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Thermochromic Luminescent Materials and Multi-Emission Bands in d^{10} Clusters

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Thermochromic luminescent compounds with d^{10} metal ions are interesting materials for applications such as sensors or display devices. However, these properties are difficult to predict prior to their synthesis. In this communication, we investigated materials with structural assemblies known to be responsible of distinct luminescence mechanisms and show that they can be interesting potential thermometers. Thus, we compared the synthesis of a zinc halide and a copper halide based compounds which only differ in their ability to create clusters with metallophilic interactions. The compounds synthesized by hydrothermal method have been structurally characterized by Single-crystal X-ray diffraction, Solid-State NMR, FTIR, UV-Visible spectroscopy, thermal analysis and EPR. The photoluminescence properties of the two materials have been characterized at different temperatures. The copper bromide compound shows luminescence thermochromism in a wide spectrum of colors owing to the formation of clusters generating multi-emission bands while the zinc bromide exhibits a single emission band and no thermochromism.

Materials exhibiting a controlled photoluminescence as a response to temperature have great promise for applications as luminescent thermometers^{1–6}. Thus, in the past decade, much effort has been focused on the synthesis of such functional materials. Irrespective of their mechanism, these phenomena can be observed by variations of color or/and intensity of the luminescence. The change in photoluminescence intensity vs. temperature is common because of the competition between non-radiative and radiative decays⁷. The color changes in a wide-range is however less common. This modification of color can either originate from the important shift of an emission band or the presence of two or several emission bands whose intensities vary differently with temperature.

Among the thermochromic luminescent materials, the ones built with transition-metal complexes are of particular interest^{8–12}. The analysis of the ordered molecular assemblies allows understanding and rationalizing the mechanisms that are involved in these phenomena. For example, compounds with d^{10} metals can exhibit luminescence which can be either of ligand centered, charge transfer or metal-centered nature. When these compounds are of cluster-type, the emitting states will depend on the nature of the polynuclear assemblies (depending on the number of transition metals and of the metal-metal interactions). Thus, the bonding interactions between metal cations with a closed-shell configuration have been extensively investigated owing to their roles in some thermochromic properties^{13–17}.

As the thermochromic behaviour of luminescence can be of different origins, a rational design of materials exhibiting such properties is difficult. In this communication, we show that one strategy is to focus on compounds in which several photoluminescence bands have different origins. Thus, if different photoemissions arising from distinct mechanisms occur in a single-phase, the responses (especially the intensities) to temperature likely differ. The resulting luminescence color will, thus, be changed. As an illustration, we synthesized two d^{10} metal halides (the new compound $[\text{C}_6\text{H}_{16}\text{N}_2]_3[\text{Cu}_4\text{Br}_6][\text{Cu}_2\text{Br}_6]$ (**1**) and the previously reported $[\text{C}_6\text{H}_{16}\text{N}_2]\text{ZnBr}_4$ (**2**)) with the same ammonium cation and studied their responses with the change of temperature (Fig. 1)¹⁸.

The d^{10} metal halide compounds are built of anionic $[\text{MX}_n]^{z-}$ inorganic units, or clusters isolated between the templating molecules of 1,4-dimethylpiperazine-1,4-dium. The new copper bromide crystallizes in the non-centrosymmetric space-group $P2_12_12_1$, and is built from two clusters: a dinuclear $[\text{Cu}_2\text{Br}_6]^{4-}$ and a tetranuclear $[\text{Cu}_4\text{Br}_6]^{2-}$ (Fig. 2). For this compound, the Cu-Br bond lengths range between $2.406(3) \text{ \AA} < d < 2.553(3) \text{ \AA}$ in $[\text{Cu}_2\text{Br}_6]^{4-}$ clusters and $2.358(3) \text{ \AA} < d < 2.478(3) \text{ \AA}$ for $[\text{Cu}_4\text{Br}_6]^{2-}$ clusters. The Cu-Cu distance is $2.686(3) \text{ \AA}$

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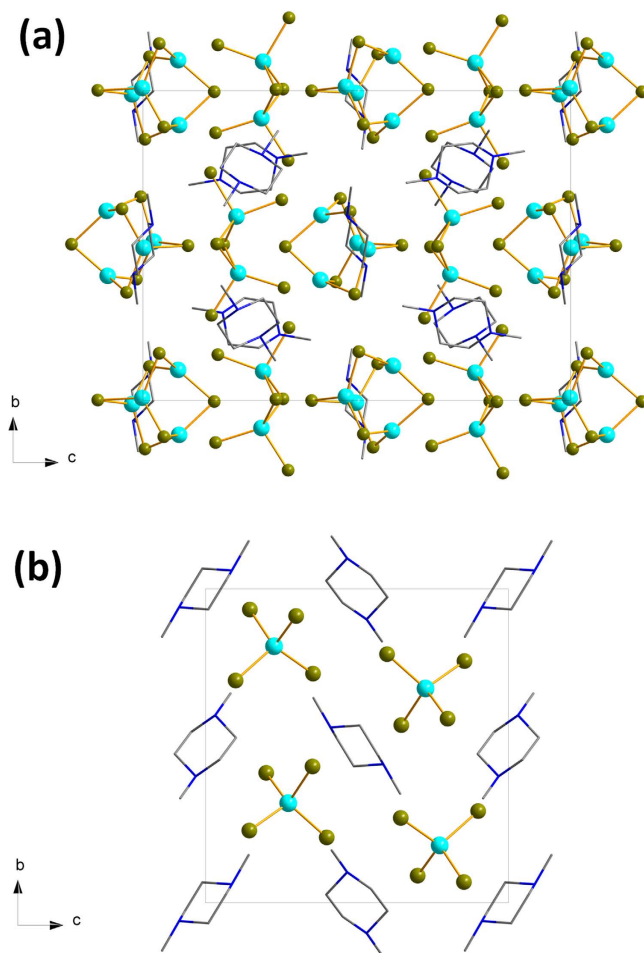


Figure 1. Representation of the two d^{10} metal halides (a) $[\text{C}_6\text{H}_{16}\text{N}_2]_3[\text{Cu}_4\text{Br}_6][\text{Cu}_2\text{Br}_6]$ (1) and (b) $[\text{C}_6\text{H}_{16}\text{N}_2]\text{ZnBr}_4$ (2).

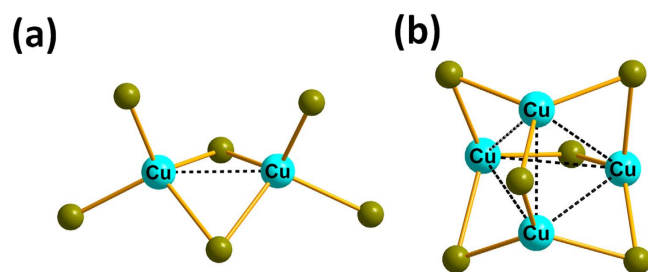


Figure 2. Representation of the copper bromide structures: (a) View of the copper dinuclear $[\text{Cu}_2\text{Br}_6]^{4-}$ and, (b) the copper tetranuclear $[\text{Cu}_4\text{Br}_6]^{2-}$.

in $[\text{Cu}_2\text{Br}_6]^{4-}$ cluster and $2.667(3) \text{ \AA} < d < 3.089(3) \text{ \AA}$ in $[\text{Cu}_4\text{Br}_6]^{2-}$ cluster. The crystal structure determined at 100 K shows that the molecular packing is little affected by temperature. At low temperature, a free water molecule with a large thermal agitation can also be observed. The previously reported zinc bromide crystallizes in the space-group $P2_1/c$ and is built from tetrahedral $[\text{ZnBr}_4]^{2-}$ units¹⁸. FTIR, UV-Visible spectroscopy, and thermal analysis were performed on both compounds. To investigate the origin of the large anisotropic displacement parameters on the atoms of the organic molecules, $\{^1\text{H}\}$ - ^{13}C CP-MAS Solid-state NMR has been performed for compound 1 (Supporting Information). The EPR spectroscopy confirmed the absence of Cu(II) in the copper bromide sample. Ground State (GS) and triplet Excited State (ES) computations have been performed on the $[\text{Cu}_4\text{Br}_6]^{2-}$ moiety (computations details are provided in Supporting Information). The optimized GS geometry enforcing the T_d symmetry is in good agreement with experimental data (Cu-Cu = 2.749 Å; Cu-Br = 2.456 Å). In this geometry, each Cu atom is in a trigonal planar configuration bonded to three Br atoms. On the other hand, the halides are linked to two metals.

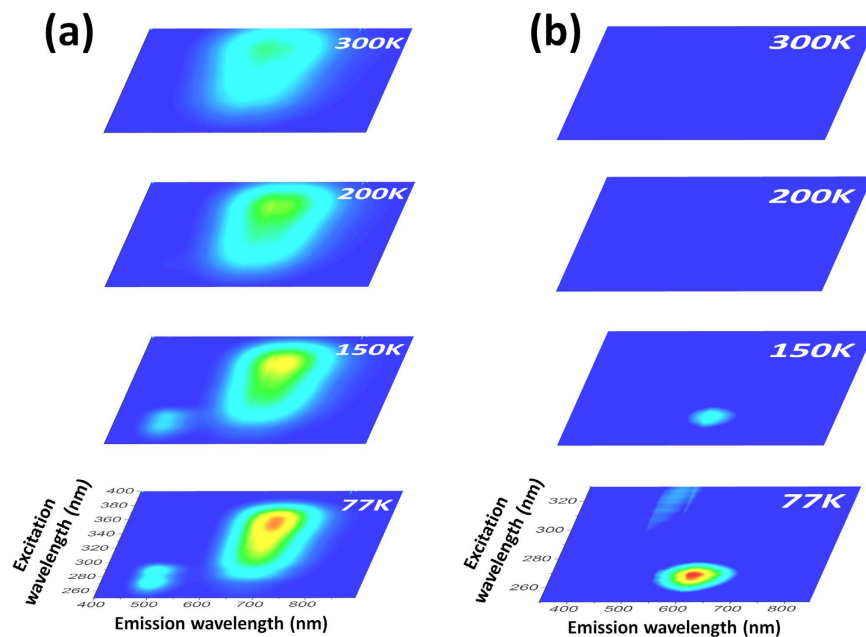


Figure 3. Solid-state photoluminescence (excitation vs. emission) for (a) compound **1** and (b) compound **2** in function of temperature.

Experimentally, the absorption spectrum exhibits a very sharp and intense peak around 370 nm (Supporting Information). This information is well reproduced in the calculations since the absorption has been computed using TD-DFT at 366 nm ($f=0.06$, T_2). Two electronic transitions are computed around 300 nm (305 nm, $f=0.02$, T_2 ; 297 nm, $f=0.03$, T_2), constituting the second band on Figure S1. Finally, an intense peak around 260 nm is constituted of multiple transitions, together with a very intense one computed at 264 nm ($f=0.07$).

The steady-state photoluminescence properties of the two compounds have been characterized at different temperatures (Fig. 3). Mapping of the excitation vs. emission allowed identifying the different photoemissions and following their evolutions with temperature. At 77 K, compound **1** shows two emission bands at high energy (about 500 nm) and lower energy (about 700 nm) (Fig. 3(a)). The relative ratio of these two emissions differ for different temperatures. The high energy emission disappears at room temperature. A blueshift of the low energy emission is also observed in function of the temperature. Its maximum is located at 650 nm at 300 K. The luminescence properties of copper halide clusters have been extensively studied. The low energy emission is well-known for cluster Cu_4 and has been previously attributed to be of cluster centered nature with a combination of halide to metal charge transfer and copper-centered transition¹³. For the zinc bromide, a weak photoemission at 620 nm was observed at low temperature (Fig. 3(b)) but this compound do not show any photoluminescence properties at 300 K.

Time-resolved photoluminescence has been performed on compound **1** (Fig. 4) in order to confirm the presence of triplet excited states (ES) prior to their computation. Figure 4(a) shows the luminescence on the whole visible area. We clearly observed a strong luminescence centered at about 620 nm in the 100 μs time-range and a second band at 490 nm. Photoluminescence decays for the two bands have been averaged based on the red square and on the green square, respectively. The lifetime of the low energy and the high energy emissions are 21 μs and 0.28 μs , respectively. These lifetimes were obtained by fitting the experimental data with a monoexponential decay convoluted with the laser pulse. As in steady-state PL, the time-resolved PL spectrum integrated over the entire 200 μs window at 77 K (Figure S7) exhibits an intense emission band located at 497 nm compared to the intensity of the band located at 657 nm. As in 300 K, the PL decays monoexponentially for the two bands but their lifetimes increase (47.5 μs and 37.7 μs for the bands at 497 nm and 657 nm respectively). This trend, with comparable lifetime values, has been well observed in literature¹⁹, and one key to understand this behaviour is probably the temperature-induced rigidity of those systems²⁰.

Optimizations of the ES on the Cu_4 clusters have been performed without symmetry constraint and have been checked to be true minima on the potential energy surface. As a matter of fact they allow an investigation of the origin of the two photoemission bands. Indeed, the computed emission wavelengths (phosphorescence) were around 490 nm and 850 nm. The computed one of lowest energy slightly overestimates the experimental one whereas the second one fits with the observed photoemission. For the lowest energy triplet ES, the metallic tetrahedron is opened and one Cu-Cu interaction is weakened (3.91 Å). One Br atom becomes singly bonded to a Cu with a short Cu-Br distance (2.38 Å). The other Cu-Cu distances are also dramatically affected with a strong shortening from GS to ES ($\approx 10\%$, 2.51–2.66 Å). In the second ES, the tetrahedron is maintained with six short metal-metal contacts (2.51 to 2.73 Å). In order to assess the strength of the interaction between copper atoms, a coupled NAO (Natural Atomic Orbital) and Wiberg indices analysis has been performed²¹. As a matter of fact, it

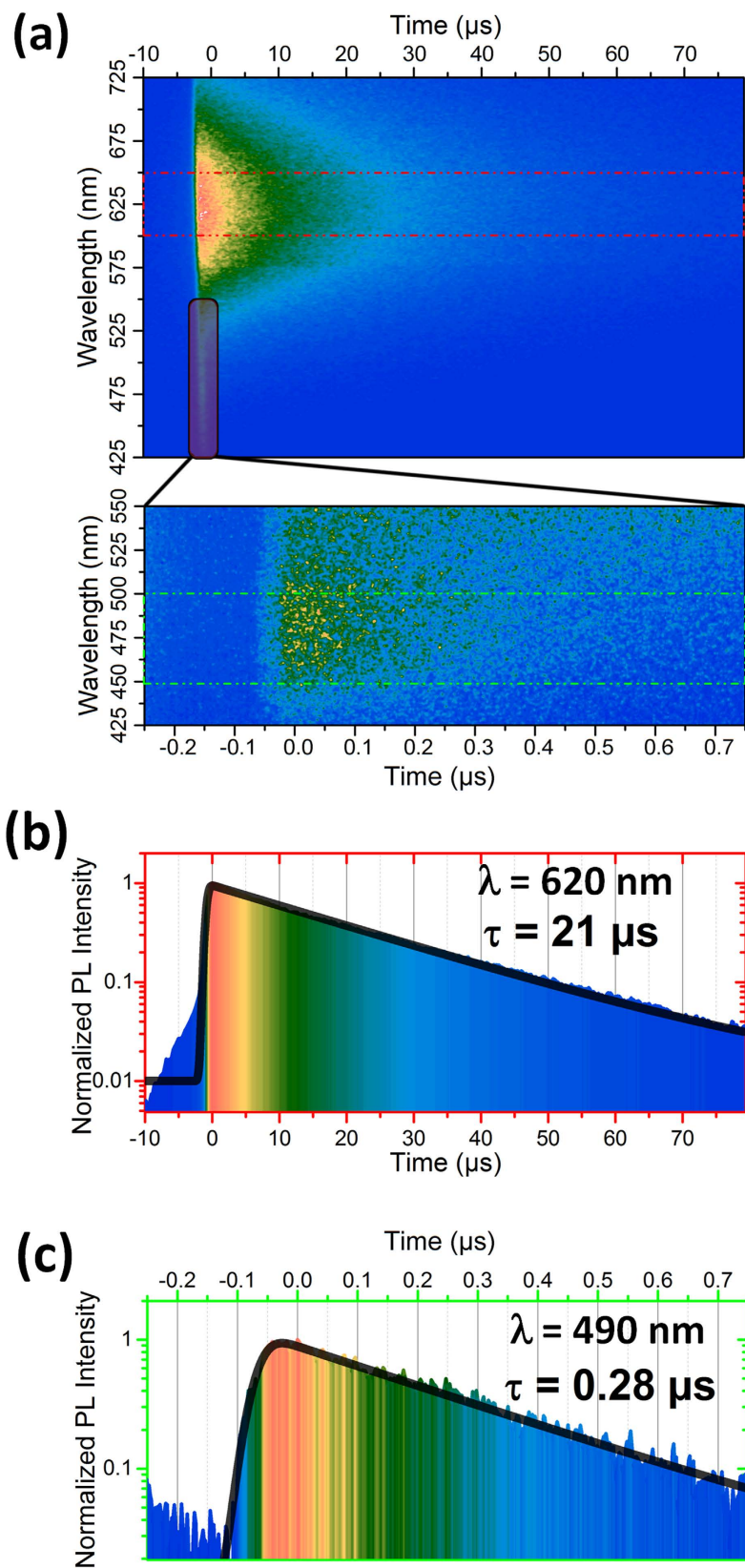


Figure 4. Time-resolved photoluminescence on compound 1 at $\lambda_{\text{exc}} = 267 \text{ nm}$: (a) Overview of the emission vs. decay time in the visible spectrum, (b) Decay of the emission band at 620 nm, and (c) Decay of emission band at 490 nm.

appears that the “closed” cluster possesses all its Cu-Cu Wiberg bond indices ranging in the same order of magnitude (0.08–0.18). However, in the case of the “opened” cluster, there is one Cu-Cu Wiberg bond index which is one order of magnitude smaller than the others (0.009). This result allows us to confirm that for the opened cluster the d^{10} - d^{10} interaction is weak^{10,11,22–26}.

Owing to the competition between non-radiative and radiative decays, most luminescent materials show a significant change in the intensity of photoluminescence vs. temperature variation. Thus, our two compounds show a decrease of intensity of the photoemission bands when the temperature increases. This phenomenon is of interest for the materials showing several emission bands originating from different mechanisms because the evolution of intensity with temperature for each of these emission bands is more likely to be different. It results that the mixing of the emission bands and, thus, the resulting colour of the photoluminescence will change gradually with temperature. This thermochromic behaviour of the photoluminescence will also be more common in the materials for which distinct mechanisms of the photoluminescence are considered. Thus, one could focus on materials in which the different categories of mechanisms (such as ligand-centered, or metal-centered) are possible. The appropriate choice of ligands can favor ligand-centered or ligand-metal charge transfer mechanism. The selection of metals favoring the formation of different clusters by metallophilic interactions can also be a key to obtain materials with multi-band emissions. Indeed, in our compounds, only the inorganic component is different (absence and presence of polynuclear assemblies for Zn and Cu halides, respectively).

In summary, mechanisms of photoluminescence in materials built from transition metal complexes are usually difficult to identify or predict prior to the synthesis. However, the crystal structures built from assemblies exhibiting multi-emission bands are more likely to exhibit thermochromic behaviour of the photoluminescence. Thus, the copper halide material built from polynuclear assemblies exhibits multi-emission bands that evolve differently with temperature. It also results a thermochromic photoluminescence from 77 K to room temperature in a wide spectrum of colors.

Methods

Synthesis. The materials have been hydrothermally synthesized using mixtures of metals (Cu (Merck, 99.7%), Zn (Merck, 99.9%)) and acid (HBr (Alfa Aesar, 48%)). The compounds are synthesized by hydrothermal method (heating at 180 °C during 24 h and slowly cooled down to room temperature at the rate of 10 °C/h) using a 23 mL Teflon-lined stainless steel autoclave. Single crystals suitable for X-ray diffraction were recovered by filtration. CCDC 1482035 and CCDC 1532971 contain the supplementary crystallographic data at room temperature and 100 K, respectively. These data are provided free of charge by The Cambridge Crystallographic Data Centre. Copper bromide (**1**) $[C_6H_{16}N_2]_3[Cu_4Br_6][Cu_2Br_6]$ was synthesized from a mixture of 7.87 mmol of Cu metal, 8.86 mmol of N,N'-dimethylpiperazine and 3 ml of HBr 48%. Zinc bromide (**2**) $[C_6H_{16}N_2]ZnBr_4$ was synthesized from a mixture of 7.65 mmol of Zn metal, 2.95 mmol of N,N'-dimethylpiperazine and 3 ml of HBr 48%.

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Author Contributions

R.G. performed the synthesis and structural characterization of the materials. C.L. carried out DFT calculations. M.P. and F.M. performed solid-state NMR and time resolved photoluminescence, respectively. R.G. and F.M. carried out temperature dependant photoluminescence experiments. R.G. directed the analysis and wrote the paper with the contributions of all authors.

Additional Information

Supplementary information accompanies this paper at <http://www.nature.com/srep>

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