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Electrosynthesis and X-ray crystallographic structure of Zn(II) *meso*-triaryltriphenylphosphonium porphyrin and structural comparison with Mg(II) *meso*-triphenylphosphonium porphine

Abdou K. D. Dimé,^[a,b] H  l  ne Cattey,^[a] Dominique Lucas^[a] and Charles H. Devillers^{*[a]}

Abstract: The electrochemical oxidation of Zn(II) 5,15-bis(*p*-tolyl)-10-phenylporphyrin (**1-Zn**) in the presence of triphenylphosphine (PPh₃) leads to the *meso*-substituted triphenylphosphonium porphyrin, coordinated by one P(O)Ph₃ molecule (**1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃**) in good yield (84%). This cationic porphyrin was characterized by NMR, MALDI-TOF mass spectrometry, UV-Vis. absorption spectroscopy and X-ray diffraction analyses. The molecular structure of this compound was compared with already published *meso*- and β -triphenylphosphonium porphyrins but also with an unpublished X-ray crystallographic structure of Mg(II) 5-triphenylphosphoniumporphine (**MgP-PPh₃⁺,PF₆⁻**) previously obtained by electrosynthesis.

Introduction

Porphyrins have been the subject of intense researches in numerous fields due to their essential role in natural processes such as photosynthesis and oxygen transport in blood.^[1] These macrocycles find applications in material science, (electro)catalysis and energy converting devices.^[2] Considerable efforts have been devoted to the functionalization of the porphyrin periphery in β and/or *meso* position(s) with halides, heteroatoms, metals or with other C-C coupled (hetero)aromatic fragments, in order to provide novel functionalities. Among the different heteroatoms, phosphorus-based porphyrins are relatively scarce in the literature, as very recently reviewed by Matano.^[3] Four phosphorus-based functional groups exist: phosphino (R₂P-), phosphonio (R₃P⁺-), phosphoryl (R₂P(E)- with R = aryl, alkoxy, E = O, S, Se) and phosphono ((HO)₂P(O)-). The introduction of the phosphorus atom is mainly performed using Pd(II) or O₂^[4] or Cu(I)^[4h, 5] coupling with "activated" (-OTf, -I, -Br) β and/or *meso* substituted porphyrins. Recently, Osuka and co-workers synthesized a *meso*-substituted diphenylphosphino-porphyrin nickel(II) complex *via* an aromatic nucleophilic substitution of *meso*-chloroporphyrin with LiPPh₂.^[4i] Among these phosphorus-

porphyrins lies on the chemical/electrochemical oxidation of the porphyrin macrocycle in the presence of a tertiary phosphine. These porphyrins may find medical applications as tumor imaging agents and mitochondrial targeting molecules as arylphosphonium derivatives have both lipophilic and cationic character. This feature allows for their facile transport through plasma membranes or cell walls to accumulate in the cytoplasm or cell mitochondria.^[6] In a pioneer work, Smith and co-workers reported that triphenylphosphine (PPh₃) can react as a nucleophile in *meso* position with the cation radical of the Zn(II) octaethylporphyrin complex (ZnOEP) preliminary generated by oxidation with magic blue (tris(4-bromophenyl)ammonium hexachloroantimonate)^[7] while Shine and co-workers^[8] described the β -position reactivity of the zinc(II) tetraphenylporphyrin (ZnTPP) cation radical (chemically generated by oxidation with I₂/AgClO₄) with PPh₃. Later, Latos-Grazyński and co-workers reported the chemical β -functionalization of preformed (TPP)Fe^{III}(ClO₄)₂ with PPh₃ that afforded the [(β -PPh₃⁺-TPP)Fe^{III}]₂O(ClO₄)₂ dimer.^[9] In all cases, *meso*- and β -phosphonium-based porphyrins were generated in modest yields (\leq 37.7%). A significant progress in the yield of phosphonium-based porphyrins was achieved by Giraudeau *et al.* when they continuously generated the porphyrin π -cation radical in presence of PPh₃ by controlled potential electrolysis. In these conditions, β -PPh₃⁺-ZnTPP,ClO₄⁻ was generated in 50% yield.^[10] The same group further exploited the electrogenerated ZnTPP cation radical reactivity with a bisphosphine base (bis(diphenylphosphino)acetylene) providing the mono phosphonium porphyrin and/or the bisphosphonium porphyrin dimer in fair to good yield, depending on the amount of the nucleophile.^[11] This work was extended to other bis- and trisphosphine nucleophiles leading to bis(β -phosphonium)-ZnTPP dimers and tris-(β -phosphonium)-ZnTPP trimers in good yield (60-93%).^[12] As part of our on-going interest focused on the redox reactivity of porphine, the fully unsubstituted porphyrin,^[13] we functionalized magnesium(II) porphine (**MgP**) with PPh₃ by controlled potential electrolysis and obtained the *meso*-triphenylphosphonium porphyrin (**MgP-PPh₃⁺,PF₆⁻**) in good yield (78%). This latter was the first X-ray crystallographic structure example of *meso*-phosphonium porphyrin.^[14] In this particular case, a large excess of phosphine was necessary to prevent the competitive **MgP** oxidative polymerization from occurring.^[15] In order to decrease the propensity to polymerize and to favor the anodic nucleophilic substitution reaction, more recently we synthesized the 5,15-bis(*p*-tolyl)-10-phenylporphyrin (**1-H₂**) which has only one available free *meso* position. The electrochemical oxidation in the presence of PPh₃ of the corresponding Ni(II)

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based porphyrins, the sole synthetic pathway to phosphonio-

complex (**1-Ni**) led to the *meso*-PPh₃⁺-porphyrin (**1-Ni-PPh₃⁺,PF₆⁻**) in 72% yield.^[16]

In this manuscript will be presented the extension of this work to the Zn(II) complex, **1-Zn**, providing the novel *meso*-PPh₃⁺-porphyrin (**1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃**) in good yield. The electrolysis is monitored by cyclic voltammetry and UV-Vis. absorption spectroelectrochemistry. X-ray crystallographic analyses reveal that the zinc cation is O-coordinated by one triphenylphosphine oxide molecule. Comparison of this X-ray crystallographic structure with previously published (**1-Ni-PPh₃⁺,PF₆⁻**)^[16] and **MgP-PPh₃⁺,PF₆⁻** (unpublished (**MgP-PPh₃⁺,PF₆⁻**, other crystallization conditions) *meso*-triphenylphosphonium porphyrins molecular structures will be presented.

Results and Discussion

Electrochemical investigations

Cyclic voltammetry

The cyclic voltammogram (CV) of **1-Zn** in CH₂Cl₂/CH₃CN mixture (4/1 v/v) 0.1 M TEAPF₆ (solid blue line, Fig. 1) exhibits one monoelectronic reversible oxidation leading to the cation-radical at $E_{1/2}$ (O1/R1) = 0.86 V ($\Delta E_p = 75$ mV) and one reversible monoelectronic reduction leading to the radical anion species $E_{1/2}$ (O2/R2) = -1.40 V ($\Delta E_p = 90$ mV, all potentials given in this manuscript will be given vs. the saturated calomel electrode (SCE); in these experimental conditions, ferrocene is oxidized at 0.40 V).^[17]

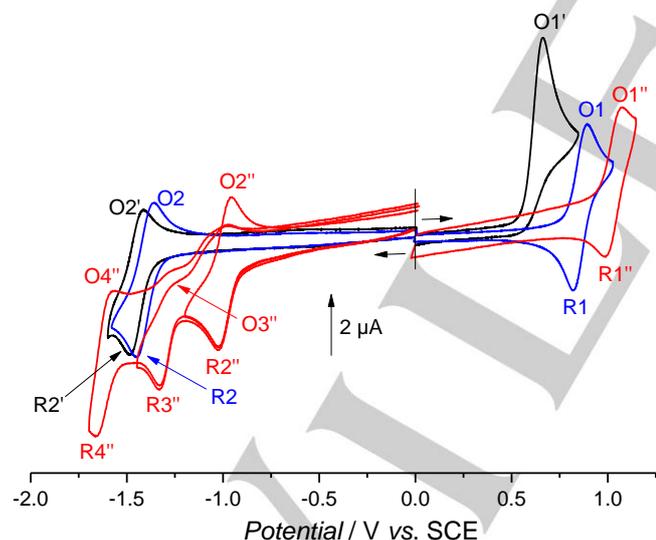


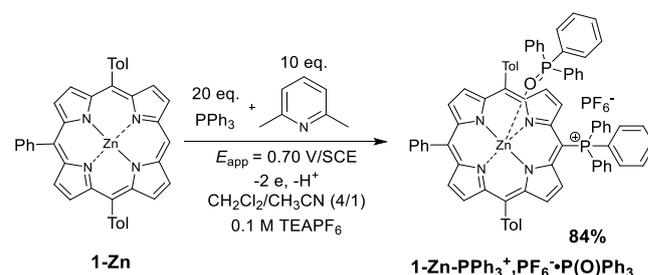
Figure 1. Cyclic voltammograms of **1-Zn** with (solid black line) or without (solid blue line) 20 eq. of PPh₃ and **1-Zn-PPh₃⁺,PF₆⁻** with 20 eq. of PPh₃ (solid red line) in CH₂Cl₂/CH₃CN (4/1 v/v) 0.1 M TEAPF₆. Concentration: 5×10^{-4} M; WE: Pt $\varnothing = 2$ mm, $\nu = 100$ mV s⁻¹.

The second reduction stage of the porphyrin ring is not observed since it falls beyond the solvent reduction. Indeed, when 20 eq. of

PPh₃ are added to the solution, the first oxidation process ($E_{pa}(O1') = 0.66$ V) becomes irreversible with a two-fold more intense peak current than the one observed for the reversible reduction ($E_{1/2}(O2'/R2') = -1.45$ V, $\Delta E_p = 70$ mV). Moreover, peak O1' is shifted towards less positive potential ($E_{pa}(O1) - E_{pa}(O1') = -0.23$ V). This behavior arises from two phenomena: 1) the initially generated cation radical **1-Zn⁺** undergoes a fast chemical reaction with PPh₃ followed by a subsequent one electron uptake (ECE-type mechanism)^[14, 18]; 2) the coordination of PPh₃ to the Zn(II) cation which increases the porphyrin electronic density and makes it easier to oxidize and more difficult to reduce, as it can be seen on the first reduction peak (compare peaks R2 and R2'). The CV of a pure sample of **1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃** in the same experimental conditions (*i.e.* with 20 eq. PPh₃) exhibits a monoelectronic reversible oxidation leading to the cation-radical ($E_{1/2}(O1''/R1'') = 1.03$ V, $\Delta E_p = 90$ mV). In the negative potential direction, the first monoelectronic reversible peak ($E_{1/2}(O2''/R2'') = -0.99$ V, $\Delta E_p = 70$ mV) may be attributed to the first reduction of the porphyrin ring since the potential difference between the first oxidation and reduction (HOMO–LUMO gap, $\Delta E_{1/2} = 2.02$ V) fall within the empirical value of 2.15 ± 0.15 V previously reported for octaethylporphyrin complexes.^[19] The CV exhibits two others reductions at $E_{pc}(R3'') = -1.33$ V and $E_{pc}(R4'') = -1.66$ V which plausibly correspond to the reduction of the phosphonium and the porphyrin ring moieties, respectively.

Electrosynthesis

Electrolyses were carried out under argon, at room temperature, in a CH₂Cl₂/CH₃CN mixture (4/1 v/v) containing 0.1 M of TEAPF₆, **1-Zn**, 20 eq. of PPh₃ and 10 eq. of 2,6-lutidine. This non-nucleophilic base is added to trap the protons released in the course of the reaction; otherwise demetalation of Zn(II) porphyrin could occur. Starting with 20.0 mg of **1-Zn** the potential was set at $E_{app} = 0.70$ V. The current progressively decreased and finally reached the residual current value when 2.4 F per mol of **1-Zn** were transferred. After removal of the supporting electrolyte, purification of the crude product on silica gel and crystallization from toluene, the *meso*-triphenylphosphonium porphyrin **1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃** was isolated in 84% yield (Scheme 1).



Scheme 1. Electrochemical synthesis of **1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃**.

We tried to diminish the amount of phosphine nucleophile up to two equivalents. Though the reaction was still efficient, the competing oxidative *meso-meso* dimerization process was observed on the UV-Vis. absorption spectrum, as attested by the

characteristic Soret band of the *meso-meso*-linked dimer **2-Zn** at 459 nm.^[17a] The presence of coordinated phosphine oxide was unexpected. However, the purity of the initial PPh₃ was 99% and contained less than 1% triphenylphosphine oxide P(O)Ph₃. Thus, in our conditions, less than 0.2 eq. of P(O)Ph₃ is initially present in the solution to be electrolyzed. The missing P(O)Ph₃ amount may be formed during the electrolysis (oxidation of PPh₃ in presence of trace amount of water leads to P(O)Ph₃^[20]) and/or might be produced during the work-up (extraction of the supporting electrolyte with several aqueous washings under an air atmosphere).

Spectroelectrochemistry

In the course of the electrolysis, the UV-Vis. absorption spectrum of the reacting solution was monitored. It was not possible to see the entire Soret band at this concentration given its high molar extinction coefficient. However, it was observed that the Soret band foot underwent a progressive bathochromic shift, in agreement with the functionalization of the porphyrin by an electrowithdrawing group (Fig. 2).^[21] Noteworthy, in CH₂Cl₂ the Soret band of the pure **1-Zn** and **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃** appears at 414 nm and 432 nm, respectively.

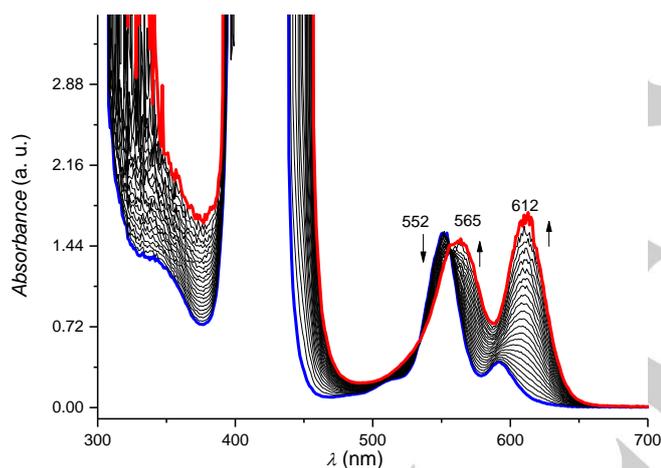


Figure 2. Electrolysis of a 5×10^{-4} M solution of **1-Zn** with 20 eq. PPh₃ and 10 eq. of 2,6-lutidine followed by UV-Vis. spectroscopy ($l = 2$ mm, 0.1 M TEAPF₆ in CH₂Cl₂/CH₃CN (4/1 v/v), $E_{app} = 0.70$ V vs. SCE, -2.4 F, WE: Pt wire, WE and counter-electrode are in separated compartments).

The initial Q band of **1-Zn** ($\lambda_{max} = 552$ nm) progressively disappears at the expense of two new bands at 565 and 612 nm. The presence of an isosbestic point located at 557 nm confirms the simple transformation of the initial porphyrin into another well-defined product, **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃**, with no intermediate observed between.

NMR and Maldi-Tof mass spectrometry characterization of **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃**

From 1D and 2D NMR analyses, the isolated compound was unambiguously identified as the *meso*-substituted triphenylphosphonium porphyrin coordinated at the Zn(II) center

by one triphenylphosphonium oxide (P(O)Ph₃) molecule (Fig. 3). In particular, the presence of four β -pyrrolic doublets between 8.7 and 8.2 ppm (H_d, H_e, H_i and H_j) as well as the disappearance of the singlet signal corresponding to the *meso* proton confirms the functionalization of the porphyrin ring in *meso* position. Moreover, the global integration value (28 H) for the signals located within the common chemical shifts of the phenyl/tolyl groups (7.4-8.0 ppm) is in agreement with the mono-*meso*-substitution.

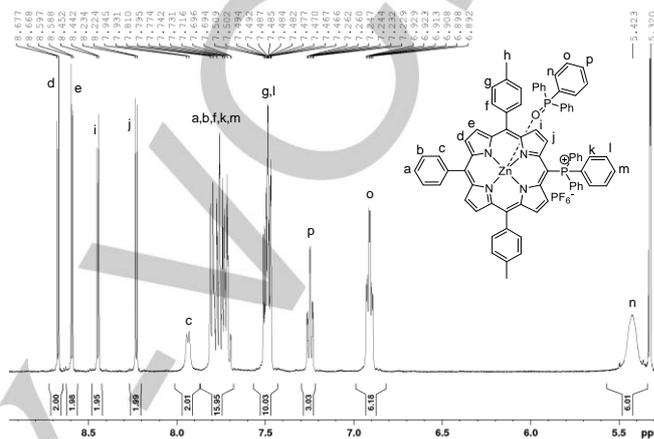


Figure 3. Partial ¹H NMR spectra of **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃** (300 MHz, 298 K, CD₂Cl₂).

As compared to the initial **1-Zn** porphyrin,^[17a] an important shielding of the β -pyrrolic proton signals H_i (from 9.39 to 8.23 ppm) and H_j (from 9.09 to 8.44 ppm) is observed. This behavior confirms the substitution at the *meso* position by the triphenylphosphonium group since H_i and H_j are under the influence of the shielding cone of the phenyl groups from the phosphonium moiety. Three unexpected signals located between 7.3 and 5.3 ppm are attributed to the aromatic protons from an O-coordinated P(O)Ph₃ molecule. As a consequence of the strong influence of the shielding cone of the porphyrin ring, these proton signals (H_n, H_o, H_p) are also highly shielded. In the ³¹P NMR spectrum, three signals are observed at 26.5, 23.9 and -144.5 ppm for the phosphorus atoms of P(O)Ph₃, -PPh₃⁺ and PF₆⁻, respectively (see ESI).

Analysis of **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃** by MALDI-TOF mass spectrometry confirms the functionalization of the porphyrin periphery by the phosphonium group. However, the initially coordinated P(O)Ph₃ ligand and the PF₆⁻ anion are not seen on the mass spectrum (see ESI).

X-ray Crystallography

A definite proof of the **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃** molecular structure was given from the X-ray diffraction analysis. Suitable crystals were obtained by slow diffusion of *n*-hexane into a concentrated CH₂Cl₂ solution of **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃**. The three-dimensional molecular views of **1-Zn-PPh₃⁺,PF₆⁻-P(O)Ph₃** are presented in Fig. 4 and Fig. 5. The zinc(II) atom is coordinated by one apical triphenylphosphine oxide molecule ($d(\text{Zn-O1}) = 2.010(4)$ Å). This Zn-O distance is slightly shorter than the one

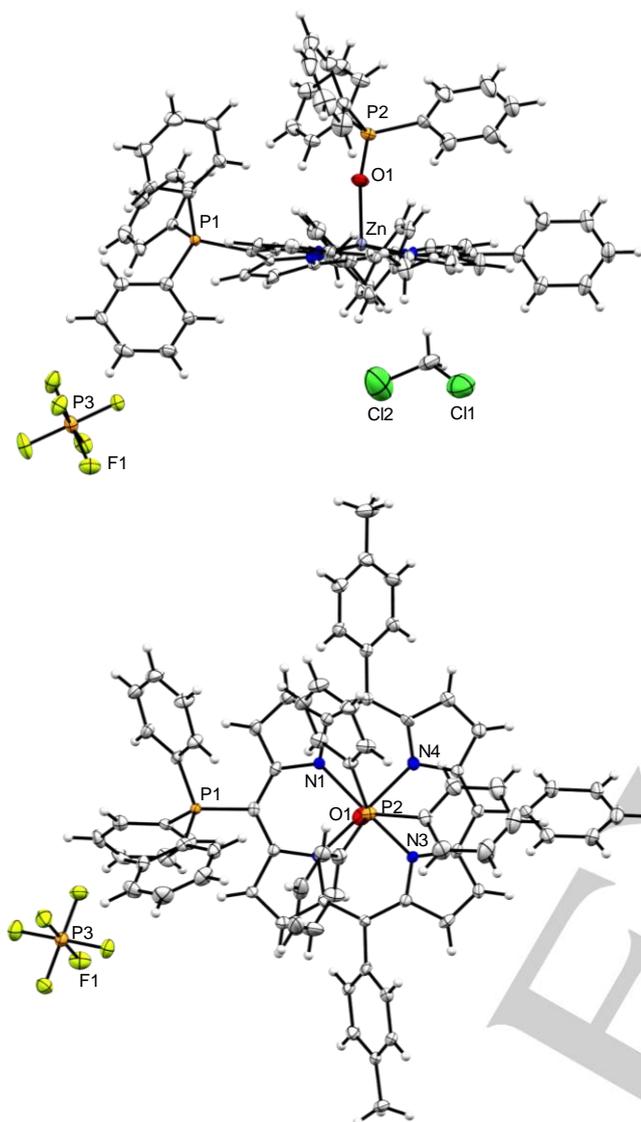


Figure 4. Side (top) and front (bottom) Mercury views of **1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃·CH₂Cl₂** X-ray crystallographic structure. CH₂Cl₂ molecule was omitted on the front view for clarity. Thermal ellipsoids are scaled to the 50% probability level.

observed for the one-dimensional coordination polymer [10,20-diphenylporphyrinatozinc(II)-5,15-diyl]-bis-[P(O)Ph₂]_n·(1.5H₂O)_n ($d(\text{Zn}-\text{O}) = 2.0224(14) \text{ \AA}$) reported by Arnold and coworkers.^[4a] The P2-O1 distance ($d(\text{P2}-\text{O1}) = 1.490(4) \text{ \AA}$) is very close to the one reported in reference [4a] ($d(\text{P}-\text{O}) = 1.4899(17) \text{ \AA}$) and the Zn-O1-P2 angle ($165.5(3)^\circ$) is significantly lower than the one found in reference [4a] ($177.56(9)^\circ$). The presence of one PF₆⁻ anion per porphyrin molecule confirms the cationic nature of the macrocycle. The CH₂Cl₂ solvent used for the crystallization process is also found in the crystal lattice and it interacts with the porphyrin via C-H... π interaction ($d(\text{C77}-\text{H77B}\cdots\text{centroid}(\text{N4}-\text{C16}-\text{C15}-\text{C14}-\text{C13})) = 3.496(11) \text{ \AA}$; $\text{C77}-\text{H77B}\cdots\text{centroid}(\text{N4}-\text{C16}-\text{C15}-\text{C14}-\text{C13}) = 171.90(\text{X})^\circ$). Each porphyrin interacts with six

PF₆⁻ anions via C...F and H...F interactions ($d(\text{C64}-\text{H64}\cdots\text{F1}) = 3.170(7) \text{ \AA}$; $d(\text{C63}-\text{H63}\cdots\text{F4}) = 3.304(8) \text{ \AA}$; $d(\text{C62}\cdots\text{F1}) = 3.127(8) \text{ \AA}$; $d(\text{C62}-\text{H62}\cdots\text{F2}) = 3.207(9) \text{ \AA}$; $d(\text{C61}-\text{H61}\cdots\text{F5}) = 3.398(8) \text{ \AA}$; $d(\text{C74}\cdots\text{F4}) = 3.091(8) \text{ \AA}$; $d(\text{C23}\cdots\text{F5}) = 3.169(8) \text{ \AA}$; $d(\text{C24}-\text{H24}\cdots\text{F3}) = 3.257(9) \text{ \AA}$; $d(\text{C34}-\text{H34}\cdots\text{F6}) = 3.231(7) \text{ \AA}$; $d(\text{C29}-\text{H29}\cdots\text{F2}) = 3.138(9) \text{ \AA}$; $d(\text{C30}-\text{H30}\cdots\text{F3}) = 3.301(7) \text{ \AA}$; $d(\text{C30}-\text{H30}\cdots\text{F4}) = 3.511(7) \text{ \AA}$).

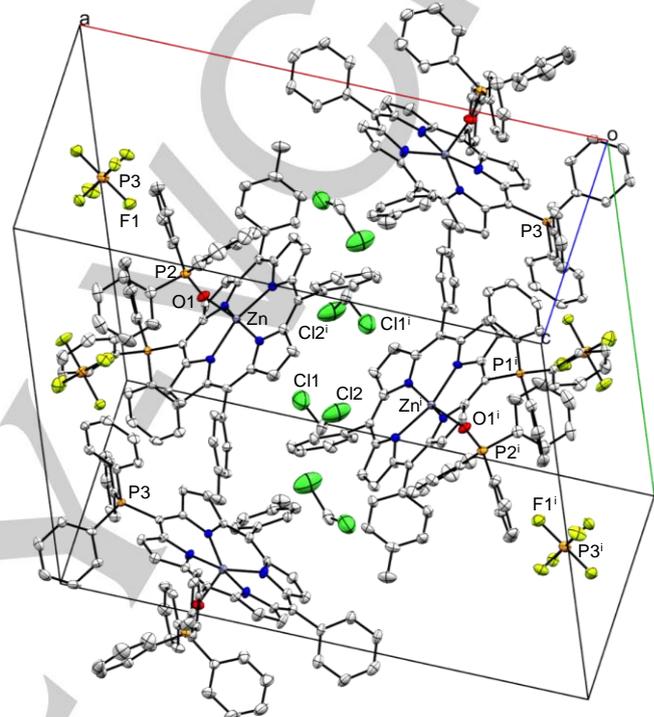
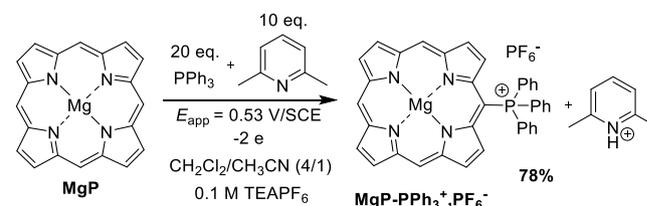


Figure 5. Unit cell of **1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃·CH₂Cl₂**. H atoms are omitted for clarity. Thermal ellipsoids are scaled to the 50% probability level. [Symmetry code: (i) 1-x, 1-y, 1-z].

Interestingly, these interactions only occur with the phenyl groups from the triphenylphosphine and triphenylphosphine oxide and not with the phenyl/tolyl moieties of the porphyrin ring. Previously, we synthesized in good yield the *meso*-triphenylphosphonium magnesium(II) porphyrin **MgP-PPH₃⁺,PF₆⁻** by electrochemical oxidation of magnesium(II) porphyrin (**MgP**) in presence of PPh₃, according to Scheme 2.^[14]



Scheme 2. Electrochemical synthesis of **MgP-PPH₃⁺,PF₆⁻** as reported in reference ^[14].

FULL PAPER

Contrary to $1\text{-Zn-PPh}_3^+\text{PF}_6^-\text{P(O)Ph}_3$, coordination of P(O)Ph_3 molecule was not observed on the ^1H and ^{31}P NMR spectra despite the fact that P(O)Ph_3 is known to strongly coordinate magnesium(II) porphyrin complexes.^[4c] As $\text{MgP-PPh}_3^+\text{PF}_6^-$ was generated at lower potential than $1\text{-Zn-PPh}_3^+\text{PF}_6^-\text{P(O)Ph}_3$ (0.53 V vs. 0.70 V), PPh_3 is plausibly not oxidized to P(O)Ph_3 in these conditions. The X-ray crystallographic structure was solved and toluene and ethanol coming from the crystallization solvents were found in the crystal lattice.^[14] Using other crystallization conditions, *i.e.* slow diffusion of diethylether into a concentrated EtOH solution of $\text{MgP-PPh}_3^+\text{PF}_6^-$, suitable crystals of $\text{MgP-PPh}_3^+\text{PF}_6^-\text{EtOH}\cdot\text{H}_2\text{O}\cdot 2\text{Et}_2\text{O}$ were obtained and this unreported X-ray crystallographic structure is presented in this manuscript for comparison purpose (Fig. 6 and Fig. 7).

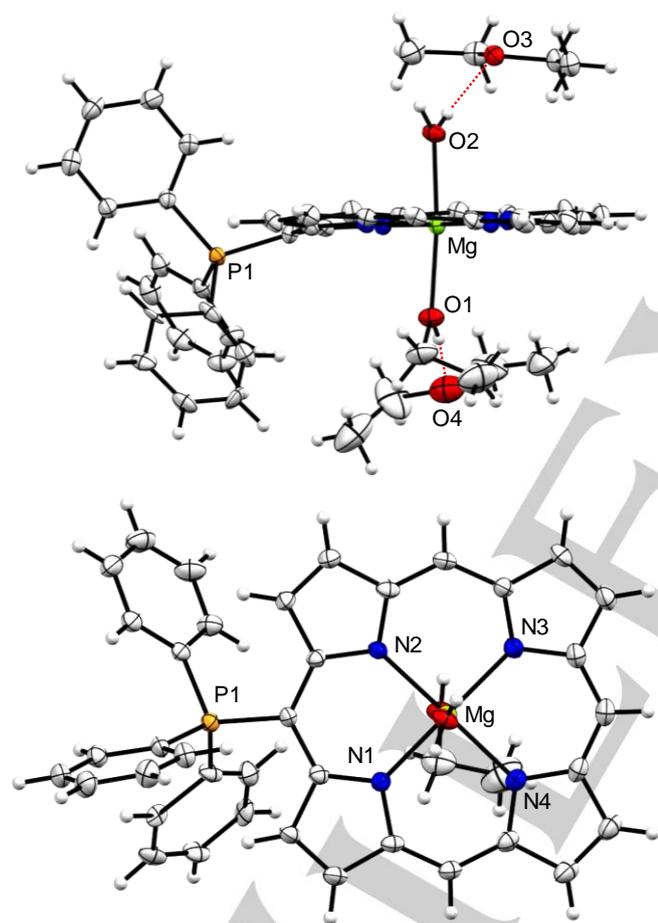


Figure 6. Side (top) and front (bottom) Mercury views of $\text{MgP-PPh}_3^+\text{PF}_6^-\text{EtOH}\cdot\text{H}_2\text{O}\cdot 2\text{Et}_2\text{O}$ X-ray crystallographic structure. PF_6^- anion is omitted in these views and both Et_2O molecules are omitted in the front view for clarity. Thermal ellipsoids are scaled to the 50% probability level.

The magnesium(II) atom is axially coordinated by one H_2O ($d(\text{Mg}-\text{O}2) = 2.145(3) \text{ \AA}$) and one EtOH molecules ($d(\text{Mg}-\text{O}1) = 2.190(4) \text{ \AA}$). Two porphyrins related by an inversion center form an H-bonded dimer involving two H_2O ligands and two Et_2O molecules

($d(\text{O}2-\text{H}2\text{A}\cdots\text{O}3) = 2.842(4) \text{ \AA}$, $d(\text{O}2-\text{H}2\text{B}\cdots\text{O}3^i) = 2.812(4) \text{ \AA}$, $\text{O}2-\text{H}2\text{A}-\text{O}3 = 176.3(2)^\circ$, $\text{O}2-\text{H}2\text{B}-\text{O}3^i = 176.5(2)^\circ$, Fig. 7). The EtOH axial ligand also exhibits an hydrogen bond with an Et_2O molecule ($d(\text{O}1-\text{H}1\cdots\text{O}4) = 2.806(7) \text{ \AA}$, $\text{O}1-\text{H}1-\text{O}4 = 174(5)^\circ$). The presence of one PF_6^- anion per porphyrin molecule confirms the cationic nature of the macrocycle. Each porphyrin interacts with four PF_6^- anions via $\text{H}\cdots\text{F}$ interactions. As for $1\text{-Zn-PPh}_3^+\text{PF}_6^-\text{P(O)Ph}_3$, three interactions between protons from the phenyl rings of the triphenylphosphonium and the fluoride atoms are observed ($d(\text{C}29-\text{H}29\cdots\text{F}1) = 3.340(5) \text{ \AA}$; $d(\text{C}35-\text{H}35\cdots\text{F}3) = 3.282(5) \text{ \AA}$; $d(\text{C}35-\text{H}35\cdots\text{F}6) = 3.393(5) \text{ \AA}$). Besides, two additional $\text{H}\cdots\text{F}$ interactions exist between *meso*-hydrogen atoms of the porphyrin ring and the hexafluorophosphate anion ($d(\text{C}15-\text{H}15\cdots\text{F}3) = 3.366(5) \text{ \AA}$, $d(\text{C}10-\text{H}10\cdots\text{F}1) = 3.449(5) \text{ \AA}$). Due to the bulkiness of the $-\text{PPh}_3$ group, the *meso* $\text{C}-\text{P}^+$ bond ($\text{C}20-\text{P}1 = 1.785(4) \text{ \AA}$) clearly bends towards the *exo*-position of the dimer and the P atom stands $0.783(2) \text{ \AA}$ above the 24 atom mean plane of the porphyrin ring.

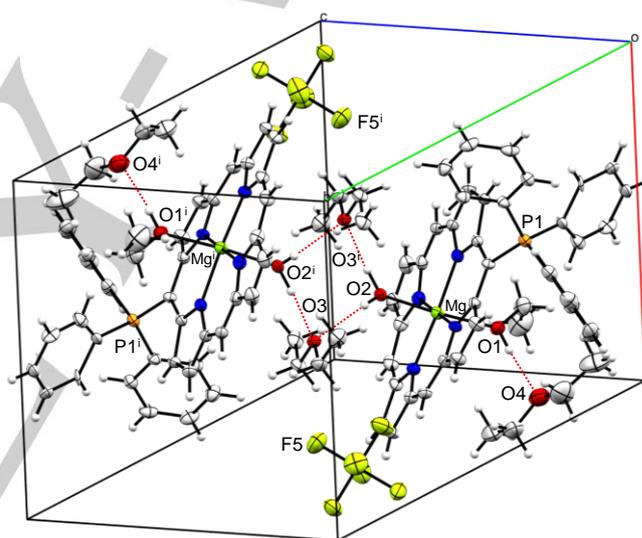


Figure 7. Unit cell of $\text{MgP-PPh}_3^+\text{PF}_6^-\text{EtOH}\cdot\text{H}_2\text{O}\cdot 2\text{Et}_2\text{O}$. Thermal ellipsoids are scaled to the 50% probability level. [Symmetry code: (i) $1-x, 1-y, 1-z$].

Selected parameters for these novel crystallographic structures are compared with both existing *meso*-triphenylphosphonium Mg(II) and Ni(II) porphyrins^[14, 16] and the unique β -substituted triphenylphosphonium Fe(III) porphyrin reported example^[9] (Table 1). For these five X-ray crystallographic structures the metal-nitrogen distances are similar for the Zn(II) and Mg(II) complexes ($2.063(3) \leq d(\text{M}-\text{N}) \leq 2.095(4) \text{ \AA}$), slightly lower for the Fe(III) complex ($2.042(6) \leq d(\text{M}-\text{N}) \leq 2.064(6) \text{ \AA}$) and much lower for the Ni(II) complex ($1.906(5) < d(\text{M}-\text{N}) < 1.919(5) \text{ \AA}$). The $\text{C}_{\text{porphyrin}}-\text{P}^+$ bond lengths are all very similar ($1.785(4) \leq d(\text{C}_{\text{porphyrin}}-\text{P}^+) \leq 1.809(6) \text{ \AA}$) and correspond to common bond lengths of aryl-triphenylphosphonium derivatives.^[22] The RMS deviation, which is a structural parameter which allows for measuring the distortion magnitude of a macrocycle, was calculated for the Mg(II), Zn(II) and Ni(II) complexes. In the case of porphyrins, the RMS

deviation corresponds to the root mean square of the distances of the 24 carbon and nitrogen porphyrin ring atoms from the mean plane formed by these atoms. The RMS deviation is low for both Mg(II) complexes (0.113 and 0.115 Å for **MgP-PPh₃⁺,PF₆⁻·EtOH·H₂O·2Et₂O** and **MgP-PPh₃⁺,PF₆⁻·EtOH·toluene** respectively) while it increases for the Zn(II) porphyrin (0.234 Å) and is much higher for the Ni(II) complex (0.407 Å), highlighting its marked deviation from the planarity.

Table 1. Selected structural parameters for the *meso* and β triphenylphosphonium-porphyrins.

Structural parameters (Å)	1-Zn-PPh₃⁺,P F₆⁻·P(O)PPh₃ h₃·CH₂Cl₂	MgP-PPh₃⁺,P F₆⁻·EtOH·H₂O O ·2Et₂O	MgP-PPh₃⁺,P F₆⁻·EtOH·toluene [14]	1-Ni-PPh₃⁺,P F₆⁻ [16]	Fe(III) β-PPh₃⁺-TPPCl₂ [9]
$d(M-N)$	2.069(4) < d < 2.091(5)	2.063(3) < d < 2.081(3)	2.072(3) < d < 2.095(4)	1.906(5) < d < 1.919(5)	2.042(6) < d < 2.064(6)
$d(C_{\text{porphyrin}}-P^+)$	1.804(5)	1.785(4)	1.794(3)	1.809(6)	1.799(9)
$d(P^+-\text{mean plane})$ [a]	0.580	0.783(2)	0.390	1.590	not reported
RMS deviation [b]	0.234	0.113	0.115	0.407	not reported

[a] mean plane calculated from the 24 carbon and nitrogen porphyrin ring atoms.

[b] The RMS deviation corresponds to the root mean square of the distances of the 24 carbon and nitrogen porphyrin ring atoms from the mean plane formed by these atoms.

Conclusions

In summary the electrochemical oxidation of zinc(II) *5,15-p*-ditolyl-10-phenylporphyrin **1-Zn** in the presence of triphenylphosphine (PPh₃) affords, in good yield and perfect regioselectivity, *meso*-substituted triphenylphosphonium porphyrin coordinated by one triphenylphosphonium oxide molecule (**1-Zn-PPh₃⁺,PF₆⁻·P(O)PPh₃**). P(O)Ph₃ is thought to come partially from the oxidation of PPh₃ during the controlled potential electrolysis. This assertion is strengthened by the fact that the more easily oxidized *meso*-triphenylphosphonium magnesium(II) porphine complex **MgP-PPh₃⁺,PF₆⁻** is not coordinated by P(O)Ph₃ although Mg(II) porphyrins are known to strongly interact with this ligand. The X-ray crystallographic structures of **Zn-PPh₃⁺,PF₆⁻·P(O)PPh₃·CH₂Cl₂** and **MgP-PPh₃⁺,PF₆⁻·EtOH·H₂O·2Et₂O** have been solved and exhibit

common structural features of aryl-phosphonium derivatives. This work is part of a more vast research program aiming at improving the control of the oxidative reactivity of metalloporphyrins. Further studies are in progress to improve the scope of the reaction.

Experimental Section

X-ray equipment and refinement

Diffraction data were collected on a Bruker Apex II CCD diffractometer equipped with an Oxford Cryosystems low-temperature device, operating at $T = 115$ K. Data were measured using ϕ and ω scans using MoK α radiation (X-ray tube, 50 kV, 32 mA). The total number of runs and images was based on the strategy calculation from the program APEX2.^[23] Cell parameters were retrieved using the SAINT^[23] software and refined using SAINT. Data reduction was performed using the SAINT software which corrects for Lorentz polarisation. A multi-scan absorption correction was performed using SADABS^[23] was used for absorption correction. Using Olex2,^[24] the structure was solved with the ShelXS^[25] structure solution program, using the direct phasing methods solution method. The model was refined with version 2018/3 of ShelXL^[26] using Least Squares minimisation.

A single violet prism-shaped crystal (0.15x0.15x0.10) mm³ of (**1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃·CH₂Cl₂**) was mounted on a mylar loop with oil. The structure was solved in the space group P2₁/c (# 14). All non-hydrogen atoms were refined anisotropically. Hydrogen atom positions were calculated geometrically and refined using the riding model. One solvent dichloromethane molecule was identified but is disordered over several positions. Only one component of the disorder was unambiguously localized with an occupation factor converged to 0.45. The data were corrected for the residual electron density due to the other components of the solvate molecule by the Squeeze procedure.^[27]

A single violet prism-shaped crystals (0.15x0.08x0.03) mm³ of (**MgP-PPh₃⁺,PF₆⁻·EtOH·H₂O·2Et₂O**) was mounted on a mylar loop with oil. The structure was solved in the space group P-1 (# 2). Most hydrogen atom positions were calculated geometrically and refined using the riding model, but hydrogen atoms bonded to oxygen atoms were refined using geometric restraints. One of the phenyl group was found disordered over two positions with occupation factors converged to 0.59:0.41, the minor component was isotropically refined. One ethanol and one water molecules are linked to the Mg atom in axial position. The ethanol molecule was found slightly disordered over two positions with occupation factors converged to 0.81:0.19, the minor component was isotropically refined. SADI restraints were applied on the two ether molecules.

CCDC-1848217 (**1-Zn-PPh₃⁺,PF₆⁻·P(O)Ph₃·CH₂Cl₂**) and CCDC-1848218 (**MgP-PPh₃⁺,PF₆⁻·EtOH·H₂O·2Et₂O**) contains the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Reagents and instrumentation

Porphyrin **1-Zn** was synthesized according to known procedures.^[28] Our data (¹H NMR, ¹³C NMR, UV-Vis. absorption, and MALDI-TOF mass spectrum) were consistent with those described in reference^[28] (see ESI). Tetraethylammonium hexafluorophosphate (TEAPF₆, Fluka puriss., electrochemical grade, $\geq 99.0\%$) and 2,6-lutidine (Aldrich, purified by redistillation, $\geq 99\%$) were used as received. CH₂Cl₂ (Carlo Erba 99.5%)

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and CH₃CN (SDS, Carlo Erba, HPLC gradient 99.9%) were distilled from P₂O₅ and CaH₂ respectively.

UV-Vis. absorption spectra were obtained with a Varian UV-vis spectrophotometer Cary 50 scan using quartz cells (Hellma). In spectroelectrochemical experiments, an UV-Vis. immersion probe (Hellma, $l = 2$ mm) was connected through optical fibers to the same spectrophotometer.

Mass spectra were obtained on a Bruker ProFLEX III spectrometer (MALDI-TOF) using dithranol as a matrix.

NMR spectra were measured on a BRUKER 300, 500 or 600 MHz spectrometer (Avance III Nanobay, Avance III, Avance II, respectively). The reference was the residual non-deuterated solvent.

All electrochemical manipulations were performed using Schlenk techniques in an atmosphere of dry oxygen-free argon at room temperature ($T = 20^\circ\text{C} \pm 3^\circ\text{C}$). The supporting electrolyte was degassed under vacuum before use and then dissolved to a concentration of 0.1 mol L⁻¹. Voltammetric analyses were carried out in a standard three-electrode cell, with an Autolab PGSTAT 302 N potentiostat, connected to an interfaced computer that employed Electrochemistry Nova software. The reference electrode was a saturated calomel electrode (SCE) separated from the analyzed solution by a sintered glass disk filled with the background solution. The auxiliary electrode was a platinum wire separated from the analyzed solution by a sintered glass disk filled with the background solution. For all voltammetric measurements, the working electrode was a platinum disk electrode ($\varnothing = 2$ mm). In these conditions, when operating in a mixture of CH₂Cl₂/CH₃CN 4/1 (0.1 M TEAPF₆) the formal potential for the Fc⁺/Fc couple was found to be +0.40 V vs. SCE.

Bulk electrolyses were performed in a two or three compartment cell with glass frits of medium porosity with an Amel 552 potentiostat coupled with an Amel 721 electronic integrator. A platinum wire spiral ($l = 53$ cm, $\varnothing = 1$ mm) was used as the working electrode, a platinum plate as the counter electrode and a saturated calomel electrode as the reference electrode. Electrolyses were followed by TLC and UV-Vis. absorption measurements.

Electrosynthesis of 1-Zn-PPH₃⁺,PF₆⁻P(O)Ph₃

Electrolyses were carried out under argon, at room temperature, in 40 mL of CH₂Cl₂/CH₃CN (4/1 v/v) containing 0.1 M of TEAPF₆, 20.0 mg (31.74 μmol) of **1-Zn**, 10 eq. of 2,6-lutidine (37 μl) and 20 eq. of PPh₃ (166.5 mg). The applied potential was $E_{\text{app}} = 0.70$ V vs. SCE. At the end of the electrolysis, 2.4 F per mol of **1-Zn** were transferred. The solution mixture was then evaporated to dryness under reduced pressure. The resulting crude solid was dissolved in a minimum of CH₂Cl₂ and this solution was washed with 3×250 mL of distilled water to remove the supporting electrolyte. The organic phase was evaporated to dryness. The crude product was then purified using silica gel column chromatography (CH₂Cl₂), recrystallized from toluene providing **1-Zn-PPH₃⁺,PF₆⁻P(O)Ph₃** in 84% yield ($m = 35.0$ mg). ¹H NMR (CD₂Cl₂, 500 MHz, 298 K): δ (ppm) 2.64 (s, CH₃, 6H), 5.42 (s, *o*-Ph, 6H), 6.88-6.95 (m, *m*-Ph, 6H), 7.46-7.52 (m, *p*-Ph, 3H), 7.46-7.52 (m, *m*-tol, *m*-Ph, 10H) 7.68-7.82 (m, *p*- and *m*-Ph, *o*-tol, 16H), 7.93-7.95 (m, *o*-Ph, 2H), 8.23 (d, ³*J* = 4.9 Hz, β -Pyrr, 2H), 8.45 (d, ³*J* = 4.9 Hz, β -Pyrr, 2H), 8.59 (d, ³*J* = 4.6 Hz, β -Pyrr, 2H), 8.67 (d, ³*J* = 4.6 Hz, β -Pyrr, 2H); ³¹P (CD₂Cl₂, 202 MHz, 298 K): δ (ppm) -144.5 (m, PF₆⁻), 23.9 (s, P⁺), 26.5 (s, P=O); RMN ¹³C (CD₂Cl₂, 125 MHz, 298 K): 126.6, 126.9, 127.3, 127.6, 128.2, 128.4, 128.5, 128.9, 129.1, 129.6, 129.6, 129.7, 130.4, 130.5, 130.6, 132.2, 133.7, 133.9, 134.4, 134.6, 134.7, 134.8, 138.1, 139.4, 142.3, 148.8, 150.2, 152.8, 153.6. λ_{max} (CH₂Cl₂)/nm (log ϵ)

432 (5.52), 567 (4.11), 608 (4.29). Maldi-Tof MS (dithranol): calcd for C₅₈H₄₂N₄PZn [M - PF₆⁻ - P(O)Ph₃]⁺ = 889.24 found 888.91.

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Keywords: Phosphorus • Porphyrinoids • Molecular electrochemistry • Electrosynthesis • Radical reactions

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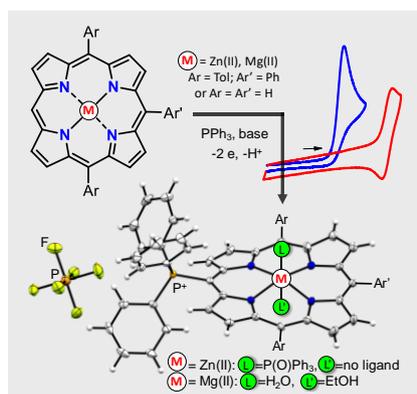
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A Zn(II) *meso*-triaryltriphenylphosphonium porphyrin was electrosynthesized in good yield via the electrochemical oxidation of Zn(II) *meso*-triarylporphyrin in the presence of triphenylphosphine. Its X-ray crystallographic structure was solved and was compared with the original Mg(II) *meso*-phosphonium porphine and the already published *meso*- and β -triphenylphosphonium porphyrins.



Phosphorus-Based Porphyrins

Abdou K. D. Dimé, H el ene Cattey,
Dominique Lucas and Charles H.
Devillers*

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Electrosynthesis and X-ray
crystallographic structure of Zn(II)
meso-triaryltriphenylphosphonium
porphyrin and structural comparison
with Mg(II) *meso*-
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