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Water nanoelectrolysis: a simple model

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A simple model of water nanoelectrolysis—defined as the nanolocalization at a single point of any electrolysis phenomenon—is presented. It is based on the electron tunneling assisted by the electric field through the thin film of water molecules (~ 0.3 nm thick) at the surface of a tip-shaped nanoelectrode (micrometric to nanometric curvature radius at the apex). By applying, e.g., an electric potential V_1 during a finite time t_1 , and then the potential $-V_1$ during the same time t_1 , we show that there are three distinct regions in the plane (t_1, V_1): one for the nanolocalization (at the apex of the nanoelectrode) of the electrolysis oxidation reaction, a second one for the nanolocalization of the reduction reaction, and a third one for the nanolocalization of the production of bubbles. These parameters t_1 and V_1 completely control the time at which the electrolysis reaction (of oxidation or reduction) begins, the duration of this reaction, the electrolysis current intensity (i.e., the tunneling current), the number of produced O_2 or H_2 molecules, and the radius of the nanolocalized bubbles. The model is in good agreement with our experiments.

I. INTRODUCTION

Water electrolysis is used for hydrogen production^{1–3} and more generally for the production of bubbles and the study of their formation, dissolution, stability, acoustic properties, etc.,^{4–6} owing to the numerous applications, e.g., in medicine.^{7–9} Microbubbles produced by electrolysis are also used to manipulate a microobject.¹⁰ Very small electrodes, with diameter $< 1 \mu\text{m}$, were used for the study of microbubble/substrate forces or the formation of nanobubbles.^{5,11,12} An effective control of the micro/nanobubbles, e.g., concerning their localization and size, is crucial for all these studies and applications. An example of such control was presented in our previous paper¹³ showing the nanolocalization of the production of bubbles at a single point, namely the apex of a tip-shaped electrode (with micrometric to nanometric curvature radius at the apex). In Ref. 14, we presented a new type of control: the immobilization of a single bubble in the liquid, at some distance from the apex of the electrode. Although electrolysis is classically described using the electric potential, the preceding phenomena revealed the role of the electric field (higher at the apex of the electrode) and the electron tunneling—assisted by the electric field—through the thin film of water layers at the electrode/solution interface. In this paper, we present a general and simple model based on this tunneling and field effect, which is in agreement with the experiments and explains nanoelectrolysis, i.e., the nanolocalization of the electrolysis reactions or the production of bubbles. It leads to a complete control of these reactions and the bubble production.

II. THE MODEL

In our experiments, one of the two electrodes—called the nanoelectrode—is tip-shaped (and made of Pt), with a curvature radius, at the apex of the electrode, ranging from $5 \mu\text{m}$ to 1nm .^{13,14} An aqueous solution of H_2SO_4 (10^{-4} to 10^{-3} mol/L) is generally used. The presence of a few water layers at the surface of the nanoelectrode^{15–18} will be modeled as a dielectric film of constant thickness d ($d \approx 0.3$ nm). Let us denote V_1 the region occupied by the dielectric film, V_2 that occupied by the solution, S_1 the surface of the nanoelectrode, S_2 that of the counter electrode, S_{12} the interface between V_1 and V_2 , and n the unit vector normal to S_{12} , oriented from V_1 to V_2 (Fig. 1). Maxwell's equations give the discontinuities of

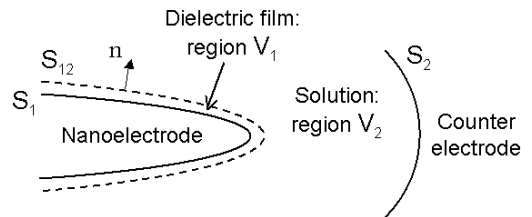


FIG. 1. The nanoelectrode, the dielectric film, the solution, and the counter electrode: general notations.

the normal components of the electric field E and the electric current j at the interface S_{12}

$$\begin{aligned} \varepsilon(E_{2n} - E_{1n}) &= \rho_s \\ j_{2n} - j_{1n} &= -\frac{\partial \rho_s}{\partial t} \end{aligned} \quad (1)$$

(assuming the same permittivity ε in V_1 and V_2 ; ρ_s is the surface charge density on S_{12}), hence

$$\gamma E_{2n} = -\varepsilon \frac{\partial}{\partial t} (E_{2n} - E_{1n}) + j_{1n}, \quad (2)$$

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with the help of Ohm's law $j_2 = \gamma E_2$ (γ being the conductivity of the solution). The term $-\varepsilon \frac{\partial}{\partial t}(E_{2n} - E_{1n})$ represents a 'charging current'—more precisely, the current due to the discharge of the interface S_{12} , according to Eq. (1)—and will be denoted j_C . Owing to the low thickness of the dielectric film, electrons can cross this film by quantum tunneling, if the electric field in the film is high enough, producing the electric current j_{1n} responsible for the electrolysis reactions. This tunneling current j_{1n} will thus be called the electrolysis current and denoted j_E . Its non linear dependence on the electric field $E = E_{1n}$ in the dielectric film—and then on the electric potential $v = E d$ applied to the dielectric film, v being the potential on a point of S_1 minus that on the neighbouring point of S_{12} —will be simply modeled using a threshold value $E_0 = v_0/d$, and a high slope γ_0 for $E > E_0$ ($\gamma_0 \gg \gamma$). For the sake of simplicity in the notations, we will use the same threshold value and slope in the region $E < 0$, i.e.,

$$\begin{aligned} j_E &= 0 & \text{if } -E_0 < E < E_0 \\ j_E &= \gamma_0(E - E_0) & \text{if } E > E_0 \\ j_E &= \gamma_0(E + E_0) & \text{if } E < -E_0 \end{aligned} \quad (3)$$

(Fig. 2).¹⁹ Denoting $j = j_{2n} = \gamma E_{2n}$ the current in the solution (at S_{12}), Eq. (2) may thus be written

$$j = j_C + j_E. \quad (4)$$

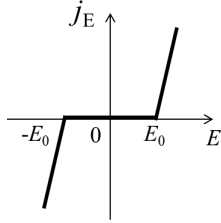


FIG. 2. Modeled dependence of the tunneling or electrolysis current j_E on the electric field E in the dielectric film.

At a given time t , all the points of S_1 (of S_2 , respectively) have practically the same potential—because the electrodes are made of metal—and we denote $V(t)$ the potential on S_1 minus that on S_2 . On the contrary, the surface S_{12} is not equipotential, because the thickness d of the dielectric film is constant but the electric field (in this film) varies, being higher at the apex of the nanoelectrode (where the curvature radius of S_1 is very small). Then, at a given time t , the potential value v varies with the position on S_1 (or on S_{12}). In the following, the nanoelectrode surface S_1 will be simply modeled as: (i) a hemisphere of radius r_1 at the apex, denoted zone a, and (ii) a cylinder (of the same radius) of length l_1 , denoted zone b (Fig. 3). As a first approximation, we assume that the electric potential field in zone a (except near the

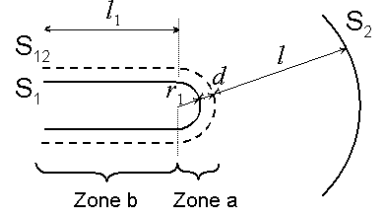


FIG. 3. Geometrical model of the nanoelectrode surface S_1 : a hemisphere (zone a) and a cylinder (zone b).

junction with the cylinder of zone b) is that produced by a spherical electrode, and that in zone b (except near the two ends of the cylinder) is that produced by a (infinite) cylindrical electrode. With the help of Maxwell's equations, we then obtain the values of E_{1n} and E_{2n} in these two zones a and b

$$\begin{aligned} E_{1n} &= \beta_1 v \\ E_{2n} &= \beta_2(V - v), \end{aligned} \quad (5)$$

the values of β_1 , β_2 , and v being different in zones a and b, and distinguished with the respective subscripts a and b, i.e.,

$$\begin{aligned} \beta_{1a} &= \frac{r_1}{(r_1 + d)d} \approx \frac{1}{d} \\ \beta_{2a} &= \frac{r_1 + d + l}{(r_1 + d)l} \approx \frac{1}{r_1} \end{aligned} \quad (6)$$

in zone a, and

$$\begin{aligned} \beta_{1b} &= \frac{1}{(r_1 + d) \log \frac{r_1 + d}{r_1}} \approx \frac{1}{d} \\ \beta_{2b} &= \frac{1}{(r_1 + d) \log \frac{r_1 + d + l}{r_1 + d}} \approx \frac{1}{r_1 \log \frac{l}{r_1}} \end{aligned} \quad (7)$$

in zone b (see Appendix A; \log refers to the natural logarithm; the approximations hold if $d \ll r_1 \ll l$).

Eq. (2) may then be written

$$\gamma \beta_2(V - v) = -\varepsilon \frac{d}{dt}(\beta_2(V - v) - \beta_1 v) + j_E,$$

i.e.,

$$V - v = \tau \left(\frac{dv}{dt} - \frac{1}{\alpha} \frac{dV}{dt} \right) + \frac{j_E}{\gamma \beta_2}, \quad (8)$$

where $\tau = \frac{\varepsilon}{\gamma} \alpha$ and $\alpha = 1 + \frac{\beta_1}{\beta_2}$, which, according to Eq. (3) (with $E = v/d$ and $E_0 = v_0/d$), leads to

$$v + \tau \frac{dv}{dt} = V + \frac{\tau dV}{\alpha dt} \quad \text{if } -v_0 < v < v_0 \quad (9)$$

$$v + \tau' \frac{dv}{dt} = V_+ + \frac{\tau dV_+}{\alpha dt} \quad \text{if } v > v_0 \quad (10)$$

$$v + \tau' \frac{dv}{dt} = V_- + \frac{\tau dV_-}{\alpha dt} \quad \text{if } v < -v_0, \quad (11)$$

where $\tau' = \frac{\tau}{1+\tilde{\gamma}\alpha'}$, $\tilde{\gamma} = \frac{\gamma_0}{\gamma}$, $\alpha' = \frac{1}{\beta_2 d} = \frac{\alpha-1}{\beta_1 d}$, $V_+ = \frac{V+\tilde{\gamma}\alpha'v_0}{1+\tilde{\gamma}\alpha'}$, and $V_- = \frac{V-\tilde{\gamma}\alpha'v_0}{1+\tilde{\gamma}\alpha'}$. Note that α , α' , τ , and τ' have different values in zones a and b, which will be distinguished with the respective subscripts a and b. For a given applied potential $t \rightarrow V(t)$, Eq. (8) or Eqs. (9), (10), and (11) determine v as a function of time (i.e., $v_a(t)$ using α_a , α'_a , τ_a , and τ'_a , and $v_b(t)$ using α_b , α'_b , τ_b , and τ'_b).

Since $j = \gamma E_{2n} = \gamma\beta_2(V - v)$, the current intensity (in the solution, in zone a or in zone b) is

$$I = \gamma\beta_2 a(V - v) = \frac{\gamma a}{\alpha' d}(V - v), \quad (12)$$

a being the corresponding area of S_{12} , i.e. $a_a = 2\pi(r_1 + d)^2 \approx 2\pi r_1^2$ for zone a, and $a_b = 2\pi(r_1 + d)l_1 \approx 2\pi r_1 l_1$ for zone b. In fact, the total intensity in the electric circuit is the sum of the contributions of zone a and zone b:

$$I = I_a + I_b. \quad (13)$$

Similarly, the electrolysis current intensity and the charging current intensity (in zone a or in zone b) are respectively $I_E = j_E a$ and $I_C = j_C a$, thus

$$\begin{aligned} I_E &= 0 & \text{if } -v_0 < v < v_0 \\ I_E &= \frac{\gamma a}{\alpha' d} \tilde{\gamma} \alpha' (v - v_0) & \text{if } v > v_0 \\ I_E &= \frac{\gamma a}{\alpha' d} \tilde{\gamma} \alpha' (v + v_0) & \text{if } v < -v_0, \end{aligned} \quad (14)$$

and

$$I = I_C + I_E \quad (15)$$

(in zone a or in zone b).

III. A SIMPLE EXAMPLE: POTENTIAL OF RECTANGULAR SHAPE

Let us show the consequences of the preceding model with the following simple example of a potential of rectangular shape (Fig. 4):

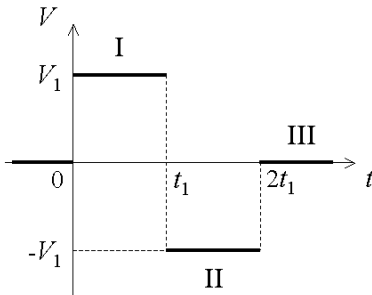


FIG. 4. Applied potential $t \rightarrow V(t)$ of rectangular shape.

$$\begin{aligned} V(t) &= 0 & \text{if } t < 0 \\ V(t) &= V_1 & \text{if } 0 < t < t_1 \quad (\text{phase I}) \\ V(t) &= -V_1 & \text{if } t_1 < t < 2t_1 \quad (\text{phase II}) \\ V(t) &= 0 & \text{if } t > 2t_1 \quad (\text{phase III}). \end{aligned}$$

We here consider $V_1 > 0$, but the case $V_1 < 0$ is exactly similar.

A. Phase I

1. Before the electrolysis reaction

During phase I, let us consider the solution v of Eq. (8) or (9) (in zone a or in zone b), assuming that $v(t)$ remains lower than v_0 . At $t = 0$, the discontinuity jump $V(0^+) - V(0^-) = V_1$ produces the term $V_1 \delta_0(t)$ in $\frac{dV}{dt}$ (δ_0 being the Dirac measure at 0), then, according to Eq. (8), the term $\frac{V_1}{\alpha} \delta_0(t)$ in $\frac{dv}{dt}$, and then the discontinuity jump $v(0^+) - v(0^-) = \frac{V_1}{\alpha}$ for v at 0. If $V_1 < \alpha v_0$ (which corresponds to our usual experimental conditions²⁰), the solution is then

$$\begin{aligned} v(t) &= 0 & \text{if } t < 0 \\ v(t) &= V_1 - (V_1 - \frac{V_1}{\alpha}) e^{-t/\tau} & \text{if } t > 0, \end{aligned} \quad (16)$$

represented in Fig. 5, with the corresponding current intensity given by Eq. (12).

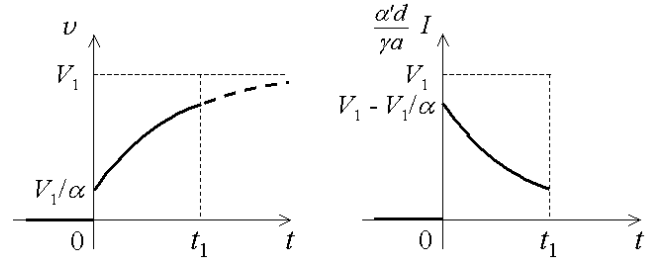


FIG. 5. Solution v of Eq. (9) during phase I (left), and corresponding current intensity I (right).

Clearly, $v(t)$ is always lower than V_1 and, if $V_1 < v_0$, $v(t)$ will always remain lower than v_0 . In the following, we suppose $v_0 < V_1 < \alpha v_0$, so that $v(t)$ will reach the value v_0 at the time

$$t' = \tau \log \frac{(1 - \frac{1}{\alpha})V_1}{V_1 - v_0} \quad (17)$$

(if the duration of phase I is large enough; Fig. 6). Note that, if $V_1 > \alpha v_0$, then $v(0^+) = \frac{V_1}{\alpha} > v_0$, so that $t' = 0$ in this case (and the solution of Eq. (10) shows that $v(t) > v_0$ for $0 < t < t_1$).

Consider now the relative position of t_1 and t' . If $t' > t_1$, $v(t)$ will not reach the value v_0 during phase

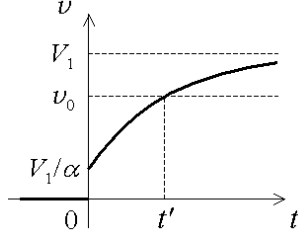


FIG. 6. Solution v of Eq. (9) during phase I. At the time t' , $v(t)$ reaches the value v_0 (if the duration of phase I is large enough).

I (i.e., for $0 < t < t_1$), and if $t' < t_1$, $v(t)$ will reach the value v_0 during phase I, at the time t' . In this last case, according to Eq. (3), there will be an electrolysis current $j_E > 0$, i.e., the oxidation reaction of electrolysis will occur, for $t' < t < t_1$. Let us represent, in the plane (t_1, V_1) , the two “oxidation” curves $t_1 = t'_a(V_1)$ in zone a (given by Eq. (17) with τ_a and α_a) and $t_1 = t'_b(V_1)$ in zone b (Eq. (17) with τ_b and α_b) (Fig. 7). In the region

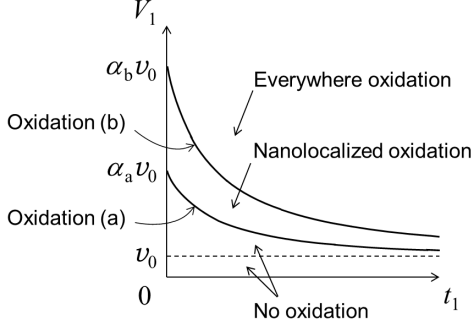


FIG. 7. During phase I, in the plane (t_1, V_1) , the two oxidation curves $t_1 = t'_a(V_1)$ in zone a and $t_1 = t'_b(V_1)$ in zone b delimit the three regions: “no oxidation” (lower region), “nanocalized oxidation” (between the two curves), and “everywhere oxidation” (upper region).

$t_1 < t'_a(V_1)$, there is no oxidation reaction (neither in zone a nor in zone b). In the region $t_1 > t'_b(V_1)$, the oxidation reaction occurs everywhere on the nanoelectrode (in zone a, after t'_a , and in zone b, after t'_b). In the region $t'_a(V_1) < t_1 < t'_b(V_1)$, the oxidation reaction is nanocalized at the apex of the nanoelectrode (it occurs in zone a, after t'_a , but not in zone b).

2. During the electrolysis reaction

If $t' < t_1$ (in zone a or in zone b), $v(t)$ reaches the value v_0 at t' and, for $t > t'$, v is the solution of the new Eq. (10), i.e.,

$$v(t) = V_{1+} - (V_{1+} - v_0)e^{-(t-t')/\tau'} \quad \text{if } t > t', \quad (18)$$

where $V_{1+} = \frac{V_1 + \tilde{\gamma}\alpha'v_0}{1 + \tilde{\gamma}\alpha'}$, represented in Fig. 8. Note that, for any finite value of $\tilde{\gamma}$, dv/dt is continuous at t' and $\frac{dv}{dt}(t') = (V_1 - v_0)/\tau$. Since $\gamma_0 \gg \gamma$, we may consider

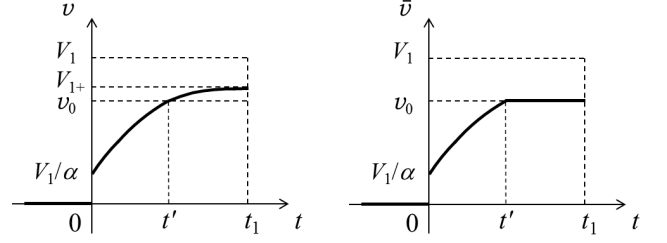


FIG. 8. Solution v during phase I (i.e., Eq. (16) for $t < t'$ and Eq. (18) for $t' < t < t_1$) for a finite value of $\tilde{\gamma}$ (left) and for $\tilde{\gamma} = +\infty$ (right). Case $t' < t_1$.

that $\tilde{\gamma}$ tends to $+\infty$ and the corresponding limit value of any quantity X will be denoted \bar{X} . Thus, $v(t)$ tends to $\bar{v}(t) = v_0$ for $t' < t < t_1$ (since V_{1+} tends to v_0 ; see Fig. 8). According to Eq. (12), the current intensity I and its limit value \bar{I} are then represented in Fig. 9. Similarly,

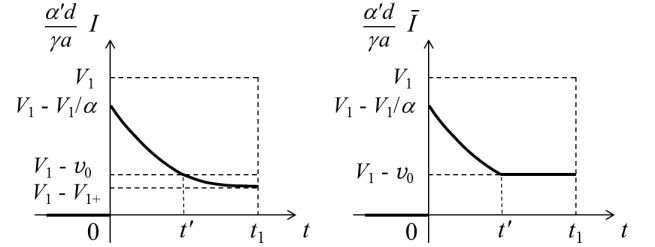


FIG. 9. Current intensity I during phase I for $\tilde{\gamma} < +\infty$ (left) and for $\tilde{\gamma} = +\infty$ (right). Case $t' < t_1$.

Eq. (14) gives the electrolysis current intensity I_E and its limit value \bar{I}_E (Fig. 10). Note that I and I_E have the

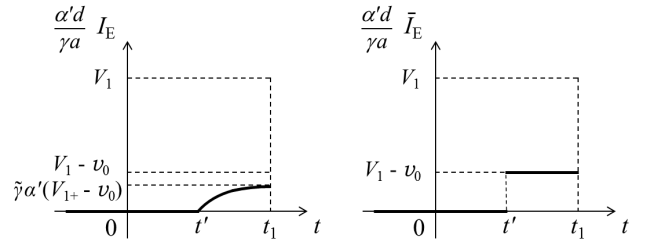


FIG. 10. Electrolysis current intensity I_E during phase I for $\tilde{\gamma} < +\infty$ (left) and for $\tilde{\gamma} = +\infty$ (right). Case $t' < t_1$.

same asymptotic value $\frac{\gamma a}{\alpha' d}(V_1 - V_{1+}) = \frac{\gamma a}{\alpha' d} \tilde{\gamma} \alpha' (V_{1+} - v_0) = \frac{\gamma a}{\alpha' d} \frac{\tilde{\gamma} \alpha'}{1 + \tilde{\gamma} \alpha'} (V_1 - v_0)$ and that $\bar{I} = \bar{I}_E = \frac{\gamma a}{\alpha' d} (V_1 - v_0)$ for $t' < t < t_1$. Thus, while before the electrolysis

reaction ($0 < t < t'$), the current intensity is equal to the charging current intensity ($I = I_C, I_E = 0$), during the electrolysis reaction ($t' < t < t_1$), owing to $\gamma_0 \gg \gamma$, the (limit) current intensity is equal to the (limit) electrolysis current intensity ($\bar{I} = \bar{I}_E, \bar{I}_C = 0$). This is a consequence of $\bar{v} = v_0 = \text{constant}$ for $t' < t < t_1$, which implies that E_{1n} and E_{2n} are constant (from Eq. (5)), and then $\bar{j}_C = -\varepsilon \frac{\partial}{\partial t}(E_{2n} - E_{1n}) = 0$.

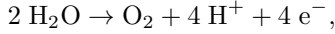
From Eqs. (14) and (18), we then obtain the electric charge given by the nanoelectrode to the solution for the electrolysis reaction (in zone a or zone b)

$$\begin{aligned} Q' &= \int_{t'}^{t_1} I_E(t) dt \\ &= \frac{\gamma a}{\alpha' d} \frac{\tilde{\gamma} \alpha'}{1 + \tilde{\gamma} \alpha'} (V_1 - v_0) (T' - \tau' (1 - e^{-T'/\tau'})), \end{aligned} \quad (19)$$

where $T' = t_1 - t'$ is the electrolysis duration, and its limit value (for $\tilde{\gamma} = +\infty$)

$$\bar{Q}' = \frac{\gamma a}{\alpha' d} (V_1 - v_0) T'. \quad (20)$$

Note that this charge is due to the electrons—tunneling through the dielectric film, from the solution to the nanoelectrode—produced by the electrolysis oxidation reaction



which gives the number of produced O_2 molecules

$$n_{\text{O}_2} = \frac{Q'}{4q_e} \quad (21)$$

(q_e the elementary charge).

B. Phase II

1. Before the electrolysis reaction

First note that the following results are based on the simple assumption of a unique threshold value v_0 (and slope γ_0).¹⁹ As previously, the following study relates to either zone a or zone b. Because of Eq. (8), the discontinuity jump $V(t_1^+) - V(t_1^-) = -2V_1$ produces the discontinuity jump $v(t_1^+) - v(t_1^-) = -2\frac{V_1}{\alpha}$ (as it occurred at $t = 0$). Since $-v_0 < v(t_1^+) < v_0$ ($v(t_1^+) > -v_0$ because $v(t_1^-) > \frac{V_1}{\alpha}$; $v(t_1^+) < v_0$ because $v(t_1^-)$ is either $< v_0$ or $< V_{1+} \approx v_0$ owing to the high value of $\tilde{\gamma}$), the solution of Eq. (9) gives

$$v(t) = -V_1 + (V_1 + v(t_1^+)) e^{-(t-t_1)/\tau} \text{ if } t > t_1 \quad (22)$$

(as long as $v(t) > -v_0$). In the case $t_1 < t'$, we know that, during phase I, $v(0^+) = \frac{V_1}{\alpha}$ and $v(t)$ increases but doesn't reach the value v_0 for $0 < t < t_1$. If $v(t_1^+)$ were equal to $-\frac{V_1}{\alpha}$, we would have the same situation (with the opposite sign for v) during phase II, so that $v(t)$ (now decreasing) would not reach the value $-v_0$ for $t_1 < t < 2t_1$.

Since $v(t_1^+) > -\frac{V_1}{\alpha}$ (because $v(t_1^-) > \frac{V_1}{\alpha}$), this implies that $v(t)$ will not reach the value $-v_0$ during phase II. In other words, if there is no electrolysis reaction (oxidation) during phase I, there will be no electrolysis reaction (reduction) during phase II.

We then consider the case $t' < t_1$. If the duration of phase II is large enough, $v(t)$ will reach the value $-v_0$ at the time $t_1 + t''$, with (from Eq. (22))

$$\begin{aligned} t'' &= \tau \log \frac{V_1 + v(t_1^+)}{V_1 - v_0} \\ &= \tau \log \frac{(1 - \frac{2}{\alpha})V_1 + v_0 + \frac{V_1 - v_0}{1 + \tilde{\gamma} \alpha'} (1 - e^{-(t_1 - t')/\tau'})}{V_1 - v_0} \end{aligned} \quad (23)$$

(using $v(t_1^-)$ given by Eq. (18)) and its limit value (for $\tilde{\gamma} = +\infty$)

$$\bar{t}'' = \tau \log \frac{(1 - \frac{2}{\alpha})V_1 + v_0}{V_1 - v_0}. \quad (24)$$

Clearly, $t'' > t'$ because $v(t_1^+) > -\frac{V_1}{\alpha}$. Thus, if $t'' < t_1$, the electrolysis duration (reduction) $\bar{T}'' = t_1 - t''$ during phase II is lower than the electrolysis duration (oxidation) $T' = t_1 - t'$ during phase I.

Strictly speaking, t'' is a function of (t_1, V_1) (Eq. (23) with t' given by Eq. (17)), but its dependence on t_1 is very low and, for large values of $\tilde{\gamma}$, $t'' \approx \bar{t}''$ which only depends on V_1 . Let us represent, in the plane (t_1, V_1) , the two “reduction” curves $t_1 = t''_a(t_1, V_1)$ in zone a (given by Eq. (23) with $\tau_a, \alpha_a, \tau'_a, \alpha'_a$, and $t'_a(V_1)$) and $t_1 = t''_b(t_1, V_1)$ in zone b (Eq. (23) with $\tau_b, \alpha_b, \tau'_b, \alpha'_b$, and $t'_b(V_1)$). Thus, in the region $t_1 < t''_a(t_1, V_1)$, there is no reduction reaction of electrolysis during phase II (neither in zone a nor in zone b). In the region $t_1 > t''_b(t_1, V_1)$, the reduction reaction occurs everywhere on the nanoelectrode (in zone a, after $t_1 + t''_a$, and in zone b, after $t_1 + t''_b$). In the region $t''_a(t_1, V_1) < t_1 < t''_b(t_1, V_1)$, the reduction reaction is nanolocalized at the apex of the nanoelectrode (it occurs in zone a, after $t_1 + t''_a$, but not in zone b).

It may be shown that $t''_a < t''_b$ in general (e.g., if $\alpha_b > 2\alpha_a$, $\alpha_a \gg 1$, and $\tilde{\gamma} > 1$) for any (t_1, V_1) such that $v_0 < V_1 < \alpha_a v_0$ and $t_1 > t'_a$. We thus have five regions in the plane (t_1, V_1) (see Fig. 11). Below the curve $t_1 = t'_a$, there is no oxidation reaction and no reduction reaction (neither in zone a nor in zone b). Between the two curves $t_1 = t'_a$ and $t_1 = t''_a$, the oxidation reaction is nanolocalized at the apex of the nanoelectrode (it occurs in zone a but not in zone b) and there is no reduction reaction (neither in zone a nor in zone b). Between the two curves $t_1 = t''_a$ and $t_1 = t'_b$, the oxidation reaction and the reduction reaction are nanolocalized at the apex of the nanoelectrode (they occur in zone a but not in zone b). Between the two curves $t_1 = t'_b$ and $t_1 = t''_b$, the oxidation reaction occurs everywhere on the nanoelectrode (in zone a and in zone b) and the reduction reaction is nanolocalized at the apex of the nanoelectrode (it occurs

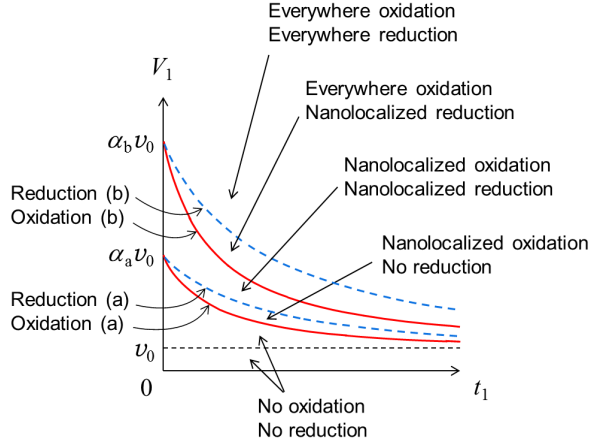


FIG. 11. In the plane (t_1, V_1) , the four curves $t_1 = t'_a$ (oxidation in zone a), $t_1 = t''_a$ (reduction in zone a), $t_1 = t'_b$ (oxidation in zone b), and $t_1 = t''_b$ (reduction in zone b) delimit the five regions: “no oxidation, no reduction” (below $t_1 = t'_a$), “nanolocalized oxidation, no reduction” (between $t_1 = t'_a$ and $t_1 = t''_a$), “nanolocalized oxidation, nanolocalized reduction” (between $t_1 = t''_a$ and $t_1 = t'_b$), “everywhere oxidation, nanolocalized reduction” (between $t_1 = t'_b$ and $t_1 = t''_b$), and “everywhere oxidation, everywhere reduction” (above $t_1 = t''_b$).

in zone a but not in zone b). Above the curve $t_1 = t''_b$, the oxidation reaction and the reduction reaction occur everywhere on the nanoelectrode (in zone a and in zone b).

2. During the electrolysis reaction

If $t'' < t_1$ (in zone a or in zone b), $v(t)$ reaches the value $-v_0$ at $t_1 + t''$ and, for $t > t_1 + t''$, v is the solution of the new Eq. (11), i.e.

$$v(t) = -V_{1+} + (V_{1+} - v_0) e^{-(t-t_1-t'')/\tau'} \quad \text{if } t > t_1 + t'', \quad (25)$$

v and \bar{v} being represented in Fig. 12. Note that, for $\tilde{\gamma} < +\infty$, dv/dt is continuous at $t_1 + t''$ and $\frac{dv}{dt}(t_1 + t'') = -\frac{dv}{dt}(t') = -(V_1 - v_0)/\tau$. According to Eq. (12), the current intensity I and its limit value \bar{I} are then represented in Fig. 13. Similarly, Eq. (14) gives the electrolysis current intensity I_E and its limit value \bar{I}_E (Fig. 14). Note that I and I_E have the same asymptotic value $\frac{\gamma^a}{\alpha^a d}(-V_1 + V_{1+}) = \frac{\gamma^a}{\alpha^a d} \tilde{\gamma} \alpha' (-V_{1+} + v_0) = -\frac{\gamma^a}{\alpha^a d} \frac{\tilde{\gamma} \alpha'}{1 + \tilde{\gamma} \alpha'} (V_1 - v_0)$ and that $\bar{I} = \bar{I}_E = \frac{\gamma^a}{\alpha^a d} (-V_1 + v_0)$ for $t_1 + t'' < t < 2t_1$ (these are the opposite values compared to those of phase I). Just as in phase I, while before the electrolysis reaction ($t_1 < t < t_1 + t''$), the current intensity is equal to the charging current intensity ($I = I_C, I_E = 0$), during the electrolysis reaction ($t_1 + t'' < t < 2t_1$), owing to $\gamma_0 \gg \gamma$, the (limit) current intensity is equal to the (limit) electrolysis current intensity ($\bar{I} = \bar{I}_E, \bar{I}_C = 0$).

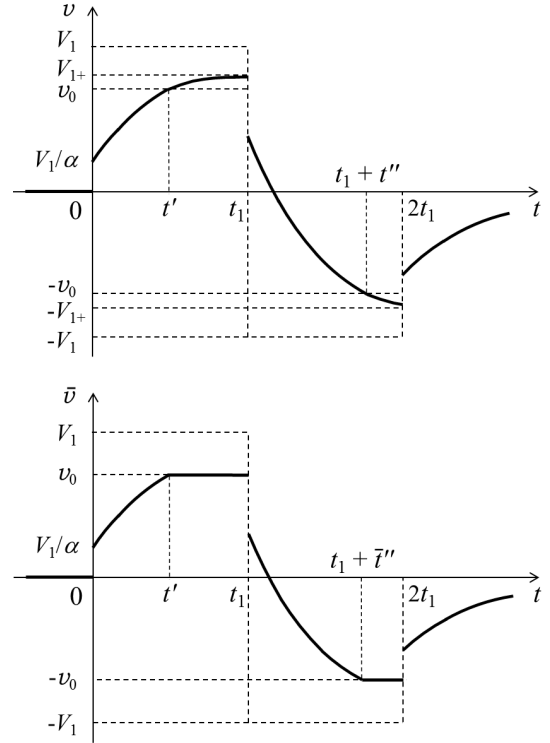


FIG. 12. Solution v during phase II (i.e., Eq. (22) for $t_1 < t < t_1 + t''$ and Eq. (25) for $t_1 + t'' < t < 2t_1$) for $\tilde{\gamma} < +\infty$ (top) and for $\tilde{\gamma} = +\infty$ (bottom). Case $t'' < t_1$. Phase I and phase III are also represented.

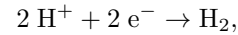
From Eqs. (14) and (25), we then obtain the electric charge given by the nanoelectrode to the solution for the electrolysis reaction (in zone a or in zone b)

$$\begin{aligned} Q'' &= \int_{t_1+t''}^{2t_1} I_E(t) dt \\ &= -\frac{\gamma^a}{\alpha^a d} \frac{\tilde{\gamma} \alpha'}{1 + \tilde{\gamma} \alpha'} (V_1 - v_0) (T'' - \tau' (1 - e^{-T''/\tau'})), \end{aligned} \quad (26)$$

where $T'' = t_1 - t''$ is the electrolysis duration, and its limit value (for $\tilde{\gamma} = +\infty$)

$$\bar{Q}'' = -\frac{\gamma^a}{\alpha^a d} (V_1 - v_0) \bar{T}'' \quad (27)$$

($\bar{T}'' = t_1 - t''$). Note that this charge is due to the electrons tunneling through the dielectric film, from the nanoelectrode to the solution, and producing the electrolysis reduction reaction



which gives the number of produced H_2 molecules

$$n_{\text{H}_2} = \frac{|Q''|}{2q_e}. \quad (28)$$

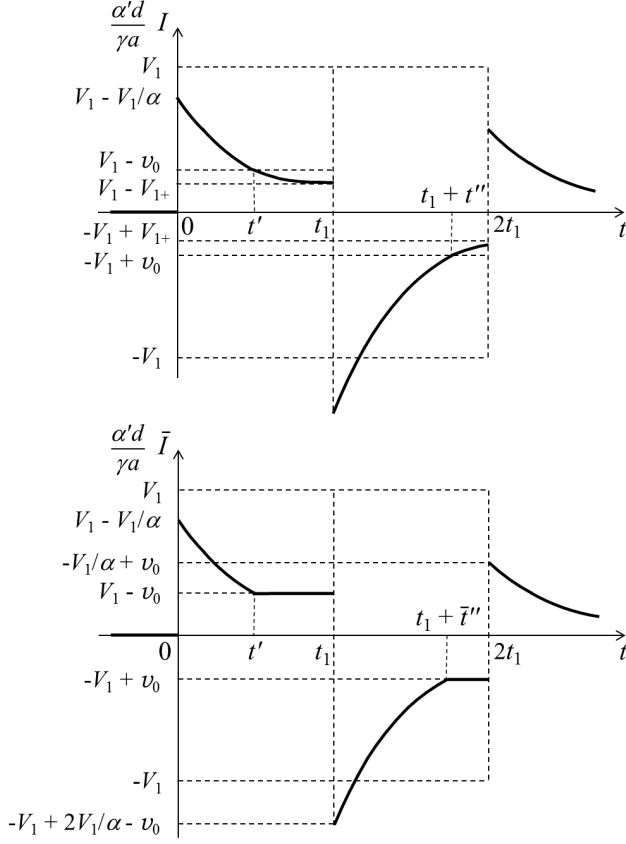


FIG. 13. Current intensity I during phase II for $\tilde{\gamma} < +\infty$ (top) and for $\tilde{\gamma} = +\infty$ (bottom). Case $t'' < t_1$. Phase I and phase III are also represented.

As noted above, $T'' < T'$, which, according to Eqs. (19) and (26), implies that $|Q''| < Q'$: the absolute value of the electric charge used for the electrolysis reaction during phase II (in our case, reduction) is lower than that used for the electrolysis reaction during phase I (in our case, oxidation). In our case, this means that the number of produced H_2 molecules is lower than twice the number of produced O_2 molecules.

C. Phase III

Because of Eq. (8), the discontinuity jump $V(2t_1^+) - V(2t_1^-) = V_1$ produces the discontinuity jump $v(2t_1^+) - v(2t_1^-) = \frac{V_1}{\alpha}$. Since $-v_0 < v(2t_1^+) < 0$ ($v(2t_1^+) > -v_0$ because $v(2t_1^-)$ is either $> -v_0$ or $> -V_{1+} \approx -v_0$ owing to the high value of $\tilde{\gamma}$; $v(2t_1^+) < 0$ is proved in Appendix B), the solution of Eq. (9) gives

$$v(t) = v(2t_1^+) e^{-(t-2t_1)/\tau} \text{ if } t > 2t_1, \quad (29)$$

which has the form represented in phase III of Fig. 12. Thus, there is no electrolysis reaction during phase III.

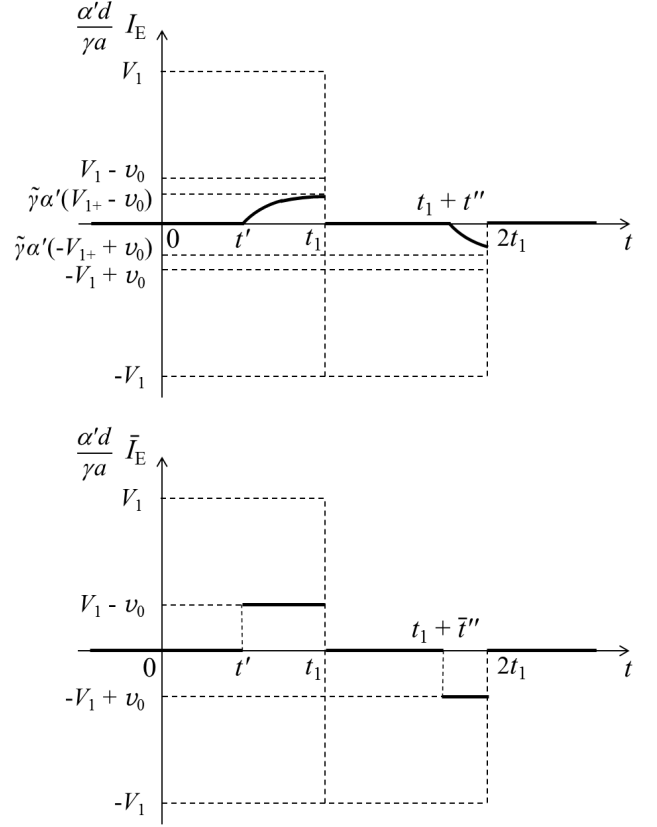


FIG. 14. Electrolysis current intensity I_E during phase II for $\tilde{\gamma} < +\infty$ (top) and for $\tilde{\gamma} = +\infty$ (bottom). Case $t'' < t_1$. Phase I and phase III are also represented.

IV. BUBBLE PRODUCTION

If $t_1 > t'$ (in zone a or in zone b), the electrolysis oxidation reaction (during phase I, between t' and t_1) produces O_2 molecules which diffuse in the solution. If the volume mole density ρ of these molecules in the solution exceeds the saturation value $H p_a$ (H the Henry's constant for O_2 in water, p_a the atmospheric pressure) in some region of the solution, at some time, a bubble of O_2 will be produced. In the simple case $\tilde{\gamma} = +\infty$, the flux of O_2 moles produced at the surface of the nanoelectrode (more precisely, at the interface S_{12}), in zone a or in zone b, is

$$j_m = \frac{\bar{j}_E}{4N_A q_e} = \frac{\gamma}{4N_A q_e \alpha' d} (V_1 - v_0), \quad (30)$$

constant between t' and t_1 (see the above Section III A 2 and bottom of Fig. 14; N_A is the Avogadro constant). Close to the surface S_{12} (corresponding to the high values of ρ), this surface may be assimilated to its tangent plane

and an approximate solution of the diffusion equation is

$$\rho(x, t) = \frac{j_m}{(\pi D)^{1/2}} \int_0^{t-t'} \frac{e^{-x^2/(4Du)}}{u^{1/2}} du \quad \text{if } t' < t < t_1,$$

$$\rho(x, t) = \frac{j_m}{(\pi D)^{1/2}} \int_{t-t_1}^{t-t'} \frac{e^{-x^2/(4Du)}}{u^{1/2}} du \quad \text{if } t_1 < t \quad (31)$$

(see Appendix C; x is the distance to S_{12} , here assumed small with respect to r_1 , and D the diffusion coefficient for O_2 in water). Note that we have to add to the preceding value the initial constant density $\rho_a = cHp_a$ of O_2 in the solution due to the equilibrium with the atmospheric O_2 ($c p_a$ being the partial pressure of O_2 in the atmosphere).

Clearly, at fixed t , ρ is a decreasing function of x , with a maximum value at $x = 0$

$$\rho(0, t) = \frac{j_m}{(\pi D)^{1/2}} \int_0^{t-t'} \frac{du}{u^{1/2}}$$

$$= \frac{2j_m}{(\pi D)^{1/2}} (t-t')^{1/2} \quad \text{if } t' < t < t_1,$$

$$\rho(0, t) = \frac{j_m}{(\pi D)^{1/2}} \int_{t-t_1}^{t-t'} \frac{du}{u^{1/2}}$$

$$= \frac{2j_m}{(\pi D)^{1/2}} ((t-t')^{1/2} - (t-t_1)^{1/2})$$

$$= \frac{2j_m}{(\pi D)^{1/2}} \frac{t_1 - t'}{(t-t')^{1/2} + (t-t_1)^{1/2}} \quad \text{if } t_1 < t. \quad (32)$$

This explains why bubbles are produced practically at the contact with the nanoelectrode. Moreover, these equations show that $\rho(0, t)$ is an increasing function of t for $t' < t < t_1$ and a decreasing one for $t > t_1$, with a maximum value at $t = t_1$

$$\rho(0, t_1) = \frac{2j_m}{(\pi D)^{1/2}} (t_1 - t')^{1/2}. \quad (33)$$

Our simple model is based on this maximum density value: we consider that a bubble will be produced if the maximum value of the density in the solution reaches some supersaturation value sHp_a

$$\rho(0, t_1) + \rho_a = sHp_a, \quad (34)$$

i.e.,

$$(V_1 - v_0)(t_1 - t')^{1/2} = 2(s - c)Hp_a(\pi D)^{1/2} N_A q_e \frac{\alpha' d}{\gamma} \quad (35)$$

(with the help of Eqs. (30) and (33)), in which t' is the function of V_1 expressed by Eq. (17) (with $v_0 < V_1 < \alpha v_0$). In the plane (t_1, V_1) , this “bubble” curve (expressed by Eq. (35)) is situated above the oxidation curve $t_1 = t'$ but generally intersects (if the coefficient s is not too high) the reduction curve $t_1 = \bar{t}'$ at two points, so that there is a large part of the bubble curve situated

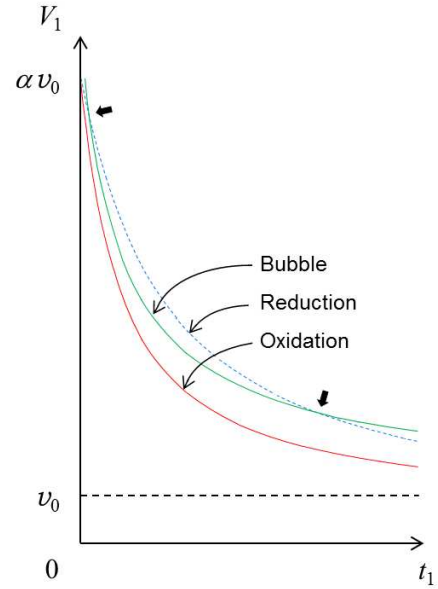


FIG. 15. In the plane (t_1, V_1) , position of the bubble curve Eq. (35) (continuous green curve) with respect to the oxidation curve $t_1 = t'$ (continuous red) and the reduction curve $t_1 = \bar{t}'$ (dashed blue) (in zone a or in zone b). The bubble curve intersects the reduction curve at two points (black arrows).

below the reduction curve (see Fig. 15). An equation similar to Eq. (35) may be obtained for H_2 molecules but the corresponding “ H_2 bubble” curve would be situated above the reduction curve $t_1 = \bar{t}''$, and then above a large part of the bubble curve Eq. (35): thus, at least in this large part, an O_2 bubble will appear but no H_2 bubble is produced. In the next section, we will see that the whole bubble curve Eq. (35) well represents the experimental bubble curve. In our experiments, with small values of t_1 ($0.5 \text{ ms} < t_1 < 0.5 \text{ s}$), we generally observe only one bubble, which probably contains both O_2 and H_2 molecules (in the case $t_1 > \bar{t}''$). This means that after an O_2 bubble is created, the H_2 molecules produced during the reduction period $[t_1 + \bar{t}'', 2t_1]$ will enter this bubble (without creating a new bubble).

Clearly, the preceding bubble curve (Eq. (35)) delimits two regions: an upper one where bubble production occurs and a lower one with no bubble production. This bubble curve—together with the oxidation and reduction ones—is represented in Fig. 16, for the two zones a and b. Thus, below the bubble curve of zone a, there is no bubble production (neither in zone a nor in zone b). Above the bubble curve of zone b, the bubble production occurs everywhere on the nanoelectrode (in zone a and in zone b). *Between the bubble curve of zone a and that of zone b, the bubble production is nanolocalized at the apex of the nanoelectrode* (it occurs in zone a but not in zone b). A numerical example will be given in the Discussion

(Section V).

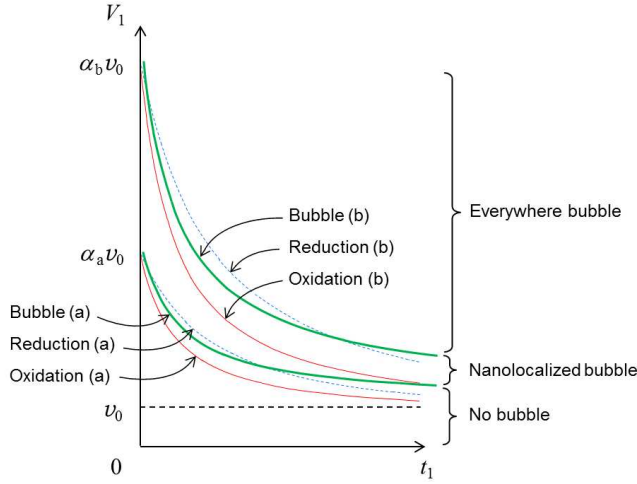


FIG. 16. In the plane (t_1, V_1) , the bubble curve of zone a and that of zone b (thick continuous green curves) delimit the three regions: “no bubble” (lower region), “nanolocalized bubble” (between the two curves), and “everywhere bubble” (upper region).

Let us consider the case of a bubble produced in zone a (i.e., at the apex of the nanoelectrode) and simply assume that when it is formed it contains all the O_2 and H_2 molecules generated by the electrolysis reactions in zone a. The radius R of the bubble is determined by the number n of these molecules

$$(p_a + \frac{2\sigma}{R}) \frac{4}{3} \pi R^3 = nkT \quad (36)$$

(p_a the atmospheric pressure, σ the surface tension, k the Boltzmann constant, T the temperature), which leads to

$$(V_1 - v_0)(t_1 - t'_a) = \frac{8q_e}{3kT\gamma r_1} (p_a + \frac{2\sigma}{R}) R^3 \quad (37)$$

if $t'_a < t_1 < \bar{t}''_a$

(with $n = n_{O_2}$ given by Eqs. (21) and (20), with α'_a and a_a ; t'_a and \bar{t}''_a given by the respective Eqs. (17) and (24), with τ_a and α_a ; $v_0 < V_1 < \alpha_a v_0$) and to

$$(V_1 - v_0)(3t_1 - t'_a - 2\bar{t}''_a) = \frac{8q_e}{3kT\gamma r_1} (p_a + \frac{2\sigma}{R}) R^3 \quad (38)$$

if $t_1 > \bar{t}''_a$

(with $n = n_{O_2} + n_{H_2}$; n_{O_2} as above, n_{H_2} given by Eqs. (28) and (27), with α'_a and a_a). In the plane (t_1, V_1) , such “radius” curves (expressed, for each value of R , by Eq. (37) below the reduction curve and by Eq. (38) above the reduction curve) are represented in Fig. 17 for three increasing values of R , i.e., R_1 , R_2 , and R_3 with $R_1 < R_2 < R_3$ (respectively denoted “radius 1”, “radius 2”, and “radius 3” curves). They either intersect the reduction curve (of zone a) at two points (indicated

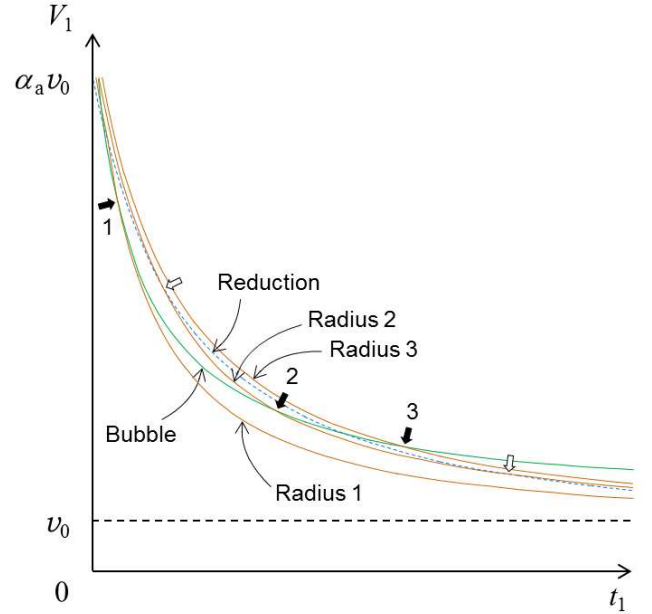


FIG. 17. In the plane (t_1, V_1) , the represented radius 1, radius 2, and radius 3 curves (brown color) respectively correspond to three increasing radius values, $R_1 < R_2 < R_3$. The two empty arrows indicate the intersections of the radius 2 curve with the reduction curve (dashed blue curve). The black arrows 1, 2, and 3 respectively indicate the intersections of the bubble curve (green) with the radius 1, radius 2, and radius 3 curves (excepting the region of very small values of t_1). As R increases from R_1 to R_3 , we can see that the value of t_1 corresponding to these intersection points increases. The oxidation curve (see Fig. 15), situated below all these curves, is not represented for the sake of clarity.

by empty arrows for the radius 2 curve in Fig. 17) or are completely situated above the reduction curve (see the radius 3 curve in Fig. 17). Moreover, they intersect the above defined bubble curve (of zone a) at one point (if we except the very small values of t_1 where the bubble curve is above the reduction curve), the corresponding value of R increasing with t_1 (see Fig. 17). Thus, excepting the very small values of t_1 and following the bubble curve of zone a, i.e., at the limit between the no bubble and the nanolocalized bubble regions, the radius of the bubble (produced at the apex of the nanoelectrode) decreases when t_1 decreases. This corresponds to the intuitive idea that a lower electric charge involved in the electrolysis reactions (then, a lower bubble radius), if released in a shorter time ($t_1 - t'$ decreases if t_1 decreases, along the bubble curve), will give the same maximal concentration $\rho(0, t_1)$ in the solution, because of diffusion. The above radius curves give the radius of the nanolocalized bubble for each point (t_1, V_1) situated above the bubble curve of zone a (and $V_1 < \alpha_a v_0$). A numerical example is given in the following section.

V. DISCUSSION

In fact, the nanoelectrode geometry shown in Fig. 3 represents the tip of the nanoelectrode. The whole geometry of the nanoelectrode surface S_1 may be modelled as: (i) the preceding two zones a and b, at the tip, (ii) a transition zone c (of 100–300 μm length), and (iii) a long cylinder of length l'_1 (~ 1 cm) and radius r'_1 (~ 50 μm), denoted zone b' (Fig. 18). As above, we assume

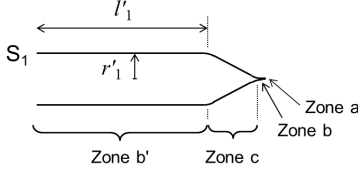


FIG. 18. Geometry of the whole nanoelectrode surface S_1 : the zones a and b of Fig. 3, at the tip, a transition zone c, and a long cylinder in zone b'.

that the electric potential field in zone b' (except near the two ends of the cylinder) is that produced by a (infinite) cylindrical electrode of radius r'_1 , and that in zone c is transitional between the values in zones b and b'. Clearly, the preceding model for zone b may be applied to zone b', using r'_1 and l'_1 instead of r_1 and l_1 .

An experimental curve of the current intensity $I(t)$ is shown in Fig. 19 (see Refs. 13 and 14 for the experimental procedure; potential of rectangular shape as in Fig. 4). This intensity represents the total intensity orig-

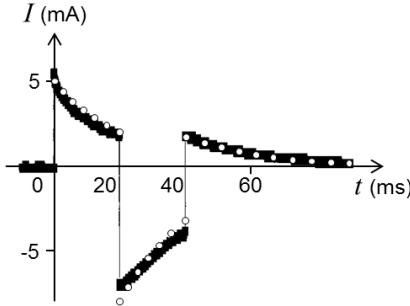


FIG. 19. Measured current intensity I as a function of time (thick black curve). Potential of rectangular shape with $V_1 = 35.6$ V and $t_1 = 20$ ms, solution with 10^{-3} mol/L of H_2SO_4 (experiment with nanolocalization of bubble production). The white dots represent the model presented in the paper.

inated from the various zones of the nanoelectrode surface. Nevertheless, according to Eq. (12) and owing to the low value of $\beta_2 a$ for zones a and b compared to that of zone b' (e.g., with $r_1 = 1$ μm , $l_1 = 10$ μm , $l = 1$ cm, $r'_1 = 50$ μm , and $l'_1 = 1$ cm: $\beta_{2a} a_a = 6.3$ μm , $\beta_{2b} a_b = 6.8$

μm , and $\beta_{2b'} a_{b'} = 1.2 \times 10^4$ μm), we may consider that I is mainly originated from zone b' (with some possible contribution of zone c). In this experiment, one bubble is produced and nanolocalized at the apex of the nanoelectrode. Assuming no electrolysis reaction in zone b', the above model applied to zone b' gives the current intensity $I(t)$ —with the help of Eqs. (12), (16), (22), and (29), and the discontinuity jumps of v at t_1 and $2t_1$ —, represented by white dots in Fig. 19, using the values $\frac{\gamma a}{\alpha d}(1 - \frac{1}{\alpha})V_1 = 5$ mA and $e^{-t_1/\tau} = 0.4$. Since $\frac{\gamma a}{\alpha d}(1 - \frac{1}{\alpha})V_1 \approx \frac{\gamma a}{\alpha d}V_1$ (owing to $\frac{\alpha-1}{\alpha} = \beta_1 d \approx 1$), this gives $\frac{a}{\alpha} \approx 0.54$ μm^2 (with $d = 0.3$ nm and $\gamma = 7.8 \times 10^{-2}$ S/m) and $\tau \approx 22$ ms, i.e., $\alpha = \frac{\tau}{\epsilon} \approx 2.4 \times 10^6$ (with $\epsilon = 80 \epsilon_0$) and $a \approx 1.3 \times 10^6$ μm^2 . These values are of the same magnitude as those corresponding to a cylinder b' of radius $r'_1 \approx 50$ μm and length $l'_1 \approx 1$ cm, i.e., $\alpha \approx 1 + \frac{r'_1}{d} \log \frac{l}{r'_1} \approx 0.9 \times 10^6$ (with $l \approx 1$ cm) and $a \approx 2\pi r'_1 l'_1 \approx 3 \times 10^6$ μm^2 , which shows that the model is in acceptable agreement with the experimental current intensity data.

The experimental black and white points in Fig. 20 correspond to the first visible occurrence of a bubble in zone a (apex of the electrode, here of micrometric curvature radius) and in zone b, respectively, while slowly increasing V_1 at constant t_1 . Since the ordinates of the bubble

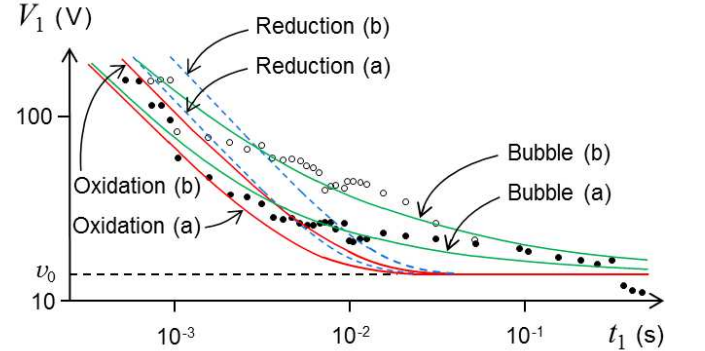


FIG. 20. Experimental first occurrence of a bubble in zone a (black points) and in zone b (white points), while slowly increasing V_1 at constant t_1 . Solution with 10^{-4} mol/L of H_2SO_4 . These experimental points are respectively compared with the bubble curves (of zone a and of zone b) of the model presented in the paper (green curves). The oxidation curves (red) and the reduction curves (dashed blue) of the model are also represented (for the two zones a and b). Logarithmic scale on both axes.

curves (of zone a and of zone b) of the above model tend to v_0 when $t_1 \rightarrow +\infty$, we obtain $v_0 \approx 14$ V from the experimental points with high values of t_1 . Such a value of v_0 applied to the 0.3 nm thick dielectric film corresponds to a typical ionization field observed in field ion microscopy²¹ and supports the model of electron transfer from water to platinum controlled by field. Moreover, we observed that the radius of the produced unique bubble (in zone a or in zone b, respectively) decreases when t_1

decreases (respectively following the black points or the white ones in Fig. 20). In the experiments with low values of t_1 , the observed bubble radius is extremely small, which probably means that these points are close to the oxidation curve (of zone a or of zone b; the oxidation curve corresponding to the first occurrence of an electrolysis reaction, when V_1 is increased at constant t_1). In Fig. 20, we thus obtain the oxidation curve of zone a (close to the experimental black points for the small values of t_1) for a value $r_1 \approx 14 \mu\text{m}$ (with $v_0 = 14 \text{ V}$, $d = 0.3 \text{ nm}$, $\varepsilon = 80\varepsilon_0$, and $\gamma = 7.8 \times 10^{-3} \text{ S/m}$; the oxidation curve being represented by $t_1 = t'$ and t' given by Eq. (17) with α_a and τ_a). It is known that the electric field at the apex of a tip of potential V and curvature radius r_1 at its apex (situated at large distance from a counter electrode) is $\approx V/(5r_1)$,^{22,23} i.e., the same as the electric field at the surface of a sphere of potential V and radius $\tilde{r}_1 \approx 5r_1$. The potential v being directly related to this electric field ($E_{1n} \approx v/d$, Eqs. (5) and (6)), our simple model of a sphere for zone a should be applied with a sphere radius $\tilde{r}_1 \approx 5r_1$ (r_1 being the curvature radius at the apex of the electrode). The above obtained value $\tilde{r}_1 \approx 14 \mu\text{m}$ thus corresponds to a curvature radius $r_1 \approx \tilde{r}_1/5 \approx 2.8 \mu\text{m}$ at the apex, in qualitative agreement with the micrometric curvature radius at the apex of our electrode. Moreover, this value $r_1 = 2.8 \mu\text{m}$ (with $l = 1 \text{ cm}$) leads to the oxidation curve of zone b represented in Fig. 20 (i.e., $t_1 = t'$ and t' given by Eq. (17) with α_b and τ_b) which, as expected and noted above, is close to the experimental white points for the small values of t_1 . The reduction curves of zones a and b (i.e., $t_1 = \bar{t}''$ and \bar{t}'' given by Eq. (24), using \tilde{r}_1 for zone a and r_1 for zone b) are also represented in Fig. 20. Note that, because of this value \tilde{r}_1 used in zone a (higher than the value r_1 of zone b), the reduction curve of zone a is close to the oxidation curve of zone b, for the values of t_1 represented in Fig. 20.

The bubble curves of zones a and b represented in Fig. 20 (i.e., Eq. (35), using \tilde{r}_1 for zone a and r_1 for zone b; $c = 0.21$, $H = 1.3 \text{ mol}/(\text{m}^3 \text{ atm})$ and $D = 2.5 \times 10^{-9} \text{ m}^2/\text{s}$ at 25°C), with a coefficient $s = 20$ for zone a and $s = 30$ for zone b, show a rather good agreement with the experimental points (i.e., the black points for the bubble curve of zone a and the white points for the bubble curve of zone b). Note that the production of a visible bubble is based on its nucleation and growth up to a micrometric size, which requires that a sufficient domain (in space and time) of the solution becomes supersaturated. This probably explains the obtained high values $s = 20\text{--}30$ which, according to Eq. (34), correspond to the maximum supersaturation of the solution at only the single point $(x, t) = (0, t_1)$.

In some experiments with similar conditions (same type of electrode, solution with 10^{-4} mol/L of H_2SO_4), the initial radius of the nanolocalized bubble (i.e., the unique bubble produced at the apex of the electrode; this radius then tends to decrease because the solution becomes undersaturated and the bubble dissolves in the

solution) was measured (green points in Fig. 21). These measurements are in qualitative agreement with the radius curves of the present model represented in Fig. 21 (i.e., for each radius value R , Eq. (37) below the reduction curve of zone a and Eq. (38) above this reduction curve, using the above values of v_0 , d , ε , γ , \tilde{r}_1 , and $\sigma = 7.2 \times 10^{-2} \text{ N/m}$ at 25°C). This confirms that the nanolocalized bubble is approximately only formed by the O_2 and H_2 molecules produced in zone a (and not those produced in zone b, the experimental points of Fig. 21 being situated above the oxidation and reduction curves of zone b shown in Fig. 20). Also in agreement with the observations, and as mentioned above after Fig. 17, the model shows that, following the bubble curve of zone a, in Fig. 21, the bubble radius decreases when t_1 decreases.

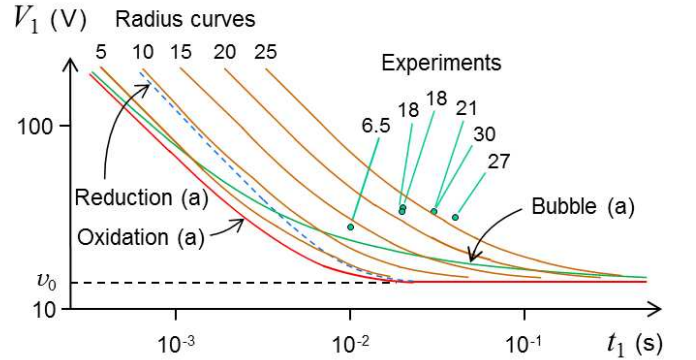


FIG. 21. Measured radii of nanolocalized bubbles (experimental green points; for each experiment, the number indicates the measured radius in μm). Solution with 10^{-4} mol/L of H_2SO_4 . These values may be compared with the radius curves of the model presented in the paper (brown curves; above each curve, the number indicates the radius value in μm). The oxidation (red), reduction (dashed blue), and bubble (green) curves of the model are also represented (for zone a). Logarithmic scale on both axes.

VI. CONCLUSION

We define nanoelectrolysis as the nanolocalization (at a single point) of electrolysis phenomena such as the electrolysis reactions or the production of bubbles. This paper presents a simple model of water nanoelectrolysis based on (i) the higher value of the electric field at the apex of a tip-shaped electrode (called the nanoelectrode; $5 \mu\text{m}$ to 1 nm of curvature radius at the apex) and (ii) the electron tunneling assisted by the electric field through the thin film of water molecules ($\sim 0.3 \text{ nm}$ thick) at the surface of the electrode. Using a finite time t_1 (related to the amplitude of the potential and to the geometry of the nanoelectrode) during which a positive (respectively, a negative) electric potential V_1 is applied between the

electrodes, we show that the electrolysis oxidation reaction which produces O_2 molecules (respectively, the reduction reaction which produces H_2 molecules) can be nanolocalized at the apex of the nanoelectrode. These two parameters t_1 and V_1 determine the time at which the electrolysis reaction begins, the electrolysis current intensity (i.e., the tunneling current intensity), the electric charge delivered by this current, and then the number of produced O_2 or H_2 molecules (O_2 if $V_1 > 0$, H_2 if $V_1 < 0$). It is also shown that, during the electrolysis reaction (oxidation or reduction), the current intensity is equal to the electrolysis current intensity (assuming a high “tunneling conductivity” $\gamma_0 = dj/dE$ above a threshold value E_0 of the electric field). The model then determines the concentration of O_2 or H_2 molecules in the solution close to the surface of the nanoelectrode (as a consequence of the electrolysis current and the diffusion of these molecules in the solution). A bubble is then produced if the maximum of this concentration reaches some supersaturation value.

This general model is applied to the simple case of a potential (between the electrodes) of rectangular shape: $V(t) = V_1$ for $0 < t < t_1$, $V(t) = -V_1$ for $t_1 < t < 2t_1$, and $V(t) = 0$ for $t < 0$ and $t > 2t_1$. If, e.g., $V_1 > 0$, it is shown that the duration of the electrolysis reduction reaction is lower than that of the oxidation reaction (when these two reactions occur) and the number of produced H_2 molecules is lower than twice the number of produced O_2 molecules (if a unique threshold value E_0 of the electric field is used¹⁹). Moreover, in the plane (t_1, V_1) , the region for the nanolocalization of the reduction reaction is shifted toward higher values of t_1 or V_1 , with respect to that for the nanolocalization of the oxidation reaction. In the same plane, we show that there is a third distinct region, that for the nanolocalization of the production of bubbles (also shifted toward higher values of t_1 or V_1 , with respect to that for the nanolocalization of the oxidation reaction). From the number of produced O_2 and H_2 molecules, the model also determines the radius of the produced (unique) nanolocalized bubble, as a function of (t_1, V_1) . In the plane (t_1, V_1) , following the “bubble curve of the apex zone”, i.e., at the limit between the no bubble and the nanolocalized bubble regions, the radius of the nanolocalized bubble decreases when t_1 decreases.

In addition to the nanolocalization of the production of bubbles (first observed in Ref. 13), the model also explains why there may be nanolocalization of an electrolysis reaction (oxidation or reduction) whereas no bubble is produced (as was assumed in our previous paper about the immobilization of a bubble¹⁴). Moreover, the above results of the model are in rather good agreement with our recent experiments with a potential of rectangular shape, concerning the current intensity, the region for the nanolocalization of the production of bubbles (in the plane (t_1, V_1)), and the radii of the nanolocalized bubbles. We also currently try to measure the current intensity originated only from the apex of the nanoelectrode, which, according to the present model, is equal to

the electrolysis current intensity (i.e., the tunneling current; during the electrolysis reaction). In conclusion, the model shows that, by means of the two parameters t_1 and V_1 , we can control (i) the nanolocalization of the oxidation reaction, (ii) the nanolocalization of the reduction reaction, or (iii) the nanolocalization of the production of bubbles, but also the time at which the electrolysis reaction (of oxidation or reduction) begins, the duration of this reaction, the electrolysis current intensity, the number of produced O_2 or H_2 molecules, and the radius of the nanolocalized bubble.

Appendix A: Spherical and cylindrical electrodes

1. Spherical electrodes

In this case, S_1 , S_{12} and S_2 are concentric spheres, of respective radii r_1 , $r_1 + d$ and $r_1 + d + l$ (in fact, if $l \gg r_1 + d$, the form of S_2 has no significant effect on the potential field in the neighbourhood of S_1 and S_{12}). We may consider that the potential is zero on S_2 . According to Maxwell’s equations, at a given time t , the potential φ satisfies to

$$\Delta\varphi = \frac{d^2\varphi}{dr^2} + \frac{2}{r}\frac{d\varphi}{dr} = \frac{1}{r}\frac{d^2(r\varphi)}{dr^2} = 0$$

(r being the distance to the centre of the spheres), hence

$$\frac{d(r\varphi)}{dr} = \varphi + r\frac{d\varphi}{dr} = a$$

(a constant; thus, $r\varphi = ar + b$, with b constant, i.e., $\varphi = a + \frac{b}{r}$). The preceding equation applied in region V_1 , between S_1 and S_{12} ,

$$\left(\varphi + r\frac{d\varphi}{dr}\right)_1 = \frac{\delta(r\varphi)}{\delta r}$$

(the subscript 1 indicates the value on the V_1 side of S_{12} , and δ the value on S_{12} minus that on S_1), i.e.,

$$\varphi_{12} + (r_1 + d)\left(\frac{d\varphi}{dr}\right)_1 = \frac{(r_1 + d)\varphi_{12} - r_1\varphi_1}{d}$$

($\varphi_1 = V$ and $\varphi_{12} = V - v$ respectively being the potentials on S_1 and S_{12}), leads to

$$E_{1n} = -\left(\frac{d\varphi}{dr}\right)_1 = \frac{r_1(\varphi_1 - \varphi_{12})}{(r_1 + d)d} = \frac{r_1}{(r_1 + d)d} v.$$

Similarly, the same equation applied in region V_2 , between S_{12} and S_2 ,

$$\left(\varphi + r\frac{d\varphi}{dr}\right)_2 = \frac{\delta(r\varphi)}{\delta r}$$

(the subscript 2 indicates the value on the V_2 side of S_{12} , and δ the value on S_2 minus that on S_{12}), i.e.,

$$\varphi_{12} + (r_1 + d)\left(\frac{d\varphi}{dr}\right)_2 = \frac{-(r_1 + d)\varphi_{12}}{l},$$

leads to

$$E_{2n} = -\left(\frac{d\varphi}{dr}\right)_2 = \frac{(r_1 + d + l)\varphi_{12}}{(r_1 + d)l} = \frac{r_1 + d + l}{(r_1 + d)l} (V - v).$$

2. Cylindrical electrodes

Here, S_1 , S_{12} and S_2 are coaxial (infinite) cylinders, of respective radii r_1 , $r_1 + d$ and $r_1 + d + l$ (here also, if $l \gg r_1 + d$, the form of S_2 has no significant effect on the potential field in the neighbourhood of S_1 and S_{12}). We may consider that the potential is zero on S_2 . According to Maxwell's equations, at a given time t , the potential φ satisfies to

$$\Delta\varphi = \frac{d^2\varphi}{dr^2} + \frac{1}{r} \frac{d\varphi}{dr} = \frac{1}{r} \frac{d}{dr} \left(r \frac{d\varphi}{dr} \right) = 0$$

(r being the distance to the axis of the cylinders), hence

$$r \frac{d\varphi}{dr} = \frac{d\varphi}{d \log r} = a$$

(a constant; thus, $\varphi = a \log r + b$, with b constant). The preceding equation applied in region V_1 , between S_1 and S_{12} ,

$$\left(r \frac{d\varphi}{dr} \right)_1 = \frac{\delta\varphi}{\delta \log r}$$

(same notations as above, for spherical electrodes), i.e.,

$$(r_1 + d) \left(\frac{d\varphi}{dr} \right)_1 = \frac{\varphi_{12} - \varphi_1}{\log \frac{r_1 + d}{r_1}},$$

leads to

$$E_{1n} = - \left(\frac{d\varphi}{dr} \right)_1 = \frac{v}{(r_1 + d) \log \frac{r_1 + d}{r_1}}.$$

Similarly, the same equation applied in region V_2 , between S_{12} and S_2 ,

$$\left(r \frac{d\varphi}{dr} \right)_2 = \frac{\delta\varphi}{\delta \log r}$$

(same notations as for spherical electrodes), i.e.,

$$(r_1 + d) \left(\frac{d\varphi}{dr} \right)_2 = \frac{-\varphi_{12}}{\log \frac{r_1 + d + l}{r_1 + d}},$$

leads to

$$E_{2n} = - \left(\frac{d\varphi}{dr} \right)_2 = \frac{V - v}{(r_1 + d) \log \frac{r_1 + d + l}{r_1 + d}}.$$

Appendix B: Proof of $v(2t_1^+) < 0$

1. Case $t_1 < t'$

According to Eqs. (16) and (22), and the discontinuity jumps at t_1 and $2t_1$, we obtain

$$v(2t_1^+) = -(V_1 - \frac{V_1}{\alpha})(1 - e^{-t_1/\tau})^2 < 0.$$

2. Case $t' < t_1 < t''$

According to Eqs. (18) and (22), and the discontinuity jumps at t_1 and $2t_1$, we obtain

$$v(2t_1^+) = -V_1 + \frac{V_1}{\alpha} + (V_1 + V_{1+} - 2\frac{V_1}{\alpha} - (V_{1+} - v_0)e^{-(t_1-t')/\tau'})e^{-t_1/\tau},$$

which is a function of t_1 (at V_1 constant). Then

$$\begin{aligned} \frac{d(v(2t_1^+))}{dt_1} &= \frac{e^{-t_1/\tau}}{\tau} \times \\ &((V_1 - v_0 + V_{1+} - v_0)e^{-(t_1-t')/\tau'} - V_1 - V_{1+} + 2\frac{V_1}{\alpha}) \\ &< -\frac{e^{-t_1/\tau}}{\tau} 2(v_0 - \frac{V_1}{\alpha}) < 0 \end{aligned}$$

(using $e^{-(t_1-t')/\tau'} < 1$), which shows that $v(2t_1^+)$ is a decreasing function of t_1 . Thus, $v(2t_1^+)$ (at t_1) is lower than its value at $t_1 = t'$ (since $t_1 > t'$):

$$\begin{aligned} v(2t_1^+) &< -(1 - \frac{1}{\alpha})V_1 + ((1 - \frac{2}{\alpha})V_1 + v_0)e^{-t'/\tau} \\ &= -\frac{(v_0 - \frac{V_1}{\alpha})^2}{(1 - \frac{1}{\alpha})V_1} < 0 \end{aligned}$$

(using the value of t' given by Eq. (17)).

3. Case $t'' < t_1$

In this case, $v(2t_1^-) < -v_0$, then $v(2t_1^+) < -v_0 + \frac{V_1}{\alpha} < 0$.

Appendix C: Solution of the diffusion equation

The one-dimensional diffusion equation

$$\frac{\partial \rho}{\partial t}(x, t) - D \frac{\partial^2 \rho}{\partial x^2}(x, t) = \delta_0(x) \delta_{t_0}(t)$$

(δ_u denotes the Dirac measure at u) has the known elementary solution in $\mathbf{R} \times]t_0, +\infty[$

$$\rho(x, t) = \frac{e^{-x^2/(4D(t-t_0))}}{(4\pi D(t-t_0))^{1/2}}.$$

In our case, the surface S_{12} being assimilated to the plane $x = 0$ and the aqueous solution occupying the region $x > 0$, ρ does not depend on the other coordinates y and z , and the elementary problem of diffusion

$$\frac{\partial \rho}{\partial t}(x, t) - D \frac{\partial^2 \rho}{\partial x^2}(x, t) = j_m dt_0 \delta_0(x) \delta_{t_0}(t)$$

has the similar solution in $\mathbf{R}_+^* \times]t_0, +\infty[$

$$\rho(x, t) = 2j_m dt_0 \frac{e^{-x^2/(4D(t-t_0))}}{(4\pi D(t-t_0))^{1/2}},$$

with the coefficient 2 because the produced molecules diffuse in the half space $x > 0$. The integration on $t_0 \in [t', t_1]$ then leads to Eq. (31) for $t_1 \leq t$. The case $t' < t < t_1$ is simply obtained by applying this equation at $t = t_1$ and then replacing t_1 with t .

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