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A systematic first principles study of binary metal hydrides

Natacha Bourgeois, Jean-Claude Crivello,* Pierre Cenedese, and Jean-Marc Joubert

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Abstract

First principles calculations were systematically performed for 31 binary metal-hydrogen (M–H) systems on a set of 30 potential crystal structures selected on the basis of experimental data and possible interstitial sites. For each M–H system, the calculated enthalpies of formation were represented as functions of H composition. The zero point energy correction was considered for the most stable hydrides via additional harmonic phonon calculations. The sequence of stable hydrides (ground-state) given by the convex hull was found in satisfactory agreement with the experimental data. Besides, new high pressure dihydrides and trihydrides were predicted, providing orientations for new materials synthesis. The overall results provide a global overview of hydride relative stabilities and relevant input data for thermodynamic modeling methods.

Introduction

Predicting the formation of reversible metal hydrides is essential in a wide range of environmental issues such as hydrogen storage for energetic use,\textsuperscript{1,2} prevention of structural...
alloys embrittlement and electrode development for nickel-metal-hydride batteries. The enthalpy of formation $\Delta H_{\text{for}}$ is the key physical quantity because it is linked with the equilibrium pressure via the van’t Hoff relation and represents therefore the main contribution to the stability of a given hydride relatively to the pure elements. For this reason, the search for models describing hydride enthalpy of formation induced many studies. The first models in the 1970s were empirical. The progress in computational science allowed calculations based on the electronic properties, able to provide the enthalpy of formation of stable as well as metastable hydrides. More recently, Density Functional Theory (DFT) systematic studies were carried out considering different hydrides structures, stable as metastable. Smithson et al. and Miwa and Fukumoto introduced simultaneously the idea that the formation enthalpy can be divided into three contributions: a positive term resulting from the expansion of the lattice, a positive term needed for a potential structure change and a negative term arising from hydrogen insertion in an expanded lattice.

The previous DFT studies were however limited to a restricted number of crystal structures. The present work intends to screen all the potential binary hydride crystal structure prototypes, on a composition range up to $y = 3$, which was not done in previous work. The major interest of our approach is to predict the stable crystal structures and to estimate the metastability of all hydrides.

The selected set of crystal structure prototypes includes (a) the ones experimentally observed among the considered systems and (b) the ones corresponding to all the possible filling of the interstitial sites in basic solid solutions such as face-centered cubic (fcc), body-centered cubic (bcc), hexagonal close-packed (hcp) and simple cubic structures. The 30 considered prototypes are listed in Table 1 and illustrated in supplementary materials.

High-throughput investigation by massive scale DFT calculations of the energy of formation was performed for 31 metals $M$ distributed in the 30 different $MH_y$ crystal structures prototypes, resulting in 930 hydrides. For each system, the ground-state was determined by the convex hull of the energies of formation as functions of hydrogen composition.
The present paper is organized as follows: after the methodology description, the predicted ground-states and energies of formation are presented and compared with the experimental data. Finally, the results are discussed in a separate section.

Table 1: Structure prototypes included in the study: H/M ratio, Prototype name, Pearson Symbol, Space Group (number), Wyckoff Positions, Interstitial site, M–H system where the structure is observed. When the prototype appears only at high pressure or is considered as metastable for a given M–H system, the metal is highlighted in italic in the last column. The interstitial sites are described as octahedral (o), tetrahedral (te), triangular (tri) or by the coordination number (CN) when the geometry is more complex. The structures called PdH₉ are superstructures of the NaCl prototype.

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<td>P₆₃/mmc (194)</td>
<td>M: 2c, 2a H : 4f</td>
<td>o</td>
<td>Fe</td>
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<td>Fm₃m (225)</td>
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<td>Mg</td>
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<td>tI₆</td>
<td>I₄/mmm (139)</td>
<td>M: 2a, H : 4d</td>
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<td>Ti, Zr, Hf</td>
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<td>CN=2</td>
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<td>BiF₃</td>
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<td>Fm₃m (225)</td>
<td>M: 4a, H : 4b, 8c</td>
<td>o, te</td>
<td>-</td>
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<tr>
<td>3</td>
<td>ReB₃</td>
<td>hP₈</td>
<td>P₆₃/mmc (194)</td>
<td>M: 2c, H : 2a, 4f</td>
<td>o, te</td>
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<td>3</td>
<td>Ho₃</td>
<td>hP₂₄</td>
<td>P₃c₁ (165)</td>
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<td>tr, te</td>
<td>Gd</td>
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</table>
Methodology

The DFT method\textsuperscript{17,18} which has proven to be one of the most accurate methods for the computation of the electronic structure of hydrides\textsuperscript{19-22} was implemented using the Vienna Ab-initio Simulation Package (VASP).\textsuperscript{23,24} The generalized gradient approximation (GGA) was used for the exchange and correlation energy functional with the Perdew-Burke-Ernzerhof functional (PBE).\textsuperscript{25,26} The number of valence electrons considered for each element may be found in supplementary materials. Calculations involving Co, Cr, Fe, Mn and Ni elements were carried out with spin polarization. A high energy cutoff of 800 eV was used for the plane wave basis set and a dense grid of k-points in the irreducible wedge of the Brillouin zone was generated with the Monkhorst-Pack scheme.\textsuperscript{27}

The cell volume and shape and the atomic positions were relaxed stepwise, using a quasi-Newton algorithm. The convergence criterion of $10^{-6}$ eV was chosen for the total energy. The charge distribution on the atoms was investigated using Bader topological analysis.\textsuperscript{28,29}

The energy of gaseous dihydrogen was estimated by the total energy of two hydrogen atoms in an asymmetric box $11 \times 12 \times 13 \AA^3$.

For each of the 930 hydrides $M\text{H}_y$, the total energy of formation at 0 K was systematically calculated from the total energies $E_{\text{DFT}}$ of $M\text{H}_y$ ($E_{\text{DFT}}^{\text{MH}_y}$), $\text{H}_2$ molecule ($E_{\text{DFT}}^{\text{H}_2}$) and the pure metal $M$ ($E_{\text{DFT}}^M$) in its reference state at 0 K, according to equation (1). All energies are given in kJ/mol of metal atom.

At zero pressure:

$$
\Delta H_{\text{for}}^{\text{MH}_y} = \Delta E_{\text{for,DFT}}^{\text{MH}_y} = E_{\text{DFT}}^{\text{MH}_y} - y E_{\text{DFT}}^{\text{H}_2} - E_{\text{DFT}}^M
$$

(1)

The zero-point energy (ZPE), arising from atomic displacements around equilibrium positions at 0 K, was determined in the frame of the theory of lattice dynamics, with phonon calculations in the harmonic approximation (HA). Using the frozen phonon (supercell) method\textsuperscript{30,31} the vibrational modes for the metals, hydrides and $\text{H}_2$ molecule were
computed using the phonopy code\textsuperscript{32} coupled with VASP calculations. For the force calculations, a criterion on forces of $10^{-6}$ eV/Å was used in order to eliminate any residual strain. The determined angular frequencies, were summed over all the wave vectors for each mode to obtain the ZPE given by equation within the quantum theory of harmonic crystals. The total contribution to the energy of formation (given in kJ/mol of metal atom) is the difference between the ZPE of the hydride $MH_y$ and the pure elements $H_2$ molecule and the pure metal, as:

$$\Delta E_{\text{ZPE}}^{MH_y} = E_{\text{ZPE}}^{MH_y} - \frac{y}{2}E_{\text{ZPE}}^{H_2} - E_{\text{ZPE}}^{M}. \quad (2)$$

The corrected formation energy $\Delta E_{\text{for,cor}}^{MH_y}$ at 0 K is thus defined as:

$$\Delta E_{\text{for,cor}}^{MH_y} = \Delta E_{\text{DFT}}^{MH_y} + \Delta E_{\text{ZPE}}^{MH_y}. \quad (3)$$

The hydride formation pressures at 300 K were determined using the equations of state of Joubert for gaseous hydrogen, valid up to $10^{11}$ Pa and 1000 K\textsuperscript{33} and approximating the entropy of formation by the opposite of the gas entropy. Considering the compressibility of the solid phases (hydride and metal) requires heavy calculations and it has been observed in a previous study on Ni–H\textsuperscript{34} that the contributions of the hydride and the metal tend to offset each other. Therefore, the resulting energy contribution is not predominant compared to that of the compressibility of the gas phase and was neglected in the present study.

Besides, when a hydride $MH_y$ is formed from a lower hydride $MH_x$ $(x < y)$ we considered the partial formation enthalpy from the intermediate reaction:

$$MH_x + \frac{y-x}{2}H_2 \rightleftharpoons MH_y. \quad (4)$$
Results

The results for each system are presented below, organized by columns of the periodic table. Only pertinent systems are represented in the present paper. However the figures corresponding to each system are available in supplementary materials. Besides, the calculated energies of formation and charge distributions for each system are compiled in supplementary materials.

In the present section, the DFT calculated energies of formation in kJ/mol-$M$ are represented on the figures as functions of hydrogen composition. The energies of the compounds belonging or close to the convex hull, within an enthalpy difference lower than 5kJ/mol, were corrected with the ZPE. The corrected energies of formation are represented by black filled squares on the graphs whereas the non corrected values are represented by gray open squares. Experimentally observed prototypes are underlined.

In Table 2 all the experimentally observed prototypes are listed in the second column, for each system. They can be compared to the prototypes predicted by the present calculations, listed in the third column. The experimental and calculated enthalpies of formation are specified. The detailed comparative analysis is done in the discussion section. The calculated total formation energies were compared with two types of experimental data: first, calorimetric measurements, second, formation enthalpy calculated from hydride formation/decomposition pressures measured at different temperatures, via the van’t Hoff relation, abbreviated "vH" in the following. Generally, only partial formation enthalpies of hydrides are available in the literature. Therefore, to allow comparisons with the calculated DFT total formation energies, the experimental partial enthalpies of formation were integrated, neglecting the solid solution domains. The list of experimental and calculated values available in the literature are compiled in supplementary materials.
Table 2: Experimentally observed prototypes compared with the predicted ground-state within 0.5 kJ/mol. Predicted but never observed prototypes are in italic. Negative phonon frequencies are highlighted with an asterisk. Total experimental and calculated formation enthalpies are indicated in brackets after the prototype.

<table>
<thead>
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<th>M</th>
<th>Experimentally observed prototypes (ΔE_{for exp}^{M, H})</th>
<th>Calc. ground-state (ΔE_{for cor}^{M, H})</th>
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<td>Li</td>
<td>NaCl (−89.3, 36 −90.6, 36 −97.2, 37 −99.1, 38 −94.6)</td>
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<td>NaCl (−56.9, 35 −56.4, 36 −57.0, 40 −58.4)</td>
<td>NaCl (−38.8), NaH₃*</td>
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<tr>
<td>K</td>
<td>NaCl (−63.4, 35 −57.8, 36 −57.3, 42 −59.1)</td>
<td>NaCl (−38.4), NaH₃*</td>
</tr>
<tr>
<td>Rb</td>
<td>NaCl (−54.4, 41 −44.3, 43 −47.3)</td>
<td>NaCl (−32.0), NaH₃*</td>
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<td>Be</td>
<td>BeH₂</td>
<td></td>
</tr>
<tr>
<td>Mg</td>
<td>TiO₂ (−76.2, 45 −67.0, 46 −74.4, 47 −77.0, 48 −77.4, 49 −81.2, 50 −76.8, 50 −70.0, 51 −70.7, 52), PbO₂, PdF₂</td>
<td>TiO₂ (−42.3), Na₃As*, HoD₃*</td>
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<td>CaCO₂Si (−188.7, 53 −174.3, 54)</td>
<td>CaCO₂Si (−155.9), ReB₃*</td>
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<td></td>
<td>SrCO₂Si (−199.1, 55 −188.7)</td>
<td>SrCO₂Si (−152.5), ReB₃*</td>
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* extrapolated at T_{amb} by the authors from dissociation pressure measurements carried out between 873 and 1173 K
* integrated from partial enthalpies given up to y = 1.85 - extrapolation up to y = 2.0
* integrated by the authors
* integrated from partial enthalpies given up to y = 1.75 - extrapolation up to y = 2.0
* integrated from partial enthalpies given up to y = 2.6 - extrapolation up to y = 3.0
* for y = 2.792

including the pressure work calculation
Alkaline metals

The alkaline metals Li, Na, K and Rb form with hydrogen a stable NaCl hydride, corresponding to the occupancy of the octahedral (o) sites of a face centered cubic host metallic lattice. Figure 1-left shows a typical picture of relative stabilities predicted by our DFT calculations for the alkaline metals with Na as an example. All the calculated ground-states included correctly the NaCl prototype and the dihydride formation is predicted to be very endothermic. For \( y \geq 1 \) a two-phase equilibrium is predicted involving NaCl and a trihydride called NaH\(_3\) in the present study. However, negative phonon frequencies are calculated for this latter, corresponding to a mechanical instability without pressure work.

![Graph showing enthalpy of formation vs. \( H/M \) for Na-H and Ca-H ground-states.]

Figure 1: Left: Na–H ground-state. Comparison with calorimetric data of Gunn and Green\(^{36}\) and \( vH \) data of Banus \( et \ al. \),\(^{40}\) Hérold,\(^{41}\) Messer \( et \ al. \)\(^{35}\) Right: Ca–H ground-state. Comparison with calorimetric data of Bronsted\(^{53}\) and \( vH \) data of Curtis and Chiotti.\(^{54}\)

Alkaline earth metals

While alkaline metals form a stable monohydride, the alkaline earth metals are known to form stable dihydrides with various prototypes: body-centered orthorhombic BeH\(_2\) for Be, tetragonal TiO\(_2\) for Mg in addition to orthorhombic PbO\(_2\) and cubic PdF\(_2\) at high pressure and orthorhombic Co\(_2\)Si for both Ca (Figure 1-right) and Sr. The calculated ground-states were found in agreement with these observations. Besides, the trihydride formation was
found very endothermic and the corresponding phonon spectra present negative frequencies at the equilibrium volume.

Scandium and yttrium

The Sc–H and Y–H systems exhibit very similar behaviors (Figure 2), both forming a dihydride with the CaF$_2$ fluorite type cubic structure via a very exothermic reaction.

Increasing H-composition at $y = 3$, the hexagonal HoD$_3$ prototype is observed and above $7.7 \times 10^9$ Pa, a phase transition from HoD$_3$ to a BiF$_3$-type hydride is experimentally reported.$^{90,91}$

The ScH$_{2.9}$ hydride observed by Antonov et al.$^{92}$ at high pressure in Sc–H was not included in the present study because it does not present any long range order. However the short range order is similar to YH$_3$. These features are well predicted by the DFT calculation (Figure 2). The HoD$_3$ prototype is also predicted for ScH$_3$.

![Figure 2: Y–H ground-state. Comparison with calorimetric data of Dantzer and Kleppa$^{59}$ and Akhachinskii$^{61}$ and vH data of Yannopoulos et al.$^{60}$](image)

Titanium, zirconium, hafnium

Ti, Zr and Hf form a non stoichiometric CaF$_2$ dihydride which deforms to a face-centered tetragonal ThH$_2$ dihydride as the composition approaches $M$H$_2$.$^{93,94}$ Moreover, a monohy-
Dride was identified in Ti–H and Zr–H systems with the same PtS prototype, corresponding to the occupation of the tetrahedral (te) sites on alternate (110) planes of a fcc metallic lattice.\textsuperscript{95,96} However, no monohydride has ever been observed for Hf–H. Our predicted ground-states include as expected the ThH\textsubscript{2} dihydride prototype, slightly more stable than the CaF\textsubscript{2} prototype. The PtS prototype lies just on the ground-state line for Ti, Zr and unexpectedly also for Hf although this prototype has never been observed (Figure 3). One may note that: on the one hand, for Ti–H, the calculated enthalpy of formation of the orthorhombic NbH prototype was found very close to the one of the PtS prototype, both presenting tetragonal interstitial sites and, on the other hand, for ZrH\textsubscript{2}, the enthalpy of formation of the PbO\textsubscript{2} prototype was found very close to the one of the ThH\textsubscript{2} prototype.

![Figure 3: Left: Ti–H ground-state. Comparison with calorimetric data of Luo et al.,\textsuperscript{64} Sorokin et al.,\textsuperscript{63} Dantzer,\textsuperscript{65} Stalinski and Bieganski.\textsuperscript{62} Right: Hf–H ground-state.]

**Vanadium, niobium, tantalum**

Both vanadium and tantalum form low concentrated hydrides, in the AuTe\textsubscript{2} and Ta\textsubscript{2}H\textsuperscript{97} prototypes, respectively, whereas Nb\textsubscript{2}H has never been observed. Contrary to vanadium, both niobium and tantalum form a monohydride with the same prototype NbH.\textsuperscript{98} In addition, the CaF\textsubscript{2} prototype is observed for V–H and Nb–H, however not for Ta–H. No trihydride
has been reported in the literature. The calculations reproduce well these trends (Fig 4). Besides, several features were pointed out for each system:

- for V–H, numerous monohydrides were calculated near the ground-state,
- for Nb–H, the Ta_2H prototype was found on the ground-state,
- for Ta–H, the dihydride CaF_2 is predicted at a pressure of 3.2 \times 10^9 \text{Pa}, the same prototype as VH_2 and NbH_2.

![Figure 4: Left: V–H ground-state. Comparison with calorimetric data of Luo et al., vH data of Veleckis, Kofstad and Wallace, and Griffiths et al. Right: Ta–H ground-state.](image)

Columns 6 to 9

From the 6th column, the hydride enthalpy of formation increases drastically (Figure 5). In fact, for Cr, Mo and Mn a hexagonal hydride with the NiAs prototype is experimentally observed at high pressure, corresponding to a hexagonal-closed-packed (hcp) metallic lattice (P6_3/mmc space group) with the octahedral interstitial sites being occupied. Whereas for Ru, Co and Rh, the NaCl prototype is observed in recent high pressure studies at 14 \times 10^9, 9 \times 10^9 and 4 \times 10^9 \text{Pa}, respectively. For Tc and Mn specifically the prototype CdI_2 was identified at y = 0.45 and y = 0.65 respectively corresponding to a superstructure
of NiAs. Besides, the FeH hydride has the characteristic of having a double hcp (dhcp) structure.\textsuperscript{107} High pressure studies also revealed the presence of metastable hcp and fcc hydrides whose compositions are close to FeH.\textsuperscript{108} Dihydrides were only observed for Cr and Rh, both with the prototype CaF\textsubscript{2} as the early transition metals. All the experimentally observed prototypes were found on the ground-states, except CaF\textsubscript{2} for CrH\textsubscript{2}, which is 2.1 kJ/mol less stable than PbO\textsubscript{2}. The calculations confirm all the hydrides which have been previously observed and predict that the higher hydride of Tc has the NiAs structure.

![Figure 5: Cr–H, Mo–H, Mn–H and Tc–H ground-states. Comparison with calorimetric data of Sieverts and Gotta,\textsuperscript{77} Randzio and Bojarski,\textsuperscript{78} fugacity measurements of Baranowski and Bojarski.\textsuperscript{79}](image-url)
Nickel, palladium, platinum

Experimentally, the NaCl prototype is observed for both Pd and Ni at atmospheric pressure and above $3.2 \times 10^8$ Pa respectively. Two platinum monohydrides were revealed at a pressure of $2.7 \times 10^{10}$ Pa at room temperature, one having a hcp structure, the other an unidentified complex structure. A transition from octahedral to tetrahedral sites with pressure is suggested. For both Ni–H and Pd–H systems (Figure 6) the NaCl prototype is predicted by the DFT calculations. Three other phases are competing with NaCl in the Pd–H system. However the NaCl prototype is stabilized by the ZPE correction. For $0 < y < 1$ ordered fcc superstructures are predicted in the Pd–H system but not in the Ni–H system (see the thermodynamic description of this system).

Figure 6: Left: Ni–H ground-state. Comparison with calorimetric data of Tkacz and Baranowski, Baranowski and Czarnota, Czarnota and Baranowski, vH data extracted from the plateau pressures of Tkacz, Antonov et al. Right: Pd–H ground-state. Comparison with calorimetric data of Flanagan et al. and Luo and Flanagan vH data of Wicke and Nernst, Clewley et al. and Lässer and Klatt.

Columns 11 and 12

The hydrides of Cu, Ag, Zn and Cd present only positive energies of formation. A monohydride has been observed only for Cu with the ZnO prototype with an equilibrium pressure measured at $8.4 \times 10^9$ Pa. The formation of a dihydride ZnH$_2$ whose structure is unknown
was also reported. The prototype ZnO is correctly predicted by the calculation for CuH, and also for AgH. Contrary to Cu–H and Ag–H, according to the calculation, no monohydride lies on the ground-states of the Zn–H and Cd–H systems and a two phase equilibrium between the metal and TiO$_2$ is predicted, whose enthalpy of formation is very close to the one of PbO$_2$.

**Discussion**

**General trends**

The evolution of the calculated formation enthalpy along the periodic table reproduces the observed trends and confirms the different mechanisms, leading either to increase or to decrease the hydride stability. First, the important stability of alkaline, alkaline earth up to Sc and Y hydrides is driven in particular by the negative enthalpy contribution gained from the charge transfer from the metal cation to hydrogen. The electronegativity difference between hydrogen and the metal represented indeed one of the two parameters in the empirical model of Miedema$^6$ for the hydride formation enthalpy. For the first two columns, in agreement with the experimental measurements, the predicted exothermicity of the hydride formation globally decreases with increasing period. This trend may be linked with the decrease of the electronegativity along the same column. The more electronegative the element is, the more energetically favorable the formation of a $M$–H bond is. Second, considering the early transition metals, it has been observed by Beavis$^{114}$ that, from the 3rd to the 5th column, the $4d$ element forms the most stable dihydride and trihydride. Our calculations confirmed this trend for these 3 columns, in relation with the decrease of the lattice parameter from the $4d$ to the $5d$ metals. Third, the shift towards positive enthalpies of formation with increasing atomic number has been linked with the increase of the cohesion of the metals up to the 8th column (Fe) where the $d$ bands are half filled.$^{14,16}$ Consequently, in the middle of the $d$ elements the metal total enthalpy $E_{\text{DFT}}^M$ reaches a minimum and the cost of the
lattice expansion for hydrogen insertion a maximum. Hence, the interatomic distances tend to decrease along the d-metals up to the 8th column. According to the calculations of Miwa and Fukumoto\textsuperscript{16} this is a dominant factor, leading to a decrease of the negative chemical enthalpy contribution arising from the $M$–$H$ bonds formation, irrespective of the nature of the metal.

**Ground-state structure prediction**

The present methodology was able to predict correctly all the experimentally observed hydrides, except for Pt. In addition, new metastable hydrides and high pressure hydrides were found on the ground-state at 0 K. The predicted prototypes are summarized in Table 2, compared with the experimental ground-states. Key results are discussed below, starting with the early transition metals. The monohydrides HfH and YH were found stable with the prototypes NbH and PtS, whose formation energies were equal to within 0.5 kJ/mol. These two prototypes have not been observed for Hf and Y yet but they are stable for other early transition metals: the PtS prototype for Ti\textsuperscript{95,96} and Zr and the NbH prototype is well known for Nb and Ta and was also observed for Ti. Interestingly, a hcp monohydride VH was found stable at 0 K, in competition with other structures. This result implies that a monohydride VH may be easily stabilized by adding chemical defects. Moreover, like all the other early transition metals, a dihydride TaH$_2$ is predicted with the CaF$_2$ prototype at $3.2 \times 10^9$ Pa.

The columns 6 to 12 of the periodic table are now discussed, starting with the $M$H stoichiometry. The o-site NiAs prototype, observed for Cr, Mo and Mn, is correctly predicted by the calculations. In agreement with this trend, the same prototype is also predicted for Tc at $2.1 \times 10^8$ Pa, in accordance with the prevalence of o-sites for monohydrides after the 6th column. One may note that the prototype CdI$_2$ observed for Tc–H is also predicted for Mn–H, with a formation enthalpy close to that of AuTe$_2$, although it has never been observed. The 8th column appeared as a transition from hexagonal towards cubic monohydrides. For FeH, the metastable NiAs prototype was found 1.4 kJ/mol less stable than the dhcp FeH.
prototype, in agreement with the observation of Antonov *et al.*

In contrast with the two previous columns, RuH is predicted with the NaCl prototype at $5.9 \times 10^9$ Pa, in agreement with the recent high pressure study of Kuzovnikov and Tkacz and with the calculations of Liu *et al.* The calculated equilibrium pressure was found close to the measured decomposition pressure of $6.2 \times 10^9$ Pa. The NaCl prototype is also predicted for CoH, RhH, NiH, and PdH in agreement with the experimental observations. Besides, the lower composition hydrides predicted for Pd–H for $0 < y < 1$ correspond to different orderings of hydrogen atoms on the interstitial sites of the fcc host lattice. Two of them are indeed experimentally observed at low temperature. Interestingly, no monohydride is predicted for the Pt–H system, contrary to the experimental observations at high pressure. The error is probably due to the pressure work which may considerably affect the relative stabilities at the considered pressures. A complementary study is necessary to include this effect.

Besides, the hydride AgH is predicted with the prototype ZnO, which was experimentally observed for CuH. From the 6th to 12th column, only two dihydrides were experimentally observed: CrH$_2$ and RhH$_2$, both having the CaF$_2$ prototype. They are correctly predicted by the calculations. Furthermore, a competition between the CaF$_2$ and PbO$_2$ prototypes is predicted for the 3d metal dihydrides and between the Co$_2$Si and CaF$_2$ prototypes for the 4d, the TiO$_2$ prototype appearing only for the 12th column, as summarized in Table 3.

Table 3: Dihydride structure prototypes predicted on the ground-state for each 3d and 4d metal from the 6th to the 12th column, associated with the predicted formation pressure.

<table>
<thead>
<tr>
<th></th>
<th>Cr</th>
<th>Mn</th>
<th>Fe</th>
<th>Co</th>
<th>Ni</th>
<th>Cu</th>
<th>Zn</th>
</tr>
</thead>
<tbody>
<tr>
<td>PbO$_2$</td>
<td>$1.0 \times 10^{10}$ Pa</td>
<td>$5.3 \times 10^{9}$ Pa</td>
<td>$9.8 \times 10^{9}$ Pa</td>
<td>$9.3 \times 10^{9}$ Pa</td>
<td>$8.7 \times 10^{9}$ Pa</td>
<td>$2.0 \times 10^{10}$ Pa</td>
<td>$8.8 \times 10^{9}$ Pa</td>
</tr>
<tr>
<td>Mo</td>
<td>Tc</td>
<td>Ru</td>
<td>Rh</td>
<td>Pd</td>
<td>Ag</td>
<td>Cd</td>
<td></td>
</tr>
<tr>
<td>Co$_2$Si</td>
<td>$9.0 \times 10^{9}$ Pa</td>
<td>$2.4 \times 10^{10}$ Pa</td>
<td>$1.5 \times 10^{10}$ Pa</td>
<td>$2.1 \times 10^{9}$ Pa</td>
<td>$6.6 \times 10^{9}$ Pa</td>
<td>$2.8 \times 10^{10}$ Pa</td>
<td>$1.2 \times 10^{10}$ Pa</td>
</tr>
</tbody>
</table>

Within the set of metals included in the study, trihydrides are experimentally observed only for Y and Sc with the HoD$_3$ prototype – without long range order for ScH$_3$. As
expected, trihydrides of column 1 and 2 were predicted highly unstable. The prototype \( \text{HoD}_3 \) is also predicted without negative phonon frequencies for \( \text{NbH}_3 \), \( \text{TaH}_3 \), \( \text{TiH}_3 \) and \( \text{ZrH}_3 \) at \( 1.3 \times 10^{10} \text{ Pa} \), \( 1.3 \times 10^{10} \text{ Pa} \), \( 8.3 \times 10^9 \text{ Pa} \) and \( 3.3 \times 10^9 \text{ Pa} \), respectively.

**Quantitative agreement**

The ZPE correction \( \Delta E_{\text{ZPE}}^{M\text{H}_y} \) was found to be non negligible. For te-sites, its value was always found positive, destabilizing, reaching 22 kJ/mol for \( \text{VH}_2 \) whereas for octahedral and other sites it may be either positive or negative. Neglecting the ZPE contribution leads thus to overestimate the stability of te-site prototypes, as it will be detailed in a future publication. For instance, the ZPE contribution is responsible for the stabilization of the NaCl PdH hydride against 3 tetrahedral hydrides PtS, NbH and ZnO, predicted as more stable by the raw DFT calculations. Again, for the Rh–H system, the raw DFT calculations predict the CaF\(_2\) hydride as the first one to be formed. As the NaCl hydride is stabilized by the ZPE contribution whereas the CaF\(_2\) is destabilized, the correction brings the ground-state in agreement with the experimental observations.\(^\text{305}\)

Comparing with the available experimental data, the magnitude of the hydride formation enthalpies seems to be globally underestimated by the calculations, in particular for alkaline and alkaline earth hydrides. The difference reaches about 20 kJ/mol for KH and softens from the left to the right of the periodic table. This difference could originate from the fact that experiments are carried out at non zero temperatures whereas the calculations are performed at 0 K. Indeed, because the reference states of the metals and hydrides (solid) and the \( \text{H}_2 \) molecule (gaseous) are different, the Kopp-Neumann rule cannot be applied: the hydride heat capacity cannot be estimated by an average between that of gaseous dihydrogen and of the corresponding pure metal. As a consequence, the variation of the formation enthalpy \( \Delta H_{\text{for}}^{M\text{H}_y} \) with the temperature should be larger than for reactions involving only metals. The difference between experimental data at non zero temperature and calculations at 0 K is thus expected to be larger. To make meaningful comparisons, extrapolations should be
carried out on the basis of heat capacity values, requesting an additional study. Besides, some experiments are carried out under high hydrogen pressure whereas the calculations are performed without mechanical pressure. Therefore, without corrections, the high pressure calorimetric data tend to be shifted towards positive values, because of the neglected pressure works affecting especially the gas phase. In order to compare with the DFT, the calorimetric values have to be corrected, as it has been done for nickel.\textsuperscript{80} When the volumes of the condensed phases are affected, more complex models are needed, which involve the corresponding volume variations with $T$ and $P$ (see e.g. the treatment done in [34]).

**Conclusion**

For each investigated $M$–$H$ system, the experimentally observed prototypes were correctly predicted on the ground-state. Besides, the calculated total hydride formation enthalpies at $0$ K, were found consistent with the available experimental values. The temperature variation of the enthalpies represents a point to be investigated in future studies.

Our systematic approach is complementary to both the experimental methods or evolutionary algorithms,\textsuperscript{118} as the latter predicts the stable crystal structures. Whereas the present database includes the metastable hydrides, giving an exhaustive picture of the relative stabilities between the different considered crystal structures. Trends which are only outlined by experimental results can be confirmed by broadening the set of hydrides. Thus, prototypes which are close to be stable at $0$ K can be highlighted, as the monohydride VH, as well as higher hydrides TaH$_2$ or ZrH$_3$ whose formation pressure can be estimated. Experiments are needed to confirm these predictions.

The obtained results represent useful input data to supply thermodynamic database usable in the frame of prediction algorithms, like the machine learning.\textsuperscript{119} They can also be used in thermodynamic modeling, such as the Calphad approach,\textsuperscript{120} allowing to easily combine the binary systems to obtain predictions in the ternary and higher systems. Such
outlooks represent promising tools for the alloy design of hydrogen storage alloys or to improve the alloy durability towards hydrogen embrittlement.

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Supporting Information Available

The following files are available free of charge.

- 1-Crystal-structures: information and illustration of the considered hydride crystal structures including hydrogen composition, prototype, Pearson symbol, space group and Wyckoff positions
- 2-Formation-enthalpies-per-system_graphs: graphical representation of the calculated formation enthalpies as function of hydrogen composition for each considered system, ground-states
- 3-Formation-enthalpies-per-system_values: calculated formation enthalpies, charge distributions and magnetic moments of the 30 crystal prototypes for each of the 31 considered systems
- 4-Literature-review: experimental and calculated hydride formation enthalpies available in the literature

References


(21) Reshak, A. MgH$_2$ and LiH metal hydrides crystals as novel hydrogen storage material: Electronic structure and optical properties. *Int. J. Hydrogen Energy* 2013, 38, 11946–11954.

(22) Reshak, A. NaMgH$_3$ a perovskite-type hydride as advanced hydrogen storage systems: Electronic structure features. *Int. J. Hydrogen Energy* 2015, 40, 16383–16390.


(39) Simandl, R. F. *Equilibrium-pressure isotherms for the system lithium-lithium hydride-hydrogen at 823, 873 and 923 K*; 1981.


(52) Selvam, P. Studies on the thermal characteristics of hydrides of Mg, Mg$_2$Ni, Mg$_2$Cu and Mg$_2$Ni$_{1-x}$M$_x$ (M = Fe, Co, Cu or Zn; 0 < x < 1) alloys. *Int. J. Hydrogen Energy* **1988**, *13*, 87–94.


(66) Rundle, R.; Wilson, A.; Nottorf, R.; Raeuchle, R. *The crystal structures of ThH₂ and ZrH₂;* 1948.


(80) Tkacz, M.; Baranowski, B. Heats of formation and decomposition of nickel and Ni$_{0.8}$Cu$_{0.2}$ hydrides measured in high pressures of gaseous hydrogen. J. Less-Common Met. 1985, 113, 83–87.


High-throughput DFT and phonon calculations were systematically performed for 31 binary metal-hydrogen systems distributed for a set of 30 potential crystal structures. The calculated enthalpies of formation were represented as functions of H composition, \( y \). The sequence of stable hydrides \( \text{M}_y \text{H}_x \) (ground-state) given by the convex hull was found in agreement with the experimental observations. Besides, predictions could be made on metastable and high pressure hydrides, providing orientations for new materials synthesis.