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Finite-element simulations of the pH-ElecFET microsensors

N. Aoun, F. Echouchene, A. K. Diallo, J. Launay, P. Temple-Boyer and H. Belmabrouk

Abstract—This paper presents a COMSOL Multiphysics 2D axisymmetric model of an pH-ElecFET (pH-sensitive electrochemical field effect transistor) microsensor. This device combines an integrated microelectrode with a pH-sensitive chemical field effect transistor (pH-ChemFET). Thus, by triggering electrolysis phenomena owing to the integrated microelectrode, associated local pH variations in microvolumes are monitored thanks to the pH-ChemFET microdevice. Taking into account (electro) chemical reactions and diffusion phenomena in liquid phase, the proposed model points out the role of the ElecFET geometrical design (microelectrode width w , gate sensitive radius r_e and distance between the pH-ChemFET gate and the microelectrode d), as well as polarization parameters, (polarization voltage V_p and time t_p), on the microsensor response. It is first applied to water electrolysis in order to validate pH impulsional variations in microvolume. Then, oxidation of hydrogen peroxide in phosphate buffer (PBS, $\text{pH}_0=7.2$) solutions is studied, evidencing the H_2O_2 potentiometric detection in the [10–100mM] concentration range. This developed model paves new ways for sensor applications, opening several new opportunities for pH-ElecFET devices for H_2O_2 -related enzymatic detection of biomolecules.

Index Terms—Modelling, ElecFET, microelectrode, pH-ChemFET, water electrolysis, hydrogen peroxide detection.

I. INTRODUCTION

In the last decade, the electrochemical microsensors have received an increasing interest in a wide range of applications such as clinical diagnostics, food analysis, environmental monitoring due to their low cost, simple operation, small size, and rapidity, sensitivity and real-time [1-3]. The electrochemical sensors can be divided into three groups depending on the measured electrical signal [4-6]: amperometric, potentiometric, and conductometric. Even so, the combination of amperometric and potentiometric techniques is a very promising method in terms of detection [7-10]. Diallo et al. have developed an electrochemical field effect transistor (ElecFET) microsensor based on this technique [8, 9]. This device is achieved through the integration of a planar noble metal electrode around the dielectric gate area of a pH-sensitive ChemFET microdevice. By triggering pH-related electrochemical reactions thanks to the microelectrode polarization and by monitoring the so-

obtained pH variations thanks to the pH-ChemFET, electrolysis phenomena and pH measurement are closely embedded at the microscale, enabling new electrochemical detection potentialities. Among the ElecFET microdevice applications, the most frequent is the manufacturing of a pH-related enzyme sensor. In this case, the ChemFET detects the pH change resulting from the enzymatic reaction in the membrane that covers the sensor [11]. The output voltage of the ChemFET controls the current flowing through the sensor-actuator system [12]. ElecFET has been successfully used to determine acid or base concentration [13], to form the heart of a carbon dioxide sensor [14] and detect different biomolecules [8, 11].

Various mathematical models of microsensors have been developed and successfully used to study and optimize analytical characteristics of microsensors [15-18].

Meena and Rajendran [15] have derived analytical expressions of concentration and current in order to describe and evaluate the performances of amperometric and potentiometric biosensors using homotopy perturbation method. A numerical study of potentiometric and amperometric electrochemical gas sensors based in a solid-state ion conducting electrolyte has presented by López-Gándaraa et al. [16] in order to optimize the diffuse layers covering one of their catalytic electrodes. The model describes the current-voltage characteristics in the system layer/electrode/electrolyte/electrode. A theoretical study and numerical simulation of potentiometric and amperometric enzyme electrodes and of enzyme reactors have been developed by Morf et al. [17, 18]. The response characteristics of potentiometric and amperometric sensor systems, as well as the product release from enzyme reactors, are analyzed, and the influence of the relevant parameters on the steady-state response is demonstrated and discussed [17].

In this paper, we apply the concept of the combination of amperometric and potentiometric techniques to simulate the ElecFET microdevice detection principles, in order to monitor water (H_2O) electrolysis phenomena in a first stage and hydrogen peroxide (H_2O_2) electrochemical detection in the second stage. We then focus on the study of the influences of the main parameters, i.e. (i) polarization voltage V_p and time t_p on the integrated microelectrode, (ii) characteristic of the microelectrode width w , (iii) distance between the gate sensitive radius and the integrated microelectrode d and (iv) gate sensitive radius r_e .

II. PRESENTATION OF THE SIMULATION MODEL

The modelling approach used takes into account the different chemical, electrochemical and physical phenomena occurring in the frame of the ElecFET detection principles:

- Oxido-reduction on the integrated microelectrode;

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- Diffusion phenomena;
- (H₃O⁺/H₂O) and (H₂O/OH⁻) and acid-base analysis of aqueous solutions at pH₀=7.2.

The ElecFET detection properties were thus modeled by studying the diffusion phenomena of the main chemical species in the electrolyte, focusing on the H₃O⁺/OH⁻ water-based ions, and finally by analyzing the pH detection properties of the silicon nitride Si₃N₄ ChemFET gate.

A. Geometry of ElecFET: comparison between real and simulated model

Since their fabrication process used silicon-based microtechnologies and therefore photolithography, the real ElecFET microdevices are based on a rectangular concentric geometry [9]. In order to decrease calculation times, the simulated model was simplified by using *r-z* cylindrical geometry, leading to a 2D axisymmetric geometry (Fig.1).

The geometrical parameters are: the gate sensitive radius *r_e*, the distance between the sensitive gate and the electrode *d*, and *w* the microelectrode ring width. The depth of the sensitive gate *z_e* is 0.5μm and the height of the modeled geometry is *h*=300 μm. The standard values of the parameters (*r_e*, *d*, *w*) are (10μm, 50μm, 100 μm). At the end of the paper these values will be modified in order to enhance the performance of pH-ElecFET. The dimensions of electrolyte domain are indicated in Fig. 1.

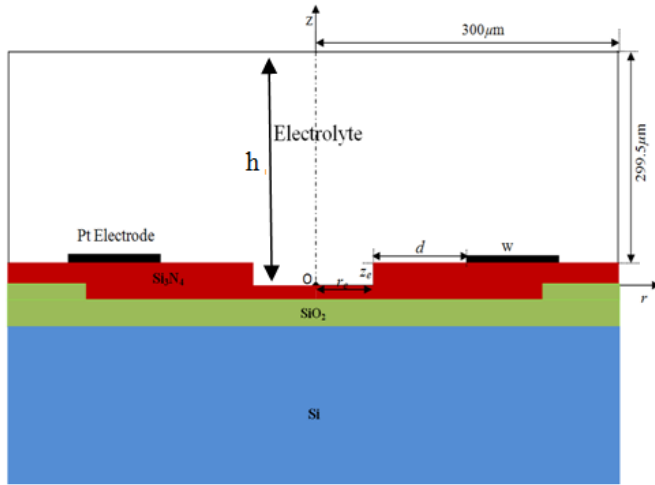
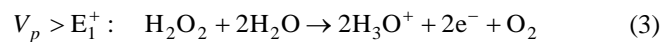
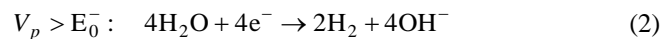
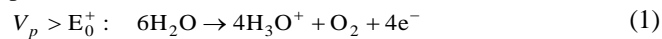


Fig. 1. Cross-section on the ElecFET device in *r-z* cylindrical coordinates.

B. Modelling of the electrochemical reactions (H₂O & H₂O₂)

The ElecFET concept was initially used for the monitoring of water (H₂O) hydrolysis and secondly of the hydrogen peroxide (H₂O₂) oxidation in water based solutions:



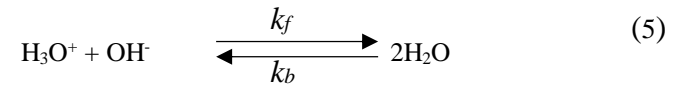
where *E₀⁺*, *E₀⁻*, and *E₁⁺* are equilibrium potentials of the O₂/H₂O, the H₂/H₂O and O₂/H₂O₂ redox couple, respectively. The values of the constants are: *E₀⁺* = 1.2V, *E₀⁻* = -0.8V [9] and *E₁⁺* = 0.7V [8].

C. Modelling of the diffusion phenomena in water

Diffusion phenomena of the most influential chemical species, i.e. H₃O⁺ and OH⁻ ions in case of H₂O detection and H₃O⁺, OH⁻ and H₂O₂ species in case of hydrogen peroxide detection, into water, have been modeled using the Fick law:

$$\frac{\partial c_i}{\partial t} = D_i \left(\frac{\partial^2 c_i}{\partial r^2} + \frac{1}{r} \frac{\partial c_i}{\partial r} + \frac{\partial^2 c_i}{\partial z^2} \right) + R_i \quad (4)$$

where *t* denotes the time, *c_i* is the concentration of the studied chemical species, *D_i* is its diffusion coefficient and *R_i* is an additional term accounting for the production/consumption rate of the ion species through a possible chemical reaction. The reaction term can be deduced using the water protolysis reaction:



where *k_f* and *k_b* are the forward and backward reaction kinetic rates, respectively.

One can thus express the reaction for the production of proton and hydroxide ions as [19]:

$$\begin{aligned} R_{H_3O^+} &= R_{OH^-} = k_b[H_2O] - k_f[H_3O^+][OH^-] \\ &= k_f \left(K_e - [H_3O^+][OH^-] \right) \end{aligned} \quad (6)$$

where *K_e* = *k_b*[H₂O]/*k_f* is the equilibrium constant of water.

The forward rate constant *k_f* of the water protolysis reaction is estimated by values taken from literature [20] while the backward rate constants *k_b* is obtained from where *K_e* denotes the water dissociation constant.

Buffer solution is of great importance in biosensor applications. Experimentally, the buffer concentration could be optimized. So, a lot of experiments have been reported on the buffer properties which enhance the response and the stability of the biosensor [21, 22]. In our kinetic model the chemical effect of the buffer solution was neglected and was not taken into account. In the frame of a kinetic approach, the influences of acid/bases other than the water-based ones are limited by diffusion phenomena (Eq. 4). In another way, we have been interested only on the protonation reaction because, as a solvent, water is in excess in solution. As a result, we can assume that the buffer solution does not have an effect on the biosensor response which can be expressed with the chemical effect modeled with physical equations.

D. Initial and boundary conditions

Boundary conditions are crucial for a correct description of phenomena with partial differential equations. For our model, several important conditions have to be set. The initial condition related to mass transport of species is given by:

$$\begin{cases} c_{H_3O^+}(r, z; t=0) = 10^{-pH_0} \\ c_{OH^-}(r, z; t=0) = 10^{-pK_e + pH_0} \\ c_{H_2O_2}(r, z; t=0) = 50mM \end{cases} \quad \begin{cases} 0 \leq r \leq 300\mu m \\ 0 \leq z \leq h \end{cases} \quad (7)$$

On the axis of symmetry, the boundary condition reads:

$$\frac{\partial c_i}{\partial r}(r=0, z, t) = 0 \quad (8)$$

At the upper interface of the electrolyte ($z=h$) and the lower interface ($z=0$ or z_e), the interface is assumed to be impermeable:

$$\frac{\partial c_i}{\partial z}(r, z=h \text{ or } 0 \text{ or } z_e, t) = 0 \quad (9)$$

In order to do not disturb the diffusion phenomena operating from the surface of the microelectrode, the outer surface of the device is assumed to be at a constant concentration:

$$c_i(r = 300 \mu\text{m}, z, t) = c_{i,0} \quad (10)$$

An oxido-reduction reaction takes place at the electrode for $t < t_p$. Thus, the boundary condition reads:

$$D_i \frac{\partial c_i}{\partial z}(r, z=z_e, t \leq t_p) = \frac{v_{i,j}}{N_j F} i_j ; \quad r_e + d \leq r \leq r_e + d + w \quad (11)$$

where N_j is the number of electron transferred for reaction "j", $v_{i,j}$ is the stoichiometric coefficient, F is Faraday's constant and t_p is the polarization time. i_j denotes the current density given from the kinetic equations for the electrochemical reactions at the electrode surface based on the Butler–Volmer expression[23]:

$$i_j = i_{0j,ref} \left\{ \prod_i \left(\frac{c_i}{c_{i,ref}} \right)^{p_{i,j}} \exp\left(\frac{\alpha_{a,j} F}{RT} \eta_j\right) - \prod_i \left(\frac{c_i}{c_{i,ref}} \right)^{q_{i,j}} \exp\left(-\frac{\alpha_{c,j} F}{RT} \eta_j\right) \right\} \quad (12)$$

where $i_{0j,ref}$ is the exchange current density due to reaction "j" at the reference concentrations in A/cm^2 , c_i is the concentration of species "i" adjacent to the surface of electrode in mol/cm^3 , $c_{i,ref}$ is the reference concentration of species "i" in mol/cm^3 , α_{aj} is the anodic transfer coefficient for reaction "j", α_{cj} is the cathodic transfer coefficient for reaction "j", $p_{i,j} = v_{i,j}$ is the anodic reaction order of species "i" in reaction "j", $q_{i,j} = -v_{i,j}$ is the cathodic reaction order of species "i" in reaction "j", R is the gas constant, T is the temperature and η_j is the overpotential of reaction "j" in volts (V), and it is measured with respect to a reference electrode of a given kind in a solution at the reference concentrations.

The overpotential for electrochemical reaction "j", (η) in Eq. 12 is given by:

$$\eta_j = V_p - E_{0,j} \quad (13)$$

where V_p is the applied potential surface in V, $E_{0,j}$ is the equilibrium potential for reaction "j".

Total current is obtained by numerical integration of the local current density over the entire ElecFET area. It is given by:

$$I(t) = \int_A i_j dA \quad (14)$$

where A is the surface area of the electrode.

E. Numerical method

The governing equations with initial and boundary conditions are solved numerically using the finite element software COMSOL Multiphysics 4.3b [24]. An electroanalysis module has been chosen for solving the mass transport of diluted species in electrolytes using the diffusion equation.

The overall computational domain is discretized using an unstructured triangular mesh Fig.2. It is noticed that the region nearby the reaction surface is refined with a better mesh quality.

For the time discretization we have employed First-order backward differentiation formula (BDF), with the time steps controlled by the numerical solver during the computations.

Total current, given by Eq. (14), is performed using a fourth order integration method.

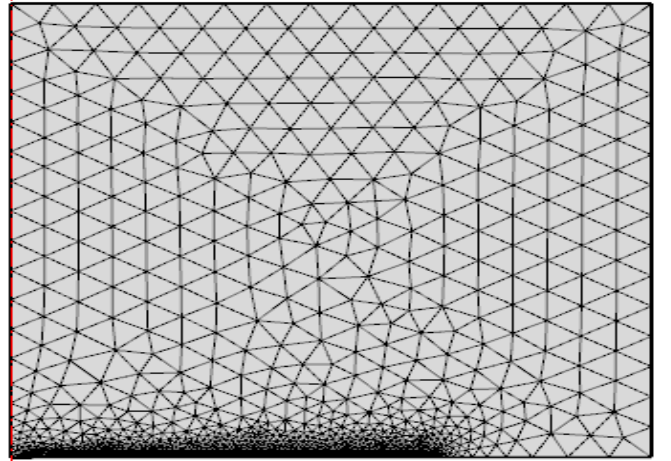


Fig. 2. Two-dimensional unstructured mesh with triangular elements.

III. RESULTS AND DISCUSSION

To describe the *pH-ChemFET response*, we can use the threshold voltage variation. The pH-ElecFET threshold voltage variation is related to the pH at the silicon nitride Si_3N_4 surface according to the following equation [20]:

$$V_T(t) = V_{T0} + s_0 \times (\text{pH}(0,t) - \text{pH}_{\text{pzc}}) \quad (15)$$

where $\text{pH}(0,t)$ is the pH at the Si_3N_4 sensitive surface when the diffusion phenomena "steady state" is reached and V_{T0} is a constant parameter depending on the $\text{SiO}_2/\text{Si}_3\text{N}_4$ pH-ChemFET technology [25], pH_{pzc} is the point zero of charge which is estimated around 4 for Si_3N_4 [23, 26]). In the following, since the V_{T0} value is only related to the pH-ChemFET technological fabrication, it is of no influence concerning the ElecFET detection properties and it will not be taken into account, i.e. it will be chosen equal to zero [27, 28]. s_0 is an ideal sensitivity of ISFET based pH sensor when pH is measured by the means of Nernst potential given by the following equation [29]:

$$s_0 = \frac{kT}{q} \ln(10) = 59.2 \text{ mV} / \text{pH} \quad (16)$$

where $k=1.38 \times 10^{-23} \text{J/K}$ is the Boltzmann constant, $T=300\text{K}$ is temperature and $q=1.6 \times 10^{-19} \text{C}$ the elementary charge.

A. Mesh sensitivity

To ensure that the convergence has been obtained and the numerical results are independent of the mesh size, Fig. 3 shows the pH local at $z=150 \mu\text{m}$ of the electrolyte for several mesh grids. There was no significant difference between the curves obtained using these grids (6818, 13643, 27719, 53534 elements) and we may conclude that the numerical convergence has been reached with all the grids. In the results section, all simulations were done with a total elements number 13645 [30, 31].

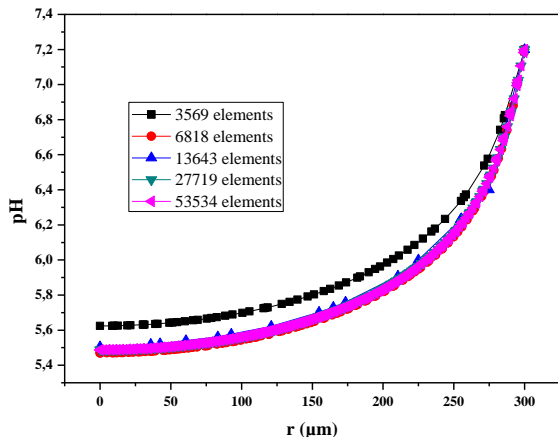


Fig. 3. pH at $z = 150 \mu\text{m}$ of the electrolyte for several mesh grids.

B. Study of the most influential parameters

The kinetic parameters and the constants used in this simulation are listed in Table I and Table II [9, 32, 33]. ElecFET detection/transduction principles were first studied for water electrolysis monitoring H_2O in aqueous solutions ($\text{pH}_0=7.2$). Then, the ElecFET detection properties were tested for the hydrogen peroxide H_2O_2 also in aqueous solutions ($\text{pH}_0=7.2$) solutions containing different stabilized concentrations $[\text{H}_2\text{O}_2]$ ranging from 10 to 100mM because of their properties in terms of activity, stability and cost. According to theoretical equations, the most influential parameters are the polarization voltage V_p , the time of polarization t_p , the concentration $[\text{H}_2\text{O}_2]$, the distance between the gate sensitive radius and the microelectrode (parameter d), the surface of the microelectrode (associated with the parameter w) and gate sensitive radius (parameter r_e).

All series of simulation have been devoted to modelling the sensor response, i.e. the dependence of the threshold voltage versus time. The curves of this threshold voltage, measured at $\text{pH}_0=7.2$, are shown in Figs.4-9.

TABLE I
PARAMETERS USED IN THE SIMULATIONS

	Reaction 1	Reaction 2	Reaction 3
$i_{0j,ref} (\text{A/m}^2)$	1.115×10^{-6}	10^{-6}	9.64×10^{-6}
$E_{0j} (\text{V})$	1.2	-0.8	0.7
$\alpha_{a,j}$	0.5	0.5	0.5
$\alpha_{c,j}$	0.5	0.5	0.5
$V_{\text{H}_3\text{O}^+}$	+1		+1
V_{OH^-}		+1	
$V_{\text{H}_2\text{O}_2}$			-1
N_j	4	4	2

TABLE II
PHYSICO-CHEMICAL INPUT VALUES USED IN THE SIMULATION MODEL

Parameter	Value	Unit
F	96352	C/mol
$D_{\text{H}_3\text{O}^+}$	9.3×10^{-9}	m^2/s
D_{OH^-}	5.3×10^{-9}	m^2/s
$D_{\text{H}_2\text{O}_2}$	3.1×10^{-10}	m^2/s
$c_{\text{H}_3\text{O}^+,ref}$	10^{-pH_0}	mol/L
$c_{\text{OH}^-,ref}$	10^{pH_0}	mol/L
$c_{\text{H}_2\text{O}_2,ref}$	50×10^{-3}	mol/L
pH_0	7.2	
k_f	1.5×10^{11}	L/mol/s
K_e	1×10^{-14}	mol^2/L^2

Case1: water electrolysis monitoring:

Before testing the sensor performance for the H_2O_2 detection, one should ensure that the sensor is able to provide a response in the case of water electrolysis.

Fig. 4 presents the temporal evolution of the threshold voltage V_T in the case of H_2O redox phenomena for different polarization voltages while maintaining a constant the polarization time ($t_p = 5\text{s}$). Since V_p is positive (respectively negative), oxidation (respectively reduction) of H_2O takes place near the platinum microelectrode.

Therefore, a production of hydronium H_3O^+ (respectively hydroxide OH^-) ions, so for a local pH changes sharply and finally the threshold voltage V_T increases (respectively decreases). In this case, phenomena occur at higher than 1.2V (positive polarization) and/or lower than -0.8V (negative polarization) polarization voltages V_p according to the associated equilibrium potential values E_0^+ and E_1^+ .

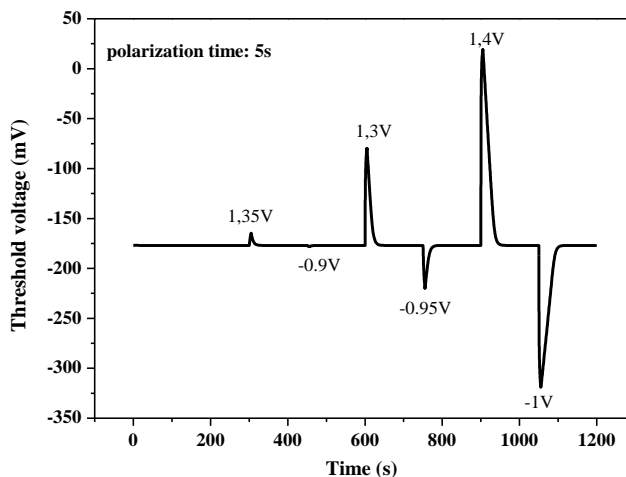


Fig. 4. Temporal variations of the pH-ChemFET threshold voltage for different polarization voltages.

The amplitude of the peak is so much greater than the value of the applied potential increases. When the polarization is interrupted, the ElecFET microsensor response follows a return to equilibrium of the system because of the diffusion laws.

Subsequently, the influence of polarization time t_p was studied. Fig. 5 illustrates the temporal variations threshold voltage for different polarization times ($t_p = 5, 10, 10$ and 3 s) for two values of the polarization voltage on the integrated microelectrode ($V_p = 1.4$ or $-1V$). In agreement with the Equation (11), the polarization time increase is responsible for a local pH decrease and therefore a pH-ChemFET threshold voltage increases (Eq. 15). Nevertheless, threshold voltage variations are lower and tend to reach saturation. Fig. 4 and Fig. 5 show clearly that the microsensor has a response in the case of water electrolysis. Therefore, we can continue our analysis and assess the performance of the sensor for H_2O_2 detection.

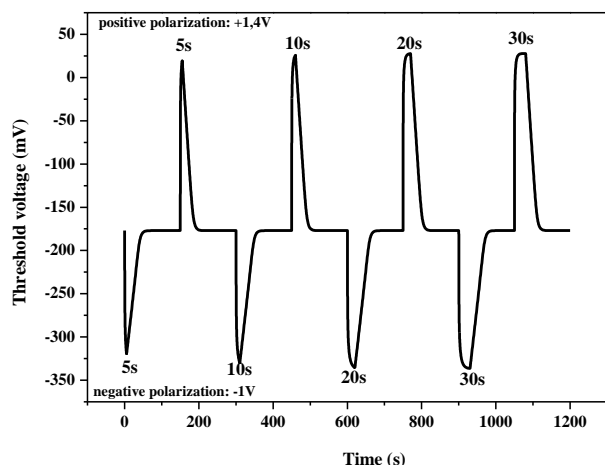


Fig. 5. Temporal variations of the pH-ChemFET threshold voltage for different polarization times.

Case2: hydrogen peroxide detection:

The ElecFET concept was studied by taking into account the H_2O_2 oxidation on the integrated microelectrode. According to the oxidation reaction (Eq.3), in presence of hydrogen peroxide H_2O_2 , a positive polarization V_p on the platinum microelectrode produces hydronium H_3O^+ ions. Therefore, the local pH decreases around the sensitive area of the microsensor and finally the threshold voltage V_T increases. For this purpose, the polarization voltage V_p was chosen less than E_0^+ to avoid any interference with water hydrolysis. Since the E_1^+ is roughly equal to $0.7V$, a similar polarization voltage V_p was applied on the platinum microelectrode at the different polarization times ($t_p = 5, 10, 15, 20, 30, 40$ and $60s$) while keeping constant $[H_2O_2]$ concentration at $50mM$ in a solution of $pH_0=7.2$, with a constant polarization voltage (Fig. 6). V_T tends toward increasingly positive values (i.e local pH decreases). As previously for the water electrolysis, the amplitude of the peaks increases with t_p . The maximum threshold voltage increases steeply over a short period of time and thereafter reaches a plateau value. It is clear that the

influence of polarization time appears only in transient state. The equilibrium state is reached for $t_p=10s$.

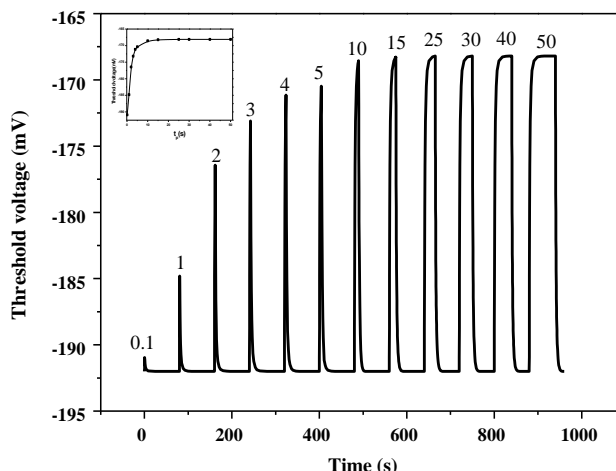


Fig. 6. Temporal variations of the pH-ChemFET threshold voltage for the different polarization times at $[H_2O_2] = 50$ mM, $V_p = 0.7V$, $d=50\mu m$, $w=100\mu m$ and $r_e=10\mu m$ and (inset) ElecFET response versus t_p .

In the same way, Fig. 7 represents the H_2O_2 -ElecFET response as a function of time for different polarization voltages V_p . This figure has been obtained with $[H_2O_2] = 50mM$ in a solution of $pH_0=7.2$ and $t_p = 30s$.

We found that the asymptotic value of threshold voltage is essentially linearly related to the polarization voltage V_p . As soon as the polarization is interrupted, the ElecFET microsensor response tends to turn to back equilibrium of the system because of diffusion phenomenon. In the case of hydrogen peroxide detection, the Butler-Volmer theory emphasizes on the influence of the $[H_2O_2]$ concentration of H_3O^+ ion production kinetics (Eq. 3).

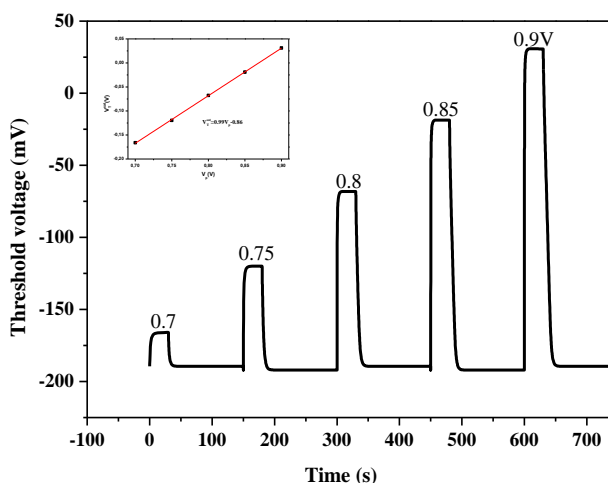


Fig. 7. Temporal variations of the pH-ChemFET threshold voltage for the different polarization voltages at $[H_2O_2] = 50mM$, $t_p=30s$, $d=50\mu m$, $w=100\mu m$ and $r_e=10\mu m$ and (inset) ElecFET response V_p .

Fig. 8 shows temporal variations of the pH-ChemFET threshold voltage for different H_2O_2 concentrations. Simulation results prove that the amplitude of the threshold voltage increases with the $[H_2O_2]$ concentration in solution, and saturates for the highest values. This saturation

phenomenon is related to the previously discussed ElecFET detection principles [33].

Fig. 8 inset presents the evolution of the asymptotic threshold voltage V_T^{Sat} versus the concentration $[H_2O_2]$ and at $V_p=0.7V$ and $t_p=30s$. As illustrated in Fig. 8 inset, the simulation shows a linear response through a large range of $[H_2O_2]$ concentration, as well as a very insensitive behavior. The influence of the microelectrode positions, the surface of the microelectrode and the gate sensitive radius on the terms were also simulated.

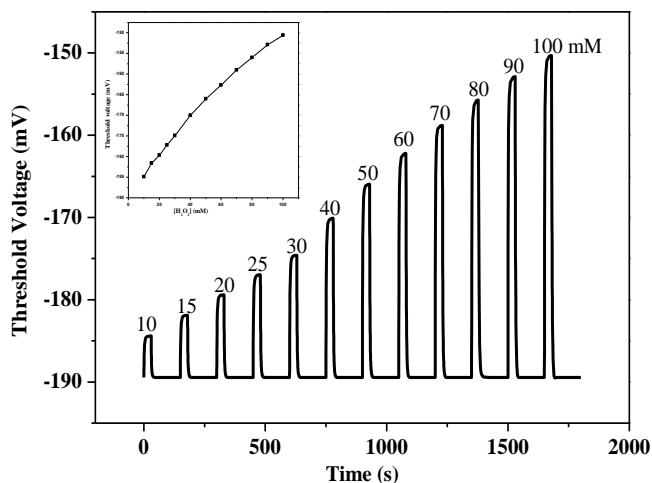


Fig. 8. Temporal variations of the pH-ChemFET threshold voltage for the different H_2O_2 concentrations at $0.7V$, $t_p=30s$, $d=50\mu m$, $w=100\mu m$ and $r_e=10\mu m$ and (inset) H_2O_2 response of the ElecFET microdevice ($V_p=0.7V$)

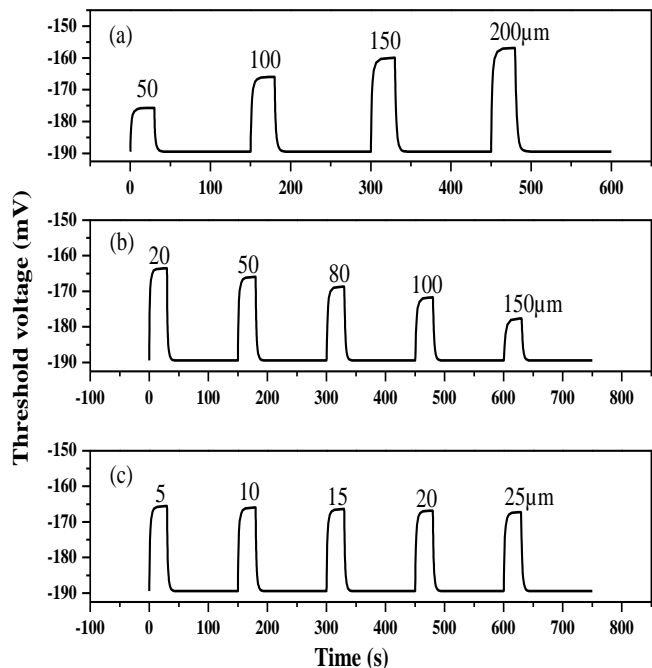


Fig. 9. Temporal variations of the pH-ChemFET threshold voltage for various: (a) surface of the microelectrode at $d=50\mu m$ and $r_e=10\mu m$ (b) positions of the microelectrode at $w=100\mu m$ and $r_e=10\mu m$, and (c) gate sensitive radius at $d=50\mu m$ and $w=100\mu m$ and to all at $V_p=0.7V$, $t_p=30 s$, $[H_2O_2]=50 mM$.

The results of these simulations are shown in Fig. 8. It is worth noting that in Fig. 9, each of the three curves includes the results of simulations for $V_p=0.7 V$, $t_p=30s$ and $[H_2O_2]=50mM$.

In Fig. 9a, the change of the threshold voltage is imposed by the variation of the microelectrode width (w), from 50 to $200 \mu m$ with a step of $50\mu m$. In this case, the maximal variation V_T^{Sat} was increased during the simulation of this stage. Far from the microelectrode ($d=20, 50, 80, 100$ and $150 \mu m$), once hydronium ions are created, they take longtime to diffuse into the solution causing a reduction in their number at very large distances leading to low response of the sensor as illustrated in Fig. 9b.

Finally, Fig. 9c shows the threshold voltage variation with time at various gate sensitive radius ($r_e=5, 10, 15, 20$ and $25\mu m$). When the gate sensitive radius increases, the V_T^{Sat} value also decreases slightly. This proves that the ElecFET detection principles depend mainly on the spatial integration of both microelectrode and pH-ChemFET microdevices.

C. Comparison with experimental results

This section presents a comparative study of experimental results and those obtained by simulation (Fig.8) in the case of hydrogen peroxide detection. It is worth noting that our ChemFET-meter measurement interface was used to measure the pH-ChemFET gate-source voltage V_{GS} , also called output voltage V_{out} , enabling the local pH monitoring in the microelectrode surroundings depending on the polarization conditions. V_{out} shift provides an identical threshold voltage shift and therefore thinking in term of V_{out} or threshold voltage is the same.

Fig.10 shows the evolution of the output voltage versus time at various $[H_2O_2]$ concentrations, while inset represents the maximum of V_{out} versus $[H_2O_2]$ concentrations. We can see that the curve of sensor response for experimental studies shows similar shape as in theoretical studies. As the $[H_2O_2]$ concentrations increases, the sensor response increases too, with V_{out} shifted towards positive values.

Our numerical results show that the amplitude of the impulsional voltage variations increases with the $[H_2O_2]$ concentration in solution, and saturates for the highest values. This saturation phenomenon is related to the ElecFET detection principles as previously mentioned: in the presence of hydrogen peroxide, a positive bias on the platinum microelectrode produces hydronium H_3O^+ ions. Therefore the local pH decreases and finally the pH-ChemFET-meter output voltage V_{out} increases. When the $[H_2O_2]$ concentration increases, the H_3O^+ ions production also increases. Nevertheless, the H_3O^+ ion increase is limited by diffusion and electromigration phenomena, leading to a new chemical equilibrium at the microscale. Hydrogen peroxide H_2O_2 potentiometric detection was finally demonstrated in the $[10-100mM]$ range.

The response from our simulation is not in the same range of the value obtained from the experience. This is due to the effect of the buffer properties. Indeed, Chong et al. have shown that when concentration of the PBS solution decreases, the ionic strength of the PBS solution also decreases causing

the drain current I_d to decrease. The dependence of the I_d to the pH value of the testing pH buffer solutions is related to the variation of the threshold voltage [34]. In the future work, we intend to take into account the effect of the buffer.

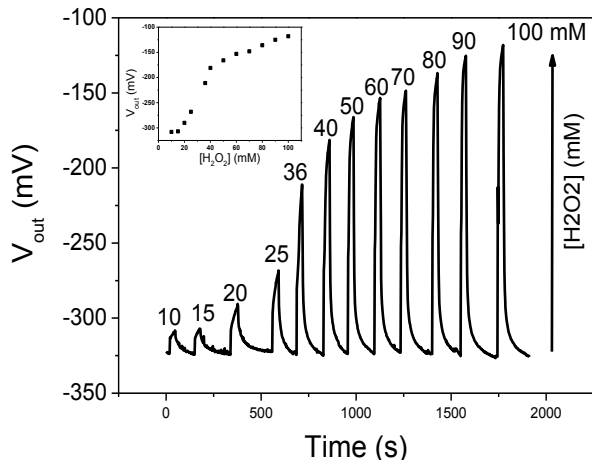


Fig. 10. Detection of impulsional pH variations related to the H_2O_2 electrolysis in phosphate buffer solutions: influence of the H_2O_2 concentration and (inset) ElecFET response for the H_2O_2 detection.

IV. CONCLUSION

In this study, we have investigated the modelling of the ElecFET microdevice, taking into account the mass transport for different species, electrochemical reactions on the microelectrode and acid/basic reactions for couples (H_3O^+/H_2O) and (H_2O/OH^-). The pH-ElecFET techniques were used for monitoring the water-based electrolysis by pH-ChemFET and extended to the hydrogen peroxide in order to obtain the threshold voltage (i.e local pH) variation. The influence of the ElecFET geometrical design, i.e. characteristic width of the microelectrode w , gate sensitive radius r_e and distance between the gate sensitive radius and the microelectrode d , as well as polarization parameters, i.e. polarization voltage V_p and time t_p , on its response behavior was studied. So, the obtaining of pH impulsional variations in microvolumes was clarified and the potentiometric detection of hydrogen peroxide H_2O_2 was evidenced in the [10–100mM] concentration range. In the near future, we intend to model the ElecFET device for the enzymatic detection of lactate and glucose in cylindrical coordinates.

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