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## Carbonatation and Decarbonation Kinetics in the La<sub>2</sub>O<sub>3</sub>-La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> System under CO<sub>2</sub> Gas Flows

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## Research Article

# Carbonatation and Decarbonation Kinetics in the $\text{La}_2\text{O}_3$ - $\text{La}_2\text{O}_2\text{CO}_3$ System under $\text{CO}_2$ Gas Flows

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The carbonatation of  $\text{La}_2\text{O}_3$  oxide and the decarbonation of lanthanum carbonate phase  $\text{La}_2\text{O}_2\text{CO}_3$  are investigated using thermal and thermogravimetry analyses under  $\text{CO}_2$  gas flow. The initial phase  $\text{La}_2\text{O}_3$  is first elaborated from pyrolysis of a  $\text{LaOHCO}_3$  precursor. Then, thermal and thermogravimetry analyses are carried out under  $\text{CO}_2$  flow, as temperature increases then decreases. The carbonatation kinetics of  $\text{La}_2\text{O}_3$  is determined at three fixed temperatures. Electrical impedance spectroscopy is performed to determine the electrical responses associated with ionic mobilities and phase changes, in the temperature range 25 to 900°C. The electrical conduction during heating under  $\text{CO}_2$  gas flow should be linked to two regimes of ionic conduction of the carbonate ions. From these electrical measurements, the ionic mobility of carbonate ions  $\text{CO}_3^{2-}$  is found to be close to  $0.003 \cdot 10^{-4} \text{ cm}^2 \text{ s}^{-1} \text{ V}^{-1}$  at 750°C for the monoclinic  $\text{La}_2\text{O}_2\text{CO}_3$  phase.

## 1. Introduction

The lanthanum-based system  $\text{La}_2\text{O}_3$ - $\text{CO}_2$ - $\text{H}_2\text{O}$  is characterized by successive phases  $\text{LaOHCO}_3$ ,  $\text{La}_2\text{O}_2\text{CO}_3$ , and  $\text{La}_2\text{O}_3$ , stable in various temperature ranges [1–9], depending on the partial pressures of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  [10–15]. It has been established that the decomposition of the hydroxycarbonate  $\text{LaOHCO}_3$  under air generally gives the  $\text{La}_2\text{O}_2\text{CO}_3$  dioxycarbonate phase [6]. However, this last phase exists under three polymorphic structural varieties with tetragonal, hexagonal, and monoclinic crystal lattices [16–19]. The carbonatation of  $\text{La}_2\text{O}_3$  under pure  $\text{CO}_2$  generally gives the main  $\text{La}_2\text{O}_2\text{CO}_3$  phase; however, the obtained system can be complex, with presence of several polymorphic structures depending on the experimental synthesis conditions [10, 18, 20]. In their work concerning the TG analysis of the  $\text{La}_2\text{O}_2\text{CO}_3$  phase under  $\text{CO}_2$  gas, the authors of [2] showed that a small amount of  $\text{La}_2\text{O}(\text{CO}_3)_2$  was probably formed as an additional phase. This phase was also studied by authors of [21].

In room conditions, the lanthanum oxide is highly sensitive to environmental water. In a previous study [22], we have established correlations between the thermal decomposition and the electrical responses of compacted pellets of this  $\text{LaOHCO}_3$  phase, subjected to pyrolysis under air: we have shown that strong variations in conductances accompanied these phase changes. We have also established that these  $\text{LaOHCO}_3$ ,  $\text{La}_2\text{O}_2\text{CO}_3$ , and  $\text{La}_2\text{O}_3$  phases have the capacity to convert carbon monoxide into  $\text{CO}_2$  at relatively low temperature: at 200–300°C, the L phase is a good catalyst converting  $\text{CO}$  into  $\text{CO}_2$ , while it might be sensitive to  $\text{CO}_2$  only above 500°C.

In the present study, we focus our attention on phase changes during carbonatation and decarbonation processes, respectively, of the  $\text{La}_2\text{O}_3$  phase and of the  $\text{La}_2\text{O}_2\text{CO}_3$  phase. The main objective of this approach should reside in connecting the weight variations due to these phase transformations with electrical responses, in order to appreciate their potential efficiency in gas sensing devices. These correlations between mass losses and electrical responses are

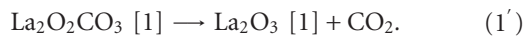
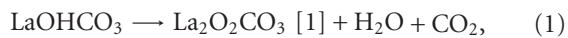
not known, and they could deliver interesting information on the electrical sensitivity of such systems.

## 2. Experimental Details

The LaOHCO<sub>3</sub> hydroxycarbonate was first prepared via a specific route [22, 23] based on a thermal treatment at 80°C of three aqueous solutions of La(NO<sub>3</sub>)<sub>3</sub>·6H<sub>2</sub>O, urea CO(NH<sub>2</sub>)<sub>2</sub>, and polyvinyl-pyrrolidone (PVP) polymer. The La<sub>2</sub>O<sub>3</sub> oxide was obtained by pyrolysis of this LaOHCO<sub>3</sub> precursor.

The various chemical steps can be summarized as follows:

- (i) First initial decomposition processes under air as temperature increases (25–1200°C):



- (ii) Carbonatation and decarbonatation under pure CO<sub>2</sub> as temperature increases (25–1200°C):



- (iii) Recarbonatation under pure CO<sub>2</sub> as temperature decreases (1200 to 25°C):



In the previous equations, in bracket [1 to 3] we have designated phases obtained after a transformation process (decomposition, carbonatation, and decarbonatation). These phases have not the same characteristics (various morphologies and specific surfaces).

The polycrystalline samples were systematically analyzed by X-ray diffraction, using a D5000 Siemens-Bruker diffractometer, equipped with a copper X-ray source (wavelength  $\lambda = 1.54 \cdot 10^{-10}$  m), and with a monochromator eliminating K $\beta$  radiation. The experiments were carried out using classical  $\theta - 2\theta$  configuration.

Thermal and Thermogravimetric analyses (DTA-TG) were carried out using SETARAM DSC 92 equipment, with a thermal rate of 10°C/minute, under CO<sub>2</sub> pure gas (rate of flow of 33 cm<sup>3</sup>·s<sup>-1</sup>).

Electrical measurements under CO<sub>2</sub> gas flow were performed using a Solartron electrical impedance spectrometer working with a maximal tension of 1V, in the frequency range 100 to 10<sup>7</sup> Hz. A reactive homemade cell was used to perform experiments under various gas flows (air, CO<sub>2</sub>) at various temperatures ranging between 25 and 900°C. The spectrometer delivers Nyquist representations of the resulting impedances recorded at fixed temperatures: the resistance value is classically obtained by extrapolation of the experimental Nyquist circles, and using electrical equivalent circuits (parallel R-C circuits) generated by the software. We

have selected specific electrical circuits with a resistance ( $R$ ) parallel to a constant phase element  $\text{CPE} = (jC^*\omega)^n$  where the exponent  $n$  is comprised between 1 and 0, and  $C^*$  is a term similar to a capacitance for  $n = 1$  (the unit of  $C^*$  depends on  $n$ ).

To obtain electrical analyses of sample surfaces reacting with gas flows, the powder samples were first compacted under a pressure of 5 kbar in a cylindrical cell. Then, the obtained cylindrical pellet was cut in form of a rectangular plate, with platinum electrodes fixed on two parallel faces (dimensions 2.3 × 8 mm). The distance between the electrodes is 9 mm. This configuration (adapted to the reactive cell) allows a determination of the electrical properties of a significant material surface exposed to gas action. In a later step, these results might be used to test a hypothetical gas sensor sensitive to CO<sub>2</sub>.

## 3. Results

### 3.1. Carbonatation-Decarbonatation Processes

**3.1.1. Heating Process under CO<sub>2</sub> Flow.** The La<sub>2</sub>O<sub>3</sub> sample, initially obtained from thermal decomposition of LaOHCO<sub>3</sub>, has been subjected to thermal and thermogravimetry analyses under CO<sub>2</sub> gas flow, with temperature increasing from 25 to 1200°C. The resulting TG-DTA curves are reported on Figure 1. A strong exothermic DTA peak is observed at 525°C: it is related to the carbonatation of La<sub>2</sub>O<sub>3</sub> with formation of the La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> phase. Then, at 960°C, we observe an endothermic feature corresponding with the decomposition of the carbonate phase. Above 980°C the La<sub>2</sub>O<sub>3</sub> phase stabilizes. A small endothermic feature is observed at 375°C: it might be associated with a partial dehydration of the sample due to the high sensitivity to environmental water of La<sub>2</sub>O<sub>3</sub>. The progressive mass evolution observed in the TG curve of Figure 1, as temperature increases, is directly associated with the classical buoyancy. A similar effect will be observed during the cooling process.

**3.1.2. Cooling Experiments under CO<sub>2</sub> Flow.** Using cooling experiments, we have analyzed the carbonatation of La<sub>2</sub>O<sub>3</sub> from 1200°C to 25°C. The results are represented on Figure 2. The formation of La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> starts from 820°C and is maximum at 750°C. The exothermic peak associated with the crystallization of La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> carbonate is observed at 790°C. This temperature of carbonatation is strongly different from the one obtained during the heating process.

At each step involving a stabilized phase, we have carried out X-ray diffraction analyses to identify the obtained phases. We have confirmed that, in the case of thermal decomposition under air of LaOHCO<sub>3</sub> phase, two different tetragonal and hexagonal La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> structures are simultaneously observed. In the case of carbonatation of the La<sub>2</sub>O<sub>3</sub> phase in the temperature range 500 to 700°C, we observe the formation of the La<sub>2</sub>O<sub>2</sub>CO<sub>3</sub> phase. The La<sub>2</sub>O(CO<sub>3</sub>)<sub>2</sub> phase was not observed in our experiments. This fact was previously reported by other authors [10, 18]. On Figure 3, we have reported the X-ray diffraction pattern characteristic

TABLE 1: Parameters extracted from Avrami's model:  $k_1$  and  $k_2$  kinetics parameters and  $p_1$  and  $p_2$  exponents, respectively, associated with the fast and slow regimes (1st and 2nd periods).

$T$ ( $^{\circ}\text{C}$ )	1st period			2nd period		
	Time range (min)	$k_1$	$p_1$	Time range (min)	$k_2$	$p_2$
450	45 $\rightarrow$ 112	$1.5 \cdot 10^{-7}$	3.3	167 $\rightarrow$ 320	0.02	0.9
480	6 $\rightarrow$ 17	$1.4 \cdot 10^{-5}$	3.9	30 $\rightarrow$ 90	0.54	0.3
500	3 $\rightarrow$ 7	$1.2 \cdot 10^{-3}$	3.4	30 $\rightarrow$ 90	1.10	0.2
Activation energies $E_1$ ( $k_1$ )		7.6 (eV)		Activation energies $E_2$ ( $k_2$ )		2.8 (eV)

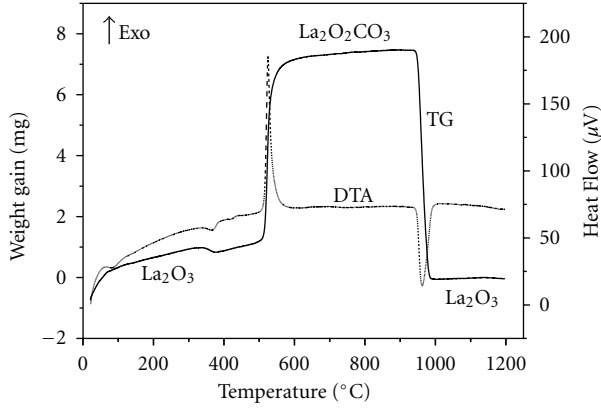


FIGURE 1: Weight gain associated with carbonatation of  $\text{La}_2\text{O}_3$  (first step): exothermic peak at  $520^{\circ}\text{C}$  linked with weight gain due to formation of the monoclinic  $\text{La}_2\text{O}_2\text{CO}_3$  phase; endothermic peak due to decarbonation of  $\text{La}_2\text{O}_2\text{CO}_3$  phase and formation of final  $\text{La}_2\text{O}_3$  (second step).

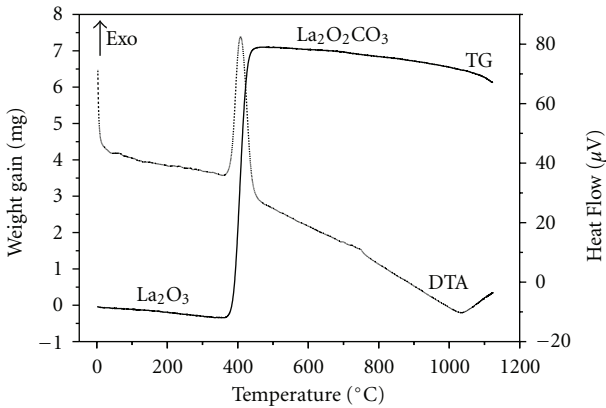


FIGURE 2: Evolution of  $\text{La}_2\text{O}_3$  weight during cooling process, under  $\text{CO}_2$  flow: formation of  $\text{La}_2\text{O}_2\text{CO}_3$  phase (exothermic peak) then relative stabilization of this phase as temperature decreases.

of the monoclinic  $\text{La}_2\text{O}_2\text{CO}_3$  phase heated at  $520^{\circ}\text{C}$  under  $\text{CO}_2$  flow, during 3 hours. The refined cell parameters are  $a = 0.4073 \pm 0.0003$  nm;  $b = 1.3503 \pm 0.0008$  nm;  $c = 0.4079 \pm 0.0005$  nm;  $\beta = 90.89^{\circ}$ . In the pattern, a weak trace of the hexagonal phase (noted as  $*$ ) is observed.

**3.2. Kinetics Study of Carbonatation of  $\text{La}_2\text{O}_3$  at Fixed Temperatures.** We have performed a weight analysis of the  $\text{La}_2\text{O}_3$

powder, obtained from the thermal decomposition of the initial  $\text{LaOHCO}_3$  phase, under  $\text{CO}_2$  gas flow at three constant temperatures. The  $\text{CO}_2$  gas flow rate was  $33 \text{ cm}^3 \cdot \text{s}^{-1}$ . In the SETARAM equipment, a fast temperature increase is first applied to the sample, and then, the temperatures are successively fixed to  $450$ ,  $480$ , and  $500^{\circ}\text{C}$ . The three initial masses of  $\text{La}_2\text{O}_3$  are successively (at  $T = 450$ ,  $480$ , and  $500^{\circ}\text{C}$ )  $m_0 = 74.36$  mg,  $70.73$  mg, and  $39.26$  mg. The data evolutions have been interpreted in terms of an elemental Avrami's model [24] (using a single mechanism approach):

$$\Delta m = \Delta m_0 [1 - \exp(-k \cdot t^p)], \quad (4)$$

- (i)  $t$  is the reaction time;
- (ii)  $\Delta m_0$  is the limit mass of  $\text{CO}_2$  involved in the carbonate formation  $\text{La}_2\text{O}_2\text{CO}_3$  from a mass  $m_0$  of  $\text{La}_2\text{O}_3$ ;
- (iii)  $\Delta m$  is the  $\text{CO}_2$  mass having reacted with  $\text{La}_2\text{O}_3$  at the time  $t$ ;
- (iv)  $k$  is a kinetics parameter depending of temperature;
- (v)  $p$  is the exponent characteristic of the reaction mechanism ( $p > 2$  for complex mechanisms,  $p < 1$ , for example, for mechanisms involving diffusion barriers).

To test the degree of validity of this Avrami's model, we have reported the function  $Y$  versus  $\ln(t)$  on Figure 4:

$$Y = \ln \left[ -\frac{\ln(\Delta m_0 - \Delta m)}{\Delta m_0} \right] = \ln(k) + p \ln(t). \quad (5)$$

For a single crystal growth mechanism, the variation of  $Y$  versus  $\ln(t)$  should have been linear. Presently, the representation of Figure 4 is not linear: this should be mainly due to the existence of at least two different crystal growth mechanisms, with two periods of mass gain corresponding to two behaviors.

In Table 1, we have reported the values of the kinetics parameters  $k_1$  and  $k_2$  and exponents  $p_1$  and  $p_2$ , corresponding with the two different behaviors in which a linear correlation might be observed. The parameters  $k_1$ ,  $p_1$  are relative to the first period depending on temperature, and the parameters  $k_2$ ,  $p_2$  are relative to the second period. The  $k_1$  and  $k_2$  are thermally activated with activation energies of, respectively,  $7.6$  and  $2.8$  eV. The  $p_1$  exponent is quasi-constant, while the  $p_2$  exponent is close to 1 at  $450^{\circ}\text{C}$  and becomes very weak at higher temperatures.

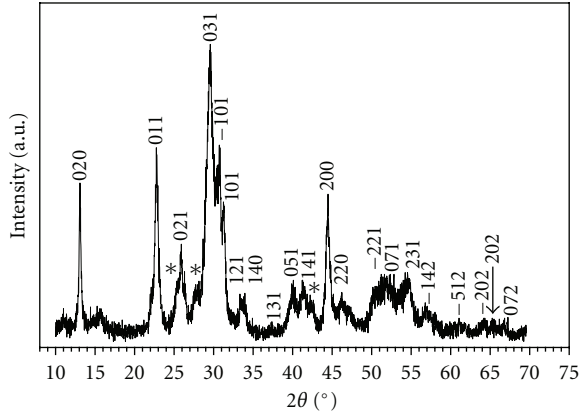


FIGURE 3: X-ray diffraction pattern of  $\text{La}_2\text{O}_2\text{CO}_3$  (monoclinic) obtained by heating  $\text{La}_2\text{O}_3$  at  $520^\circ\text{C}$  under  $\text{CO}_2$  flow. Trace of hexagonal phase (noted \*).

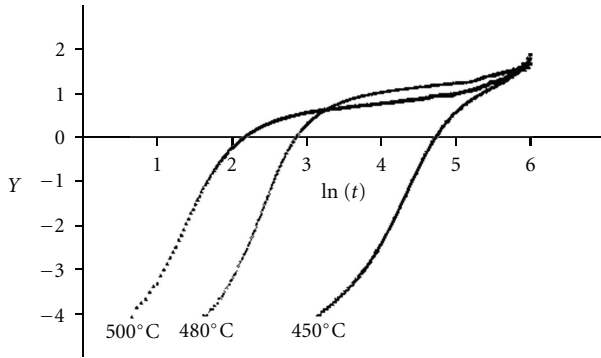


FIGURE 4: Test of Avrami's model validity: the representation of  $Y = \ln[-\ln(\Delta m_0 - \Delta m)/\Delta m_0]$  versus  $\ln(t)$  ( $t$  in mn) shows that two main types of behaviors can be observed as time increases.

The first growth regime should be associated with a fast carbonatation of grain surfaces associated with complex diffusion mechanisms. During this period, a carbonate shell enveloping oxide grains probably should be formed. The second growth regime should be associated with reaction and diffusion in grain cores, with a decrease of the reaction rate due to the carbonate shell: the resulting slow diffusion regime could govern the global reaction speed.

**3.3. Electrical Analyses under  $\text{CO}_2$  Gas Flows.** To correlate the phase modifications to electrical behaviors, we have analyzed compacted powder samples in the electrical cell. In this experiment, a rectangular compacted sample resulting from the total decomposition of the initial  $\text{LaOHCO}_3$  sample has been subjected to a progressive heating, under pure  $\text{CO}_2$  gas flow. Between  $600$  and  $700^\circ\text{C}$ , carbonatation occurs, thus involving a strong increase in conductance mainly due to the ionic mobility of  $\text{CO}_3^{2-}$  carbonate ions. Then, above  $750^\circ\text{C}$  decarbonation occurs, involving a decrease of conductance due to  $\text{CO}_3^{2-}$  carbonate ions elimination and formation of  $\text{La}_2\text{O}_3$ . This oxide should be formed at  $950^\circ\text{C}$ .

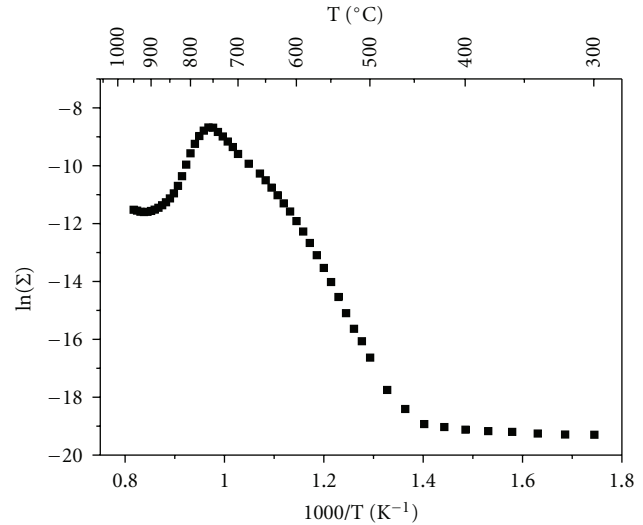


FIGURE 5: Evolution of  $\ln(\Sigma)$  versus  $10^3/T$  of initial  $\text{La}_2\text{O}_3$  sample during its carbonatation and decarbonation, in the temperature range  $300$  to  $950^\circ\text{C}$ , under a constant  $\text{CO}_2$  gas flow: (a) starting step of carbonatation between  $450$  and  $750^\circ\text{C}$  with two conduction regimes; (b) decarbonation step above  $750^\circ\text{C}$ .

On Figure 5, we have reported the  $\ln(\Sigma)$  values versus temperature (total  $\Sigma$  values). We observe a strong decrease of  $\ln(\Sigma)$  between  $450^\circ\text{C}$  and  $750^\circ\text{C}$ : the carbonatation of  $\text{La}_2\text{O}_3$  should start from  $450^\circ\text{C}$ , with a first regime up to  $600^\circ\text{C}$  and a second regime up to  $750^\circ\text{C}$ . Two activation energies for the conduction behavior can be determined:  $2.5$  (first regime) and  $1.4$  eV (second regime). In this carbonatation domain, the ionic conduction plays a major role with mobile species  $\text{CO}_3^{2-}$ .

Above  $750^\circ\text{C}$ , we observe a strong decrease in the  $\ln(\Sigma)$  values: in this temperature range, decarbonation occurs in a continuous way, with the elimination of  $\text{CO}_3^{2-}$  ions. As  $\text{La}_2\text{O}_3$  phase stabilizes, the resistance reaches a stabilized value.

Using the observed value  $\Sigma = 1.7 \cdot 10^{-4} \Omega^{-1}$  at  $750^\circ\text{C}$  (on Figure 5: maximum value of  $\Sigma$  just before decomposition), and considering as negligible the conductance of  $\text{La}_2\text{O}_3$  at the same temperature (close to  $10^{-8}$  to  $10^{-9} \Omega^{-1}$  at  $750^\circ\text{C}$ ), we have evaluated an ionic conductance due to  $\text{CO}_3^{2-}$  ions to  $\Delta\Sigma = 1.7 \cdot 10^{-4} \Omega^{-1}$ . From this evaluation of  $\Delta\Sigma$ , we have determined the order of magnitude of the carbonate ion mobility  $u(\text{CO}_3^{2-})$  at  $750^\circ\text{C}$ . Other values could be derived from the data obtained in the temperature range  $450$  to  $700^\circ\text{C}$ . The concentration of carbonate ions  $C_{\text{ion}}$  has been calculated from the effective density of the sample  $\mu = 5.1 \text{ g}\cdot\text{cm}^{-3}$  (for a theoretical crystal density of  $6.51 \text{ g}\cdot\text{cm}^{-3}$ ) and using the sample volume  $V = 0.1656 \text{ cm}^3$ . A value of  $C_{\text{ion}} = 0.0155 \text{ mol}\cdot\text{cm}^{-3}$  has been obtained. To determine the mobility, we have used the classical relations:

$$\Delta\Sigma = \sigma \frac{S}{L}, \quad \sigma = Q \cdot C_{\text{ion}} \cdot u(\text{ion}), \quad (6)$$

where  $\sigma$  is the conductivity,  $S$  and  $L$  are the surface and separation distance of the two electrodes, and where

$Q = 193\,000\text{ C}\cdot\text{mol}^{-1}$ . The relation giving the conductivity assumes an activity coefficient of 1: it only delivers an order of magnitude for the mobility.

We have obtained an order of magnitude of  $u(\text{CO}_3^{2-}) = (0.003 \pm 0.001) 10^{-4}\text{ cm}^2\text{ s}^{-1}\text{ V}^{-1}$  for a carbonate ion moving at  $750^\circ\text{C}$  mainly along grain boundaries (or grain surfaces), and partly in the grain cores. This relatively high mobility can be associated with the activation energy of  $1.4\text{ eV}$  (in the temperature range  $600$  to  $750^\circ\text{C}$ ) as calculated above.

#### 4. Discussion-Conclusions

The carbonation kinetics of  $\text{La}_2\text{O}_3$  has been determined at various temperatures. In the case of mass gain analyses, an elemental Avrami's approach has allowed determining a complex two-step mechanism of growth: (i) a fast surface carbonation with carbonate shell formation and (ii) a diffusion mechanism in grain cores with slower kinetics. The electrical analyses argue in favor of two different conduction mechanisms: during carbonation at increasing temperature, the first activation energy ( $2.5\text{ eV}$ ) should be associated with ionic conduction at grain surfaces, and the second activation energy ( $1.4\text{ eV}$ ) should be due to an increasing contribution of the conduction in the bulk. Correlatively, it should be remarked that, in thermal analyses, the stability range is observed from  $500$  to  $850^\circ\text{C}$ , while in electrical analyses, this stability range is observed from  $500$  to  $750^\circ\text{C}$ . This can be explained by the two different heating kinetics conditions used in the two experiments.

Finally, we observe a relatively high ionic mobility mainly due to the  $\text{CO}_3^{2-}$  ions in  $\text{La}_2\text{O}_2\text{CO}_3$  at  $750^\circ\text{C}$ . In our evaluation, we have neglected the ionic conduction of oxygen ions.

It should be concluded that these phase modifications associated with high ionic conduction might be used as electrical sensitive material to detect  $\text{CO}_2$ , provide temperatures that could be fixed close to  $400$ – $550^\circ\text{C}$  (carbonation of  $\text{La}_2\text{O}_3$  phase) and  $750^\circ\text{C}$  to restore the initial  $\text{La}_2\text{O}_3$  phase.

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