namely, OH radicals, present on nano-TiO₂ surfaces in the presence of light to
14 emit VOCs, therefore, act as a point source of atmospheric pollutants^{9,21-23}. This is
15 troublesome since practically all paint materials contain various amounts of organic binder
16 matter or a solvent, that is, all paints include some sort and quantity of a binder because it is
17 the binder itself that keeps the pigment in place after the paint dries. In fact, an important
18 increase in acetaldehyde and formaldehyde emissions from photocatalytic paints under UV
19 irradiation have been observed^{21,23}. Other compounds that include ethylacrolein, pentanal, 1-
20 hydroxy-butanone and hexanal have also been detected²¹. As a result, the photocatalytic
21 reactions that lead to VOC emissions from wall paints are important because they may affect
22 the durability of the paint, limit its photocatalytic properties and have direct and/or indirect
23 impacts on human health, human well-being and indoor air quality. Further, the photocatalytic
24 degradation of the binder matrix may significantly reduce the advantages of this technology.
25 Consequently, to limit VOC emissions within indoor environments from the photocatalytic

1 degradation of the binder matter present in the photocatalytic paints, a new generation of
2 binders is being developed.

3 In this work, three binder matrices were investigated for possible photocatalytic VOC
4 emissions. The surface emission fluxes were measured for binders with and without nano-
5 sized TiO₂ particles. Initial experiments were realized under “dark” conditions, that is, in the
6 absence of UV irradiation, to better quantify any VOC emissions due to the inherent paint
7 chemical composition. Other experiments were carried out in the presence of UV light to see
8 if the presence of nano-sized TiO₂ particles plays an effect on the surface emission fluxes.
9 The obtained results for mineral silicate and acrylic/siloxane binders are compared with
10 acrylic binder to determine any possible impacts and effects of the paint composition. The
11 obtained results will help to improve the binder formulation materials to minimize any
12 photocatalytic VOC emissions from photocatalytic paints.

13 **2. MATERIALS AND METHODS**

14 **2.1. Preparation of paints**

15 All paints were formulated and produced by the local paint manufacturer ALLIOS (Marseille,
16 France). Three different binder pairs studied include (1) copolymer acrylic (mix butyl acrylate
17 and vinyl acetate), (2) a copolymer acrylic/siloxane and (3) mineral silicate. Each binder pair
18 included (1) one photocatalytic binder (P) that contained nano-sized TiO₂ particles and (2)
19 one reference paint (R) without any nano-sized TiO₂ particles present in the binder. All
20 mineral binders contained 5% of the organic binder to help the drying process of the selected
21 paint on the given support so that only non-volatile content of a paint remained after the
22 coating solidifies. Binder compositions are shown in Table 1.

23 Table 1: Composition of selected binders evaluated in this work. Binder with (+) and without (-) a photocatalytic
24 component.

Name	Binders	Photocatalytic
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		Component
P1	Acrylic copolymer + TiO ₂	+
R1	Acrylic copolymer	-
P2	Acrylic copolymer /siloxane (60/40) + TiO ₂	+
R2	Acrylic copolymer /siloxane (60/40)	-
P3	Mineral silicate + TiO ₂	+
R3	Mineral silicate	-

1

2 The paints were applied on one side of glass plates that were 30 cm long and 2 cm wide
3 according to a standard operating procedure developed by the manufacturer ALLIOS¹⁷. This
4 procedure allowed for production of homogeneous, uniform and reproducible wet films that
5 were 100 µm thick. The painted glass plates were stored for 21 days at T=298 K while a
6 humidified (55% RH) synthetic air was allowed to flow over the plates. Two types of TiO₂
7 particles were used: (1) the larger sized ultrafine (micro-size) TiO₂ powder was used to render
8 the paints whiter and (2) the nano-sized TiO₂ particles were used to render the paints
9 photoactive. All photocatalytic paints contained the TITANE P2 white pigment in anatase
10 form. The nano-sized TiO₂ particle content was 85% and the ultrafine (micro-size) TiO₂
11 powder had a listed specific area of 350 m² g⁻¹. All photocatalytic paints contained 13.4%
12 (w/w) of the inactive micro-size TITANE P2 particles. A mixture of ground additives (slurry),
13 was prepared with 35% (w/w) active nano-sized TiO₂. Then, other paint constituents
14 (architectural constituent of paints, i.e., CaCO₃ and micrometric TiO₂) was mixed with 10% of
15 the slurry to reach 3.5% (w/w) of photocatalytic active TiO₂ nanoparticles.

16 **2.2. Experimental setup**

17 The flow tube reactor used in this work is similar to the one used in previous studies of VOC
18 emissions from photocatalytic paints^{17,23}. Here, the horizontal flow tube reactor experiments
19 involved time-resolved detection of the selected VOCs emitted from the solid paint media
20 using the High Sensitivity-Proton Transfer Reaction-Mass Spectrometry (HS-PTR-MS)

1 (IONICON Analytik, Austria) in the presence of UV irradiation or under “dark” conditions,
2 that is without UV light irradiation.

3 The schematic diagram of the apparatus has been published elsewhere^{23,24} and is described in
4 detail in Figure S1. Some experimental details that are particularly relevant to this work are
5 given below.

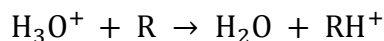
6 A borosilicate glass double-wall horizontal flow tube reactor with an internal volume of
7 approximately 131 cm³ was used to study surface VOC emission fluxes of selected solid
8 paints as a function of given experimental conditions. The flow tube reactor was maintained at
9 a constant temperature (± 1 K), using a RC6 LAUDA temperature-controlled circulating bath,
10 by allowing water to circulate through the outer jacket. All experiments were carried out at
11 $T=(298\pm 1)$ K and at 1 atm total pressure of synthetic air (Linde Gas, >99.999 stated purity).
12 The geometry of the flow tube reactor was such that it allowed for the synthetic air carrier gas
13 flow to enter at one end and the High Sensibility Proton-Transfer Reaction-Mass
14 Spectrometer (HS-PTR-MS) (IONICON Analytik) to be located downstream at the opposite
15 end. Similar to the work of Hanson and Ravishankara²⁵, the flow tube was mounted
16 horizontally, and the paint was applied on a glass plate located within the flow tube. The plate
17 was made of glass and was 30 cm long and 2 cm wide.

18 The flow tube reactor was placed within a stainless-steel protective box. Within the box, two
19 UV fluorescent lamps (Philips TL-D 18W Actinic BL, 340-400 nm, $\lambda_{\text{max}} = 368$ nm,
20 length = 60 cm) were placed side-by-side on the upper side of the reactor. A stainless-steel
21 box was used because stainless-steel exhibits good reflection properties and allows for a
22 homogeneous irradiation during experiments. The spectral irradiance of the UV lamps used in
23 this work has been determined previously²⁶ and was estimated to be 8.5 W m⁻². This value
24 represents a mean integrated irradiance, (340< λ (nm)<400), of solar light that is known to

1 enter an indoor environment²⁶. A sheath of synthetic air flow (300 sccm) was allowed to enter
2 the flow-tube system. This flow was controlled using a mass flow controller (Brooks SLA
3 Series mass flow controller, in the range 0–1 slm, $\pm 1\%$ stated accuracy). Dimensions and
4 geometry of the flow reactor and the experimental conditions allowed for the experiments to
5 be carried out at laminar flow conditions. The flow rate through the reactor was 1.1 cm s^{-1} and
6 the residence time was 27 seconds. Prior to allowing the synthetic air to enter the flow tube
7 reactor, the air flow was separated into two separate flows, one of dry air and the other
8 humidified by bubbling in deionized water (resistivity $>18 \text{ M}\Omega \text{ cm}$). Deionized water was
9 prepared by allowing tap water to pass through a reverse osmosis demineralization filter (ATS
10 Groupe Osmose) followed by a commercial deionizer (Millipore, Milli-Q50). Gas flows were
11 controlled using needle valves and the relative humidity could be changed by varying
12 throughput of these valves. Finally, the two flows were allowed to mix to obtain an
13 experimental relative humidity (RH) of 40%. The RH was measured online at the reactor exit
14 using a hygrometer “Hygrolog NT2” (Rotronic) with “HygroClip SC04” probe. The accuracy
15 of RH measurements was $\pm 1.5\%$. More details on the experimental setup is available in the
16 Supporting Information (Figure S1).

17 **2.3. VOCs measurements**

18 It is now well recognized, that Proton Transfer Reaction-Mass Spectrometry (PTR-MS)
19 technology is an excellent technique to measure rapid changes in gas phase concentration of
20 selected VOCs²³. In this work, an HS-PTR-MS system was used to carry out online
21 measurements of VOC emissions. This technique was previously described in detail
22 elsewhere²⁷. Briefly, it consists of an electron impact ion source that produces hydronium
23 ions, H_3O^+ , that react with the neutral organic molecule, R, within the drift tube to produce
24 water and the pseudo-molecular ion, $[\text{RH}]^+$, as show below:



1 This reaction is only possible if the proton affinity of R is higher than the proton affinity of
2 H₂O. The resulting ion products are then mass selected and detected using the electron
3 multiplier detector.

4 The HS-PTR-MS was operated with a drift pressure and a temperature of 2.02 mbar and
5 333 K respectively. The drift voltage was 500 V that corresponded to the E/N value of
6 124 Townsend (1 Townsend = 10⁻¹⁷ V cm²), where E is the electric field (V cm⁻¹) and N is the
7 ambient air number density within the drift tube (cm⁻³). The quantification of VOCs was
8 based on calibration with certified gas containing different mixtures of aromatic hydrocarbons
9 at 100 ppb levels (RESTEK 34423-PI).

10 The VOC emissions were analyzed in a full mass spectrum mode between 21 and 200 amu
11 and the detection rate was 1 s amu⁻¹. The resulting HS-PTR-MS time resolution was 3 min.

12 **2.4. VOCs mixing ratio and estimation of surface emission fluxes**

13 The VOC mixing ratio (ppb) was calculated using the following equation²⁸:

$$C_x = 1.65 \times 10^{-11} \times \frac{U_{drift} \times T_{drift}^2}{k \times P_{drift}^2} \times \frac{i[X]}{i[H_3O^+] + Xr \times i[H_3O^+(H_2O)]} \quad (1)$$

14 where X is the target compound, H₃O⁺ and H₃O(H₂O)⁺ are the reagent ions. i[X], i[H₃O⁺] and
15 i[H₃O⁺(H₂O)] are the ion signals in counts per second (cps) of X, H₃O⁺ and H₃O(H₂O)⁺,
16 respectively, normalized by their corresponding transmission efficiencies; U_{drift} is the drift
17 tube voltage (V); T_{drift} is the drift tube temperature (K); k is the proton-transfer reaction
18 constant (cm³ s⁻¹); P_{drift} corresponds to the pressure within the drift tube (mbar).

19 The *m/z* specific relative transmission efficiency was experimentally determined over the
20 mass range of 21-181 with a calibration gas standard, consisting of a mixture of 14 aromatic
21 organic compounds (TO-14A Aromatic Mix, Restek Corporation, Bellefonte, USA,
22 100±10 ppb in Nitrogen).

1 The factor X_r is compound specific and has been described previously²⁹. This factor reflects
2 both the difference in the rate coefficient for the proton-transfer-reactions $H_3O^+ + R$ and
3 $H_3O(H_2O)^+ + R$ and the difference in transmission efficiencies for the two reagent ions of the
4 quadrupole mass spectrometer. These values were directly taken from the works of de Gouw
5 *et al.*²⁹ and Gandolfo *et al.*²³. When no value was available in the literature, $X_r = 0.5$ was used
6 as recommended by de Gouw *et al.*²⁹.

7 The values of k have been determined elsewhere^{30,31} and when no proton-transfer-reaction
8 rate constant value for the compounds of interest were available in the literature, the canonical
9 $k = 2 \cdot 10^{-9} \text{ cm}^3 \text{ s}^{-1}$ value was used. The k and X_r values used in this work are shown in Table
10 S1 in the Supporting Information.

11 The HS-PTR-MS is known to underestimate the concentration of formaldehyde³². Therefore,
12 a correction factor of 2.7 was used to obtain the global quantity of this VOC. The correction
13 factor was previously determined in the laboratory setting and the experimental procedure is
14 detailed in the Supporting Information of Gandolfo *et al.*²³.

15 The surface emission fluxes ($\text{molecules cm}^{-2} \text{ s}^{-1}$) for different compounds were determined
16 using the following equation:

17

$$\text{Surface Emission Fluxes } X = \frac{C_x \times 2.46 \times 10^{10} \times \text{Internal reactor volume}}{(\text{Residence time} \times \text{Paint surface})} \quad (2)$$

18

19 Where the internal reactor volume, the residence time within the reactor and the paint surface
20 are expressed in cm^3 , s, cm^2 , respectively. C_x is the mixing ratio (ppb) measured at the exit of
21 the reactor as a function of time for the target compounds. 2.46×10^{10} is the conversion used to
22 convert ppb to molecule cm^{-3} .

3. RESULTS AND DISCUSSION

3.1. VOCs identification

The identification of selected VOCs was based on mother-ion masses weight, fragmentation fingerprints, bibliography and previously published data. Given the HS-PTR-MS instrument used in this work, certain mass to charge ratios, namely m/z 33, 45, 57, 59, 61, 67, 71, 79, 89, 107, 121 and 123 were easily attributed to specific VOC compounds. However, identification of eight ions, namely m/z 47, 57, 71, 73, 75, 87, 89 and 129 could not be easily attributed to a specific compound. However, earlier laboratory²³ of P1 and R1, using the Proton Transfer Reaction-Time of Flight-Mass Spectrometry (PTR-ToF-MS) attributed m/z 47, 57, 71, 73, 75, 87, 89 and 129 as formic acid, acrolein, methyl vinyl ketone (MVK), acrylic acid, propionic acid, pentanal and/or vinyl acetate, propionic acid and octanal, respectively. As a result, the above listed seventeen m/z ratios were chosen to identify and quantify the selected VOCs used in this work (see Table 2).

The formaldehyde, acetaldehyde and benzene emission results are shown individually because these compounds are known to cause adverse health effects. The other VOCs have been grouped as a function of their chemical class. The chemical classes of the selected compounds are shown below in Table 2.

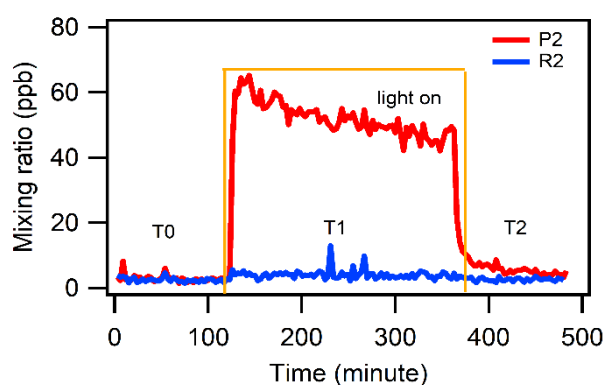
Table 2: Compounds detected and quantified using the HS-PTR-MS technique. References : (a) Manoukian *et al.*³³, (b) De Gouw *et al.*²⁹, (c) Karl *et al.*³⁴, (d) Hartungen *et al.*³⁵, (e) Jobson *et al.*³⁶, (f) Gandolfo *et al.*²³.

Most probable compound	m/z	Chemical classes	Ref
Formaldehyde	31		a, c
Methanol	33		b, c, e
Acetaldehyde	45		a, b, c, e
Formic Acid	47	Acids	c
Acrolein	57	Carbonyls	c
Acetone + Propanal	59	Carbonyls	a, b, c, e
Acetic Acid	61	Acids	b, c, d, e
Methyl Vinyl Ketone	71	Carbonyls	b, c
Acrylic Acid	73	Acids	f

Propionic Acid	75	Acids	d
Benzene	79		a, b, c, e
Pentanal + Vinyl Acetate	87	Carbonyls	a, f
Butanoic Acid	89	Acids	d
Ethylbenzene + Xylene	107	Aromatics	a, b, c
Trimethylbenzene + Ethyltoluene	121	Aromatics	a, b, c
Benzoic Acid	123	Aromatics	e, f
Octanal	111 and 129	Carbonyls	f

1 3.2. Typical Experiment

2 In a typical experiment, the studied paint on its glass support was placed in the horizontal
3 flow tube reactor and was left in the “dark” (UV lamps off) for two hours (T_0). Here, the
4 nano-TiO₂ particles present in the photocatalytic paint were not activated and the “dark” VOC
5 emissions were determined. After, the UV lights were turned on for three hours (T_1). Here, the
6 photocatalytic paints were activated and the formed reactive radicals species reacted with the
7 organic compounds present in the binder which leading to VOC formation^{9,22,23}. Finally, UV
8 lights were turned off (T_2) and the VOC emission signal was followed for another two hours.
9 A typical example of the formaldehyde temporal profile is shown in Figure 1. The observed
10 ion temporal profile is the same for all VOCs used in this work.



11
12 Figure 1: Typical temporal profile of formaldehyde emission from P2 (red line) and R2 (blue line) at $T=298$ K
13 and $RH=40\%$.

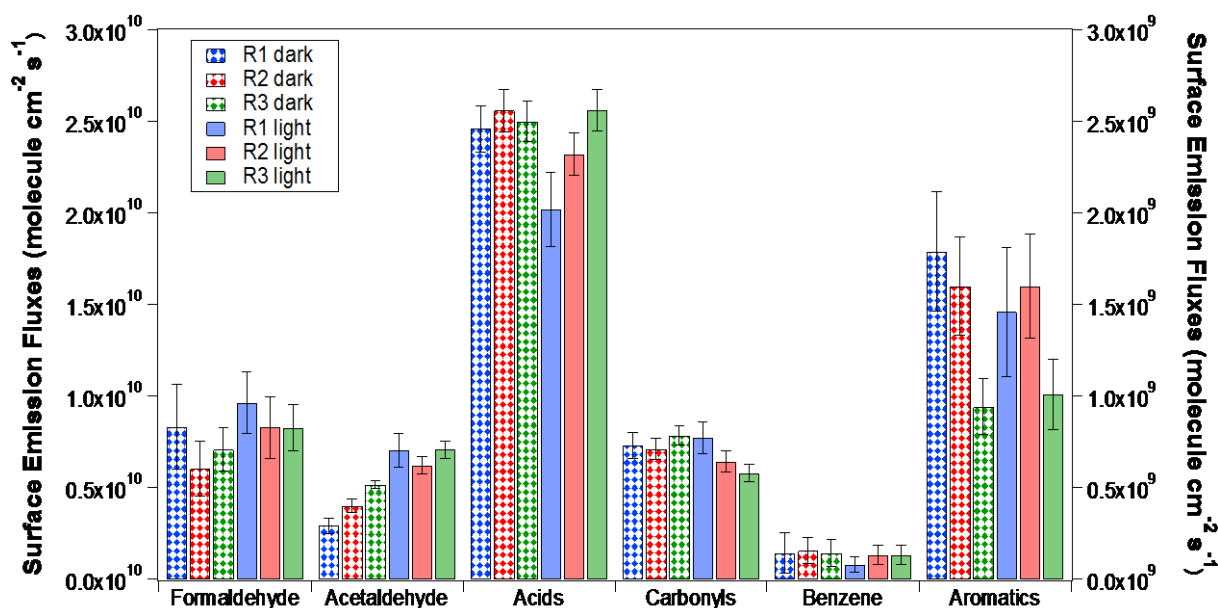
14 The observed formaldehyde temporal profiles obtained for photocatalytic (P2) and reference
15 (R2) paint are different. For R2, when the UV light is turned on, a very slight increase in the
16 formaldehyde temporal profile signal is observed and when the UV light is turned off, the ion
17 signal comes back to its original base value.

1 For P2, when the UV light is turned on the formaldehyde ion signal increases significantly.
 2 Here, in a time frame of a few minutes the formaldehyde ion signal stabilizes and remains
 3 nearly constant for the rest of the irradiation period (T1). Once the UV lamps are turned off,
 4 the intensity of the formaldehyde ion signal drops immediately and reaches its initial value
 5 within few minutes.

6 3.3. Surface Emission Fluxes

7 3.3.1. Reference paints

8 The VOCs surface emission fluxes of reference paints (R1, R2 and R3) under (1) “dark”
 9 condition (lozenge) and (2) under UV light (solid) are shown in Figure 2.



10
 11 Figure 2: VOCs surface emission fluxes (molecule cm⁻² s⁻¹) obtained for selected compounds under “dark”
 12 conditions (lozenge) and in presence of UV light (solid). Surface emission fluxes for formaldehyde,
 13 acetaldehyde, acids and the carbonyls are shown on the left axis, aromatics and benzene on the right axis. Blue
 14 bars correspond to R1, red bars to R2 and green bars R3. Errors bars are ±1σ precision based on ten
 15 experimental point values.

16 As shown in Figure 2, the surface emission fluxes of R2 and R3, were compared to R1 under
 17 “dark” conditions. For most compounds, the difference in surface emissions fluxes were
 18 inferior to 25% but in some cases an important difference was observed. For example, an
 19 increase of 37% for R2 and 77% for R3 in the acetaldehyde surface emission fluxes, have

1 been observed. The selected acids surface emission fluxes seem to be analogous for the three
2 binders. However, increases of 37% in acrylic acid, R2, and 39%, R3, in propionic acid
3 emissions have been observed. Other acids' surface emission fluxes are similar or decrease
4 down to 27% (formic acid R3). Moreover, for R3, decreases in aromatics' emissions have
5 been observed. This decrease is mainly due to 64% and 73% decrease in ethylbenzene and
6 benzoic acid emissions, respectively. In general, the surface emission fluxes values were in
7 the same range. Even though an observed maximum difference of 77% in emission fluxes was
8 observed under "dark" conditions the influence of binder composition on VOC emissions is
9 very low when the paint surfaces are not exposed to UV radiation. All surface emission fluxes
10 are shown on tables S2-S4 in Supporting Information.

11 The VOCs emissions under "dark" conditions were compared with those under UV light. As
12 shown in Figure 2, the most important variation was observed for acetaldehyde; emission
13 increases of 142% for R1, 55% for R2 and 37% for R3 have been measured under "light"
14 conditions compared to experiments carried out under "dark" experimental conditions. The
15 emissions of acids are observed to be similar under "dark" and "light" conditions. An
16 exception is the acrylic acid. Here, acrylic acid emission increases of 56% and 45% have been
17 measured for R1 and R2, respectively. On the other hand, a decrease of 12% in formic acid
18 emission and 29% in the acetic acid emission for R1 as well as a decrease of 12% in acetic
19 acid and formic acid emissions for R2, have been observed.

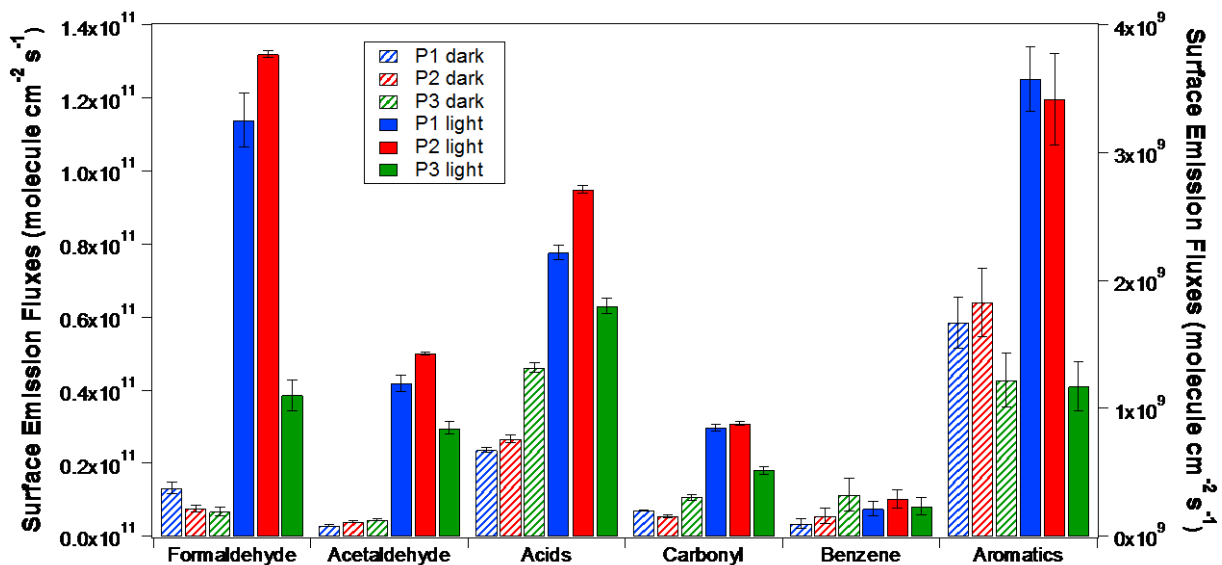
20 The surface emission fluxes of carbonyls compounds by R1 are, globally, similar even if the
21 emission of pentanal increase of 38% has been observed. Regarding R3, a decrease of
22 carbonyls emission has been observed due to 38% reduction in acetone emission. As shown in
23 Figure 2, a decrease of 44% of benzene emission by R1 has been observed.

1 With respect to the aromatics, an increase in 56% of benzoic acid emission has been observed
2 by R3 even if the global emission for this class of compounds is similar. For the rest of the
3 compounds, the difference in emission is inferior to 25%. All surface emission fluxes are
4 shown on tables S2-S4 in Supporting Information.

5 The obtained results show that near-UV light has a low impact on the VOCs surface emission
6 fluxes from reference points. Moreover, emission fluxes are similar for the three binders, that
7 is, inferior to 25%, for most emissions, for the three binders under “dark” and under UV
8 irradiation. Here, the binder composition has no impact on VOC emissions implying that a
9 limited impact of the selected reference points on indoor air pollution.

10 3.3.2. Photocatalytic paints under “dark” and UV light conditions

11 In this part, the surface emission fluxes of formaldehyde, acetaldehyde, benzene, aromatics,
12 acids and carbonyls were studied under “dark” conditions and UV irradiation for the three
13 photocatalytic paints (P1, P2 and P3). The obtained emissions results from photocatalytic
14 paints under “dark” and UV light experimental conditions are shown in Figure 3.



15
16 Figure 3: VOCs surface emission fluxes ($\text{molecule cm}^{-2} \text{s}^{-1}$) obtained for different compounds under “dark”
17 (hatched) conditions and under UV irradiation (solid). Surface emission fluxes for formaldehyde, acetaldehyde,

1 acids and carbonyls are shown on the left axis, aromatics and benzene on the right axis. Blue bars correspond to
2 P1, red bars to P2 and green bars to P3. Errors bars are $\pm 1\sigma$, precision based on ten experimental point values.

3 As shown in Figure S3, a comparison of the surface emission fluxes between photocatalytic
4 paints and reference paints under “dark” conditions shows an increase of 130% for benzene
5 and 83% for benzoic acid for P3. However, the VOC emissions from both photocatalytic and
6 reference paints are observed to be in the same concentration range. This implies that the
7 presence of nano-TiO₂ particles in the binder does not affect the VOC emissions under “dark”
8 conditions.

9 Further, VOC surface emission fluxes by P1, P2 and P3 under “dark” conditions are shown in
10 Figure 3. For formaldehyde, emissions have been observed to decrease 43% and 49% for P2
11 and P3, respectively. On the other hand, acetaldehyde emissions are observed to increase 37%
12 for P2 and 58% for P3.

13 The P3 surface emission flux of all acids under “dark” experimental conditions, increased by
14 95% compared to P1 surface emission flux. This emission increase has been observed to
15 correspond mainly to acetic acid and propionic acid emission, +219% and +131%,
16 respectively. An increase of 88% in acetone emissions have been observed for P3 but the
17 global emission of carbonyl compounds is similar for both types of paints. This compensation
18 is due to an observe decrease by 15% of MVK and pentanal surface emission fluxes. An
19 important difference is observed in the emission of benzene; an increase of 65% for P2 and
20 237% for P3 has been observed. Finally, aromatic compounds surfac0e emission fluxes have
21 been observe to decrease 33% for P3, an effect compounded by an0 induced benzoic acid
22 emission decrease of 48%. For the remaining VOCs, the difference of the emission fluxes
23 between the two binders is smaller than 30%. All surface emission fluxes are shown on tables
24 S5-S7 in Supporting Information.

1 Although P3 are observed to release 32% more VOCs compared to P1 and 44% more than P2
2 under “dark” conditions, this difference is negligible compared to the difference in emissions
3 under UV light experimental condition for photocatalytic paints. Therefore, it can be
4 concluded that the composition of the binder has a limited impact on VOC emissions under
5 “dark” conditions.

6 In the presence of UV light, the difference in surface emission fluxes was observed to be more
7 important compared to the experiments carried out under “dark” conditions as details below.

8 In addition to experiments carried out under “dark” conditions, the surface emission fluxes by
9 P1, P2 and P3 were also evaluated in the presence of UV radiation ($340 < \lambda(\text{nm}) < 400$,
10 $\lambda_{\text{max}} = 368 \text{ nm}$, 8.5 W m^{-2}). For R3, a decrease of 66% in formaldehyde, 29% in acetaldehyde,
11 40% in all carbonyl and 68% in all aromatic compounds emission as compared to P1 have
12 been observed. As shown in Figure 3, the selected compound emissions for P2 (red bars) and
13 P1 (blue bars) seem similar or at least within the shown error bars. However, it was not
14 possible to differentiate the individual emissions of selected VOCs in Figure 3. Emission
15 impact from P2 and P3 compared to P1 emission will be discussed later for individual VOCs
16 (Table 3).

17 All surface emission fluxes for each compound under different experimental conditions are
18 shown in the Supporting Information (Table S5-S7).

19 The impact of each binder on VOCs emission was evaluated in order to find out the best
20 performance of the photocatalytic paints as a remediation technology aimed to clean the air in
21 indoor environments.

22 As shown in Table 3, the selected VOC emissions from P2 and P3 are compared to P1, under
23 UV light experimental conditions.

1 Table 3: Surface emission fluxes of selected VOCs for P2 and P3 as compared to the VOC emissions observed
 2 for P1. Standard deviations are precision only and are based on ten experimental point values.

Compounds	P1 surface emission fluxes (molecule cm ⁻² s ⁻¹)	P2	P3
Formaldehyde	(1.14±0.07)·10 ¹¹	+16%	-66%
Methanol	(2.02±0.15)·10 ¹⁰	-7%	-41%
Acetaldehyde	(4.18±0.22)·10 ¹⁰	+20%	-29%
Formic Acid	(1.13±0.05)·10 ¹⁰	+7%	-27%
Acrolein	(7.15±0.65)·10 ⁹	+13%	-17%
Acetone + Propanal	(1.30±0.07)·10 ¹⁰	-14%	-25%
Acetic Acid	(2.58±0.15)·10 ¹⁰	+32%	+27%
Methyl Vinyl Ketone	(1.35±0.14)·10 ⁹	+3%	-66%
Acrylic Acid	(2.29±0.09)·10 ¹⁰	+40%	-94%
Propionic Acid	(6.46±0.44)·10 ⁹	+25%	+198%
Benzene	(2.15±0.59)·10 ⁸	+36%	+7%
Pentanal + Vinyl Acetate	(6.39±0.27)·10 ⁹	+26%	-88%
Butanoic Acid	(1.13±0.09)·10 ¹⁰	-23%	-88%
Ethylbenzene + Xylene	(1.81±0.18)·10 ⁹	-21%	-73%
Trimethylbenzene + Ethyltoluene	(1.53±0.17)·10 ⁹	+10%	-73%
Benzoic Acid	(2.41±0.57)·10 ⁸	+26%	-45%
Octanal	(1.83±0.28)·10 ⁹	+20%	-45%
Total Emission	(2.87±0.08)·10¹¹	+15%	-43%

3 It can be seen (Table 3), that the net VOC emission flux of selected compounds from P2 are
 4 15% higher as compared to the VOC emission flux from P1. The observed positive net
 5 difference in the total emission between the two paints is probably best explained by the
 6 difference in the composition or formulation of the paint binders. That is, in the paint
 7 composed of the acrylic/siloxane binder, only 40% of the copolymer acrylic was replaced by
 8 siloxane, R₂SiO, that still contains an organic group, R. As a result, the observed change in
 9 the VOC emission flux between these two paints has no positive impact on indoor air quality.
 10 Therefore, the acrylic/siloxane binders are of little interest to mitigate human health risks
 11 from indoor air pollution.

12 On the other hand, the net VOC emission flux of selected compounds from P3 under UV
 13 irradiation was 43% lower as compared to VOC emissions from P1. As shown in Table 3,
 14 majority of surface emission fluxes of selected VOCs from P3 are lower as compared to the

1 VOC emissions from P1 except for propionic acid (+198%), and acetic acid (+28%). The
2 limited emission VOC fluxes are likely due to the organic compound that is still present in the
3 mineral silicate binder. In other cases, we observed a large decrease of up to -88% for
4 pentanal and butanoic acid, as compared to P1. It may be supposed that if the photocatalytic
5 paints are formulated without any organic compounds mixed within the binder, the VOCs
6 emissions will be eliminated. However, to date, it is not possible to produce paints without
7 any organic matter present in the binder. It is necessary to have at least 5% of the organic
8 binder to obtain a good drying on the given surface support. Nevertheless, this mineral binder
9 seems to be an important innovation and a possible mitigation strategy to improve the indoor
10 air quality by considerable reduction in the VOC emissions from paints containing mineral
11 binder as compared to more traditional organic paints used in the past.

12 **3.4. Environmental Implications**

13 Sources of indoor VOCs are numerous that include emissions from building materials (wood,
14 PVC pipes), combustion processes (e.g. cooking), furnishing (carpets³⁷, floor coverings³⁸) or
15 utilization of household cleaning products³⁹⁻⁴¹.

16 Based on the experimental results obtained in this work, we estimated the formation of a
17 harmful indoor air pollutant, i.e., formaldehyde in a model room. For the model calculations,
18 we considered a room that is 2.5 m high, 5 m wide and 4 m long; hence a total volume of 50
19 m³ with floor covering³⁸, carpets³⁷, sofa⁴², furniture⁴³. In order to simulate more realistic
20 living conditions, we considered that only 4 m² of the wall is irradiated with direct sunlight of
21 8.5 W m⁻² for in the wavelength region between 300 and 400 nm^{26,44} at 298 K and 40% RH.

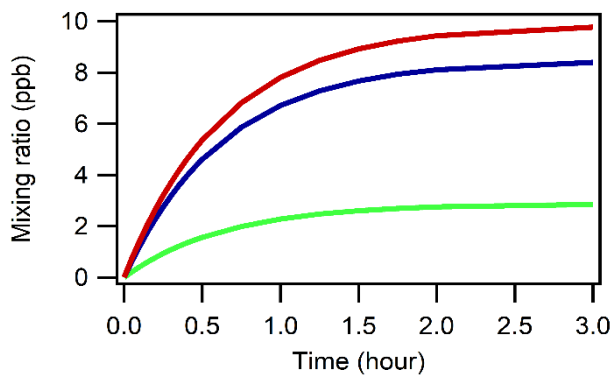
22 The elimination rate constant (k_{obs}) corresponds to the sum of air exchange rate and the
23 removal process rate. Considering an average air exchange rate of 0.56 h⁻¹ (45) and removal

1 process rate of 1.01 h^{-1} ⁽³³⁾, we obtained an elimination rate constant of 1.57 h^{-1} . The removal
 2 process encompasses the reactivity, surface adsorption, phase change and decomposition.
 3 Moreover, we assumed a constant production of formaldehyde under direct light irradiation.
 4 The formation of formaldehyde is calculated using the following equation³³:

$$C_a(t) = \frac{E_R}{k_{obs}V} + \left(C_0 - \frac{E_R}{k_{obs}V} \right) e^{(-k_{obs}t)}$$

5 where $C_0 = 0$, is the initial formaldehyde concentration (mg m^{-3}), $C_a(t)$ is the formaldehyde
 6 concentration at time (mg m^{-3}), E_R is the formaldehyde mass emission rate induced by
 7 photocatalytic paint (mg h^{-1}), k_{obs} is the elimination rate constant (h^{-1}) and V is the total
 8 volume of the room (50 m^3).

9 In Figure 4, we can see the calculated formaldehyde emission mixing ratio within the room
 10 from various photocatalytic paints in the presence of UV light.



11
 12 Figure 4: The steady state mixing ratios of formaldehyde released in the room by P1 (blue line), P2 (red line),
 13 and P3 (green line), under realistic environmental conditions.

14 Model calculation shows that the total formaldehyde mixing ratio increases with time
 15 reaching a steady state value of 10 ppb for P2, 8 ppb for P1 and 3 ppb for P3 after 3 hours.
 16 When the organic paint is replaced by the mineral silicate binder, formaldehyde emissions are
 17 calculated to decrease by a factor of 2.7 to 3.4. In comparison, model calculation of
 18 formaldehyde emission within a typical furnished room equipped from with coffee table,

1 dining table, seats and an eco-design sofa were estimated to be 18 ppb²³. Consequently, such
2 typical living room furnishings appear to be a weak source of formaldehyde in a real-life
3 indoor environment. Under UV irradiation, released mixing ratio of formaldehyde from
4 photocatalytic paint with mineral binder are close to those from reference paints (0.8 ppb)²³.

5 The emission of formaldehyde induced by different binders is rather low compared to the
6 threshold value of 100 $\mu\text{g m}^{-3}$ (80 ppb in the U.S standard conditions for temperature and
7 pressure) established by the Indoor Air Quality Guidelines (IAQG)⁴⁶ for short term exposure.

8 The mixing ratio of formaldehyde obtained by the P3 is 2.7 to 3.4 times lower than other
9 binders and 27 times lower than IAQG value implying its suitability to not degrade the indoor
10 air quality.

11 In addition to formaldehyde, P3 reduce considerably the surface emission fluxes for vinyl
12 acetate and ethylbenzene, potential human carcinogens⁴⁷.

13 **4. CONCLUSION**

14 Two new binders, organic binder with an acrylic/siloxane and mineral silicate paints were
15 developed and compared with an organic binder (100% acrylic). These paints have been
16 studied using a flow tube reactor coupled with a HS-PTR-MS instrument for real-time
17 monitoring of released VOCs. Surface emission fluxes were measured for seventeen VOCs
18 under given experimental conditions. First, the emissions of selected VOCs by reference paint
19 were compared under “dark” condition and UV irradiation. The emissions fluxes are observed
20 to be similar for the three binders under “dark” conditions and UV irradiation concluding that
21 the composition of the binder does not affect the formation of VOCs. Second, the emissions
22 for photocatalytic paints, containing nano-sized TiO_2 particles, were compared under “dark”
23 conditions. It seems that the mineral binder released more VOCs compared to organic binders,
24 but this difference is relatively low compared to the difference in VOC emissions under UV

1 irradiation. A comparative study between photocatalytic and reference paints under “dark”
2 conditions revealed that the VOCs emissions are in the same range. The presence of nano-
3 TiO₂ in the binder has a slight effect on VOCs emissions under “dark” conditions. Finally, the
4 comparison of the selected VOCs emissions from acrylic/siloxane and mineral binders and an
5 acrylic binder revealed that the paints with the acrylic/siloxane binder exhibited slightly
6 higher surface emission fluxes as compared with the acrylic binder. Thus, this binder did not
7 improve the properties of the photocatalytic paints with respect to the VOCs emissions.

8 However, the photocatalytic paints containing the mineral silicate binder exhibited very low
9 VOCs emissions in the presence of UV irradiation. For pentanal and butanoic acid, a decrease
10 in emissions of 88% was observed as compared to the emissions from paints formulated using
11 the acrylic photocatalytic binder. In fact, the use of mineral binder can reduce the quantity of
12 total VOCs emission for about 43% compared to acrylic binder. Moreover, the model
13 simulation of formaldehyde’s levels in a real-life indoor environment demonstrated a low
14 formation potential for this compound, implying the lower toxicity of the mineral binder-
15 based paint compared to the all other paints.

16 For this reason, it can be stated that the development of a mineral binder is an important
17 innovation for reducing the indoor air pollution. In the future, new experiments focused on
18 this binder will be performed to determine its capacity in improving the indoor air quality as a
19 function of time.

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24

1 **ASSOCIATED CONTENT**

2 **Supporting Information**

3 **Notes**

4 The authors declare no competing financial interest.

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9 **ASSOCIATED CONTENT**

10 **Supporting Information**

11 Experimental set up (Figure S1), Values of k , X_r and C (Table S1), Emission surface fluxes
12 for methanol under different conditions (Figure S2), comparison on surface emission fluxes
13 for photocatalytic and reference paints under “dark” conditions (Figure S3) and surface
14 emission fluxes for different VOCs (Table S2-S7).

15 **Notes**

16 The authors declare no competing financial interest.

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