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► **To cite this version:**

A.R. Galvis E, F. Leardini, J.R. Ares, Fermin Cuevas, J.-F. Fernandez. Simulation and design of a three-stage metal hydride hydrogen compressor based on experimental thermodynamic data. International Journal of Hydrogen Energy, 2018, 43 (13), pp.6666-6676. 10.1016/j.ijhydene.2018.02.052 . hal-02054365

**HAL Id: hal-02054365**

**<https://hal.science/hal-02054365v1>**

Submitted on 28 Jul 2021

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# Simulation and design of a three-stage metal hydride hydrogen compressor based on experimental thermodynamic data

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## Abstract

A semi-empirical method was developed to design a three stage Metal Hydride Hydrogen Compressor (MHHC) through the determination of thermodynamic properties of several hydrides. As a first step, three AB<sub>2</sub>-type alloys that satisfy operation conditions were selected from published thermodynamic data entailing over 200 single plateau hydrides. These alloys were synthesized by arc melting and characterized by X-Ray Powder Diffraction (XRPD), Scanning Electron Microscopy (SEM) and Energy Dispersion X-ray spectroscopy (EDX). Absorption and desorption Pressure-composition-Isotherms (P-c-I) were determined between 23 and 80 °C to characterize their thermodynamic properties. Subsequently, an algorithm that uses these experimental data and a real equation of state for gaseous H<sub>2</sub> was implemented to simulate the volume, alloy mass, pressure and temperature of operation for each compressor stage, while optimizing the compression ratio and total number of compressed H<sub>2</sub> moles. Optimal desorption temperatures for the three stages were identified within the range of 110 to 132 °C. A system compression ratio (CR) of 92 was achieved. The number of H<sub>2</sub> moles compressed, the alloy mass and volume of each stage depend linearly on the volume of the external tank in which the hydrogen is delivered.

**Key Words:** AB<sub>2</sub> intermetallic hydrides, Hydride compressor, Thermodynamics, Semi-empirical modelling

## 1. Introduction

The urgency to replace fossil fuels by sustainable renewable ones has increased the research on novel technologies. Between several choices, hydrogen excels as an energy vector that can help to reduce greenhouse gases thanks to its high energy content and diversified sources of supply [1, 2]. Key features in the hydrogen cycle are its production, transport and storage [3, 4]. On this regard, hydrogen compressors provide an excellent approach to integrate the production and storage with the distribution of hydrogen as an energy carrier for different applications [5-7]. Specifically, Metal Hydride Hydrogen Compressors (MHHC) give great improvements in comparison to other compression technologies and therefore have focused much attention in recent years [5, 6].

Several approaches have been followed to enhance MHHC performances [5]. Simulations are particularly useful to predict and design hydrogen sorption and thermal management in Metal Hydride (MH) tanks, including MHHC. Hardy *et al.* [8] developed a hierarchical methodology for modeling hydrogen storage systems, in which finite element models (FEM) were used to simulate and optimize the hydrogen sorption properties of a cylindrical hydrogen storage tank containing sodium alanate with a shell and tube type heat exchanger. Although the model gave proper results, some realistic features such as hysteresis and the volume expansion of the complex hydride during sorption were not taken into account and the system geometry required further optimization. Førde *et al.* [9] developed a one-dimensional heat and mass transfer model with an ideal equation of state (EOS) for a MH-based hydrogen storage system. It was validated with experimental results from a La<sub>0.83</sub>Ce<sub>0.10</sub>Pr<sub>0.04</sub>Nd<sub>0.03</sub>Ni<sub>4.4</sub>Al<sub>0.60</sub> alloy at different temperatures and hydrogenation pressures. This model exhibited a higher sensitivity to kinetic and thermal parameters such as the activation energy, effective thermal conductivity and heat of the reaction

than with other operational variables described. Gkanas *et al.* [10] optimized the compression ratio by combining hydride materials AB<sub>5</sub> (LaNi<sub>5</sub> and MmNi<sub>4.6</sub>Al<sub>0.4</sub>) and AB<sub>2</sub> (Zr-V-Mn-Nb and Ti-Zr-V-Fe-Cr-Mn) for the first and second stage of the compressor, while maximizing the hydrogen rate using numerical simulations to predict the real behavior of the system.

Moreover, some prototypes and experimental approaches have been developed. Laurencelle *et al.* [11], designed a three-stage AB<sub>5</sub> MHHC prototype for a hydrogen production facility using a low-pressure alkaline electrolyzer. The materials selected for the system were LaNi<sub>4.8</sub>Sn<sub>0.2</sub>, LaNi<sub>5</sub> and MmNi<sub>4.7</sub>Al<sub>0.3</sub> for the first, second and third stage respectively, where it performs a thermal cycling between 20 and 80 °C with a flow rate that reaches about 20 L (NTP) of hydrogen per hour. Vanhanen *et al.* [12] studied the feasibility of combined MHHC and heat pump through the characterization of various alloys. The used alloys for both compression and heat pump were the Hydralloy C2 and Hydralloy C0 for the first and second stage, respectively. Although, the hydrogen pressure increased from 12-18 bar up to 85 -110 bar by utilizing a very narrow temperature interval (20-60°C), if a third alloy (TiCrMn<sub>0.55</sub>Fe<sub>0.33</sub>V<sub>0.15</sub>) was added, then the hydrogen pressure could reach almost 199 bar. As a result, a three stage MHHC gave greater results than just a two stage one. Wang *et al.* [13] developed a two-stage 700 bar hydride compressor using AB<sub>2</sub> hydrides by comparing the thermodynamic behavior of different compositions of the AB<sub>2</sub> alloys Ti-Cr-Mn and Ti-Zr-Cr-Fe-V, for the first and second stage of the compressor respectively. With exchanging cold oil (298 K) and hot oil (423 K), the built compression system could convert hydrogen pressure from around 40 bar to over 700 bar. Lototsky *et al.* [14] created a prototype MH compressor operating in a temperature range 20 to 150 °C and providing compression of hydrogen from 10 to 200 bar with an average productivity up to 1 m<sup>3</sup> /h. The compressor has a two stage layout where the first part uses AB<sub>5</sub>- and the second AB<sub>2</sub>-type hydride-forming intermetallic compounds. Bhuiya *et al.* [15] made an experiment-driven design procedure in order to optimize the parameters between stages in a dual-stage hydrogen compressor and subsequently achieved an enhanced compression ratio. The compressor system was composed by a LaNi<sub>5</sub> in the first stage and a Ca<sub>0.2</sub>Mm<sub>0.8</sub>Ni<sub>5</sub> in the second one, obtaining a higher couple compression ratio (53%) than the one acquired with just the hydride of the first stage. Muthukumar *et al.* [16] performed experimental tests on a MmNi<sub>4.6</sub>Al<sub>0.4</sub> based hydrogen compressor by varying operating parameters such as supply pressure and heating temperature under constant and variable delivery pressures with different storage volumes. Hydrogen storage pressure increases with a rise in the supply pressure and heating temperature. In the best conditions studied, with a variable H<sub>2</sub> delivery a maximum isentropic efficiency of 7.3% is obtained at a CR of 8.8 (43.8/5 bar) and 95°C hot fluid temperature, while with a constant delivery, the corresponding efficiency is 14.2%, the CR is 3 (30/10 bar) and the hot fluid is at 85°C. To summarize, previous studies are focused in the best combination of alloys at different stages of compression while searching the enhancement of the kinetics through geometry optimization.

Unfortunately, several MHHC works disregard many important thermodynamic features such as sloping plateau and hysteresis during the metal-hydrogen reaction or the use of real H<sub>2</sub> equation of state (EOS). Moreover, they only considered the combination of a few materials, though the thermodynamics of many MH are reported in the literature [5,17]. On this regard, Voskuilen *et al.* [18] developed an interesting approach by creating a fairly conscious data base of MH and the use of a simple numerical model to select a combination of materials for a two-hydride thermal system, specifically a heat pump. Thermophysical and kinetic properties of the alloys selected were considered to optimize a key system parameter like the Coefficient of Performance (COP).

The novelty of the present work is twofold. Firstly, the implementation of more than 200 alloys for a selection program, aids to ensure that the materials selected are an optimum combination for MHHC design. Secondly, real H<sub>2</sub> gas EOS as well as experimental features of the alloys (*i.e.* sloping plateau, hysteresis, and variation of the

reversible capacity with temperature) were considered to account for the real behavior of these materials and their combined influence in the final outcome of the system.

As a result, a new methodology is developed for improving the selection of materials for a MHHC system and, therefore, its performance. In the first part of the study, thermophysical data from literature and operational parameters of the compressor are considered to select the proper materials through a simple model based on ideal thermodynamics. Then, the selected materials were synthesized and characterized chemically, structurally and thermodynamically. Finally, an algorithm that considers the realistic thermodynamic models and real gas H<sub>2</sub> EOS is implemented to define and optimize some parameters of the MHHC system (*i.e.*, volume and alloy mass of each stage) to deliver the greatest quantity of hydrogen compressed at the highest compression ratio.

## 2. Materials and methodology

First of all, the thermophysical properties of more than 200 alloys from a data base [18] and literature studies [19-39] were considered to implement a selection program for a three-stage MHHC. The operational parameters of the system were considered as described in section 2.1. The selected materials were synthesized and characterized as detailed in section 2.2. Finally, the Pressure-composition Isotherms (P-c-I) determined experimentally are the main input of a Matlab<sup>TM</sup> program that uses realistic thermodynamic features of the hydrides and real EOS to design each stage of compression so that the final compression ratio and number of moles compressed can be optimized, as it is described in section 2.3.

### 2.1. First numerical approach

Several single-plateau metal hydride compounds were first selected for each stage of the compressor based on the thermodynamic properties of each alloy. Also, they were constrained according to the operational parameters of the system, such as: *i*) working temperature and inlet pressure of the compressor system (between room temperature ( $RT \approx 23^\circ\text{C}$ ) -  $150^\circ\text{C}$ , and around 1 bar, respectively), *ii*) an ideal compression ratio ( $CR = [\text{desorption pressure at high temperature} / \text{absorption pressure at low temperature}] > 8$  between coupling stages, *i.e.* first and second). Similarly, some criteria were included in the algorithm: *iii*) Van't Hoff derived plateau pressure (without slope), *iv*) hysteresis ( $\ln[P_{\text{plateau}}(\text{abs}) / P_{\text{plateau}}(\text{des})]$ ). Finally, the possibility of tuning of the thermodynamic properties of the hydrides by modifying their chemical composition was the final criteria to select the materials.

The first selection of the algorithm is made by comparing for each hydride reported the temperatures of operation and related plateau pressures with the established range at which the compressor works. Then, for the selected compounds, the published enthalpy and entropy of formation [18-39] are employed to generate Van't Hoff plots both for absorption and desorption. For materials for which data is only provided on absorption or desorption, an average value of hysteresis is applied [18] depending on the hydride type (AB, AB<sub>2</sub>, AB<sub>5</sub>...) to obtain the respective absorption/desorption Van't Hoff plots. Then, several alloys were selected for the first stage based on the constraint that the inlet pressure ( $P_{\text{in}}$ ) at RT should be slightly above 1 bar. Similarly, for the third stage, alloys with desorption pressures as high as possible at temperatures  $< 150^\circ\text{C}$  were identified. Subsequently, the second stage alloy has to accomplish that its desorption pressure at high temperature should be equal or greater than the absorption pressure of the third alloy at RT. Analogously, the desorption pressure of the first stage have to be equal or greater than the absorption pressure of the second alloy at RT. Several combinations of alloys fulfilling these conditions were found. The final selection was taken by choosing alloys of close chemical composition so that by tuning a small quantity of their constituting elements, the final thermodynamic properties can be properly modified. As a result, three compounds, a low pressure (LP), a medium pressure (MP) and a high pressure (HP) hydride were selected as the ideal final combination of alloys that achieves all of the requirements proposed above.

## 2.2 Experimental approach

The three samples selected for the compressor were synthesized by arc melting of raw elements (with a purity of Ti=99.99%, Zr=99.5%, Mn=99.9%, Cr=99.2%, V=99.7%) under an argon atmosphere of 0.7 bar. 6% weight in excess of Mn was used for all synthesis. The melted ingots were turned at least 5 times to ensure homogeneous composition. Their microstructure and elemental chemical composition were analyzed by Scanning Electron Microscopy (SEM, Hitachi S-3000N OK) and Energy Dispersive X-ray spectroscopy (EDX), respectively. X-ray Powder Diffraction (XRPD, PANALYTICAL X'pert PRO  $\theta$ -2 $\theta$ ) analysis with Cu-K $_{\alpha}$  radiation in a  $\theta$ -2 $\theta$  configuration and a further Rietveld refinement were performed to determine phase amounts, lattice constants and crystal density. P-c-I absorption and desorption curves were determined with a real gas H<sub>2</sub> EOS [40] and the data measured in a thermalized home-made Sieverts-type volumetric reactor at three different temperatures, 23, 50 and 80 °C. The samples were manually grinded by agate mortar and pestle, sieved under 45  $\mu$ m in Ar atmosphere and subsequently placed in the Sieverts reactor. The activation of all samples was performed by three absorption/desorption cycles at room temperature (23 °C), and at a pressure of 25 bar for the alloy M1=Ti<sub>0.85</sub>Zr<sub>0.15</sub>Mn<sub>1.33</sub>V<sub>0.3</sub> (0.566(1) gr) and of 50 bar for the alloys M2=Ti<sub>0.8</sub>Zr<sub>0.2</sub>Mn<sub>1.2</sub>Cr<sub>0.6</sub>V<sub>0.2</sub> (0.556(1) gr) and M3=Ti<sub>0.9</sub>Zr<sub>0.1</sub>Mn<sub>1.4</sub>Cr<sub>0.4</sub>V<sub>0.2</sub> (0.516(1) gr).

## 2.3 Second numerical approach

As final step, an algorithm in Matlab™ was implemented to simulate different operational variables of the compressor. The main input to the program comprises either experimental and/or simulated P-c-I curves at the selected temperatures of MHHC operation. Simulation of P-c-I curves uses the model developed by Fang et al. [41]. This model fits accurately the three main regions of the P-c-I curves with two equations that consider their variation with temperature [41, 42]. The first equation is a modification of the Sievert's law that simulates the solid-solution branch, *i.e.* the alpha ( $\alpha$ ) and beta ( $\beta$ ) phases. The second one models the plateau pressure, *i.e.* the biphasic  $\alpha$ + $\beta$  region. For each equation, several mathematical coefficients have to be found to simulate the P-c-I at low and high temperature. Also, considerations and constraints have been taken into account for the simulation: (1) The scheme of the whole system (Fig.1) involves three stages of compression, and an external vessel volume ( $V_{ext}$ ) connected to the third stage where the H<sub>2</sub> will be compressed to; (2) the  $P_{in}$  generates directly the minimum concentration in the beta phase ( $\beta_{min}$ ) in each hydride at RT; (3) the desorption point (pressure and hydrogen content in the P-c-I) at each stage will be the one corresponding to the maximum concentration in the alpha phase ( $\alpha_{max}$ ), and, its pressure ( $P_{des}$ ) would have to be equal or greater than the absorption pressure ( $P_{abs}$ ) of the following stage at RT; (4) The final desorption point, at the third stage, will determine the optimum high temperature value (HT) that enhances the CR and final number of H<sub>2</sub> moles compressed ( $n_{H_2-C}$ ); (5) A 65% filling density of the hydrogenated material density is used taking into account former studies on these type of materials [43]; (6) the mass ( $m$ ) and volume (SVol) of the reactor at each stage are also linked to  $V_{ext}$  where the H<sub>2</sub> will be compressed, hence, this volume will be another parameter to vary in order to observed its effect in the quantity of mass and reactor volumes at each stage; (7) the hydride is in thermal equilibrium with the surrounding gas; (8) a dead volume was considered at each stage due to connection tubes, valves and others (15.4(5) cm<sup>3</sup>/stage); (9) real H<sub>2</sub> gas EOS is also implemented [40].

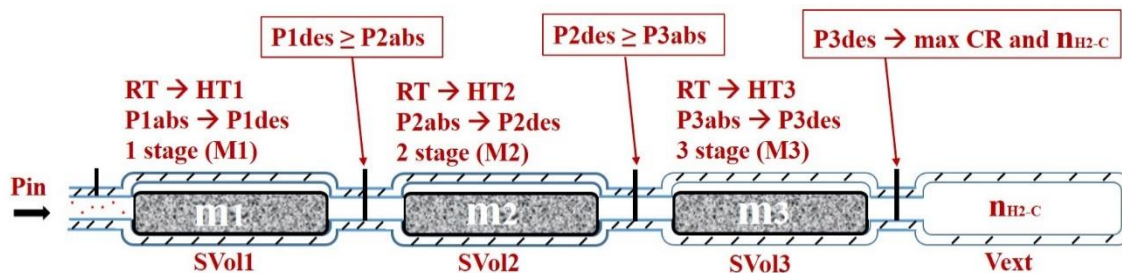


Fig.1. MHHC design scheme. Numbers indicate each consecutive stage up to  $V_{ext}$ .

### 3. Results and discussions

#### 3.1 Material selection from first numerical approach.

Fig.2 presents a Van't Hoff plot from the selected hydrides [17, 32] and the operational points of the three stages. The materials selected by the algorithm were the  $Ti_{0.9}Zr_{0.1}Mn_{1.3}V_{0.3}$  (LP), the  $Ti_{0.8}Zr_{0.2}Mn_{1.2}Cr_{0.6}V_{0.2}$  (MP) and the  $Ti_{0.9}Zr_{0.1}Mn_{1.4}Cr_{0.4}V_{0.2}$  (HP). The selected alloys are  $AB_2$  type, with compositions differing from those used by Vanhanen et al. in the design of a comparable MHHC system [12].

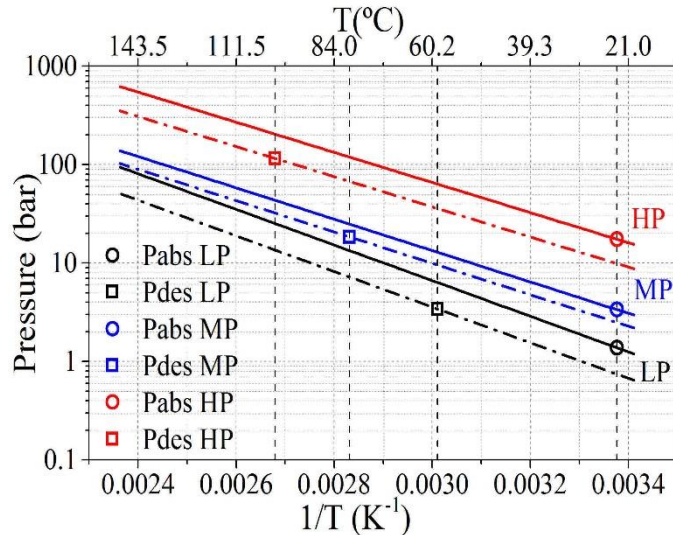


Fig.2. Van't Hoff plot of the three selected alloys. The absorption (solid lines) and desorption (dash lines) curves can be seen for the LP (black), the MP (blue) and the HP (red) compounds. The open circles in each material represent the inlet absorption pressure at each stage and at 23°C, while the open squares are the pressures of desorption at 59°C, 80°C and 100°C for the LP, MP and HP alloys respectively.

Fig.2 shows that for each alloy the desorption pressure at high temperature (open squares) matches with the absorption pressure at 23°C (open circles) of the alloy used in the subsequent stage. This ideally couples the first to the second and the second to the third stages. Also in Table 1, the main ideal values found for the three materials can be observed.

Table 1. Main results found by the selection algorithm for the three alloys

Stage / material	Absorption pressure @ 23°C (bar)	Desorption pressure (bar)	Desorption temperature (°C)	CR between stages
1 / LP	1.4	3.4	59	1-2: 13.1
2 / MP	3.3	18.4	80	2-3: 34.9
3 / HP	17.5	115.4	100	1-3: 82.4

The CR achieved between stages is always greater than 8, as established before. It is worth highlighting that the three alloys are  $TiMn_2$ -based, for which small compositional changes give rise to strong modifications of the Van't Hoff plots (Fig.2). The combination of these alloys fulfills all of the parameters established previously for this approach.

#### 3.2 Microstructure and composition of the alloys

Fig.3 shows the Rietveld refinement of the XRPD data for the three samples, showing the characteristic hexagonal C14-type Laves phase (SG  $P6_3/mmc$ ; N° 194) of  $TiMn_2$  based alloys [44-47]. The fact that the main phase of the three alloys crystallize in the same structural type concurs with previous reports showing that  $TiMn_2$   $AB_2$ -type alloys can accommodate a wide stoichiometric ratio ( $1.2 \leq B/A \leq 2$ ) without changing their structural order [48, 49].

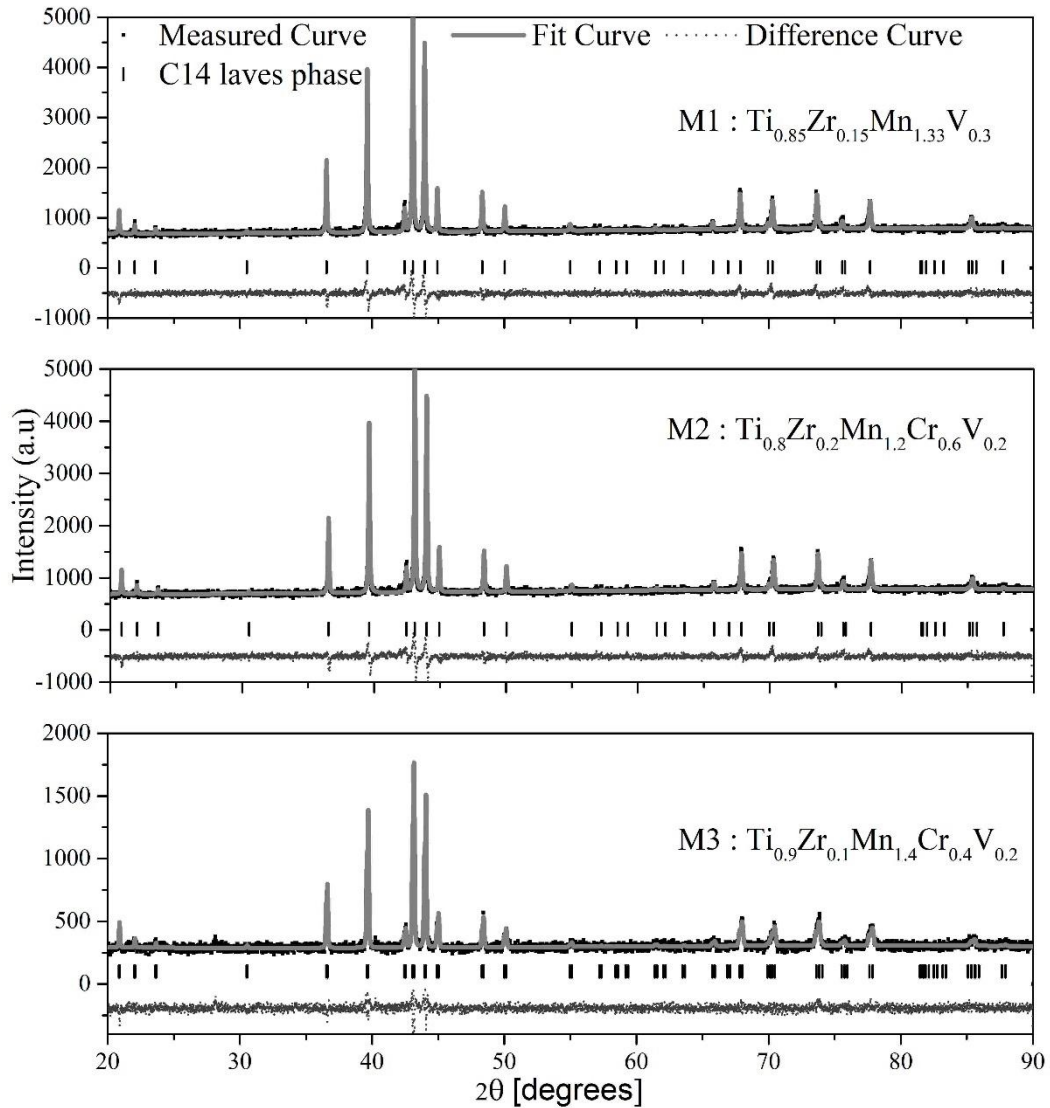


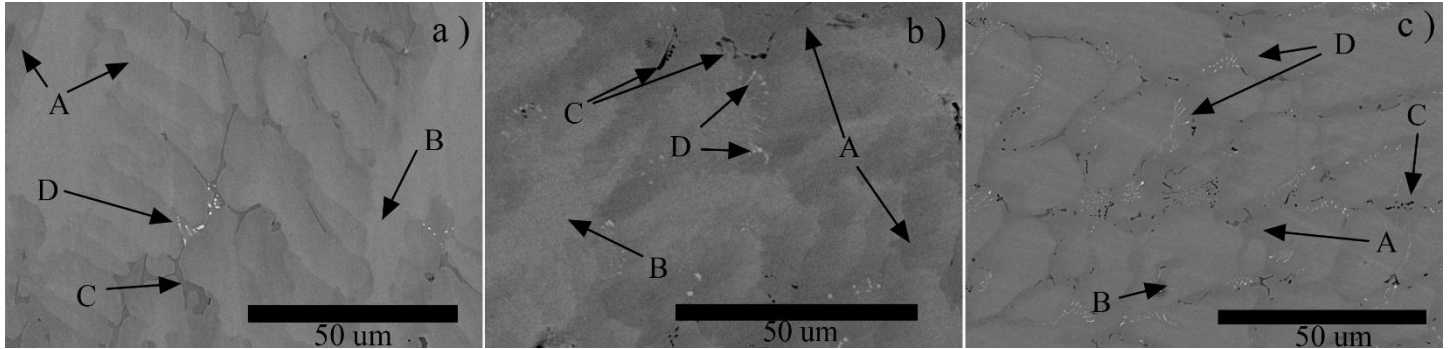
Fig.3. X-ray diffractograms with Rietveld refinement of M1, M2 and M3

The lattice parameters of the main AB<sub>2</sub> phase are gathered in Table 2. Clearly the cell volume gradually decreases as 168.6 cm<sup>3</sup>, 167.6 cm<sup>3</sup> and 165.5 cm<sup>3</sup> for M1, M2 and M3, respectively. This variation of the unit cell volume parameters agrees with empirical geometric laws on the correlation between thermodynamic properties and the cell volume of the alloys [50]. Furthermore, the alloy chemical compositions obtained by EDX analysis are gathered in Table 2 while in Fig.4 back-scattered electron images of the alloys can be seen.

Table 2. Average stoichiometry and lattice parameters of the C14-type Laves phase for M1, M2 and M3

	Light area		Dark area		Nominal (at./f.u.)	Lattice parameters	
	Average (at./f.u.)	Standard deviation (at./f.u.)	Average (at./f.u.)	Standard deviation (at./f.u.)			
<b>M1</b>	Ti	0.803	0.008	0.871	0.85	a [Å]	4.914(1)
	Zr	0.197	0.006	0.129	0.15	c [Å]	8.062(2)
	Mn	1.530	0.009	1.268	1.33	V [Å <sup>3</sup> ]	168.557(5)
	V	0.303	0.006	0.295	0.3		
<b>M2</b>	Ti	0.732	0.006	0.836	0.8	a [Å]	4,903(1)
	Zr	0.268	0.005	0.164	0.2	c [Å]	8,047(1)
	Mn	1.155	0.008	1.189	1.2	V [Å <sup>3</sup> ]	167.584(2)
	Cr	0.615	0.006	0.514	0.6		

	V	0.174	0.004	0.223	0.005	0.2		
	Ti	0.882	0.007	0.938	0.007	0.9	a [Å]	4,883(1)
	Zr	0.118	0.005	0.062	0.004	0.1	c [Å]	8,014(2)
<b>M3</b>	Mn	1.373	0.008	1.375	0.008	1.4	V [Å <sup>3</sup> ]	165.478(6)
	Cr	0.436	0.006	0.280	0.005	0.4		
	V	0.180	0.005	0.246	0.005	0.2		



**Fig.4.** SEM-BSE images showing the microstructure of the three alloys. a) M1, b) M2 and c) M3. Labels stand for: A. dark zone, B. light zone, C. AB precipitate, D. ZrO<sub>2</sub>.

As described in previous works with alloys of similar composition [43, 51], the sample is mainly composed by a mixture of two distinctive areas (light and dark presented in table 1 and shown in Fig.4) with different compositions. For M2 and M3 alloys, the two areas have close Mn content. However, for A-type elements, the light area exhibits higher Zr and lower Ti content than the nominal one. Similarly, for B-type elements, the light area has higher Cr and lower V than the nominal value. As for M1, the content of A-type elements Zr and Ti follow the same trend as in M2 and M3. However, the light area displays a high content of V leading to over stoichiometry in B-type elements as compared to the nominal value. This might be driven by the lack of Cr in this alloy. It has to be highlighted that despite of the different composition of light and dark zones, both phases crystallize in C14 laves phase with close lattice parameters as evidenced by the XRPD analysis. Furthermore, for the three alloys a ZrO<sub>2</sub> (D label in Fig.3) and an AB type precipitate (C label at Fig. 3) were observed in small amount (less than 1% of surface images). Their influence on the thermodynamic properties can be neglected as compared to the effect of chemical changes observed in the alloy matrix (light and dark zones) [48, 52]. The determined chemical composition of AB precipitate for M1 is Ti<sub>0.985(6)</sub>Zr<sub>0.104(4)</sub>Mn<sub>0.817(6)</sub>V<sub>0.191(4)</sub>, while for M2 and M3 it is Ti<sub>0.94(4)</sub>Zr<sub>0.06(4)</sub>Mn<sub>0.50(6)</sub>Cr<sub>0.15(6)</sub>V<sub>0.093(4)</sub>. These secondary phases were not observed in XRPD patterns (Fig.3).

### 3.3 Thermodynamic characterization and P-c-I modelling for numerical approach

The P-c-I of the three alloys at 23, 50 and 80 °C are displayed by symbols in Fig.5 a (M1), b (M2) and c (M3). Fits of the isotherm data to the Fang model are given both in absorption (solid line) and desorption (dash line). The experimental Van't Hoff plots determined from plateau pressures at a hydrogen content of ≈1wt% for absorption (solid line) and desorption (dash line) are displayed in Fig.5d.



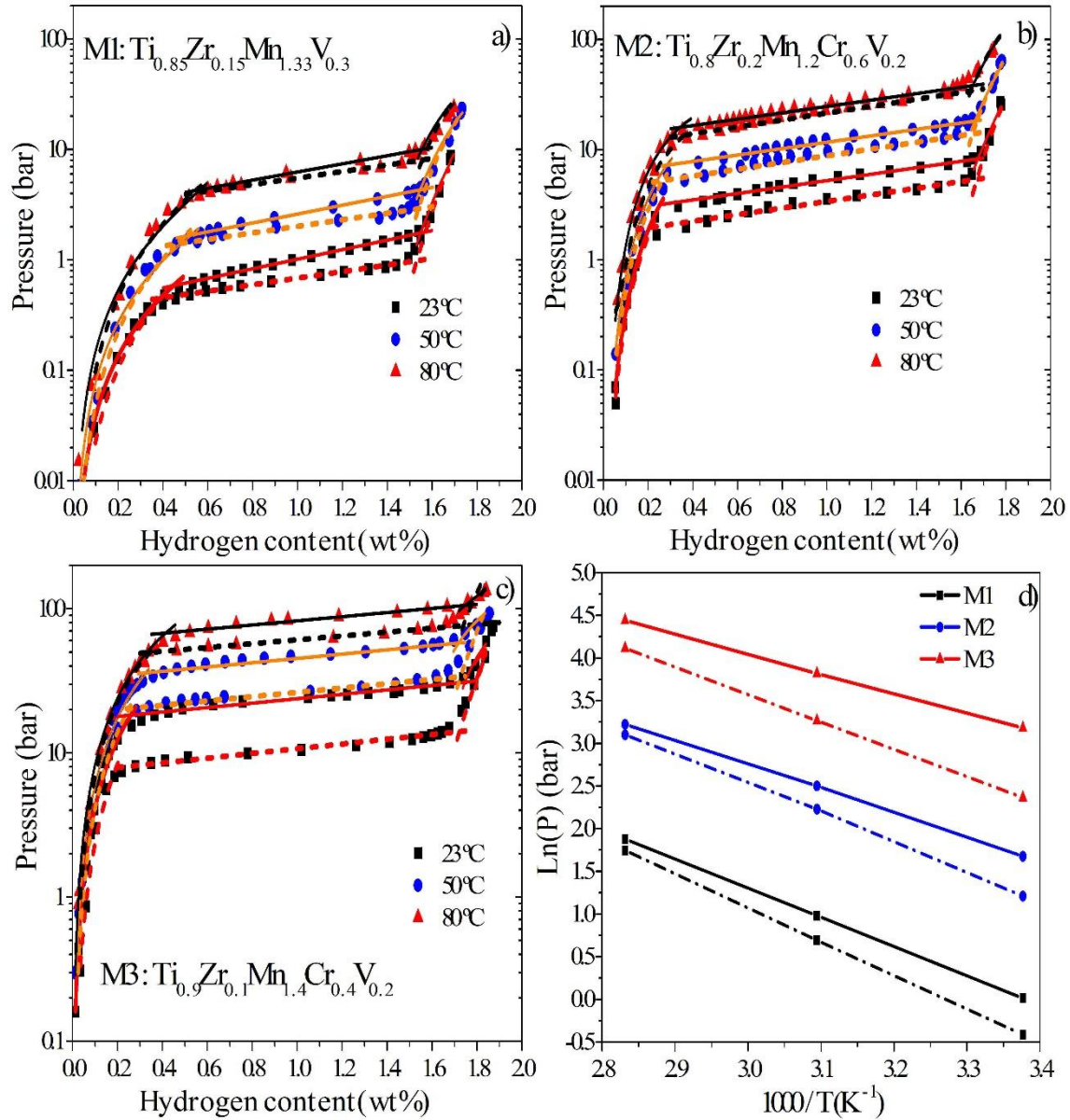


Fig.5. Thermodynamic properties of M1 (a), M2 (b) and M3 (c) alloys. Experimental P-c-I data in a-c are shown by symbols and the fits performed with the Fang model by lines. d) Derived Van't Hoff plots for absorption (solid line) and desorption (dash line).

Fig.5 shows that the Fang model provides a good fit of P-c-I data. All alloys exhibit sloping plateau which is attributed to the chemical heterogeneity of the alloys (Fig.4 and Table 2) [48] [52]. Furthermore, the hysteresis decreases with temperature for the three alloys (Fig.5 and table 3), which is in agreement with previous reports [53]. This reduction can be an advantage for the operation of the compressor, since the desorption and absorption pressure at high temperature would be closer, and this fact increases the CR and efficiency of the compressor [5]. Furthermore, the enthalpies and entropies of formation (F) and dissociation (D) can be found in Table 3.

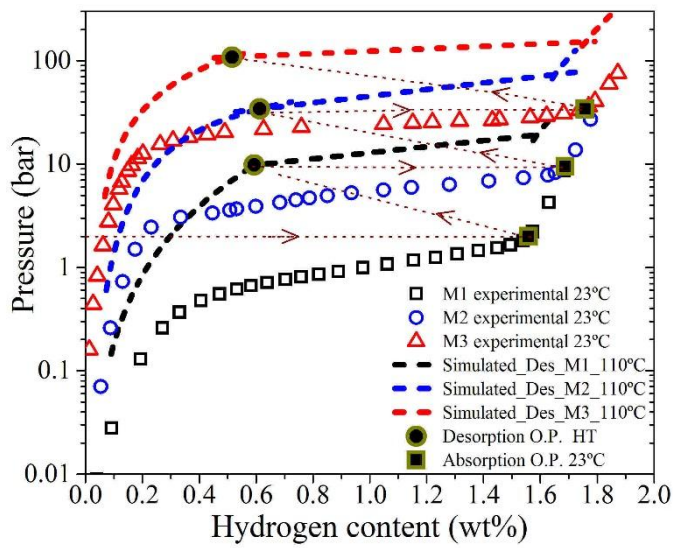
Table 3. Thermodynamic data of M1, M2 and M3

	Hysteresis @ $\approx 1\text{wt}\%$			Enthalpy <sub>F</sub> (kJ/mol)	Entropy <sub>F</sub> (J/mol*K)	Enthalpy <sub>D</sub> (kJ/mol)	Entropy <sub>D</sub> (J/mol*K)
	23°C	50°C	80°C				
M1	0.39	0.26	0.14	28±2	96±2	-33±1	-107.8±0.1
M2	0.44	0.28	0.14	24±1	94±1	-29±1	-108±3
M3	0.80	0.54	0.31	19.3±0.5	91.3±0.2	-27±1	-110±2

The results in entropy and enthalpy obtained are in general agreement with previous studies on this type of materials [38, 50, 54]. The decrease in enthalpy (lower stability) follows closely the decrease in cell volume previously mentioned.

### 3.4 Realistic thermodynamic simulation approach

Fig.6 shows the thermodynamic simulation of the coupling between the three stages of a realistic MHHC compressor, *i.e.* after considering hysteresis and sloping plateau effects. The figure shows the operation points (O.P.) for absorption (solid squares) at 23°C and desorption (solid circles) at a given HT that fulfills the operational conditions described in section 2.3. HT was initially fixed to 110 °C for the three stages. P-c-I absorption curves at 23°C are experimental data (open symbols). P-c-I desorption curves at HT are simulations using Fang model (Fig. 5) at the temperature where the operation conditions are satisfied. The dot lines with arrows indicate the thermodynamic path followed by the MHHC compressor.



**Fig.6. Thermodynamic simulation of the three stage MHHC. The solid squared points are the absorption O.P. for each stage. The solid circle points indicate the desorption O.P. The open symbols are the experimental data of each alloy at 23°C, while the dash lines are the simulations of the P-c-I zones at 110°C. The dot lines with arrows indicate the path followed by the three stage MHHC.**

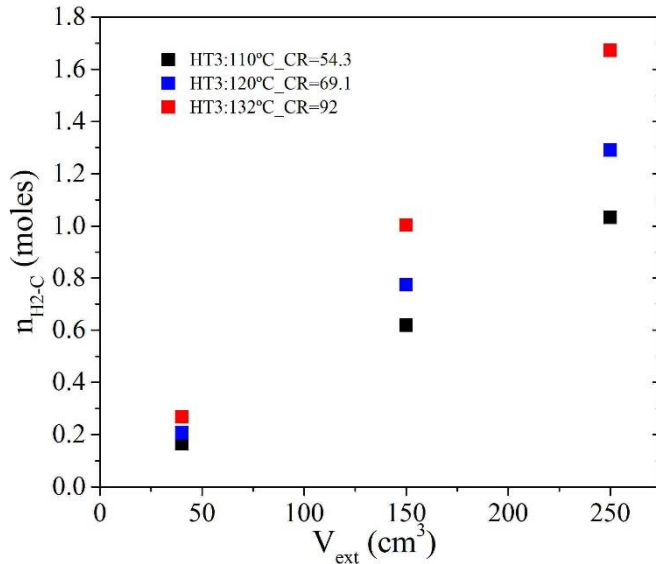
Several aspects of this simulation are worth discussing. The reversible capacity of each stage is comprised between  $\beta_{\min}$  at RT and  $\alpha_{\max}$  at HT. One should notice that  $\alpha_{\max}$  increases and  $\beta_{\min}$  decreases with temperature, so that the higher HT the lower the reversible capacity is [55]. For the first two stages, the HT should be above a certain value so that the desorption pressure of the initial stage is higher than the absorption pressure of the following stage. For instance, for HT1 = 110°C, the desorption pressure attains 9.6 bar, which is higher than the absorption pressure at 23 °C of the second stage: 9.5 bar. Likewise, the same HT value for the second stage releases hydrogen at 34.1 bar, which is higher than the 33.9 bar of absorption pressure in the third stage at 23°C. At these conditions, and, fixing also HT3 = 110 °C, the reversible capacity for M1, M2 and M3 alloys are 0.97, 1.07 and 1.2 wt%, respectively, with an output system pressure of 108.2 bar.

Furthermore, one should now consider that the desorption O.P. of the third stage influences the total CR and  $n_{H_2-C}$  of the system. To validate the influence of the HT3 temperature in these performances, three temperatures were analyzed: 110°C, 120°C and 132°C. Table 4 present the main results achieved with the different temperatures studied.

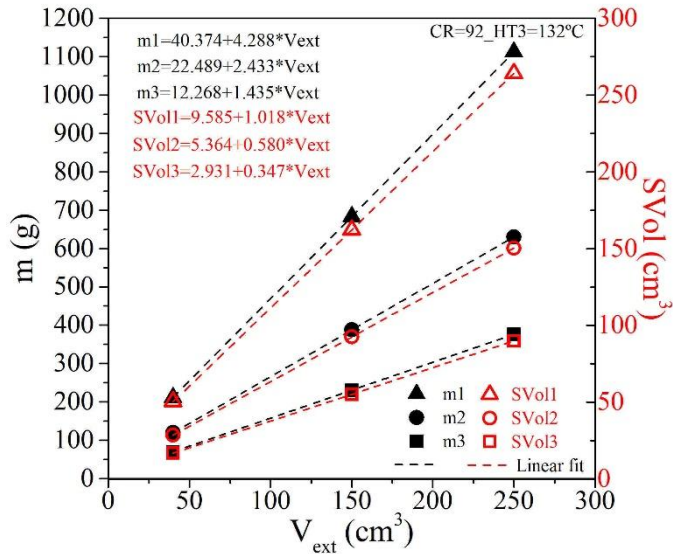
**Table 4. Main parameters achieved with the HT3 variation**

HT3 (°C)	P <sub>des</sub> (bar)	Reversible capacity (wt.%)	CR
110	108.2	1.2	54.3
120	137.7	1.15	69.1
132	183.2	0.98	92

As the HT3 temperature rises, the CR gradually increases from 54.3 to 92, considering that P<sub>des</sub> is an increasing function of temperature. The n<sub>H<sub>2</sub>-C</sub> depends mainly on two parameters of the design, the reversible capacity and the P<sub>des</sub> of the third stage, which are largely dependent on temperature and have a different effect on n<sub>H<sub>2</sub>-C</sub>. The first one decreases due to the shrinking of the two-phase  $\alpha+\beta$  miscibility on approaching the critical temperature (Fig.6) [55], leading to a modest decrease in the moles compressed. On the other hand, the P<sub>des</sub> increases exponentially with temperature, allowing for an increase in n<sub>H<sub>2</sub>-C</sub> as shown in Fig.7. Furthermore, besides the third stage temperature HT3, the V<sub>ext</sub> to which the hydrogen gas is compressed, is of critical importance as it determinates to a large extent the n<sub>H<sub>2</sub>-C</sub> of the system. This is evidenced in Fig.7, showing that this magnitude linearly depends on the V<sub>ext</sub>.

**Fig.7. n<sub>H<sub>2</sub>-C</sub> as a function of the V<sub>ext</sub> for three HT3 desorption temperatures.**

As expected, the best CR (92) and the highest amount of n<sub>H<sub>2</sub>-C</sub> (1.672) were achieved at the highest desorption temperature (132°C) and the largest V<sub>ext</sub> (250 cm<sup>3</sup>) studied. The enhancement of both performances benefit from the exponential grow of the plateau pressure with temperature even if the reversible capacity is reduced. Therefore, on increasing HT3 both the CR of the system and n<sub>H<sub>2</sub>-C</sub> increase. Indeed, the used methodology constraints the O.P. in the first and second stages to the variations in the last stage of compression. Then, the m and SVol are determined by HT3, leading to the best operational characteristics of the compressor.



**Fig.8. Dependence of alloy mass (left-axis) and vessel volume of each stage (right-axis) with  $V_{ext}$  for HT3 = 132°C. Inset equations stand for the best linear fits**

Fig.8 shows the dependence of the alloy mass and vessel volume at each stage on  $V_{ext}$  for HT3 = 132 °C. Both magnitudes can be fitted with linear equations (given in the figure) that depend on  $V_{ext}$ . The optimum alloy masses scale approximately in the ratio 3:2:1 for the first, second and third stages, respectively. The link between these parameters derive from the thermodynamic simulations shown in Fig.6. The given linear fits can be extrapolated to any external vessel volume at the given operation conditions. A summary of the optimized operating parameters of the MHHC system is displayed in Table 5.

**Table 5. Summary of the operating system parameters found during the simulation**

Stage	$P_{abs}$ (bar) // RT (°C)	$P_{des}$ (bar) // HT (°C)	Reversible capacity (wt.%)	m (g) // SVol (cm <sup>3</sup> )		
				At $V_{ext}$ (cm <sup>3</sup> )		
				40	150	250
1	1.99 // 23	9.6 // 110	0.97	212 // 50.3	684 // 162	1112 // 264
2	9.5 // 23	34.1 // 110	1.07	120 // 28.6	388 // 92	631 // 150
3	33.9 // 23	183 // 132	0.98	70.5 // 16.8	230 // 55.3	376 // 89.8

Alloy masses and stage volumes given in table 4 result from the optimization of thermodynamic parameters (e.g. compression ratio and desorption temperature) of the third stage. Nevertheless, a further kinetic analysis will be needed to better optimize the system operation and design to attain the highest efficiency and COP of the system [5].

It is not straightforward to compare our simulation results to previous works because they are very sensitive to system design (volumes and masses of alloys) and operation temperatures. However, for similar temperatures of operation, between RT and 130°C, CR values between 3 and 75 have been reported for other MHHC prototypes [5]. Our results are at the top of this range.

The results of this realistic thermodynamic approach differ in several aspects from those of the ideal first approach. First, the nominal composition of the alloys could not be exactly obtained after their synthesis due to the different volatility of the raw materials, as well as to the intrinsic tendency of TiMn<sub>2</sub>-type compounds to yield chemical inhomogeneity. This proves that simulations will be more precise when they are based in experimental data from real materials. Second, absorption pressures at each stage using real experimental data were observed to differ

from that predicted for the ideal ones in the first approach. For instance, for M1 alloy, the experimental absorption pressure at 23°C and  $\beta_{\min}$  is higher (1.99 bar) than that predicted by the first approach (1.4 bar). Analogous deviations are found for M2 and M3 alloys. This variation results from the sloping nature of the plateau pressure which is not considered in the ideal Van't Hoff plots [5, 51]. Indeed, the increase in the real absorption pressure at each stage requires a correlated increase in the desorption temperature of the previous one. As compared to an ideal thermodynamic model with flat plateau, the realistic one gives rise to a substantial increase (between 30 and 50°C, depending on the stage) in the desorption temperatures, thus affecting to the total energy consumption of the MHHC. Third, as concerns hysteresis effects, their influence on the operational parameters is lower than that induced by the sloping plateau. In fact, the hysteresis tends to be reduced when increasing the temperature [53]. Fourth, the realistic simulation takes into account the dependence of the  $\alpha_{\max}$  and  $\beta_{\min}$  phase limits with the temperature. As the  $\alpha+\beta$  region shrinks, the reversible capacity decreases, leading to the necessity of increasing the alloy mass at each compression stage. As a last remark, it is worth noting that even if the evidenced realistic features induce significant changes in the operation conditions of the system, the selection of the three initial alloys by the first approach remains valid. This results from the fact that thermodynamic correlations between the pairs of materials designated remains roughly the same.

### Conclusions

Three AB<sub>2</sub> compounds (M1: Ti<sub>0.85</sub>Zr<sub>0.15</sub>Mn<sub>1.33</sub>V<sub>0.3</sub>, M2: Ti<sub>0.8</sub>Zr<sub>0.2</sub>Mn<sub>1.2</sub>Cr<sub>0.6</sub>V<sub>0.2</sub>, M3: Ti<sub>0.9</sub>Zr<sub>0.1</sub>Mn<sub>1.47</sub>Cr<sub>0.4</sub>V<sub>0.2</sub>) have been selected from literature thermodynamic data in order to be used in a three stage MHHC. A realistic thermodynamic model that uses experimental thermodynamic data and real equation of state for H<sub>2</sub> gas were used to provide a more precise simulation of the MHHC operation parameters. This model allows to determine the optimal operational temperature, alloy mass and vessel volume at each stage. The realistic thermodynamic features of hydride formation and decomposition (sloping plateau, hysteresis and dependence of  $\alpha+\beta$  miscibility gap with temperatures) have an important effect on the optimal values of these operational parameters. In particular, for a given inlet pressure and final compression ratio, the desorption temperature and alloy mass at each compression stage are significantly higher than those obtained with an ideal model for hydride formation and decomposition (i.e., flat plateau, no hysteresis and constant miscibility gap). A further heat and mass transfer and hydrogen sorption semi-empirical kinetic analysis has to be performed in order to optimize the operability of the system in dynamic conditions.

### Acknowledgements

The authors thank the Spanish MINECO/FEDER for financial support under contract (MAT2015-65203-R) and Mr. F. Moreno and F. Couturas for technical assistance. One of the authors (A.R. Galvis E) would like to thank the Colombian Government and the Administrative Department of Science, Technology and Innovation -Colciencias-, for the financial support under contract 568.

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