



The Essential Role for Laboratory Studies in Atmospheric Chemistry

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Abstract

Laboratory studies of atmospheric chemistry characterize the nature of atmospherically relevant processes down to the molecular level, providing fundamental information used to assess how human activities drive environmental phenomena such as climate change, urban air pollution, ecosystem health, indoor air quality, and stratospheric ozone depletion. Laboratory studies have a central role in addressing the incomplete fundamental knowledge of atmospheric chemistry. This article highlights the evolving science needs for this community and emphasizes how our knowledge is far from complete, hindering our ability to predict the future state of our atmosphere and to respond to emerging global environmental change issues. Laboratory studies provide rich opportunities to expand our understanding of the atmosphere via collaborative research with the modeling and field measurement communities, and with neighbouring disciplines.

Introduction

Atmospheric chemistry is the study of the chemical processes that affect the composition of the atmosphere, encompassing societally important issues such as air pollution and its related health and ecosystem effects, as well as climate (see Figure 1). This field unifies the evaluation of both natural and anthropogenic emissions, measurements of atmospheric composition across a range of temporal and spatial scales, and assessment of deposition processes that ultimately remove chemical compounds from the air. All of these processes are impacted by, or may feedback on, climate change. At the heart of atmospheric chemistry are chemical transformations, often initiated by sunlight. Molecules absorb solar radiation, leading to a cascade of catalytic and oxidative reactions. Our molecular-level understanding of atmospheric chemistry has

106 successfully provided the foundations of air quality forecasts including those related to acid
107 rain,¹ the basis upon which the chemicals responsible for the formation of the Ozone Hole were
108 banned under the Montreal Protocol and its amendments,² and connections between human
109 impact on atmospheric composition and climate change.³ The evaluation of geoengineering as an
110 option to offset global warming also requires a detailed understanding of the underlying
111 chemistry.⁴

112 Simplified parameterizations of this detailed chemistry developed from well-controlled
113 laboratory studies, alongside representations of emissions, transport and deposition processes,
114 are incorporated into computer models to predict atmospheric composition across appropriate
115 temporal and spatial (local, regional, and global) scales. Comparisons with measurements from
116 the field are often used to evaluate model predictions or to help identify poorly represented
117 chemistry. Laboratory studies are an essential bridge between field measurements and models
118 and provide the basic (or fundamental) physics and chemistry of the underlying mechanisms of
119 phenomena observed in the field, which is the basis for generating physically meaningful
120 parameterizations for use in models. This interplay between laboratory studies of chemical,
121 photochemical, and physical processes, atmospheric modeling, and field measurements has
122 propelled the field of atmospheric chemistry forward to achieve remarkable success in
123 understanding the details of stratospheric ozone depletion, acid rain, urban air quality, and the
124 chemistry of climate-forcing agents, and in informing environmental policies (see Figure 2).

125 The increased emphasis by funding agencies on solutions-driven research, as opposed to
126 fundamental research of the atmosphere in general, inherently demotivates the development of a
127 comprehensive understanding of the core processes occurring in the atmosphere. In particular,
128 attention and resources are funneled into important but highly specific sub-fields, such as

detailed studies of the atmospheric impacts of different energy source activities including light oil extraction, hydraulic fracturing or bitumen extraction.⁵ Laboratory studies, together with field and modeling studies, constitute the long-standing “three-legged stool” of atmospheric chemistry. The fundamental science aspect of this “three-legged stool” is overshadowed when the science is solutions-driven, lessening our ability to not only address known environmental issues such as the relationship of air quality to human health but also our readiness to respond to unforeseen future environmental threats. Moreover, the complexity of the environment demands avenues both for discovery-based science and for developing and testing hypotheses under well controlled and relevant conditions. Despite the successes mentioned above, the atmosphere is far from being fully understood with regard to changes in composition and the complex interplay of chemical and physical processes, and thus predictive capabilities are limited. The importance of fundamental understanding has been stressed previously,⁶ most recently in the recent report on the Future of Atmospheric Chemistry Research prepared for the National Academy of Science in the United States.⁷

This article highlights the significant role of laboratory-based atmospheric chemistry research, pointing out that our ability to respond to a changing environment and to accurately inform policy development hinges upon a fundamental molecular-level understanding of these processes. Additionally, a commitment from universities, research centers, and governmental agencies to support such research activities in a meaningful way and to promote innovative and interdisciplinary research is needed at this time of a rapidly changing atmospheric composition. Several key challenges and opportunities facing the laboratory community are presented in this article.

Emerging Topics for Future Laboratory Studies in Atmospheric Chemistry

The atmospheric chemistry community has traditionally relied on laboratory studies to determine the rates and mechanisms of key gas-phase reactions. For example, these processes have been shown to be important for better understanding ozone depletion in the stratosphere and the coupled processes of tropospheric oxidant production and organic oxidation that control the abundance of key climate forcing agents such as methane (CH₄) and dimethyl sulfide (DMS).^{2,8} There is currently considerable focus on the chemistry that occurs within or on aerosol particles and cloud droplets, and increasingly on the interface of the atmosphere with the oceans, biosphere, cryosphere, and indoor environments.⁹⁻¹³

Recent advances that highlight how far we have come in this field but, more importantly, how much further we have to go to develop sufficiently accurate and predictive models of the atmosphere are presented below. The examples below demonstrate that chemistry is central to our understanding of the field and that we once thought to be understood, such as volatile organic carbon (VOC) oxidation mechanisms, is much more varied and complex than described by current mechanisms and included in models.

i. Do we understand how organic molecules are oxidized in the atmosphere?

Our conceptual view of organic oxidation mechanisms (see Figure 3), central to our understanding of smog formation, the lifetime of greenhouse gases such as ozone (O₃) and methane, and the climate effects of aerosol particles, has been transformed in the past few years. An example of such a transformation is illustrated by recent studies of the chemistry of key biogenic hydrocarbons, isoprene and the monoterpenes, that are emitted from vegetation. Although isoprene is the largest source of organic carbon to the atmosphere and even a component of exhaled human breath, our understanding of how it is oxidized under appropriate

radical concentration conditions is still incomplete.¹⁴ Only a short time ago, it was not known whether it could contribute to aerosol formation, but recent advances now indicate it is an important aerosol precursor (Figure 3).^{15, 16} In particular, when the OH radical adds to one of isoprene's carbon-carbon double bonds a variety of highly oxidized products form.¹⁷ The radical intermediates are reactive, and can isomerize leading to efficient autocatalytic oxidation and multi-functionalised compounds that may form aerosol particles.¹⁸ Similarly, α -pinene, a monoterpene, is converted into "highly oxidized molecules (HOMs)" (or "extremely low volatility organic compounds", ELVOCs) under atmospheric conditions on timescales of seconds to minutes, with up to 10 oxygen atoms being introduced onto the terpene backbone.¹⁹ These are all new findings.

Major uncertainties in the oxidation mechanisms of VOCs, particularly biogenic VOCs, still remain. In particular, there is a need to fully establish how much recycling of the OH radical occurs from isoprene oxidation under different atmospheric conditions. This is especially true in environments with significant input of molecules from biogenic sources, where the impact of OH recycling on oxidant levels may be significant.²⁰ Also important toward understanding the impacts of VOC oxidation are the roles of epoxides and the HOM compounds in the formation of biogenic secondary organic aerosol (SOA), new particles,²¹ and the potential health effects of particles containing these highly oxidized species. Indeed, the recognition that such organics may play a major role in atmospheric particle nucleation and growth processes nicely illustrates another connection between fundamental physical chemistry processes and atmospheric behavior.

Little is known about the chemistry of these highly oxidized multifunctional reaction products with respect to subsequent gas-phase and condensed-phase reactivity, photochemistry and light-

absorbing properties, hygroscopicity, and volatility. Further complexity arises through the addition of nitrogen-based functional groups to the molecule via reactions involving nitrogen oxide radicals, amines, or ammonia. Light-absorbing compounds can result that contribute to the colored particles collectively referred to as brown carbon aerosol and known to arise from burning processes. This has implications for climate change given that biomass burning and wildfire emission sources are expected to increase in the future.²²

Our framework for understanding the multiphase chemistry of when these oxidized products partition from the gas phase to cloud water or an aerosol particle is in its infancy. Questions that arise include: How important are radical processes compared to non-radical nucleophilic/electrophilic, hydrolytic, and addition/condensation reactions?²³ How important are non-ideal solution effects and how can they best be treated to help interpret field investigations and improve representations in models?²⁴ These issues can be addressed with focused laboratory studies coming out of the coupled atmospheric-physical-analytical chemistry laboratory communities.

ii. What controls the major oxidants in the atmosphere?

The chemistry of reactive chemical intermediates – usually radicals – is at the heart of atmospheric chemistry (see Figure 4). Although our understanding of reaction mechanisms initiated by conventional oxidants (OH, Cl, O₃) is extensive, it is not complete and the importance of new classes of reactive intermediates, e.g. Criegee intermediates, is now being recognized and studied. For example, although the rapid cycling of OH and HO₂ radicals helps to drive the chemistry that forms multifunctional organic compounds, the field still does not yet have closure between measured and modeled OH concentrations in clean environments.^{20, 25}

Beyond gas-phase OH, there has been a recent surge of activity in the chemistry of other reactive intermediates, notably nitrate radicals,²⁶ halogen oxide radicals, Criegee intermediates, and compounds active as photosensitizers,²⁷ but our understanding of these processes remains largely insufficient. The Criegee intermediates present a compelling story, having recently been observed for the first time via advanced mass spectrometry techniques.^{28, 29} These reactive species are formed when ozone reacts in the gas phase with carbon-carbon double bonds, and their involvement in the oxidation of key molecules, such as SO₂, thus affecting new particle formation, is now being studied.

iii. What is the atmospheric importance of multiphase chemistry?

Some of the largest uncertainties in our field arise from the chemistry and photochemistry that takes place involving different interfaces and media, such as aerosol particles and cloud droplets, that interact with gas phase constituents. Collectively, this is referred to as multiphase chemistry (see Figure 5).³⁰

The organic aerosol medium represents a poorly characterized component of particle chemistry that can affect partitioning of gas-phase organics and can promote chemistry distinctly different from that which occurs in other chemical environments. Aqueous phase droplets and particles can enhance the formation of organic aerosol by promoting soluble organic partitioning and subsequent reactions.³¹ Constraining advances in the field of organic particle and mixed organic-inorganic particle chemistry is our lack of detailed knowledge of the bulk composition, phase state, morphology, transport properties, interfacial composition, and photochemistry within these particles under environmental conditions.^{27, 32} The development of new methodologies, such as spectroscopic and mass spectrometric approaches, that are able to better characterize these

properties at the individual particle level is a key challenge. In addition, limitations currently exist in our ability to prepare appropriately complex model substances in the laboratory that realistically resemble those present in the environment.

Since particles can indirectly modify climate by acting as cloud condensation nuclei (CCN) and ice nuclei (IN) an important area of multiphase chemistry involves studies of the hygroscopicity of atmospheric particles.³³ Although much has been learned about the relationship of aerosol particle composition to hygroscopicity and CCN activity, the roles of interfacial processes and organic substituents are still being resolved.^{34, 35} Our understanding of the IN activity of aerosol particles is even less refined because we lack a first-principles model that can accurately predict the rates of heterogeneous ice nucleation as a function of aerosol type and composition.³⁶

The atmosphere interacts with the surfaces of the oceans, vegetation, soil, and indoor environments. To illustrate such interactions, a long-standing uncertainty in atmospheric chemistry has been related to non-gas phase sources of HONO, a photolytic source of the OH radical. To explain measured levels within many boundary layer environments, it is known that during the day there is a source that forms HONO much faster than the traditional gas-phase formation route.^{37, 38} A variety of mechanisms have been proposed, all involving interfacial chemistry.^{39, 40, 41, 42}

A highly interdisciplinary frontier lies with the impacts of atmospheric particles on human health. Reactive oxygen species (ROS), such as hydrogen peroxide or organic hydroperoxides, are a class of molecules that can give rise to oxidative stress, the state where the human body's oxidant and anti-oxidant balance is disrupted.⁴³ Also important to oxidative stress are molecules with labile oxidation states and metals that promote the formation of ROS species, either in

inhaled atmospheric particles or the body.⁴⁴⁻⁴⁶ A key question is the degree to which these molecules drive oxidative stress in the body and whether their sources are endogenous or exogenous. Moving ahead in this field requires the collaboration of atmospheric chemists and toxicologists.

Challenges and Opportunities in Laboratory Atmospheric Chemistry Research

i. Complexity

The atmosphere has many chemical constituents and processes interacting in a non-linear manner. The complexity that arises in the system has typically been addressed by a reductionist approach where every relevant rate constant and photochemical property is measured. For example, this approach is used to predict the non-linear response of urban ozone to changing NO_x emissions, a relationship now well-recognized in the air pollution control community and with connections to climate change.^{8, 47} But other complex interactions are less well understood. For example, how do chemical interactions with aerosol particles affect the aerosols' ability to nucleate liquid water and ice clouds? How does the liquid-like layer that exists at the surface of snow affect the likelihood of molecules being sorbed from the gas phase?⁴⁸ How does biology affect the indoor microbiome and its ability to change VOC levels?¹³ And, what is the interplay between biology and chemistry in controlling the fluxes of climatically-active gases such as DMS and isoprene to the atmosphere from the ocean and vegetation.^{9, 12}

A reductionist approach is necessary to arrive at the understanding required for confidence in model predictions of air quality and climate. Nevertheless, it is also fruitful to work with top-down methods where observations from laboratory experiments conducted under conditions close to those in the environment can be parameterized for inclusion in models. For example, it

is now possible to measure the rate of loss of a gas-phase reactant to ambient aerosol under real-world conditions.⁴⁹ The goal of such experiments is to obtain quantitative closure between the bottom-up and top-down approaches for assessing the rates of this chemistry.

ii. Collaboration and Interdisciplinary Research

In addition to the value of single-investigator science, atmospheric chemistry has also long benefited from collaboration. A number of schemes can enhance opportunities for collaboration. In particular, there is an advantage to incorporating laboratory projects into the funding of large field campaigns, to facilitate exchange of ideas and experimental techniques from one community to another. Another approach is for a small number of research teams to work together to simultaneously study the chemistry in one facility, to examine the nature of complex processes.⁵⁰ Recent advances in theoretical methods mean that quantum chemical calculations can significantly enhance knowledge gained from laboratory investigations. In addition, collaborations between laboratory scientists and atmospheric modelers are required to enable both direct impact-testing of laboratory data and identify areas of research for which remaining uncertainty has significant repercussions.

New interdisciplinary opportunities are continuing to arise at the interfaces of atmospheric chemistry. For example, dedicated lab experiments under controlled conditions will inform us regarding the nature of biosphere-atmosphere interactions.⁵¹ This, in turn, will help us to better understand how ecosystem health is affected via atmospheric exposures and will aid in determining the detailed mechanisms by which vegetation removes oxidants from the atmosphere. Another example is the chemistry involving atmospheric constituents and the lung-air interface. Experiments conducted jointly by atmospheric chemists and toxicologists will

307 better establish how airborne particles contribute to increased levels of oxidative stress, a
308 common hypothesis invoked to explain the epidemiological connections between increased
309 particulate loadings and negative health outcomes.⁵² There is also considerable scope for new
310 explorations in indoor environments where we spend most of our time and receive most of our
311 pollutant exposure.¹³ Indoor surface-area-to-volume ratios are high, suggestive of the importance
312 of surface chemistry, and many cleaning and personal care products whose constituents partition
313 between the gas phase and indoor surfaces are prevalent.⁵³ Furthermore, the desire to reduce
314 energy consumption in some modern buildings may lead to less ventilated indoor environments,
315 and requires better understanding of indoor air chemistry to ensure that such green buildings are
316 also healthy buildings.

317 For all these opportunities, there is the need for science funding agencies to recognize that such
318 interdisciplinary research often falls between the cracks of funding programs that are better tuned
319 to promote the value of disciplinary research.

320
321 iii. The interplay of laboratory experiments with computational and atmospheric modeling
322 scientists

323 Connecting laboratory experiments to chemical theory remains crucial to the atmospheric
324 chemistry field. At the molecular level, as computational methods allow for increasingly
325 complex chemistry to be studied on a computer,⁵⁴ interactions between laboratory and
326 computational chemists are expected to become more common.

327 Individual gas-phase rate or photochemical parameters may be easily incorporated into chemical
328 transport models and their impact evaluated. However, the challenge of transferring laboratory

results from complex systems is not as straightforward. The evaluation of physical and chemical laboratory data as provided by the NASA JPL (<http://jpldataeval.jpl.nasa.gov/>) and IUPAC (<http://iupac.pole-ether.fr/>) panels is a particularly important task that supports feedback between modelers and experimentalists in both directions.^{55, 56} Collaborative efforts between laboratory scientists and process-level atmospheric modelers can ensure accurate parameterizations of complex chemistry are incorporated into models, and the integrating nature of the models can provide information about which conditions or timescales need further experimental constraints. Such collaborations can occur as part of research grants, information collaborations, or special sessions at international conferences that emphasize the integration of modeling and laboratory results.

iv. The interplay of laboratory experiments with field measurements

Many advanced analytical technologies have been developed in the past decade that can be applied to atmospheric chemistry research, with increasingly sensitive and multiplexing instruments deployed in both the lab and atmosphere. In order to understand new field observations the laboratory community is pushed to develop a better understanding of the associated chemistry and its possible environmental impacts. For example, recent progress has arisen from the advent of advanced mass-spectrometric techniques that now allow for the identification and quantification of species present in complex chemical mixtures.^{57, 58} Such instrumentation has been used extensively in new studies of particle nucleation and growth processes.²¹ With a strong demand for sensitive and reliable measurement technologies, the atmospheric chemistry field fosters developments in advanced measurement technologies. These advances have led to a recent growth in experimental studies where the relevant chemistry is studied on genuine atmospheric materials (such as atmospheric aerosol particles) commonly

referred to as “field-in-the-lab” or “lab-in-the-field” methods.⁴⁹ These approaches are attractive to test the applicability of laboratory model materials and to provide top-down results that fundamental-based bottom-up approaches aspire to match.

Conclusions

The scientific uncertainties illustrated above are associated with some of the most central questions in atmospheric chemistry: How are atmospheric molecules transformed, and by what mechanisms? In what phase does this chemistry occur and on what timescale? How does this chemistry affect air quality and climate? How do ecosystems affect atmospheric chemistry and vice versa? How does this chemistry affect the interaction of the atmosphere with other parts of the environment? These issues and others highlight the ongoing central role for laboratory studies and a molecular-level understanding of atmospheric chemistry that enable the development of informed environmental policy.

The laboratory studies community is adapting to address far greater scientific complexity than was apparent only a decade ago, needing to embrace an interdisciplinary and collaborative research approach while at the same time continuing to focus on the measurement of fundamental properties of atmospherically relevant molecules and processes. These aspects of the field are highly attractive to early career scientists who are looking for research experiences outside of traditional disciplines. Funding agencies can facilitate these efforts, by supporting fundamental laboratory science and by promoting interdisciplinary and collaborative research and the interplay between laboratory studies and associated modeling, and field measurement activities.

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References

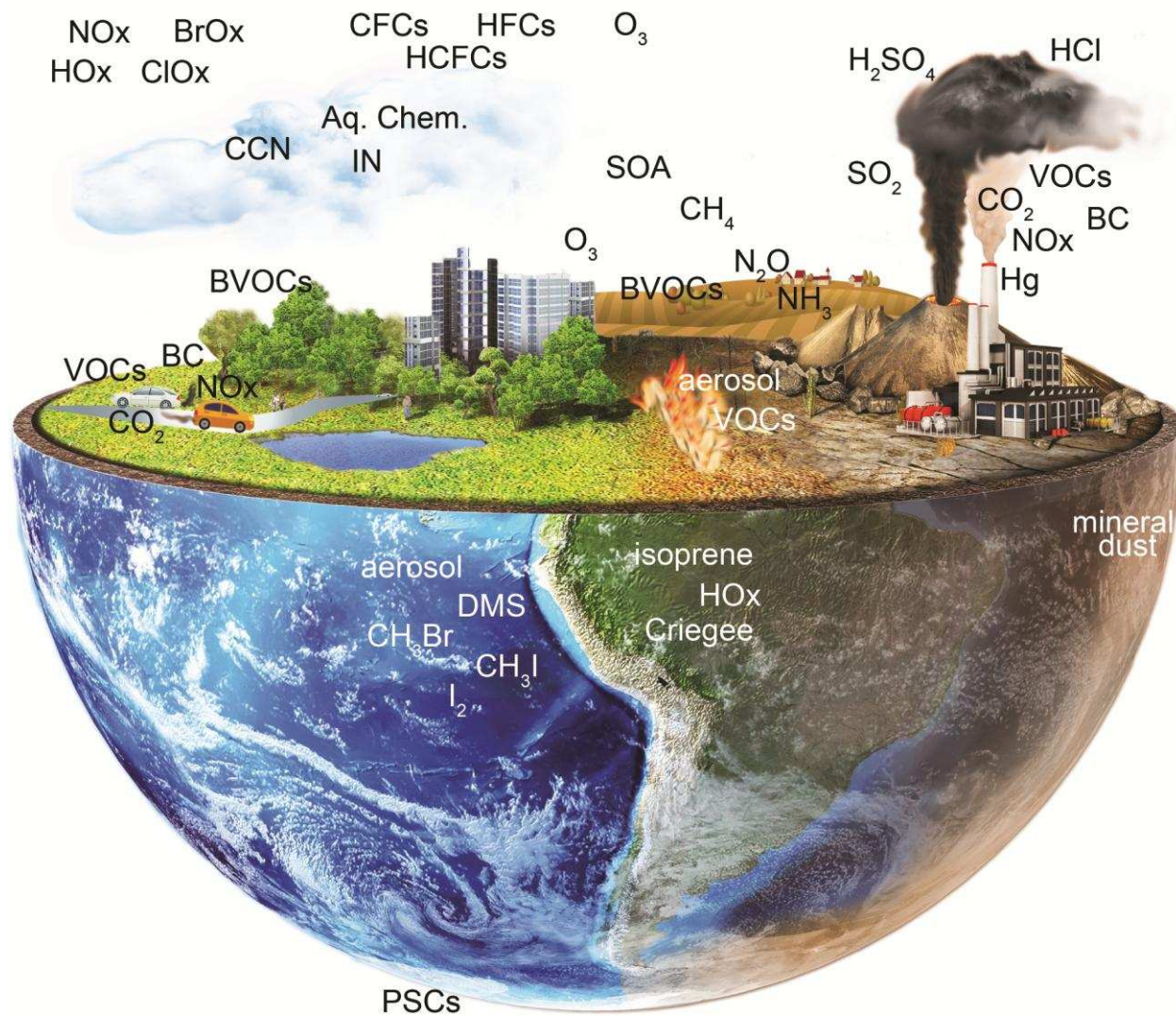
1. EPA, Environmental Protection Agency (US), <https://www.epa.gov/acidrain>.
2. WMO, World Meteorological Organization, Scientific Assessment of Ozone Depletion: 2014. *Global Ozone Research and Monitoring Project 2014, Report 55*, 55 pp., Geneva, Switzerland.
3. IPCC, Climate Change 2013: The Physical Science Basis. *Contribution of Working Group 1 to the Fifth Assessment Report of the Intergovernmental Panel on Climate Change, Stocker et al. (eds) 2013*, Cambridge University Press, Cambridge, UK and New York, USA.
4. Vaughan, N. E.; Lenton, T. M., A review of climate geoengineering proposals. *Climatic Change* **2011**, *109*, 745-790.
5. Liggio, J.; Li, S. M.; Hayden, K.; Taha, Y. M.; Stroud, C.; Darlington, A.; Drollette, B. D.; Gordon, M.; Lee, P.; Liu, P., et al., Oil sands operations as a large source of secondary organic aerosols. *Nature* **2016**, *534*, 91-95.
6. Abbatt, J.; George, C.; Melamed, M.; Monks, P.; Pandis, S.; Rudich, Y., New Directions: Fundamentals of atmospheric chemistry: Keeping a three-legged stool balanced. *Atmos. Environ.* **2014**, *84*, 390-391.
7. National Academies of Sciences, Engineering, and Medicine. 2016. The Future of Atmospheric Chemistry Research: Remembering Yesterday, Understanding Today, Anticipating Tomorrow. Washington, DC: The National Academies Press. DOI: 10.17226/235730.
8. Monks, P. S.; Archibald, A. T.; Colette, A.; Cooper, O.; Coyle, M.; Derwent, R.; Fowler, D.; Granier, C.; Law, K. S.; Mills, G. E., et al., Tropospheric ozone and its precursors from the urban to the global scale from air quality to short-lived climate forcer. *Atmos. Chem. Phys.* **2015**, *15*, 8889-8973.
9. Carpenter, L. J.; Nightingale, P. D., Chemistry and Release of Gases from the Surface Ocean. *Chem. Rev.* **2015**, *115*, 4015-4034.
10. Abbatt, J. P. D.; Thomas, J. L.; Abrahamsson, K.; Boxe, C.; Granfors, A.; Jones, A. E.; King, M. D.; Saiz-Lopez, A.; Shepson, P. B.; Sodeau, J., et al., Halogen activation via interactions with environmental ice and snow in the polar lower troposphere and other regions. *Atmos. Chem. Phys.* **2012**, *12*, 6237-6271.
11. McNeill, V. F.; Grannas, A. M.; Abbatt, J. P. D.; Ammann, M.; Ariya, P.; Bartels-Rausch, T.; Domine, F.; Donaldson, D. J.; Guzman, M. I.; Heger, D., et al., Organics in environmental ices: sources, chemistry, and impacts. *Atmos. Chem. Phys.* **2012**, *12*, 9653-9678.
12. Heald, C. L.; Spracklen, D. V., Land Use Change Impacts on Air Quality and Climate. *Chem. Rev.* **2015**, *115*, 4476-4496.
13. Weschler, C. J., Chemistry in indoor environments: 20 years of research. *Indoor Air* **2011**, *21*, 205-218.
14. Wennberg, P., Let's abandon the high NO_x and low NO_x terminology *IGAC Newsletter* **2013**, *50*, 3-4.
15. Surratt, J. D.; Chan, A. W. H.; Eddingsaas, N. C.; Chan, M. N.; Loza, C. L.; Kwan, A. J.; Hersey, S. P.; Flagan, R. C.; Wennberg, P. O.; Seinfeld, J. H., Reactive intermediates revealed in secondary organic aerosol formation from isoprene. *Proc. Nat. Acad. Sci. U.S.A.* **2010**, *107*, 6640-6645.
16. Lin, Y. H.; Zhang, Z. F.; Docherty, K. S.; Zhang, H. F.; Budisulistiorini, S. H.; Rubitschun, C. L.; Shaw, S. L.; Knipping, E. M.; Edgerton, E. S.; Kleindienst, T. E., et al., Isoprene Epoxydiols as Precursors to Secondary Organic Aerosol Formation: Acid-Catalyzed Reactive Uptake Studies with Authentic Compounds. *Environ. Sci. Technol.* **2012**, *46*, 250-258.
17. Paulot, F.; Crounse, J. D.; Kjaergaard, H. G.; Kurten, A.; St Clair, J. M.; Seinfeld, J. H.; Wennberg, P. O., Unexpected Epoxide Formation in the Gas-Phase Photooxidation of Isoprene. *Science* **2009**, *325*, 730-733.

18. Crounse, J. D.; Nielsen, L. B.; Jorgensen, S.; Kjaergaard, H. G.; Wennberg, P. O., Autoxidation of Organic Compounds in the Atmosphere. *J. Phys. Chem. Lett.* **2013**, *4*, 3513-3520.
19. Ehn, M.; Thornton, J. A.; Kleist, E.; Sipila, M.; Junninen, H.; Pullinen, I.; Springer, M.; Rubach, F.; Tillmann, R.; Lee, B., et al., A large source of low-volatility secondary organic aerosol. *Nature* **2014**, *506*, 476-480.
20. Stone, D.; Whalley, L. K.; Heard, D. E., Tropospheric OH and HO₂ radicals: field measurements and model comparisons. *Chem. Soc. Rev.* **2012**, *41*, 6348-6404.
21. Trostl, J.; Chuang, W. K.; Gordon, H.; Heinritzi, M.; Yan, C.; Molteni, U.; Ahlm, L.; Frege, C.; Bianchi, F.; Wagner, R., et al., The role of low-volatility organic compounds in initial particle growth in the atmosphere. *Nature* **2016**, *533*, 527-530.
22. Laskin, A.; Laskin, J.; Nizkorodov, S. A., Chemistry of Atmospheric Brown Carbon. *Chem. Rev.* **2015**, *115*, 4335-4382.
23. Jang, M. S.; Czoschke, N. M.; Lee, S.; Kamens, R. M., Heterogeneous atmospheric aerosol production by acid-catalyzed particle-phase reactions. *Science* **2002**, *298*, 814-817.
24. Herrmann, H.; Schaefer, T.; Tilgner, A.; Styler, S. A.; Weller, C.; Teich, M.; Otto, T., Tropospheric Aqueous-Phase Chemistry: Kinetics, Mechanisms, and Its Coupling to a Changing Gas Phase. *Chem. Rev.* **2015**, *115*, 4259-4334.
25. Rohrer, F.; Lu, K. D.; Hofzumahaus, A.; Bohn, B.; Brauers, T.; Chang, C. C.; Fuchs, H.; Haseler, R.; Holland, F.; Hu, M., et al., Maximum efficiency in the hydroxyl-radical-based self-cleansing of the troposphere. *Nat. Geosci.* **2014**, *7*, 559-563.
26. Ng, N. L.; Brown, S. S.; Archibald, A. T.; Atlas, E.; R.C., C.; Crowley, J. N.; Day, D. A.; Donahue, N. M.; Fry, J. L.; al., F., Nitrate radicals and biogenic volatile organic compounds: oxidation, mechanisms and organic aerosol. *Atmos. Chem. Phys. Discuss.* **2016**, doi:10.5194/acp-2016-734.
27. George, C.; Ammann, M.; D'Anna, B.; Donaldson, D. J.; Nizkorodov, S. A., Heterogeneous Photochemistry in the Atmosphere. *Chem. Rev.* **2015**, *115*, 4218-4258.
28. Taatjes, C. A.; Shallcross, D. E.; Percival, C. J., Research frontiers in the chemistry of Criegee intermediates and tropospheric ozonolysis. *Phys. Chem. Chem. Phys.* **2014**, *16*, 1704-1718.
29. Welz, O.; Savee, J. D.; Osborn, D. L.; Vasu, S. S.; Percival, C. J.; Shallcross, D. E.; Taatjes, C. A., Direct Kinetic Measurements of Criegee Intermediate (CH₂OO) Formed by Reaction of CH₂I with O₂. *Science* **2012**, *335*, 204-207.
30. Poschl, U.; Shiraiwa, M., Multiphase Chemistry at the Atmosphere-Biosphere Interface Influencing Climate and Public Health in the Anthropocene. *Chem. Rev.* **2015**, *115*, 4440-4475.
31. Ervens, B.; Turpin, B. J.; Weber, R. J., Secondary organic aerosol formation in cloud droplets and aqueous particles (aqSOA): a review of laboratory, field and model studies. *Atmos. Chem. Phys.* **2011**, *11*, 11069-11102.
32. Koop, T.; Bookhold, J.; Shiraiwa, M.; Poschl, U., Glass transition and phase state of organic compounds: dependency on molecular properties and implications for secondary organic aerosols in the atmosphere. *Phys. Chem. Chem. Phys.* **2011**, *13*, 19238-19255.
33. Petters, M. D.; Kreidenweis, S. M., A single parameter representation of hygroscopic growth and cloud condensation nucleus activity. *Atmos. Chem. Phys.* **2007**, *7*, 1961-1971.
34. Ruehl, C. R.; Davies, J. F.; Wilson, K. R., An interfacial mechanism for cloud droplet formation on organic aerosols. *Science* **2016**, *351*, 1447-1450.
35. Noziere, B.; Baduel, C.; Jaffrezo, J. L., The dynamic surface tension of atmospheric aerosol surfactants reveals new aspects of cloud activation. *Nat. Commun.* **2014**, *5*, 7-12.
36. Hoose, C.; Mohler, O., Heterogeneous ice nucleation on atmospheric aerosols: a review of results from laboratory experiments. *Atmos. Chem. Phys.* **2012**, *12*, 9817-9854.
37. Kleffmann, J., Daytime sources of nitrous acid (HONO) in the atmospheric boundary layer. *ChemPhysChem* **2007**, *8*, 1137-1144.

- 475 38. Li, X.; Rohrer, F.; Hofzumahaus, A.; Brauers, T.; Haseler, R.; Bohn, B.; Broch, S.; Fuchs, H.; Gomm,
476 S.; Holland, F., et al., Missing Gas-Phase Source of HONO Inferred from Zeppelin Measurements in the
477 Troposphere. *Science* **2014**, *344*, 292-296.
- 478 39. George, C.; Strekowski, R. S.; Kleffmann, J.; Stemmler, K.; Ammann, M., Photoenhanced uptake
479 of gaseous NO₂ on solid-organic compounds: a photochemical source of HONO? *Faraday Discuss.* **2005**,
480 *130*, 195-210.
- 481 40. Ye, C. X.; Zhou, X. L.; Pu, D.; Stutz, J.; Festa, J.; Spolaor, M.; Tsai, C.; Cantrell, C.; Mauldin, R. L.;
482 Campos, T., et al., Rapid cycling of reactive nitrogen in the marine boundary layer. *Nature* **2016**, *532*,
483 489-491.
- 484 41. VandenBoer, T. C.; Young, C. J.; Talukdar, R. K.; Markovic, M. Z.; Brown, S. S.; Roberts, J. M.;
485 Murphy, J. G., Nocturnal loss and daytime source of nitrous acid through reactive uptake and
486 displacement. *Nat. Geosci.* **2015**, *8*, 55-60.
- 487 42. Donaldson, M. A.; Bish, D. L.; Raff, J. D., Soil surface acidity plays a determining role in the
488 atmospheric-terrestrial exchange of nitrous acid. *Proc. Nat. Acad. Sci. U.S.A.* **2014**, *111*, 18472-18477.
- 489 43. Lushchak, V. I., Free radicals, reactive oxygen species, oxidative stress and its classification.
490 *Chem.-Biol. Interact.* **2014**, *224*, 164-175.
- 491 44. Charrier, J. G.; Anastasio, C., On dithiothreitol (DTT) as a measure of oxidative potential for
492 ambient particles: evidence for the importance of soluble transition metals. *Atmos. Chem. Phys.* **2012**,
493 *12*, 9321-9333.
- 494 45. Squadrito, G. L.; Cueto, R.; Dellinger, B.; Pryor, W. A., Quinoid redox cycling as a mechanism for
495 sustained free radical generation by inhaled airborne particulate matter. *Free Radic. Biol. Med.* **2001**, *31*,
496 1132-1138.
- 497 46. Shuster-Meiseles, T.; Shafer, M. M.; Heo, J. E.; Pardo-Levin, M.; Antkiewicz, D. S.; Schauer, J. J.;
498 Rudich, A.; Rudich, Y., ROS-generating/ARE-activating capacity of metals in roadway particulate matter
499 deposited in urban environment. *Environ. Res.* **2016**, *146*, 252-262.
- 500 47. von Schneidmesser, E.; Monks, P. S.; Allan, J. D.; Bruhwiler, L.; Forster, P.; Fowler, D.; Lauer, A.;
501 Morgan, W. T.; Paasonen, P.; Righi, M., et al., Chemistry and the Linkages between Air Quality and
502 Climate Change. *Chem. Rev.* **2015**, *115*, 3856-3897.
- 503 48. Abbatt, J. P. D., Interactions of atmospheric trace gases with ice surfaces: Adsorption and
504 reaction. *Chem. Rev.* **2003**, *103*, 4783-4800.
- 505 49. Bertram, T. H.; Thornton, J. A.; Riedel, T. P.; Middlebrook, A. M.; Bahreini, R.; Bates, T. S.; Quinn,
506 P. K.; Coffman, D. J., Direct observations of N₂O₅ reactivity on ambient aerosol particles. *Geophys. Res.*
507 *Lett.* **2009**, *36*.
- 508 50. Nguyen, T. B.; Crounse, J. D.; Schwantes, R. H.; Teng, A. P.; Bates, K. H.; Zhang, X.; St Clair, J. M.;
509 Brune, W. H.; Tyndall, G. S.; Keutsch, F. N., et al., Overview of the Focused Isoprene eXperiment at the
510 California Institute of Technology (FIXCIT): mechanistic chamber studies on the oxidation of biogenic
511 compounds. *Atmos. Chem. Phys.* **2014**, *14*, 13531-13549.
- 512 51. Mentel, T. F.; Wildt, J.; Kiendler-Scharr, A.; Kleist, E.; Tillmann, R.; Dal Maso, M.; Fisseha, R.;
513 Hohaus, T.; Spahn, H.; Uerlings, R., et al., Photochemical production of aerosols from real plant
514 emissions. *Atmos. Chem. Phys.* **2009**, *9*, 4387-4406.
- 515 52. West, J. J.; Cohen, A.; Dentener, F.; Brunekreef, B.; Zhu, T.; Armstrong, B.; Bell, M. L.; Brauer, M.;
516 Carmichael, G.; Costa, D. L., et al., "What We Breathe Impacts Our Health: Improving Understanding of
517 the Link between Air Pollution and Health". *Environ. Sci. Technol.* **2016**, *50*, 4895-4904.
- 518 53. Morrison, G., Interfacial chemistry in indoor environments. *Environ. Sci. Technol.* **2008**, *42*, 3494-
519 3499.
- 520 54. Vereecken, L.; Glowacki, D. R.; Pilling, M. J., Theoretical Chemical Kinetics in Tropospheric
521 Chemistry: Methodologies and Applications. *Chem. Rev.* **2015**, *115*, 4063-4114.

55. Cox, R. A., Evaluation of laboratory kinetics and photochemical data for atmospheric chemistry applications. *Chem. Soc. Rev.* **2012**, *41*, 6231-6246.
56. Ammann, M.; Cox, R. A.; Crowley, J. N.; Jenkin, M. E.; Mellouki, A.; Rossi, M. J.; Troe, J.; Wallington, T. J., Evaluated kinetic and photochemical data for atmospheric chemistry: Volume VI - heterogeneous reactions with liquid substrates. *Atmos. Chem. Phys.* **2013**, *13*, 8045-8228.
57. Prather, K. A.; Hatch, C. D.; Grassian, V. H., Analysis of Atmospheric Aerosols. *Annu. Rev. Anal. Chem.* **2008**, *1*, 485-514.
58. Lopez-Hilfiker, F. D.; Mohr, C.; Ehn, M.; Rubach, F.; Kleist, E.; Wildt, J.; Mentel, T. F.; Lutz, A.; Hallquist, M.; Worsnop, D., et al., A novel method for online analysis of gas and particle composition: description and evaluation of a Filter Inlet for Gases and AEROsols (FIGAERO). *Atmos. Meas. Tech.* **2014**, *7*, 983-1001.

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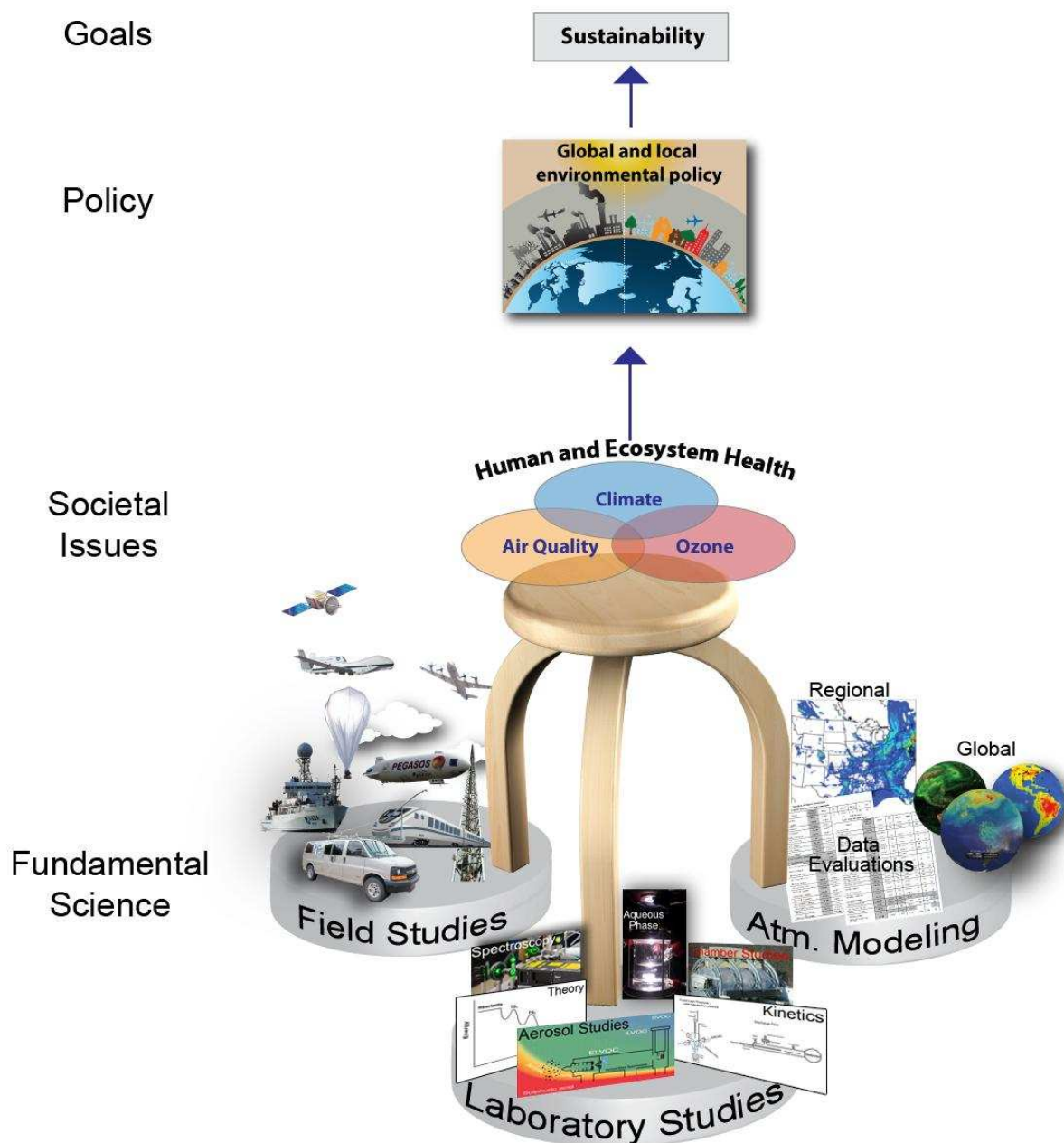
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540 **Figure 1:** A chemical view of the atmosphere highlighting biogenic and anthropogenic
 541 emissions sources and key atmospheric species.

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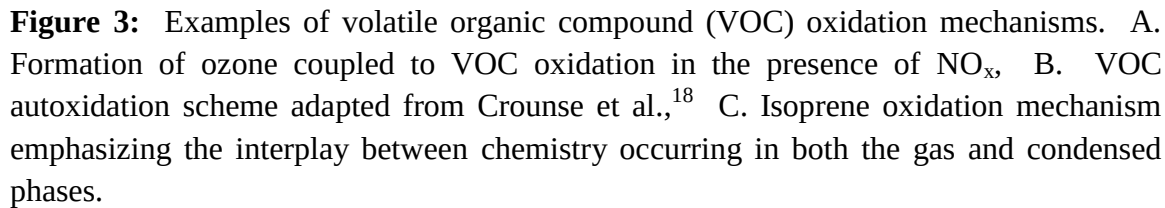
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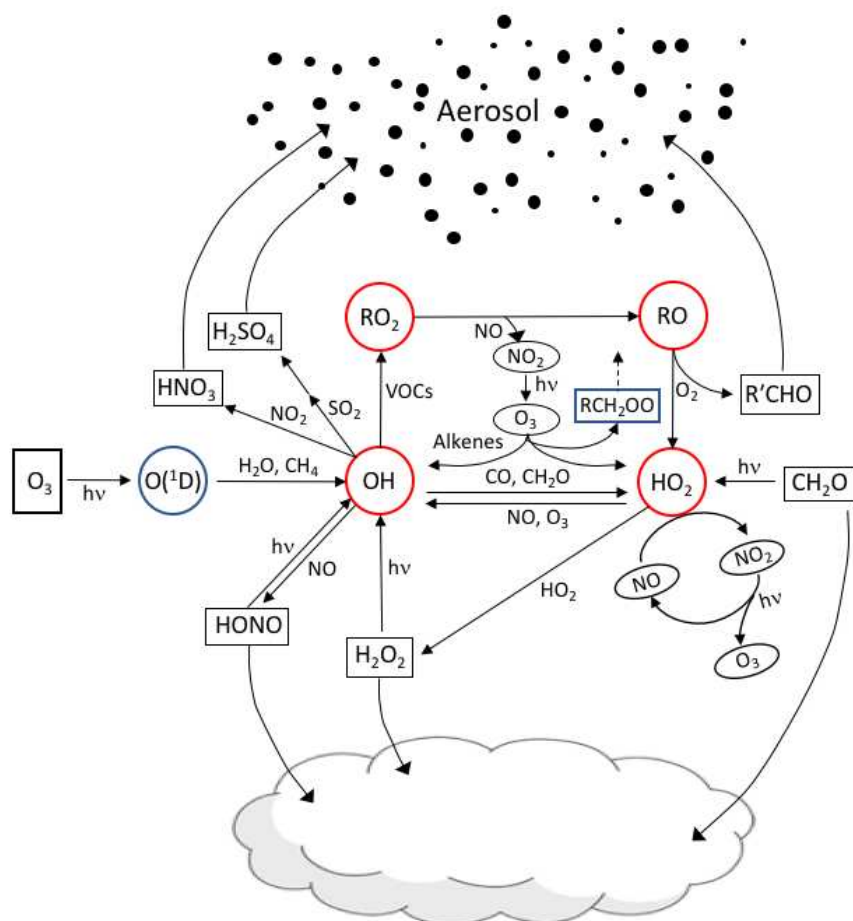
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546 **Figure 2:** The three-legged stool connecting atmospheric chemistry to sustainable policy.

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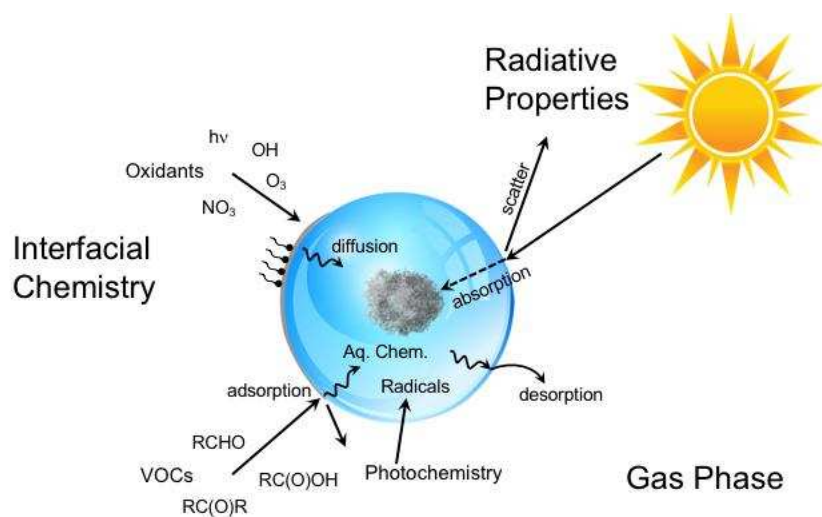
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559 **Figure 4:** Examples of radical formation, cycling, formation of stable products, and
 560 contributions to aerosol and cloud chemistry. Note that radical chemistry also occurs in cloud
 561 droplets and aerosol particles but is not shown.

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567 **Figure 5:** Representation of multiphase processes and radiative properties for a particle or
568 droplet containing a solid core. Note that other particle morphologies are possible.

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