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Oscillatory behavior in methane combustion: on the influence of the operating parameters

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Abstract

The influence of the main process parameters on the oscillatory behavior of methane oxidation was analyzed in conditions relevant for Low Temperature Combustion (LTC) processes. The investigation was performed by means of direct comparisons between experimental measurements realized in two Jet Stirred Flow Reactors used at atmospheric pressure. Coupling the operating conditions of the two systems, wide ranges of inlet

temperature (790 - 1225 K), equivalence ratio ($0.5 < \Phi < 1.5$), methane mole fraction (X_{CH_4} from 0.01 to 0.05), bath gases (i.e. He, N₂, CO₂, or H₂O) as well different overall mixture dilution levels, were exploited in relation to the identification of oscillatory regimes. Although the reference systems mainly differ in thermal conditions (i.e. heat exchange to the surroundings), temperature measurements suggested that the oscillatory phenomena occurred when systems working temperature accessed a well-identifiable temperature range.

Experimental results were simulated by means of a detailed kinetic scheme and commercial codes developed for complex chemistry processes. Simulations were also extended considering systems with different heat loss to the surroundings, thus passing from adiabatic to isothermal systems. Results highlighted the kinetic nature of the dynamic behavior. Because of predictions were consistent with experimental tests, further numerical analyses were realized to identify the kinetics responsible for the establishment of oscillatory phenomena. Temperature oscillations were predicted for a significant reactor working temperature range, where oxidation and recombination kinetic routes, involving carbon C₁₋₂ species as well as reactions of the H₂/O₂ sub-scheme, become competitive, thus boosting limit cycles behaviors. Oscillatory phenomena cease when the system working temperatures exceed characteristic threshold values with the promotion of faster oxidation routes that diminish the inhibiting effects of recombination reactions.

Keywords: Temperature oscillations, Jet Stirred Flow Reactor, chemical kinetics, diluted combustion.

Introduction

Despite much progress in the application of new combustion concepts to practical systems, such as Low-Temperature Combustion (LTC) processes, there remain unresolved issues on

the stabilization of the oxidation process¹. A problem related to LTC processes is the susceptibility to thermo-kinetic oscillations, originating from non-linear dependence of relevant variables.

Such instabilities may hamper the large-scale implementation of advanced combustion processes². Therefore, there is a need for a fundamental understanding of critical parameters that promote instabilities.

To properly investigate oscillatory phenomena, open systems are necessary. A very useful tool for the observation of combustion regimes as a function of the operating parameters is the Well Stirred Reactor (WSR). In this device, in contrast to chemical reactions occurring in a closed vessel that proceed to a unique point of thermodynamic equilibrium, it is possible to achieve a multiplicity of permanent steady states far from equilibrium as well as sustained oscillations that may also be preserved indefinitely.

A large literature exists on the analysis of dynamic phenomena taking place in open well-stirred reactors. For example, such reactors have been widely used to investigate oscillatory cool flames and complex ignition phenomena under non-adiabatic conditions for high molecular weight organic compounds at low temperature ($T < 850$ K), as a result of thermal feedback coupled to complex kinetics³⁻⁶. Scientific interpretations of these phenomena date back several decades because modeling of the low temperature oxidation of hydrocarbons is relevant to the onset of “knock” in spark-ignition (SI) engines, or to the ignition in diesel fuels⁷⁻⁹. Recently, the same oscillatory behavior was also found for dimethyl ether (DME) between 530 and 650 K¹⁰.

Oscillatory behaviors have also been recognized for low molecular weight fuels, such as hydrogen¹¹⁻¹⁷, or hydrogen and carbon monoxide mixtures¹⁸⁻²⁰. Notwithstanding exothermic contributions, the oscillations of such mixtures originate from a kinetic interaction driven by the competitive elementary reactions $H + O_2 = OH + O$ and $H + O_2 + M = HO_2 + M$. For the

latter, in fact, the rate of recombination waxes and wanes according to the third body efficiencies of species in the vessel.

Other experimental^{12,21,22} and numerical^{23,24} studies have focused on methane ignition and oxidation under MILD combustion operative conditions in flow reactors at atmospheric pressure. These studies highlighted the onset of non-isothermal oscillatory regimes similar to cool-flame oscillations but at higher temperatures (1000 - 1300 K). Sabia et al.²⁵ extended the studies by de Joannon et al.^{2,21} adding H₂ to CH₄/O₂/N₂ mixtures. These studies showed that H₂ can reduce the parameters space where oscillations were detected, with respect to methane mixtures. de Joannon et al.^{2,21} and Sabia et al.^{22,25} identified different kinetic routes promoted by high inlet temperatures and highly diluted mixtures, with respect to conventional flames. In particular, they pointed out that the competition between termolecular recombination reactions and branching reactions is crucial for the appearance of these regimes. Wada et al.²⁴ added that a key reaction pathway for the onset of oscillations is the route that converts CH₃ to CH₂O. At low temperature CH₃ is oxidized to CH₃O and CH₃O₂ that, in turns, react to form CH₂O, while at intermediates one CH₃ produces directly CH₂O by reacting with HO₂ radicals.

All these previous studies^{2,21-25} claim that oscillations have a thermo-kinetic nature, thus that they originate from a balance between the heat released by chemical reactions and the heat loss to the surroundings.

On the other hand, Sabia et al.²⁶ have recently found that dynamic behaviors can be also numerically obtained under adiabatic conditions, so they highlighted the kinetic aspect of such phenomena.

Another process, which received considerable attention, is the thermal oxidative coupling of methane (OCM). It is generally accepted that a homogeneous coupling of methyl radical, producing ethane, becomes important when the temperature is between 1000 – 1300 K²⁷,

which is the same temperature range in which non-isothermal oscillatory regimes have been highlighted for methane oxidation. Therefore, a detailed study of methane oscillations can be also utilized to improve OCM process.

Recently, Bagheri et al.²⁸ explored thermochemical oscillations of methane from a kinetic modeling perspective, also focusing on the effect of different dilution gases (N₂, H₂O, CO₂) in promoting or inhibiting such instabilities. The authors found that the variations of heat capacity associated with the different bath gases directly impact the oscillations characteristics. On the other hand, the authors noted that in WSR isothermal studies available in the literature, oscillations were not reported; however, an attempt at modeling these cases revealed the presence of cyclic oscillatory behaviors.

Given this background, the objective of the current work is to extend the fundamentals of methane combustion stability under well-stirred flowing conditions, with particular emphasis on the physical and chemical factors that control the oscillatory phenomenon. In order to seek the means to interpret events in a wider range of external parameters, experimental tests were carried out in two Jet Stirred Flow Reactors (JSFRs), characterized by different heat release. Oscillations onset, waveform, amplitude and frequency were characterized as a function of (1) inlet temperature, (2) equivalence ratio, (3) residence time, (4) methane mole fraction/mixture dilution level, and (5) the chemical nature of the bath gas (He, N₂, CO₂ or H₂O), at atmospheric pressure.

The interest in the effects of the bath gas is twofold: first, for LTC systems, the combustion occurs in the presence of non-inert species (CO₂ or H₂O) due to the strong recirculation of exhausted gases; second, such species are commonly present in Low Calorific Value (LCV) fuels. These species may have a thermal effect (i.e. higher heat capacities) or a kinetic effect, participating directly in bimolecular elementary reactions, or as third body in termolecular reactions with high collisional Chaperon efficiencies. The comparison with mixtures diluted

with inert species can reveal the role of CO₂ or H₂O in oxidation processes at low-intermediate temperatures.

Finally, the experimental measurements were supported by extensive numerical analyses of a chosen detailed thermo-kinetic model in order to identify the kinetics responsible for the onset and termination of oscillations with respect to operating parameters.

Experimental and numerical tools

Experimental apparatuses

In the present work, two spherical Jet Stirred Flow Reactors (JSFR) were used: the former is available at the Laboratoire Réactions et Génie des Procédés, Université de Lorraine, CNRS (Nancy)²⁹, while the latter at the Institute for Research on Combustion, CNR (Napoli)². Hereafter they will be referred as the “Nancy” and the “Napoli” reactors. Photos of the two JSFRs are reported as Supporting Information. They were designed under the same construction rules^{30,31}, but conceived to work under different operating conditions (i.e. residence time, mixture composition). The detailed description of the reactors is provided elsewhere^{2,29}.

The opportunity to combine the experimental tests from the two facilities implies that a large field of parameters is accessible for an exhaustive investigation of methane oxidation process. Both the JSFRs consist of a fused silica sphere where reactions take place. Quartz is commonly used to prevent catalytic reactions. The Napoli reactor consists of a sphere of 113 cm³ while the Nancy one of 85 cm³. Fresh reactants enter the reactor through an injection cross (four nozzles, 1 mm diameter for the Napoli reactor; 0.3 mm for the Nancy one) located at the center of the sphere. The well-mixing condition was verified in the past by means of pulse tracer experiments³². The Napoli reactor behaves as an ideal WSR for residence times

(τ) lower than 0.6 s, while the Nancy one lower than 10 s. Thus, the experiments were respectively carried out with residence times equal to 0.5 s and 2 s.

In both the reactors the total flow is composed by methane, oxygen and diluent. In the Napoli device, a primary flow, composed by oxygen and diluent is pre-heated through a quartz tube located within two semi-cylindrical electrically heated ceramic fiber insulated ovens supplied by Watlow Srl, regulated by PID controllers. The primary flow subsequently mixes with the secondary one, composed by fuel and diluent, in a premixing chamber before entering inside the reactor. The reactor is located within two other thermally insulated electrical fiber ovens to minimize heat loss to the surroundings. A recirculation air system provides a homogeneous temperature distribution in the oven. The homogeneity of the reactor external wall temperature is monitored by means of two movable N type thermocouples.

In the Nancy experiments, the reactor is preceded by an annular preheating zone in which the temperature of the gases is increased up to the reaction temperature before entering inside the reactor. Then the pre-mixed mixture enters the reactor. Both the spherical reactor and the annular preheating zone are heated by Thermocoax heating resistances rolled up around their walls. The external wall reactor temperature is monitored by means of a K type thermocouple.

In both experimental facilities, the preheating system allows reaching inlet temperatures (T_{in}) up to 1250 K. T_{in} is continuously monitored by a K type shielded thermocouple located in the center of the nozzles, with a precision of ± 2 K. Electronic control of the heating system gives uniform temperatures within ± 5 K over the external vessel surface and constant in time to within ± 0.5 K. This allows maintaining the reactor external wall at the same temperature as the inlet temperature.

The residence time of the fuel mixture in the preheating system is more than 100 times lower than the mean residence time in the reactor. Thus, oxidation of the fuel inside the injector is

prevented. When a temperature increase is detected by the T_{in} thermocouple, indicating the onset of ignition/oxidation reactions during the mixing process, the experimental conditions are not considered. For diluted methane mixtures this condition does not occur in the investigated temperature range.

Gases are stored in cylinders and supplied to the system by means of calibrated digital thermal mass flow controllers supplied by BronkHorst High-Tech. They are preceded by filters to retain particulate impurities. The precision of the mass flow rates corresponds to $\pm 0.5\%$ of set operating conditions.

Gases are provided with high purities: 99.999% for nitrogen, oxygen, and helium, 99.5% for carbon dioxide and methane.

The experiments carried out with steam diluted mixtures rely on a water feeding system consisting of a stainless steel cylinder, pressurized with dry nitrogen, a calibrated liquid mass flow meter and a Controlled Evaporator Mixer (CEM) provided by BronkHorst High-Tech. The CEM consists of a control valve, a mixing device and a heat exchanger whose temperature is controlled by an integrated system controller. A carrier gas (nitrogen) is used to promote the evaporation process and to transport the vapor. The temperature of gas lines downstream the evaporator is controlled by electrical coils to prevent water condensation.

The reactor pressure is continuously monitored by means of a transducer connected to the outlet from the vessels. It is kept constant in both the apparatuses by means of a needle valve located on the gas exhaust line.

Experimental procedures and measurement methodology

In each test, the total mixture mass flow rate was set for the specified residence time τ , adjusted for the inlet temperature T_{in} , and the reactor pressure P . The mixture residence time

is defined as τ = reactor volume/volume flow rate. For the Napoli tests, P was maintained at 1.1 atm and τ fixed at 0.5 s, whereas, for the Nancy ones, P = 1.05 atm and τ = 2 s.

For each test, vessels were flushed with inert gas before use. After fuel injection, a period of 5 minutes was allowed to reach thermal equilibrium. Initial test conditions refer to the parameters set prior to reactions taking place. Therefore, the inlet temperature has not been varied continuously, but each experimental point corresponds to a different test.

In order to accurately monitor the temperature changes during the oxidation process, homemade type R unshielded thermocouples gas welded from a very fine Pt-Pt 13% Rh wire (0.2 mm bead size) interfaced to a PC with a high-speed multi-channel module, supplied by National Instruments, were used. A photo of the thermocouple located inside the reactor is reported as Supporting Information. The global thermocouple response time is about 30 ms thus allowing for a faithful record of the reactor temperature during oscillatory behaviors with a precision of ± 2 K. A thin-wall alumina double-core tube covers thermocouples wires to prevent catalytic reactions. Herein, it has to be pointed out that thermocouples made using the same procedure were used in both the JSFRs.

The outlet concentration of the reactants was analyzed by gas chromatography. In the Napoli reactor, the exit gases were cooled down by means of a heat exchanger installed at the reactor outlet to quench the oxidation reactions. Then, gases passed through a silica gel trap to eliminate the moisture content before gas analyses. Therefore, results are reported on a dry basis. The outlet was connected to a gas chromatograph (Agilent 3000) equipped with a column Molsieve 5A and a pre-column Plot U and a Thermal Conductivity Detector (TCD).

In Nancy the quantification of the products from the reactor was realized with the aids of two gas chromatographs (GCs). The first GC equipped with a thermal conductivity detector was used to quantify O₂. The second GC equipped with flame-ionization detector preceded by a

methanizer and a PlotQ capillary column was used to quantify CH₄, CO, CO₂, C₂H₄, C₂H₆ and C₂H₂.

Maximum relative errors in mole fractions are estimated to be $\pm 10\%$, when the concentrations approach the detection threshold (about 10 ppm). The instrument calibration was performed daily using standard gas mixtures stored in cylinders.

Experiments were performed at least twice under all test conditions to ensure measurement reproducibility. The test repeatability was excellent under all conditions with differences among the various tests well within the reported experimental uncertainties.

Computational approach

Simulations were carried out using the OpenSMOKE++³³ PSR code, calculating the transient solutions of an open WSR³³. Geometrical characteristics were the same as those of the experimental reactors, i.e. a volume of 113 cm³ and an internal surface of 113 cm² for the Napoli reactor and a volume of 85 cm³ and an internal surface of 93 cm² for the Nancy one. Experimental mixture inlet conditions reported in Tab. 1 are declared in the input files. Since the two experimental systems are non-adiabatic, Newtonian-type heat losses were enabled. The reactor overall heat transfer coefficient (U) in the JSFRs was calculated by means of an empirical correlation available in the literature³⁴ and fixed at 1.3×10^{-3} cal/cm² s K. Reactor walls were assumed to be infinitely thin with negligible temperature gradients. The wall temperature was set equal to the inlet one.

A recently updated detailed gas-phase chemical kinetic mechanism “POLIMI” (C1C3LT_1412) was used. It was developed by Ranzi et al.³⁵ for the pyrolysis and combustion of a large variety of fuels. The mechanism consists of 2642 elementary reactions with 107 species. Thermodynamic properties are included in the database. A detailed analysis and validation of the reliability of this scheme is discussed elsewhere³⁵. This kinetic

mechanism was chosen as reference scheme because it was one of the most reliable in predicting oscillatory behaviors under MILD operating conditions, for both methane^{2,24} and propane^{36,37} mixtures.

Results

Experiments were carried out from fuel lean to fuel rich mixtures for a wide range of inlet temperatures, at atmospheric pressure. The bath gas was changed along with the fuel mole fraction and the overall molar fraction of diluent species (hereafter indicated as dilution “d”). The experimental conditions are summarized in Tab. 1. Data for cases S1, S2 and S3 are from Bagheri et al.²⁸. Furthermore, the data collected for each dataset are provided in the Supporting Information.

Case	Φ	τ (s)	X_{CH_4}	X_{O_2}	X_{He}	X_{N_2}	X_{CO_2}	X_{H_2O}	Number of tests
L1	0.5	0.5	0.020	0.080	0	0.900	0	0	22
L2	0.5	0.5	0.020	0.080	0	0	0.900	0	16
L3	0.5	0.5	0.020	0.080	0	0.495	0	0.405	21
L4	0.75	2	0.010	0.027	0.963	0	0	0	10
L5	0.75	2	0.010	0.027	0	0	0.963	0	10
L6	0.75	2	0.025	0.067	0.908	0	0	0	10
L7	0.75	2	0.025	0.067	0	0	0.908	0	10
L8	0.75	2	0.050	0.134	0.816	0	0	0	10
L9	0.75	2	0.050	0.134	0	0	0.816	0	10
S1	1	0.5	0.033	0.067	0	0.900	0	0	24
S2	1	0.5	0.033	0.067	0	0	0.900	0	15
S3	1	0.5	0.033	0.067	0	0.495	0	0.405	22

S4	1	2	0.010	0.020	0.970	0	0	0	10
S5	1	2	0.010	0.020	0	0	0.970	0	10
S6	1	2	0.025	0.050	0.925	0	0	0	10
S7	1	2	0.025	0.050	0	0	0.925	0	10
S8	1	2	0.050	0.100	0.850	0	0	0	10
S9	1	2	0.050	0.100	0	0	0.850	0	10
R1	1.25	2	0.010	0.016	0.974	0	0	0	10
R2	1.25	2	0.010	0.016	0	0	0.974	0	10
R3	1.25	2	0.025	0.040	0.935	0	0	0	10
R4	1.25	2	0.025	0.040	0	0	0.935	0	10
R5	1.25	2	0.050	0.080	0.870	0	0	0	10
R6	1.25	2	0.050	0.080	0	0	0.870	0	10
R7	1.5	0.5	0.043	0.057	0	0.900	0	0	25
R8	1.5	0.5	0.043	0.057	0	0	0.900	0	17
R9	1.5	0.5	0.043	0.057	0	0.495	0	0.405	23

Tab. 1 Experimental conditions studied in the JSFRs for CH₄ mixtures. 790 K < T_{in} < 1225 K, P = 1.1 atm when $\tau = 0.5$ s (Napoli) and P = 1.05 atm when $\tau = 2$ s (Nancy).

In order to classify the combustion regimes as a function of T_{in}, as exemplifying case, the reactor temperature (T), and the CH₄ and O₂ concentration, for a stoichiometric CH₄/O₂ mixture diluted in N₂ at 90% (Napoli tests case S1) are reported in Fig. 1.

In this figure, vertical lines identify the transition conditions from one state to another in response to a stepwise change of T_{in}. At low temperatures (T < 1000 K), non-reactive conditions (I), characterized by zero temperature increase and zero fuel conversion, are identified. For increasing T_{in}, a secondary stationary state (II) with very small extents of

oxidation is recognized for $1000 \text{ K} < T_{in} < 1100 \text{ K}$ with a maximum temperature increase of about 5 K. An oscillatory region (III) establishes in the T_{in} range 1100 - 1170 K with periodic temperature variations. For this oxidation regime, two T values are reported for the same T_{in} in Fig. 1 corresponding to the maximum and minimum values of temperature T detected within the oscillation cycle. In this case, the measured concentration values are averaged over time because of the time elapsed for the gas sampling before the chemical analyses, thus only one concentration value is reported in Fig. 1. To better identify these conditions, the closed symbols are substituted with open ones.

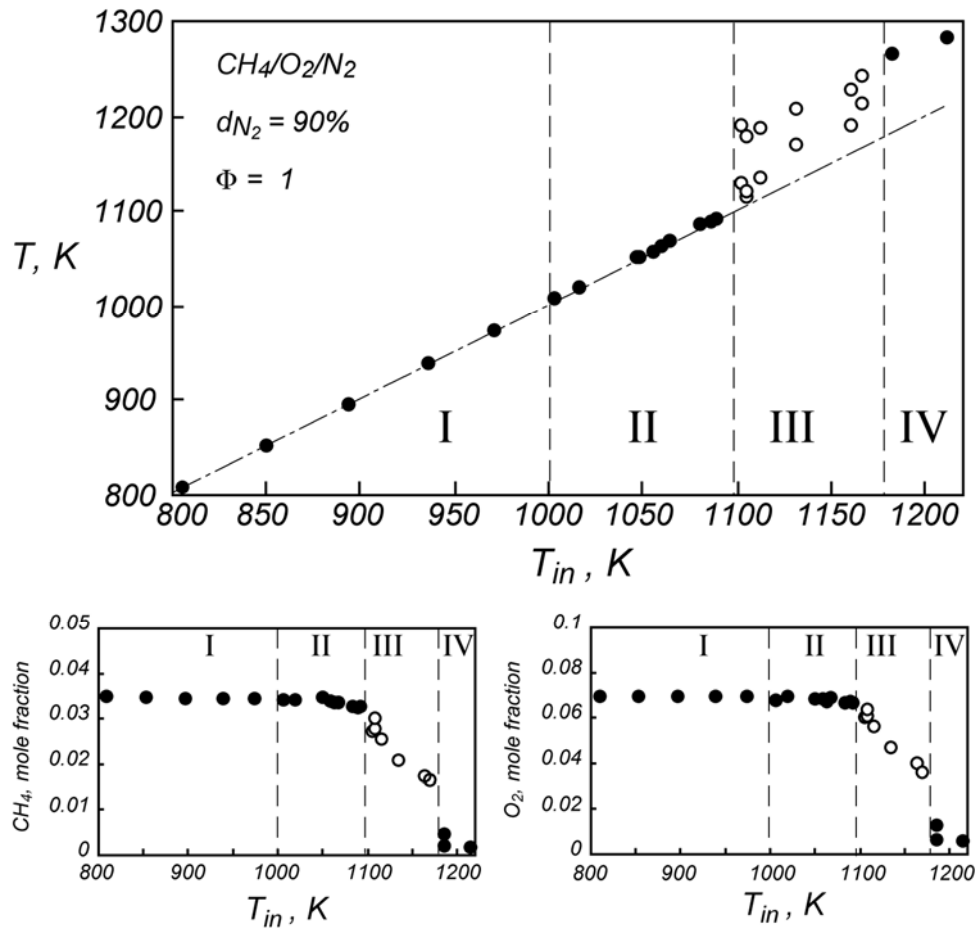


Fig. 1 Experimental reactor temperature (T) and reactants concentrations versus T_{in} . $\Phi = 1$, $\tau = 0.5 \text{ s}$, $P = 1.1 \text{ atm}$, N_2 bath gas, $X_{CH_4} = 0.033$. Closed symbols represent steady states, open symbols oscillatory regimes. See text for description of I, II, III, IV zones.

For $T_{in} > 1170$ K, the mixtures ignite and the system reaches a stationary state (IV) with temperature increase between 70 and 80 K and almost complete conversion. This steady state regime can be attained either monotonically or via damped oscillations.

The experimental condition reported in Fig. 1 is used as a reference case because it includes all the observed critical phenomena. However, the stationary state IV has been observed only for relatively short residence times (namely in the Napoli reactor). This can be explained considering the relative importance of sensible heat removal by system losses at the vessel walls and the heat convective transport associated with the outflow by calculating the ratio between two characteristic times, the mixture mean residence time τ and the Newtonian cooling time (t_c). The latter depends on the vessel characteristic dimensions V/S (V and S are respectively the vessel volume and surface), and on the ratio $\rho c_p/U$ (ρ and c_p represent the mixture density and specific heat capacity), which does not vary significantly with the temperature. When an exothermic reaction is in progress, if $\tau > t_c$, heat losses to the walls predominate over heat transport in the outflow. Therefore, the longer the residence times, the lower the system temperature. For both reactors, t_c can be assumed to be equal to 0.07 ± 0.01 s. Considering the residence times used in the two JSFRs, τ/t_c is about 7 for the Napoli reactor and 30 for the Nancy one.

Thus, both systems are far from the adiabatic condition. In particular, in the Nancy configuration, temperature increases due to oxidation reactions are not measurable at stationary states because the system approaches the isothermal condition. However, interestingly, temperature oscillations were still observed.

The analysis of the complex waveform periodic events could elucidate the role of parameters in triggering oscillations, and hence it could provide some useful hints for the study of oscillatory phenomena in practical systems.

Based on the above classifications, $T_{in} - \Phi$ maps were constructed (Figs. 2 and 3) considering all experimental conditions investigated in both reactors (see Tab. 1). In particular, each point on a map corresponds to an experimental condition. Given that the present work is mainly concerned with oscillatory states, the demarcation of the different reaction modes is only provided on the maps for oscillatory and non-oscillatory regimes. Therefore, regimes (I), (II), and (IV) are all identified as “no oscillation” regimes with full circles, whereas dynamic conditions are represented by open circles. Maps are also provided with numerical results, where the temperature ranges of oscillating conditions are indicated by grey rectangles.

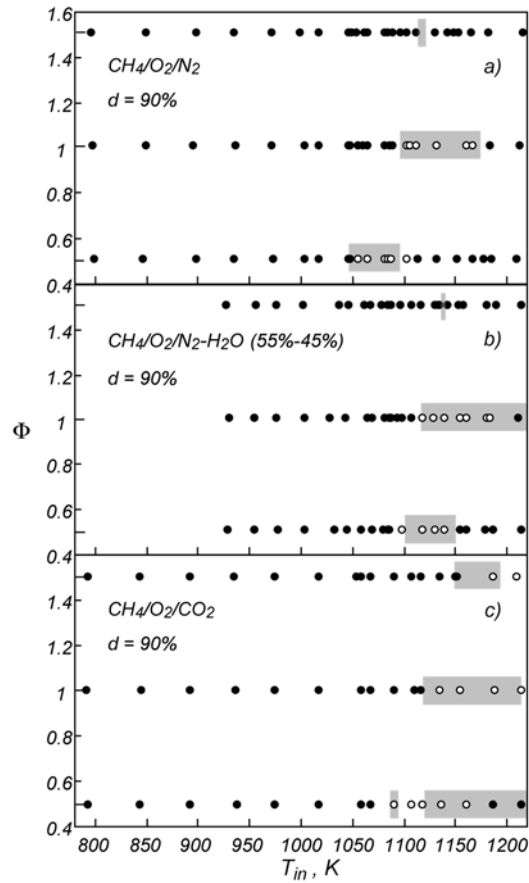


Fig. 2 Experimental (symbols) and simulated (rectangles) $T_{in} - \Phi$ maps. $\tau = 0.5$ s, $P = 1.1$ atm, $X_{bath\ gas} = 0.9$. (a) N_2 bath gas, (b) N_2-H_2O bath gas, (c) CO_2 bath gas. Closed circles: “no oscillation”; open circles: “oscillation”; grey rectangles: numerical oscillatory conditions simulated with the “POLIMI” model.

Fig. 2a summarizes the Napoli experimental tests carried out for mixtures of CH₄/O₂ diluted in N₂ at 90% for several equivalence ratios ($\Phi = 0.5, 1, 1.5$). For $\Phi = 0.5$ (L1), temperature oscillations occur for T_{in} between 1055 K and 1100 K. For $\Phi = 1$ (S1), the dynamic behavior starts for $T_{in} > 1100$ K and ends at $T_{in} = 1180$ K. Finally, for the fuel rich condition (R7), the dynamic behavior was not observed in the whole investigated T_{in} range.

An analogous identification of regimes can also be made when H₂O or CO₂ are used as diluent species.

Fig. 2b shows experimental results for a mixture of CH₄/O₂ diluted in N₂ and H₂O with a relative concentration of 55% and 45% and a global mixture dilution level of 90% (L3, S3 and R9). It can be seen that steam leads to a shift of the oscillations regions towards higher T_{in} with respect to the N₂-diluted system. This shift is even more evident when mixtures are diluted in CO₂ (Fig. 2c, L2, S2 and R8). In this case, temperature oscillations also appeared for the fuel rich condition.

Fig. 3 shows the $T_{in} - \Phi$ maps for the systems diluted in He (a, c, e) and CO₂ (b, d, f). Tests were repeated for three specific methane mole fractions, i.e. $X_{CH_4} = 0.01$ (a, b), 0.025 (c, d), 0.05 (e, f), at $\tau = 2$ s and for three Φ values (0.75, 1, 1.25) with T_{in} stepwise of 25 K (Nancy tests). Fig. 3a identifies oscillatory regimes in the $T_{in} - \Phi$ plane for $X_{CH_4} = 0.01$ and He bath gas. The “oscillation” region covers a quite narrow temperature range of 50 K ($1075 \text{ K} < T_{in} < 1125 \text{ K}$) for the fuel lean condition (L4). Oscillations for stoichiometric mixtures (S4) involve a wider range of T_{in} ($T_{in} > 1100 \text{ K}$). For the fuel rich conditions (R1), temperature oscillatory behaviors were not detected in the whole temperature range considered.

As the methane mole fraction increases (Figs. 3c, 3e), the oscillatory region on the map widens under stoichiometric conditions (S6 and S8), and shrinks for lean ones. In particular, at $\Phi = 0.75$, this region covers a temperature range from 1100 K to 1125 K when $X_{CH_4} = 0.05$

(L8) and vanishes when $X_{CH_4} = 0.025$ (L6). Oscillations were never observed under rich conditions, for mixtures diluted in He.

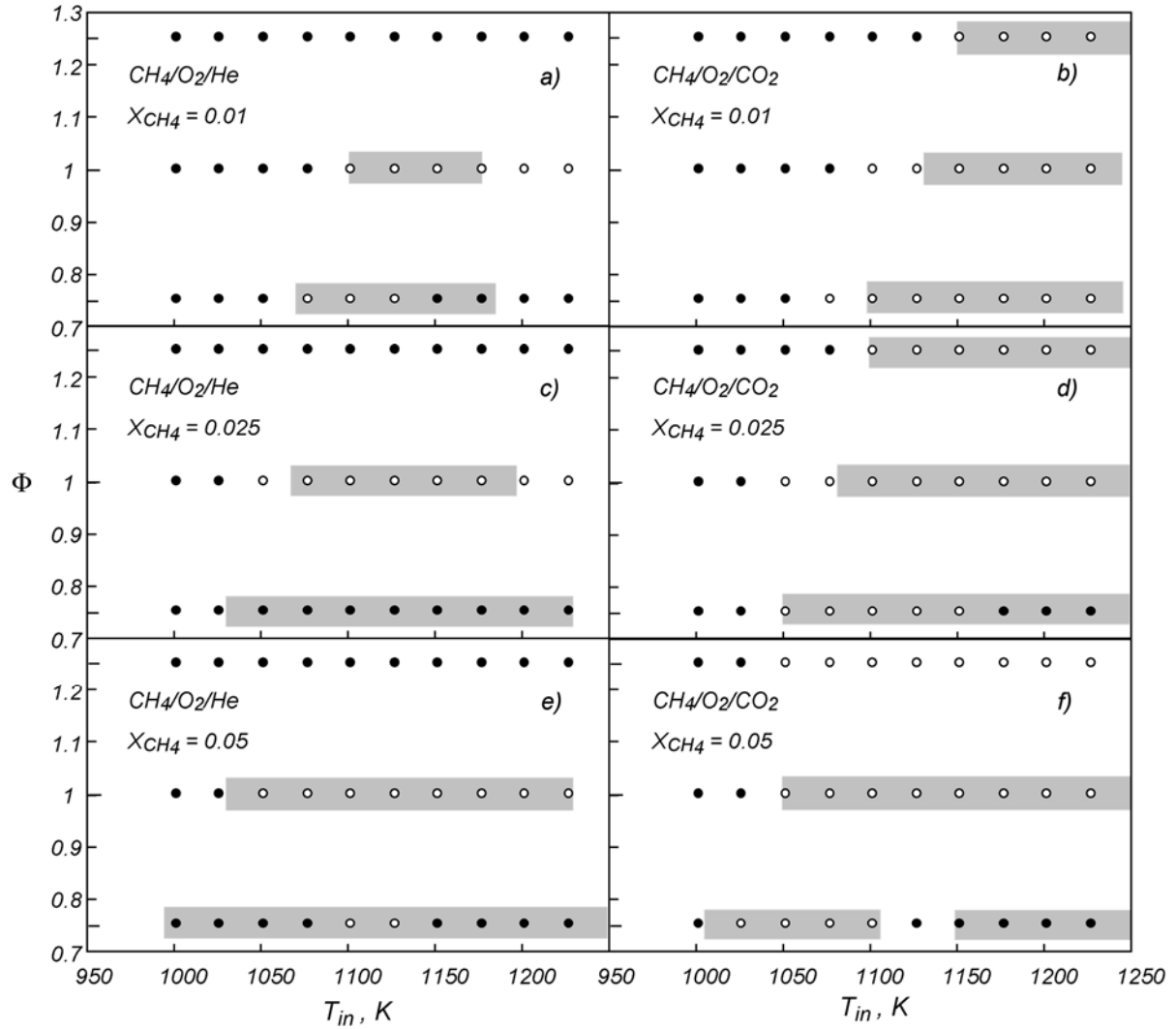


Fig. 3 Experimental (symbols) and simulated (rectangles) T_{in} - Φ maps. $\tau = 2$ s, $P = 1.05$ atm. (a) He bath gas and $X_{CH_4} = 0.01$, (b) CO_2 bath gas and $X_{CH_4} = 0.01$, (c) He bath gas and $X_{CH_4} = 0.025$, (d) CO_2 bath gas and $X_{CH_4} = 0.025$, (e) He bath gas and $X_{CH_4} = 0.05$, (f) CO_2 bath gas and $X_{CH_4} = 0.05$. Closed circles: “no oscillation”; open circles: “oscillation”; grey rectangles: numerical oscillatory conditions simulated with the “POLIMI” model.

A similar identification of regimes can also be made when CO_2 is used as bath gas (Figs. 3b, 3d, 3f). Oscillatory regions are clearly wider than in the He diluted cases. More specifically,

in agreement with the data obtained at a lower residence time in the Napoli reactor, the oscillatory region extends up to fuel rich conditions, indicating that the nature of temperature oscillations with CO₂ bath gas under fuel rich conditions is more complex, in agreement with the observations of Sabia et al.²⁶

The performance of the chosen detailed kinetic scheme to predict experimental data was tested by means of systematic runs. As it can be seen from the location of the grey rectangles on the maps (Figs. 2 and 3), the adopted kinetic model is able to reproduce the experimental $T_{in} - \Phi$ with relatively minor differences on the location of boundaries between regions of different behavior. Accuracy checks of the numerical procedures ensured that the predicted oscillations are not due to numerical integration problems. For validation, the simulations were also performed using ChemKin PRO³⁸, yielding the same results.

In particular, the detailed kinetic scheme is able to reproduce quite accurately the influence of the equivalence ratio and of the bath gas on the oscillation region boundaries. Moreover, the kinetic model correctly predicts the absence of oscillations for fuel rich conditions when using N₂, He, or H₂O as bath gases. It has to be pointed out that in the case of Napoli experiments, the kinetic model predicts the occurrence of oscillations under fuel rich conditions when N₂ and N₂-H₂O are used as bath gases, but the T_{in} range is very narrow (less than 5 K). However, disagreements exist in model predictions for several conditions. In particular, in Figs. 2c and 3f it is possible to observe that the POLIMI model predicts a discontinuity in the oscillation T_{in} range under fuel lean conditions. Such a behavior was not experimentally observed. Furthermore, in Figs. 3d and 3e, for fuel lean conditions, the numerically predicted T_{in} range for temperature oscillations is wider compared to the experimental one, and, in Fig. 3c, the model predicts oscillations under fuel lean conditions that were not experimentally detected. In addition, in Fig. 3f oscillations are not predicted under rich conditions whereas they were detected experimentally.

In order to show some characteristic features of temperature oscillations as a function of T_{in} , four typical temporal temperature profiles are displayed in Figs. 4 and 5.

Fig. 4 refers to a stoichiometric CH_4/O_2 diluted in N_2 at 90% (S1) preheated to $T_{in} = 1100$ K (Fig. 4a) and $T_{in} = 1165$ K (Fig. 4b). The periodic profiles are reported after ignition, thus the initial value is arbitrary. In Fig. 4a, temperature increases slowly at first then abruptly up to a maximum value, followed by an instantaneous temperature decrease. The oscillation period is about 6.7 s with an amplitude of about 60 K. The minimum temperature is about 25 K higher than T_{in} . As T_{in} is increased (Fig. 4b), the temporal temperature profile has a different shape with a lower period (less than 1 s).

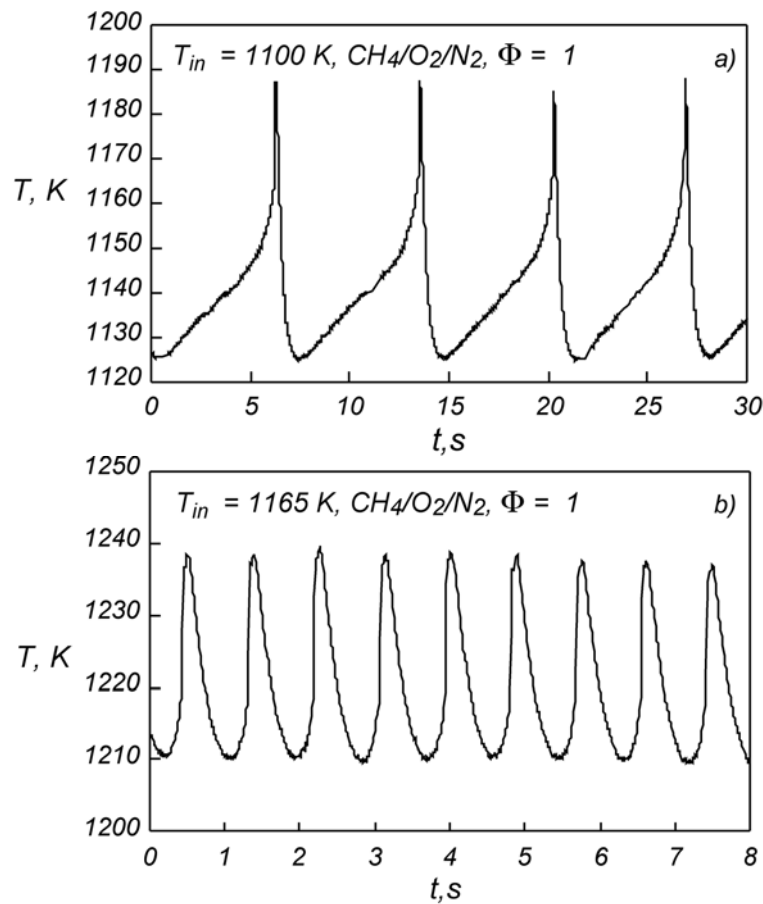


Fig. 4 Representative experimental temporal temperature profiles obtained at $\Phi = 1$ in N_2 bath gas with $X_{CH_4} = 0.033$ and $\tau = 0.5$ s at (a) $T_{in} = 1100$ K and (b) $T_{in} = 1165$ K. The initial time is arbitrary.

These two profiles delineate a characteristic feature of the oscillatory behaviors: for increasing T_{in} , temporal profiles become more symmetric, while the oscillation amplitude decreases and the frequency increases. Furthermore, the difference between the minimum temperature and the inlet one also increases. These observations rule out the possibility of classifying the identified periodic phenomenon as multiple ignitions. The periodicity should be a feature of the kinetic mechanism itself and is not an artifact of flow operation.

Fig. 5 refers to CH_4/O_2 mixtures diluted in CO_2 , obtained at $\tau = 2$ s. In particular, Fig. 5a is for a stoichiometric mixture preheated to $T_{in} = 1100$ K (S7), while Fig. 5b refers to a fuel rich mixture ($\Phi = 1.25$) preheated to $T_{in} = 1125$ K (R4).

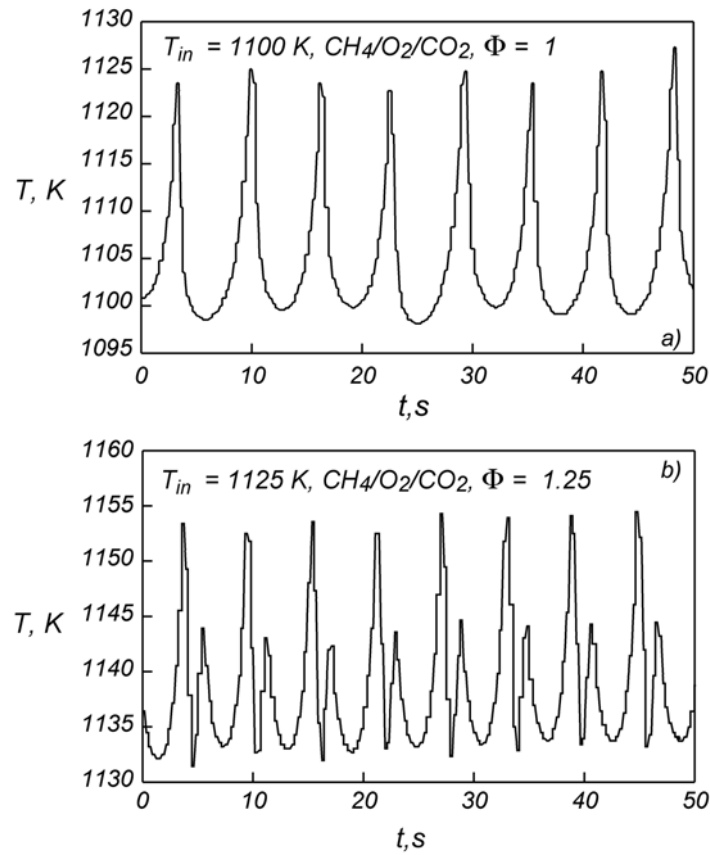


Fig. 5 Representative experimental temporal temperature profiles obtained in CO_2 bath gas with $X_{\text{CH}_4} = 0.025$ and $\tau = 2$ s at (a) $\Phi = 1$ and $T_{in} = 1100$ K and (b) $\Phi = 1.25$ and $T_{in} = 1125$ K. The initial time is arbitrary.

In Fig. 5a, the temporal profile is symmetric and the oscillation period is about 6 s with an amplitude of about 25 K. The minimum temperature is equal to the inlet one. As the equivalence ratio is increased (Fig. 5b), the temporal temperature profile has the same period, but a different shape. In particular, two maxima can be identified within each oscillation cycle with higher amplitude for the first peak. This confirms that the nature of temperature oscillations with CO₂ under fuel rich conditions is more complex, in agreement with Sabia et al.²⁶. Other examples of temporal temperature profiles are provided as Supporting Information.

The computed temporal temperature profiles are not reported in Figs. 4 and 5 because, although the overall features of the computed temperature history resemble the experimental measurements very closely, the predicted temperature increase is substantially larger than the experimental (usually within 200 K). This occurs because the predicted temperature oscillations are characterized by spike variations in less than 10 ms and, even though the response time of the thermocouples is very short ($\cong 30$ ms), the temperature increase cannot be properly resolved in time. A detailed analysis of numerical temperature profiles can be found in Bagheri et al.²⁸.

Fig. 6 shows the experimental (symbols) and numerical (lines) oscillations frequencies as a function of T_{in} , in two representative sets of experiments. In particular, Fig. 6a refers to stoichiometric mixtures diluted in N₂, N₂-H₂O and CO₂, with $X_{CH_4} = 0.033$ and $\tau = 0.5$ s (S1, S2 and S3), whereas Fig. 6b is related to stoichiometric mixtures diluted in He and CO₂, with $X_{CH_4} = 0.05$ and $\tau = 2$ s (S8 and S9).

For all bath gases, frequencies are relatively low at first. As T_{in} increases, oscillations frequencies increase. In both figures (Figs. 6a and 6b), CO₂ leads to a lower frequency with respect to other bath gases. The data related to N₂-H₂O dilution (Fig. 6a) lie between those obtained using N₂ bath gas and those related to CO₂ bath gas for $T_{in} < 1140$ K. For $T_{in} > 1140$

K this trend is reversed, and the N_2 - H_2O -diluted system displays higher frequencies than the other two mixtures. Overall Fig. 6 suggests that the oscillation frequencies are significantly modified by the presence of different bath gases and they are strongly dependent on T_{in} .

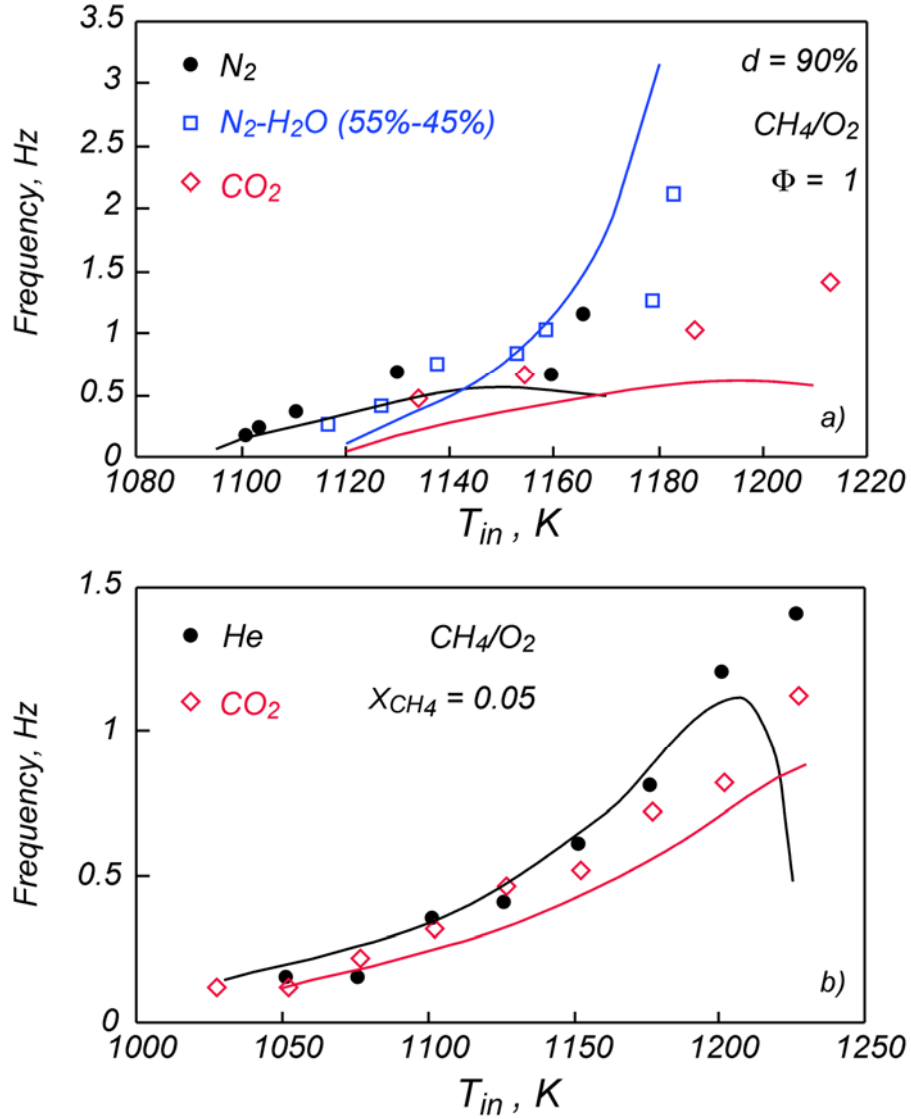


Fig. 6 Experimental (symbols) and numerical (lines) oscillation frequency as a function of T_{in} at $\Phi = 1$ in different bath gases with (a) $X_{CH_4} = 0.033$ and $\tau = 0.5$ s and (b) with $X_{CH_4} = 0.05$ and $\tau = 2$ s.

The POLIMI kinetic model satisfactorily predicts the experimental trends. However, depending on the parameters, a significantly longer period is predicted by the kinetic model at the highest T_{in} compared to experimental results, with a maximum deviation of 1.33 s.

Discussion

The factors that govern the stability of stationary states and the establishment of temperature oscillations need to be understood using numerical analysis. The reference model gives a satisfactory, qualitative and quantitative account for the $\Phi - T_{in}$ regimes along with oscillatory frequencies, as T_{in} and mixture composition are changed. Therefore, numerical analyses of complex dynamic behavior can be used to understand the role of elementary reactions in determining the observed behaviors.

It is well known that a central feature of oscillatory regimes is a negative *feedback mechanism* (i.e. the reaction products influence the rates of earlier reaction steps in the mechanism) that restricts the runaway acceleration of an intermediate concentration and differentiates the oscillatory reaction from an explosion⁴, where the feedback is always positive. Negative feedback can be caused by a number of different phenomena, some purely physical, others of chemical nature⁴. As reported in the introduction, oscillatory regimes may be significantly influenced by heat losses. Therefore, as a first step it is appropriate to analyze the influence of the heat loss on the observed phenomenon (Fig. 7).

Figure 7a shows the effect of modifications to the heat transfer coefficient (U) on the steady-state reaction temperatures as a function of T_{in} for a stoichiometric $CH_4/O_2/N_2$ mixture. The different conditions analyzed allow to sensibly vary the system working temperature, thus representing an indirect way to parametrically evaluate the operating conditions that boost the onset of oscillatory behavior. In addition, the system was modeled under the asymptotic

assumptions of adiabatic (Fig. 7b) and isothermal (Fig. 7c) conditions in order to highlight crucial aspects of the problem.

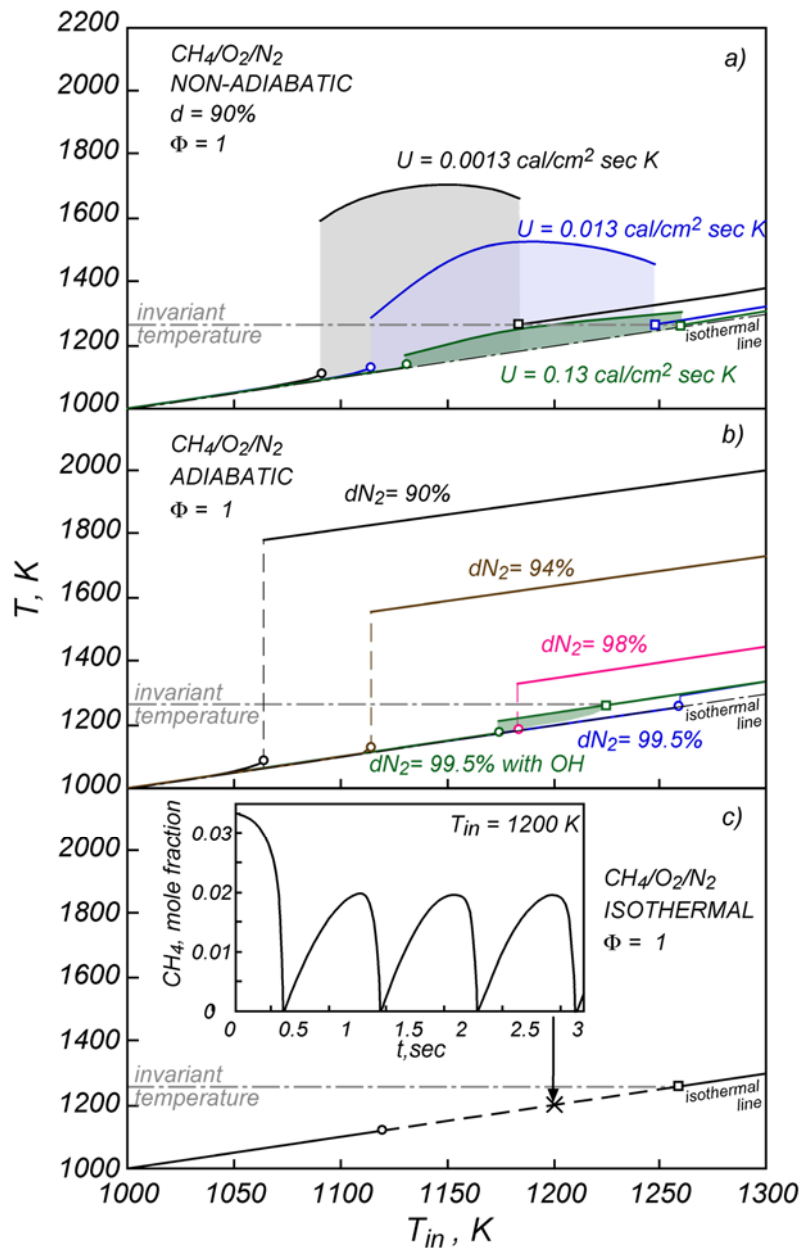


Fig. 7 Computed reactor temperature (T) versus inlet temperature (T_{in}) under a) non-adiabatic conditions for different values of the heat transfer coefficient; b) adiabatic conditions for different dilution; c) isothermal conditions. $\Phi = 1$, $\tau = 0.5 \text{ s}$, $P = 1.1 \text{ atm}$, N_2 bath gas. Open circles: ignition condition. Open squares: end of oscillations. Shadowed area: maximum and minimum values of temperature oscillations.

Figure 7a shows the reactor temperature for a stoichiometric mixture diluted in N₂ at 90% by varying the heat transfer coefficient (U) from 0.0013 to 0.13 cal/cm² sec K in steps of one order of magnitude. As expected, the temperature values are inversely proportional to the heat loss coefficients. For $U = 0.0013$ cal/cm² sec K, the system temperature remains equal to T_{in} up to $T_{in} = 1080$ K, then for $1080 < T_{in} < 1090$ K, the temperature increases very slightly, and finally, at about $T_{in} = 1090$ K, the mixture ignites and temperature oscillations are predicted. The ignition condition is identified on the figure by an open circle. In Fig. 7a, the shadowed area under the T_{max} curves identifies the maximum and minimum values of temperature oscillations. For these inlet conditions, the minimum reactor temperature during oscillations is equal to the inlet one. Oscillatory behavior is identifiable up to $T_{in} = 1180$ K (square symbol). For $T_{in} > 1180$ K, the system shows a steady stationary state behavior with a temperature increase of about 80 K, in agreement with the experimental results reported in Fig. 1.

For $U = 0.013$ cal/cm² sec K the behavior is very similar but the mixture ignition occurs at higher T_{in} (about 1120 K), then temperature oscillations establish. They have lower amplitudes with respect to the previous case. Temperature oscillations die out at $T_{in} = 1250$ K, when a steady state condition is reached. Similar observations apply for the case $U = 0.13$ cal/cm² sec K. Ignition occurs for $T_{in} = 1130$ K with the onset of temperature oscillations that cease at $T_{in} = 1260$ K. Interestingly, these oscillations are observed when the temperature decreases below a fixed value, i.e. $T = 1260$ K, reported in Fig. 7 with a horizontal dot-dashed line, thus suggesting that kinetics is controlling the suppression of oscillations.

Fig. 7b shows the behavior of the ideal PSR under adiabatic conditions for a stoichiometric CH₄/O₂ mixture diluted in N₂ at different levels (from 90% to 99.5%).

For N₂ = 90%, the ignition occurs at about $T_{in} = 1060$ K, then the system reaches a steady stationary state. No oscillatory behavior is predicted. As the dilution level of the mixture

increases ($dN_2 = 94\%$ and 98%) the auto-ignition event occurs at higher T_{in} with lower temperature increase. None of the system exhibits a periodic dynamic behavior because of the larger temperature increase. In fact, when ignition occurs, the working temperature T is above 1260 K , thus no oscillations are observed, in agreement with Figure 7a. In the case of $dN_2 = 99.5\%$, the ignition occurs at 1260 K with a temperature increase of 35 K . In this case the ignition occurs when T_{in} exceeds the identified threshold temperature, thus no oscillations are observed. However, for the latter system a further simulation has been considered. In particular, the auto-ignition event was forced by doping the inlet mixture with a low concentration of OH (200 ppm). For this case, the auto-ignition occurs for $T_{in} = 1200\text{ K}$ and oscillations are observed. The periodic dynamic behavior can no longer be sustained when the system temperature exceeds 1260 K (horizontal dot-dashed line). It is worth noting that the addition of OH does not significantly change (1 K) the system temperature in the high temperature stationary solution.

On the other hand, transition from low-temperature steady states to periodic oscillatory states is attainable when the mixture residence time is higher than the mixture ignition delay time. Since an increase of the heat transfer coefficient or of X_{N_2} leads to an increase of the ignition delay time, a shift of the appearance of oscillations with T_{in} is observed in Figs. 7a and 7b.

In addition, it was noted that an oscillatory dynamic regime is also observed under isothermal conditions, as shown in Fig. 7c. Although the temperature is not allowed to vary in this latter case (the shadow area is replaced by a dashed line), oscillations in methane mole fraction over time are highlighted in the insert of Fig. 7c, referring to inlet temperature conditions of 1200 K . Interestingly, methane concentration oscillations occur when T_{in} is in the temperature range $1120 - 1260\text{ K}$. These simulations showed some peculiarity of the oxidation process at intermediate temperature for stoichiometric CH_4/O_2 mixtures:

- the system is prone to exhibit periodic dynamic behavior when the working temperature is approximately in the range 1120 - 1260 K, even under adiabatic or isothermal conditions.
- the high temperature steady state always occurs for $T \geq 1260$ K.

Further simulations were performed for lean mixtures ($\Phi = 0.5$) diluted in N_2 at 90%. They confirmed the presence of an invariant temperature that marks the shift from oscillatory behaviors to high temperature steady states. This temperature differs from the one identified for a stoichiometric mixture, thus further highlighting the chemical nature of the oscillatory phenomenon.

As discussed in the introduction, oscillatory cool flames in LTC conditions arise from the interaction between heat release, complex chemical kinetics, and heat losses, therefore they could not be identified in adiabatic or isothermal conditions. Moreover, cool flame regions for higher hydrocarbons occur at lower temperatures, where peroxide species play the major role in favoring branching or termination. The results discussed in the present study indicate a rather different type of phenomenon occurring in diluted mixtures, at intermediate inlet temperatures (1050 - 1200 K).

As mentioned in the introduction to this study the oscillations observed in the oxidation of hydrogen, or of syngas mixtures, originate from a kinetic interaction between free radical chain branching and termination reactions (i.e. $H + O_2 = OH + O$ vs. $H + O_2 + M = HO_2 + M$). The last reaction is promoted by the high collisional efficiency of water formed during mixture oxidation. In particular, the POLIMI mechanism proposes a collisional efficiency for H_2O 14 times higher than that of nitrogen. To decouple possible relevant effects of reactions belonging to the H_2/O_2 sub-system, further simulations were performed setting H_2O and CO_2 third body efficiencies equal to the nitrogen one. Interestingly, temperature oscillations were

still predicted, thus suggesting that oscillations observed for methane mixtures arise from a different kinetic subset.

Rate of Reaction analyses (RR), flux diagrams, and first-order temperature sensitivity analyses were used to determine the rate limiting reactions for different T_{in} , considering a stoichiometric mixture diluted at 90% in N_2 . The numerical analysis has been performed for simplicity under isothermal conditions.

The most significant and representative results (RR analyses) are reported in Fig. 8, showing the rate of the controlling reactions as a function of the inlet temperature. The aim of Fig. 8 is not to show the whole kinetic routes that lead to methane conversion as a function of T_{in} , but to show how kinetics change from slow combustion to high temperature conditions. In previous papers^{2,26}, the authors highlighted the strong competition between the methyl oxidation channels and methyl recombination to ethane, occurring at intermediate temperatures. Given this information, only the reaction rates involving methyl radicals are reported in Figure 8a.

The highlighted grey area represents the T_{in} range in which oscillations are predicted. The main reaction paths change substantially moving from the region of slow reaction to the oscillatory region and from the oscillatory region to the high temperature region.

At low temperatures ($1000\text{ K} < T_{in} < 1100\text{ K}$), methane is converted to methyl radicals by means of H-abstraction reactions (mostly by OH and HO_2 not reported in Figure 8) which is subsequently oxidized to CH_3O , formaldehyde, formyl radical and CO. Alternatively, two methyl radicals can either react to form ethane ($CH_3 + CH_3 + M = C_2H_6 + M$) terminating the radical chain or enhance reactivity through propagation reactions such as $CH_3 + HO_2 = CH_3O + OH$. This latter reaction in particular, despite being formally a propagation step, converts two relatively unreactive radicals (CH_3 and HO_2) to very reactive radicals (CH_3O and OH).

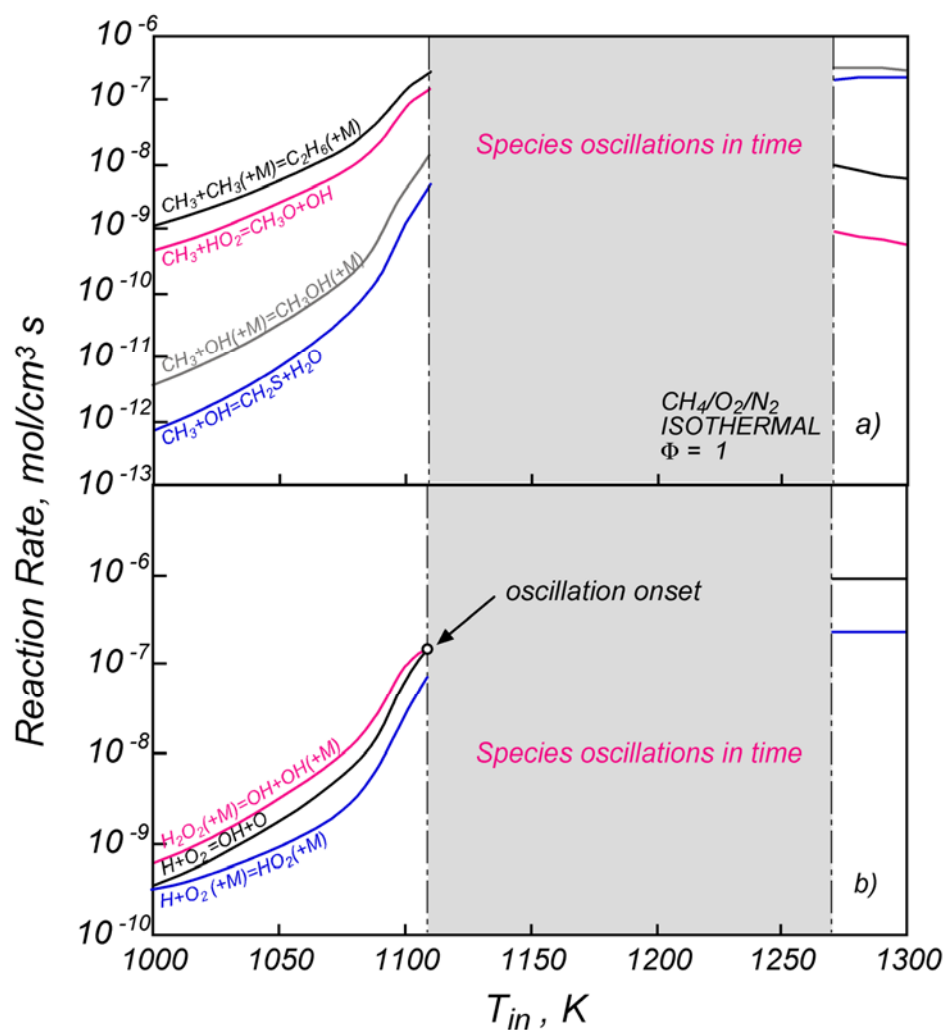


Fig. 8 Rate of reaction analysis for key reactions involving (a) methyl radical and (b) H_2/O_2 subset as a function of T_{in} under isothermal conditions. $\Phi = 1$, $\tau = 0.5$ s, $P = 1.1$ atm, N_2 bath gas. The grey area represents the T_{in} range in which oscillations are predicted.

Due to the relative stability of HO_2 in the temperature window ($1000 \text{ K} < T_{\text{in}} < 1100 \text{ K}$), large amounts of H_2O_2 are formed via H-abstractions. H_2O_2 is rapidly decomposed to OH radicals through the third body reaction $\text{H}_2\text{O}_2 + \text{M} = \text{OH} + \text{OH} + \text{M}$. This strongly enhances reactivity, providing radicals that produce methyl from methane.

At high temperatures ($T_{\text{in}} > 1260 \text{ K}$), methyl radicals mainly react with OH via the reactions $\text{CH}_3 + \text{OH} + \text{M} = \text{CH}_3\text{OH} + \text{M}$ and $\text{CH}_3 + \text{OH} + \text{M} = \text{CH}_2(\text{S}) + \text{H}_2\text{O}$. It is worth noting that, at high temperatures, the RR of these reactions are one order of magnitude higher than the

methyl recombination reaction to ethane. The radicals necessary to sustain the branching mechanism are provided by the reaction $\text{H} + \text{O}_2 = \text{OH} + \text{O}$.

Fig. 8b shows how the onset of oscillations occurs at a noticeable point, i.e. when the low and high temperature branching reaction rates involving H radical and O_2 become equal, better explaining the shift as a function of T_{in} of the transition from low-temperature steady states to periodic oscillatory states observed in Figs. 7a and 7b. Temperatures between 1120 K and 1260 K mark the regime shift between intermediate- and high-temperature chemistries. It has to be pointed out that in Fig. 8b the curve relative to the rate of reaction of the hydroperoxide decomposition is not reported at $T_{\text{in}} > 1260$ K because it becomes negative (thus not diagrammable on log-scale) with negligible values.

For completeness, the same reaction rates are reported in Fig. 9 for one oscillation cycle at $T_{\text{in}} = 1200$ K, for the same reference mixture reported in Fig. 7c. Such operating conditions lead to the establishment of species concentration oscillations in time. Note that methane profile is also reported on the secondary axis.

It is possible to note that during an oscillation period, when methane concentration is relatively high, the methyl recombination to ethane and the oxidation to CH_3O by HO_2 radicals are the most important reactions. The reactions between CH_3 and OH radicals ($\text{CH}_3 + \text{OH} + \text{M} = \text{CH}_3\text{OH} + \text{M}$, $\text{CH}_3 + \text{OH} + \text{M} = \text{CH}_2(\text{S}) + \text{H}_2\text{O}$) become important only when methane is fully converted, due to higher availability of radicals. After methane reaches its minimum value, species accumulated within the reactor mix with the incoming unburned mixture. However, the methyl recombination reaction is still predominant, and subtracts CH_3 radicals to the oxidation pathways thus suppressing the conversion of methane. It accumulates in the reactor restarting the new cycle.

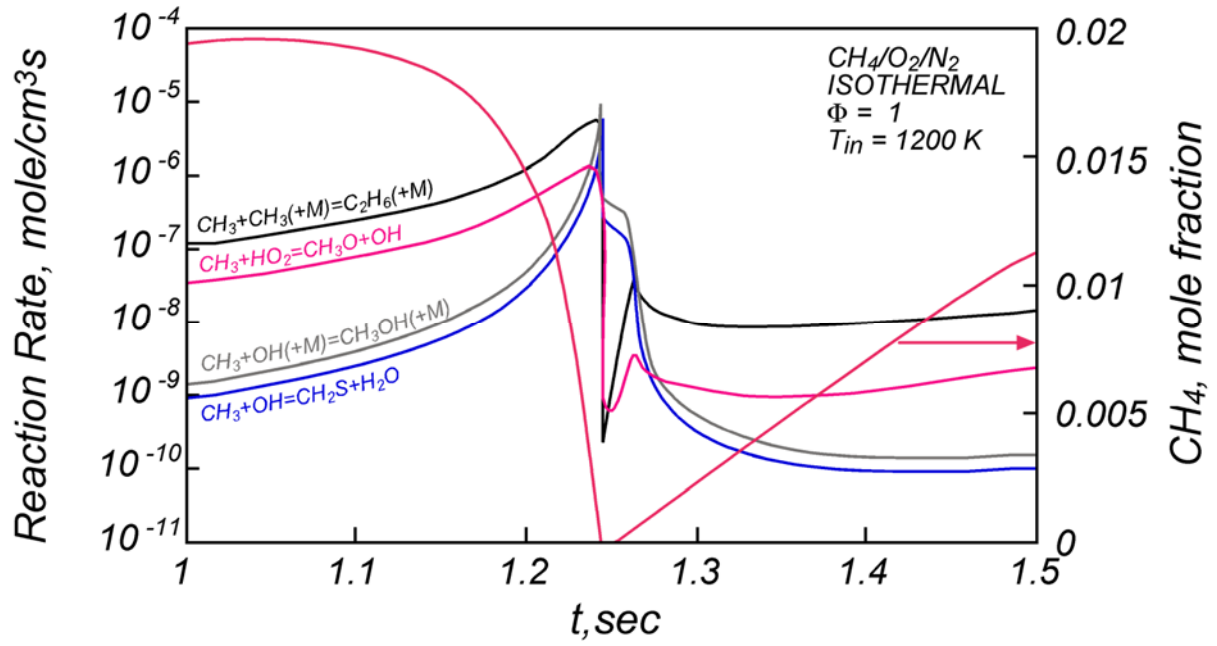


Fig. 9 Rate of reaction analysis for key reactions involving methyl radical (left axis) and methane mole fraction (right axis) as a function of time for one oscillation cycle under isothermal conditions. $\Phi = 1$, $\tau = 0.5$ s, $P = 1.1$ atm, N_2 bath gas, $T_{in} = 1200$ K.

Summarizing, the RR analyses suggest that the oscillations occur when the high temperature branching pathway overcomes the hydrogen peroxide one, thus producing a large amount of radical species that can promote methane oxidation reactions. On the other hand, the recombination channel is very active and subtracts CH_3 radicals to the oxidative channel, thus decreasing the system reactivity.

The dynamic behavior dies out when the reactions between methyl radicals and OH ($CH_3 + OH + M = CH_3OH + M$, $CH_3 + OH + M = CH_2(S) + H_2O$) are fast enough to relieve the system by the inhibiting effect of the recombination reaction.

Therefore, it is clear that diluted combustion conditions force the reactive system to work in a very narrow temperature window where the competition between oxidation and recombination is stressed. This competition results in the peculiar behavior discussed above which is strongly dependent on temperature, equivalence ratio and bath gas composition.

As a matter of fact, ad hoc experimental measurements aiming at testing the effect of equivalence ratio, inlet temperature, and bath gas provide direct evidence for the kinetic foundations discussed above.

In particular, the displacement of the oscillations to a higher inlet temperature at a constant equivalence ratio, when the diluent is changed from He or N₂ to H₂O or CO₂, is consistent with an increase of the overall third-body efficiencies. In fact, the role of termolecular reactions is reinforced by the high collisional efficiencies of such species postponing the shift between the H₂ sub-system branching reactions ($\text{H}_2\text{O}_2 + \text{M} = \text{OH} + \text{OH} + \text{M}$ vs. $\text{H} + \text{O}_2 = \text{OH} + \text{O}$), thus the onset of oscillations, to higher inlet temperatures, and reinforcing the role of methyl recombination to ethane, thus forcing the end of the oscillatory behavior at higher inlet temperatures. In the case of non-isothermal system, it has also to be considered that such species have higher heat capacities with respect to N₂ or He, thus in the same operating conditions, also the reactor working temperature is lower. This can affect the relative weight of competitive reactions.

Despite the discussion here presented mostly refers to chemical kinetics aspects, none of the oscillatory phenomena can be solely attributed to isothermal kinetics. Heat release and dissipation are integral to the occurrence of such events, and are of particular relevance in real devices³⁹. Moreover, the competition between exothermic oxidation and endothermic pyrolysis must be also taken into account.

The kinetic model here used to interpret the experimental measurements shows good agreement with boundaries on the $\Phi - T_{\text{in}}$ maps and for the oscillatory frequencies. However, some disagreements were observed. It has to be pointed out that the evaluation of the heat transfer coefficient used in the adopted model is affected by uncertainties. However, it was observed that such an uncertainty does not significantly change the predictions in the explored experimental conditions. Indeed, a similar satisfactory agreement is observed for

isothermal conditions. However, under isothermal conditions the predicted boundaries diverge towards slightly higher inlet temperatures compared to the measurements. This further confirms the predominant nature of chemical effects in the occurrence of such phenomena.

An indirect proof of this statement is given by the fact that simulations using different kinetic mechanisms available in the literature (see Supporting Information) lead to noteworthy disagreement among model predictions themselves. This suggests that some additional effort should be dedicated to the development of detailed kinetic schemes to correctly describe fuels oxidation processes under diluted combustion conditions, at intermediate temperature.

The general importance of recombination is particularly significant and represents a major limitation in comparing experimental and modeling results since the rate constants for such reactions can vary with the type of bath gas used. Therefore, in order to match modeling to experiment with any accuracy it seems that the third body collisional efficiencies have to be carefully evaluated. Such a consideration is often either very difficult or is neglected.

For all the above reasons, there is still much scope in developing and refining models for the simulation of the large variety of observable dynamic behaviors of methane combustion.

In this regard, the present work provides a safe guidance to reliabilities of modeling, opportunities for further research, and rational employment of efforts.

Conclusion

Two Jet Stirred Flow Reactors were used to characterize the oscillatory behavior of CH_4/O_2 mixtures as a function of system external parameters, such as inlet temperature, equivalence ratio, residence time, methane mole fraction, and the chemical structure of the bath gas (He, N_2 , CO_2 or H_2O), at atmospheric pressure. The agreement between results obtained in the two

employed experimental reactors is certainly highly valuable and inspires considerable confidence of the reliability of results.

The experimental observations revealed several novel features of methane oxidation at intermediate temperatures in presence of high concentration of different bath gases:

- (i) Oscillatory regimes were established for almost all the CH_4/O_2 considered mixtures in both the reactors, when inlet temperatures are approximately between 1050 and 1250 K.
- (ii) The critical inlet temperature for the onset of temperature oscillations diminishes as the mixture equivalence ratio is gradually reduced. As the methane mole fraction is increased, the oscillation region on $\Phi - T_{\text{in}}$ maps widens under stoichiometric conditions, but shrinks under fuel lean conditions.
- (iii) When the equivalence ratio is increased above the stoichiometric value, temperature oscillations are no longer observed for He, N_2 and $\text{N}_2\text{-H}_2\text{O}$ as bath gases, whereas for CO_2 diluted mixtures, they also appear for fuel rich conditions.
- (iv) H_2O and CO_2 lead to a general shift of the oscillation regions toward higher inlet temperatures with respect to N_2 - and He-diluted mixtures.
- (v) The waveforms, frequency and amplitude of oscillations are strongly dependent on inlet experimental conditions.

Summarizing, temperature oscillations occur when the reactor working temperature accesses a determined range. They strongly depend on system intensive and extensive inlet parameters as well as reactor adiabatic features. The nature of the dynamic behavior is essentially related to the chemistry promoted under intermediate working temperatures.

Numerical analyses identified such kinetic features. They originate from the competition between methyl oxidation and recombination kinetic pathways and between low and high temperature branching reactions of the H_2/O_2 sub-system. The end of oscillations occurs at

high temperatures when the inhibiting role of methyl recombination pathways is weakened by the promotion of different methane oxidation routes. It was found that bath gases can play a crucial role to the establishment of oscillations because of different heat capacities, thus different reactor working temperatures, and because of high collisional efficiencies in termolecular reactions, thus altering the relative importance of the identified competitive pathways responsible for temperature oscillations. A more fundamental definition of collisional efficiencies and their dependence on temperature and pressure would also be beneficial to improvements of existing kinetic models.

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Supporting Information.

- Photos of the JSFRs used and of the thermocouple located inside the reactor.
- Experimental data.
- Comparisons of model predictions with experimental measurements using different kinetic schemes.

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