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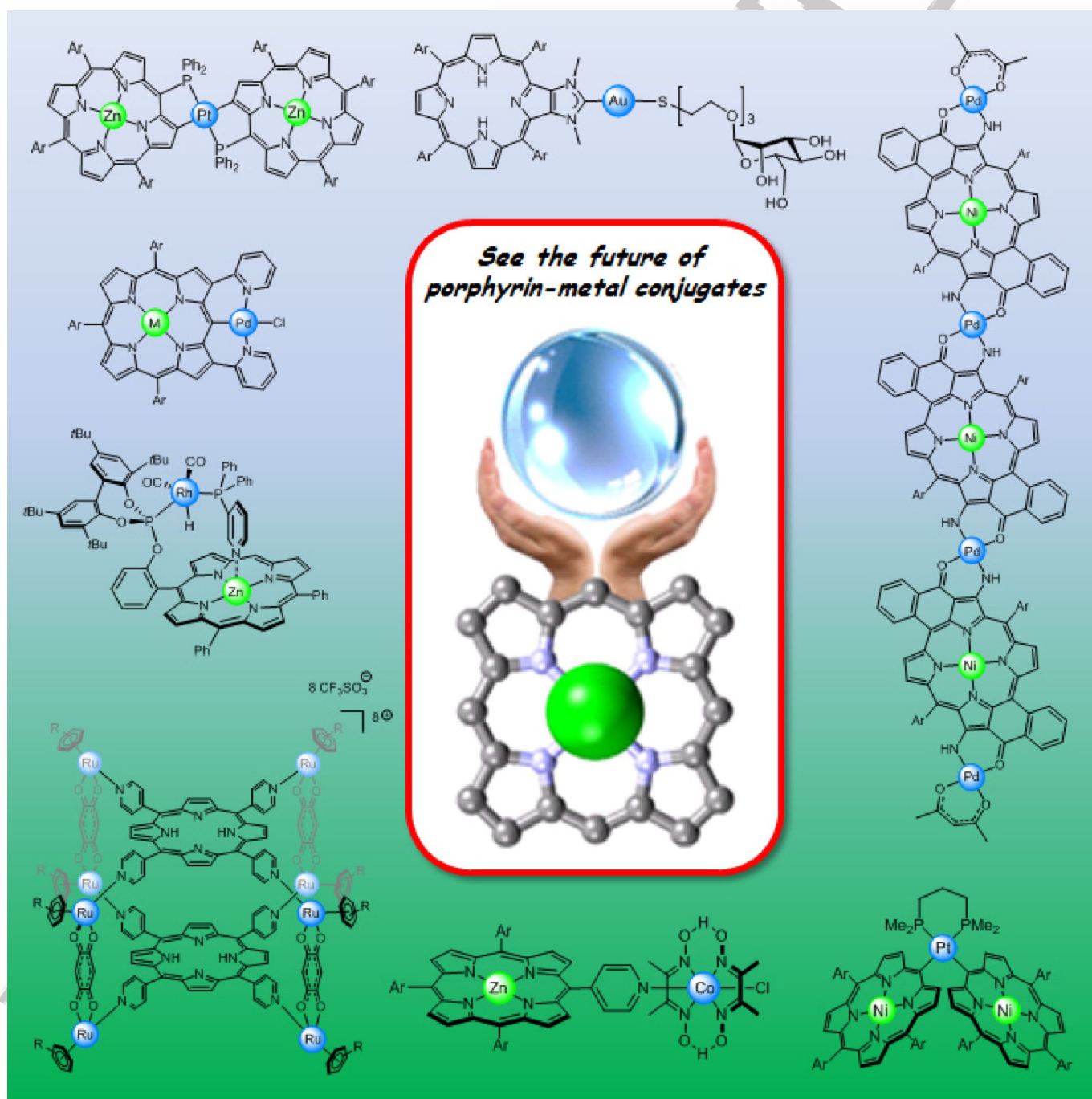
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Peripherally Metalated Porphyrins with Applications in Catalysis, Molecular Electronics and Biomedicine

Jean-François Longevial,^[a] Sébastien Clément,^[a] Jennifer A. Wytko,^[b] Romain Ruppert,^[b] Jean Weiss,^[b] and Sébastien Richeter*^[a]

Dedicated to Henry J. Callot on the Occasion of his 75th Birthday



Abstract: Porphyrins are conjugated, stable chromophores with a central core that binds a variety of metal ions and an easily functionalized peripheral framework. By combining the catalytic, electronic or cytotoxic properties of selected transition metal complexes with the binding and electronic properties of porphyrins, enhanced characteristics of the en-

semble are generated. This review article focuses on porphyrins bearing one or more peripheral transition metal complexes and discusses their potential applications in catalysis or biomedicine. Modulation of the electronic properties and intramolecular communication through coordination bond linkages in bis-porphyrin scaffolds is also presented.

1. Introduction

Natural porphyrins and their congeners, such as chlorins and bacteriochlorins, are considered as “pigments of life” because they are known to play significant roles in important biological processes such as photosynthesis (role of chlorophylls in photosynthetic systems),^[1] the transport and storage of molecular oxygen (role of hemoporphyrins in blood hemoglobin and myoglobin, respectively)^[2] or catalytic oxidation reactions (role of cytochrome P-450 enzymes).^[3] Inspired by Nature, the chemistry of tetrapyrrolic macrocycles, such as porphyrins, phthalocyanines, porphyrazines or corroles, has rapidly grown during the last fifty years and a large variety of porphyrin derivatives are now reported in the literature. The term “metalloporphyrins” (and metal complexes of related phthalocyanines, porphyrazines or corroles) usually refers to compounds where the metal ions are coordinated within the internal cavity. The inner metal in the porphyrin core can play various roles depending on the targeted applications. It can be used as a catalytic site or as a coordination site for the construction of multiporphyrinic systems ranging from finite dimeric species to infinite polymeric materials.^[4]

Continuous advances in the functionalization of the β -pyrrolic and *meso* positions of porphyrins led to the synthesis of numerous porphyrinoids bearing peripheral coordination sites in addition to the classical inner coordination site.^[5] These compounds offer new opportunities to anchor metal complexes at the periphery of porphyrins. Some core-modified porphyrinoids, such as *N*-confused porphyrins^[6] and porphyrazines,^[5] possess nitrogen atoms in place of carbon atoms at their periphery. Although these nitrogen atoms can be used to anchor peripheral metal complexes, these fascinating compounds are not discussed in this review that focuses on porphyrins only. Moreover, only porphyrins containing *peripheral transition metal complexes* are presented. Therefore, porphyrins contain-

ing, for example, appended crown ethers complexing alkali and alkaline earth metal ions are not included since they are discussed elsewhere.^[7] Here, we present selected examples of peripherally metalated porphyrins where the outer transition metal complexes can play key roles in catalysis, molecular electronics and biomedicine (Figure 1).

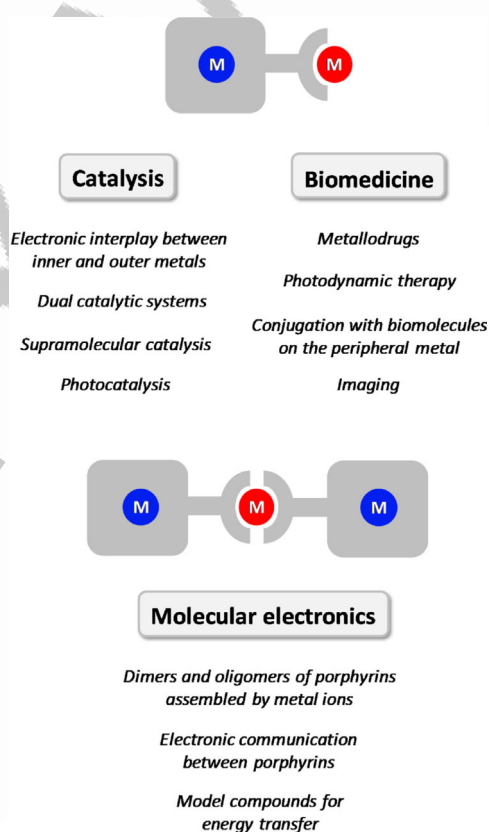


Figure 1. Topics discussed in this review.

2. Catalysis with Peripherally Metalated Porphyrins

To explain enzymatic processes, iron porphyrins were actively investigated as mimics of cytochrome P-450 to understand the enzyme’s catalytic activity. Besides these iron porphyrins, different metalloporphyrins catalyzed other reactions, including amination (with Co^[8]), cyclopropanation (Rh,^[9] Co,^[10] Ir^[11]) or olefination (Fe^[12]) reactions. In this context, the use of peripherally metalated porphyrins for catalysis is a relatively recent field of research (less than two decades old).^[13] In this case, the periph-

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eral transition metal complexes play the role of catalytic sites, whereas the neighboring porphyrin can act: (i) as a modulating unit to tune the electronic and catalytic properties of the peripheral catalytic sites, (ii) as a secondary coordination site for supramolecular or dual catalysis, or (iii) as a light harvesting unit for dual photoredox and transition metal catalysis. These aspects are discussed hereafter.

2.1. Modulation of the electronic and catalytic properties

It was expected that the electronic and catalytic properties of the peripheral metal catalyst(s) could be modulated by the porphyrin due to possible electronic interplay between the outer and inner metal ions. Mizoroki-Heck coupling reactions between aryl halides and alkenes were extensively used by several research groups as "benchmark reactions" to evaluate the catalytic properties of porphyrins with peripheral Pd^{II} complexes. In their pioneering work, Klein, Gebbink and co-workers

reported the synthesis of a series of symmetrical (metallo)porphyrins with four meso aryl groups bearing pincer-type ligands with NCN, SCS or PCP donor atoms and the corresponding Pd^{II} complexes.^[14] The SCS-Pd^{II} complexes **1-M** (M = Mg^{II}, Ni^{II}, Mn^{III}Cl and 2H) were tested as pre-catalysts in the Mizoroki-Heck reaction between iodobenzene and styrene (Scheme 1).^[15] Catalytic activities depended on the nature of the metal ions inside the porphyrin, and increased in the order **1-Mn^{III}Cl** < **1-2H** < **1-Ni^{II}** < **1-Mg^{II}**. Complex **1-Mg^{II}** was the most active catalyst with a conversion of >95% after 5 hours. Complex **1-Mn^{III}Cl** exhibited the lowest activity with a 92% conversion after 45 hours. These catalytic activities are in line with the relative electron-donating abilities of the porphyrins. Control experiments confirmed that the differences in catalytic activities arose from intramolecular interactions rather than intermolecular interactions. The main outcome of the study was that leached Pd particles were the possible catalytically active species. Following this observation, it was proposed that the nature of the inner

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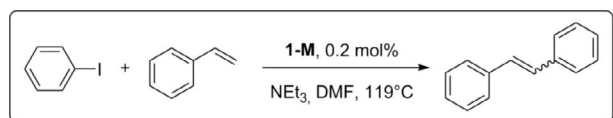
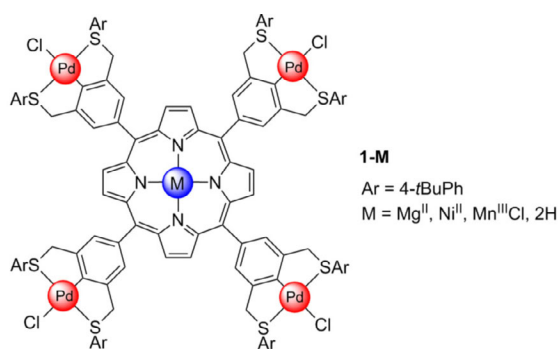
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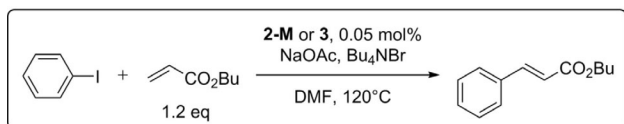
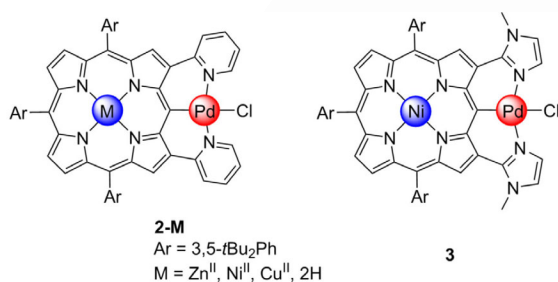
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Scheme 1. Mizoroki-Heck reaction catalyzed by Pd complexes **1-M**.

metal ion in the porphyrin directly influenced the electron-donating ability of the porphyrin and, hence, facilitated (or not) the leaching of Pd in the reaction mixture.

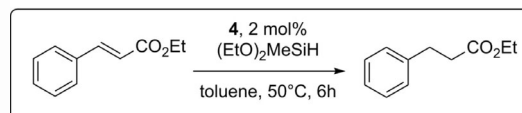
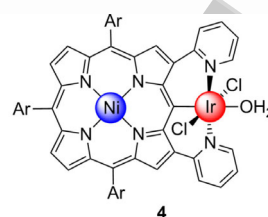
Stronger electronic interactions between the porphyrin and peripheral metal complexes are expected in the absence of orthogonal *meso* aryl groups. For example, the *meso* triarylporphyrin containing two pyridyl groups at the β -pyrrolic positions adjacent to the free *meso* carbon (see Scheme 2)^[16] pro-



Scheme 2. Mizoroki-Heck reaction catalyzed by Pd complexes **2-M** and **3**.

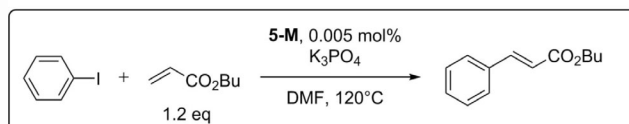
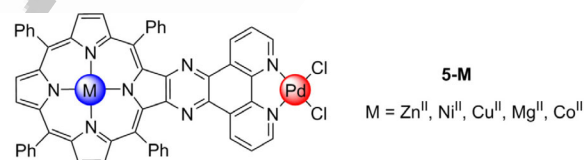
vides an NCN pincer-type ligand was used for the synthesis of the corresponding Pd^{II} complexes **2-M** (M = Zn^{II}, Ni^{II}, Cu^{II} and 2H). When these complexes were used as pre-catalysts in the Mizoroki-Heck reaction between iodobenzene and butyl acrylate, the catalytic activity strongly depended on the nature of the inner metal ion and increased in the order **2-2H** < **2-Cu^{II}** < **2-Ni^{II}** < **2-Zn^{II}**. This observation was not discussed at that time but it was clear that the inner metal ion had a strong impact on the catalytic activity. Interestingly, when pyridyl ligands were replaced by *N*-methyl imidazole, the apparent catalytic activity of complex **3** in the same reaction was enhanced compared to that of **2-Ni^{II}**.^[17] This enhancement showed that a fine-tuning of the catalytic properties was achieved by modu-

lating the nature of the ligands at the β -pyrrolic positions (Scheme 2). The same pincer ligand with pyridyl groups was used for the synthesis of the Ir^{III} complex **4**, which was catalytically active towards the hydrosilylation of ethyl cinnamate (Scheme 3).^[18]



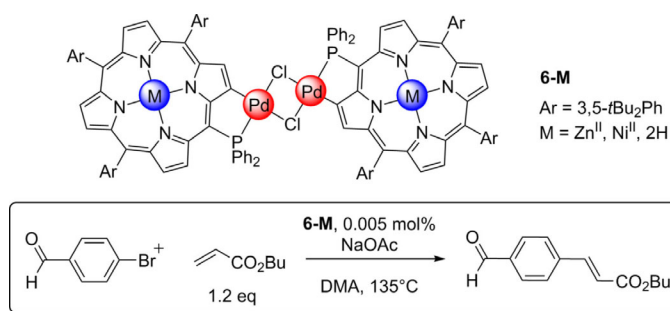
Scheme 3. Hydrosilylation reaction catalyzed by Ir complex **4**.

A porphyrin conjugated with a phenanthroline ligand was used in the synthesis of Pd^{II} complexes **5-M** (M = Zn^{II}, Ni^{II}, Cu^{II}, Mg^{II} and Co^{II}).^[19] These complexes were used as pre-catalysts for the Mizoroki-Heck reaction between iodobenzene and butyl acrylate. Similarly, to the results previously described, the observed catalytic activity was found to significantly depend on the inner metal ion. The complex **5-Zn^{II}** appeared to be the most active catalyst with a 100% conversion reached after 30 minutes. All other complexes were less active and **5-Co^{II}** was the least active.



Scheme 4. Mizoroki-Heck reaction catalyzed by Pd complexes **5-M**.

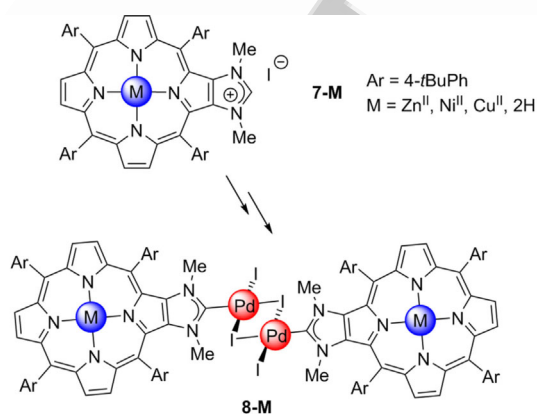
A Zn^{II} *meso* triarylporphyrin, where the fourth *meso* position was functionalized with a diphenylphosphanyl group,^[20] was used to prepare a Zn^{II} bis(μ -chloro)-bridged porphyrin appended phosphapalladacycle **6-Zn^{II}**, which was further converted to the free-base- and Ni-porphyrin analogs **6-2H** and **6-Ni^{II}**. These complexes were tested as pre-catalysts in the Mizoroki-Heck reaction between *p*-bromobenzaldehyde and butyl acrylate (Scheme 5). The observed catalytic activity significantly depended on the metal in the porphyrin and increased in the order **6-2H** < **6-Zn^{II}** < **6-Ni^{II}**. However, this order does not coincide with the electron-donating ability of the porphyrin π -system, estimated from the first oxidation potentials. However,



Scheme 5. Mizoroki-Heck reaction catalyzed by Pd complexes **6-M**.

this order did correlate with the σ -donating ability of the phosphorus atom (estimated by ³¹P NMR spectroscopy), which increases in the order: **6-Ni^{II}** < **6-2H** < **6-Zn^{II}**. Comparisons with the previously described Mizoroki-Heck reactions are difficult because aryl halides and catalyst loadings were different.

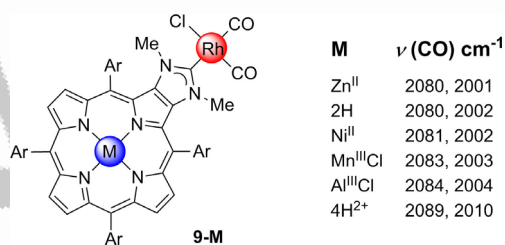
Since palladacycles readily decompose during Mizoroki-Heck reactions (see the previous examples), other types of peripheral ligands were studied. *N*-heterocyclic carbenes (NHCs) are known to form very stable bonds with transition metal ions in low and high oxidation states and thus are actively used in organometallic chemistry and catalysis. Porphyrins fused to imidazolium salts **7-M**^[21] serve as NHC precursors for the synthesis of bis(μ -iodo)-bridged Pd^{II} complexes **8-M** (M = Ni^{II}, Cu^{II}, Zn^{II} and 2H).^[22] These complexes were tested as pre-catalysts in the Mizoroki-Heck reaction between iodobenzene and styrene (Scheme 6) and the observed catalytic activity depended on the absence or presence of a metal ion in the porphyrin. A conversion between 80–90% of iodobenzene was observed with compounds **8-M** (M = Ni^{II}, Cu^{II} and Zn^{II}), whereas only 30% was observed with the free-base porphyrin **8-2H** after 10 hours. Pd leaching occurred during the reaction and it was shown that leached Pd could be trapped by the inner coordination site of the porphyrin, thus decreasing the concen-



Scheme 6. Mizoroki-Heck reaction catalyzed by Pd complexes **8-M**.

tration of free Pd particles available for the coupling reaction. Interestingly, in almost all Mizoroki-Heck reactions presented so far in this section, the pre-catalysts incorporating free-base porphyrins are among the least active ones. Of course, steric and electronic properties are important, but other factors (axial coordination, solubility, etc.) should also be taken into account, including the possible trapping of Pd in the porphyrin macrocycle.

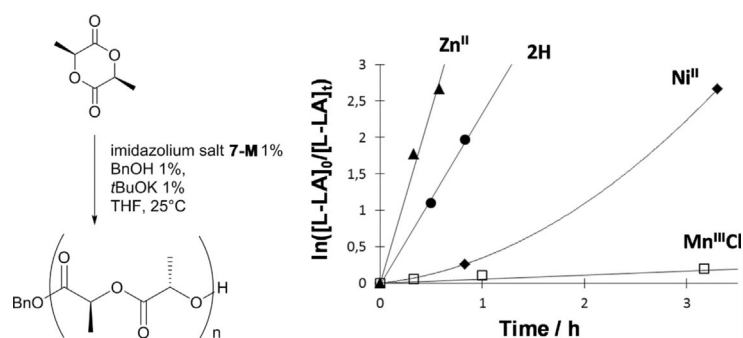
Many experimental observations made in the course of these investigations suggest that Pd^{II} complexes decompose and release Pd⁰ colloids at elevated temperature, including when NHC-Pd^{II} complexes are used as pre-catalysts.^[23] The Mizoroki-Heck reaction was probably not the best choice to demonstrate the electronic interplay between the inner metal in the porphyrin and the peripheral catalytic site. Therefore, other reactions and catalysts were investigated. In the course of their studies, Richeter and co-workers synthesized the Rh^I complexes **9-M** (M = Ni^{II}, Zn^{II}, Mn^{III}Cl, Al^{III}Cl, 2H and 4H²⁺, see Scheme 7).^[24] The electron donating abilities of the peripheral



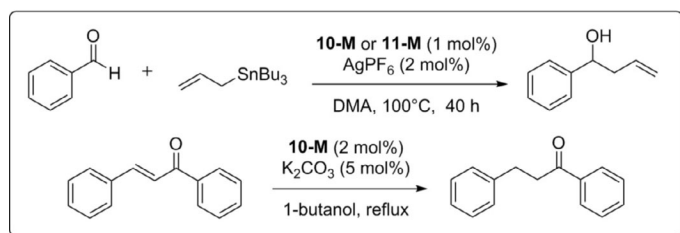
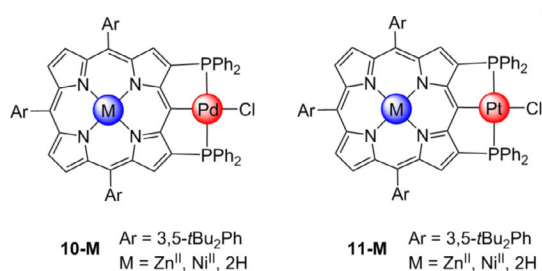
Scheme 7. Structures of Rh^I complexes **9-M**, and $\nu(\text{CO})$ values measured by IR spectroscopy.

NHCs were evaluated from the stretching vibrations of the CO ligand on Rh^I complexes **9-M**. The σ -donating ability of the NHCs increased in the order M = 4H²⁺ < Al^{III}Cl \approx Mn^{III}Cl < Ni^{II} \approx 2H \approx Zn^{II}.^[24,25] IR spectroscopy was not sensitive enough to distinguish between **9-Zn^{II}**, **9-Ni^{II}** and **9-2H**. However, the very small differences observed are in good agreement with the π -donating ability of the porphyrins, estimated from the first oxidation potentials of imidazolium salts **7-M**. These potentials are shifted anodically in the order **7-Zn^{II}** (+0.36 V) < **7-2H** (+0.57 V) < **7-Ni^{II}** (+0.64 V). Imidazolium salts **7-M** (M = Ni^{II}, Zn^{II}, Mn^{III}Cl, 2H; see Scheme 6) were used as NHC precursors to catalyze the ring-opening polymerization (ROP) of L-lactide (L-LA) (Scheme 8).^[24] The observed catalytic activity strongly depended on the metal ion inside the porphyrin and increased in the order **7-Mn^{III}Cl** < **7-Ni^{II}** < **7-2H** < **7-Zn^{II}**. This order coincides with the σ -donating ability of the NHCs as well as the π -donating ability of the porphyrins; the more electron-rich the porphyrin, the more catalytically active the peripheral NHC is towards the ROP of L-LA.

A phosphorus-containing PCP pincer-type ligand and the corresponding Pd^{II} **10-M** and Pt^{II} **11-M** complexes (M = Ni^{II}, Zn^{II} and 2H)^[26] were tested in the allylation reaction of benzaldehyde with allyltributyltin, because this reaction is known to proceed without significant decomposition of the catalyst (Scheme 9). However, the obtained data showed similar cata-



Scheme 8. Polymerization of L-LA catalyzed by porphyrin-NHCs generated in situ by deprotonating imidazolium salts **7-M**. Plots of $\ln([L-LA]_0/[L-LA]_t)$ versus time for $[L-LA]_0 = 2$ M and a targeted degree of polymerization of 100.



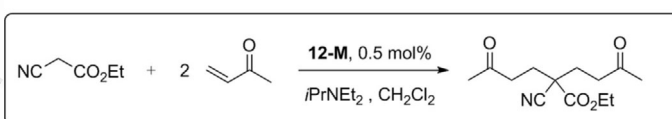
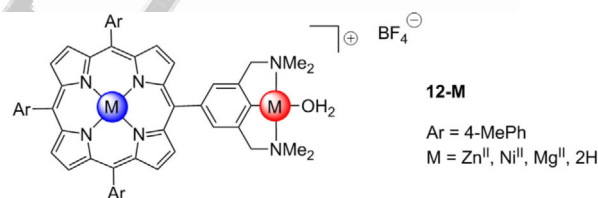
Scheme 9. Allylation reaction of benzaldehyde with allyltributyltin and 1,4-reduction of chalcone through transfer hydrogenation catalyzed by Pd^{II} complexes **10-M** and Pt^{II} complexes **11-M**.

lytic activity for all catalysts, implying that the inner metal had little or no influence on the peripheral catalytic sites. The Pd^{II} complexes **10-M** (M = Zn^{II}, Ni^{II} and 2H) were also tested for the 1,4-reduction of chalcone through transfer hydrogenation (Scheme 9). In this case, the catalytic activity depended on the nature on the inner metal and increased in the order **10-2H** < **10-Ni^{II}** < **10-Zn^{II}**. The higher activity observed for **10-Zn^{II}** is in agreement with its lower oxidation potentials compared to those of **10-Ni^{II}** and **10-2H**, that is, the Zn^{II} porphyrin is more electron-rich than the other two. It is important to note here that mercury poisoning tests excluded the formation of Pd nanoparticles in the course of the reaction.

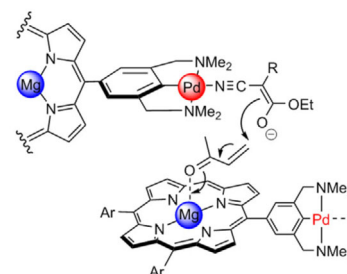
2.2. Dual catalytic systems and supramolecular assemblies

As seen above, the modulation of the electronic properties of the aromatic core by the nature of the metal ion inside the porphyrin can have significant impact on the electronic and/or catalytic properties of peripheral metal complexes. However, the central metal ion can also be used as a secondary catalytic site for dual catalysis. Cationic Pd^{II} complexes **12-M** (M = Mg^{II},

Ni^{II}, Zn^{II} and 2H; Scheme 10) constructed from porphyrins bearing an NCN pincer-type ligand were tested as catalysts in the double Michael addition of ethyl- α -cyanoacetate with methyl vinyl ketone.^[27] The observed catalytic activity did not depend of the porphyrin when M = Zn^{II}, Ni^{II} or 2H and was similar to the activity of the corresponding non-porphyrin complex. Interestingly, the catalytic activity increased by a factor of six when M = Mg^{II}. Control experiments suggest that simultaneous activation of methyl vinyl ketone by coordination of the carbonyl group to the Lewis acidic Mg^{II} site and activation of ethyl- α -cyanoacetate bound to the cationic Pd^{II} complex within supramolecular assemblies may explain the enhanced activity. Interestingly, combining Mg^{II}TTP and cationic [Pd(NCN)OH₂]⁺BF₄⁻ into one molecule **12-Mg^{II}** led to a three-fold increase of its catalytic activity compared to that found for a 1:1 mixture of its constituents. Indeed, the Mg^{II} porphyrin and the peripheral Pd^{II} complex act in a cooperative way when they are connected together.



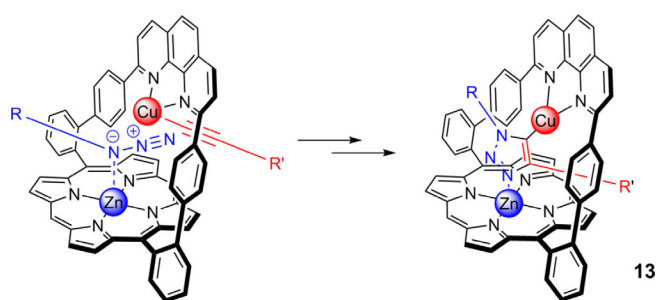
Possible mechanism:



Scheme 10. Double Michael addition of ethyl- α -cyanoacetate with methyl vinyl ketone catalyzed by complexes **12-M**. Below: possible mechanism to explain the increased catalytic activity when M = Mg^{II}.

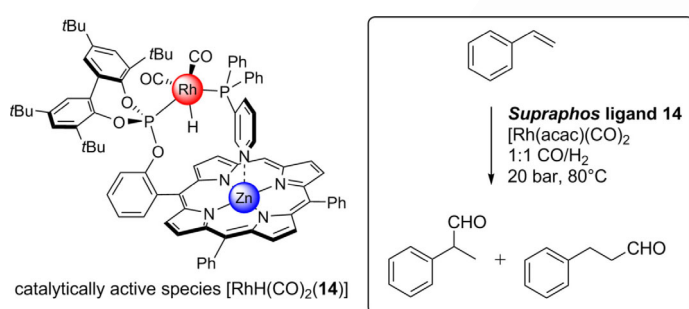
Other reported examples of such “cooperative dual catalytic systems” include the synthesis of phenanthroline-strapped porphyrin [2]rotaxanes by a copper-catalyzed alkyne-azide cycloaddition.^[28] The cooperative and simultaneous activation of both the azide, by the Zn^{II} porphyrin of **13**, and of the alkyne, by the phenanthroline-bound Cu^I, dramatically improved the yield of rotaxanes (Scheme 11).

Metalloporphyrins are also attractive molecules in the field of supramolecular chemistry. Well-defined finite species can be obtained by self-assembly via axial coordination to metal ions such as Zn^{II} in the porphyrins. In catalysis, this supramolecular



Scheme 11. Synthesis of phenanthroline-strapped porphyrin [2]rotaxane **13** through a Cu-catalyzed alkyne-azide cycloaddition reaction.

approach was applied to construct bidendate ligands by combining two phosphorus-based ligands, one with a pyridyl ligand as donor and another appended to a Zn^{II} porphyrin (Scheme 12).^[29] The self-assembled *Supraphos* ligand **14** was



Scheme 12. Schematic representation of a Rh^I complex with *Supraphos* ligand **14** for the Rh-catalyzed hydroformylation of styrene.

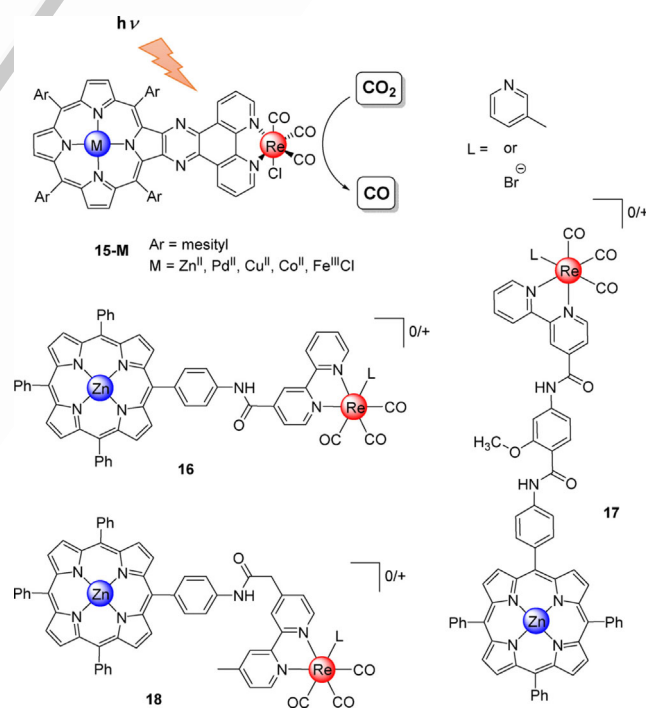
successfully used for the challenging Rh-catalyzed, hydrogenation of enamide substrates, the Pd-catalyzed kinetic resolution of racemic cyclohexenyl acetate, and the Rh-catalyzed hydroformylation of styrene. The powerful aspect of this supra-molecular approach is that a vast library of 48 bidendate ligands was obtained from only six porphyrin derivatives and eight monodentate phosphorus ligands.

2.3. Photocatalysis

Generally, Ru and Ir complexes are at the forefront of photoredox catalysis because of their high redox potentials in the excited states.^[30] Visible light is used to perform photoinduced electron transfer with substrates through oxidative/reductive quenching pathways. Combining a photosensitizer with a second metal catalyst leads to dual catalytic systems with novel reactivity that is sometimes not accessible using either of the compounds independently.^[31] Most of the time, the two components are not linked together (bimolecular systems). Dual catalytic systems where the two components are connected, that is, in close proximity (unimolecular systems), and able to act in a cooperative way are less common. Some challenging reactions could surprisingly be achieved or enhanced with such systems under visible light irradiation ("photo-assisted" catalysis). Porphyrin derivatives, which have redox poten-

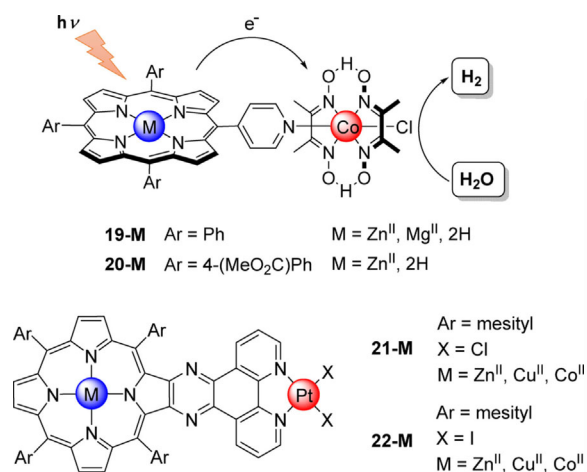
tials in their excited states and their ground states close to those of [Ru(bpy)₃]²⁺,^[32] can be used for photocatalytic applications such as reduction of CO₂ to CO or H₂ evolution.

Converting CO₂ into valuable chemicals is of interest especially if solar energy is employed. Towards this goal, a series of Re^I complexes **15-M** (M = Cu^{II}, Pd^{II}, Zn^{II}, Co^{II} and Fe^{III}Cl) based on a porphyrin conjugated with a phenanthroline ligand.^[33] were tested as catalysts for the photochemical reduction of CO₂ to CO in DMF using triethylamine as a sacrificial electron donor (Scheme 13). For these heterodinuclear complexes, the active catalytic center is the peripheral Re^I complex and not the (metallo)porphyrin. The nature of the metal center in the porphyrin, as well as the irradiation conditions, strongly influences the reduction of CO₂ to CO. The Re^I complex conjugated with the Zn^{II} porphyrin **15-Zn** is the most active photocatalyst with a turnover number (TON_{CO}) of 13 after 24 hours of irradiation at λ > 375 nm. All other photocatalysts show minimal activity under these conditions. Investigation of the photoreduction of CO₂ by three types of Zn^{II} porphyrin-Re^I dyads **16**, **17** and **18** showed that the TON_{CO} obtained with **18** was 360 after 7 hours of irradiation at λ > 520 nm (Scheme 13).^[34] This value is one order of magnitude higher than those of **16** and **17**. It was also noticed that the Zn^{II} porphyrins were reduced to Zn^{II} chlorins (and Zn^{II} isobacteriochlorins), which are also active sensitizers. Complete bleaching renders the dyads inactive. Interestingly, the use of a saturated bridge between the porphyrin and the Re^I complexes was beneficial and in accord with results obtained for Ru-Re complexes.^[35]



Scheme 13. Structures of porphyrin Re^I complexes **15-M**, **16**, **17** and **18** for the photoreduction of CO₂ to CO.

Photochemical production of H₂ with noble-metal-free molecular species is still a hot topic for the development of future cost-effective, environmentally benign sources of energy. Complexes belonging to the class of cobaloximes are active catalysts for the photoinduced generation of H₂. Sun and co-workers reported the synthesis of porphyrins **19-M** (M = Zn^{II}, Mg^{II} and 2H) with one *meso* 4-pyridyl group that acts as an apical ligand for the complex [Co^{III}(dmgH)₂(Cl)]^{±0} (dmg = dimethylglyoxime, Scheme 14).^[36] Photoinduced (λ > 400 nm) H₂ gener-



Scheme 14. Structures of porphyrin cobaloxime complexes **19-M** and **20-M**, and Pt^{II} complexes **21-M** and **22-M** for photoinduced H₂ generation.

ation, in a THF–H₂O solution containing NEt₃, with **19-M** as photocatalysts, strongly depends on the inner metal of the porphyrin. A TON_{H₂} of 22 was observed after 5 hours with **19-Zn^{II}**, whereas only a value of 3 was reached for **19-Mg^{II}** after 3 hours. Only trace amounts of H₂ were detected for **19-2H** under these conditions. Guldi, Coutsolelos, Gryko and co-workers investigated photophysical and catalytic properties of similar dyads **20-M** (M = Zn and 2H) and discussed the possible photochemical processes leading to the formation of catalytically active species (Scheme 14).^[37] Again, the photocatalyst **20-Zn^{II}** was the most efficient for H₂ evolution. In other photocatalytic studies with Pt^{II} complexes **21-M** and **22-M**^[38] based on a porphyrin conjugated with a phenanthroline ligand, the catalytic evolution of H₂ depended on the nature of M in the inner core of the porphyrin and the halides bound to the Pt^{II} center. The optimal photocatalyst was **21-Zn^{II}** with X = Cl (TON_{48h} = 25 and 2, respectively for **21-Zn^{II}** and **22-Zn^{II}**; λ = 470 nm in CH₃CN/trimethylamine/H₂O 6:3:1).

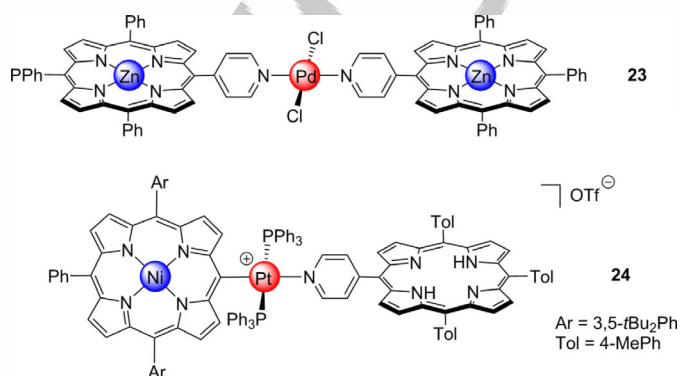
3. Molecular Electronics with Porphyrins Assembled by Transition Metal Ions

Mimicking the properties of natural systems with synthetic porphyrin derivatives has been the subject of intense research during the last few decades. Porphyrin dimers and oligomers assembled by covalent or non-covalent bonds have been investigated to better understand the electronic interactions between the chromophores in their ground and excited states. In

this section, some representative examples of porphyrin dimers and oligomers assembled by transition metal ions are presented along with the factors important for efficient electronic communication between the porphyrins.

3.1. Factors governing electronic interactions between porphyrins in the ground state

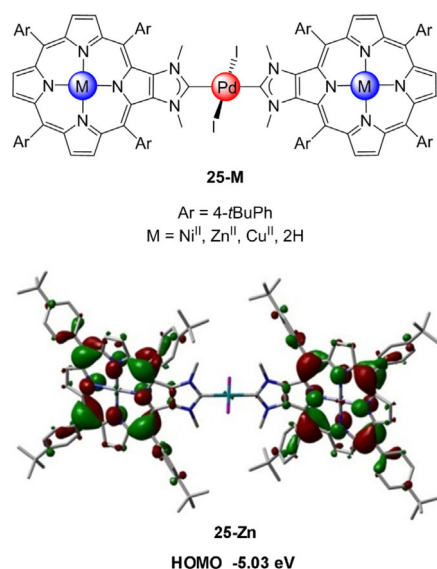
An early example of a porphyrin dimer assembled by a transition metal was reported by Lehn and Drain (Scheme 15).^[39] In dimer **23**, *meso* 4-pyridyl units are used as peripheral ligands and square-planar Pd^{II} ions with a *trans* geometry are used as



Scheme 15. Structures of dimeric species **23** and **24**.

assembling metals. The electronic interactions between the porphyrins through the assembling metal were evaluated in the ground state by UV-visible spectroscopy. Soret absorption bands of such porphyrin dimers are moderately red-shifted (6–8 nm) and extinction coefficients *per* porphyrin (≈80%) are substantially lower than that of the parent porphyrin. These observations suggest that electronic communication between the porphyrins is quite weak. The lack of interactions between the porphyrins of dimer **23** can be attributed to the fact that the *meso* pyridyl ligands are nearly orthogonal to the porphyrin planes. Although porphyrin-porphyrin interactions are weak, this is an extremely powerful approach to build complex multiporphyrinic architectures.^[40] Porphyrin dimer **24** and other oligomers with similar pyridyl bridging units (Scheme 15)^[41] contain a platinum ion directly linked to the porphyrin core. Despite the direct σ-bond between one porphyrin and the Pt^{II} complex, only weak electronic interactions were observed between the porphyrins (Soret bands are red-shifted by ≈1–5 nm when inner M = Ni^{II}). As stated by the authors, the bridging metal fragment “acts more as a structural component than a conjugative linking group” in this case.

With this in mind, porphyrins with coplanar and conjugated peripheral coordination sites were expected to display larger electronic interactions between the porphyrins within dimeric structures. However, this is not always the case, as seen for *trans*-[Pd^{II}]₂(NHC)₂ complexes **25-M** in which two porphyrin-NHC ligands are coordinated to Pd^{II} (Scheme 16).^[42] Clearly, the UV-visible spectra of the dimeric structures **25-M** correspond to the sum of the spectra of the individual porphyrins, and the

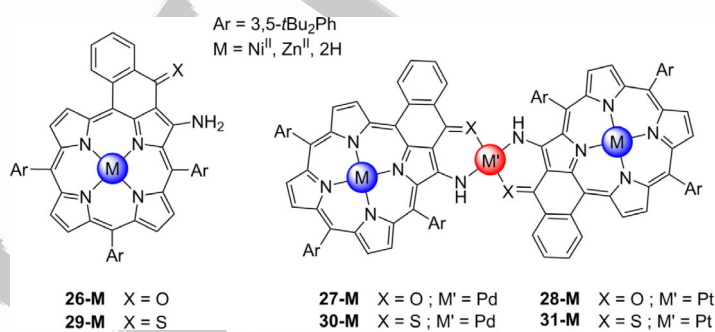


Scheme 16. Structures of Pd^{II} complexes **25-M** and representation of the HOMO orbital of the **25-Zn** complex calculated by DFT.

cyclic voltammograms (CVs) of complexes **25-M** are similar to the sum of the CVs of the individual porphyrins. These data suggest that there is no electronic communication between the porphyrins despite the conjugation between the porphyrin core and the NHC ligands. DFT calculations corroborated the experimental findings: no significant atomic contribution of the coordination unit (Pd^{II} and carbene) was found in the frontier orbitals, thus preventing orbital overlap between the porphyrins through the assembling metal.^[43]

In previous systems **23–25** porphyrin cores, peripheral ligands and/or assembling metal (Pd^{II} or Pt^{II}) are not coplanar. The planarity of the entire assembly can be enforced by introducing peripheral chelating sites conjugated with the porphyrin π -system. Porphyrins conjugated with an enamino-ketone **26-M** were used to construct fully planar dimeric species where the porphyrins are assembled by square-planar metal ions such as Pd^{II} or Pt^{II} (Scheme 17).^[44] In these dimers, **27-M** and **28-M**, strong electronic interactions were observed in the ground state between the porphyrins.^[45]

The UV-visible spectra of the dimers **27-M** do not correspond to the sum of the UV-visible spectra of the individual



Scheme 17. Structure of porphyrins conjugated with enamino(thio)ketone **26-M** and **29-M** and dimeric species with Pd^{II} (**27-M** and **30-M**), and Pt^{II} (**28-M** and **31-M**) as assembling metal ions.

porphyrins **26-M**. The intense Soret absorption bands are broad and split and the Q absorption bands (500–700 nm) are bathochromically shifted with respect to the monomeric porphyrins (See Figure 2a for UV-visible absorption spectra of monomer **26-Ni** and dimer **27-Ni**). The electronic communication between the porphyrins through the assembling metal was

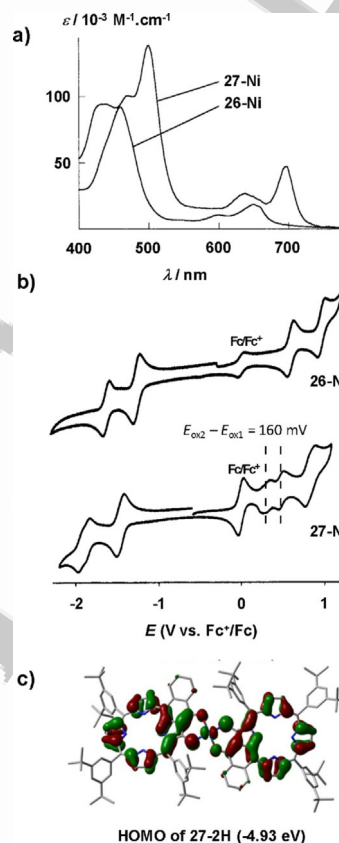


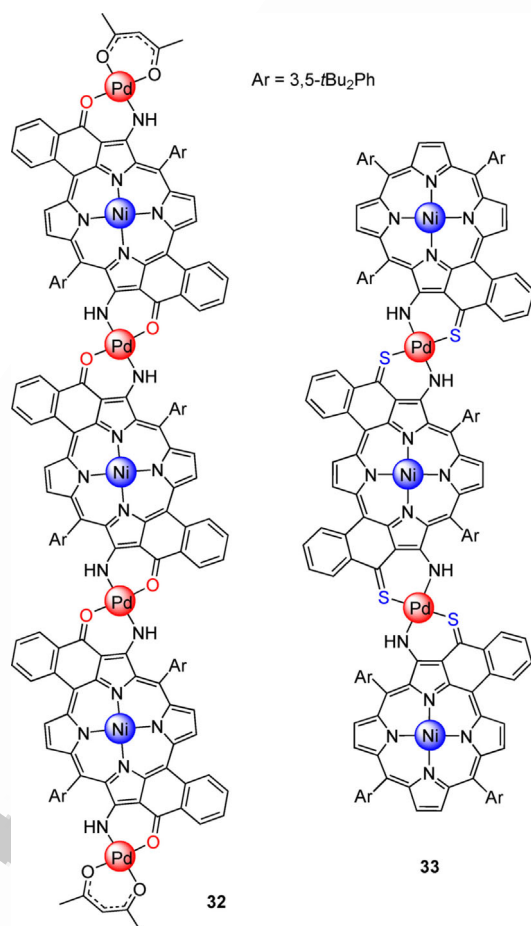
Figure 2. a) UV-visible absorption spectra of **26-Ni** and **27-Ni**. b) Cyclic voltammograms of **26-Ni** and **27-Ni**. c) Representation of the HOMO of Pd^{II} complex **27-2H** calculated by DFT.

confirmed by electrochemical studies. For monomeric porphyrins **26-M**, two monoelectronic reversible oxidation and two monoelectronic reversible reduction processes were observed by cyclic voltammetry. The cyclic voltammograms of the dimeric species **27-M** did not correspond to the sum of the electrochemical properties of individual porphyrins and some significant changes were observed. First, electrochemical gaps ($\Delta E = E_{\text{ox}1} - E_{\text{red}1}$) decreased when going from monomers **26-M** to dimers **27-M**. Second, some redox potentials are split, notably the first oxidation waves, which are split into two well-defined one-electron processes with splitting values ($\Delta E = E_{\text{ox}2} - E_{\text{ox}1}$) ranging from 120–180 mV for dimers **27-M** (splitting value is 160 mV for **27-Ni**, see Figure 2b). This phenomenon is also observed in conjugated dimeric species where porphyrins are linked together by covalent organic spacers (80 mV and 260 mV for a C=C and C \equiv C, respectively).^[46,47] The highest splitting value reported so far is 410 mV

for C_{β} - C_{meso} - C_{β} triply fused Ni^{II} porphyrin dimers reported by Osuka and co-workers.^[48] The splitting of the first oxidation wave is due to the delocalization of the resulting radical cation over both porphyrins. This delocalization renders the oxidation of the second porphyrin more difficult. DFT calculations showed a significant atomic contribution of the d orbitals of the linking metal to the frontier molecular orbitals, allowing electronic communication between the two porphyrins (Figure 2c).

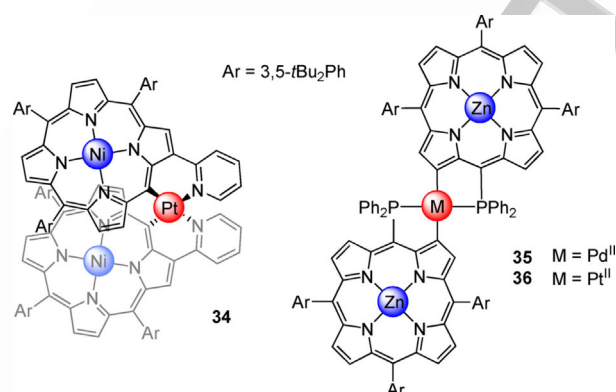
Interestingly, replacement of the oxygen atom of the chelating site by a sulfur atom was feasible with Lawesson's reagent (see structure of **29-M**).^[49,50] This small change led to enhanced electronic interactions between the porphyrins in the dimers **30-M** (splitting value is 190 mV for **30-Ni**). Replacement of Pd^{II} by Pt^{II} also dramatically enhanced the electronic communication between the porphyrins in dimers **28-M** and **31-M** as demonstrated by splitting values of the first oxidation waves of 220 and 280–290 mV for **28-Ni** and **31-Ni**, respectively. Higher oligomers, such as trimers **32** and **33** linked by coordination bonds (Scheme 18) were constructed from porphyrins bearing two peripheral chelating sites.^[51] Again, electronic interactions between the individual porphyrins were observed.^[52]

Meso triarylporphyrins in which the fourth *meso* position was unsubstituted and one of the two neighboring β -pyrrolic positions was functionalized with a pyridyl group were used as



Scheme 18. Structures of trimers **32** and **33**.

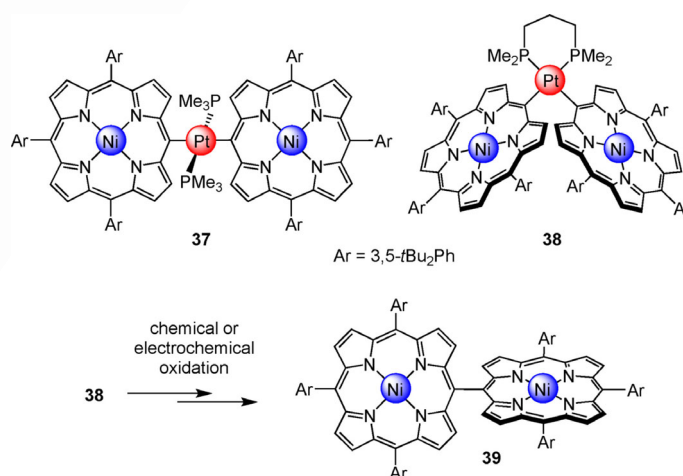
building blocks for the synthesis of a cofacial dimer **34** with a C_{meso} - Pt^{II} - C_{meso} linkage in a *cis* geometry (Scheme 19).^[53] The two porphyrins strongly interact through the *cis* Pt^{II} , as shown by electrochemical studies. A splitting value of 230 mV be-



Scheme 19. Structures of dimeric species **34–36**.

tween the first two one-electron oxidation waves was observed. Interestingly, this dimer shows helical chirality, which may be useful for asymmetric catalysis. In contrast to **34**, the electronic communication is quite weak in coplanar C_{β} -to- C_{β} Pd^{II} and Pt^{II} porphyrin dimers **35** and **36** with a *trans* geometry (Scheme 19).^[54] A *meso* diphenylphosphanyl group was used to construct these coplanar dimers with C_{β} - M - C_{β} linkages ($M = Pd^{II}$ or Pt^{II}). Electrochemical studies revealed small splitting values, ranging from 40 mV (Pd^{II} in **35**) to 60 mV (Pt^{II} in **36**), between the first two one-electron oxidation waves. A similar splitting value of 60 mV was reported for a series of butadiyne- and Pt^{II} diacetylide-bridged conjugated porphyrin dimers.^[55]

In C_{meso} -to- C_{meso} Pt^{II} -bridged porphyrin dimers **37** and **38** with *trans* and *cis* geometry, respectively (Scheme 20),^[56] strong electronic communication between the two porphyrins was demonstrated by UV-visible spectroscopy and electrochemical



Scheme 20. Structures of *trans*-dimer **37** and *cis*-dimer **38**. Structure of the *meso-meso* linked dimer **39** obtained after chemical or electrochemical oxidation of the *cis*-dimer **38**.

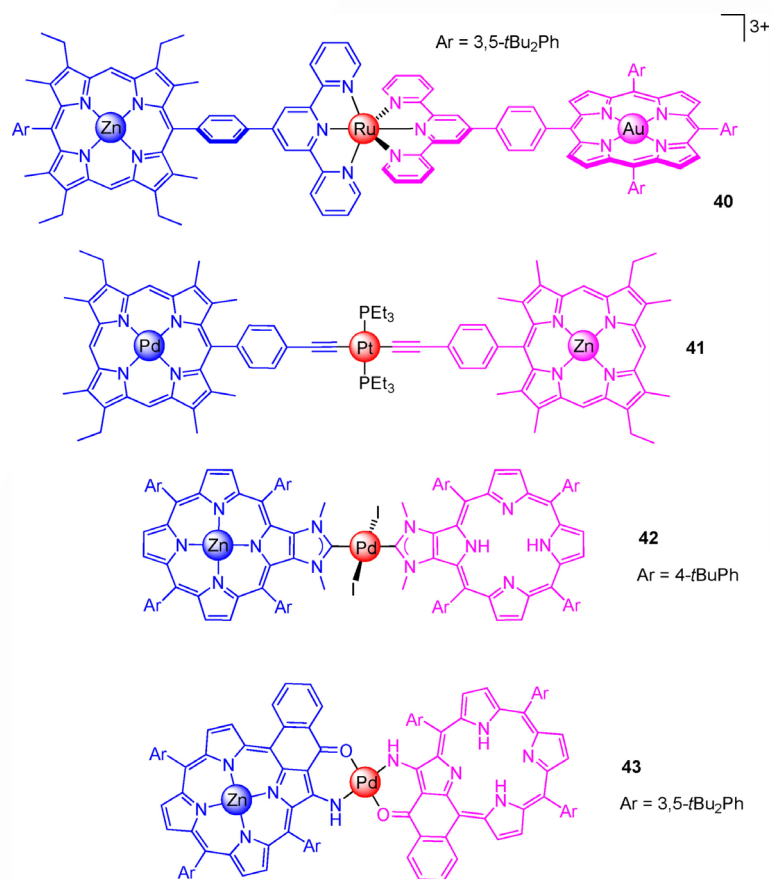
studies. For the *trans*-dimer **37**, the splitting value of the first oxidation wave was 170 mV, which is in the range of values found for dimers **27–M**. DFT calculations showed a clear contribution of the Pt^{II} *d* orbitals in the *trans* geometry to the frontier molecular orbitals. A through-bond conjugation between the porphyrins was postulated. For the *cis*-dimer **38**, DFT calculations showed through-space interactions between the porphyrins in the frontier molecular orbitals, notably between the two carbon atoms linked to Pt^{II}. Interestingly, the electrochemical behavior of the *cis*-dimer **38** was very different because a chemical transformation occurred after the first oxidation, leading to the formation of the *meso-meso* linked dimer **39** through a surprising reductive elimination. This transformation could also be achieved by a chemical oxidation.

3.2. Electronic interactions between the porphyrins in the excited state

Natural porphyrin analogues such as chlorophylls and bacteriochlorophylls ensure energy and electron transfer processes in photosynthetic systems with high efficiencies and rates (as fast as 100–150 femtoseconds for some energy transfers).^[57] Several donor-acceptor (D-A) systems with different geometries, orientations and distances between the chromophores were studied in order to better understand energy and/or electron transfer processes. However, transition-metal complexes as bridging units in D-A assemblies are less studied than covalently linked systems. Pioneering works by several authors including J.-P. Sauvage and co-workers studied electron transfers in bis-porphyrinic D-A systems like **40**, where the individual units were connected by metal ions (Scheme 21).^[58] Later, other systems were built to study the faster energy transfer processes (see dyads **41–43** in Scheme 21).^[43,59,60] In this section are summarized some important results concerning S₁ energy transfer rates within molecular D-A systems, where D and A are assembled with metal complexes.

There are two possible processes of S₁ energy transfer between D and A within dyads, namely Förster and Dexter mechanisms (Figure 3). The Förster resonance energy transfer (FRET) mechanism is a “through space” coulombic energy transfer that necessitates a close proximity (≈1–10 nm) between D and A (Figure 3a). The Dexter mechanism is a “through bond” exchange energy transfer occurring only over short distances (<1 nm) between D and A. As this mechanism corresponds to a double exchange of electrons between D and A, molecular orbital overlap between them is essential (Figure 3b). As shown in the next paragraphs, the S₁ energy transfer rates between D and A within molecular D-A systems vary dramatically by modifying the nature of the peripheral ligands and the bridging transition metal ions.

Photophysical studies of the dyad **41**, where a Pd^{II} porphyrin (D) and a Zn^{II} porphyrin (A) are linked together by an organo-



Scheme 21. Structures of dyads **40–43** with donors in blue and acceptors in pink.

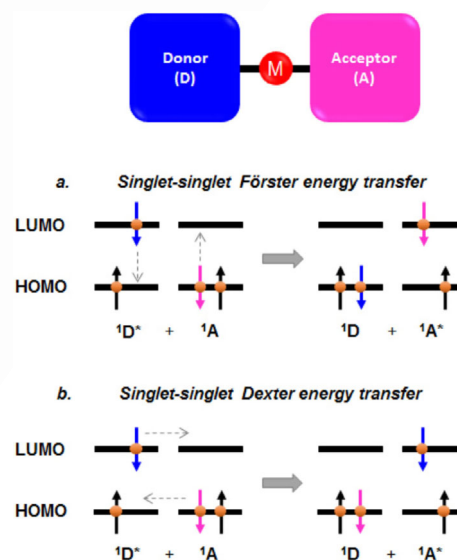


Figure 3. Schematic diagrams of Förster (a) and Dexter (b) energy transfer processes between a donor (D) and an acceptor (A).

metallic spacer (Scheme 21),^[59] showed that the energy transfer rates were quite slow (S₁ energy transfer rate of ≈500 ps). Energy transfer rates were also slow (S₁ energy transfer of ≈90–330 ps) in dyad **42** containing a Zn^{II} porphyrin (D) and a free-base porphyrin (A) fused to peripheral NHC ligands and

linked together by a *trans* Pd^{II}₂ complex.^[43] DFT calculations for **42** clearly showed the lack of Pd^{II} atomic contributions to the π -system of the frontier molecular orbitals (Figure 4, left side). Therefore, the Dexter mechanism cannot operate and S₁

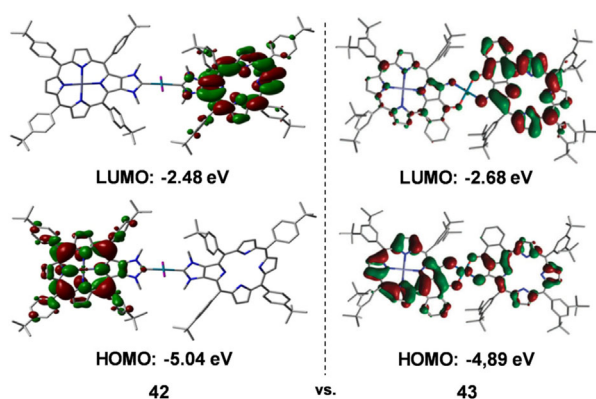
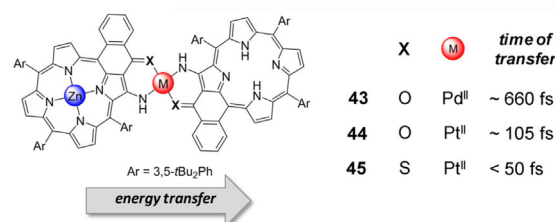


Figure 4. Representation of the frontier molecular orbitals HOMO and LUMO for dyads **42** and **43**.

energy transfer is only due to Förster resonance energy transfer (FRET). However, the S₁ energy transfer rate is much slower than the predicted values when considering only FRET. This slow transfer may be explained by an electric field screening of the donor and the acceptor transition moments by the electron-rich *trans* Pd^{II}₂ complex, which slows down the energy transfer rate.

Energy transfer rates in the two previous dyads **41** and **42** suggest that the presence of a transition metal in the bridge between the porphyrins has a negative impact on the energy transfer rates between the porphyrins, but this is not always the case. For example, a Zn^{II} porphyrin (D) and a free-base porphyrin (A), each bearing an enaminoketone, were assembled around a square planar Pd^{II} to give dyad **43** that exhibited a very fast S₁ energy transfer that takes place in ≈ 660 fs.^[60] This value is comparable to the one found for covalently *meso-meso* linked bis-porphyrins (≈ 550 fs).^[61] This fast rate is due to a strong electronic communication between the porphyrin through the assembling Pd^{II}. DFT calculations show a significant atomic contribution of the *d* orbital to the frontier molecular orbitals (MO) (4.6% for the HOMO) allowing orbital overlap between the two porphyrins. Consequently, the Dexter mechanism operates efficiently (Figure 4, right side). Moreover, the absence of additional ligands on the square planar Pd^{II} is probably beneficial since there is no electric field screening effect of ancillary ligands. Replacing Pd^{II} by Pt^{II} significantly decreases the time needed for the S₁ energy transfer, which is ≈ 105 fs for the dyad **44** because Pt^{II} allows stronger molecular orbitals coupling between the two porphyrins (Scheme 22).^[62] DFT calculations showed higher atomic contribution of the Pt^{II} *d* orbital to the frontier MO compared to Pd^{II} (11.2% for the HOMO). The fastest S₁ energy transfer (< 50 fs) was reached in dyad **45** where oxygen atoms were replaced by sulfur atoms in the peripheral chelating sites.^[50] Again, DFT calculations showed higher atomic contribution of the Pt^{II} *d* orbital to the



Scheme 22. S₁ energy transfer in dyads **43–45**.

frontier MO (13.5% for the HOMO) leading to increased molecular orbitals coupling between the porphyrins and better electronic communication.

In conclusion, the use of transition metal ions as linkers between a donor and an acceptor is feasible and dyads with ultrafast energy transfer rates can be obtained by judiciously choosing the ligands and the metal ion. One disadvantage of these systems compared to covalently linked dyads is that the heavy atom effect of the bridging metal favors the formation of the T₁ state by depleting the S₁ state (intersystem crossing, ISC). Nevertheless, this ISC is of great importance when considering other applications, such as photodynamic therapy, discussed in the next section.

4. Photodynamic Therapy Applications

In the field of biomedical applications, porphyrin derivatives are mainly used as photosensitizers (PS) for photodynamic therapy (PDT) applications. For effective PDT, light and molecular oxygen (³O₂) are the two essential elements in addition to a PS. Ideally, the PS is first accumulated in cancer cells and then irradiated with light. As shown in the simplified Jablonski diagram in Figure 5, the PS absorbs light at a suitable wavelength (ideally in the photo-therapeutic window 650–900 nm for better depth penetration through tissues) to promote the S₀ → S₁ excited state. Then, the singlet excited state S₁ is converted to the lower energy excited triplet state T₁ through ISC. The PS in its T₁ state can transfer its energy to ³O₂ to generate reactive singlet oxygen (¹O₂), which is a highly oxidizing and cytotoxic agent. PDT offers spatial and temporal control because the generated ¹O₂ has lifetimes of ≈ 3 μ s and limited diffusion coefficients of $\approx 2\text{--}4 \times 10^{-6}$ cm² s⁻¹ in biological media.^[63,64] In this section, representative examples of porphyrin-based PS periph-

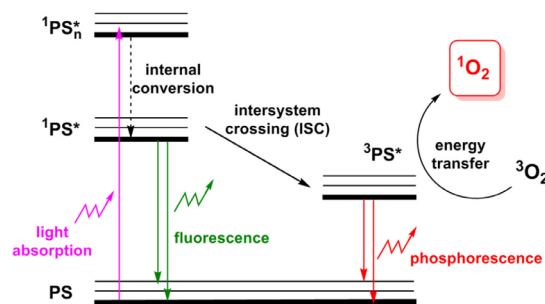
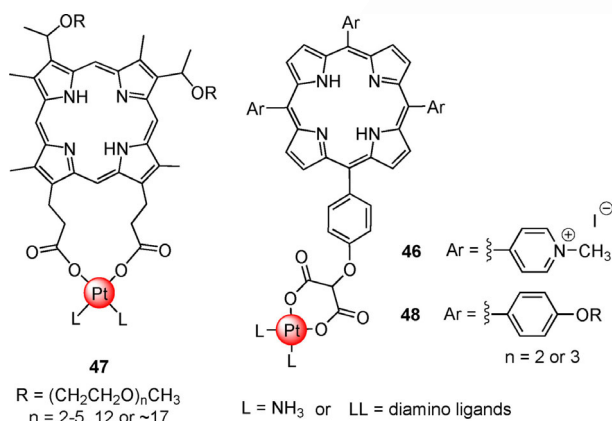


Figure 5. Simplified Jablonski diagram explaining the photophysical processes leading to the formation of ¹O₂.

erally metalated with transition metal ions are presented. The benefits obtained from these metal complexes for therapy and imaging applications will be emphasized.

4.1. Porphyrins with peripheral platinum complexes

Platinum complexes are powerful anticancer agents and one identified mode of action is their ability to crosslink purine bases of DNA. These stable complexes interfere with DNA repair machinery and ultimately lead to death of cancer cells. Cisplatin and carboplatin are active towards bladder, ovarian, head, neck, testicular and lung cancers. However, many side effects are associated with the use of these complexes. To improve their efficiency and limit the side effects, Pt complexes are combined with molecules known to preferentially accumulate in tumor tissues. Song and co-workers synthesized free-base porphyrins **46** with three *meso* 4-methylpyridinium groups that ensure water solubility (Scheme 23).^[65] The fourth *meso* position was functionalized with an aryl group bearing Pt^{II} complexes. The anticancer properties of such Pt^{II} porphyrin conjugates are superior to those of cisplatin due to a higher accumulation in tumors because of the porphyrin.

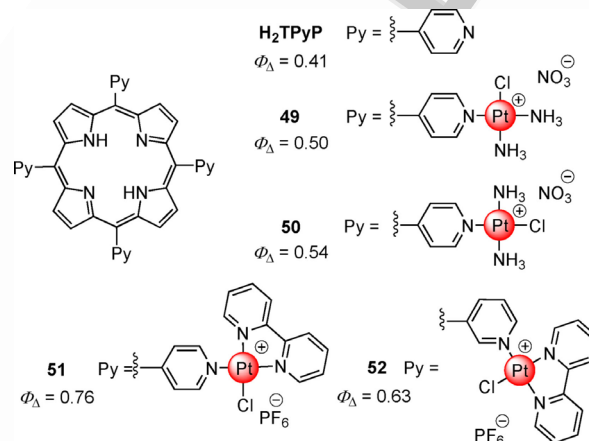


Scheme 23. Structures of Pt^{II}-porphyrin conjugates **46–48**.

Porphyrin derivatives with peripheral Pt^{II} complexes can be used for dual chemo- and phototherapy because they combine within the same molecular species the antitumor activity of Pt^{II} complexes and the phototoxicity of porphyrins upon irradiation with light. Brunner and co-workers investigated in detail the use of hematoporphyrin derivatives **47** as PS for PDT with *cis* Pt^{II} complexes coordinated to the two carboxylate groups of the propionate side chains (Scheme 23).^[66] The most active compounds are those with two amine ligands bound to Pt^{II}, and with polyethyleneglycol units (PEG) for solubility in water. The same group synthesized free-base porphyrins with three *meso* aryl groups bearing PEG chains, while the fourth *meso* aryl group was functionalized with Pt^{II} complexes.^[67] A representative example is the Pt^{II} porphyrin conjugate **48** in Scheme 23.^[67a] As for Pt^{II} hematoporphyrin conjugates, the free-base porphyrin improved the anticancer activity of the Pt^{II} moiety by an additional photodynamic effect upon irradiation

with red light. The light-induced toxicity is improved compared to their hematoporphyrin analogues when irradiated under the same conditions ($\lambda = 600\text{--}730\text{ nm}$; 10 min; 24 J cm^{-2}).

Tetraplatinated porphyrins **49** and **50** (Scheme 24), obtained by reacting the free-base *meso*-tetra(4-pyridyl)porphyrin with Pt^{II} complexes, were used for photodynamic therapy applica-



Scheme 24. Structures of H₂TPyP and Pt^{II}-porphyrin conjugates **49–52**. ¹O₂ quantum yields Φ_{Δ} indicated determined in DMF for H₂TPyP, **49** and **50** and in CH₃CN for **51** and **52**.

tions.^[68] The ¹O₂ quantum yields (Φ_{Δ}) of **49** (*cis* geometry) and **50** (*trans* geometry) are higher than that of the corresponding non-platinated porphyrin, suggesting that the Pt^{II} ions promote the formation of the T₁ state through a heavy atom effect. In vitro studies were performed on different cell lines. Interestingly, Pt^{II}-porphyrin conjugates **49** and **50** show very limited toxicity on HeLa cells in the dark compared to cisplatin, but the toxicity upon irradiation (incubation time: 4 hours, irradiation conditions: $\lambda = 420\text{ nm}$, 15 mn, 6.95 J cm^{-2}) is 600–1200 times higher. The factor corresponding to the phototoxic index (PI = IC₅₀ in the dark/ IC₅₀ irradiated) for PS **50** is >5000 in cisplatin-resistant cell line CP70 suggesting that dual chemo- and phototherapy is a very promising approach for cancer therapy. Subcellular colocalization studies showed improved accumulation of these PS in the nucleus (30 times more than cisplatin). This is a great advantage because DNA is the main identified biological target for Pt^{II} complexes. Viscometry experiments, emission and absorption spectroscopy, circular dichroism studies and NMR spectroscopy showed that PS **50** can interact with DNA; although no damage of DNA was observed in the dark, significant cleavage and damage of the DNA occurred upon light irradiation.

The use of *meso*-3-pyridyl ligands instead of *meso*-4-pyridyl ligands has a significant impact on the PDT efficiency (Scheme 24).^[69] Both tetraplatinated porphyrins **51** and **52** have high singlet oxygen quantum yields, and are phototoxic at the nanomolar scale. However, PS **52** (3-pyridyl) is twice as phototoxic as **51** (4-pyridyl) despite a lower Φ_{Δ} . The overall better efficiency of PS **52** was attributed to its higher intracellular concentration compared to PS **51**. The different charge distribution in **51** and **52** confers more amphiphilicity to **52**

and thus better membrane affinity or permeability and cellular uptake. Therefore, the PS with the highest Φ_{Δ} is not always the best PS for PDT. The number, the nature and the localization of the peripheral metal complexes are important factors to consider when designing a PS for dual chemo- and photo-therapy. The observed benefits of these Pt^{II} complexes are due to synergistic effects (the overall effect is greater than the sum of individual effects): the porphyrin macrocycle improves the cellular uptake of Pt^{II} complexes possibly via low density lipoproteins (LDL) receptors over-expressed by cancer cells, whereas the Pt^{II} complexes enhanced the water solubility and the photodynamic properties of the porphyrin through a heavy atom effect.

Interaction of Pt^{II} porphyrin conjugates with DNA is property necessary for dual chemo- and photo-therapy. Particularly interesting DNA sequences include G-rich DNA sequences of the telomeres at the end of chromosomes. These sequences are able to form G-quadruplexes (G4, see Figure 6) by π - π stack-

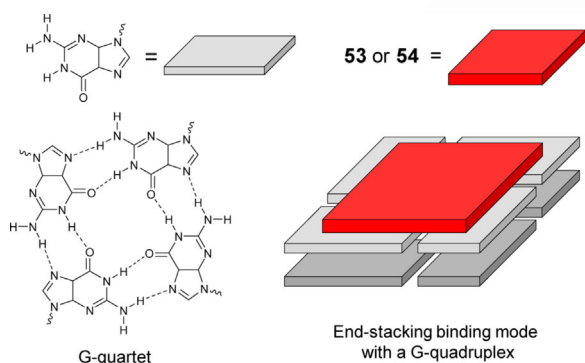
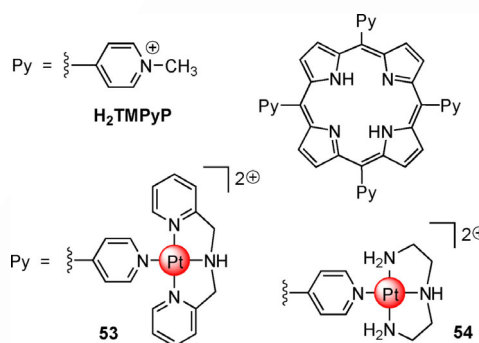


Figure 6. Structures of a G-quartet and a G-quadruplex. Illustration of the end-stacking binding mode of **53** or **54** with a G-quadruplex.

ing of G-quartets, which are formed by four guanines assembled through complementary hydrogen bonds. These G4 structures seem to play an important role in cell aging processes. Flat cationic molecules such as H_2TMPyP (Scheme 24) can stabilize G4 structures and display anticancer properties by inhibiting the activity of telomerase, an enzyme responsible for maintaining the integrity of the telomeres. Mao et al. showed that the methyl group of the pyridinium units of H_2TMPyP could be replaced by cationic Pt^{II} complexes in clover-like compounds **53** and **54** (Scheme 25).^[70] Both compounds display high affinity for G4 structures and, more importantly, a higher selectivity for G4 rather than for double stranded DNA with an end-stacking binding mode (Figure 6). They display anticancer activities on HeLa cells by inhibiting the telomerase activity and also by repressing oncogene expression. The photodynamic properties of these compounds have not yet been investigated.

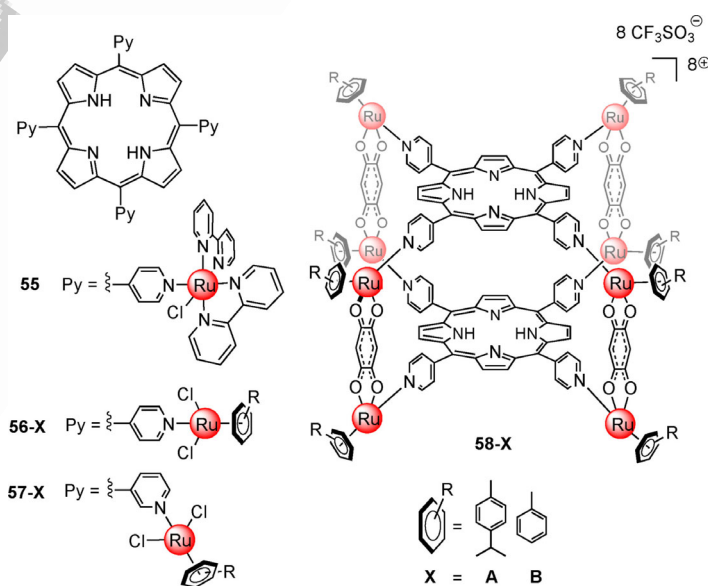


Scheme 25. Structures of H_2TMPyP and Pt^{II} -porphyrin conjugates **53** and **54**.

4.2. Porphyrins with peripheral ruthenium complexes

Ruthenium-based compounds appear to be an attractive class of metallodrugs able to interact with DNA and proteins because they are usually less toxic than their Pt^{II} counterparts. Moreover, ruthenium polypyridyl complexes can generate reactive oxygen species upon irradiation, leading to DNA damage. As with Pt^{II} complexes, the combination of ruthenium complexes with porphyrins was investigated for dual chemo- and photo-dynamic therapy. In an early example, four Ru^{II} complexes were anchored to a free-base *meso*-tetrakis(4-pyridyl)porphyrin (**55**: Scheme 26), which was able to photoinduce oxidative cleavage and modification of DNA.^[71]

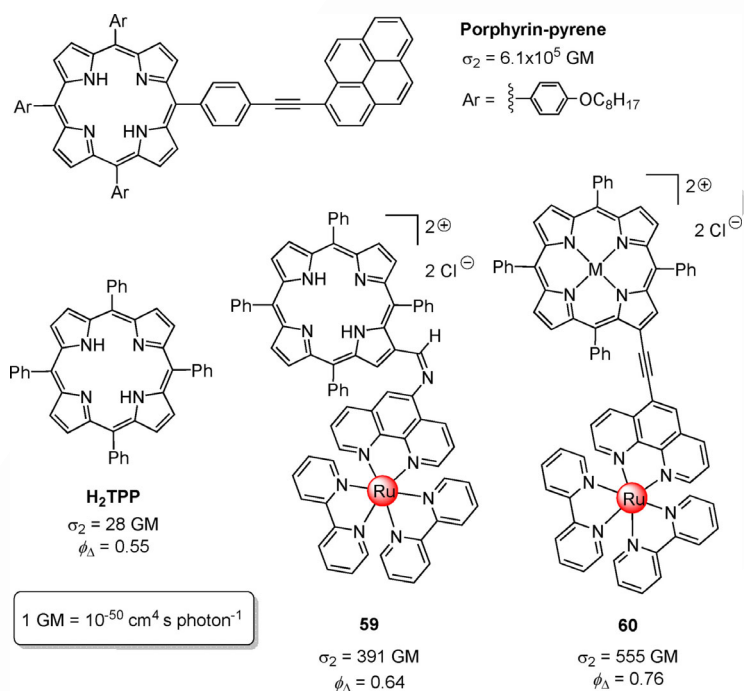
The anticancer and photodynamic properties of porphyrins **56-X** and **57-X** functionalized with arene Ru^{II} complexes were investigated as potential photosensitizing chemotherapeutic agents (Scheme 26).^[72] These complexes accumulated in Me300 melanoma cells cytoplasm and were moderately toxic in the dark ($IC_{50} \approx 10 \mu M$ for **56-A** and $20 \mu M$ for **56-B**, **57-A** and **57-B**, after 72 h incubation). Complexes **56-A** and **57-A** were proven to shut down DNA synthesis by inhibiting incorporation of thymidine. Therefore, protein synthesis and cell



Scheme 26. Structures of Ru^{II} -porphyrin conjugates **55**–**58**.

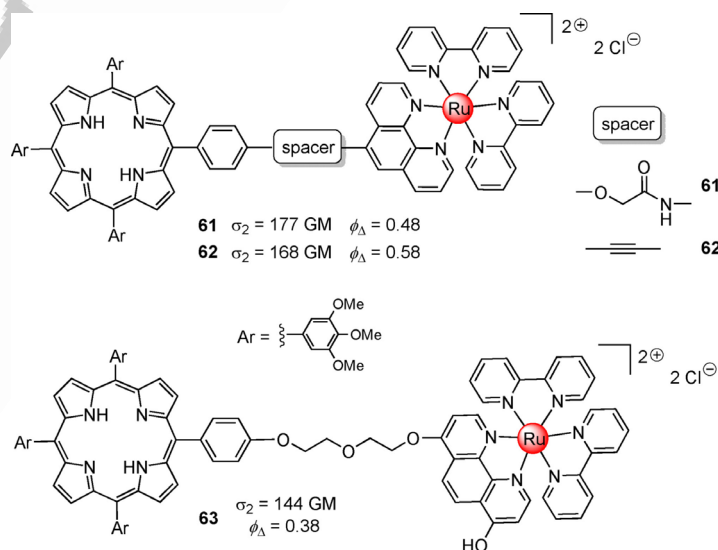
survival were diminished. Upon irradiation with red light ($\lambda = 652$ nm), excellent phototoxicities were observed. Photodynamic studies revealed that Ru^{II} complexes with 3-pyridyl ligands were more efficient than those with 4-pyridyl ligands. Interestingly, the same observation was made with Pt^{II} complexes **51** and **52**, suggesting that attaching the metal complexes in the *meta* positions of the *meso* aryl groups was the best choice. The potential of cationic octa-nuclear Ru^{II}-porphyrin metalla-cubes **58-X** for chemo- and phototherapy was also studied.^[73] Both compounds displayed excellent cytotoxicity towards melanoma cells Me300 (IC₅₀ \approx 5 and \approx 9 μ M for **58-A** and **58-B**, respectively) and were more toxic than the corresponding tetra-Ru^{II} complexes **56-A** and **56-B**. The metalla-cubes **58-X** were not toxic towards pulmonary (A549) and cervical (HeLa) cancer cells (IC₅₀ > 70 μ M), but exhibited good phototoxicity upon irradiation at 652 nm (\approx 75% cell death at 1 μ M concentration and a light dose of 10 J cm⁻²).

A current focus point in the field of PDT is the use of two-photon absorption (TPA) to generate ¹O₂. This technique consists in the simultaneous interaction of a PS with two photons of low energy. TPA can be performed with lasers at wavelengths higher than 800 nm, which allow deeper penetration of the light through tissues. Moreover, spatial resolution of the treatment is excellent because TPA occurs at the diffraction-limited beam focus, resulting in minimal side damage along the beam path. Monomeric porphyrins have relatively low TPA cross-sections ($\sigma_2 < 30$ GM) and are not suitable PS for TPA. To increase the cross-section value of porphyrin-based sensitizers, their chemical structure can be modified by: (i) the introduction of electron-donating (D) or electron-accepting (A) groups, (ii) the extension of their π -system, or (iii) the synthesis of symmetrical (D- π -D, A- π -A) or asymmetrical (D- π -A) systems. Following this strategy, compounds with very high TPA cross-sections could be obtained ($\sigma_2 \approx 6.1 \times 10^5$ GM for **Porphyrin-pyrene** in Scheme 27).^[74] Other candidates for TPA-PDT include the amphiphilic Ru^{II}-porphyrin conjugates **59** and **60** with a hydrophilic cationic head (Ru^{II} complex) and a hydrophobic tail (porphyrin) for enhanced cellular uptake (Scheme 27).^[70] Ru^{II} and porphyrin cores were linked together via conjugated spacers such as an acetylide^[75a] or a Schiff base^[75b] at a β -pyrrolic position. TPA cross-sections of these Ru^{II} conjugates and their ¹O₂ quantum yields (ϕ_Δ) were significantly increased compared to those of **H₂TPP** (Scheme 27). These compounds were effective sensitizers for PDT under mono-photonic excitation, but were also effective under bi-photonic excitation at $\lambda_{exc} = 850$ and 800 nm for **59** and **60**, respectively. Therefore, up to 70% of HeLa cells were killed upon laser excitation for 30 minutes (conditions of irradiation for **59**: 5 μ M, 4 hours incubation time, $\lambda_{exc} = 850$ nm; for **60**: 1 μ M, 3 hours incubation time, $\lambda_{exc} = 800$ nm).



Scheme 27. Structures, two-photon absorption cross-sections σ_2 and ¹O₂ quantum yields ϕ_Δ of **H₂TPP**, **Porphyrin-pyrene** and Ru^{II} complexes **59** and **60** in CHCl₃.

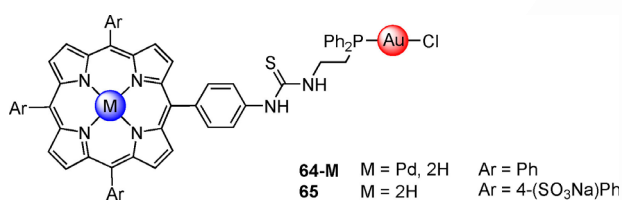
Other Ru^{II} porphyrin conjugates were reported in the literature where the Ru^{II} complex is connected to *meso* aryl^[76] or pyridyl^[77] groups of the porphyrin. TPA cross-section values are lower due to the perpendicular orientation of the *meso* aryl group with respect to the porphyrin plane, preventing efficient π -conjugation between the porphyrin and the Ru^{II} complex (see Ru^{II} porphyrin conjugates **61–63** represented in Scheme 28).^[78] Interestingly, the nature of the spacers used to link Ru^{II} complexes and porphyrins together have a strong influence on the subcellular localisation of the Ru^{II}-porphyrin conjugates and hence, on the PDT efficiency.^[78]



Scheme 28. Structures, two-photon absorption cross-sections σ_2 and ¹O₂ quantum yields ϕ_Δ determined in CHCl₃ of Ru^{II}-porphyrin conjugates **61–63**.

4.3. Porphyrins with peripheral gold complexes

Gold complexes are an interesting class of metallodrugs with a great potential for cancer therapy with modes of action different from those of platinum- or ruthenium-based drugs. Modes of action of gold metallodrugs include mitochondrial damages, inhibition of thioredoxine reductase, proteasome inhibition, and modulation of specific kinase.^[79] Although the potential of Au^{III} porphyrin complexes (with Au^{III} in the macrocycle) as anti-cancer agents is now quite well established (they are able to induce apoptosis and cell cycle arrest),^[80] the use of porphyrins peripherally functionalized with gold complexes is in its infancy and few examples are reported. Among these, neutral complexes **64-M** porphyrins with peripheral Au^I-phosphine complexes (Scheme 29)^[81] proved to be non-toxic (IC₅₀ > 200 μM)



Scheme 29. Structure of Au^I porphyrin conjugates **64-M** and **65**.

towards HCT116 and SW480 colon cancer cells due to their poor water solubility. The water-soluble complex **65** is more toxic, with IC₅₀ ≈ 56 μM on HCT116 and ≈ 72 μM on SW480, and the observed cytotoxicity is due to the peripheral Au^I complex. The water-soluble complex **65** was also an effective PS for PDT because its cytotoxicity increased up to four times upon illumination for 30 minutes with white light.

Although NHCs are best known as ligands for applications in catalysis, the stability of the NHC-metal bond is an interesting property for biomedical applications to prevent the premature release of the metal ions and to improve the bioavailability of the metallodrugs.^[82] Free-base porphyrins fused to NHC-Au^I-X complexes **66–68** (X = Cl or thiolates, see Figure 7) were investigated for their photodynamic properties on MCF7 breast cancer cells.^[83] Compared to the corresponding imidazolium salt **8-2H**, the Au^I porphyrin complex **66** is a better photosensitizer, because the efficiency of **7-2H** to generate ¹O₂ is about 60% that of Au^I complex **66**. This difference can be attributed to the heavy atom effect exerted by the peripheral Au^I complex, as suggested by the lower fluorescence quantum yield and lifetime of the imidazolium salt **8-2H** compared to the Au^I porphyrin complex **66**. The synthesis of thiolato-Au^I complexes **67** and **68** was achieved by substituting the chloride of Au^I complex **66** for thiolates. The Au^I complex **68** is particularly interesting for targeted PDT because it contains a mannose group susceptible of targeting MCF7 breast cancer cells, which are known to overexpress mannose receptors at their surface. The mannose group also dramatically improves the solubility of compound **68** in water. As illustrated in Figure 7 (right side), all compounds presented limited cytotoxicity in the dark (less than 20% cell death at a concentration of 10 μM). Upon irradi-

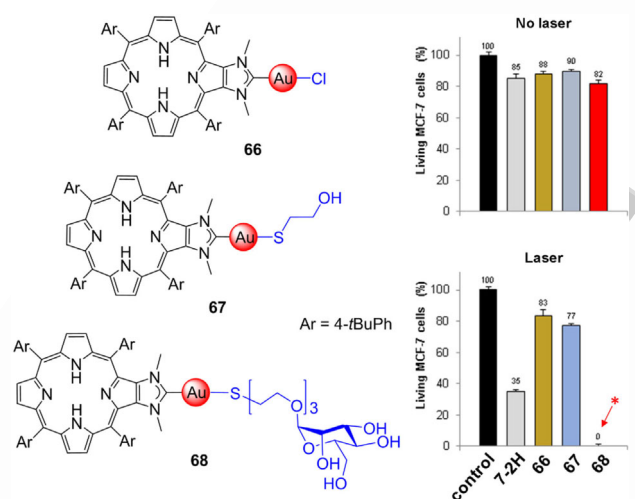


Figure 7. Left: structure of Au^I porphyrin complexes **66–68**; Right: Cytotoxic (no laser) and photodynamic effect (laser) of imidazolium salt **7-2H** and Au^I complexes **66–68**. The cells were incubated or not (control experiment in black) with 10 μM of PS for 4t and then submitted or not to laser irradiation (λ = 405 nm, 18.75 J cm⁻², 10 min). Cells were allowed to grow for 2 days and cell viability was quantified with MTT assay. Data are mean values standard deviation from three independent experiments. *No living cells detected.

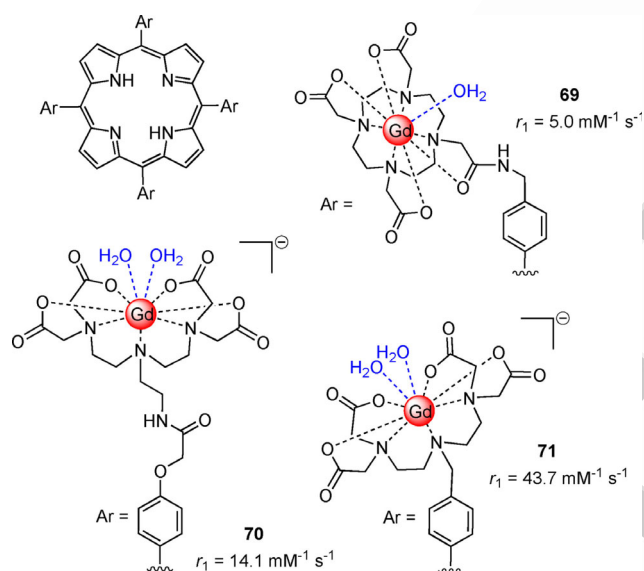
ation, imidazolium salt **7-2H** induced 65% cell death due to the strong interaction of cationic species like **8-2H** with negatively charged cell membranes. Whereas Au^I complexes **66** and **67** showed limited phototoxicity, compound **68** functionalized with mannose became very toxic upon irradiation and no living MCF7 cells were detected. It was demonstrated that the mannose group of **68** promoted active endocytosis in MCF7 cells via mannose receptors. The low photodynamic effects observed with **66** and **67** may be explained by the absence of specific targeting, thus limiting active endocytosis.

4.4. Porphyrins with peripheral gadolinium complexes

Theranostic agents allow simultaneous imaging and therapy with a single molecular system. Therefore simple water soluble porphyrins, such as H₂TMPyP (Scheme 25), and most of the peripherally metalated porphyrins presented in this section can be considered as theranostic agents. They can be detected by fluorescence imaging and, at the same time, used as a PS for PDT. Adding different imaging modalities to theranostic agents is of interest to accurately detect tumors. Magnetic resonance imaging (MRI) is a diagnostic technique offering several advantages, including an excellent spatial resolution without using ionizing radiation. MRI relies on the use of contrast agents (CAs) such as gadolinium(III) complexes.^[84] The CA improves the quality of the images by shortening the relaxation rate of protons of H₂O molecules that interact with Gd^{III}. The longitudinal relaxivity r₁ is directly related to the relaxation rate: the higher the r₁ value, the better the CA will be. In clinical use, Gd^{III} complexes with macrocycles like DOTA (H₄DOTA = 1,4,7,10-tetraazacyclododecane-*N,N,N',N''*-1,4,7,10-tetraacetic acid) or chelating ligands like diethylenetriaminepentaacetate are extensively used, but have relatively low r₁ values (≈ 3–

6 $\text{mm}^{-1}\text{s}^{-1}$ per Gd^{III}) at clinical magnetic fields (1–3 T). Improving the r_1 value with porphyrin macrocycles was first investigated.^[85] However, the inner cavity of porphyrins is too small to accommodate Gd^{III} adequately, and Gd^{III} porphyrins are too unstable for safe in vivo applications. The use of porphyrins bearing peripheral and thermodynamically stable Gd^{III} complexes was a better option to combine within the same framework PDT, fluorescence imaging and MRI. Moreover, synergistic effects were expected. First, Gd^{III} porphyrin conjugates should present higher r_1 values compared to commercially available CAs because of their high molecular weights. Second, the hydrophilic peripheral Gd^{III} complexes should improve the water solubility. Finally, this hydrophilicity combined with the hydrophobic porphyrin confers an amphiphilic character to the conjugates and promotes endocytosis.

Among the Gd^{III} porphyrin conjugates reported in the literature,^[85–90] only those with four Gd^{III} at the porphyrin periphery (like those represented in Scheme 30) at 25 °C and 20 MHz are



Scheme 30. Structures of Gd^{III} porphyrins **69–71** and corresponding relaxivities r_1 ($\text{mM}^{-1}\text{s}^{-1}$) per Gd^{III} measured at 25 °C and 20 MHz.

discussed for better comparison. The tetranuclear Gd^{III} complex **69** is a free-base porphyrin bearing four *meso* DOTA-like Gd^{III} complexes.^[87] The r_1 value of $5.0\text{ mM}^{-1}\text{s}^{-1}$ per Gd^{III} is in the range of commercially available Gd^{III} based CAs. The tetranuclear Gd^{III} complex **70** contains diethylenetriamine-*N,N,N',N'*-tetraacetate (DTTA^{4-}) as peripheral chelating sites for Gd^{III} .^[88] The higher r_1 value of $14.1\text{ mM}^{-1}\text{s}^{-1}$ per Gd^{III} emphasizes the advantage of using such peripheral $[\text{Gd}^{\text{III}}\text{DTTA}]^-$ complexes for better contrast. It is reasonable to assume that the higher r_1 value is due to the presence of two inner-sphere H_2O molecules per Gd^{III} as for non-porphyrinic $[\text{Gd}^{\text{III}}\text{DTTA}]^-$ complexes^[89] (only one H_2O molecule per Gd^{III} in complex **69**). Complex **70** can also be used as a PS for PDT because it is able to generate $^1\text{O}_2$ upon irradiation with red light $\lambda = 650 \pm 20\text{ nm}$. For the free-base porphyrin **71** with four peripheral $[\text{Gd}^{\text{III}}\text{DTTA}]^-$ complexes (identical to those of complex **70**) a

very high r_1 value of $43.7\text{ mM}^{-1}\text{s}^{-1}$ per Gd^{III} was observed^[90] Comparing the chemical structures of complexes **70** and **71** highlights the advantage of using a shorter and more rigid benzyl spacer instead of the longer and more flexible *N*-ethyl-2-phenoxyacetamide. Complex **71** produces $^1\text{O}_2$ in high yield ($\Phi_{\text{A}} = 0.45$ in H_2O) and kills 50% of Hela cancer cells upon irradiation with red light (incubation with $6\text{ }\mu\text{M}$ of **71** for 24 hours, $\lambda = 635\text{ nm}$, light dose = 21 J cm^{-2}). These compounds are all promising theranostic agents for future biomedical applications.

5. Summary and Outlook

Continuous synthetic efforts to synthesize new chromophores have led to innovative porphyrin derivatives containing peripheral coordination sites in addition to their classical inner cavity. Numerous homo- and hetero-multimetallic species have been reported during the last twenty years. Initially, many of these beautiful structures were seen as laboratory curiosities without clearly identified applications. However, combining the optic and electronic properties of porphyrins with the electronic, catalytic or biological properties of metal complexes can lead to compounds with enhanced or new properties.

In the field of catalysis, the peripheral metal complex can play the role of the catalytic site and the resulting catalytic activity may be finely tuned by the porphyrin macrocycle. Axial coordination to inner metal also offers an additional opportunity to develop supramolecular catalysts or to perform dual catalysis (both inner and outer metals are catalytic sites and play complementary roles). Moreover, the porphyrin core can be used as a light-harvesting unit for molecular photocatalytic systems, thus avoiding the use of traditional Ru- or Ir-based photoredox catalysts.

Porphyrins bearing peripheral coordination sites are suitable building blocks for the synthesis of dimers or oligomers of porphyrins linked by metal ions. Depending of the peripheral ligands and the assembling metal ions, strong electronic communication between the porphyrins was observed both in the ground and in the excited states.

Beside catalysis, peripherally metalated porphyrins have biomedical applications because they offer a suitable combination or properties for dual chemo- and photodynamic therapy. For some compounds presented in this review, the overall observed benefit is due to synergistic effects between the porphyrin and the peripheral metal complexes: the porphyrin improves the cellular uptake of the metal complexes, while the latter may improve the water solubility and/or the photodynamic properties. Additional imaging modes such as MRI (Gd^{III}) may also be considered to develop new theranostics.

This summary highlighted the potential of peripherally metalated porphyrins for catalysis, biomedicine and molecular electronics. We presume that other applications, such as molecular sensors, will be considered in the future. One of the next challenges is the synthesis of porphyrins with peripheral non-noble metal complexes, since Pd, Pt and Ru have been mainly used up until now. Such compounds might be interesting for applications in (photo)catalysis or biomedicine.

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Conflict of interest

The authors declare no conflict of interest.

Keywords: antitumour agents · catalysis · molecular electronics · photochemistry · porphyrinoids

- [1] J. Deisenhofer, O. Epp, K. Miki, R. Huber, H. Michel, *Nature* **1985**, *318*, 618–624.
- [2] R. N. Pittman in *Regulation of Tissue Oxygenation, Chapter 4-Oxygen Transport* (Eds.: D. N. Granger, J. Granger), Morgan & Claypool Publishers Life Science series, **2011**, pp. 19–30.
- [3] P. R. Ortiz de Montellano in *Cytochrome P450: Structure, Mechanism, and Biochemistry Chapter 4-Substrate Oxidation by Cytochrome P450 Enzymes* (Eds.: P. R. Ortiz de Montellano), Springer, **2015**, pp. 111–176.
- [4] a) X. Chi, A. J. Guerin, R. A. Haycock, C. A. Hunter, L. D. Sarson, *J. Chem. Soc. Chem. Commun.* **1995**, 2563–2565; b) R. Takahashi, Y. Kobuke, *J. Am. Chem. Soc.* **2003**, *125*, 2372–2373; c) A. Tsuda, T. Nakamura, S. Sakamoto, K. Yamaguchi, A. Osuka, *Angew. Chem. Int. Ed.* **2002**, *41*, 2817–2821; *Angew. Chem.* **2002**, *114*, 2941–2945; d) I.-W. Hwang, T. Kamada, T. K. Ahn, D. M. Ko, T. Nakamura, A. Tsuda, A. Osuka, D. Kim, *J. Am. Chem. Soc.* **2004**, *126*, 16187–16198.
- [5] S. Richeter, C. Jeandon, J.-P. Gisselbrecht, R. Ruppert in *Handbook of Porphyrin Science, Vol. 3* (Eds.: K. M. Kadish, K. M. Smith, R. Guilard), World Scientific, Singapore, **2010**, pp. 429–483.
- [6] a) A. Srinivasan, H. Furuta, *Acc. Chem. Res.* **2005**, *38*, 10–20; b) M. Toganoh, H. Furuta in *Handbook of Porphyrin Science, Vol. 2* (Eds.: K. M. Kadish, K. M. Smith, R. Guilard), World Scientific, Singapore, **2010**, pp. 295–367.
- [7] a) P. Even, B. Boitrel, *Coord. Chem. Rev.* **2006**, *250*, 519–541; b) N. Jux in *Handbook of Porphyrin Science, Vol. 24* (Eds.: K. M. Kadish, K. M. Smith, R. Guilard), World Scientific, Singapore, **2012**, pp. 181–269.
- [8] H. Lu, H. Jiang, Y. Hu, L. Wojtas, X. P. Zhang, *Chem. Sci.* **2011**, *2*, 2361–2366.
- [9] J. L. Maxwell, K. C. Brown, D. W. Bartley, T. Kodadek, *Science* **1992**, *256*, 1544–1547.
- [10] A. Penöni, R. Wanke, S. Tollari, E. Gallo, D. Musella, F. Ragaini, F. Demartin, S. Cenini, *Eur. J. Inorg. Chem.* **2003**, 1452–1460.
- [11] B. J. Anding, A. Ellern, L. K. Woo, *Organometallics* **2012**, *31*, 3628–3635.
- [12] G. Cheng, G. A. Mirafzal, L. K. Woo, *Organometallics* **2003**, *22*, 1468–1474.
- [13] G. S. Ananthnag, V. S. Shetti, *Dalton Trans.* **2017**, *46*, 14062–14082.
- [14] B. M. J. M. Suijkerbuijk, M. Lutz, A. L. Spek, G. van Koten, R. J. M. Klein Gebbink, *Org. Lett.* **2004**, *6*, 3023–3026.
- [15] B. M. J. M. Suijkerbuijk, S. D. Herreras Martínez, G. van Koten, R. J. M. Klein Gebbink, *Organometallics* **2008**, *27*, 534–542.
- [16] S. Yamaguchi, T. Katoh, H. Shinokubo, A. Osuka, *J. Am. Chem. Soc.* **2007**, *129*, 6392–6393.
- [17] J. Yamamoto, T. Shimizu, S. Yamaguchi, N. Aratani, H. Shinokubo, A. Osuka, *J. Porphyrins Phthalocyanines* **2011**, *15*, 534–538.
- [18] K. Yoshida, T. Nakashima, S. Yamaguchi, A. Osuka, H. Shinokubo, *Dalton Trans.* **2011**, *40*, 8773–8775.
- [19] R.-S. Lin, M.-R. Li, Y.-H. Liu, S.-M. Peng, S.-T. Liu, *Inorg. Chim. Acta* **2010**, *363*, 3523–3529.
- [20] Y. Matano, K. Matsumoto, T. Shibano, H. Imahori, *J. Porphyrins Phthalocyanines* **2011**, *15*, 1172–1182.
- [21] J.-F. Lefebvre, D. Leclercq, J.-P. Gisselbrecht, S. Richeter, *Eur. J. Org. Chem.* **2010**, 1912–1920.
- [22] J.-F. Lefebvre, J.-F. Longevial, K. Molvinger, S. Clément, S. Richeter, *C. R. Chimie* **2016**, *19*, 94–102.
- [23] A. V. Astakhov, O. V. Khazipov, A. Y. Chernenko, D. V. Pasyukov, A. S. Kashin, E. G. Gordeev, V. N. Khrustalev, V. M. Chernyshev, V. P. Ananikov, *Organometallics* **2017**, *36*, 1981–1992.
- [24] J.-F. Lefebvre, M. Lo, J.-P. Gisselbrecht, O. Coulembier, S. Clément, S. Richeter, *Chem. Eur. J.* **2013**, *19*, 15652–15660.
- [25] J.-F. Lefebvre, M. Lo, D. Leclercq, S. Richeter, *Chem. Commun.* **2011**, *47*, 2976–2978.
- [26] K. Fujimoto, T. Yoneda, H. Yorimitsu, A. Osuka, *Angew. Chem. Int. Ed.* **2014**, *53*, 1127–1130; *Angew. Chem.* **2014**, *126*, 1145–1148.
- [27] B. M. J. M. Suijkerbuijk, D. J. Schamhart, H. Kooijman, A. L. Spek, G. van Koten, R. J. M. Klein Gebbink, *Dalton Trans.* **2010**, *39*, 6198–6216.
- [28] Y. Miyazaki, C. Kahlfuss, A. Ogawa, T. Matsumoto, J. A. Wytko, K. Oohora, T. Hayashi, J. Weiss, *Chem. Eur. J.* **2017**, *23*, 13579–13582.
- [29] V. F. Slagt, M. Röder, P. C. J. Kamer, P. W. N. M. van Leeuwen, J. N. H. Reek, *J. Am. Chem. Soc.* **2004**, *126*, 4056–4057.
- [30] C. K. Prier, D. A. Rankic, D. W. C. MacMillan, *Chem. Rev.* **2013**, *113*, 5322–5363.
- [31] a) K. L. Skubi, T. R. Blum, T. P. Yoon, *Chem. Rev.* **2016**, *116*, 10035–10074; b) B. L. Tóth, O. Tischler, Z. Novák, *Tetrahedron Lett.* **2016**, *57*, 4505–4513; c) X. Lang, J. Zhao, X. Chen, *Chem. Soc. Rev.* **2016**, *45*, 3026–3038.
- [32] K. Rybicka-Jasińska, W. Shan, K. Zawada, K. M. Kadish, D. Gryko, *J. Am. Chem. Soc.* **2016**, *138*, 15451–15458.
- [33] C. Matlachowski, B. Braun, S. Tschierle, M. Schwalbe, *Inorg. Chem.* **2015**, *54*, 10351–10360.
- [34] a) C. D. Windle, M. V. Câmpian, A.-K. Duhme-Klair, E. A. Gibson, R. N. Perutz, J. Schneider, *Chem. Commun.* **2012**, *48*, 8189–8191; b) C. D. Windle, M. W. George, R. N. Perutz, P. A. Summers, X. Zhong Sun, A. C. Whitwood, *Chem. Sci.* **2015**, *6*, 6847–6864.
- [35] B. Gholamkhash, H. Mametsuka, K. Koike, T. Tanabe, M. Furue, O. Ishitani, *Inorg. Chem.* **2005**, *44*, 2326–2336.
- [36] P. Zhang, M. Wang, C. Li, X. Li, J. Dong, L. Sun, *Chem. Commun.* **2010**, *46*, 8806–8808.
- [37] K. Peuntinger, T. Lazarides, D. Dafnomili, G. Charalambidis, G. Landrou, A. Kahnt, R. P. Sabatini, D. W. McCamant, D. T. Gryko, A. G. Coutsolelos, D. M. Guldi, *J. Phys. Chem. C* **2013**, *117*, 1647–1655.
- [38] P. Lang, J. Habermehl, S. I. Troyanov, S. Rau, M. Schwalbe, *Chem. Eur. J.* **2018**, *24*, 3225–3233.
- [39] C. M. Drain, J.-M. Lehn, *J. Chem. Soc. Chem. Commun.* **1994**, 2313–2315.
- [40] C. M. Drain, F. Nifiatis, A. Vasenko, J. D. Batteas, *Angew. Chem. Int. Ed.* **1998**, *37*, 2344–2347; *Angew. Chem.* **1998**, *110*, 2478–2481.
- [41] R. D. Hartnell, D. P. Arnold, *Organometallics* **2004**, *23*, 391–399.
- [42] S. Richeter, A. Hadj-Aïssa, C. Taffin, A. van der Lee, D. Leclercq, *Chem. Commun.* **2007**, 2148–2150.
- [43] M. Abdelhameed, P.-L. Karsenti, A. Langlois, J.-F. Lefebvre, S. Richeter, R. Ruppert, P. D. Harvey, *Chem. Eur. J.* **2014**, *20*, 12988–13001.
- [44] S. Richeter, C. Jeandon, R. Ruppert, H. J. Callot, *Chem. Commun.* **2001**, 91–92.
- [45] S. Richeter, C. Jeandon, J.-P. Gisselbrecht, R. Ruppert, H. J. Callot, *J. Am. Chem. Soc.* **2002**, *124*, 6168–6179.
- [46] H. Higuchi, K. Shimizu, J. Ojima, K.-I. Sugiura, Y. Sakata, *Tetrahedron Lett.* **1995**, *36*, 5359–5362.
- [47] V. S.-Y. Lin, S. G. DiMaggio, M. J. Therien, *Science* **1994**, *264*, 1105–1111.
- [48] A. Tsuda, H. Furuta, A. Osuka, *J. Am. Chem. Soc.* **2001**, *123*, 10304–10321.
- [49] S. Richeter, C. Jeandon, J.-P. Gisselbrecht, R. Graff, R. Ruppert, H. J. Callot, *Inorg. Chem.* **2004**, *43*, 251–263.
- [50] H. Dekkiche, A. Buisson, A. Langlois, P.-L. Karsenti, L. Ruhlmann, P. D. Harvey, R. Ruppert, *Inorg. Chem.* **2016**, *55*, 10329–10336.

- [51] S. Richeter, C. Jeandon, J.-P. Gisselbrecht, R. Ruppert, H. J. Callot, *Inorg. Chem.* **2007**, *46*, 10241–10251.
- [52] a) S. Richeter, C. Jeandon, R. Ruppert, H. J. Callot, *Chem. Commun.* **2002**, 266–267; b) S. Richeter, C. Jeandon, C. Sauber, J.-P. Gisselbrecht, R. Ruppert, H. J. Callot, *J. Porphyrins Phthalocyanines* **2002**, *6*, 423–430; c) M.-A. Carvalho, H. Dekkiche, L. Karmazin, F. Sanchez, B. Vincent, M. Kanessato, Y. Kikkawa, R. Ruppert, *Inorg. Chem.* **2017**, *56*, 15081–15090.
- [53] S. Yamaguchi, T. Katoh, H. Shinokubo, A. Osuka, *J. Am. Chem. Soc.* **2008**, *130*, 14440–14441.
- [54] a) Y. Matano, K. Matsumoto, Y. Nakao, H. Uno, S. Sakaki, H. Imahori, *J. Am. Chem. Soc.* **2008**, *130*, 4588–4589; b) Y. Matano, K. Matsumoto, H. Hayashi, Y. Nakao, T. Kumpulainen, V. Chukharev, N. V. Tkachenko, H. Lemmetyinen, S. Shimizu, N. Kobayashi, D. Sakamaki, A. Ito, K. Tanaka, H. Imahori, *J. Am. Chem. Soc.* **2012**, *134*, 1825–1839.
- [55] Y.-J. Chen, S.-S. Chen, S.-S. Lo, T.-H. Huang, C.-C. Wu, G.-H. Lee, S.-M. Peng, C.-Y. Yeh, *Chem. Commun.* **2006**, 1015–1017.
- [56] N. Fukui, H.-W. Jiang, A. Osuka, *Org. Chem. Front.* **2017**, *4*, 767–772.
- [57] J. R. Durrant, G. Hastings, D. M. Joseph, J. Barber, G. Porter, D. R. Klug, *Proc. Natl. Acad. Sci. USA* **1992**, *89*, 11632–11636.
- [58] A. Harriman, F. Odobel, J.-P. Sauvage, *J. Am. Chem. Soc.* **1994**, *116*, 5481–5482.
- [59] D. Bellows, S. M. Aly, C. P. Gros, M. El Ojaima, J.-M. Barbe, R. Guillard, P. D. Harvey, *Inorg. Chem.* **2009**, *48*, 7613–7629.
- [60] M. Abdelhameed, A. Langlois, P.-L. Karsenti, S. Richeter, R. Ruppert, P. D. Harvey, *Chem. Commun.* **2014**, 50, 14609–14612.
- [61] H. S. Cho, D. H. Jeong, M.-C. Yoon, Y. H. Kim, Y.-R. Kim, D. Kim, S. C. Jeoung, S. K. Kim, N. Aratani, H. Shinmori, A. Osuka, *J. Phys. Chem. A* **2001**, *105*, 4200–4210.
- [62] H. Dekkiche, A. Buisson, A. Langlois, P.-L. Karsenti, L. Ruhlmann, R. Ruppert, P. D. Harvey, *Chem. Eur. J.* **2016**, *22*, 10484–10493.
- [63] S. Hatz, J. D. Lambert, P. R. Ogilby, *Photochem. Photobiol. Sci.* **2007**, *6*, 1106–1116.
- [64] S. Hatz, L. Poulsen, P. R. Ogilby, *Photochem. Photobiol.* **2008**, *84*, 1284–1290.
- [65] R. Song, Y.-S. Kim, C. O. Lee, Y. S. Sohn, *Tetrahedron Lett.* **2003**, *44*, 1537–1540.
- [66] a) H. Brunner, K.-M. Schellerer, B. Treitinger, *Inorg. Chim. Acta* **1997**, *264*, 67–79; b) C. Lottner, K.-C. Bart, G. Bernhardt, H. Brunner, *J. Med. Chem.* **2002**, *45*, 2064–2078; c) H. Brunner, K.-M. Schellerer, *Inorg. Chim. Acta* **2003**, *350*, 39–48; d) C. Lottner, R. Knuechel, G. Bernhardt, H. Brunner, *Cancer Lett.* **2004**, *203*, 171–180; e) C. Lottner, R. Knuechel, G. Bernhardt, H. Brunner, *Cancer Lett.* **2004**, *215*, 167–177.
- [67] a) C. Lottner, K.-C. Bart, G. Bernhardt, H. Brunner, *J. Med. Chem.* **2002**, *45*, 2079–2089; b) H. Brunner, N. Gruber, *Inorg. Chim. Acta* **2004**, *357*, 4423–4451.
- [68] A. Naik, R. Rubbiani, G. Gasser, B. Spingler, *Angew. Chem. Int. Ed.* **2014**, *53*, 6938–6941; *Angew. Chem.* **2014**, *126*, 7058–7061.
- [69] T. T. Tasso, T. M. Tsubone, M. S. Baptista, L. M. Mattiazzi, T. V. Acunha, B. A. Iglesias, *Dalton Trans.* **2017**, *46*, 11037–11045.
- [70] X.-H. Zheng, Q. Cao, Y.-L. Ding, Y.-F. Zhong, G. Mu, P. Z. Qin, L.-N. Ji, Z.-W. Mao, *Dalton Trans.* **2015**, *44*, 50–53.
- [71] J. Onuki, A. V. Ribas, M. H. G. Medeiros, K. Araki, H. E. Toma, L. H. Catalani, P. Di Mascio, *Photochem. Photobiol.* **1996**, *63*, 272–277.
- [72] a) F. Schmitt, P. Govindaswamy, G. Süß-Fink, W. H. Ang, P. J. Dyson, L. Juillerat-Jeanneret, B. Therien, *J. Med. Chem.* **2008**, *51*, 1811–1816; b) F. Schmitt, P. P. Govindaswamy, O. Zava, G. Süß-Fink, L. Juillerat-Jeanneret, B. Therien, *J. Biol. Inorg. Chem.* **2009**, *14*, 101–109.
- [73] F. Schmitt, N. P. E. Barry, L. Juillerat-Jeanneret, B. Therien, *Bioorg. Med. Chem. Lett.* **2012**, *22*, 178–180.
- [74] N. Sheng, D. Liu, J. Wu, B. Gu, Z. Wang, Y. Cui, *Dyes Pigm.* **2015**, *119*, 116–121.
- [75] a) J. X. Zhang, K.-L. Wong, W.-K. Wong, N.-K. Mak, D. W. J. Kwong, H.-L. Tam, *Org. Biomol. Chem.* **2011**, *9*, 6004–6010; b) H. Ke, W. Ma, H. Wang, G. Cheng, H. Yuan, W.-K. Wong, D. W. J. Kwong, H.-L. Tam, K.-W. Cheah, C.-F. Chan, K.-L. Wong, *J. Lumin.* **2014**, *154*, 356–361.
- [76] a) T. Gianferrara, A. Bergamo, I. Bratsos, B. Milani, C. Spagnul, G. Sava, E. Alessio, *J. Med. Chem.* **2010**, *53*, 4678–4690; b) J. Pan, L. Jiang, C.-F. Chan, T.-H. Tsoi, K.-K. Shiu, D. W. J. Kwong, W.-T. Wong, W.-K. Wong, K.-L. Wong, *J. Lumin.* **2017**, *184*, 89–95.
- [77] a) T. Gianferrara, B. Serli, E. Zangrando, E. Iengo, E. Alessio, *New J. Chem.* **2005**, *29*, 895–903; b) P. Sweigert, Z. Xu, Y. Hong, S. Swavey, *Dalton Trans.* **2012**, *41*, 5201–5208.
- [78] J.-X. Zhang, J.-W. Zhou, C.-F. Chan, T. C.-K. Lau, D. W. J. Kwong, H.-L. Tam, N.-K. Mak, K.-L. Wong, W.-K. Wong, *Bioconjugate Chem.* **2012**, *23*, 1623–1638.
- [79] S. Nobili, E. Mini, I. Landini, C. Gabbiani, A. Casini, L. Messori, *Med. Res. Rev.* **2010**, *30*, 550–580.
- [80] a) A. D. Lammer, M. E. Cook, J. L. Sessler, *J. Porphyrins Phthalocyanines* **2015**, *19*, 398–403; b) Y. Wang, Q.-Y. He, R. W.-Y. Sun, C.-M. Che, J.-F. Chiu, *Cancer Res.* **2005**, *65*, 11553–11564.
- [81] S. Tasan, C. Licon, P.-E. Doulain, C. Michelin, C. P. Gros, P. Le Gendre, P. D. Harvey, C. Paul, C. Gaiddon, E. Bodio, *J. Biol. Inorg. Chem.* **2015**, *20*, 143–154.
- [82] M.-L. Teyssot, A.-S. Jarrousse, M. Manin, A. Chevry, S. Roche, F. Norre, C. Beaudoin, L. Morel, D. Boyer, R. Mahiou, A. Gautier, *Dalton Trans.* **2009**, 6894–6902.
- [83] J.-F. Longevial, K. El Cheikh, D. Aggad, A. Lebrun, A. van der Lee, F. Tielens, S. Clément, A. Morère, M. Garcia, M. Gary-Bobo, S. Richeter, *Chem. Eur. J.* **2017**, *23*, 14017–14026.
- [84] a) E. Tóth, A. E. Merbach in *The Chemistry of Contrast Agents in Medical Magnetic Resonance Imaging*, 2nd ed. (Eds.: J. Wiley and Sons), New York, **2013**, Relaxivity of Gadolinium(III) Complexes, pp. 25–82, Chap. 2; b) P. Caravan, *Chem. Soc. Rev.* **2006**, *35*, 512–523; c) E. Terreno, D. D. Castelli, A. Viale, S. Aime, *Chem. Rev.* **2010**, *110*, 3019–3042.
- [85] M. J. F. Calvete, S. M. A. Pinto, M. M. Pereira, C. F. G. C. Geraldès, *Coord. Chem. Rev.* **2017**, *333*, 82–107.
- [86] a) C. P. Gros, A. Eggenspiller, A. Nonat, J.-M. Barbe, F. Denat, *MedChemComm* **2011**, *2*, 119–125; b) J. Schmitt, V. Heitz, A. Sour, F. Bolze, P. Kessler, L. Flamigni, B. Ventura, C. S. Bonnet, E. Tóth, *Chem. Eur. J.* **2016**, *22*, 2775–2786.
- [87] A. Eggenspiller, C. Michelin, N. Desbois, P. Richard, J.-M. Barbe, F. Denat, C. Licon, C. Gaiddon, A. Sayeh, P. Choquet, C. P. Gros, *Eur. J. Org. Chem.* **2013**, 6629–6643.
- [88] J. Luo, L.-F. Chen, P. Hu, Z.-N. Chen, *Inorg. Chem.* **2014**, *53*, 4184–4191.
- [89] A. Sour, S. Jenni, A. Ortí-Suárez, J. Schmitt, V. Heitz, F. Bolze, P. Loureiro de Sousa, C. Po, C. S. Bonnet, A. Pallier, E. Tóth, B. Ventura, *Inorg. Chem.* **2016**, *55*, 4545–4554.
- [90] a) J. B. Livramento, E. Tóth, A. Sour, A. Borel, A. E. Merbach, R. Ruloff, *Angew. Chem. Int. Ed.* **2005**, *44*, 1480–1484; *Angew. Chem.* **2005**, *117*, 1504–1508; b) J. Costa, E. Tóth, L. Helm, A. E. Merbach, *Inorg. Chem.* **2005**, *44*, 4747–4755.

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