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Organometallic Mediated Radical Polymerization of Vinylidene Fluoride

Sanjib Banerjee,[a,b]# Vincent Ladmiral,[a] Antoine Debuigne,[b] Christophe Detrembleur,[b] Rinaldo Poli*[c,d] and Bruno Améduri*[a]

Abstract: An unprecedented level of control for the radical polymerization of vinylidene fluoride (VDF), yielding well-defined PVDF (up to 14,500 g/mol) with low dispersity (1±3.2), was achieved using organometallic mediated radical polymerization (OMRP) with an organocobalt compound as initiator. The high chain-end fidelity was demonstrated by the synthesis of PVDF- and PVAc-containing di- and tri-block copolymers. DFT calculations rationalize the efficient reactivation of both head and tail chain end dormant species.

Poly(vinylidene fluoride), PVDF, is the second largest commercially available fluoropolymer after polytetrafluoroethylene (PTFE). It exhibits remarkable properties such as hydrophobic and oleophobic properties, chemical resistance to acids and solvents, low dielectric constants and excellent weathering.[1] Hence, it finds uses in many high-tech applications in aerospace and aeronautical engineering, coatings,[2] piezoferroelectric devices,[3] backsheets for photovoltaic devices and functional membranes for water treatment.[4] VDF has a reactivity close to those of TFE[5] and chlorotrifluoroethylene,[6] but it is not explosive and much less toxic.[7]

Reversible deactivation radical polymerization (RDRP) of VDF has only been achieved using two techniques: (i) iodine transfer polymerization (ITP)[1c, 5] and (ii) reversible addition fragmentation chain transfer (RAFT) polymerization.[6] Both techniques require the use of a radical initiator and a suitable chain transfer agent (CTA), which ensures the reversible degradative transfer (DT) at the heart of ITP and RAFT. Both were shown to produce relatively well-defined diblock[1c, 5a, 5c, 6c, 7] and triblock[1c, 5d] copolymers. However, the chain defects resulting from head-to-head (HH, -CH₂CF₂-CH₂F-) and tail-to-tail (TT, -CF₂CH₂F-) monomer additions impose limits on the molar masses (Mn) and dispersities (D) attainable with these techniques, as demonstrated in both ITP[5c] and RAFT[6c] polymerization. Upon transfer to the CTA, the HH additions generate less reactive species that lead to a slowdown of the degenerative chain transfer, a broadening of the molar mass distribution and a degradation of the control.[6c, 6d, 8] Ideally, a RDRP technique should reactivate both PVDF chain ends at the same rate. Asandei et al.[9] were able to reactivate the -CF₂CH₂I terminated PVDF chains and to synthesize block copolymers using Mn₂(CO)₁₀ and a photoradiation process in ITP. However, this was only possible using conventional radical polymerization. Liepins et al. reported the synthesis of PVDF with a low content of HH chain defects using a modified Ziegler-Natta Catalyst, while Chung’s group[11] used a trialkylborane/oxygen mixture as the initiating system, resulting in significant chain defects. However, the controlled nature of these polymerizations was not addressed. In terms of the coordination-insertion strategy for fluorinated alkenes, Jordan et al.[12] described the copolymerization of vinyl fluoride and ethylene using (phosphine-arenesulfonate)Pd(Me)(py) and phosphinebis(arenesulfonate)PdMe(py) catalysts, while Rieger and coworkers reported the copolymerization of 3,3,3-trifluoropropene with ethylene using (phosphine-sulfonate)Pd(Me)(DMSO) as catalyst.[13] However, these catalysts did not provide a very good control and, in addition, were not applied to the VDF homopolymerization or VDF-olefin copolymerization.

Among other RDRP techniques, organometallic-mediated radical polymerization (OMRP), in which a transition-metal complex reversibly traps the propagating polymer radicals,[14] has been successfully employed for the radical (co)polymerization of less reactive monomers, including vinyl acetate (VAc)[14c-d] and VAc-ethylene,[15] with Co(C₅H₅)₂(acac) = acetylacetone as controlling agent. However, there is no report of the application of this technique to homopolymerize fluorinated alkenes. We report here the first OMRP of a fluorinated monomer (VDF) under mild experimental conditions, leading to unprecedented control over the homopolymerization of this monomer (Scheme 1). Furthermore, macromolecular engineering was achieved using the in situ generated PVDF-Co(acac)₂ macrorinitiator to prepare well-defined PVAc and PVDF-containing diblock- and triblock copolymers (Scheme S1).

VDF homopolymerizations were carried out in dimethyl carbonate (DMC, a solvent with a low aptitude to transfer H atoms to PVDF radicals and which swells PVDF[15] in a Hastelloy (HC 276) autoclave at 40 °C using [R₂(VAc)-Co]₅ as initiator (R₀ = primary radical generated by V-70)),[15] abbreviated as R-Co[3], at three [VDF]₀/[R-Co]₀ initial molar ratios (100, 200 and 300). Notably, no polymerization was observed when V-70/Co/acac₂ was used at 40 °C or even at 50 °C. This shows that azo initiators are not able to initiate the polymerization of VDF[15] and justifies the choice of R-Co as initiating species. The selected temperature was the mildest one, allowing to attain a reasonable control. These polymerizations produced PVDFs with higher targeted molar masses (Mn = 5,900 to 14,500 g/mol) and lower dispersities (D < 1.32) (Table S1 and Figure S1), than those achieved by ITP or RAFT under the same conditions.

¹H, ¹³C and ¹⁹F NMR enabled the structural analysis of the synthesized PVDFs, including the chain ends (see details in S1).
The recovered polymers, isolated after quenching with TEMPO (see SI, Figures S2-S4, P2, Table S1), showed the expected features (normal Head-to-Tail and reversed Head-to-Head and Tail-to-Tail monomer additions) of radically polymerized VDF, the -CF<sub>3</sub>H and -CH<sub>2</sub> end-groups originating from transfer reactions to solvent, the characteristic signals of the oligo(VAc) group from the R-Co initiator (1H NMR, Figure S2), and the TEMPO chain end ([<sup>19</sup>F NMR, Figure 2B and Figure S3]). Notably, only the R<sub>2</sub> (VAc)<sup>-</sup>CH<sub>2</sub>CF<sub>2</sub>- and not the R<sub>1</sub> (VAc)<sup>-</sup>CF<sub>2</sub>CH<sub>2</sub>- chain end could be identified, whereas both -CF<sub>3</sub>CH<sub>2</sub>CH=CFCH=CF<sub>2</sub>- TEMPO and -CH<sub>2</sub>CF<sub>2</sub>CH=CFCH=CF<sub>2</sub>- TEMPO ω-chain ends were visible at -66.5 and -61.5 ppm, respectively. The SI offers a more detailed discussion of these assignments.

Meanwhile, the proportions of intrachain TT and HH reverse additions (equations S4–S5), which remain close to each other as expected<sup>[16]</sup>, increase gradually with VDF conversion (Figure S9), reaching a plateau at ca. 4%. The polymerization pursual at approximately the same rate after stagnation of the terminal HH and TT inversions demonstrates that both types of dormant chains, PVDF-CH<sub>2</sub>CF<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub> and PVDF-CF<sub>2</sub>CH<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub>, are reactivated. Based on this result, we propose that the polymerization proceeds by a typical reversible termination mechanism with Co<sup>III</sup>(acac)<sub>2</sub> as the radical trapping species (Scheme 1).

The aptitude of both dormant species to reactivate may be evaluated from the results of a recent DFT study of the Co-C homolytic bond dissociation enthalpies (BDEs) in various (CH<sub>2</sub>CF<sub>2</sub>)<sub>3</sub>Co(acac)<sub>2</sub> molecules, with all possible F substitutions at the α and β positions.<sup>[19]</sup> The CH<sub>2</sub>CF<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub> and CHF<sub>2</sub>CH=Co<sup>III</sup>(acac)<sub>2</sub> compounds, which may be considered as models of the head and tail dormant chains, yield BDE<sub>H</sub> = 27.4 kcal/mol and BDE<sub>T</sub> = 26.0 kcal/mol, respectively, for a BDE difference of 1.4 kcal/mol, which is very small. Even more strikingly, the tail chain-end yields a weaker bond than the head chain-end, predicting an easier reactivation of the tail dormant species. In the case of the CH<sub>2</sub>CF<sub>2</sub>-X and CHF<sub>2</sub>CH=CH<sub>2</sub>-X models of the 'head' and 'tail' dormant species in ITP (X = I) and RAFT (X = SMe<sub>2</sub>O), on the other hand, the BDE is greater for the latter [ABDE = BDE<sub>T</sub> - BDE<sub>H</sub> = 4.2 kcal/mol for I or 6.2 kcal/mol for SC(S)OMe]<sup>[18, 19]</sup>. In order to further improve the model, the chain simplification was implemented only at the level of the C atom. Thus, PVDF-CF<sub>2</sub>CH<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub> was modeled by CHF<sub>2</sub>CF<sub>2</sub>-CH=Co<sup>III</sup>(acac)<sub>2</sub>, for which the calculated BDE is 25.2 kcal/mol, and PVDF-CF<sub>2</sub>CH<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub> was modeled by CHF<sub>2</sub>CF<sub>2</sub>-CH=Co<sup>III</sup>(acac)<sub>2</sub> (product of a regular HT addition, BDE = 27.9 kcal/mol) and CHF<sub>2</sub>CF<sub>2</sub>-Co<sup>III</sup>(acac)<sub>2</sub> (product of a TT addition, BDE = 27.4 kcal/mol). The computational details are given in the SI. Qualitatively, the bond strengths remain in the order PVDF-Co<sup>III</sup>(acac)<sub>2</sub> < PVDF<sub>r</sub>-Co<sup>III</sup>(acac)<sub>2</sub>, predicting no accumulation of the PVDF-Co<sup>III</sup>(acac)<sub>2</sub> medium (contrary to RAFT and ITP). In reality, the observation of a greater proportion of PVDF<sub>r</sub>-Co<sup>III</sup>(acac)<sub>2</sub> dormant chains (Figure S8)
relative to the in-chain errors means that the experimental bond strengths must be in the opposite order (PVDF-Co\((acac)\)_2 > PVDF-Co\((acac)\)_2), but the difference must be much smaller than in the corresponding ITP and RAFT dormant species.

Figure 2. A) SEC traces of the PVDF prepared by Co(acac)_2-mediated OMRP (trace a, after 24 h), of the PVDF-b-PVAc diblock copolymer (trace b, after 24+24 h) obtained by chain extension with VAc, and of the PVDF-b-PVAc-b-PVDF triblock copolymer (trace c, after 24+24+16 h) prepared by chain coupling of the previous diblock copolymer using isoprene. B) 19F NMR spectra of PVDF (bottom), of the PVDF-b-PVAc diblock copolymer (middle), and of the PVDF-b-PVAc-b-PVDF triblock copolymer (top).

In addition to improved control in VDF homopolymerization, the labile carbon–metal bond in PVDF-Co\((acac)\)_2 allowed reactivation for chain extension with VAc, leading to PVDF-b-PVAC diblock copolymers and increasing the scope of available PVDF-containing block copolymers. The SEC chromatograms (Figure 2A) show a shift of the PVDF-Co\((acac)\)_2 distribution (trace a, M_n = 4,100 g/mol, D = 1.27) toward higher molar masses upon chain extension to form the PVDF-b-PVAc diblock, while maintaining low D values (trace b, M_n = 10,300 g/mol, D = 1.28). As expected, the RI-SEC chromatograms of the PVDFs were negative (fluoropolymers have low refractive index).\(^{[16, 66, 21]}\) 1H and 19F NMR spectra (Figures S10 and 2B, respectively) show the characteristic PVAc –CH(OAc)– signal at 4.8 ppm and the VDF=VAc dyad –CF2= signal at -93.3 ppm, confirming the product structure.\(^{[66]}\)

Co\((acac)\)_2-terminated polymers prepared by OMRP are known to undergo rapid chain–chain coupling upon addition of a conjugated diene, affording double-molecular weight terminated products\(^{[22]}\) and allowing the facile synthesis of symmetric ABA triblock copolymers when starting from Co\((acac)\)_2-terminated AB diblock copolymers.\(^{[22]}\) Application of this technique to PVDF-b-PVAc-Co\((acac)\)_2 using isoprene as coupling agent yielded a symmetrical PVDF-b-PVAc-b-PVDF triblock copolymer (Scheme S1). The SEC chromatograms revealed that the M_n of the coupled product approximately doubled (19,900 g/mol vs 10,300 g/mol) while the dispersity remained low (Table S3).

As expected, the thermal stability of the produced PVDFs increases with increasing M_n (P1-P3, Table S1), as shown by the thermogravimetric analyses (TGA) (Figure S11). The melting points \(T_m\) of these samples, determined by differential scanning calorimetry, were similar (ca. 163 °C) (Figure S12). Expectedly, the degree of crystallinity increased from 34 to 54% (calculated using equation S3) with increasing M_n.

The thermal stabilities (TGA, Figure S13) of the PVAc-containing di-/tri-block copolymers were close to those of the PVDF homopolymers as evidenced by their \(T_{d,10\%}\) decomposition temperatures (Table S3). However, the di-and tri- block copolymer displayed a \(T_g\) characteristic of the PVAc block (at ca. 35 °C)\(^{[23]}\) and a reduced degree of crystallinity (<8%) relative to neat PVDF (Table S3, Figures S15-S16).

In conclusion, this communication reports the first example of an organometallic-mediated radical polymerization (OMRP) of VDF and demonstrates an unprecedented level of control for this monomer, including the first synthesis of a block copolymer with successful transition to a second RDRP process. This was possible thanks to the facile reactivation of the dormant species formed after an inverted monomer addition, PVDF-X, when X = Co\((acac)\)_2, whereas the reactivation of these chains is inefficient in ITP (X = I) and RAFT (X = xanthate). Due to the remarkable properties of the fluorinated groups (low dielectric constant and interesting thermal, electroactive and surface properties), the resulting copolymers might find applications in high value added materials (e.g., coatings, binders for lithium ion batteries, piezoelectric devices, and membranes).

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Keywords: alkyl cobalt species • block copolymer • organo-metallic mediated radical polymerization • one pot synthesis • vinylidene fluoride


Towards new possibilities: An unprecedented organometallic mediated radical polymerization (OMRP) of VDF yielded well-defined PVDF (at least up to 14,500 g/mol) with low dispersity (≤1.32). The experimental findings show that both head and chain ends exhibits reactivation and these experimental findings are supported by DFT calculations.