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Protonation of Methyluracils in The Gas Phase:

The Particular Case of 3-Methyluracil

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Dedicated to Terrance B. McMahon on the occasion of his 70th birthday.

Abstract

The gas-phase structure of the protonated isomeric methyl uracils (1-Me-, 3-Me- and 6-Me-) was examined using mid-infrared multiple photon dissociation (IRMPD) spectroscopy performed at CLIO, the Orsay (France) Free Electron Laser facility. Experimental infrared spectra were recorded for protonated species generated by electrospray ionization, isolated and irradiated in a quadrupole ion trap, and compared to calculated infrared absorption spectra of the different low-lying isomers computed at the B3LYP/6-31++G(d,p) level of theory.

For both protonated 1-Me-uracil and 6-Me-uracil, the global energy minima correspond to enolic tautomers, whose infrared absorption spectra were found to match very well with the experimental IRMPD spectra. A small fraction of another low energy lying keto tautomer is also present under electrospray conditions. Protonation of 3-Me-uracil by electrospray results exclusively in the formation of a keto form. 3-Me-uracil, which therefore constitutes a particular case in the series of pyrimidine nucleobases studied so far. Methylation of the N3 position of uracil prevents the interconversion between the keto and enol forms, as encountered for uracil and uracil methylated on N1, C5, or C6. These data also give new insights about the unimolecular reactivity of protonated uracils.

Keywords: IRMPD spectroscopy, photodissociation, protonated Methyluracils, mass spectrometry, DFT calculations.

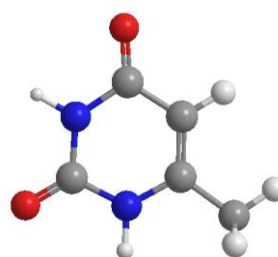
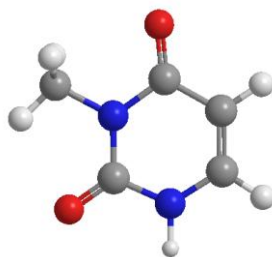
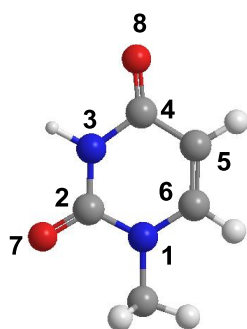
1. Introduction

Within DNA, the strong hydrogen bonds (HBs) between the canonical forms of purine and pyrimidine nucleobases ensure the structural integrity of the biopolymer, and subsequently the encoding and expression of the genetic information. Thymine, for instance, is found in its canonical structure, a dioxo tautomer, in the Watson-Crick base-pair made with adenine. Besides the formation of canonical AT and GC base pairs, other recognition patterns are possible. Through internal proton transfers, which are correlated to the intrinsic acid-base properties of their electronegative centers, different tautomeric forms of the nucleobases can be formed, leading to the establishment of non-canonical hydrogen bonds. The role of tautomerization in genetic expression has therefore been extensively studied. Some of these studies have determined that the production of these tautomers can lead to mispairs and eventual mutations during nucleic acid replication. [1-4]

Because of the probable relationship between the occurrence of these rare tautomeric forms and point mutations, the evaluation of the tautomeric behaviour of nucleobases and the characterization of their acid-base properties, is of fundamental importance. To address these questions, working in the gas phase is particularly appropriate because it allows one to obtain information on their intrinsic properties by eliminating any influence from solvent and conformational averaging effects. Furthermore, such studies allow for a direct comparison between experimental and theoretical data. Professor Terrance B McMahon has followed this strategy all along his extraordinary career dedicated to the study of the reactivity of gaseous ions, and recently provided new insights about the proton transport within protonated nucleobases, notably catalyzed by ammonia. [5]

Over the past decade, mass spectrometry has been increasingly useful in gas-phase studies of nucleobases, thanks to the combination of two technical developments that have occurred during the last thirty years. The first one is the advent of electrospray ionization, [6] which

allows for facile formation of ions in the gas phase from non-volatile compounds, and greatly simplifies the study of nucleic acid building blocks (from nucleobases to double strands) by mass spectrometry. The second one is the development of structure-sensitive activation techniques, which enables one to obtain direct structural information on gaseous ions. Among those, Infrared Multiple Photon Dissociation spectroscopy (IRMPD) of mass-selected ions, is now established as a very powerful approach to probe the structure of gaseous ions of moderate size, [7-10], and different groups have used IRMPD spectroscopy to study the structure and tautomerization of protonated DNA and RNA building blocks generated by electrospray [11-24]. In our continuing effort to study the structure and unimolecular reactivity of gaseous nucleobases, either protonated [11-13,25,26] or complexed to metals [27-32], the present paper reports the IRMPD study in the fingerprint region between 1000 and 2000 cm^{-1} of three protonated methyluracils, namely 1-Me, 3-Me and 6-Me-uracil (Scheme 1), generated by electrospray ionization and trapped in a quadrupole ion trap. The study is complemented by DFT and *ab-initio* electronic structure calculations that provide information regarding the relative stability of the different conformers and their vibrational modes. Several recent theoretical studies have been dedicated to the effect of methylation on tautomeric equilibria of nucleobases.[33-37] Our work shows that 3-Me-uracil constitutes a particular case in the series of pyrimidine nucleobases we have studied so far. Its protonation under electrospray conditions results in the exclusive formation of a protonated keto form, methylation of the N3 position of uracil preventing the interconversion between the keto and enol forms.



<Scheme

1>

2. Experimental and theoretical methods

2.1 Experimental details.

IRMPD spectroscopy experiments have been performed using an experimental platform employing a quadrupole ion trap (Bruker Esquire 3000+) which has been previously described in detail.[38-40] Tunable mid-infrared radiation is produced by the Free Electron Laser (FEL) of CLIO (Centre Laser Infrarouge d'Orsay).[41] This laser is based on a 16-45 MeV linear electron accelerator, whereby bunches of electrons are injected in an undulator, a periodic magnetic field, which is placed in the optical cavity. Using a fixed electron energy, the energy of the emitted photons can be continuously scanned by adjusting the undulator gap. The mean average power was monitored and found to be quite stable ($\sim 1200 \rightarrow 400$ mW) over the $900 \rightarrow 1900$ cm^{-1} energy range scanned using electrons accelerated at 45 MeV.

A conical hole (2.0 mm of diameter) in the ring electrode of the trap was made in order to allow the optical access to the centre of the trap. The IR-FEL beam is mildly focussed using a 1m focal length spherical mirror, and enters the vacuum chamber through a ZnSe window oriented near the Brewster angle. Multistage mass spectrometry was carried out using the standard Bruker Esquire Control (v5.2) software. After mass selection (MS1), mass-selected ions were irradiated with the IR FEL light using the MS2 step, after which the resulting ions

were mass-analyzed. During the MS2 step, the excitation amplitude was set to 0 to avoid any CID-like process. Depending on the nucleobase, an irradiation time ranging from for 300 to 750 ms has been used.

Mass spectra were recorded after 10 accumulations, using the standard mass range (m/z 50-3000) and the normal scan resolution (13000 Th/s), the accumulation time being typically of 10-25 ms, depending on the sample. This sequence was repeated 10 times for each photon energy.

10^{-4} M aqueous solutions of the nucleobases were prepared in a 1/1 water/methanol mixture (v/v), and were introduced in the source using direct infusion with a syringe pump. Typical ESI conditions were a flow rate of 3 $\mu\text{L}/\text{min}$, capillary spray voltage at -4.0 kV, nebulizer at 14 PSI, drying gas flow at 5 L/min, and drying gas temperature at 200°C .

IRMPD spectra are obtained by plotting the photofragmentation yield R ($R = -\ln[I_{\text{parent}}/(I_{\text{parent}} + \sum I_{\text{fragment}})]$), where I_{parent} and I_{fragment} are the integrated intensities of the mass peaks of the precursor and of the fragment ions, respectively) as a function of the frequency of the IR radiation.[42]

2.2 Theoretical calculations.

DFT calculations were carried out using the B3LYP density functional,[43,44] as implemented in the Gaussian 09 set of programs.[45] The different protonated forms considered were first optimized at the B3LYP/6-31++G(d,p) level, without any symmetry constraint. Harmonic vibrational frequencies were calculated at this level to characterize the stationary points as local minima or saddle points, and to estimate the zero-point vibrational energy (ZPE) corrections. The infrared absorption spectra were calculated within the harmonic approximation. It has been previously demonstrated[46,47] that DFT and MP2 methods provide similar geometries for nucleic acid bases. In particular, the hybrid B3LYP

functional has been shown to accurately describe both position and relative intensities of IR bands.[46-49] As far as the positions are concerned, a scaling factor value of 0.96 has been chosen, as in our previous study on protonated uracil and thymine.[11] Finally, to be consistent with the experimental spectral resolution, the calculated spectra have been convoluted by a 15 cm⁻¹ wide large Lorentzian function.

Relative energies for the different conformers were refined by single-point calculations on top of the B3LYP-optimized geometries using a larger 6-311++G(3df,2p) basis set. The reliability of the B3LYP relative energies was assessed by additional single-point M06-2X and MP2(fc)/6-311+G(3df,2p) calculations, and completed by CCSD(T)/6-31++G(d,p) estimates. This set of data is available as a Supporting Information. The non-covalent interactions, which are in part responsible for the relative stabilities of the different tautomers of the species investigated, have been analysed through the use of the NCI approach. This approach takes advantage of the fact that non-covalent interactions are characterized by low-density and low reduced density gradient values, so these weak interaction regions can be located by using gradient isosurfaces, in which a blue-green-red color code one to distinguish between strong attractive non-covalent interactions (blue) and strong repulsive non-covalent interactions (red). Green color denotes very weak interactions (attractive or not) in the van der Waals range. These representations have been created with NCIPLOT (Non-Covalent Interaction Plot). [50,51]

3. Results, discussion

3.1 Overview of the IRMPD experiments

The photo-fragments observed by IRMPD are similar to those obtained by low-energy collision induced dissociation. Each compound exhibits a specific set of photo-fragments. Protonated 1-Methyl-uracil is characterized by elimination of H₂O (*m/z* 109), as well as [CHNO] (*m/z* 84) and [C₂H₃NO] (*m/z* 70). Elimination of ammonia (*m/z* 110) is observed

only when each nitrogen bears a hydrogen atom, namely with 6-Methyluracil. This is consistent with recent chemical dynamics simulations which suggest that this process proceeds mostly through proton transfer between the two nitrogen atoms. [25] For this compound, the other fragments are associated with elimination of water and 43 mass units [CHNO]. The m/z 70 ion is no longer observed, unlike the case of 3-Methyluracil. This strongly suggests that one of the nitrogen atoms has to be methylated to generate this ion. The most intense photo-fragment observed with 3-Methyluracil is detected at m/z 96 and is specific of this molecule. This ion formally corresponds to elimination of CH_3NH_2 . This result is in line with previous studies on protonated uracil[25,26,52], which demonstrated, notably by isotopic labelling, that the loss of ammonia involves almost exclusively the N3 centre.

The IRMPD spectra of protonated 1-Me-Uracil, 3-Me-Uracil and 6-Me-uracil are given in Figure 1. By examining this Figure, one can clearly observe two different behaviours under IRMPD conditions. On the one hand, protonated 1-Me- and 6-Me-uracils exhibit similar spectra, characterized notably by an overwhelming feature around 1600 wavenumbers, a band around 1480 cm^{-1} , and a small absorption near 1800 cm^{-1} . This IRMPD profile has been already observed for protonated uracil and thymine (5-Me-uracil)[11] and suggests that these four protonated species share structural similarities. On the other hand, the IRMPD spectrum obtained for protonated 3-Me-uracil is markedly different, as attested by the very intense absorptions near 1200 and 1800 cm^{-1} , as well as intense features between 1450 and 1600 cm^{-1} . This finding clearly indicates that the structure of protonated 3-Methyluracil differs from those of both protonated uracil and the other Methyluracils.

The assignment of the IRMPD spectrum of the various compounds is based on the comparison with the spectra computed for the various low energy-lying isomers. As will be discussed in detail in the forthcoming sections, there is a quite good agreement between the IRMPD spectrum and the calculated IR absorption spectrum of the lowest energy isomer of

each protonated nucleobase. For both 1-Me- and 6-Me-uracil, this isomer corresponds to an enolic form, with the exception of the relatively *weak* IRMPD signal observed at about 1800 cm^{-1} (Figure 1a and 1c). As shown in our previous studies on uracil and thymine, this particular band is the result of a resonant absorption through the infrared active C=O stretch of a protonated keto form. Considering the strong infrared cross-section associated with the carbonyl group, this weak signal tends to suggest a minor contribution of a keto form for both systems. By contrast, the C=O stretching mode results in a very intense absorption for 3-Methyluracil, pointing to the prominent formation of a keto form (*vide infra*) for this ion.

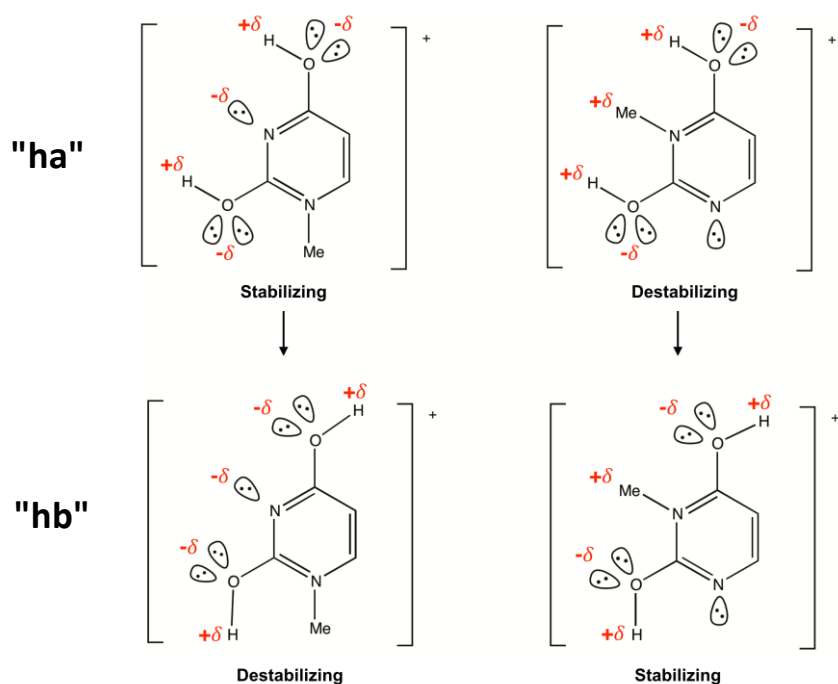
3.2 DFT study of protonated methyluracils

Figure 2 gathers the four most stable protonated forms for 1-Me and 6-Me protonated uracils, whereas for the 3-Me structure the three most stable forms plus a fourth one selected for the sake of comparison are shown. Energetics and structure of all the protonated forms optimized in this study are provided in sections S1-3 of the Supporting Information, along with the labelling for the different isomers.

As can be observed in Figure 2, similar results are obtained for protonated 1-Me- and 6-Me-uracil. The corresponding global minima correspond to dienolic forms, namely **1Me2_ha** and **6MeU2_ha**, respectively. These results are also comparable to those obtained during previous theoretical studies regarding protonated uracil[11,53,54] which, regardless of the method of calculation (*ab initio* or DFT), describe the enolic structure as the most stable one. For both systems, the second most stable structure is the O8-protonated keto form (**1Me1_hd** and **6MeU1_hd**). The third most stable form (the **XMeU1_hc** rotamers) lie 12 kJ/mol above the global minimum. One can notice that the energy difference between the most stable structures is small (about 1 kJ/mol) and reduced with respect to protonated uracil. A detailed study of protonated cytosine by Tureček and co-workers demonstrated that use of diffuse functions

and extended basis sets was essential to accurately evaluate the relative energies between the various protonated forms.[55] All calculations at the different levels of theory (see Section 2.2) confirmed that the **XMeU2_ha** structures are systematically more stable than the **XMeU1_hd** structures. In particular, the CCSD(T)/6-31++G(d,p) calculations provide an energy gap even more pronounced than DFT and MP2, with an estimated value of 8.4 kJ/mol.

DFT calculations point out a different situation for protonated 3-Me-uracil. For this system, the global minimum corresponds to the keto form **3MeU1_hd**. The **3MeU1_hc** rotamer is slightly less stable. More importantly, dienolic forms appear strongly destabilized with respect to 1-Me- and 6-Me-uracil. Whereas **1MeU2_ha** and **6MeU2_ha** forms are the global minima of the PES, the **3MeU2_ha** form, is 90.2 kJ/mol less stable than the global minimum. This large destabilization is due to the replacement of the attractive electrostatic interactions between the positively charged hydroxyl hydrogen atoms and the N lone-pair that are present in the dienolic **1MeU2_ha** and **6MeU2_ha** forms, to repulsive interactions in **3MeU2-ha** between the aforementioned hydroxyl hydrogen atoms and the hydrogens of the methyl group (see Scheme 2, Figure S4 and text therein).



Scheme 2

Note that even the most stable enolic form (**3MeU2b_hb**) characterized by the two hydroxyl hydrogens pointing away from the methyl group is already high in energy (+29.5 kJ/mol). Hence, whereas in the previously described 1-Me and 6-Me forms the stability is governed by the subtle balance between the acidity properties of the different protonation sites (N, O), the presence of the Me group completely modifies this balance.

Scheme 2 also shows why structure **hb** is a high-lying local minimum for 1-Me and 6-Me isomers, because of the repulsive interactions between the oxygens lone-pairs and the N3 lone-pair, whereas the analogous structure for 3-Me-uracil is a low-lying minimum stabilized by the interactions of the oxygen lone-pairs with the positively charged atoms of the methyl group (see Figure S5). The system is further stabilized through a OH...N1 interaction. The acidity of this hydroxyl group at position C2 in structure **hb** is enhanced by the fact that it is acting as a proton acceptor with respect to the neighbouring CH₃ group. Accordingly, the tautomeric enolization from **3MeU1_hd** and **3MeU1_hc** to **3MeU2b_hb** is clearly

disfavoured because the carbonyl group at C2, in the first two tautomers, is already acting as a HB acceptor with respect to one of the H atoms of the methyl group.

Finally, protonation onto C5 (**XMeU1_he**) and C6 centers (**XMeU1_hf**) have also been considered. The resulting structures are also very high in energy. Note that **3MeU1_hf** and **6MeU1_hf** do not correspond to minima on the potential energy surface. In conclusion, both IRMPD experiments and DFT calculations clearly indicate the peculiarity of 3-Methyluracil with respect to both uracil and other methylated (1-Me, 5-Me and 6-Me) derivatives.

3.3 IRMPD spectra of protonated 1-Me- and 6-Me-uracil

The experimental IRMPD spectrum of protonated 1-Methyluracil is given in Figure 3a. This spectrum is dominated by a broad band (fwhm~70 cm^{-1}) centred at 1590 cm^{-1} . This band displays an asymmetric profile with a maximum at 1580 cm^{-1} and a shoulder at 1620 cm^{-1} . This suggests that this signal probably corresponds to the convolution of two strongly active IR modes. Two other intense signals are detected, namely a band at c.a. 1490 cm^{-1} , and a broad feature centred around 1190 wavenumbers. Finally, two relatively weak bands are also detected around 1320 and 1785 cm^{-1} .

All the IRMPD bands (except for the signal detected at 1785 cm^{-1}) can be assigned by considering the IR active modes of the **1MeU2_ha** structure, which are summarized in Table 1. The asymmetric profile of the large signal centred at 1590 cm^{-1} can be ascribed to two intense computed modes at 1566 cm^{-1} and 1631 cm^{-1} , associated to the C4-C5 stretch and the combined stretches of C5-C6 and C2-N3 bonds, respectively. The former is only slightly red-shifted as compared to the strong IR absorption detected at 1580 cm^{-1} , while the latter is in good agreement with the shoulder observed experimentally at the blue side of the band (1620 cm^{-1}). The intense band detected at 1490 cm^{-1} can be attributed to the CO stretch of both C2O7 and C4O8 bonds, while the weak feature detected at 1317 cm^{-1} may be assigned to the

bending modes of the vinylic hydrogens. Finally, two computed modes can account for the broad feature detected around 1200 cm^{-1} , and correspond to various bending modes of the two hydroxyl groups. This experimental signal appears to be characteristic of the presence of enolic tautomers as it cannot be reproduced by the **1Me1_hd** and **1Me1_hc** keto forms (Figures 3c-d).

The computed spectrum for the global minimum **1MeU2_ha** cannot account, however, for the small signal observed experimentally at 1785 cm^{-1} . This signal (position, intensity) has already been observed for protonated uracil and thymine,[11,12] as well as for protonated uridine [21] and thymidine,[23] and results from a resonant absorption through an active C=O stretch. This IR mode is characteristic of the presence of an oxo protonated form, as attested by the agreement observed with the computed C=O stretch of oxo forms such as **1MeU1_hd** (Figure 3c) and **1MeU1_hc** (Figure 3d). Note that the computed spectrum of **1MeU1_hd**, and to a lesser extent that of **1MeU1_hc**, can also account for the experimental bands detected at 1621 , 1590 and 1317 cm^{-1} . Consequently, as for protonated uracil and thymine, the experimental IRMPD spectrum suggests the presence of two tautomeric forms of protonated 1Me-uracil.

Given the multiple photon character of the IRMPD process, the relative IRMPD band intensities should be taken with caution, and rigorous estimates of the relative proportions of each protonated form cannot be made on the sole basis of the experimental spectra. However, the calculated infrared cross section of the **1MeU1_hd** isomer at 1779 cm^{-1} is very large and at least as large as the IR modes around it. As a result, given the fact that the residual laser power at 1800 cm^{-1} was still significant ($\sim 500\text{ mW}$), if the **1MeU1_hd** oxo isomer would be selectively formed, one would expect a stronger IRMPD signal at 1780 cm^{-1} , i.e. of similar order of magnitude as those detected $\sim 1600\text{ cm}^{-1}$, as for 3-Me-uracil. Consequently, we may reasonably conclude that the most abundant tautomer generated under our experimental

conditions corresponds to the enolic structure. A similar conclusion can be drawn for protonated 6-Me-uracil and will not be discussed in detail. As already mentioned (Figure 2), the IRMPD spectrum obtained for protonated 6-Me-uracil and 1-Me-uracil share the same characteristics. Comparison with computed spectra for **6MeU2_ha**, **6Me1_hd** and **6Me1_hc** is provided in the Supporting Information.

These data globally suggest that under electrospray conditions, tautomerization of either the neutral or protonated forms of these molecules occur in order to account for the formation of dienolic tautomers of the protonated species. The results obtained for uracil support this assumption. For neutral uracil, it has been shown that the presence of one or several molecules of water considerably lowers the activation barrier associated with tautomerization.[11,56] The same has been found for protonated uracil, the tautomerization towards enolic forms being greatly facilitated by the presence of protic solvents like ammonia[5] or water.[12] Finally, it is also worth noting that activation of the ions by sequential absorption of IRMPD photons may promote the isomerization of the enolic forms prior to dissociation. Our recent chemical dynamic simulation study of the collision induced dissociation of protonated uracil[25] demonstrated that the enolic form of protonated uracil, strictly similar to **1MeU2_ha** and **6-MeU2_ha**, is totally unreactive upon CID activation, and that interconversion towards more reactive protonated forms has to take place to explain the CID fragments observed experimentally.

3.4 IRMPD spectra of protonated 3-Me-uracil

Figure 4 compares the experimental IRMPD spectrum of protonated 3-Me-uracil (Figure 4a) to the computed vibrational spectrum of the two most stable optimized keto forms, and of the **3MeU2_ha** structure (Figure 4b-d). The experimental IRMPD spectrum is strikingly different from those just previously discussed, indicating that protonation of 3-Me-uracil leads to a

different structure. In spite of the decrease in the FEL power above 1700 cm^{-1} , one can observe a particularly intense signal around 1780 cm^{-1} , which contrasts markedly with the weak feature observed for the two other nucleobases. The IRMPD spectrum is also characterized by a sharp signal at 1206 cm^{-1} . The IRMPD signature between 1450 and 1600 wavenumbers also differs significantly from what was measured for uracil, 1-Me-uracil, 6-Me-uracil and thymine. Comparison of Figure 4a and 4b shows an excellent agreement between the experimental trace and the DFT-computed vibrational spectra of the calculated global minimum, namely **3MeU1_hd**. Assignment of all the IRMPD can be achieved by considering the IR active modes of **3MeU1_hd** (Table 2). The experimental band at 1780 cm^{-1} can be attributed to the strongly active C2=O7 carbonyl stretch. The signal detected at 1206 cm^{-1} can result from the superposition of two IR modes close in energy (1202 and 1230 cm^{-1}), corresponding to the combination of CH and NH bending modes, and the bending of the hydroxyl group, respectively. Finally, the shape of the IRMPD spectrum between 1440 and 1600 cm^{-1} is also in very good agreement with the infrared cross section of **3MeU1_hd** isomer. This feature can be decomposed in four distinct signals at 1446 , 1510 , 1546 and 1597 cm^{-1} which can be ascribed to CH (1439 cm^{-1}) and NH (1450 cm^{-1}) bending modes, the C4O8 (1492 cm^{-1}), C4C5 (1551 cm^{-1}) and C5C6 (1597 cm^{-1}) stretching modes, respectively. The rotamer **3MeU1_hc** can also reproduce some of the signals detected experimentally but the agreement with experiment is much better for **3MeU1_hd**, which is consistent with the lower stability of the **3MeU1_hc** form as discussed above. Finally, one can see (Figure 4c) that the enolic form **3MeU2_ha**, whose structure is similar to that observed for 1-Me- and 6-Me-uracil, can be reasonably precluded. This is again consistent with the expected low stability of this form as we have discussed above. The same applies to form **3MeU2b_hb** (Figure 2), which lies 29.5 kJ/mol above the global minimum, and whose calculated normal modes do not match the experimental trace. Consequently, one can conclude from this set of data that

the protonation of 3-Me-uracil under electrospray conditions results exclusively in the formation of a keto form, which according to our computational survey should correspond to the **3MeU1_hd** structure.

Methylation of uracil onto the N3 atom therefore prevents the tautomerization process. The presence of the methyl group on N3 avoids the attractive stabilization between the nitrogen lone pairs and positively-charged hydrogens from the neighboring hydroxyl groups. Furthermore, this finding is consistent with the solvent-assisted mechanism as proposed in different studies, accounting for the interconversion of the keto form into an enolic form[5,12], which presently cannot occur due the presence of this methyl group. By contrast, the proposed mechanism is possible for the other uracil compounds.

Conclusion

Comparison of the present results with those deduced from previous studies demonstrates that 3-Me-uracil constitutes a particular case in the series of pyrimidic nucleobases we have studied so far. On the one hand, its protonation under electrospray conditions results in the formation of a protonated keto form. On the other hand, the most abundant structure generated under electrospray conditions for cytosine, uracil, 1-Me-uracil, 5-methyluracil (thymine) and 6-Me-uracil corresponds to an enolic form, a minor proportion of a keto form being also evidenced. This study also shows that the free C=O stretching mode of protonated nucleobases can be particularly intense under IRMPD conditions. These new data give further support to the interpretation provided in our previous work and are consistent with the solvent-assisted mechanism proposed to explain the tautomerization of nucleobases under electrospray conditions.

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Supporting information

Structures and energetics of the different optimized structures are provided as Supporting Information, together with IRMPD and computed IR absorption spectra for protonated 6-Me-uracil.

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Schemes, Tables and Figures captions

Scheme 1: Most stable calculated neutral tautomers of 1-Me, 3-Me and 6-Me uracil.

Scheme 2: Interactions that contribute to the stabilization and destabilization of conformers **ha** and **hb** of **1MeU2_ha** and **6MeU2_ha** and **3MeU2-ha** systems.

Table 1: Experimental IRMPD spectrum of protonated 1-Methyluracil compared with computed vibrational modes for the **1MeU1_hd** and **1MeU2_ha** structures

Table 2: Experimental IRMPD spectrum of protonated 3-Methyluracil compared with computed vibrational modes for the **3MeU1_hd** and **3MeU1_hc** structures

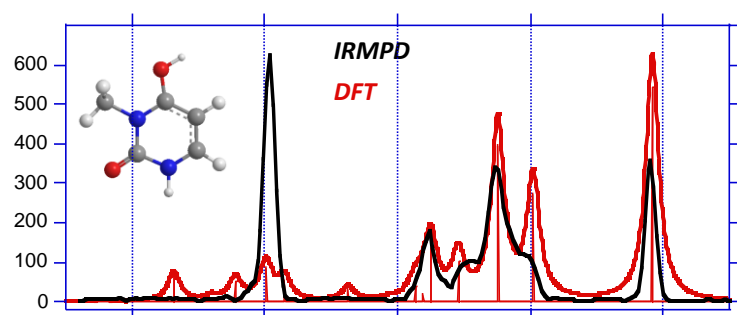
Figure 1: Experimental IRMPD spectra recorded for (a) 1-Methyluracil, (b) 3-Methyluracil and (c) 6-Methyluracil.

Figure 2: Structure and B3LYP/6-311++G(3df,2p)+ZPE relative energies (kJ/mol) of the lowest energy structures obtained of protonated 1-Me-, 3-Me- and 6-Me-uracil . See Sections S-1-3 of the Supporting information for a complete survey of the various structures considered.

Figure 3: IRMPD spectrum of protonated 1-Methyluracil. DFT calculated IR absorption spectra of the **1MeU2_ha** (b), **1MeU1_hd** (c) and **1MeU1_hc** (d) isomers compared to the experimental IRMPD spectrum (a) recorded with an irradiation time of 300 ms

Figure 4: IRMPD spectrum of protonated 3-Methyluracil. DFT calculated IR absorption spectra of the **3MeU1_hd** (b), **3MeU2_ha** (c) and **3MeU1_hc** (d) isomers compared to the experimental IRMPD spectrum (a) recorded with an irradiation time of 750 ms

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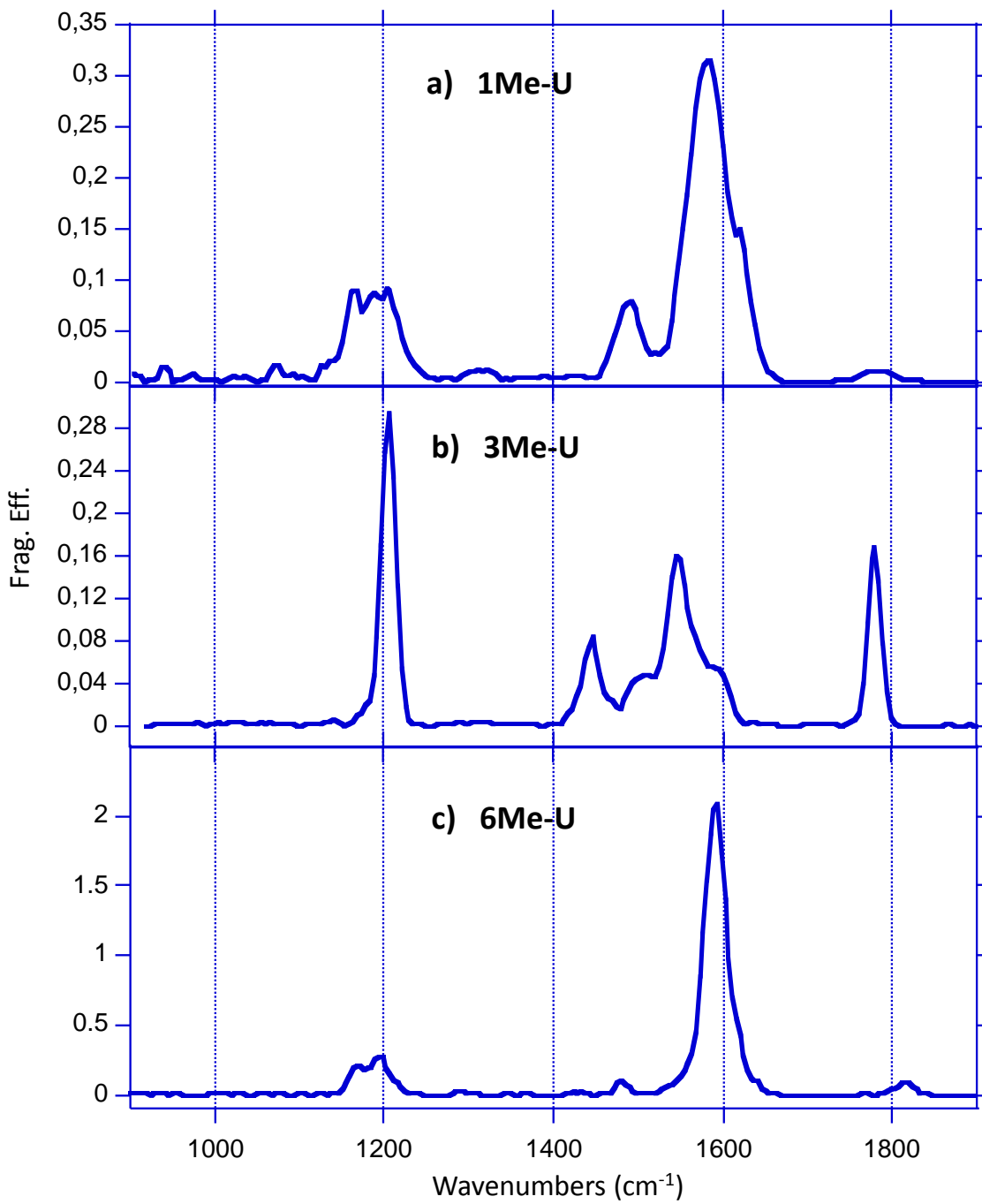


Figure 1

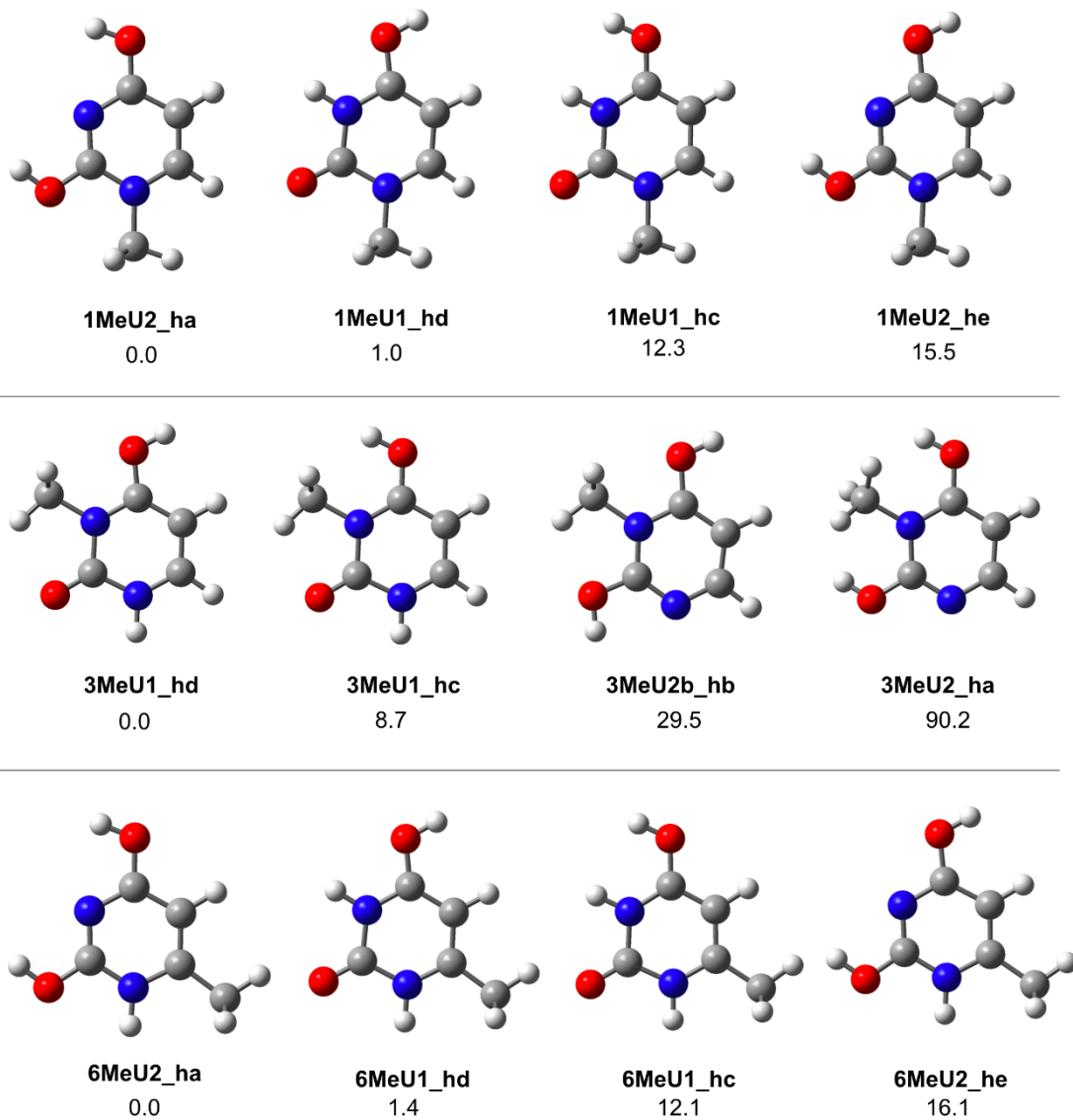


Figure 2

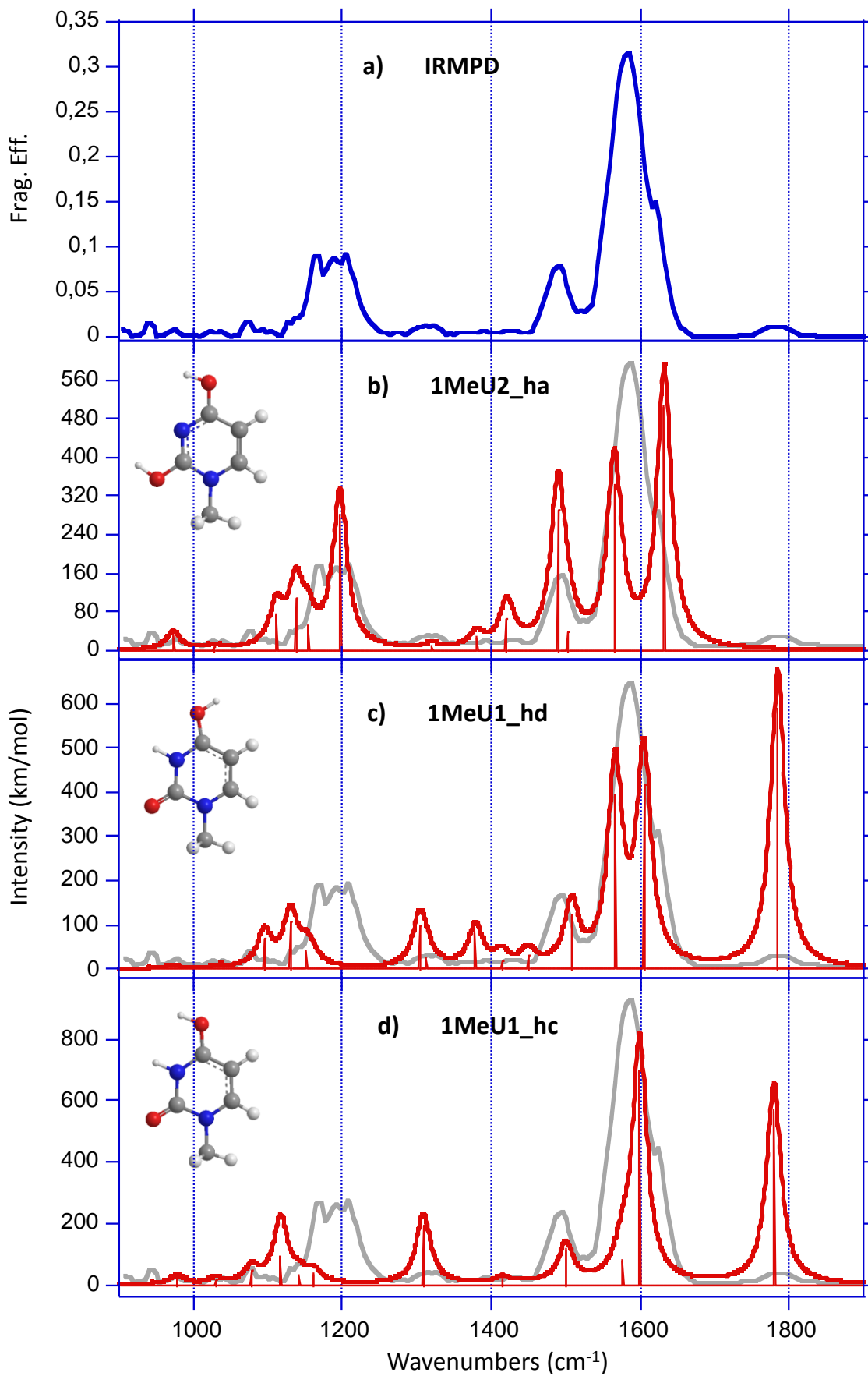


Figure 3

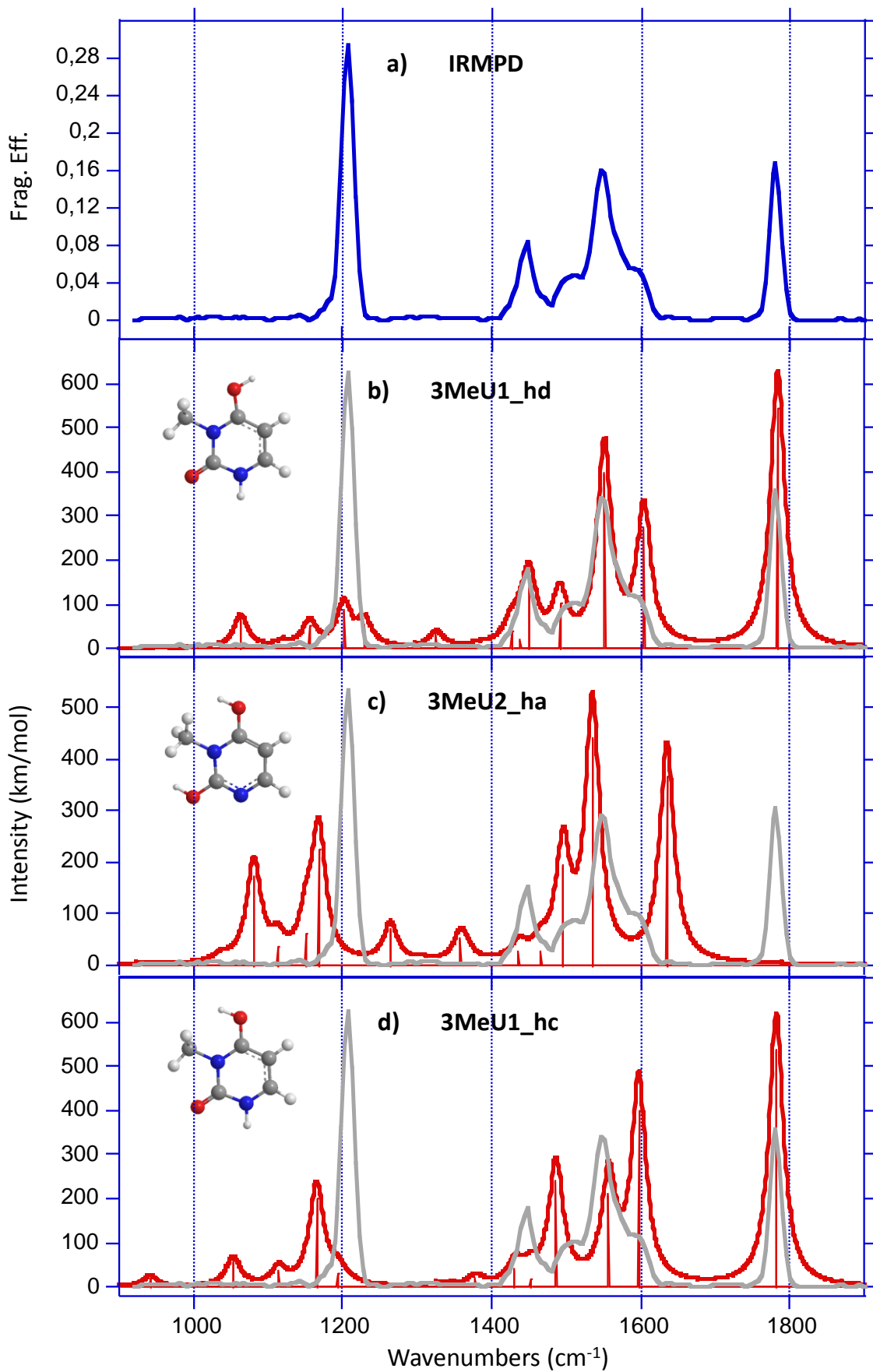


Figure 4

Table 1: Experimental IRMPD spectrum of protonated 1-Methyluracil compared with computed vibrational modes for the **1MeU1_hd** and **1MeU2_ha** structures

Wavenumber (cm ⁻¹)		Calculated intensities (km.mol ⁻¹)		Vibrational mode
Exp.	Calc.	1MeU1_hd	1MeU2_ha	
1150-1220	1153	40		δ C5H
	1154		50	δ C2OH + ν N1-CH ₃
	1196		281	Combined δ COH modes (rock.)
1317	1304	99		δ N3H
	1322		10	δ C5H+ δ C6H
1490	1490		291	ν C4O8
	1502		38	ν C2O7 +ν N3C4
	1508	122		ν N3C4
1580	1566		342	ν C4C5
	1567	391		ν C4C5
1620	1605	416		ν C5C6
	1631		504	ν C5C6 + ν C2N3
1785	1779	585		ν C2O7

Table 2: Experimental IRMPD spectrum of protonated 3-Methyluracil compared with computed vibrational modes for the **3MeU1_hd** and **3MeU1_hc** structures

Wavenumber (cm ⁻¹)		Calculated intensities (km.mol ⁻¹)		Vibrational mode
Exp.	Calc.	3MeU1_hd	3MeU1_hc	
1206	1193		32	v C2-N3
	1202	86		δ C6H + δ N1H
	1230	52		δ C4OH
1446	1427	38		δ CH ₂ sciss. of CH ₃
	1430		41	δ CH ₂ sciss. of CH ₃
	1439	20		δ CH of CH ₃
	1451	139		δ N1H
	1453		20	δ CH of CH ₃
1510	1486		240	v C4O8
	1492	101		v C4O8
1546	1551	394		v C4C5
	1557		209	v C4C5
1597	1597		401	v C5C6
	1604	271	504	v C5C6
1780	1782		537	v C2O7
	1784	545		v C2O7