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Thermal, Spectroscopic and Chemical Characterization of Biofield Energy Treated Anisole

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Abstract

The objective of the present study was to evaluate the impact of biofield energy treatment on the thermal, spectroscopic, and chemical properties of anisole by various analytical methods such as gas chromatography-mass spectrometry (GC-MS), high performance liquid chromatography (HPLC), differential scanning calorimetry (DSC), Fourier transform infrared (FT-IR) spectroscopy, and ultraviolet-visible (UV-Vis) spectroscopy. The anisole sample was divided into two parts, control and treated. The control part was remained same while the other part was treated with Mr. Trivedi's unique biofield energy treatment. Mass spectra showed the molecular ion peak with five fragmented peaks in control and all treated samples. The isotopic abundance ratio of ²H/¹H, and ¹³C/¹²C [(PM+1)/PM] in treated sample was increased by 154.47% (T1) as compared to the control [where, PM- primary molecule, (PM+1)-isotopic molecule either for ¹³C or ²H]. The HPLC chromatogram showed retention time of treated anisole was slightly decreased as compared to the control. Moreover, the heat change in the sharp endothermic transition of treated anisole was increased by 389.07% in DSC thermogram as compared to the control. Further, C-C aromatic stretching frequency of treated sample was shifted by 2 cm⁻¹ to low energy region in FT-IR spectroscopy. The UV-Vis spectra of control sample showed characteristic absorption peaks at 325 nm, which was red shifted and appeared as shoulder in the treated sample. These results suggested that biofield treatment has significantly altered the physical and spectroscopic properties of anisole, which could make them stable solvent for organic synthesis and as a suitable reaction intermediate in industrial applications.

Keywords: Biofield energy treatment; Anisole; Gas chromatography-Mass spectrometry; High performance liquid chromatography

Abbreviations: GC-MS: Gas chromatography-Mass spectrometry; PM: Primary molecule; PM+1: Isotopic molecule either for ¹³C/¹²C or ²H/¹H

Introduction

Anisole, phenyl ether (Ph-O-Me, Figure 1) usually used as a starting material for various pharmaceutical/flavonoid products and as solvent in organic synthesis and physical studies [1]. It is used for the synthesis of raw material for drugs such as cyclofenil, in treating Raynaud's phenomenon in people with scleroderma [2], and tramadol hydrochloride, that is used for the relief of moderate or severe pain [3]. 4-hydroxyanisole and t-butylhydroxyanisole are the most popular derivatives of anisole and used extensively as depigmenting agents and antioxidant respectively [4,5]. The anisole was used selectively in various steps as a solvent in the synthesis of cefoxitin [6] and latamoxef an antibiotic administered intravenously has a broad spectrum of activity against Gram-positive and Gram-negative bacteria [7]. Anisole is a weakly polar aprotic solvent with lower electric permittivity ($\epsilon_s=4.33$), which allows its usability in cyclic voltammetric studies as an alternative to tetrahydrofuran and dichloromethane [8]. The use of hazardous and toxic solvents in synthesis and characterization is regarded as a very important point for the safety of lab-workers and pollution. Nowadays, a relatively green solvent, anisole, have been successfully applied to process organic/polymer solar cells [9]. Hence, the stability of anisole is important to perform organic reactions at moderate temperature, and its applicability in pharmaceutical products. The chemical and physical property could be altered by Mr. Trivedi's unique biofield energy treatment which is well known to modify the physical, and structural characteristics of living and non-living substances [10,11]. The electrical current generates through internal physiological processes like blood flow, brain activity, and heart function *etc.* exists inside the human

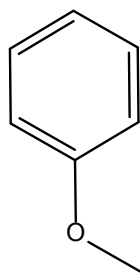
body in the form of vibratory energy particles like ions, protons, and electrons and they generate magnetic field in the human body [12]. The energy fields that purportedly surround the human body is called the biofield. Currently, researchers have been exploring the potential benefits of integrative energy medicine in a variety of situations to promote overall health and wellness of individuals. The energy medicine is regarded as one of the complementary and alternative medicine (CAM) and defined under the subcategory of energy therapies by National Center for Complementary and Alternative Medicine (NCCAM) [13]. The practitioner of energy medicine can harness the energy from the environment/universe and can transmit into any object around the globe. The object(s) receive the biofield energy and responded in a useful way. This process is called as biofield energy treatment. Mr. Trivedi's unique biofield treatment is also called as The Trivedi Effect[®]. The Trivedi Effect[®] has been well studied in various research fields like microbiology [10,14] biotechnology [15,16] and agricultural research [17]. Based on the outstanding results achieved by biofield treatment on microbiology and biotechnology, an attempt was made to evaluate the impact of biofield treatment on various properties of anisole.

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Anisole

Figure 1: Structure of anisole.

Experimental

Materials and methods

Anisole was procured from Genuine Chemical, India. The samples were characterized using, gas chromatography-mass spectrometry (GC-MS), high performance liquid chromatography (HPLC), differential scanning calorimetry (DSC), Fourier transform infrared (FT-IR) spectroscopy, and ultraviolet-visible (UV-Vis) spectroscopy.

Biofield treatment modalities

Anisole was taken in this experiment for biofield treatment. The compound was divided into two parts named as control and treated. No treatment was given to the control set. The second set of anisole was handed over to Mr. Trivedi for biofield energy treatment under laboratory conditions. Mr. Trivedi provided the biofield treatment through his energy transmission process to second sets of samples without touching. After treatment, the treated samples were stored at standard conditions for GC-MS analysis as per the standard protocol. The experimental results in treated groups were analyzed and compared with the untreated (control) set.

GC-MS

The GC-MS analysis was done on Perkin Elmer/auto system XL built with Turbo mass, USA. The detection limit of the detector is upto 1 picogram. For GC-MS analysis, the treated sample was further divided into three parts as T1, T2, and T3. The GC-MS spectrum was plotted as the % abundance vs. mass to charge ratio (m/z). The isotopic abundance ratio of $^{13}\text{C}/^{12}\text{C}$ or $^2\text{H}/^1\text{H}$, (PM+1)/PM, and $^{18}\text{O}/^{16}\text{O}$, (PM+2)/PM was expressed by its deviation in the treated sample as compared to the control. The percentage changes in isotopic ratio (PM+1)/PM and (PM+2)/PM was calculated on a percentage scale from the following formula:

$$\text{Percentage changes in isotopic ratio (PM + 1 / PM)} = \frac{R_{\text{Treated}} - R_{\text{Control}}}{R_{\text{Control}}} \times 100$$

Where, R_{Treated} and R_{Control} are the ratios of intensity at (PM+1) to PM in mass spectra of treated and control samples respectively.

HPLC

The HPLC analysis was performed on a Knauer High Performance Liquid Chromatograph (Berlin, Germany), equipped with Smartline Pump 1000 and a UV 2600 detector. Chromatographic separation was performed on a C_{18} column (Eurospher 100) with a dimension of 250×4 mm and $5 \mu\text{m}$ particle size. The mobile phase used was methanol with a flow rate of 1 mL/min at 25°C . The solutions of standard and the sample were prepared in methanol for both GC-MS and HPLC studies.

The method development for this assay of anisole based on its chemical properties. Anisole is polar molecule and therefore, a polar solvent methanol was used as the diluent.

DSC

The DSC was done with Perkin Elmer/Pyris-1, USA, a heating rate of $10^\circ\text{C}/\text{min}$ and nitrogen flow of 5 mL/min was used. The change in latent heat (ΔH) of control and treated ethanol was recorded from their respective DSC curves. The percent change in boiling point and latent heat of vaporization was computed using following equations: Percent change was calculated using following equations:

$$\% \text{ change} = \frac{[T_{\text{Treated}} - T_{\text{Control}}]}{T_{\text{Control}}} \times 100$$

Where, T_{Control} and T_{Treated} are the peak point of degradation control and treated samples, respectively.

FT-IR spectroscopy

For FT-IR spectra, Shimadzu's Fourier transform infrared spectrometer (Japan) was used in the frequency region of $500\text{-}4000 \text{ cm}^{-1}$. The samples were prepared by crushing anisole crystals with spectroscopic grade KBr into fine powder and then pressed into pellets.

UV-Vis spectroscopy

UV-Vis spectra of control and treated samples of anisole were acquired from Shimadzu UV-2400 PC series spectrophotometer within the wavelength region of $200\text{-}400 \text{ nm}$. Quartz cell with 1 cm and a slit width of 2.0 nm were used for analysis.

Results and Discussion

GC-MS analysis

The GC-MS spectra of control and treated samples of anisole are presented in Figure 2. For GC-MS study of the treated sample was divided into three parts T1, T2, and T3. Mass spectra showed the PM peak at $m/z=108$ in control and all the treated anisole samples (T1, T2, and T3) same intensities. The intensity ratio of (PM+1)/PM is presented in Table 1 and percent change in isotopic abundance ratio was calculated and shown in Figure 3. Five major peaks at $m/z=108, 93, 78, 65, 51,$ and 39 were observed in both control and treated samples of anisole due to the following ions respectively: $\text{C}_7\text{H}_8\text{O}^+, \text{C}_6\text{H}_5\text{O}^+, \text{C}_6\text{H}_6^+, \text{C}_5\text{H}_5^+, \text{C}_4\text{H}_3^+,$ and C_3H_3^+ ions. Peaks at $m/z=93, 78, 65, 51,$ and 39 were observed due to the fragmentation of anisole to phenol, benzene, cyclopentadiene, buten-3-yne and propyne ions. All peaks were same for both treated and control samples and well matched with reported literature [18]. The isotopic abundance ratio of (PM+1)/PM of anisole sample was increased in all the treated samples T1=154.47%, T2=21.71%, and T3=34.44%. However, the isotopic abundance ratio of (PM+2)/PM was not seen in the mass spectrum. The increased isotopic abundance ratio of (PM+1)/PM in the treated anisole may increase the effective mass (μ) and binding energy in this molecules with heavier isotopes. This may alter the property of the chemical bond and eventually property of the molecule. The transformation may be happened in nuclear level due to the biofield energy treatment. It is expected that some of ^1H and ^{12}C may be interconverted to ^2H , and ^{13}C atoms inside the molecule, respectively. The bond strength also plays an important role in kinetic effects due to the greater strength of the isotopic bonds such as $^2\text{H}\text{-}^{12}\text{C}$ bond relative to a $^1\text{H}\text{-}^{12}\text{C}$ bond [19].

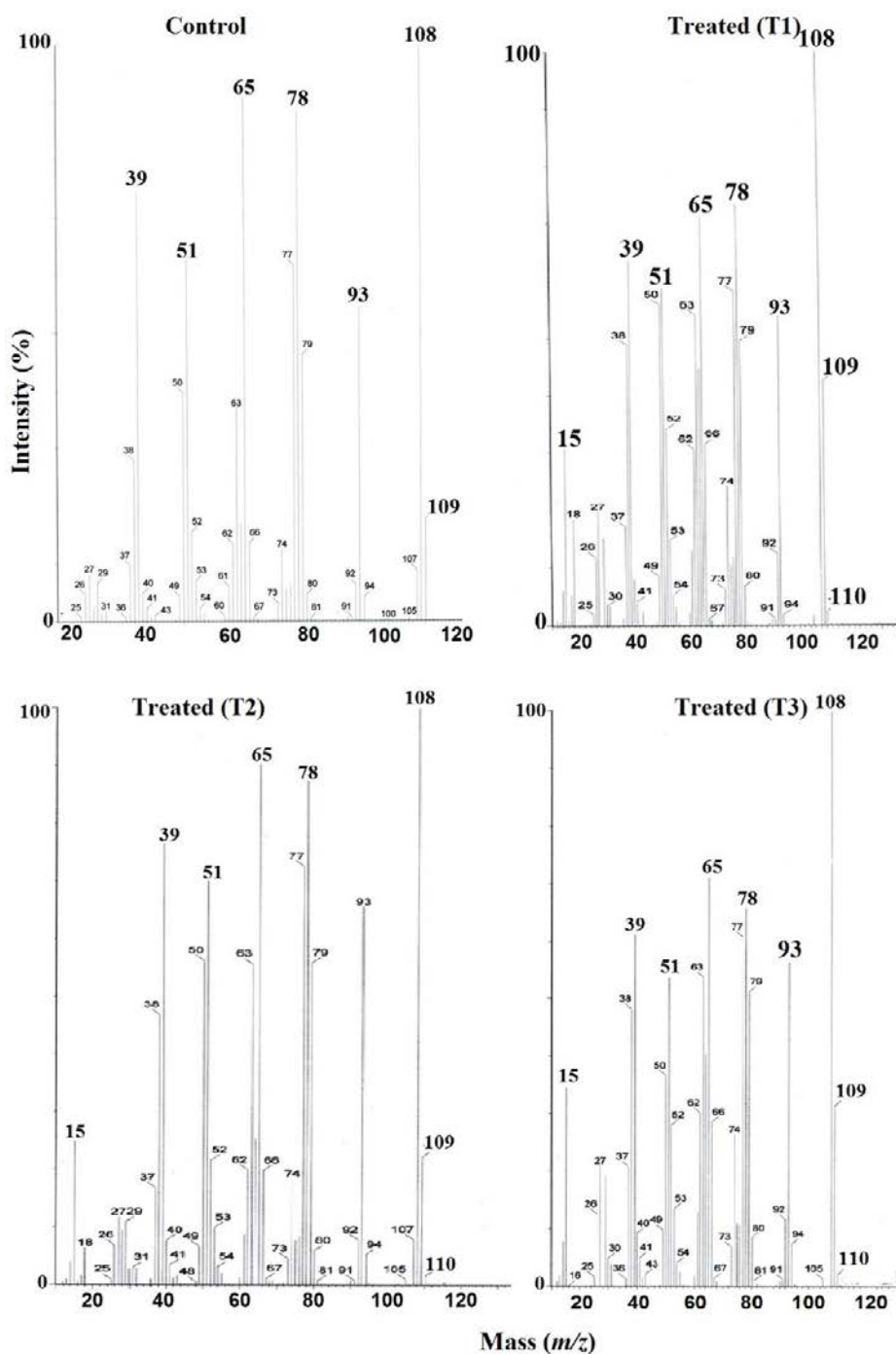


Figure 2: GC-MS spectra of control and treated anisole samples.

Peak Intensity	Control	Treated		
		T1	T2	T3
$m/z=(PM)$	100	100.00	82.62	98.19
$m/z=(PM+1)$	23.11	23.76	13.00	20.25
$m/z=(PM+2)$	1.05	2.05	1.00	1.80

Table 1: GC-MS isotopic abundance analysis result of anisole.

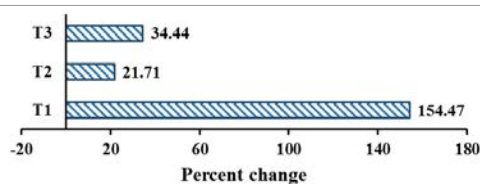


Figure 3: Percent change in isotopic abundance (PM+1)/PM and (PM+2)/PM of anisole under biofield treatment as compared to control.

HPLC analysis

HPLC chromatogram is the detector response as a function of time is shown in Figure 4 (control and treated anisole). The compounds have been separated into two peaks in both control and treated samples. The retention time (T_R) of control sample was 2.58 min and 2.96 min with relative intensity 22.51% and 77.49%, respectively and the T_R of treated anisole showed at 2.41 min and 2.71 min with relative intensity 29.85% and 70.10%, respectively. These decreased in T_R may be attributed to the increased polarity of anisole after biofield energy treatment.

DSC analysis

DSC was used for thermal analysis of control and treated anisole samples. The change in heat (ΔH) of control and treated samples is shown in Figure 5. One strong endothermic transition was observed at 114.09°C in control anisole sample. However, it was observed at 112.27°C after biofield energy treatment. The heat absorbed in this process was found to be 550.35 J/g in treated anisole sample. The increase in ΔH in the treated sample was significantly increased by 389.08% as compared to the control, (Table 2). The increase in ΔH after biofield treatment could be due to alteration of intermolecular interaction in anisole that might increase the thermal stability of the treated samples.

FT-IR spectroscopic analysis

The FT-IR spectrum of control and biofield energy treated anisole is shown in Figure 6. The vibrational peaks at 3003 cm^{-1} and 2955 cm^{-1} were assigned to aromatic C-H stretching of the phenyl ring for both control and treated samples. The vibrational peak at 2904 cm^{-1} was observed due to the methyl (aliphatic) C-H stretching for both the control and treated samples. The C-C stretching frequency (in-ring, aromatic) was observed at 1603 cm^{-1} and 1496 cm^{-1} for control sample whereas it was shifted to 1601 cm^{-1} and 1494 cm^{-1} (the lower energy region) after biofield energy treatment. The characteristic C-O stretching for ethers was seen at 1302 cm^{-1} and 1077 cm^{-1} for aliphatic (R-O) and aromatic (Ar-O) parts respectively in both control and treated samples. The shifting of aromatic stretching frequency to the lower wavenumber may be attributed to the increased number of higher isotopes (PM+1) after biofield energy treatment. The vibration energy of a bond (in wavenumbers) in a molecule is represented by the following equation

$$\bar{\nu} = \frac{1}{2\pi C} \sqrt{\frac{K}{\mu}} \quad (1)$$

Where, K is a constant that varies from one bond to another, C is the speed of light, and $\mu = m_1 m_2 / (m_1 + m_2)$, is the reduced mass or effective mass (m_i is the mass of atom i). According to the equation (1), $\bar{\nu}$ is inversely proportional to the reduced mass [20]. It exhibited that reduced mass is higher in case of heavier isotope as compared to lighter one. This showed that anisole with heavier isotope after biofield energy treatment has low vibration energy as compared to lighter in control. Due to the increased isotopic abundance ratio of (PM+1) in the treated anisole, the effective mass (μ) was increased and subsequently wavenumber was decreased. The FT-IR results after biofield energy treatment as compared to the control have shown changes in chemical properties of the molecule, which may lead to the changes in polarity as well as thermal property of anisole molecule.

UV-Vis spectroscopic analysis

The UV-vis spectra of anisole in methanol are shown in Figure 7. The UV spectrum of control sample showed characteristic absorption

peaks at $\lambda_{\text{max}} = 325 \text{ nm}$, however after biofield treatment the peak was red shifted and exhibited a shoulder. Both the control and treated samples showed characteristic strong absorption in the range of 200-275 nm, which was saturated at that particular concentration. Reported

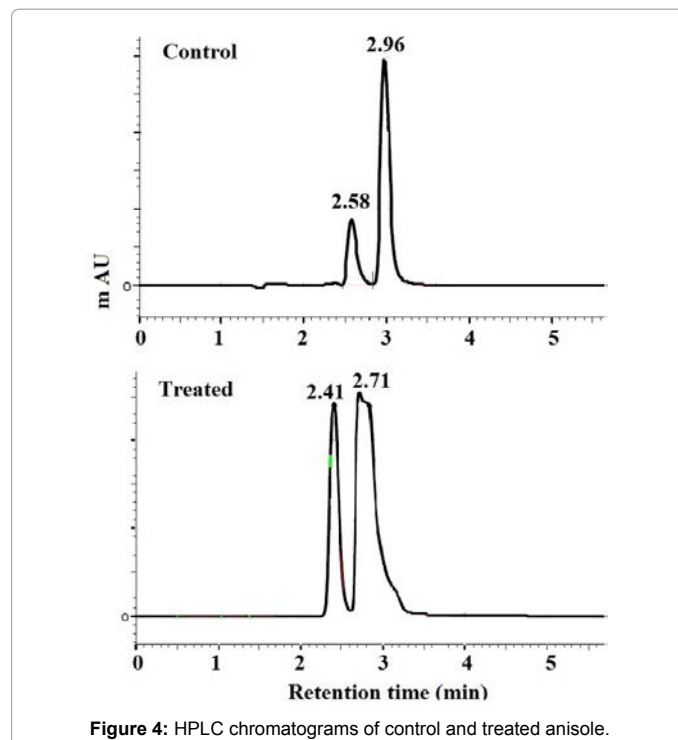


Figure 4: HPLC chromatograms of control and treated anisole.

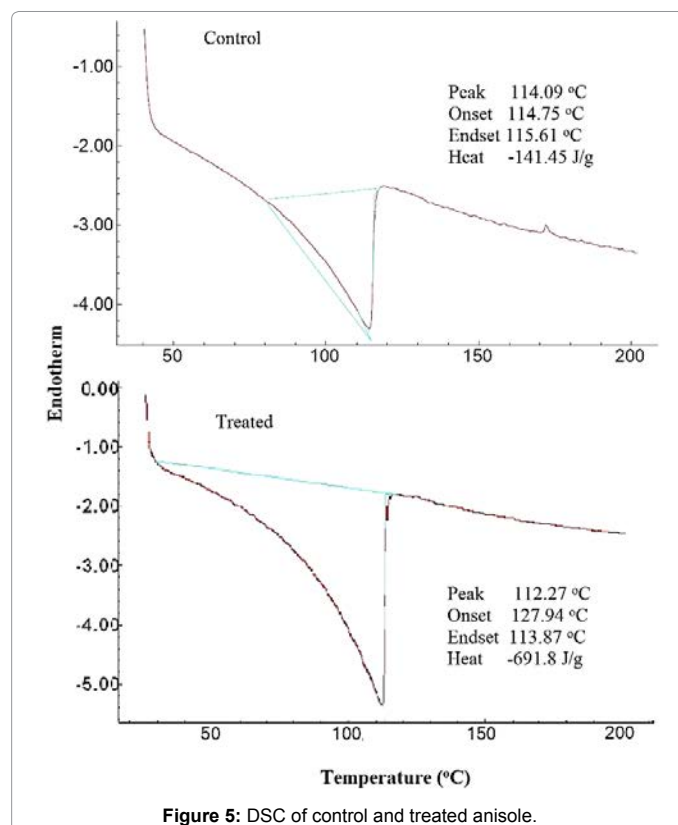


Figure 5: DSC of control and treated anisole.

	Peak position (°C)	Heat change J/g
Control	114.09	-141.45
Treated	112.27	-691.80
Percent Change	-1.59	389.07

Table 2: DSC analysis of anisole.

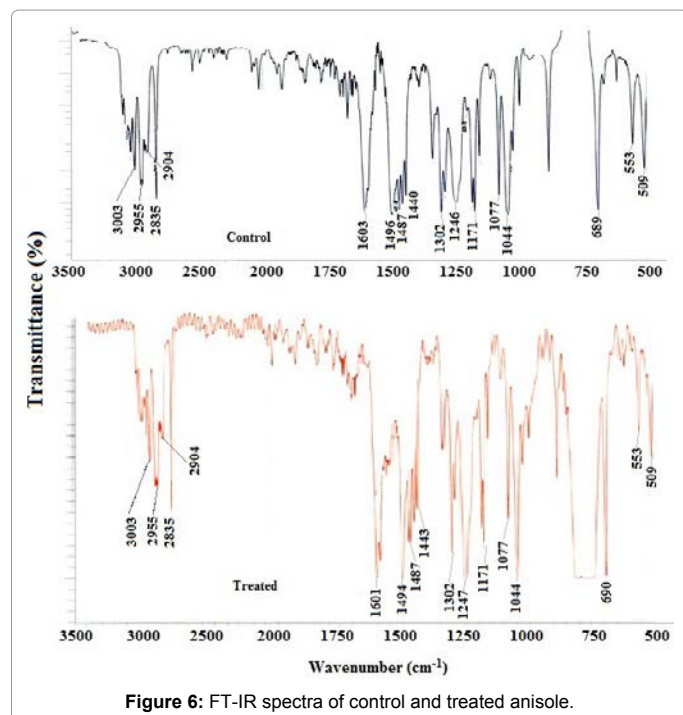


Figure 6: FT-IR spectra of control and treated anisole.

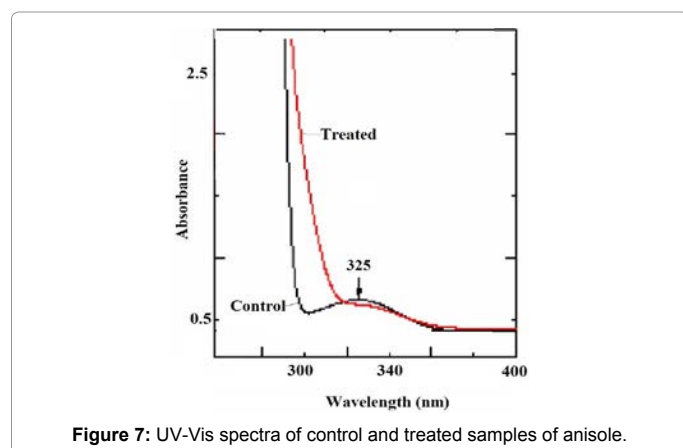


Figure 7: UV-Vis spectra of control and treated samples of anisole.

literature have suggested that the first strong band in the region of 300-350 nm that corresponds to π - π^* transitions and the weak n - π^* band in the region of 320-370 nm [21]. The n - π^* bands show a distinct shift in the treated sample, it was observed with the varied polarity of the medium/solvent and reported in the literature [22]. Hence, it is hypothesized that biofield treatment might induce some polarity to the system that makes shift of the absorption positions. This observation is also supported by the lower T_r of the treated sample in HPLC chromatogram.

Conclusions

Anisole was studied to see the influence of biofield energy treatment

and isotopic abundance ratio of treated anisole was compared with the control sample. After performing various analytical studies on biofield treated anisole, it was found that anisole has shown a strong response to biofield energy treatment. GC-MS data showed the isotopic abundance ratio of (PM+1)/PM was increased by 154.47% in the treated sample as compared to the control. Biofield energy may also alter the polarity of the molecule, which eventually alters the absorption position in UV-Vis and the retention time peak in HPLC chromatogram. Higher isotopic ratio (PM+1)/PM in treated samples correlated with the higher ΔH in DSC thermogram and shifting of aromatic C-C stretching frequency in FT-IR spectra to the lower wavenumber region as compared to the untreated sample. It is assumed that biofield energy treatment may enhance the stability of the organic small molecule of industrial importance, which ultimately affects the shelf-life and efficacy of the final product.

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