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Trung-Kien Pham, Ana Alexandra Cameirao, Jean-Michel Herri. Experimental flowloop study on methane hydrate formation and agglomeration in high water cut emulsion systems. INTERNATIONAL CONFERENCES ON EARTH SCIENCES AND SUSTAINABLE GEO-RESOURCES DEVELOPMENT -ESASGD 2016-, Hanoi University of Mining and Geology (HUMG), Nov 2016, Hanoi, Vietnam. pp.187 à 198. hal-01472719

HAL Id: hal-01472719

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International Conference on Integrated Petroleum Engineering (IPE)

Experimental flowloop study on methane hydrate formation and agglomeration in high water cut emulsion systems

Trung-Kien PHAM^{a,b}, Ana CAMEIRAO^{a,*}, Jean-Michel HERRI^a

^a Gas Hydrate Dynamics Centre, Ecole Nationale Supérieure des Mines de Saint-Etienne, 158 Cours Fauriel, Saint-Etienne 42023, France

^b Department of Oil Refining and Petrochemistry, Oil and Gas Faculty, Hanoi University of Mining and Geology, 18 Pho Vien, Duc Thang, Bac Tu Liem, Ha Noi, Viet Nam

Abstract

As an oil or a gas field gradually matures, the amount of water produced (water cut) increases, and the hydrate risk also increases. Especially in the offshore systems, operating at low temperature and high pressure, conditions are favourable to the formation of gas hydrate, from the combination of liquid water and gas molecules, under the form of a solid phase. It is a serious issue in the flow assurance; it may cause many troubles, up to plugging.

This work brings new understanding on hydrate kinetics of crystallization, nucleation and growth, agglomeration, breakage and deposition under flowing. We focused on high water cut emulsion systems under flowing and we studied the role of commercial anti-agglomerants or low dosage hydrate inhibitors (AA-LDHIs). These additives can disperse hydrate particles to prevent plugging. More specifically, the topic of this study concerns methane hydrate crystallization, agglomeration and hydrate slurry transportation in a flowloop at high water cut systems, with and without AA-LDHI. We have been used up-to-date measurements and calculations among FBRM (Focused Beam Reflectance Measurement), PVM (Particle Video Microscope), pressure, pressure drop, flowrate/density, hydrate conversion and hydrate volume fraction. In this study, we performed the experiments at very high water volume fractions (80-90-100%) in laminar regime (150L/h) with and without anti-agglomerants. The results showed that using gas-lift system (riser), dramatic events were witnessed with a high rate of methane hydrate formation, followed with a quick agglomeration and leading to plugs, even with additive. But we also observed a dramatic foaming event, even at low AA-LDHI concentration of 0.01%.

We concluded our work with a discussion of the agglomeration mechanisms, based on a quick hydrate formation and agglomeration rates.

Keywords: Emulsion, Gas bubble, Gas hydrate formation, Agglomeration, Anti-agglomerants, Hydrate slurry transport, Suspension, Flow assurance, and Multiphase flow.

1. Introduction

Gas hydrates (or clathrate hydrates) are ice-like crystalline structures composed of molecular water cages stabilized by light hydrocarbon gases (methane, ethane, propane, etc) at low temperature and high pressure conditions (Sloan and Koh, 2008). Gas hydrate formation is a great concern in flow assurance which faces a potential hazard (plug) to oil and natural gas transportation once they form in deep-sea pipelines. There are several strategies to manage this risk including: addition of thermodynamic inhibitors (THIs), kinetic hydrate inhibitors (KHIs) and commercial anti-agglomerants or low dosage hydrate inhibitors (AA-LDHIs). The most traditional method to prevent gas hydrates is thermodynamic inhibition which shifts the hydrate equilibrium conditions to higher pressure and lower temperature. However, the severe conditions (higher pressures, lower temperatures, higher acid gas contents, higher water cuts) in oil and gas transportation make thermodynamic prevention less acceptable because of high amount of THIs used. To reduce the cost, KHIs and/or AA-LDHIs are used at a very small amount compared to THIs. In fact, KHIs may delay the nucleation or growth of hydrate formation while AA-LDHIs allow the hydrate formation but help to maintain hydrate particles dispersed in flowlines.

Low water cut systems have been studied previously. Camargo and Palermo et al., (2002) developed a rheological model for suspension and agglomeration of hydrate particles in an asphaltic crude oil based on the data obtained from laboratory and flow loop scales. In their work, water in oil emulsion was prepared at low

water cut (15-50%) and without anti-agglomerants. Turner et al., (2005) also studied gas hydrate formation in emulsion, both autoclave and flowloop, and water cut of (5%-35%) and without anti-agglomerants. Other publications concerning the understanding of kinetics and mechanisms of methane gas hydrate formation and agglomeration, as well as flow characteristics and rheological properties of natural gas hydrate slurry, in low water fraction, with and without anti-agglomerants are: Colombel et al., 2008; Lachance et al., 2008; Didier et al., 2009; Bo-Hui Shi et al., 2011; Liang Mu et al., 2013; Ke-Le Yan et al., 2013.

In high water cut systems, insights into gas hydrate formation and dissociation in emulsion were developed by Greaves et al., 2008. This contribution mainly focused on the effects of hydrate formation and dissociation on the stability of high water content emulsion. Afterwards, the hydrate formation and agglomeration and rheology of hydrate slurry were investigated in high water fraction with and without anti-agglomerants by the following authors : Gao et al., 2009; Shiva Talatori et al., 2011; Hossein Moradpour et al., 2011; Prasad et al., 2011; Joshi et al., 2013; Jun Chen et al., 2014, 2015.

The study concerning gas hydrate formation and agglomeration in the Archimede flowloop (this apparatus simulates the actual pipeline conditions) at Ecole des Mine de Saint-Etienne was firstly conducted by Fidel-Dufour et al.,(2006). This work showed the rheological properties of hydrate slurry of water in dodecane emulsion at low water cuts (5-30%) and in the presence of anti-agglomerants (0-0.5%). A model was also developed to describe crystallization and rheology based on the flow loop experimental data. Leba et al., (2010) studied the formation and agglomeration of gas hydrate particles in water in oil (Kerdane[®]) emulsion with low water cuts (5-30%) and with the presence of anti-agglomerants (0.5-1%) in flow loop using the Focused Beam Reflectance Measurement (FBRM probe). Recently, Mendes-Melchuna et al., (2016) investigated gas hydrate formation and agglomeration in emulsion with different water cuts (from low to high water cuts), flowrates and different amount of additives (AA-LDHIs) by using FBRM and PVM measurements.

In general, most of the previous studies on gas hydrate formation and agglomeration and rheological characteristics of hydrate slurry merely focused on the low water cut, in small scale (autoclave) and without anti-agglomerants. In contrast, at high water cut, the gas hydrate formation and agglomeration mechanisms in the presence of AA-LDHI in flowlines are poorly understood. The aim of this work is to go deeper in this understanding of hydrate crystallization and agglomeration in pipelines at high water cut, with emulsion systems, and effects of AA-LDHIs.

2. Experimental Methodology

2.1. Flowloop description

The experiments were performed in the Archimede flow loop at Ecole des Mines de Saint Etienne. A schematic of the flow loop is shown in the Fig. 1. The flow loop is constructed with the following conditions: maximum pressure is 100 bar, the temperature is controlled between 0°C and ambient temperature, the flow rate is between 0 and 500 L/h. The flow loop is designed with 30 m of horizontal pipe with 1 cm diameter, a riser and a descending pipe (both of them have a height of 12m and a diameter of 1.5 cm), a gas injection system, and one sapphire window. The total liquid volume is 11.5 L and gas volume is 15L.

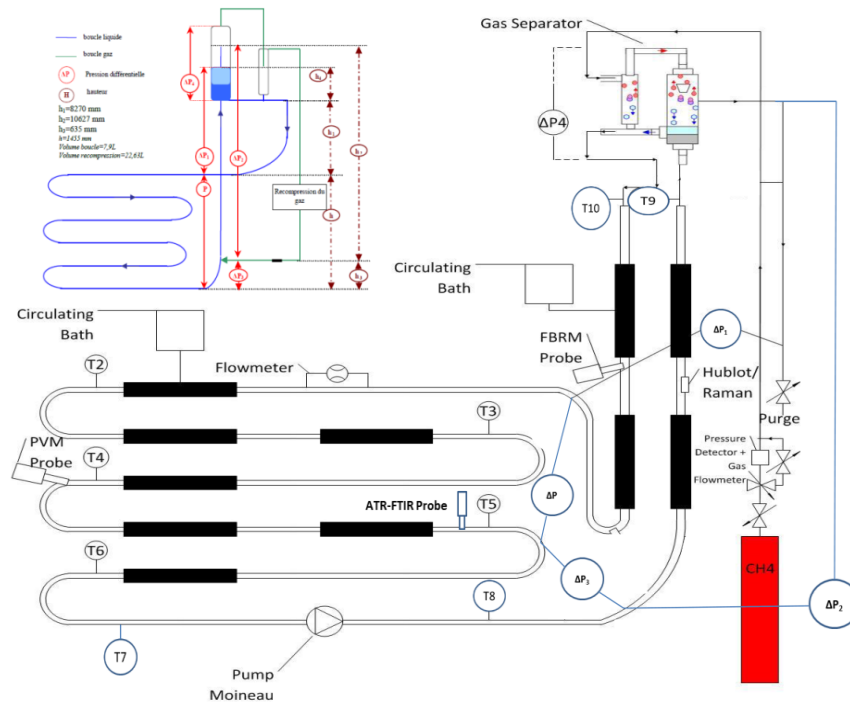


Fig. 1. Schematic of Archimede experimental flow loop: simplified with ballast system (above) and details with pump Moineau (below).

In this experimental apparatus, some probes are equipped to monitor the gas hydrate formation and agglomeration, including: the sensors of temperature (thermocouples) and pressure; the probes of pressure drop (differential pressure transducer) for evaluating the viscosity afterwards, and identifying the plugging troubles; the Focused Beam Reflectance Measurement (FBRM) probe for detecting particle size distribution of emulsion droplets and hydrates (via their chord lengths); the Particle Video Microscope (PVM) probe to observe the emulsion, the gas hydrate formation and agglomeration phenomena; the Attenuated Total Reflection (ATR) probe for determining the concentration of methane and phase inversion of emulsion in fluid flow. Also, the mass flow meter is installed for measuring the flow rate and density. A pump Moineau and/or a gas-lift system are used to circulate the fluids in the flow loop. In this experimental set-up, we used a gas-lift or ballast system which is described in [Fidel-Dufour et al., 2006](#). The cooling systems control the temperature in the flow loop at (4-5°C) by using the water/ethanol circulation jacket attached to the flowloop. The gas compensation system (pressure controller) keeps the total pressure at a constant value (75 bar).

2.2. Materials and experimental procedure

Emulsions are formed by mixing water (ultrapure water, provided from Millipore filter) and oil (Kerdane[®], provided by TOTAL FLUID) with or without a small amount of commercial anti-agglomerants (0.01%). The composition of Kerdane and anti-agglomerants is shown in [Table A.1, Appendix A](#). Because of foam formed in the mixture at 80%WC-0.01%AA-LDHI, 5 ppm of silicone oil (M.40.165.10 provided by Huber) was added only for this experiment. The mixture was charged into the flow loop and circulated at the flow rate of 150L/h to induce emulsification or dispersion. This circulation in the loop is carried out by a gas lift by injecting the compressed air (at 4-5°C and atmospheric pressure) at the bottom of the riser. The flowloop tests are shown in the [Table 1](#).

Table 1. Experimental conditions for flowloop tests.

No.	WC (%)	Temp. (°C)	Press. (bar)	Flowrate (L/h)	AA-LDHI (%wt)
1	100	4-5	75	150	0
2.rp	100	4-5	75	150	0
3	100	4-5	75	150	0.01
4	90	4-5	75	150	0
5	90	4-5	75	150	0.01
6	80	4-5	75	150	0
7	80	4-5	75	150	0.01

Firstly, the mixture in the flow loop system was circulated at a desired flowrate (150L/h) and cooled down to 4-5°C. Then it was pressurized up to 75 bar by the injection of methane (99.99 % purity supplied by AIR LIQUID) and ballast system was started. During the gas solubilization, we measured the total pressure decrease versus time to evaluate the gas transfer rate. Then, we performed a second injection up to 75 bar, and gas was consumed again by solubilization but in smaller quantity. Then, pressure stabilized and we waited for the gas hydrate formation (and agglomeration) which will again consume gas. During the hydrate formation, we monitored the gas consumption rate, the pressure drop in the horizontal section and in the separator, the temperature, the FBRM and the PVM signals. At the end of experiment (when crystallization is finished or/and pipe is plugged), the liquid and gas were depressurized to the ambient pressure and the cooling system was stopped. Finally, the remaining mixture was taken out of flow loop, which was cleaned two times with tap water and finally with pure water for the next experiment. All experimental data were online recorded by LABVIEW, V819 PVM software from METTLER-TOLEDO, LASENTEC[®] products and LASENTEC[®] FBRM data acquisition software. The experimental data were treated and analyzed on a computer.

3. Results and discussion

3.1. Emulsification

[Fig. 2](#) shows the pressure drop and average chord length during emulsification for the experiments without AA-LDHI. It can be concluded from [Fig. 2](#) that the emulsion is quite homogeneous regarding its entire volume. In fact, we observed a constant pressure drop, so the emulsion was not partitioned under sections of different water cut. However, the droplet size of emulsion was fairly stable even if the signal had little fluctuations at the beginning but it was steady at the end of experiment. It can be explained by an un-stationary dispersion/agglomeration process. The size was controlled mechanically from the shear stress only. Chord length distribution appeared in the mixture of 100%WC due to some oil still may stick in flowloop during cleaning or/and the presence of gas bubble in the flow.

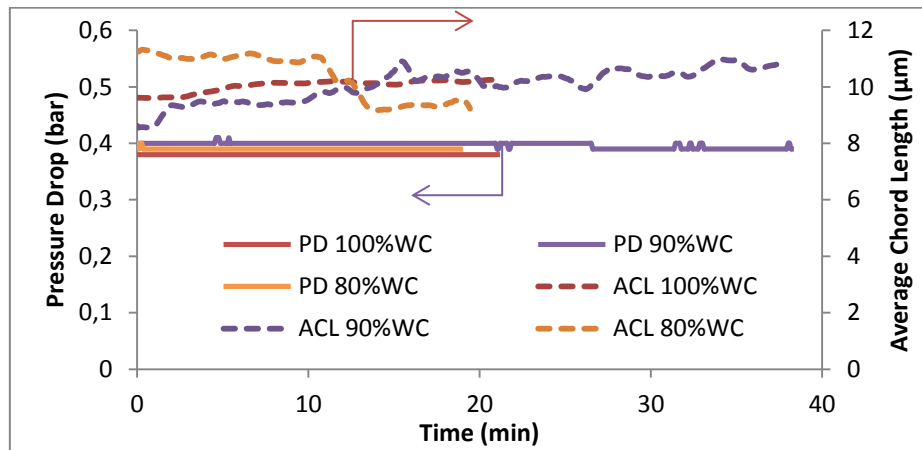


Fig. 2. Pressure drop (solid line) and average chord length (dashed line) during emulsification for the experiments without AA-LDHI.

By addition of AA-LDHI (Fig. 3), although pressure drop measurements still remained stable, the average chord length decreased significantly with the mixture at 100%WC. In contrast, a slight fluctuation of mean droplet size was witnessed for mixture at 80%WC and 90%WC. This can be explained because oil attracts AA-LDHI more than water and the formation of oil in water emulsion may reduce the AA-LDHI aggregation in water. As a result, AA-LDHI dispersed more uniformly in water once oil is added and this helps emulsion more stable (Minwei Sun et al., 2014). From Fidel-Dufour's work, it can be concluded that concentration of 0.01%AA-LDHI in all mixtures is higher than critical micelle concentration (CMC). Furthermore, the concentration of additives is very low, and it is not enough to stabilize the emulsion chemically.

Some foam was also formed in the mixtures with AA-LDHI. It is assumed that the low interfacial tension between oil, water and compressed air leads to foaming.

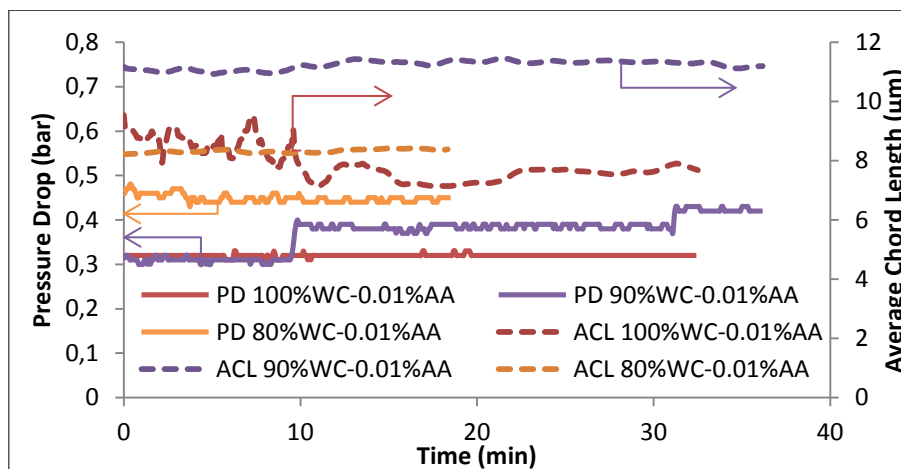
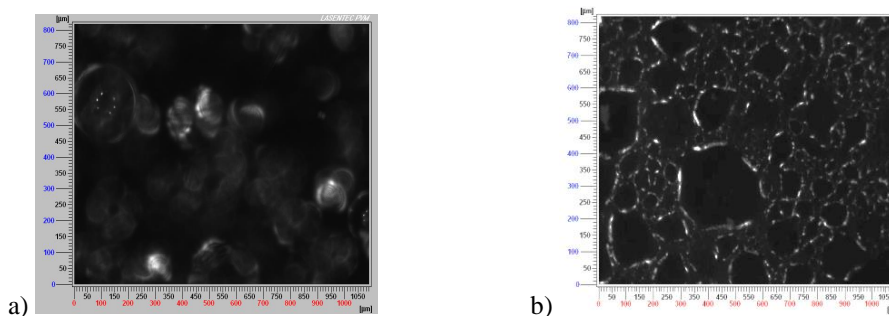


Fig. 3. Pressure drop (solid line) and average chord length (dashed line) during emulsification for the experiments with 0.01% AA-LDHI.

PVM images (Fig. 4) show numerous droplets in the emulsion during emulsification. In this case, in very high water cut emulsion systems, it is supposed that they are oil droplets in a water continuous phase.



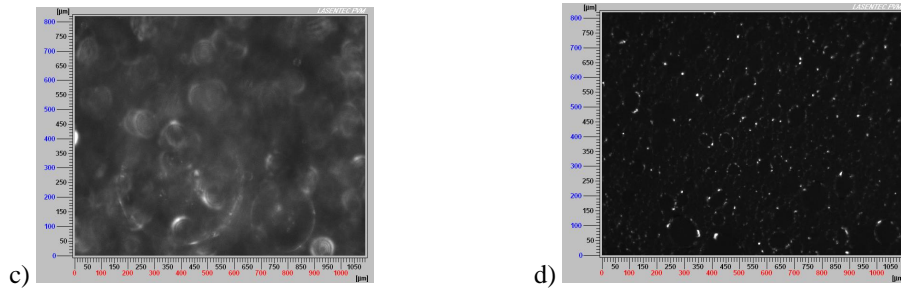


Fig. 4. PVM images of different mixtures during emulsification: (a) 90% WC without AA-LDHI; (b) 80% WC without AA-LDHI; (c) 90% WC with AA-LDHI; (d) 80% WC with AA-LDHI.

In general, addition of AA-LDHI caused foam and instable signals in terms of average chord length and pressure drop but the average chord length is more stable at 80% WC and 90% WC thanks to oil. We assume that it is due to lower interfacial tension between oil, water and compressed air. Mean droplet size of emulsion is approximately 10 μm from the FBRM measurements in all experiments.

3.2. Crystallization

3.2.1. Onset of gas hydrate formation and plugging

The beginning of gas hydrate formation can be identified by an increase of temperature (crystallization is exothermic), an increase of pressure drop in the horizontal pipeline and in the separator, a sharp decrease of static pressure, and also from the pattern of the PVM images (see Fig. 5, 6, 7 and 9).

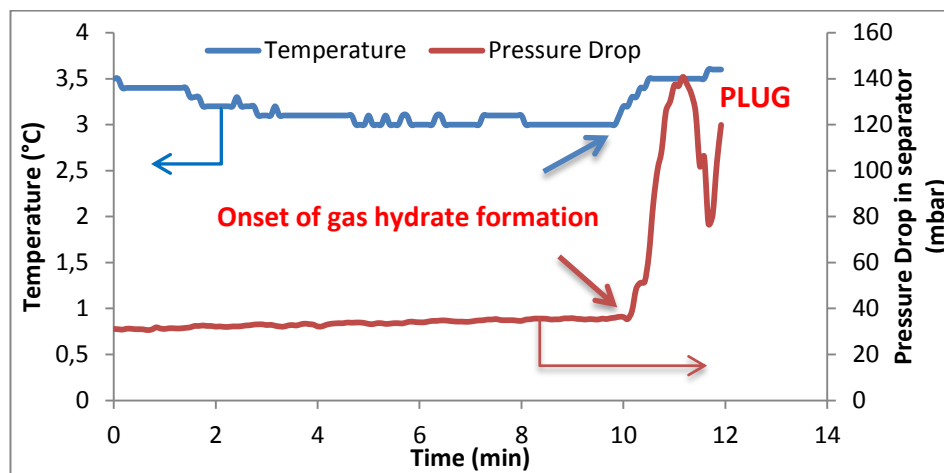


Fig. 5. Typical temperature and pressure drop (in separator) profile during crystallization experiment (90% WC without AA-LDHI).

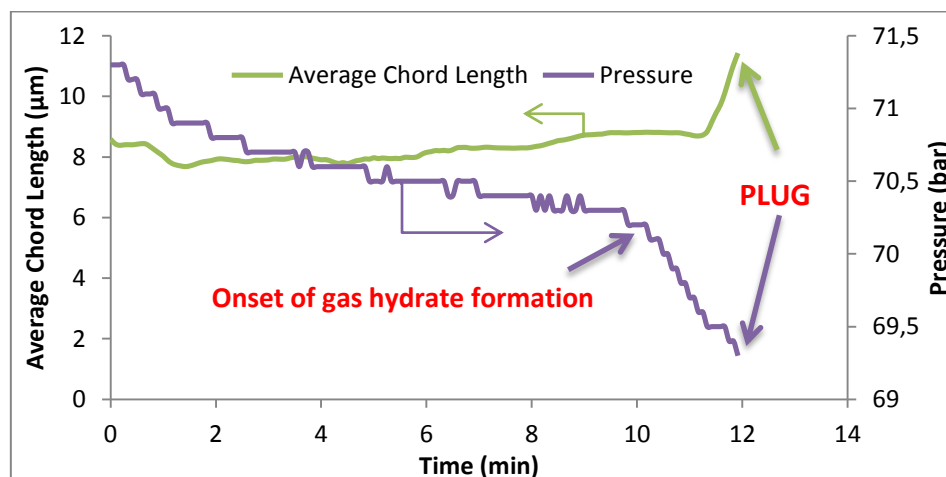


Fig. 6. Typical average chord length and pressure profile during a crystallization experiment (90% WC without AA-LDHI).

Results from Table 2 show that, despite the small hydrate volume in the flowloop, in each test the plug happened very quickly (some tens of seconds) after the beginning of the gas hydrate formation. This time elapsed between the nucleation and the plugging is named “flowing time” in this study. We observed a general trend that the higher the amount of oil the longer the flowing time, both with and without AA-LDHI.

Table 2. Summary of the experimental results for high water cut flowloop tests.

No.	WC (%)	AA-LDHI (%)	Conversion (%)	Hydrate Volume (%)	Nucleation time (min)	Flowing time (min)
1	100	0	1.27	1.59	1.30	1.60
2.rp	100	0	0.99	1.24	10.50	1.60
3	100	0.01	0.50	0.64	12.90	1.10
4	90	0	2.17	2.44	11.50	2.10
5	90	0.01	2.06	2.33	24.60	2.00
6	80	0	1.07	1.07	7.10	3.50
7	80	0.01	3.60	3.60	1.50	8.80

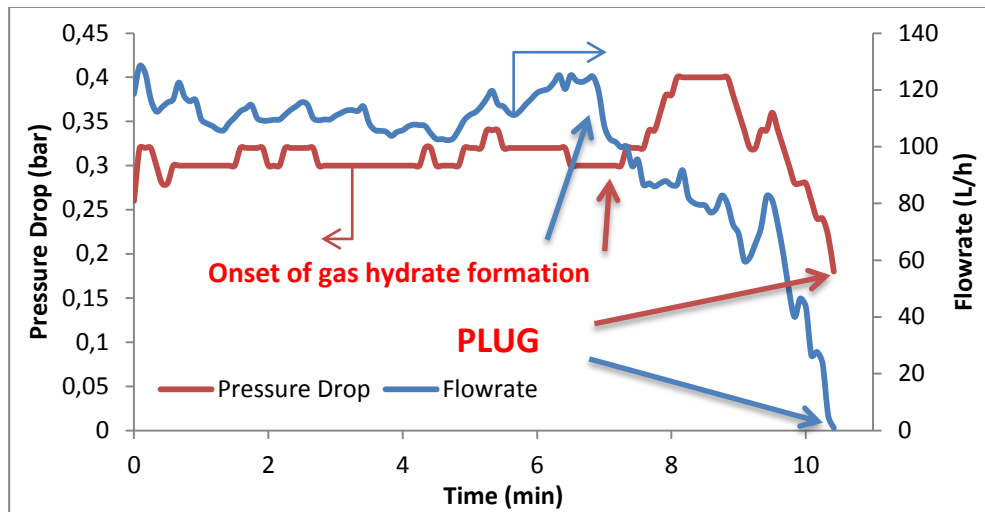


Fig. 7. Typical flowrate and pressure drop (horizontal line) profiles during crystallization experiments (80%WC without AA-LDHI).

Plug time is defined in this work when the flow rate falls to zero (Fig. 7). At the same time, we observed a dis-functioning of gas-lift system which began working abnormally. When hydrates form, they may agglomerate and it is evidenced by an increase of pressure drop, and the increase of the average chord length. The increases of both are observed on Fig. 6 and 7 and are also reported by (Leba et al., 2010; Jun Chen et al., 2014). In all tests, there is a slight increase of pressure drop in horizontal sections (Fig. 7) once hydrate is formed. The hydrate volume (shown in Table 2) is not important enough to yield to a big increase in pressure drop in the horizontal parts of flowloop. Also, we know from previous work (Young hoon Sohn et al., 2015 and Joshi et al., 2013) that agglomeration may not produce sharp increase in pressure drop but high pressure drops were observed only once hydrate deposited on the wall of pipe.

3.2.2. Effect of water cut on gas hydrate formation and agglomeration

Fig. 5 and 6 showed the experimental results of mixture with 90%WC, without AA-LDHI. It is a representative experimental test, showing a quick increase of pressure drop in separator, and also an important decline of static pressure once hydrate formation starts. Fig. 8 shows the hydrate conversion at different high water cuts (80-90-100%) with and without AA-LDHI. On each curve, the end of plots corresponds to the plug formation time. The longer the plot the longer the flowing time. The results illustrated that without AA-LDHI, higher water cut mixtures exhibited lower capability to prevent plugs or higher risk of hydrate blockage. And it is also observed for the mixtures in the presence of AA-LDHI. At the beginning, the rate of gas consumption is almost equal in all experiments. But after approximately one minute, the conversion rate of mixture at 90%WC-0.01%AA-LDHI increased sharply and remained nearly unchanged. On the contrary, hydrate conversion of the mixture at 100%WC-0.01%AA-LDHI and 80%WC was stable and it had slight increment afterwards.

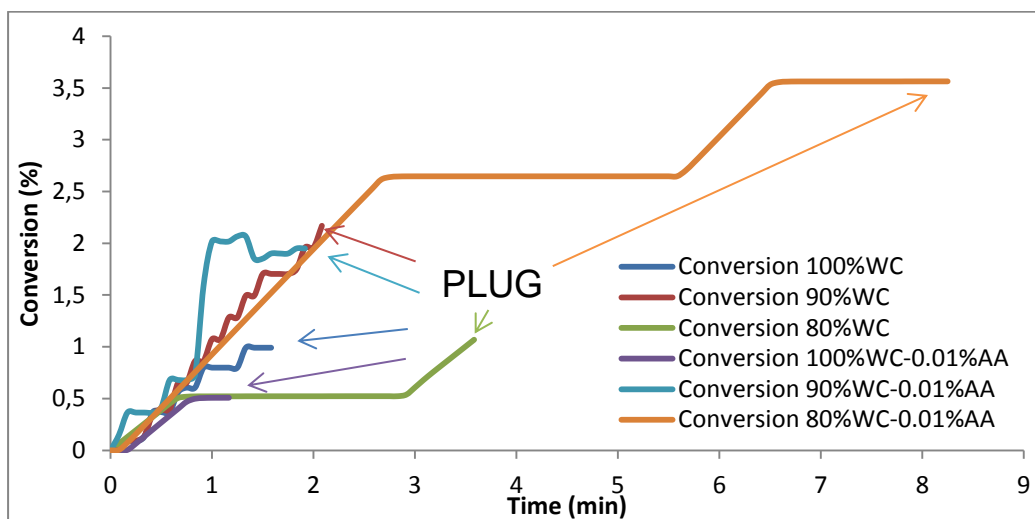


Fig. 8. Conversion of gas hydrate for experiments at different water cut with and without AA-LDHI.

From PVM results (Fig. 9), we observed clearly hydrate particles with hydrate film covering droplets. In some cases, we can observe some agglomeration. In the experiments with AA-LDHI, PVM images had a bad quality and therefore, it was impossible to follow the crystallization processes.

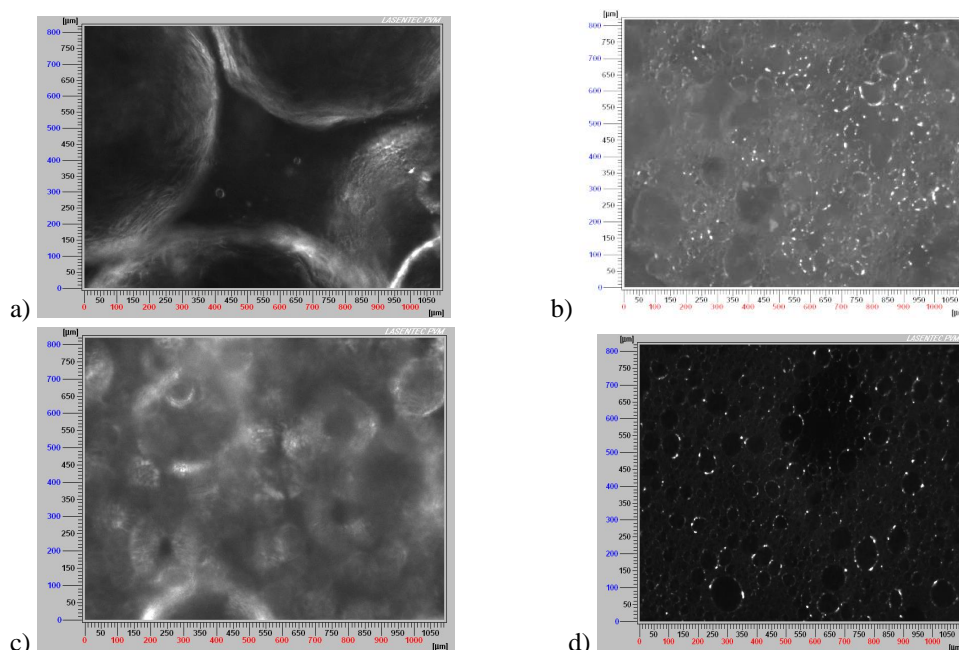


Fig. 9. PVM images of gas hydrate formation for different mixtures: (a) 90%WC without AA-LDHI; (b) 80%WC without AA-LDHI; (c) 90%WC with 0.01% AA-LDHI; (d) 80%WC with 0.01% AA-LDHI.

3.2.3. Effect of additives on the gas hydrate formation and agglomeration

The results of mixture with AA-LDHI compared to mixtures without AA-LDHI, at the same water cut (Table 2), demonstrated that the addition of low quantity of AA-LDHI is not sufficient to prevent plugging. The addition of AA-LDHI also seemed to inhibit hydrate nucleation in the experiments with 100%WC and 90%WC. Interestingly, mixture with 80%WC-0.01%AA showed very short nucleation time compared to others with AA-LDHI. It is reported that the amount of oil enhanced effectively the performance of AA-LDHI (Minwei Sun et al., 2015).

3.3. Role of AA-LDHI and water continuous phase for hydrate formation and agglomeration

All experiments at high water cut systems showed higher rate of hydrate growth, and faster agglomeration and plug events, compared to previous studies at low and high water cut (Fidel-Dufour et al.,

2006; Leba et al., 2010; Mendes-Melchuna et al., 2015). Our experiments differ due to the presence of a gas lift section. It can be hypothesized that gas hydrate is formed in the riser at a higher rate due to the good quality of the gas dispersion. We also suspect an accumulation of the particles in the separator (once a significant surge in the pressure drop of the separator was observed). Moreover, all of the experiments have been performed at high water cut. Due to the gas-lift system, it can promote a different and quicker growth and agglomeration rates when compared to literature at low water cut where the systems are mainly oil dominated. At high water cut, our system becomes water continuous, and it modifies the activity of bubbles, and interaction between continuous phase and the surface of pipelines. Furthermore, the formation of a water-gas bubble bridge between hydrate particles may be more dynamic.

Firstly, it can be attributed to gas hydrate formation at the surface of gas bubbles which has been previously observed by Hideo Tajima et al., 2009 and Y.-T.Luo et al., 2006. Moreover, we have to take into account an un-expected formation of foam. For the experiment at 80%WC with AA-LHDI, we observed more foam in comparison with other experiments with AA-LDHI, and also a very short time of nucleation. Both could be attributed to the high contact surface area between oil, water and gas, because a dispersed and fine population of gas bubbles and foam is formed.

Secondly, hydrate bridges between hydrate particles also contributes as one of principal reasons for a rapid agglomeration. Here, hydrate film formed on the gas and/or foam surfaces perhaps favors links between hydrate particles. In this explanation, we assumed that thanks to a lot of bubbles in the riser, hydrates can readily form on the surface of bubbles, and these bubbles, surrounded by water, may play a vital role as the water-gas bubble bridge, which later are transformed into hydrate-gas bubble easily. This created a higher rate of agglomeration resulting in a blockage.

Thirdly, at high water cut, water is the continuous phase and it also means that water covers almost the pipeline surface. As a result, when hydrate formed, in the presence of water film between the pipe surface and hydrate particles, it modified the interaction between the hydrate particle and the pipe wall. It is a different case when oil is the continuous phase at low water fraction. G.Aspenes et al., 2009 considered that the adhesion force of hydrate particles and pipe surface with water droplets is much higher than hydrate-hydrate adhesion force or hydrate-pipe wall interaction without water layer. It can explain that hydrates deposit on the pipe wall once they formed leading to plug quickly.

In addition, low flow rate and small diameter of pipe may contribute to the gas hydrate deposition. Moreover, from the average chord length results, using gas lift systems, mean droplet size of emulsion is quite small ($\sim 10\mu\text{m}$) compared to that of previous work (Mendes-Menchuna et al., 2016) using protocol with pump Moineau. It is perhaps another reason which causes higher gas hydrate formation due to larger contact surface between oil and water. Because of lower velocity of bubbles at higher height in the riser, hydrate covering bubbles catch up at the top of riser and they agglomerate below the bulk gas-liquid interface in separator which leads to a sudden increase of pressure drop in separator and causes plug (Y.-T.Luo et al., 2006). Beside, bubbles in the riser not only enhance gas transfer to emulsion to reach faster gas saturation in emulsion but also promote gas transfer once hydrate formed causing higher rate of gas hydrate formation.

The role of oil to prevent plugging in the presence of AA-LDHI has been enlightened by (Minwei Sun et al., 2014, 2015). They supposed that surfactants absorbed on the hydrate surface, and they created hydrophobic layers which attract oil to cover on the hydrate surface, thus it led to prevent plugging. The similar explanation to the role of oil on emulsion stabilizing in the presence of AA-LDHI (discussed in section 3.1) may be applied to anti-agglomeration capacity of AA-LDHI with the support of oil. By this way, more AA-LDHI can absorb on hydrate particle surface and prevent agglomeration. In our case, the plug still happened because of low dosage of AA-LDHI. In cases without AA-LDHI, perhaps, oil dispersed into water phase slightly hindered adherence forces between hydrate particles.

3.4. Proposed mechanism

According to (Fidel-Dufour et al., 2006), flow in the riser of our flowloop shows a slug flow pattern ($\rho_L U_L^2 = 15 \text{ kg/m.s}^2$; $\rho_G U_G^2 = 3 \text{ kg/m.s}^2$ with $U_G/(U_G+U_L)=0.64$), which is encountered in small diameter riser ($<0.05 \text{ m}$). From proposed mechanism (Fig. 10), it is considered that once hydrate formed on the gas bubble surface in the riser, it can merge with other hydrate covering bubbles, creating a bigger one. Also, hydrate film covering gas bubbles can be broken into free hydrate particles and the free gas and it can be recovered itself by a new hydrate film, and/or can agglomerate by dint of adherence forces between hydrate particles (Hideo Tajima et al., 2009; Yang et al., 2004; Y.-T.Luo et al., 2006).

From experimental results, we witnessed very high rate of hydrate growth (gas consumption) at different water cut experiments. It can be explained by the dominance of gas bubbles in the riser comparing to oil droplets. The oil droplets may contain some dissolved gases which lead to the formation of a hydrate layer covering their surface but maybe fewer and smaller than hydrate covering bubbles. The main reason explaining the quick formation of plugs immediately after the nucleation may be the high rate of gas hydrate formation around the gas bubble, cracks from hydrate shell (this helps promoting mass transfer) and their rapid agglomeration due to water-gas bubble bridge between hydrate particles and hydrate deposition as described above (section 3.3).

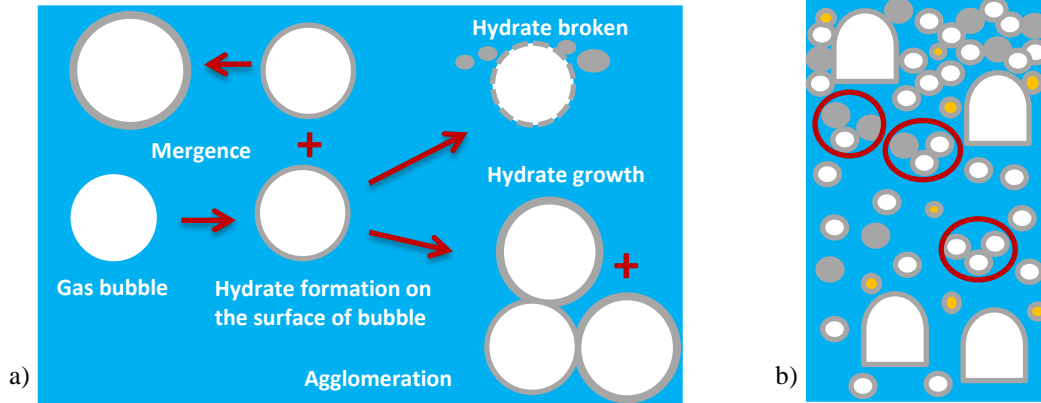


Fig. 10. (a) Hydrate formation on the surface of bubble, growth, agglomeration (by water-gas bubble bridge) and breaking in the riser; (b) slug flow regime and typical hydrate agglomeration in the riser.

4. Conclusions

In high water systems using gas-lift, we are confronted to a high risk of hydrate blockade because of high rate of hydrate formation (high dispersion of gas bubbles) and high risk of agglomeration, leading to a quick plug formation, even with AA-LDHI. Main reason for this phenomenon is attributed to the hydrate film formation around the surface of gas bubbles in the riser. Very low dosage of AA-LDHI is not sufficient to prevent their agglomeration, and plugging occurs rapidly. A mechanism based on a quick agglomeration rate was proposed to explain the rapid plugging. In all experiments with and without AA-LDHI, oil partially made contribution to maintain flow after gas hydrate formation.

Nomenclature & abbreviations			
n_{erw}	number of mol of water converted to hydrate [mol]	U_L	superficial velocity of liquid [m/s]
n_{CH4}	number of mol of methane consumption to hydrate [mol]	ρ_G	mass density of the gas phase [kg/m ³]
n_{water}	initial number of mol of water in mixture [mol]	U_G	superficial velocity of gas [m/s]
η	conversion [%]	rp	repeated test
n_{cell}	number of cells of hydrate structure I [-]	AA-LDHI	anti-agglomerant - low dosage hydrate inhibitor
V_{cell}	volume of one cell [m ³ /cell]	KHI	kinetic hydrate inhibitor
V_{Hyd}	volume of hydrate [m ³]	THI	thermodynamic inhibitor
V_{total}	total volume of mixture [m ³]	PD	pressure drop [bar]
$\%vol_{Hyd}$	hydrate volume fraction [%]	ACL	average chord length [μ m]
WC	water cut or water fraction [%]	CMC	critical micelle concentration [%]
M_w	molecular weight of water [g/mol]	FBRM	Focused Beam Reflectance Measurement
m_w	initial mass of water in mixture [g]	PVM	Particle Video Microscope
ρ_L	mass density of the flowing phase [kg/m ³]	ATR	Attenuated Total Reflection

Acknowledgements

The authors are grateful to TOTAL E&P for the financial support in the Archimede-3 project and their approval for publishing this paper. We also would like to acknowledge the contribution on technical support and reviewing this manuscript from the members of TOTAL and the Gas Hydrate Dynamics Centre team at Ecole des Mine de Saint Etienne.

Appendix A

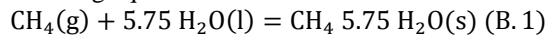
The composition of Kerdane and anti-agglomerants is shown in **Table A.1**

Table A.1. Chemical composition and properties of Kerdane and AA-LDHI.

Substances/Properties	Kerdane®	Anti-agglomerants (AA-LDHI)
Chemical Composition	Hydrocarbons, C11-C14, n-alkanes, isoalkanes, cyclical, < 2% aromatics	Natural oil based surfactant, methyl alcohol, ethanol, 2-(2-propoxyethoxy), ethylene glycol, diethylene glycol, < 2% ethylene glycol monopropyl ether
Appearance	Clear light yellow	Clear light amber
Boiling Point/Interval	180-260°C	No data available
Flash Point	>64°C	15.6°C
Density	790-825 kg/m ³	863 kg/m ³
Kinematic Viscosity	2.1 mm ² /s at 25°C	No data available
Solubility	Not applicable	Water soluble

Appendix B

To calculate the hydrate conversion, we used these following equations:



$$n_{crw} = n_{\text{CH}_4} * 5.75 \quad (\text{B. 2})$$

$$\eta = \frac{n_{crw}}{n_{iwater}} 100\% \quad (\text{B. 3})$$

$$n_{iwater} = \frac{m_w}{M_w} \quad (\text{B. 4})$$

Appendix C

To calculate the hydrate volume, these following equations are applied:

$$n_{crw}(\text{molecules}) = n_{crw}(\text{moles}) * N(\text{Avogadro}) \quad (\text{C. 1})$$

$$n_{cell} = \frac{n_{crw}(\text{molecules})}{46} \quad (\text{C. 2})$$

$$V_{Hyd} = n_{cell} * V_{cell} \quad (\text{C. 3})$$

$$\%vol_{Hyd} = \frac{V_{Hyd}}{V_{total}} 100\% \quad (\text{C. 4})$$

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