



HAL
open science

High Brightness and Easy Color Modulation in Lanthanide-Based Coordination Polymers with 5-Methoxyisophthalate as Ligand: Toward Emission Colors Additive Strategy

Insa Badiane, Stéphane Freslon, Yan Suffren, Carole Daiguebonne, Guillaume Calvez, Kevin Bernot, Magatte Camara, Olivier Guillou

► **To cite this version:**

Insa Badiane, Stéphane Freslon, Yan Suffren, Carole Daiguebonne, Guillaume Calvez, et al.. High Brightness and Easy Color Modulation in Lanthanide-Based Coordination Polymers with 5-Methoxyisophthalate as Ligand: Toward Emission Colors Additive Strategy. *Crystal Growth & Design*, 2017, 17 (3), pp.1224-1234. 10.1021/acs.cgd.6b01607 . hal-01485636

HAL Id: hal-01485636

<https://hal.science/hal-01485636v1>

Submitted on 4 Jul 2017

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

1
2
3 **High brightness and easy color modulation in lanthanide-**
4 **based coordination polymers with 5-methoxy-isophthalate**
5
6 **as ligand: Toward emission colors additive strategy**
7
8
9
10
11
12
13
14
15
16
17
18

19 Insa Badiane^{a,b}, Stéphane Freslon^a, Yan Suffren^a, Carole Daiguebonne^{a*},
20
21 Guillaume Calvez^a, Kevin Bernot^a, Magatte Camara^b and Olivier Guillou^{a*}.
22
23
24
25

26
27 ^a INSA Rennes, UMR 6226 "Institut des Sciences Chimiques de Rennes", 20 Avenue des
28
29 buttes de Coësmes, F35708 Rennes
30

31 ^b Université Assane Seck de Ziguinchor, LCPM – Groupe "Matériaux Inorganiques: Chimie
32
33 Douce et Cristallographie". BP. 523 Ziguinchor – Sénégal
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48

49 * To whom correspondence should be addressed.
50
51
52
53
54
55
56
57
58
59
60

ABSTRACT

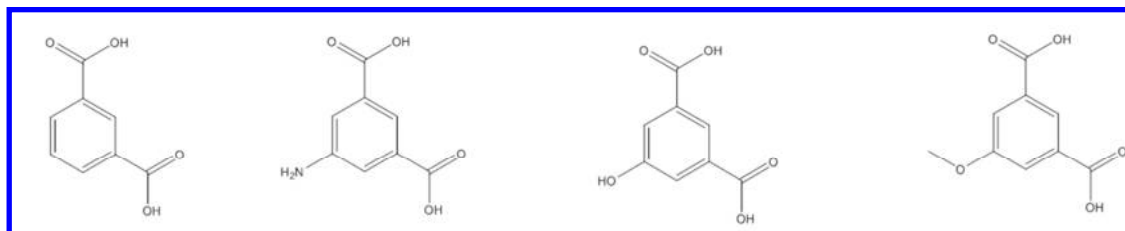
Reactions in water between lanthanide chlorides and the di-sodium salt of 5-methoxy-isophthalic acid, $\text{Na}_2(\text{mip}) \cdot 7\text{H}_2\text{O}$, lead to the first three series of lanthanide-based coordination polymers based on this ligand. The first series contains only one compound with chemical formula $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5 \cdot 2\text{H}_2\text{O}]_\infty$. The second series has general chemical formula $[\text{Ln}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ with $\text{Ln} = \text{La} - \text{Ce}$. The last family is constituted by compounds with general chemical formula $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $\text{Ln} = \text{Sm} - \text{Er}$ plus Y. Luminescent properties of the compounds that belongs to this family have been studied. The Tb-based compound presents one of the brightest luminescence reported to date for a lanthanide-based coordination polymer. The weak intermetallic energy transfer evidenced on the third family of compounds allows easy and predictable color modulation of the hetero-bimetallic powders through additive colors strategy.

KEYWORDS

Lanthanide • Coordination polymer • Crystal structure • Luminescence • Energy Transfer

INTRODUCTION.

Lanthanide-based coordination polymers have attracted great interest in the last two decades because of their potentially interesting porosity,^[1-9] magnetic^[10-13] or optical^[14-22] properties for example. More recently, lanthanide-based coordination polymers have proved their interest as taggants for fight against counterfeiting^[23] because they can present highly flexible luminescence.^[24, 25] Some of us are engaged in the synthesis and study of lanthanide-based coordination polymers for more than twenty years^[26] mainly focusing their attention on benzene-poly-carboxylate ligands.^[27] In recent years we have studied several series with isophthalate derivatives as ligands^[28-32] (See Scheme 1).



Scheme 1. From left to right, schematic representation of the ligand that have been investigated by our group: isophthalic acid (H_2ip),^[32] 5-amino-isophthalic acid (H_2aip),^[28, 31] 5-hydroxy-isophthalic acid (H_2oip)^[30] and 5-methoxy-isophthalic acid (H_2mip) (This work).

These series that present almost isorecticular crystal structures (Figure 1) exhibit promising optical properties that arise from competing transfers (ligand-to-metal and metal-to-metal energy transfers, photo-induced electron transfer...). These promising optical properties have encouraged us to undertake the study of lanthanide-based coordination polymers with 5-methoxy-isophthalate ligand hereafter referred as (mip)²⁻.

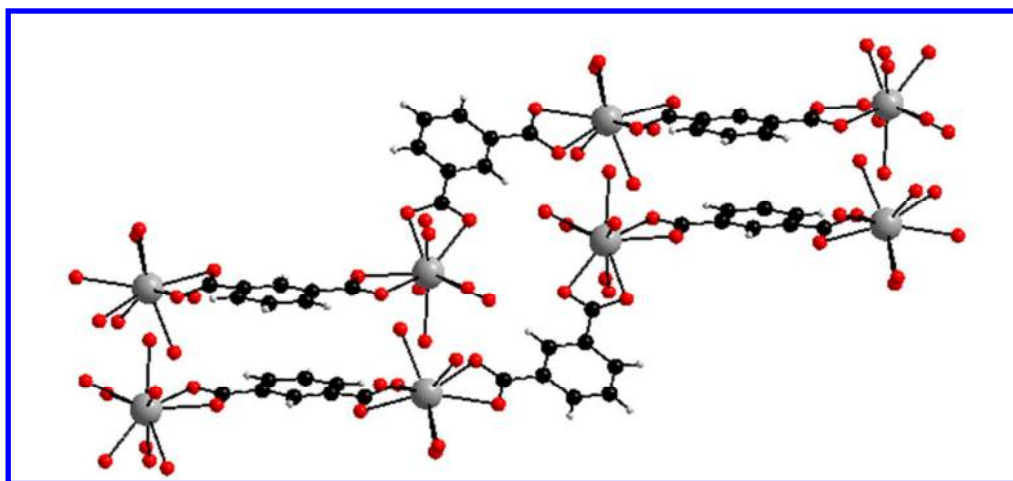


Figure 1. Projection view of a zig-zag double-chains molecular motif of $[\text{Gd}_2(\text{ip})_3(\text{H}_2\text{O})_9 \cdot 6\text{H}_2\text{O}]_\infty$. Redrawn from reference [32]. This molecular motif is also encountered in crystal structures of 5-amino-isophthalate- and 5-hydroxy-isophthalate-based lanthanide coordination polymers.^[28, 30, 31]

To the best of our knowledge, some coordination polymers based on $(\text{mip})^{2-}$ ligand and involving transition metal ions have recently been reported^[33-35] but, to date, none of them involves lanthanide ions. We want to report here the first three series of lanthanide-containing coordination polymers based on this ligand.

EXPERIMENTAL SECTION.

Synthesis and characterization of the ligand.

5-methoxy-isophthalic acid has been purchased from TCI (98%) and used without further purification. Two equivalents of sodium hydroxide are added to an aqueous suspension of 5-methoxy-isophthalic acid. The obtained clear solution is then evaporated to dryness and the residual solid is dissolved in ethanol and refluxed for one hour. Then precipitation is provoked by addition of an excess of ethoxy-ethane. The white precipitate is filtered and washed with ethoxy-ethane. After re-crystallization in deionized water, the powder is filtered and dried under ambient conditions. The yield is about 90%. Re-

crystallization has provided single crystals suitable for X-ray structure determination (Figure 2). Selected crystal and final structure refinement data are listed in Table 1.

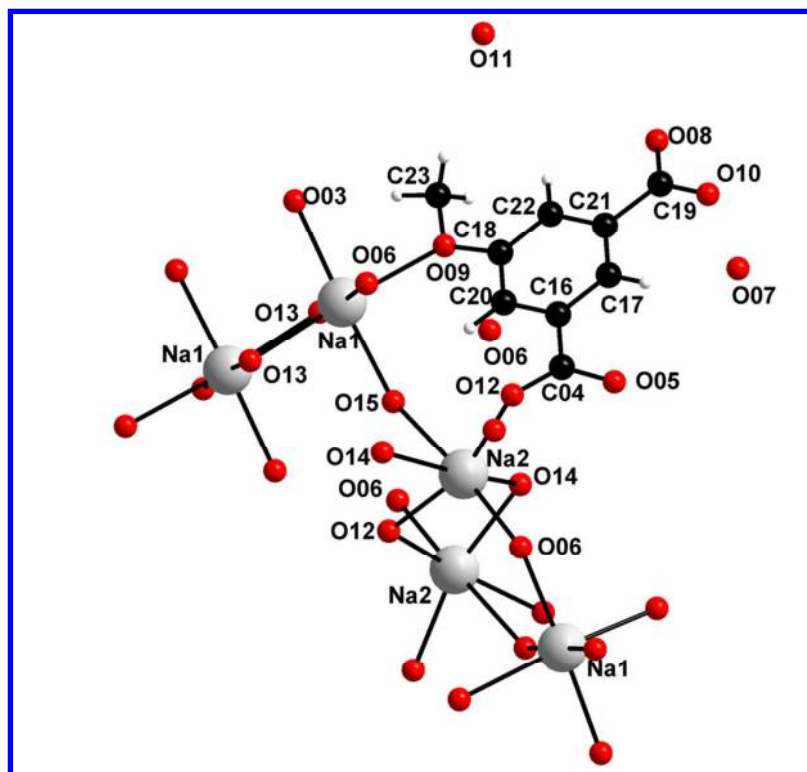


Figure 2. Projection view of an extended asymmetric unit of $\text{Na}_2(\text{mip}) \cdot 7\text{H}_2\text{O}$.

Table 1. Crystal and final structure refinement data for $\text{Na}_2(\text{mip}) \cdot 7\text{H}_2\text{O}$

Molecular formula	$\text{Na}_2\text{O}_{12}\text{C}_9\text{H}_{20}$
System	monoclinic
a (Å)	12.4922(6)
b (Å)	7.1710(3)
c (Å)	17.6871(8)
β (°)	96.114(2)
V (Å ³)	1575.43(12)
Z	4
Formula weight (g mol ⁻¹)	366.12
Space group (No.)	$P2_1/n$ (14)
D_{calc} (g cm ⁻³)	1.485
μ (mm ⁻¹)	0.185
R (%)	6.20
R_w (%)	22.20
GoF	0.999
N° CCDC	1501139

Powder X-ray diffraction pattern indicates that the powder and the single crystal are isostructural (Figure S1). Chemical analysis for $\text{Na}_2\text{C}_9\text{H}_{20}\text{O}_{12}$ (MW = 366 $\text{g}\cdot\text{mol}^{-1}$) found (calc.): Na 12.4% (12.5%); C 29.4% (29.5%); H 5.6% (5.5%); O 52.6% (52.5%).

IR spectrum does not show any characteristic band of protonated carboxylate functions (1410 cm^{-1}). Thermal analyses confirm the departure of 7 water molecules (Exp. 34.0%; Calc. 34.4%). First, there is a first weight loss between 50°C and 150°C that correspond to six water molecules (Exp. 29.3%; Calc. 29.5%). The seventh one is removed between 150°C and 200°C (Exp. 4.7%; Calc. 4.9%). This assumption is supported by IR spectra of the exhausted gas that have been recorded all along the thermal analysis (Figure S2). Liquid UV-visible absorption spectrum of a diluted aqueous solution ($2.73 \cdot 10^{-4}\text{ mol}\cdot\text{L}^{-1}$) of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ present a maximum at 299 nm with $\epsilon_{299} = 4320\text{ L}\cdot\text{mol}^{-1}\cdot\text{cm}^{-1}$ (Figure S3).

Synthesis and characterization of the coordination polymers as single crystals.

Lanthanide oxides (99.99%) were purchased from the AMPERE company. The chlorides were prepared according to literature procedures.^[36] Tetramethylorthosilicate (TMOS) and agarose gels were purchased from Acros Organics and used without further purification. They were jellified according to established procedures.^[37-39] Dilute aqueous solutions ($0.1\text{ mol}\cdot\text{L}^{-1}$) of the lanthanide chloride on one hand and of the di-sodium salt of methoxy-isophthalate on the other hand were allowed to slowly diffuse through a gel bridge in U-shaped tubes. After a few weeks, single crystals suitable for X-ray structure determination were obtained (See Figure 3 for an example).

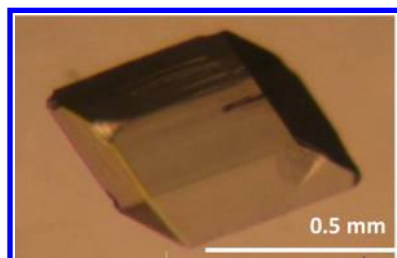


Figure 3. Picture of a single crystal of $[\text{La}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5\cdot\text{H}_2\text{O}]_\infty$.

Three types of crystal structures were identified depending on the involved lanthanide ion. Results are summarized in Table 2.

Table 2. Structural classification of the microcrystalline powders and of the single crystals.

Ln	La	Ce	Pr	Nd	Sm	Eu	Gd	Tb	Dy	Ho	Y	Er	Tm	Yb	Lu
Single crystal		■	■	■			■				■				
Microcrystalline powder	■	■			■	■	■	■	■	■	■	■	■	■	■
■	$[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$														
■	$[\text{Ln}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5\cdot\text{H}_2\text{O}]_\infty$														
■	$[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8\cdot 4\text{H}_2\text{O}]_\infty$														
□	unknown														

Synthesis and characterization of the coordination polymers as microcrystalline powders.

Microcrystalline powders of the coordination polymers were obtained by mixing stoichiometric aqueous solutions of a lanthanide chloride (1 mmol in 20 mL of deionized water) and of di-sodium methoxy-isophthalate (1.5 mmol in 20 mL of deionized water). Precipitations immediately occurred. Microcrystalline powders were filtered and dried in air. Yields are closed to 90%. Chemical analyses of the microcrystalline powders are listed in Table S1. On the basis of their powder X-ray diffraction diagrams, microcrystalline powders were classified in two series (See Figures S4 and S5): The first series contains the compounds that have been obtained with lanthanum or cerium while the second one contains compounds obtained with one of the lanthanide ion comprised between samarium and ytterbium plus yttrium. Other lanthanide ions lead to amorphous powders. Compounds that belong to the first family are isomorphous to $[\text{Gd}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5\cdot\text{H}_2\text{O}]_\infty$ and those that belong to the second family are isomorphous to $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8\cdot 4\text{H}_2\text{O}]_\infty$. Both crystal structures are described hereafter. It can be noticed that no microcrystalline powder present a powder X-ray diffraction diagram that corresponds to $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$ (Table 2).

Single crystal X-ray diffraction.

Single crystals of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ and of $[\text{Gd}(\text{mip})(\text{Hmip})](\text{H}_2\text{O})_5\cdot \text{H}_2\text{O}]_\infty$ were mounted on a APEXII AXS-Bruker diffractometer (150K) with Mo $K\alpha$ radiation ($\lambda = 0.71073 \text{ \AA}$) and single crystal of $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8\cdot 4\text{H}_2\text{O}]_\infty$ and $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$ on a D8Venture Bruker AXS diffractometer with multilayers monochromatized Mo $K\alpha$ radiation and equipped with a CMOS PHOTON100 detector (300K). Data reduction and cell refinement were performed with Denzo and Scalepack programs^[40] or Apex3(2015), Saint(V8.37a) and Sadabs(2014/5)^[41-43] according to the used diffractometer. The crystal structures have been solved by direct methods using SIR97 program,^[44] and refined with full matrix least-square methods based on F^2 (SHELX97^[45]) with WINGX program.^[46] All non-hydrogen atoms were refined anisotropically using SHELXL program. Hydrogen atoms bound to the organic ligand were located at ideal positions except those bound to the disordered ligand in the crystal structure of $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$ that have not been located. Hydrogen atoms of the water molecules were not located. Crystal and final structure refinement data of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ are listed in Table 1 and those of $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$, $[\text{Gd}(\text{mip})(\text{Hmip})](\text{H}_2\text{O})_5\cdot \text{H}_2\text{O}]_\infty$ and of $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8\cdot 4\text{H}_2\text{O}]_\infty$ in Table 3. Full details of the X-ray structure determination of the four crystal structures have been deposited with the Cambridge Crystallographic Data Center under the depository numbers CCDC-1501139 ($\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$), CCDC-1505544 ($[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5\cdot 2\text{H}_2\text{O}]_\infty$), CCDC-1501131 ($[\text{Gd}(\text{mip})(\text{H}(\text{mip}))(\text{H}_2\text{O})_5\cdot \text{H}_2\text{O}]_\infty$) and CCDC-1501130 ($[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8\cdot 4\text{H}_2\text{O}]_\infty$). They can be obtained free of charge at <http://www.ccdc.cam.ac.uk/conts/retrieving.html> [or from the Cambridge Crystallographic Data Centre, 12, Union Road, Cambridge CB2 IEZ, UK; fax: (internat.) +44-1223/336-033; E-mail: deposit@ccdc.cam.ac.uk], on request, from the authors and the reference to this publication.

Table 3. Crystal and final structure refinement data for $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5 \cdot 2\text{H}_2\text{O}]_\infty$, $[\text{Gd}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ and $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$.

Molecular formula	$\text{Ce}_2\text{C}_{27}\text{H}_{46}\text{O}_{29}$	$\text{GdO}_{16}\text{C}_{18}\text{H}_{25}$	$\text{Y}_2\text{O}_{27}\text{C}_{27}\text{H}_{18}$
System	monoclinic	triclinic	monoclinic
a (Å)	10.8750(2)	7.4437(3)	17.4214(11)
b (Å)	10.6070(1)	10.9493(4)	10.7884(8)
c (Å)	18.4790(2)	14.5247(6)	20.3688(15)
α (°)	-	89.238(2)	-
β (°)	100.700(1)	83.423(2)	104.136(2)
γ (°)	-	78.619(2)	-
V (Å ³)	2094.51(1)	1152.84(8)	3712.4(5)
Z	2	2	4
Formula weight (g.mol ⁻¹)	1114.88	654.53	952.23
Space group (No.)	$P2_1/c$ (14)	$P-1$ (2)	$P2_1/n$ (14)
D_{calc} (g.cm ⁻³)	1.768	1.851	1.704
μ (mm ⁻¹)	2.242	2.955	3.211
R (%)	5.14	2.86	3.33
R_w (%)	16.30	8.16	10.19
GoF	1.083	1.101	1.055
N° CCDC	1505544	1501131	1501130

Powder X-ray diffraction.

The diagrams have been collected using a PanalyticalX'Pert Pro diffractometer equipped with an X'Celerator detector. Calculated patterns were produced using POWDERCELL and WINPLOTR software programs.^[47, 48]

Thermal dependent X-ray diffraction diagrams have been produced with the same diffractometer. The samples are heating from room temperature to 1000°C using an Anton Parr HTK 1200 furnace under nitrogen atmosphere.

Solid state luminescent measurements.

Solid state emission and excitation spectra have been measured on a Horiba Jobin-Yvon Fluorolog III fluorescence spectrometer equipped with a Xe lamp 450 W and a UV-Vis photomultiplier (Hamamatsu R928, sensitivity 190-860 nm). Most of the luminescence

1
2
3 spectra were recorded between 350 nm and 725 nm at room temperature. The quantum yield
4
5 measurements were performed using a Jobin-Yvon integrating sphere ($\Phi = (E_c - E_a)/(L_a - L_c)$)
6
7 with E_c being the integrated emission spectrum of the sample, E_a the integrated “blank”
8
9 emission spectrum, L_a the “blank” absorption and L_c the sample absorption at the excitation
10
11 wavelength). The emission/excitation spectra and quantum yield recordings were realized on
12
13 powder samples introduced in cylindrical quartz cells of 0.7 cm diameter and 2.4 cm height,
14
15 which it is placed directly inside the integrating sphere. The luminescence of the Gd-based
16
17 microcrystalline powder has been measured at 77 K, the sample is introduced in a quartz
18
19 capillary tube, which it is placed inside a small Dewar containing nitrogen liquid.
20
21 Luminescence decays have also been measured at room temperature using this apparatus with
22
23 a Xe flash lamp (phosphorescence mode). Lifetimes and quantum yields are averages of three
24
25 independent determinations. Appropriate filters were used to remove the residual excitation
26
27 laser light, the Rayleigh scattered light and associated harmonics from spectra. All spectra
28
29 were corrected for the instrumental response function.
30
31
32

33
34 Comparative solid state luminescent spectra have been measured on a Perkin-Elmer
35
36 LS-55 spectrometer with a Xe flash lamp and equipped with a UV-Vis photomultiplier
37
38 (Hamamatsu R928) between 450 and 725 nm under identical operating conditions and
39
40 without turning the lamp off to ensure a valid comparison between the emission spectra. All
41
42 spectra have been recorded with the phosphorescence mode, not with the standard
43
44 fluorescence mode (delay and gate times are applied to remove residual light from sample
45
46 holder). Automatic filters were also used during the measurements as for the Fluorolog III.
47
48

49
50 Luminescence intensities of the samples expressed in $\text{Cd}\cdot\text{m}^{-2}$ have been measured with
51
52 a Gigahertz-Optik X1-1 optometer with an integration time of 200 ms on 1.5 cm^2 pellets. The
53
54 intensity of the UV flux at sample location, $0.385\text{ mW}\cdot\text{cm}^{-2}$, has been measured with a
55
56 VilberLourmat VLX-3W radiometer. $[\text{Tb}_2(\text{bdc})_3\cdot 4\text{H}_2\text{O}]_\infty$ where bdc^{2-} stands for terephthalate
57
58
59
60

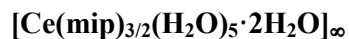
1
2
3 was used as a standard. Its luminance is 142 Cd.m^{-2} under these operating conditions
4
5 $(\lambda_{\text{exc}} = 312 \text{ nm; flux} = 0.385 \text{ mW.cm}^{-2})$.^[49]
6
7

8 9 10 **Colorimetric measurements.**

11 The CIE (Commission Internationale de l'Eclairage) (x,y) emission color
12 coordinates^[50, 51] were obtained using a MSU-003 colorimeter (Majantys) with the
13 PhotonProbe 1.6.0 Software (Majantys). Color measurements: 2°, CIE 1931, step 5 nm, under
14 312 nm UV light. $X = k \times \int_{380\text{nm}}^{780\text{nm}} I_{\lambda} \times x_{\lambda}$, $Y = k \times \int_{380\text{nm}}^{780\text{nm}} I_{\lambda} \times y_{\lambda}$ and $Z = k \times$
15 $\int_{380\text{nm}}^{780\text{nm}} I_{\lambda} \times z_{\lambda}$ with k constant for the measurement system I_{λ} sample spectrum intensity,
16 wavelength depending, x_{λ} , y_{λ} , z_{λ} trichromatic values $x = X/(X+Y+Z)$, $y = Y/(X+Y+Z)$ and $z =$
17 $Z/(X+Y+Z)$. Mean xyz values are given for each sample, which act as light sources
18 (luminescent samples). Standards from Phosphor Technology used, calibrated at 312 nm: red
19 phosphor $\text{Gd}_2\text{O}_2\text{S:Eu}$ ($x = 0.667$, $y = 0.330$) and green phosphor $\text{Gd}_2\text{O}_2\text{S:Tb}$ ($x = 0.328$, $y =$
20 0.537).
21
22
23
24
25
26
27
28
29
30
31
32
33
34
35
36
37

38 **Coupled Thermal and IR analyses**

39 Coupled thermal (ATG/DSC) and IR analyses have been performed using a Perkin
40 Elmer STA6000 thermal analyzer coupled to a Perkin Elmer Frontier IR spectrophotometer
41 analyzer. Coupling was insured by a Perkin Elmer TL8000 transfer line. This allows
42 simultaneous recording of the weight of the sample, the heat flux and the IR spectrum of the
43 exhausted gases versus temperature. Measurements were performed in ceramic crucibles
44 under a nitrogen atmosphere between room temperature and 900°C with a $20^\circ\text{C.min}^{-1}$ heating
45 rate. At the end of the experiment, the compound was maintained for one hour at 900°C under
46 air atmosphere in order to complete the combustion.
47
48
49
50
51
52
53
54
55
56
57
58
59
60

RESULTS AND DISCUSSION.

This compound has been obtained as single crystals only. Its crystal structure can be described as the superimposition of 1D molecular chains in which Ce^{3+} ions and $(\text{mip})^{2-}$ ligands alternate (Figure 4). There is only one independent Ce^{3+} ion in the crystal structure that is nine coordinated by five oxygen atoms from coordination water molecules and four oxygen atoms from two carboxylate clips that form a slightly distorted muffin (Table S2). A fully deprotonated ligand with a half occupancy factor is located between the chains and ensures the electro-neutrality of the crystal packing. Stability of the crystal packing is ensured by a complex hydrogen-bonds network that involves crystallization water molecules as well as π -stacking interactions. Indeed, phenyl groups of the free and coordinated ligands stack parallel to the a -axis. Free $(\text{mip})^{2-}$ ligands are disordered. The disorder model is drawn in scheme 2. Intra-molecular chain Ce-Ce shortest distances are 10.2 Å. Shortest distances between Ce^{3+} ions that belong to neighboring chains are 10.8 Å along the a -axis and 6.0 Å and 6.9 Å in the (bc) plane.

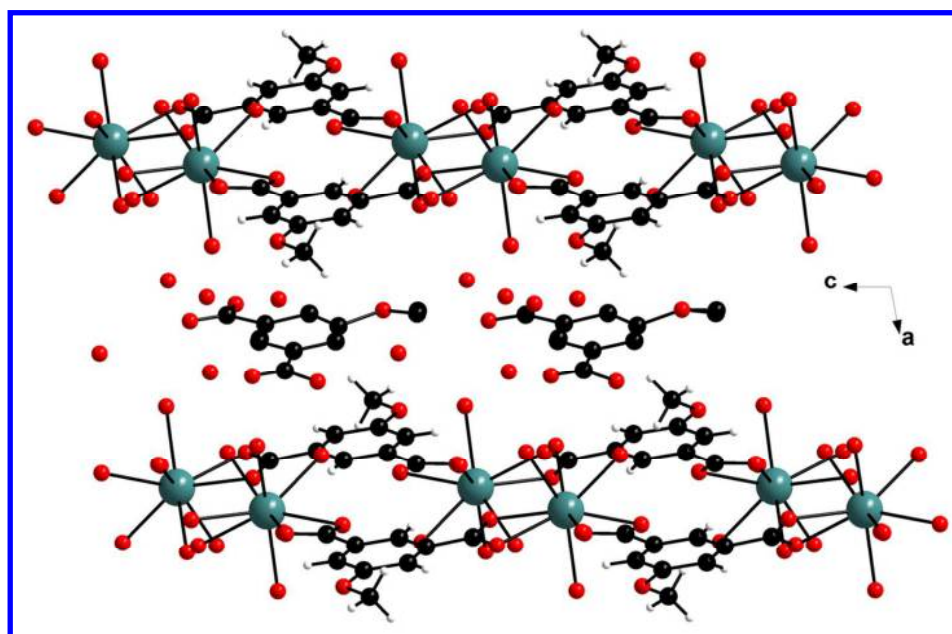
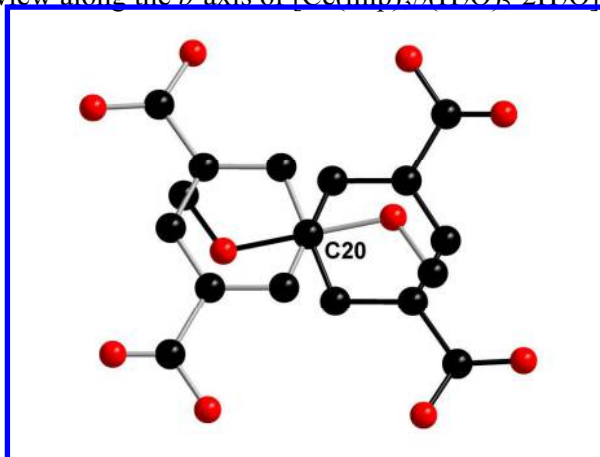


Figure 4. Projection view along the *b*-axis of $[\text{Ce}(\text{mip})_3(\text{H}_2\text{O})_5 \cdot 2\text{H}_2\text{O}]_n$.



Scheme 2. Schematic representation of the uncoordinated $(\text{mip})^{2-}$ ligand. The two equiprobable configurations of the free ligand, the one with light grey bonds and the one with black bonds, share the C20 atom.

$[\text{Ln}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ with Ln = La, Ce, Pr or Gd

Compounds that belong to the second series have general chemical formula $[\text{Ln}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ where Ln symbolizes a lanthanide ion. These compounds have been obtained in gels as single crystals for Ln = La, Pr, Gd but they have only been obtained as microcrystalline powders for Ln = La or Ce. Crystal structure has been solved on the Gd-containing compound and isostructurality of the La-, Ce- and Pr-derivatives has been assumed on the basis of X-ray diffraction measurements (Figure S4). This crystal structure can be described as the juxtaposition of 1D ribbon-like molecular motifs that spread parallel to the *b*-axis (Figure 5). There is only one crystallographically independent Gd^{3+} ion that is nine coordinated by four oxygen atoms from carboxylate functions and five oxygen atoms from coordination water molecules that form a slightly distorted capped square antiprism (Table S3)^[52]. Coordination modes of the $(\text{mip})^{2-}$ ligand that links the Gd^{3+} ions are drawn in Figure 5.

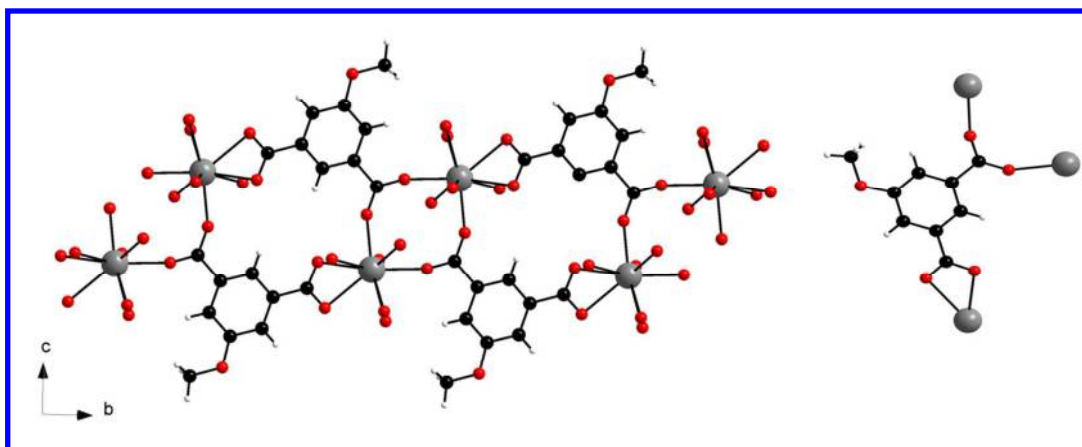
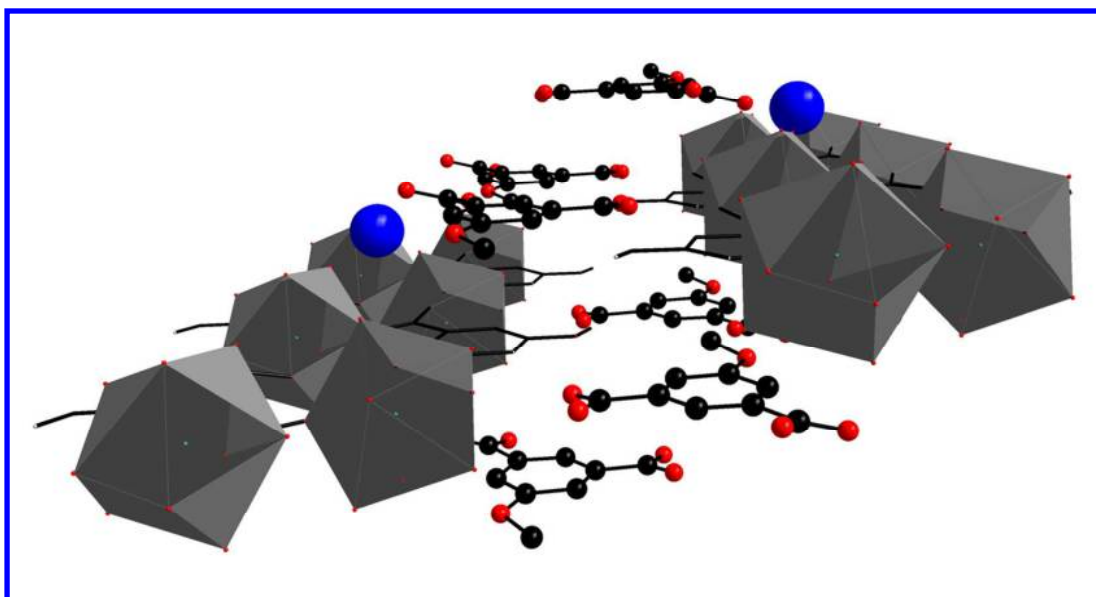


Figure 5. Left: Projection view along the a -axis of a ribbon-like molecular motif of $[\text{Gd}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$. Right: coordination modes of the ligand $(\text{mip})^{2-}$.

In addition, there is one crystallization water molecule and one half-protonated ligand that lay between the ribbon-like molecular motifs (Scheme 3). Presence of an uncoordinated ligand between molecular motifs, that insures electro-neutrality of the crystal structure, has already been observed in lanthanide-based coordination polymers obtained with 5-hydroxy-benzene-1,3-di-carboxylate.^[30] A complex network of hydrogen bonds and π -stacking interactions insure the stability of the crystal structure.

Coupled thermal (ATG/DSC) and IR analyses have been performed (Figure S6). They show the departure of the six water molecules per lanthanide ion at about 100°C (Exp. 20% - Calc. 17.3%)



Scheme 3. Perspective view of the half-protonated ligand (ball-and-stick model) between two ribbon-like molecular motifs in $[\text{Gd}(\text{mip})(\text{H}(\text{mip}))(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$. Blue balls symbolize crystallization water molecules. Lanthanide ions coordination polyhedrons have been drawn and hydrogen atoms have been omitted for clarity.

It can be noticed that this crystal structure differs from the crystal structures of the lanthanide-based coordination polymers obtained with isophthalate, 5-amino-isophthalate or 5-hydroxy-isophthalate.^[28, 30-32]

$[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with Ln = Sm-Yb plus Y.

Compounds that have been obtained with lanthanide ions comprised between samarium and ytterbium plus yttrium constitute a family of isostructural compounds. Isostructurality of the compounds has been assumed on the basis of their powder X-ray diffraction patterns (Figure S5). The crystal structure has been solved on the basis of an Y-containing single crystal. Once again this mono-dimensional crystal structure, is completely different from those that had been obtained with isophthalate, 5-amino-isophthalate or 5-hydroxy-isophthalate.^[28, 30-32] There are two independent Y^{3+} ions in this crystal structure. Y1 and Y2 are both nine-coordinated by five oxygen atoms from three carboxylate functions and

four coordination water molecules that form slightly distorted capped square antiprisms^[52] (Figure 6 and Table S4). Two out of the three independent ligands bind Y1 and Y2 in a bis-bidentate manner while the third one binds them in a bis-monodentate way (Figure 7).

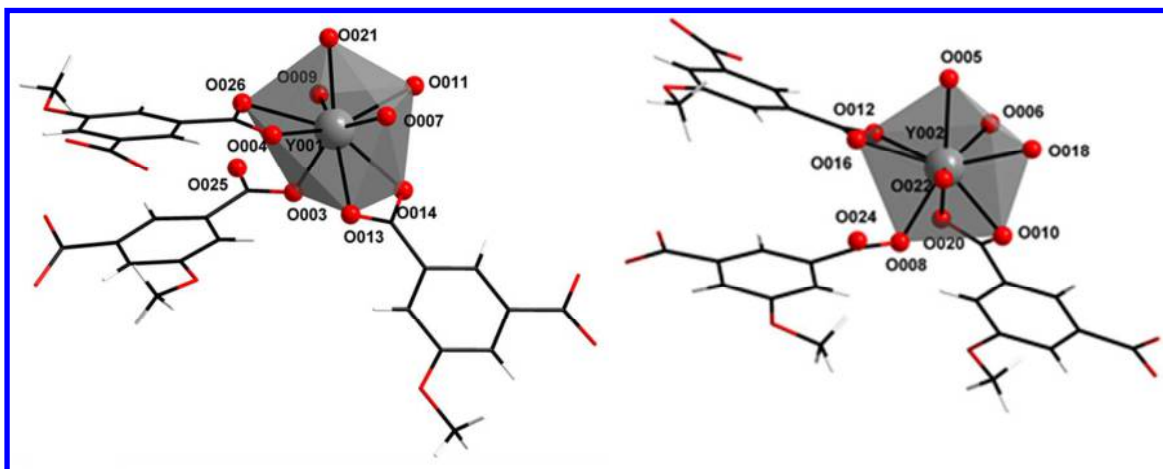


Figure 6. Extended coordination polyhedrons of Y1 (left) and Y2 (right) in $[Y_2(mip)_3(H_2O)_8 \cdot 4H_2O]_\infty$.

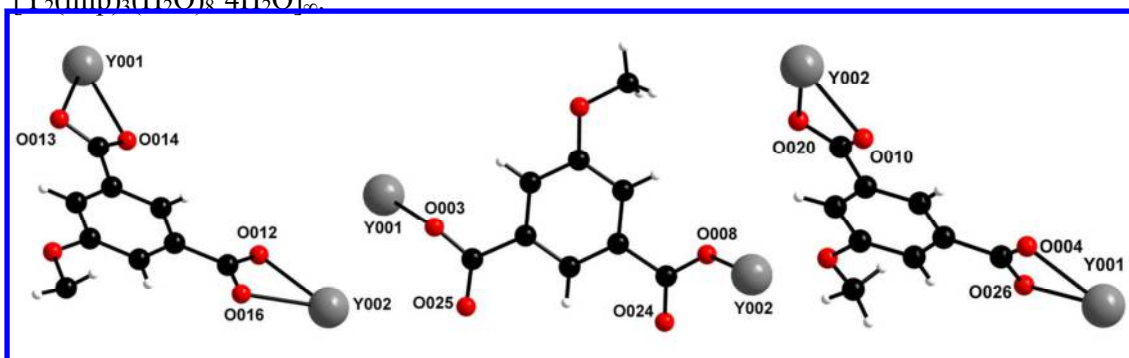


Figure 7. Coordination modes of the three independent ligands in $[Y_2(mip)_3(H_2O)_8 \cdot 4H_2O]_\infty$.

The crystal structure can be described as the juxtaposition of wrapped ladder-like molecular motifs that spread parallel to the *a*-axis. Uprights are constituted by an alternation of Y^{3+} ions linked together by bis-bidentate ligands. Bis-monodentate ligands constitute the rungs of the ladder (Figure 8). Intermetallic distances between adjacent lanthanide ions that belong to the same molecular motif are close to 10 Å. These quite long intermetallic distances are rather rare in lanthanide-based coordination polymers made of poly-carboxylates as ligands because carboxylate functions adopt commonly a bridging bis-monodentate

coordination mode that leads to short Ln-Ln distances ($\sim 4.5 \text{ \AA}$).^[24] However, ladder-like molecular motifs stack in such a way that the shortest intermetallic contacts between lanthanide ions that belong to neighboring motifs are only 6 \AA (See scheme 4).

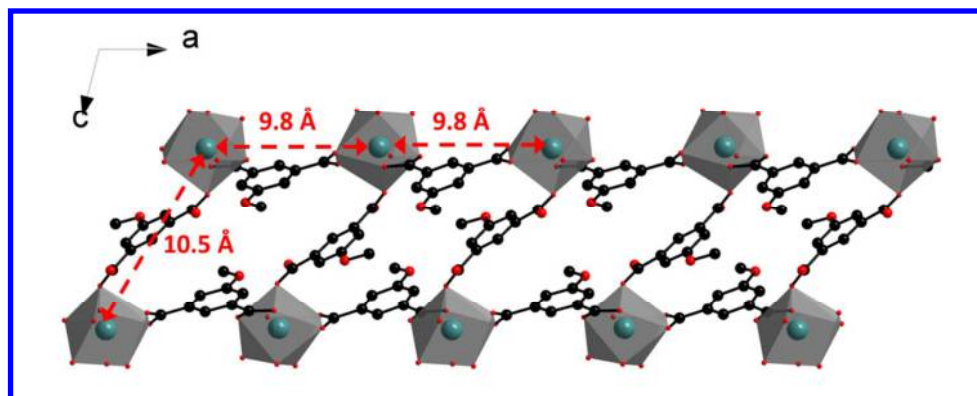
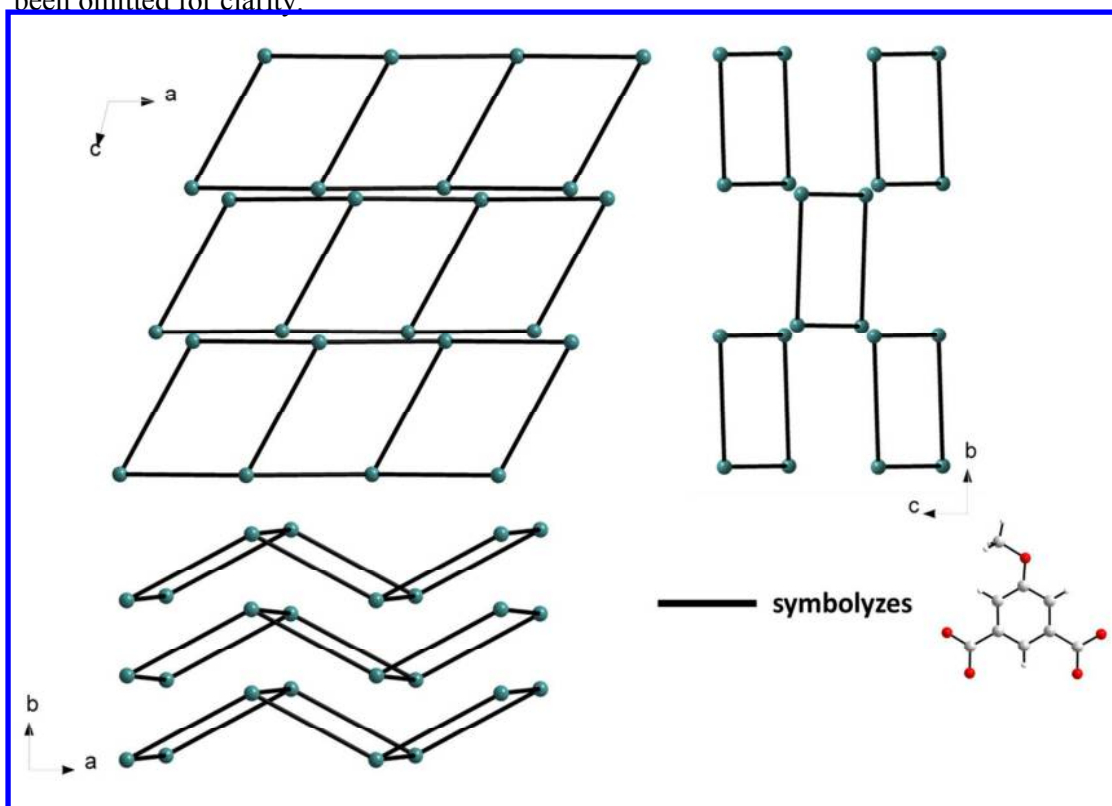


Figure 8. Ladder-like molecular motif in $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$. Hydrogen atoms have been omitted for clarity.



Scheme 4. Schematic projections views of the crystal packing of $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$.

1
2
3 With the rough model that has been described previously,^[53] the mean volume
4 occupied by an Y^{3+} ion in this crystal structure is $\bar{v} = 464 \text{ \AA}^3$, that is, the mean distance
5 between two lanthanide ions is about 9.6 \AA ($r = \sqrt[3]{\frac{3\bar{v}}{4\pi}}$). This is quite interesting as far as
6
7
8
9
10
11 luminescence properties are targeted. Indeed, it is well known^[54] that intermetallic energy
12 transfers are less efficient when the lanthanide ions are more than 10 \AA far from each other.
13
14
15 Therefore, in this crystal structure these energy transfers are expected to be weak.
16

17
18 Stability of the crystal packing is ensured by a strong hydrogen-bonds network that
19 involves coordination and crystallization water molecules as well as oxygen atoms that belong
20 to the ligands (See Table S5).
21
22
23

24
25 Coupled thermal analyses that have been performed on the basis of the Y-containing
26 compound indicate a first weight loss centered at 125°C that corresponds to the departure of
27 the twelve water molecules per formula unit (Exp.: 21.9% - Calc.: 22.6%). The dehydrated
28 phase that is obtained remains stable up to about 500°C . At last the ligand decomposes and
29 Y_2O_3 is obtained (See Figure S7). Thermo-dependent powder X-ray diffraction shows that the
30 dehydrated phase observed between 100°C and 500°C is different from the initial hydrated
31 one (Figure S8). Moreover, even when exposed to wet atmosphere, this phase does not
32 rehydrate (Figure S9). The irreversibility of the structural change and the robustness of the
33 dehydrated phase could be related to the departure of four coordination water molecules per
34 lanthanide ion and the presence of a non-coordinated oxygen atom from a carboxylate
35 function in the vicinity of the under-crowded lanthanide ion that is available for coordination.
36
37
38
39
40
41
42
43
44
45
46
47

48
49 This series of compounds is of particular interest as far as strongly luminescent
50 compounds are targeted. Indeed, the quite long mean intermetallic distance is supposed to
51 favor luminescence because negligible intermetallic energy transfer is expected. Colorimetric
52 coordinates and luminance of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with Ln = Sm-Dy are reported in
53 figure 9 and table S6.
54
55
56
57
58
59
60

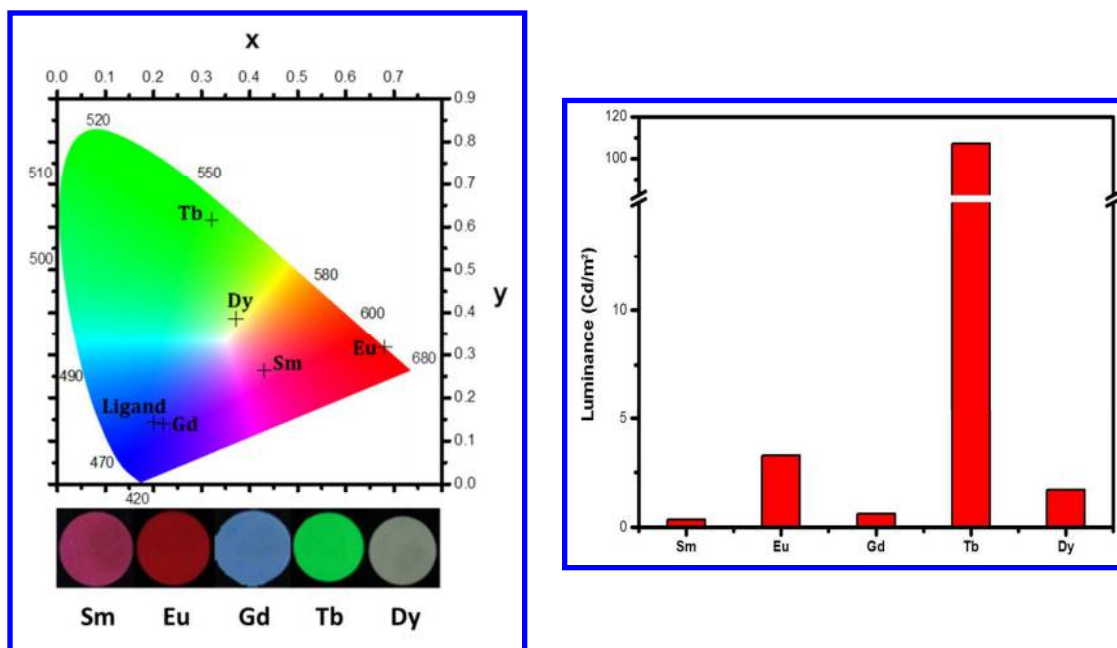


Figure 9. Left: Colorimetric coordinates and pictures under UV irradiation of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $\text{Ln} = \text{Sm-Dy}$ ($\lambda_{\text{exc}} = 312 \text{ nm}$). Right: Luminance of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $\text{Ln} = \text{Sm-Dy}$ under UV excitation ($\lambda_{\text{exc}} = 312 \text{ nm}$).

These measurements indicate that the luminance of the Tb-containing compound is significant (107 cd.m^{-2}). Actually, to the best of our knowledge, among all the lanthanide-containing coordination polymers with benzene-poly-carboxylate derivatives as ligands whose luminance has been measured to date, only $[\text{Tb}_2(\text{bdc})_3 \cdot 4\text{H}_2\text{O}]_\infty$ (where bdc^{2-} stands for terephthalate) exhibits a higher luminance (142 cd.m^{-2} under similar operating conditions).^[49]

Excitation and emission spectra of the compounds that involve Sm^{3+} , Eu^{3+} , Tb^{3+} or Dy^{3+} have been recorded (See Figure 10) and their calculated luminescence lifetimes and overall quantum yields are given in Table 4.

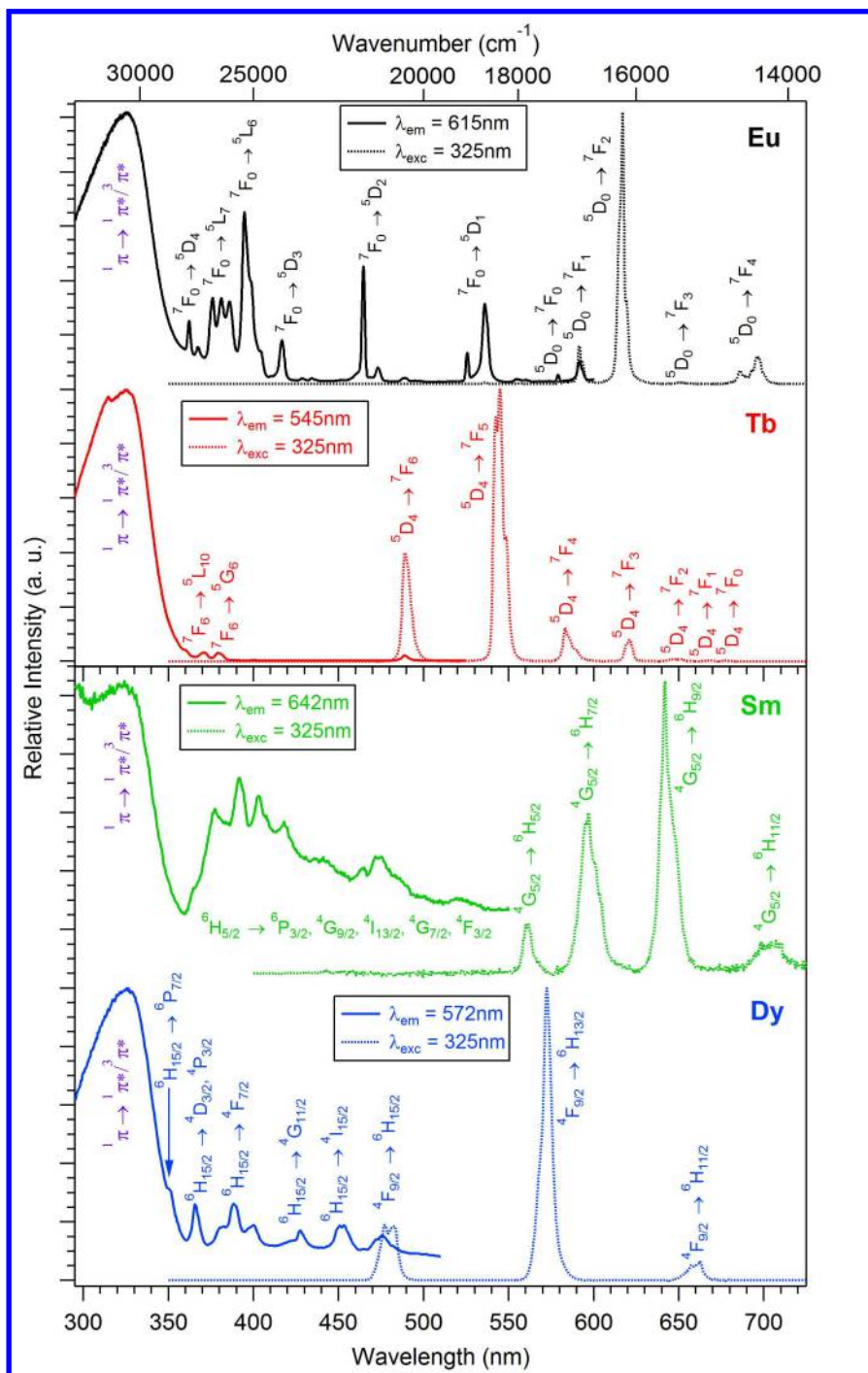


Figure 10. Excitation and emission spectra of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $\text{Ln} = \text{Eu}$ (black curves), Tb (red curves), Sm (green curves) and Dy (blue curves). $\lambda_{\text{exc}} = 325 \text{ nm}$.

Table 4. Overall quantum yields and luminescence lifetimes of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with Ln = Sm, Eu, Tb or Dy ($\lambda_{\text{exc}} = 325\text{nm}$).

	$Q_{\text{Ln}}^{\text{Ligand}} (\%)$	τ_{obs}
$[\text{Sm}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$	0.18(5)	< 10 μs
$[\text{Eu}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$	2.8(2)	0.21(1) ms
$[\text{Tb}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$	41(3)	0.71(1) ms
$[\text{Dy}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$	0.68(8)	< 10 μs

These measurements indicate that the antenna effect is effective for all compounds. Indeed the four compounds can be excited at the same wavelength that corresponds to ligand absorption. This was expected because the energy of the first excited singlet state of the ligand (29000 cm^{-1} estimated on the basis of the excitation spectrum (Figure S10) of $[\text{Gd}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$) and of its first excited triplet state (24400 cm^{-1} estimated on the basis of the emission spectrum (Figure S10) recorded at 77K of $[\text{Gd}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ ^[55]) are supposed, according to Latva's^[56] and Reinholdt's^[57] empirical rules, to favor efficient antenna effect ($\Delta E(^1\pi\pi^* - ^3\pi\pi^*) = 4600 \text{ cm}^{-1} < 5000 \text{ cm}^{-1}$). Additionally, it can be noticed that the phosphorescent lifetime of the triplet state is much longer ($\tau_{\text{obs}} = 39(3) \text{ ms}$) than the one measured for other coordination polymers ($\tau_{\text{obs}} = 7.0(1) \text{ ms}$ for $[\text{Gd}_2(\text{bdc})_3 \cdot 4\text{H}_2\text{O}]_\infty$).^[49]

From a global point of view, it can be noticed that the overall quantum yields of the Eu-based compounds is very low compared with the one of the Tb-based compound (a factor 15 approximately). Because $\text{CH}_3\text{-O-}$ is a +M and -I group, a photo-induced electron transfer (PET) may occur for the Eu-based compound as observed for 5-hydroxy-isophthalate and 5-amino-isophthalates derivatives^[28, 30, 31]. On the other hand, intrinsic quantum yield is low ($Q_{\text{Eu}}^{\text{Eu}} = 9.0(2)\%$ measured with $\lambda_{\text{exc}} = 464.5 \text{ nm}$ that corresponds to the $^7\text{F}_0 \rightarrow ^5\text{D}_2$ transition) which induces a low overall quantum yield despite sizeable sensitization ($\eta_{\text{sens}} = 30\%$ ^[54]).

Unfortunately, it has not been possible to evaluate accurately the intrinsic quantum yield of the terbium derivative because the direct excitation transition ($^7\text{F}_6 \rightarrow ^5\text{G}_6$) overlaps with the excitation band of the ligand. Furthermore, an excitation of the following $^7\text{F}_6 \rightarrow ^5\text{D}_4$

1
2
3 transition at 489 nm does not allow accurate calculation of the intrinsic quantum yield of the
4
5 Tb³⁺ ion because of a weak absorption. However, tentative measurements that have been
6
7 performed seem to indicate that the sensitization of the Tb³⁺ ion by the ligand is very high and
8
9 likely close to 100%.

10
11 For Sm- and Dy-based compounds, the overall quantum yield are lower than 1%
12
13 indicating a weak antenna effect compared to Tb-based compounds (a factor 230 and 60
14
15 approximately with Sm and Dy, respectively). This can be attributed to the small energy gap
16
17 between the emitting levels and the receiving levels of these two lanthanide ions.^[54]

18
19
20 In the last years we have undertaken the investigation of heteronuclear lanthanide-
21
22 based coordination polymers of formula [Ln_{2-2x}Ln'_{2x}(L)₃(H₂O)_y]_∞ (with L= benzene-
23
24 polycarboxylate ligand) in order to afford materials where both brightness and color can be
25
26 modulated.^[14, 16, 30, 49] Such modification of the emissive properties when compared with
27
28 homonuclear-based compounds is possible if strong intermetallic energy transfers are
29
30 observed. In fact, for a given emissive compound, dilution with optically inactive lanthanide
31
32 ion (Gd³⁺, ...) can enhance the brightness up to 20%^[49] whereas dilution with optically active
33
34 ones lead to color change. This last modification is however rather tricky to control as the
35
36 obtained color depends of the relative position of the emissive energy levels as well as the
37
38 intrinsic quantum yield of each lanthanides. As a result, one lanthanide can overwhelm the
39
40 emission of the other as observed on [Tb_{2-2x}Eu_{2x}(bdc)₃(H₂O)₄]_∞ where full red emission is
41
42 visible for x as low as 0.2.^[49]

43
44
45 The high Ln³⁺-Ln³⁺ mean distance in family 2, makes it a nice platform to test color
46
47 modulation procedure with more predictable emission, ideally “additive color synthesis” as
48
49 performed with classic dyes.

50
51 We have investigated two series of hetero-lanthanide coordination polymers with
52
53 respective chemical formula [Tb_{2-2x}Gd_{2x}(mip)₃(H₂O)₈·4H₂O]_∞ and

[$\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}$] $_{\infty}$ with $0 \leq x \leq 1$. The first series aims to quantify the intermetallic energy transfers and the second one to investigate the color modulation ability. (Figure 11). All these compounds present the same crystal structure than the homo-lanthanide compounds (Figures S11 and S12). Accurate metallic composition of each compound has been measured by EDS measurements (Tables S7 and S8). Powder X-ray diffraction patterns show neither evidence of segregation of the lanthanide ions nor bi-phasic character of the microcrystalline powders.^[49]

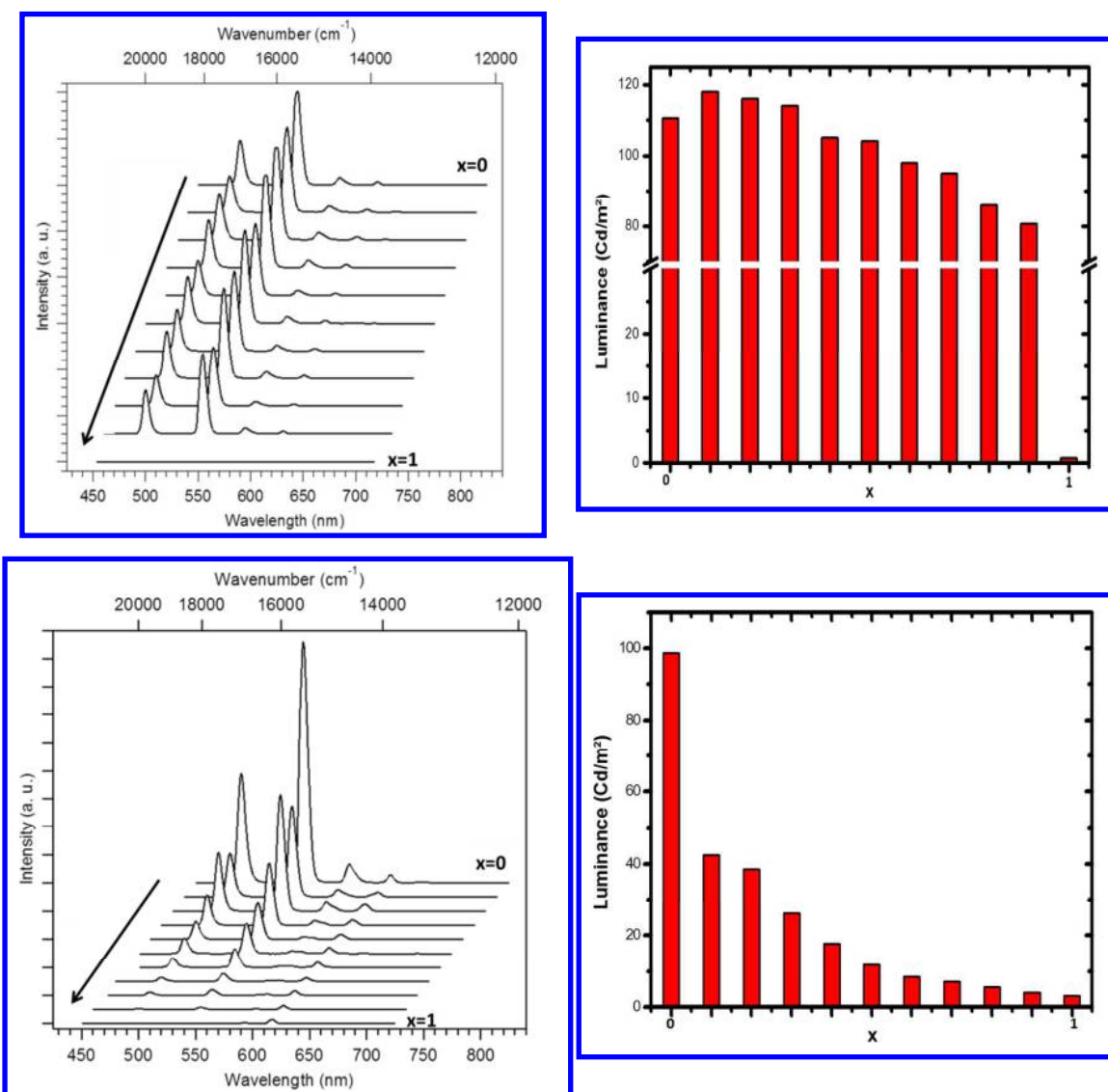


Figure 11. Emission spectra (left) and luminance measurements (right) under UV radiation ($\lambda_{\text{exc}} = 325 \text{ nm}$) of [$\text{Tb}_{2-2x}\text{Gd}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}$] $_{\infty}$ with $0 \leq x \leq 1$ (top) and [$\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}$] $_{\infty}$ with $0 \leq x \leq 1$ (bottom).

1
2
3
4
5 As anticipated, data that concern $[\text{Tb}_{2-2x}\text{Gd}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ clearly show that
6
7 there is little influence of the dilution of the Tb^{3+} ions by optically inactive Gd^{3+} ions. Indeed,
8
9 luminance increases by only 7% for $x = 0.1$ which corresponds to a mean terbium-to-terbium
10
11 distance of 10.7 Å. Then, it decreases progressively with the terbium content. This clearly
12
13 evidences the weak Tb-Tb intermetallic transfer within the series. In order to quantify this
14
15 observation we have calculated the intermetallic energy transfer rate ($\eta_{\text{ET}} = 31(3) \%$) using
16
17 the relationship $\eta_{\text{ET}} = 1 - \frac{\tau_{\text{obs}}}{\tau_0}$ (where τ_{obs} and τ_0 are the luminescent lifetimes in presence or
18
19 in absence of an acceptor respectively^[54]) on the basis of the compounds
20
21 $[\text{TbEu}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ ($\tau_{\text{obs}} = 0.45(3)$ ms) and $[\text{TbGd}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ ($\tau_0 = 0.72(2)$
22
23 ms). As a matter of comparison η_{ET} was 95% for $[\text{TbEu}(\text{bdc})_3(\text{H}_2\text{O})_4]_\infty$.^[49]
24
25
26
27

28 The second series, $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$, show a significant color change
29
30 upon Eu addition but with a trend that is significantly different from the
31
32 $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{bdc})_3(\text{H}_2\text{O})_4]_\infty$ series (figure 12).^[24, 49] As foreseen above, a regular variation from
33
34 green to red is observed on the powders. This can be related to the weak Tb-to-Eu energy
35
36 transfer. It is worth notice that similar trend is visible on single crystals that have been
37
38 obtained by slow evaporation of the filtrates retrieved after the microcrystalline powders
39
40 preparation. This further strongly suggests the monophasic character of the microcrystalline
41
42 powders and the random distribution of the lanthanide ions over the metallic sites of the
43
44 crystal structure.
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

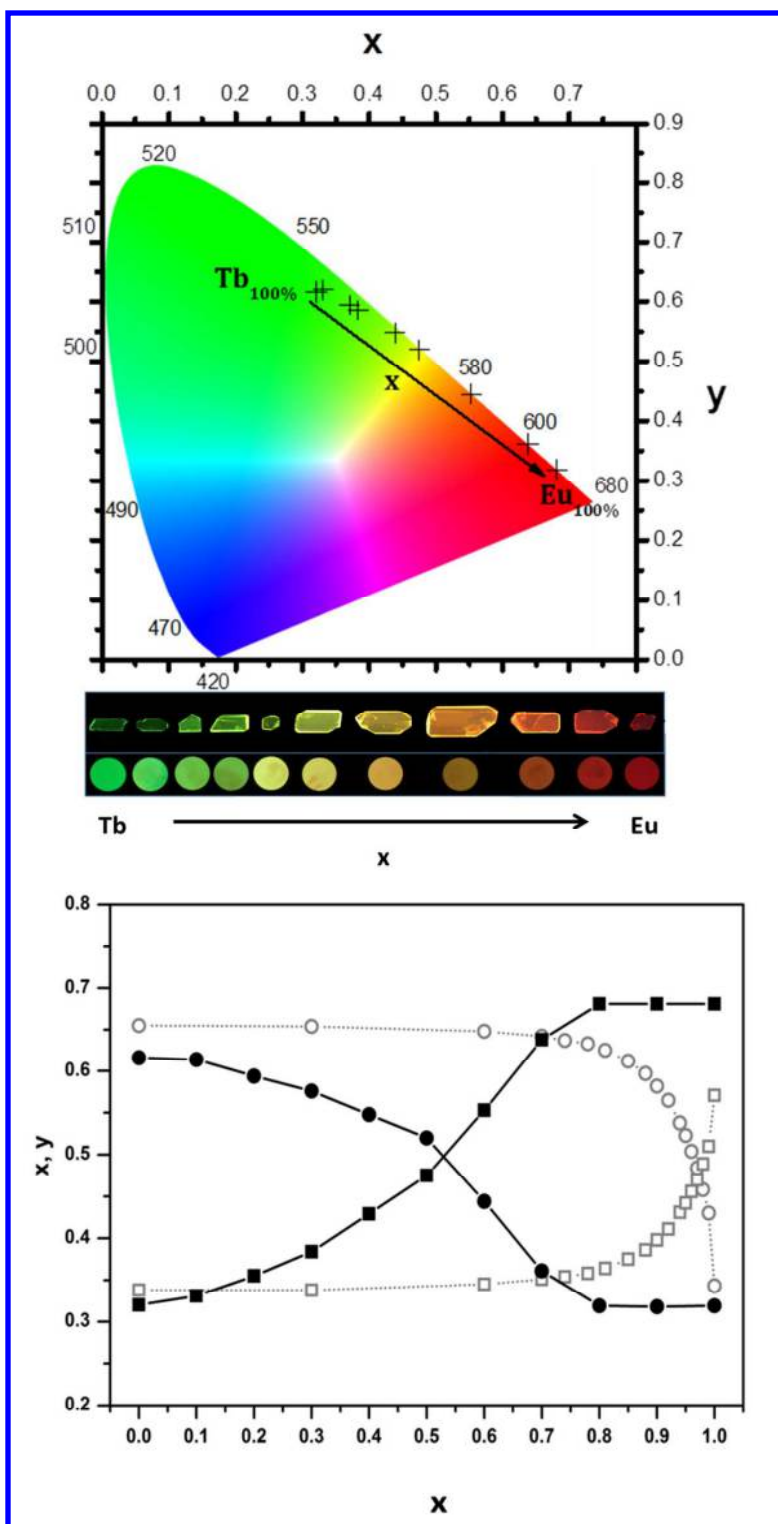


Figure 12. Colorimetric coordinates (top) and pictures of pellets of microcrystalline powders and of single crystals (middle) under UV irradiation ($\lambda_{\text{exc}} = 312\text{nm}$) of $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $0 \leq x \leq 1$. Bottom; comparison of the evolution of the colorimetric coordinates versus x of $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ (low intermetallic energy transfer) in black and of $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{bdc})_3(\text{H}_2\text{O})_4]_\infty$ (strong intermetallic transfer) in

1
2
3 light grey. x and y colorimetric coordinates are symbolized by circles and squares
4 respectively. Curves that correspond to $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{bdc})_3(\text{H}_2\text{O})_4]_\infty$ series are redrawn from
5 reference^[49].
6
7

8 9 **CONCLUSION AND OUTLOOKS**

10
11 In this paper, we described the first three series of lanthanide-based coordination
12 polymers with 5-methoxy-isophthalate as ligand. One of the families is particularly interesting
13 as far as potential applications are targeted. Indeed, the bright luminescence exhibited by
14 some of the compounds that constitute this series associated with the weak intermetallic
15 transfer allow new lanthanide ions combinations with new luminescent properties. In
16 particular, the color modulation within a heteronuclear family is particularly efficient and
17 predictable and can be based on a simple colors addition technique. This strategy could lead
18 to molecular bar codes exhibiting luminescent properties in different spectral domains. We are
19 convinced that this opens new opportunities in the field of anti-counterfeiting taggants. Our
20 group is currently working along this line.
21
22
23
24
25
26
27
28
29
30
31
32
33
34

35 **ACKNOWLEDGEMENTS**

36
37 The French Cooperation Agency in Senegal is acknowledged for financial support.

38
39 The CDIFX of Rennes is acknowledged for single crystal X-ray diffraction data collection.
40
41
42
43

44 **SUPPORTING INFORMATION**

45
46 Experimental and simulated powder X-ray diffraction patterns of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ (Figure S1).
47
48 ATG/DSC of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ between room temperature and 300°C under N_2 flux. IR spectra
49 of the exhausted gas during thermal analyses versus temperature. IR spectrum recorded at
50 $T=90^\circ\text{C}$ (Figure S2). Liquid state UV-visible absorption spectrum of a diluted aqueous
51 solution of $\text{Na}_2(\text{mip})\cdot 7\text{H}_2\text{O}$ (Figure S3). Chemical analyzes for compounds that constitute
52 families **(1)** and **(2)** (Table S1). Experimental and simulated powder X-ray diffraction patterns
53
54
55
56
57
58
59
60

1
2
3 of compounds with general chemical formula $[\text{Ln}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ with Ln = La or
4
5 Ce (Figure S4). Experimental and simulated powder X-ray diffraction patterns of compounds
6
7 that have general chemical formula $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with Ln = Sm-Er plus Y
8
9 (Figure S5). Continuous Shape Measurements (CSM) performed on the Ce center of
10
11 $[\text{Ce}(\text{mip})_{3/2}(\text{H}_2\text{O})_5 \cdot 2\text{H}_2\text{O}]_\infty$ (Table S2). Continuous Shape Measurements (CSM) performed on
12
13 the Gd center of $[\text{Gd}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$ (Table S3). ATG/DSC of
14
15 $[\text{La}(\text{mip})(\text{Hmip})(\text{H}_2\text{O})_5 \cdot \text{H}_2\text{O}]_\infty$. IR spectra versus temperature of the exhausted gas during
16
17 thermal analysis. IR spectrum recorded at 90°C (Figure S6). Continuous Shape Measurements
18
19 (CSM) performed on the Y centers of $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ (Table S4). Selected
20
21 hydrogen-bond distances in $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ (Table S5). ATG/DSC of
22
23 $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$. IR spectra versus temperature of the exhausted gas during thermal
24
25 analysis. IR spectrum recorded at 90°C (Figure S7). Thermo-dependent powder X-ray
26
27 diffraction patterns of $[\text{Y}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ (Figure S8). Powder X-ray diffraction
28
29 patterns of $[\text{Y}_2(\text{mip})_3]_\infty$: (a) as obtained by dehydration at 200°C ; (b) after exposure to
30
31 ambient air ; (c) after immersion in water (Figure S9). Numerical values for colorimetric
32
33 coordinates and luminance of $\text{Na}_2(\text{mip}) \cdot 7\text{H}_2\text{O}$ and $[\text{Ln}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with Ln = Sm-
34
35 Dy under UV excitation ($\lambda_{\text{exc}} = 312 \text{ nm}$) (Table S6). Solid state UV-vis absorption spectrum
36
37 of $[\text{Gd}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$. Solid state excitation and emission spectra recorded at 77K of
38
39 $[\text{Gd}_2(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ (Figure S10). Powder diffraction diagrams of $[\text{Tb}_{2-2x}\text{Gd}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$
40
41 with $0 \leq x \leq 1$ (Figure S11). Powder diffraction diagrams of $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$
42
43 with $0 \leq x \leq 1$ (Figure S12). Metallic contents measured by EDS
44
45 for $[\text{Tb}_{2-2x}\text{Gd}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $0 < x < 1$ (Table S7). Metallic contents measured
46
47 by EDS for $[\text{Tb}_{2-2x}\text{Eu}_{2x}(\text{mip})_3(\text{H}_2\text{O})_8 \cdot 4\text{H}_2\text{O}]_\infty$ with $0 < x < 1$ (Table S8).

50
51
52 The Supporting Information is available free of charge on the ACS Publications website.
53
54
55
56
57
58
59
60

REFERENCES

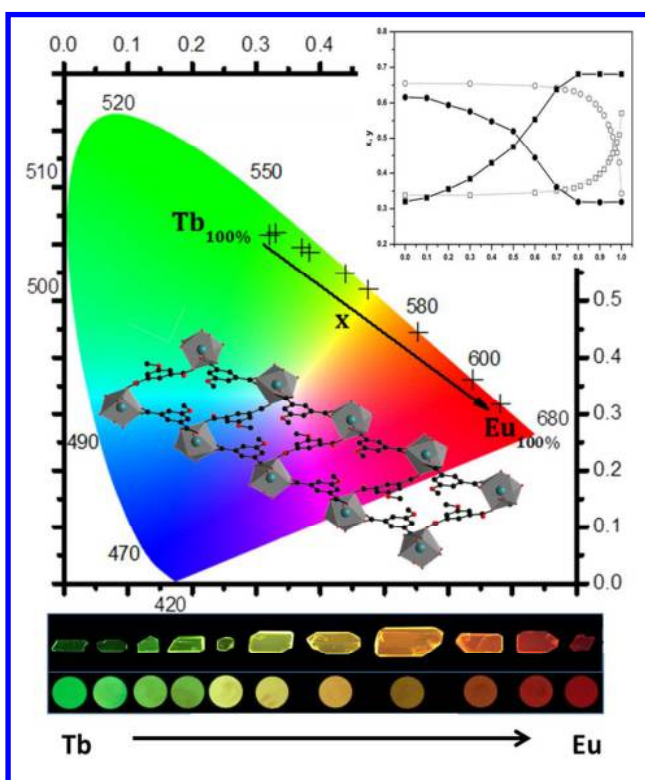
- Guillerm, V.; Weselinski, L.; Belmabkhout, Y.; Cairns, A. J.; D'Elia, V.; Wojtas, L.; Adil, K.; Eddaoudi, M., *Nature Chemistry* **2014**, *6*, 673-680.
- Eddaoudi, M.; Kim, J.; Rosi, N.; Vodak, D.; Wachter, J.; O'Keeffe, M.; Yaghi, O. M., *Science* **2002**, *295*, 469-472.
- Yaghi, O. M.; Li, G.; Li, H., *Nature* **1995**, *378*, 703-706.
- Yaghi, O. M.; Li, H. L., *J. Am. Chem. Soc.* **1995**, *117*, 10401-10402.
- Surlblé, S.; Serre, C.; Millange, F.; Férey, G., *Solid State Sci.* **2006**, *8*, 413-417.
- Millange, F.; Serre, C.; Marrot, J.; Gardant, N.; Pelle, F.; Férey, G., *J. Mater. Chem.* **2004**, *14*, 642-645.
- Serre, C.; Millange, F.; Thouvenot, C.; Gardant, N.; Pelle, F.; Férey, G., *J. Mater. Chem.* **2004**, *14*, 1540-1543.
- Kustaryono, D.; Kerbellec, N.; Calvez, G.; Daiguebonne, C.; Guillou, O., *Cryst. Growth Des.* **2010**, *10*, 775-781.
- Kerbellec, N.; Daiguebonne, C.; Bernot, K.; Guillou, O.; Le Guillou, X., *J. Alloys Compd.* **2008**, *451*, 377-383.
- Calvez, G.; Bernot, K.; Guillou, O.; Daiguebonne, C.; Caneschi, A.; Mahé, N., *Inorg. Chim. Acta* **2008**, *361*, 3997-4003.
- Jeon, J. R.; Clérac, R., *Dalton Trans.* **2012**, *41*, 9569-9586.
- Bernot, K.; Luzon, J.; Caneschi, A.; Gatteschi, D.; Sessoli, R.; Bogani, L.; Vindigni, A.; Rettori, A.; Pini, M. G., *Phys. Rev. B* **2009**, *79*, 134419 1-11.
- Yi, X.; Calvez, G.; Daiguebonne, C.; Guillou, O.; Bernot, K., *Inorg. Chem.* **2015**, *54*, 5213-5219.
- Fan, X.; Freslon, S.; Daiguebonne, C.; Le Polles, L.; Calvez, G.; Bernot, K.; Guillou, O., *Inorg. Chem.* **2015**, *54*, 5534-5546.
- Le Natur, F.; Calvez, G.; Daiguebonne, C.; Guillou, O.; Bernot, K.; Ledoux, J.; Le Polles, L.; Roiland, C., *Inorg. Chem.* **2013**, *52*, 6720-6730.
- Freslon, S.; Luo, Y.; Daiguebonne, C.; Calvez, G.; Bernot, K.; Guillou, O., *Inorg. Chem.* **2016**, *55*, 794-802.
- Lan, A. J.; Li, K. H.; Wu, H. H.; Olson, D. H.; Emge, T. J.; Ki, W.; Hong, M. C.; Li, J., *Angew. Chem. Int. Ed.* **2009**, *48*, 2334-2338.
- Feng, J.; Zhang, H. J., *Chem. Soc. Rev.* **2013**, *42*, 387-410.
- Cui, Y.; Yue, Y.; Qian, G.; Chen, B., *Chem. Rev.* **2012**, 1126-1162.
- Cui, Y.; Xu, H.; Yue, Y.; Guo, Z.; Yu, J.; Chen, Z.; Gao, J.; Yang, Y.; Qian, G.; Chen, B., *J. Am. Chem. Soc.* **2012**, *134*, 3979-3982.
- Serre, C.; Millange, F.; Thouvenot, C.; Gardant, N.; Pelle, F.; Férey, G., *J. Mater. Chem.* **2004**, *14*, 1540-1543.
- Reddy, M. L. P.; Sivakumar, S., *Dalton Trans.* **2013**, *42*, 2663-2678.
- Guillou, O.; Daiguebonne, C.; Calvez, G.; Bernot, K., *Accounts Chem. Res.* **2016**, *49*, 844-856.
- Kerbellec, N.; Kustaryono, D.; Haquin, V.; Etienne, M.; Daiguebonne, C.; Guillou, O., *Inorg. Chem.* **2009**, *48*, 2837-2843.
- Yang, Q. Y.; Pan, M.; Wei, S. C.; Li, K.; Du, B. B.; Su, C. Y., *Inorg. Chem.* **2015**, *54*, 5707-5716.
- Daiguebonne, C.; Gérault, Y.; Guillou, O.; Lecerf, A.; Boubekeur, K.; Batail, P.; Kahn, M.; Kahn, O., *J. Alloys Compd.* **1998**, *275-277*, 50-53.
- Guillou, O.; Daiguebonne, C., Lanthanide ions containing coordination polymers. In *Handbook on the Physics and Chemistry of Rare Earths (vol 34)*, Gschneider, K. A.; Bünzli, J. C. G.; Pecharsky, V. K., Eds. Elsevier: Amsterdam, 2005; Vol. 34, pp 359-404.
- Freslon, S.; Luo, Y.; Calvez, G.; Daiguebonne, C.; Guillou, O.; Bernot, K.; Michel, V.; Fan, X., *Inorg. Chem.* **2014**, *53*, 1217-1228.
- Fan, X.; Daiguebonne, C.; Guillou, O.; Camara, M., *Acta Crystallogr. E* **2014**, *E70*, m181-m182.

- 1
2
3 30. Fan, X.; Freslon, S.; Daiguebonne, C.; Calvez, G.; Le Polles, L.; Bernot, K.; Guillou, O., *J. Mater. Chem. C* **2014**, 5510-5525.
- 4
5 31. Luo, Y.; Calvez, G.; Freslon, S.; Bernot, K.; Daiguebonne, C.; Guillou, O., *Eur. J. Inorg. Chem.* **2011**, 3705-3716.
- 6
7 32. Daiguebonne, C.; Kerbellec, N.; G rault, Y.; Guillou, O., *J. Alloys Compd.* **2008**, 451, 372-376.
- 8
9 33. Mc Cormick, L. J.; Morris, S. A.; Teat, S. J.; Mc Pherson, M. J.; Slawin, A. M. Z.; Morris, R. E., *Dalton Trans.* **2015**, 44, 17686-17695.
- 10
11 34. Zhang, H. J., *Z. Kristallogr. NCS* **2011**, 226, 629-630.
- 12
13 35. Tian, C.; Lin, Z.; Du, S., *Cryst. Growth Des.* **2013**, 13, 3746-3753.
- 14
15 36. Desreux, J. F., In *Lanthanide Probes in Life, Chemical and Earth Sciences*, Choppin, G. R.; B nzli, J. C. G., Eds. Elsevier: Amsterdam, 1989; Vol. Elsevier, p 43.
- 16
17 37. Henisch, H. K., *Crystals in Gels and Liesegang Rings*. Cambridge University Press: Cambridge, 1988.
- 18
19 38. Henisch, H. K.; Rustum, R., *Crystal Growth in Gels*. The Pennsylvania State University Press: 1970; p 1-196.
- 20
21 39. Daiguebonne, C.; Deluzet, A.; Camara, M.; Boubekour, K.; Audebrand, N.; G rault, Y.; Baux, C.; Guillou, O., *Cryst. Growth Des.* **2003**, 3, 1015-1020.
- 22
23 40. Sheldrick, G. M., *Acta Crystallogr. A* **2008**, 64, 112-122.
- 24
25 41. Inc., B. A. *SAINT*, V837A; Bruker: Madison, Wisconsin, USA, 2014.
- 26
27 42. Inc, B. A. *SADABS*, 2014/5; Bruker: Madison, Wisconsin, USA, 2014.
- 28
29 43. Inc, B. A. *APEX3*, Bruker: Madison, Wisconsin, USA, 2015.
- 30
31 44. Altomare, A.; Burla, M. C.; Camalli, M.; Carrozzini, B.; Cascarano, G.; Giacovazzo, C.; Guagliardi, A.; Moliterni, A. G. G.; Polidori, G.; Rizzi, A. C., *J. Appl. Crystallogr.* **1999**, 32, 339-340.
- 32
33 45. Sheldrick, G. M.; Schneider, T. R., *Macromol. Crystallogr. B* **1997**, 319-343.
- 34
35 46. Farrugia, L. J., *J. Appl. Crystallogr.* **2012**, 45, 849-854.
- 36
37 47. Roisnel, T.; Rodriguez-Carjaval, J., *J. Mater. Sci. Forum* **2001**, 378, 118-123.
- 38
39 48. Kraus, W.; Nolze, G., *J. Appl. Crystallogr.* **1996**, 29, 301-303.
- 40
41 49. Haquin, V.; Etienne, M.; Daiguebonne, C.; Freslon, S.; Calvez, G.; Bernot, K.; Le Polles, L.; Ashbrook, S. E.; Mitchell, M. R.; B nzli, J. C. G.; Guillou, O., *Eur. J. Inorg. Chem.* **2013**, 3464-3476.
- 42
43 50. Wyszecski, G., Colorimetry. In *Handbook of Optics*, Driscoll, W. G.; Vaughan, W., Eds. Mac Graw-Hill Book Company: New-York, 1978; pp 1-15.
- 44
45 51. CIE, *International Commission on Illumination - Technical report*. CIE: 1995; Vol. 13-3, p 16.
- 46
47 52. Casanova, D.; Llundell, M.; Alemany, P.; Alvarez, S., *Chem. - Eur. J.* **2005**, 11, 1479-1494.
- 48
49 53. Luo, Y.; Zheng, Y.; Calvez, G.; Freslon, S.; Bernot, K.; Daiguebonne, C.; Roisnel, T.; Guillou, O., *Cryst. Eng. Comm.* **2013**, 15, 706-720.
- 50
51 54. B nzli, J. C. G.; Eliseeva, S. V., Basics of lanthanide photophysics. In *Lanthanide Luminescence*, H nninen, P.; H rm , H., Eds. Springer Berlin Heidelberg: 2010; pp 1-45.
- 52
53 55. Gutierrez, F.; Tedeschi, C.; Maron, L.; Daudey, J.-P.; Poteau, R.; Azema, J.; Tisnes, P.; Picard, C., *Dalton Trans.* **2004**, 1334-1347.
- 54
55 56. Latva, M.; Takalo, H.; Mukkala, V.-M.; Matachescu, C.; Rodriguez-Ubis, J. C.; Kankare, J., *J. Lumin.* **1997**, 75, 149-169.
- 56
57 57. Steemers, F. J.; Verboom, W.; Reinhoudt, D. N.; Van der Tol, E. B.; Verhoeven, J. W., *J. Am. Chem. Soc.* **1995**, 117, 9408-9414.
- 58
59
60

FOR TABLE OF CONTENTS USE ONLY.

High brightness and easy color modulation in lanthanide-based coordination polymers with 5-methoxy-isophthalate as ligand: Toward emission colors additive strategy

Insa Badiane, Stéphane Freslon, Yan Suffren, Carole Daiguebonne, Guillaume Calvez, Kevin Bernot, Magatte Camara and Olivier Guillou.



Some hetero-lanthanide-based coordination polymers with 5-methoxy-isophthalate as ligand present very weak intermetallic energy transfers and therefore exhibit original luminescence properties.