



HAL
open science

Dual cation- and anion-based redox process in lithium titanium oxysulfide thin film cathodes

Vincent Dubois, Brigitte Pecquenard, Samantha Soulé, Hervé Martinez,
Frédéric Le Cras

► **To cite this version:**

Vincent Dubois, Brigitte Pecquenard, Samantha Soulé, Hervé Martinez, Frédéric Le Cras. Dual cation- and anion-based redox process in lithium titanium oxysulfide thin film cathodes. ACS Applied Materials & Interfaces, 2017, 9 (3), pp.2275-2284. 10.1021/acsami.6b11987. hal-01450521

HAL Id: hal-01450521

<https://hal.science/hal-01450521v1>

Submitted on 12 Apr 2021

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Article

Dual cation- and anion-based redox process in lithium titanium oxysulfide thin film cathodes for all-solid-state lithium-ion batteries

Vincent Dubois, Brigitte Pecquenard, Samantha Soulé, Herve Martinez, and Frédéric Le Cras

ACS Appl. Mater. Interfaces, **Just Accepted Manuscript** • DOI: 10.1021/acsami.6b11987 • Publication Date (Web): 21 Dec 2016Downloaded from <http://pubs.acs.org> on January 5, 2017**Just Accepted**

“Just Accepted” manuscripts have been peer-reviewed and accepted for publication. They are posted online prior to technical editing, formatting for publication and author proofing. The American Chemical Society provides “Just Accepted” as a free service to the research community to expedite the dissemination of scientific material as soon as possible after acceptance. “Just Accepted” manuscripts appear in full in PDF format accompanied by an HTML abstract. “Just Accepted” manuscripts have been fully peer reviewed, but should not be considered the official version of record. They are accessible to all readers and citable by the Digital Object Identifier (DOI®). “Just Accepted” is an optional service offered to authors. Therefore, the “Just Accepted” Web site may not include all articles that will be published in the journal. After a manuscript is technically edited and formatted, it will be removed from the “Just Accepted” Web site and published as an ASAP article. Note that technical editing may introduce minor changes to the manuscript text and/or graphics which could affect content, and all legal disclaimers and ethical guidelines that apply to the journal pertain. ACS cannot be held responsible for errors or consequences arising from the use of information contained in these “Just Accepted” manuscripts.

Dual Cation- and Anion-Based Redox Process in Lithium Titanium Oxysulfide Thin Film Cathodes for All-Solid-State Lithium-Ion Batteries

Vincent Dubois^{a,b}, Brigitte Pecquenard^{*a}, Samantha Soulé^c, Hervé Martinez^c, and Frédéric Le Cras^{*d,e}

^aCNRS, Université de Bordeaux, ICMCB UPR 9048 and Bordeaux INP, F-33600 Pessac, France.

^bST Microelectronics, 16 rue Pierre et Marie Curie, F-37071 Tours, France

^cIPREM-ECP CNRS UMR 5234, Université de Pau, 2 avenue Pierre Angot, F-64053 Pau, France

^dCEA LETI, Minatec Campus, 17 rue des Martyrs, F-38054 Grenoble, France.

^eUniversité Grenoble Alpes, F-38000 Grenoble, France

*Corresponding authors: frederic.lecras@cea.fr, brigitte.pecquenard@icmcb.cnrs.fr

Abstract: A dual redox process involving $\text{Ti}^{3+}/\text{Ti}^{4+}$ cation species and $\text{S}^{2-}/(\text{S}_2)^{2-}$ anion species is highlighted in oxygenated lithium titanium sulfide thin film electrodes during lithium (de)insertion, leading to a high specific capacity. These cathodes for all-solid-state lithium-ion microbatteries are synthesized by sputtering of LiTiS_2 targets prepared by different means. The limited oxygenation of the films that is induced during the sputtering process favors the occurrence of the $\text{S}^{2-}/(\text{S}_2)^{2-}$ redox process at the expense of the $\text{Ti}^{3+}/\text{Ti}^{4+}$ one during the battery operation, and influences its voltage profile. Finally, a perfect reversibility of both electrochemical processes is observed, whatever the initial film composition. All-solid-state lithium microbatteries using these amorphous lithiated titanium disulfide thin films and operated between 1.5-3.0 V/ Li^+/Li deliver a greater capacity (210-270 mAh g^{-1}) than LiCoO_2 , with a perfect capacity retention (-0.0015 % cycle⁻¹).

Keywords: LiTiS_2 , thin films, all-solid-state lithium batteries, titanium sulfide, disulfide pairs

1. Introduction

The development of self-powered nomadic, wearable or implantable miniaturized electronics devices, boosted by the growing interest for the Internet of things (IoT), is made possible by the progress in the development of highly integrated and low power consumption components (MEMS, CMOS, transmitters), but is finally often faced with the lack of suitable energy storage solutions. Indeed, these new applications demand particular features for the embedded power source in terms of thickness, footprint, flexibility, safety, capacity and voltage, which cannot be fulfilled in particular by conventional batteries. Optimized solutions for part of these needs should be found among thin film batteries (primary or secondary) based on different chemistries (Zn/Polyaniline, Zn/MnO₂, Li/LiCoO₂,...) or obtained by different manufacturing processes (ink-jet printing, screen printing, tape casting, sputtering,...).^{1,2} All-solid-state thin film batteries, manufactured by sequential deposition of thin films using vacuum deposition techniques such as sputtering, are particularly attractive when a reduced footprint, a very low thickness, a thermal and mechanical resistance and/or a high integration with microelectronic components is required. In this context, all-solid-state thin film batteries, also called microbatteries, are becoming emerging power sources in the microelectronics industry nearly 30 years after the achievement of the first proofs of concept.³⁻⁹ Contrary to bulk-type lithium batteries, metallic lithium is usually used as the negative electrode in all-solid-state lithium microbatteries, since the solid glassy electrolyte (typically LiPON) is able to prevent the formation of lithium dendrites during the charge.¹⁰ Nevertheless, the deposition of lithium thin films at the industrial scale is actually a real issue. In addition, the presence of metallic lithium in the cells limits their thermal resistance, making them for instance incompatible with the reflow soldering process used to connect the electronic component on the printed circuit board. Therefore, lithium-ion or 'lithium-free' cell designs,^{11,12} and consequently lithiated materials at the positive electrode, are preferably envisaged for microbatteries with improved performance. In this context, the aim of this study was to develop a new lithiated positive electrode material with possible

1
2
3 advantages over the widespread LiCoO₂ cathode material.¹³⁻¹⁷ The latter comprise in particular a
4
5 lower operating voltage (~ 2V) more adapted to energy harvesting components and low power
6
7 consumption CMOS circuits, the absence of a compulsory post-deposition annealing and a slightly
8
9 higher volumetric capacity ($\geq 67 \mu\text{Ah}\cdot\text{cm}^{-2} \mu\text{m}^{-1}$). Previous work achieved on thin film cathodes
10
11 synthesized by sputtering of a titanium sulfide (TiS₂) target has already highlighted that insertion/de-
12
13 insertion of lithium in the resulting amorphous materials is highly reversible in the voltage range [1-3
14
15 V/Li⁺/Li] and lead to a high volumetric capacity.¹⁸ This work is therefore devoted to the synthesis of
16
17 lithiated titanium disulfide thin films electrodes (theoretical capacity: 225 mAh g⁻¹, 68 $\mu\text{Ah cm}^{-2} \mu\text{m}^{-1}$
18
19 ¹) prepared by sputtering of a LiTiS₂ target, then to their physico-chemical characterization, and to
20
21 the investigation of their electrochemical behavior in all-solid-state lithium microbatteries.
22
23
24
25
26
27

28 **2. Experimental Section**

29
30 *Preparation of sputtering targets:* LiTiS₂ sputtering targets were synthesized by three different ways.
31
32 The first route (referred as R1) consisted in the chemical intercalation of lithium into TiS₂ using a n-
33
34 butyl lithium reagent.¹⁹ This preparation was performed at room temperature in an argon-filled
35
36 glovebox containing less than 1 ppm H₂O. A solution of 1.6 M n-butyl lithium in anhydrous hexane
37
38 was added dropwise into a vessel containing dry hexane and dispersed TiS₂ powder (Aldrich,
39
40 99.9 %), so that a 50% excess of n butyl-lithium at a concentration of 0.2 M was reached at the end.
41
42 Then, the suspension remained stirred during one week. The resulting fine powder was then filtered,
43
44 rinsed using anhydrous hexane and dried under vacuum at 100°C. About 8 g of LiTiS₂ were prepared
45
46 at a time. Solid state reaction of a mixture of TiS₂ (Aldrich, 99.9%, 1 μm), Ti powder (Alfa Aesar,
47
48 99.9%, 150 mesh) and Li₂S (Alfa Aesar, 99.9%) with a molar ratio 3:1:2 was used as an alternative
49
50 route (referred as R2) and allowed the preparation of larger batches of LiTiS₂.²⁰ Precursors were
51
52 ground together using a planetary ball-mill and an argon-filled agate jar, prior being annealed under
53
54 argon. All material transfers were carried out without any air exposure. Both sputtering targets (50
55
56
57
58
59
60

1
2
3 mm diameter, ~ 3 mm thick) were prepared from about 15 g of LiTiS₂ powder prepared either by
4
5 chemical reaction or solid state reaction. Targets were obtained by cold pressing at 60 MPa in a
6
7 cylindrical stainless steel die inside the glove box. The one from LiTiS₂ precursor prepared by solid
8
9 state reaction was further annealed at 425°C under argon during 24 h. A third route (referred as R3)
10
11 consisted in applying the spark plasma sintering (SPS) technique directly to a mixture of TiS₂, Ti and
12
13 Li₂S precursors in a graphite mold (all the precursors were introduced in the mold inside a glove-
14
15 box).
16
17

18
19
20
21 *Preparation of thin film electrodes:* Lithium titanium sulfide thin films were prepared from these
22
23 three different types of targets by radio-frequency magnetron sputtering, using an equipment
24
25 (Plassys) connected to an argon-filled glove box. Before deposition, the sputtering chamber was
26
27 evacuated to a background pressure of less than 6.10^{-7} Pa, then filled with high purity argon
28
29 (99.9999%). Prior to each deposition, sputtering of a titanium target was performed on the shield to
30
31 trap oxygen traces present in the chamber, then a pre-sputtering of the LiTiS₂ target was
32
33 systematically carried out for 1 hour. All depositions were performed with no intentional
34
35 heating/cooling of the substrate.
36
37

38
39 In order to compare the physicochemical and then the electrochemical properties of lithium titanium
40
41 sulfide thin films prepared from the different LiTiS₂ targets, standard deposition conditions were
42
43 defined as follows: pure argon as the discharge gas, a power of 1.27 W cm^{-2} , an argon flow of 20
44
45 $\text{cm}^3 \cdot \text{min}^{-1}$, a total pressure of 0.5 Pa and a target-substrate distance of 100 mm. These conditions
46
47 were chosen to limit the incorporation of oxygen in the film and to facilitate the deposition of Li
48
49 species. The latter are based on theoretical considerations, previous work dealing with the synthesis
50
51 of TiO_xS_y thin films from TiS₂ targets,²¹ and of a preliminary study assessing the influence of the
52
53 total pressure on the electrochemical behavior of the films synthesized from LiTiS₂ prepared by
54
55 chemical reaction (see supplementary information Figure S1).
56
57
58
59
60

1
2
3
4
5 *Manufacture of all-solid-state microbatteries:* All-solid-state microbatteries having an active area of
6
7 7 or 25 mm² were manufactured by sequential deposition of thin films either on passivated 4'' silicon
8
9 wafer (Siltronix) or on 4'' glass wafer (Schott) substrates. The typical active stack was W (250 nm)/
10
11 Li-Ti-S (0.5-1.0 μm)/LiPON (1.4 μm)/Li (3 μm), using the deposition conditions for W, LiPON and
12
13 Li reported previously.²² Photolithography was used to pattern the first level (W current collector);
14
15 subsequent materials were deposited using shadow masking. At the end, wafers were capped with a
16
17 polymer and a glass plate to protect the microbatteries from air and moisture.
18
19

20
21
22 *Physicochemical characterization of materials:* The chemical composition of Li-Ti-S thin films was
23
24 determined by combining Rutherford backscattering spectroscopy (RBS) and Inductive coupled
25
26 plasma (ICP) measurements (Table 1). RBS analysis was achieved at a backscattering angle of 150°
27
28 using an incident beam of 4He⁺ ions with an energy of 2 MeV. The spectra were analyzed with the
29
30 SIMNRA software.²³ For that purpose, 100 nm thick films were deposited onto vitreous carbon
31
32 substrates. As lithium is not detected by RBS, the Li/Ti lithium content in the film was obtained
33
34 using an inductively coupled plasma-optical emission spectrometer ICP-OES (Varian 720ES) using
35
36 the following emission lines: λ = 670.78 nm for Li, at λ = 368.52 nm for Ti. Thin film samples
37
38 deposited on the glass substrate were dissolved in 10 ml of HCl, then a small amount of the resulting
39
40 solution was introduced in the spectrometer. The homogeneity of the film composition over the
41
42 thickness was checked by Auger electron spectroscopy (VG Microlab 310 F) for samples deposited
43
44 onto silicon wafers.
45
46
47

48
49 XPS measurements were carried out with a Thermo Scientific K-Alpha X-ray photoelectron
50
51 spectrometer using a focused monochromatic Al Kα radiation (hν = 1486.6 eV) in order to determine
52
53 the nature and the local environment of mainly Ti and S species. Besides, air-tight glass containers
54
55 were systematically used to transfer thin film samples from the sputtering chamber to the glove-box
56
57
58
59
60

1
2
3 filled with purified argon and connected directly to the XPS spectrometer, in order to avoid the
4
5 contamination of the surface. For the Ag 3d_{5/2} line the full width at half-maximum (FWHM) was
6
7 0.50 eV under the recording conditions. The X-ray spot size was 400 μm. Peaks were recorded with
8
9 constant pass energy of 20 eV. The pressure in the analysis chamber was less than 2.10⁻⁸ Pa. Short
10
11 acquisition time spectra were recorded at the beginning and at the end of each experiment to check
12
13 that the samples did not suffer from degradation during the measurements. Peak assignments were
14
15 made with respect to reference compounds analyzed in the same conditions. The binding energy
16
17 scale was calibrated from the hydrocarbon contamination using the C 1s peak at 285.0 eV. Core
18
19 peaks were analyzed using a nonlinear Shirley-type background.²⁴ Peak positions and areas were
20
21 optimized by a weighted least-squares fitting method using 70% Gaussian and 30% Lorentzian line
22
23 shapes. Quantification was performed on the basis of Scofield's relative sensitivity factors.²⁵
24
25 Analyses of lithiated/de-lithiated thin films were performed on electrode materials deposited on an
26
27 aluminum foil and charged at 3.2 V/Li⁺/Li or discharged at 1.5 V/Li⁺/Li in button cells using a liquid
28
29 electrolyte (EC:DMC 1:1, 1M LiPF₆, Novolyte). Mechanical erosion under ultrahigh vacuum was
30
31 achieved before carrying out the analyses in order to remove the SEI (Solid Electrolyte Interphase)
32
33 covering the surface of the thin-film electrode. This method is used instead of ionic bombardment
34
35 because the latter is likely to induce a reduction of elements present in the electrode material.
36
37
38 X-ray diffraction was carried out with a Bruker D8 Advance diffractometer in Bragg-Bretano
39
40 geometry using an air-tight cell filled with argon.
41
42
43
44
45
46

47 *Electrochemical characterization of materials:* Electrochemical measurements were performed using
48
49 a VMP3 galvanostat-potentiostat (Bio-Logic) with channels equipped for electrochemical impedance
50
51 spectroscopy (EIS) analysis. Preliminary electrochemical characterization of lithiated titanium
52
53 sulfide thin films was carried out with a liquid electrolyte using 2032 button cells. These cells were
54
55 assembled in a glove box under argon atmosphere using the thin film deposited on an aluminum disc
56
57
58
59
60

1
2
3 as the positive electrode, a polypropylene non-woven (Viledon FS2123, Freudenberg) and a
4
5 microporous polypropylene film (Celgard 2400) as separators, a lithium foil (battery grade,
6
7 Chemetall Foote) as the negative electrode and a liquid electrolyte 1M LiPF₆ in EC:DMC 1:1
8
9 (Novolyte). Aluminum substrates were weighted in the glove box before and after thin film
10
11 deposition using a microbalance (0.1 μg accuracy).

12
13
14 All-solid-state microbatteries were electrically connected to the apparatus using shielded miniature
15
16 probe heads (Cascade Microtech) and tungsten tips. Considering the microbattery design,
17
18 measurements were performed using a two electrode set-up. Both the apparatus and microbatteries
19
20 were kept at a controlled temperature of 25°C during measurements. The mass of electrode materials
21
22 used in these microbatteries (~ 50 μg) were estimated *a priori* from the area of the pattern and the
23
24 positive electrode thickness. Nevertheless, due to shadowing effects, the actual film thickness is
25
26 actually lower at the edges of the pattern, hence the actual mass was systematically under-estimated.
27
28 In order to determine more precisely the latter, a calibration was performed *a posteriori* on the basis
29
30 of the specific capacity obtained for the first charge performed in button cells. The evolution of the
31
32 open circuit voltage and the impedance of the microbatteries as a function of their state of charge (or
33
34 the lithium content in the lithium titanium thin film electrode) were measured using the
35
36 Galvanostatic Intermittent Titration Technique (GITT). The latter consisted in iterating sequences
37
38 comprising a constant current pulse (2 μA cm⁻², ~5% capacity increment), followed by a 4 hours rest
39
40 period and then by EIS measurements (voltage amplitude ΔE = ± 10 mV, frequency range 100 MHz
41
42 -10 mHz).
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

3. Results and discussion

3.1. LiTiS₂ compounds and related sputtering targets

Sputtering targets of LiTiS₂ composition were synthesized via three distinct routes. The first one required the preparation of a LiTiS₂ powder by chemical lithiation of TiS₂ using n-butyl lithium (route R1). The characterization of the resulting product by X-Ray diffraction (Figure S2a) confirmed the synthesis of a pure 1T-Li_xTiS₂ phase, having the following refined hexagonal cell parameters $a = b = 3.4501(2) \text{ \AA}$, $c = 6.1928(3) \text{ \AA}$ (P-3m1 space group). Therefore, the latter values confirm the full lithiation of the TiS₂ precursor and the achievement of the LiTiS₂ composition in agreement of the ICP analysis.²⁶ Besides, it was found that this LiTiS₂ powder retained the same plate-like shape of the pristine TiS₂ particles. As a consequence, the cold compaction carried on this anisotropic powder for shaping the target leads to a maximum compactness of only 61% (1.94 g cm⁻³). Then, although LiTiS₂ is known to be thermally stable at least up to 800°C either under its 1T or 3R structural type,^{20,27} additional conventional sintering aiming at increasing the compactness of the target was not conceivable due to the high chemical reactivity of the material.

An alternative route for the preparation of a LiTiS₂ target by solid state reaction of TiS₂, Li₂S and Ti powdered precursors (route R2) was also investigated in the prospect for the preparation of larger targets. The course of the reaction was followed by *ex situ* XRD on pelletized samples annealed under argon during 20 hours between 400°C and 600°C. The corresponding diffraction patterns (Figure S2b) show the complete transformation of TiS₂ into a 1T-LiTiS₂ phase at 400°C, despite the presence of small amounts of unreacted Ti and Li₂S. Additional peaks, which correspond to the 3R-Li_xTiS₂ phase,²⁰ progressively grow when the temperature is further increased. At 600°C, the transformation of 1T-LiTiS₂ into 3R-LiTiS₂ is almost complete and is accompanied by the formation of a lithium titanium oxide impurity (likely Li₄Ti₅O₁₂). Finally, a thermal treatment of the target at 425°C during 40 hours was chosen as a trade-off to get the most complete reaction without the

1
2
3 formation of oxide impurities. The resulting density of the R2 target was estimated to $\sim 1.60 \text{ g.cm}^{-3}$,
4
5 which corresponds to a compactness of 52%. Finally, we had the opportunity to prepare a third
6
7 target by spark plasma sintering using the same precursors. As expected, the rapid thermal annealing
8
9 at 600°C of the TiS_2 , Ti, Li_2S blend then led to the preparation of a denser target (2.8 g cm^{-3} , i.e. \sim
10
11 91% compactness).
12
13
14
15

16 **3.2. Physico-chemical properties of lithium titanium sulfide thin films**

17
18 Lithium titanium sulfide thin film deposition using the same standard sputtering conditions (see
19
20 experimental section) for each target leads to conventional deposition rates between 4.8 and 5.4
21
22 nm.min^{-1} , the highest one being achieved consistently with the denser target. SEM images (Figure
23
24 S3) show that the resulting thin films exhibit a dense morphology and a very smooth surface that are
25
26 well-adapted to the sequential process of manufacture of all-solid-state thin film batteries. All the
27
28 films exhibit an amorphous character (XRD) and similar densities of $2.8 \pm 0.1 \text{ g cm}^{-3}$. Chemical
29
30 analyses by RBS and ICP-OES lead to the following film compositions (Table 1): $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$
31
32 (target R1), $\text{Li}_{0.6}\text{TiO}_{0.6}\text{S}_{1.8}$ (target R2) and $\text{Li}_{1.0}\text{TiO}_{0.3}\text{S}_{2.0}$ (target R3). An example of RBS spectrum is
33
34 given on Figure S4. The chemical homogeneity of these films was confirmed by Auger spectroscopy
35
36 depth profiling. Therefore, as for thin films deposited by sputtering from a TiS_2 target,²¹ it was found
37
38 again that, despite all precautions taken to avoid oxygen contamination, a small oxygen amount is
39
40 actually incorporated in the films. Titanium is indeed highly reactive even under high vacuum, a
41
42 residual fraction of 0.7% of O_2 in the discharge gas being typically sufficient to fully oxidize Ti into
43
44 TiO_2 .²⁸ Nevertheless, since the amount of oxygen in the R3 film is only half of the oxygen content in
45
46 the R2 film, it is clear that the density of the sputtering target influences also the oxygen
47
48 incorporation in the film, a denser target leading to a lower oxygen content.
49
50
51
52
53
54
55
56
57
58
59
60

Table 1. Chemical composition of the $\text{Li}_x\text{TiO}_y\text{S}_z$ thin films deduced from ICP and RBS experiments

Thin films	Li/Ti (ICP)	O/Ti (RBS)	S/Ti (RBS)	Chemical formula
R1 (Chemical lithiation)	1.18 \pm 0.02	0.5 \pm 0.1	2.07 \pm 0.04	$\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$
R2 (Solid state reaction)	0.62 \pm 0.02	0.6 \pm 0.1	1.80 \pm 0.04	$\text{Li}_{0.6}\text{TiO}_{0.6}\text{S}_{1.8}$
R3 (SPS of precursors)	1.03 \pm 0.02	0.3 \pm 0.1	2.04 \pm 0.04	$\text{Li}_{1.0}\text{TiO}_{0.3}\text{S}_{2.0}$

XPS analysis of the films deposited from the different targets was performed. The corresponding Ti2p and S2p spectra are shown in Figure 1. According to a previous study of TiO_yS_z films,²¹ Ti2p_{3/2-1/2} core peaks for R1 films (Figure 1a) were fitted with three doublets: i) a doublet attributed to Ti^{4+} ions in a sulfur environment as in TiS_2 (456.1 – 462.2 eV) or in TiS_3 (456.0 – 462.2 eV), ii) a doublet assigned to Ti^{4+} ions in a majority oxygen environment considering the binding energy close to the reference for rutile TiO_2 (458.5 – 464.1 eV), iii) an additional doublet required to fit the experimental curve, located at intermediate binding energies between the two previous ones corresponding to titanium ions into a mixed oxygen-sulfur environment. Then, only Ti^{4+} species were evidenced in the R1 sample, with relative amounts of 24%, 45% and 31%, respectively for TiS_2 - TiS_3 , TiOS and TiO_2 -like environments (Table S1). The S2p_{3/2-1/2} XPS core peak displays only one doublet at 161.0-162.2 eV assigned to S^{2-} ions as in the TiS_2 reference compound (160.9 – 162.1 eV) or TiS_3 (161.0 – 162.1 eV).²¹ These results are actually fully consistent with the charge balance for the $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ composition, and means that all the Ti^{3+} present initially in the LiTiS_2 target is oxidized (by trapping oxygen traces) during the sputtering process.

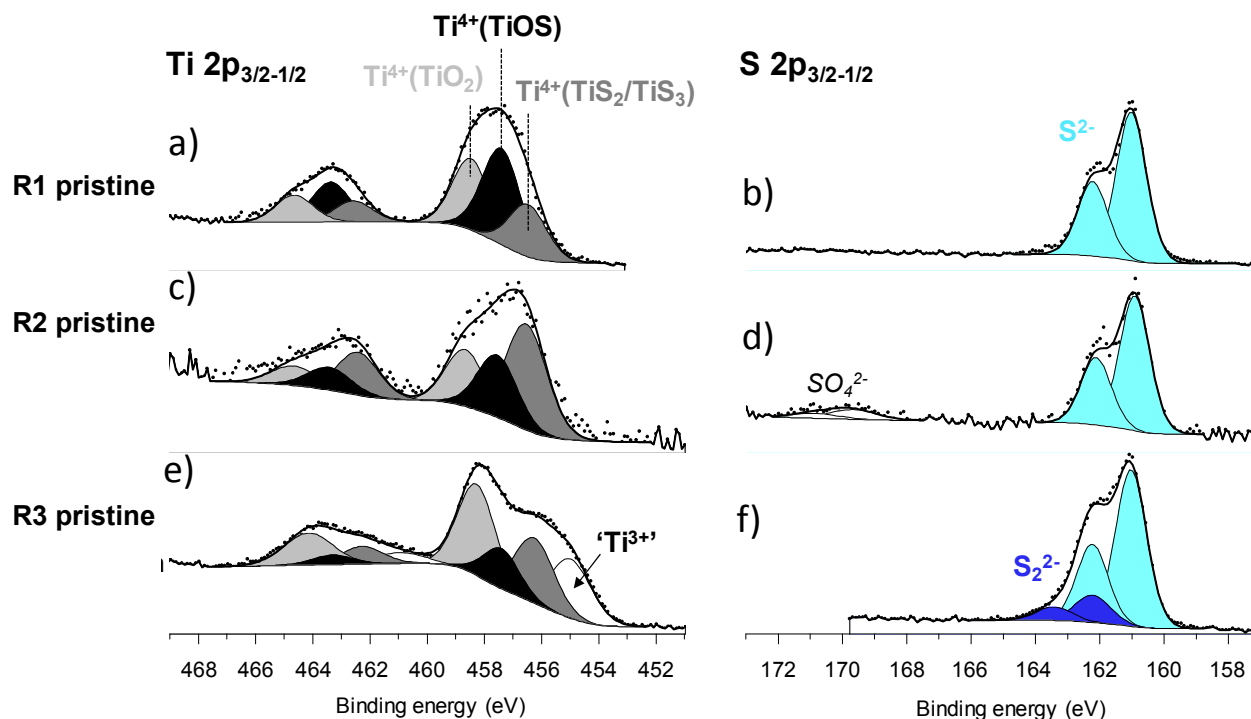


Figure 1. Ti 2p and S 2p core peak spectra of ‘LiTiOS’ thin films prepared from (a,b) R1, (c,d) R2 and (e,f) R3 targets using standard sputtering conditions.

The Ti2p spectrum of the R2 film exhibits the three same environments for titanium (Figure 1c) (sulfur, mixed oxygen-sulfur and oxygen environment) in relative amounts of 48%, 27% and 25% (Table S1). As for the R1 film, the S2p_{3/2-1/2} spectrum (Figure 1b) clearly evidences the presence of only S²⁻ ions. Note that a small component attributed to sulfate ions is detected at 169.7-170.9 eV, probably corresponding to contamination. As expected from the Li_{0.6}TiO_{0.6}S_{1.8} film composition, XPS analysis is also consistent with a full Ti⁴⁺ content. Consequently, as in Li_{2-x}FeS₂ materials,²⁹⁻³¹ further electrochemical lithium de-insertion from these electrodes should necessarily involve the oxidation of anionic species.

In the case of the R3 film, in addition to the doublets corresponding to Ti⁴⁺ ions in the three types of environments (TiS₂-TiS₃ 25%, TiOS 14%, TiO₂ 38%), the Ti2p_{3/2-1/2} XPS core peak (Figure 1e) exhibits a doublet located at 455.0-461.0 eV (23%) assigned to a more reduced state titanium species,

1
2
3 probably Ti^{3+} ions. Moreover, the $\text{S}2\text{p}_{3/2-1/2}$ core peak (Figure 1f) displays a doublet at 161.0-162.2
4 eV corresponding to S^{2-} ions (83%) and another one at 162.2-163.4 eV attributed to a small amount
5 of sulfur in S_2^{2-} disulfide pairs (17%). The relative percentages clearly reveal that the titanium in
6 sulfur environment is closer to TiS_2 rather than TiS_3 . Then, further redox processes accompanying
7 the lithium de-insertion from this material should necessary involve both titanium and sulfur species.
8
9

10 11 12 13 14 15 16 17 **3.3. Electrochemical behavior of Li-Ti(O)S thin films – Sulfide anions participating to** 18 **the redox process** 19

20 The electrochemical behavior of the different $\text{Li}_x\text{TiO}_y\text{S}_z$ thin films was firstly assessed in liquid
21 electrolyte cells. After assembly, the open circuit voltage of the series of cells prepared in standard
22 conditions from R1, R2 and R3 targets was respectively 2.1, 2.2 and 1.6 V. Cyclic voltammetry
23 (Figures 2a-c) was performed at first, starting by the oxidation (delithiation) of the material, in order
24 to reveal the different electrochemical steps. The first oxidation of the films prepared with R1 and
25 R2 targets proceeds with a single step at $\sim 2.6 \text{ V/Li}^+/\text{Li}$, which is reversible (Figure 2a and Figure
26 2b). As the latter leads to the appearance of a new doublet on the $\text{S}2\text{p}$ spectra (Figure 3) for both R1
27 and R2 electrodes, respectively located at 162.2-163.4 eV and 162.3-163.5 eV, which corresponds to
28 S_2^{2-} disulfide pairs, it can be clearly attributed to the oxidation of S^{2-} in S_2^{2-} . Besides, quantitative
29 analyses (Table S2) show that the $\text{S}_2^{2-}/\text{S}^{2-}$ ratio is higher in the fully oxidized R1 electrode than in the
30 R2 one; this is consistent with the higher delithiation capacity theoretically expected from the initial
31 film compositions ($\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ for R1, $\text{Li}_{0.6}\text{TiO}_{0.6}\text{S}_{1.8}$ for R2) and also from the actual capacity
32 measured during the first charge (Figure 2d). Compared to pristine materials, only a slight evolution
33 of the $\text{Ti}2\text{p}_{3/2-1/2}$ core peaks occurs during the charge, the latter being connected to a small variation
34 of the proportion of the three Ti^{4+} environments.
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

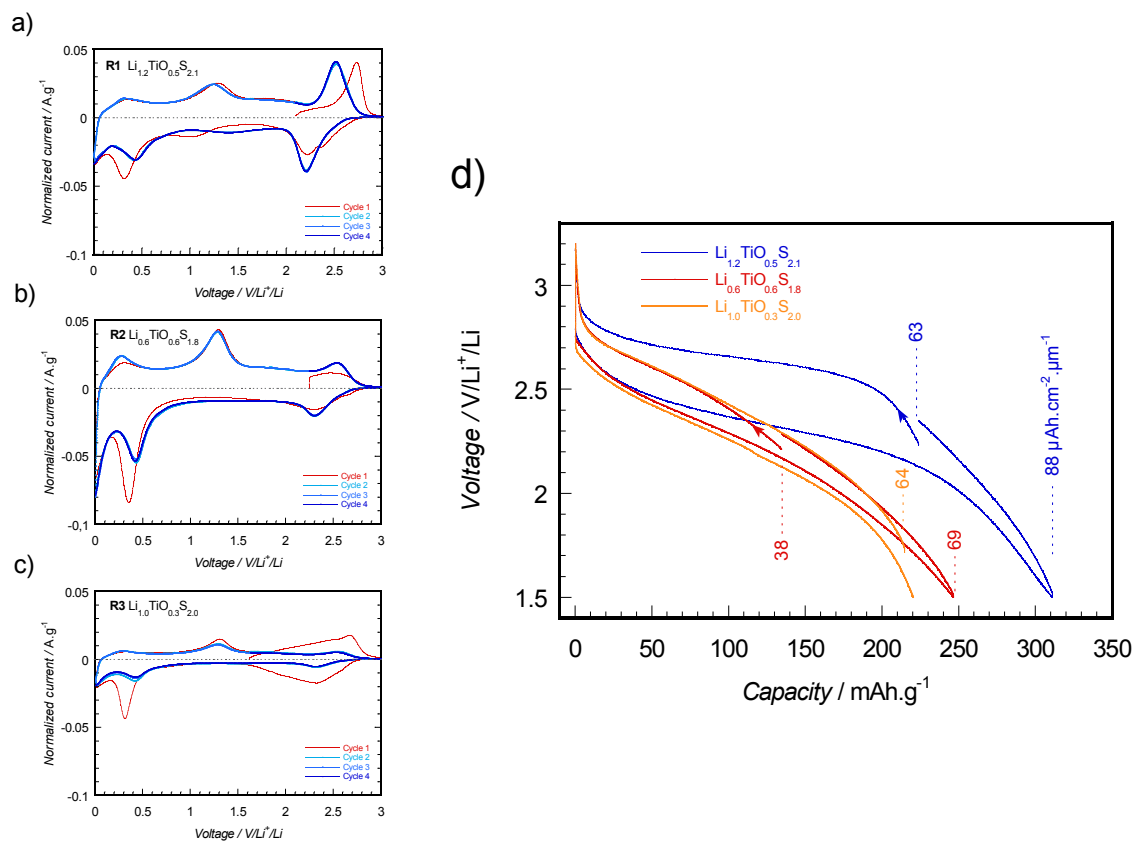


Figure 2. Electrochemical behavior of thin films in button cells (1M LiPF₆, EC:DMC 1:1 electrolyte). Cyclic voltammetry of thin film electrodes prepared from R1 (a), R2 (b) and R3 (c) targets (sweep rate 5 μV s⁻¹). Comparison of the first galvanostatic cycle (2 μA cm⁻²) for the three materials (d), with indication of the volumetric capacity.

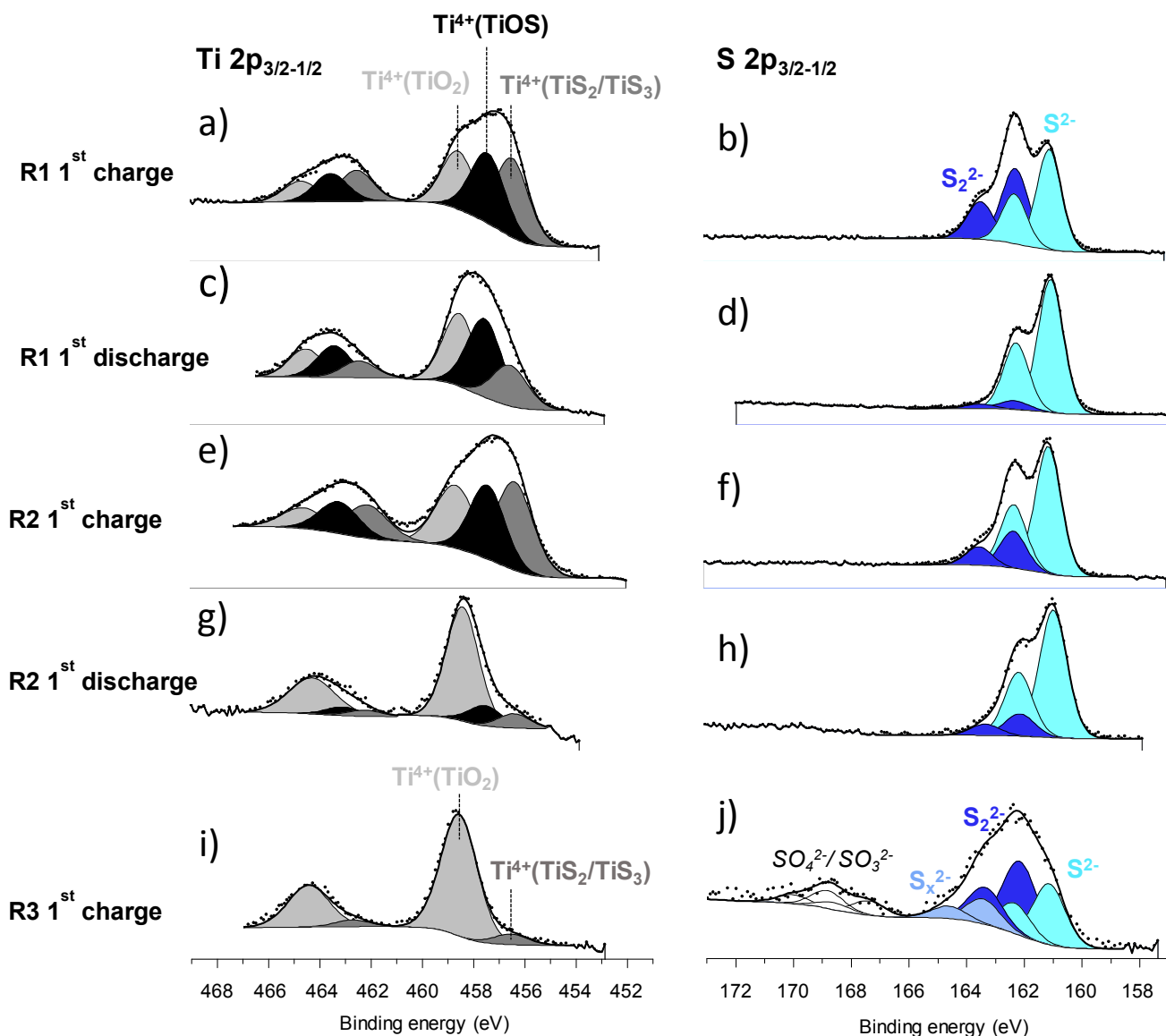


Figure 3. Evolution of Ti2p (a, c, e, g, i) and S2p (b, d, f, h, j) core peaks spectra between the end of the first charge at 3.2 V/Li⁺/Li and the end of the subsequent discharge at 1.5 V/Li⁺/Li, for the three types of films (R1, R2 and R3).

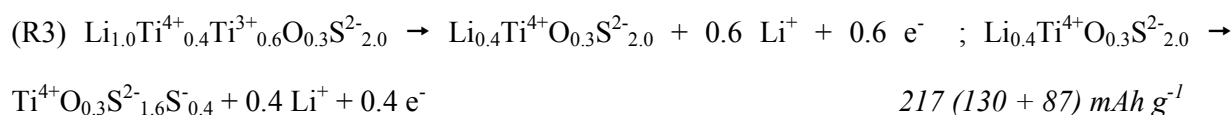
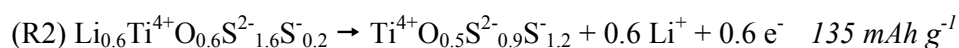
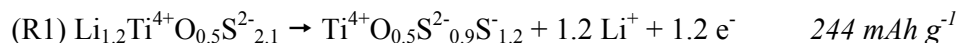
The first oxidation of the electrode prepared with the R3 target, starting from a significantly lower OCV, hence from a more reduced state compared to the two others, is characterized by a quite progressive increase of the current ending with a more marked peak at around 2.6 V/Li⁺/Li attributable to the S²⁻/S₂²⁻ reaction (Figure 2c). Indeed, as deduced from the S2p_{3/2-1/2} core peak

(Figure 3), it is also found here that the proportion of S^{2-} species is significantly higher in the fully charged sample (40%) than in the pristine material (17%). Therefore, the first oxidation stages must be related to the oxidation of Ti^{3+} initially present in the material into Ti^{4+} ; then, titanium and sulfide-based processes partially overlap during the continuation of the oxidation and also during the subsequent reduction. Regarding the $Ti2p_{3/2-1/2}$ core peak spectrum, it reveals the disappearance of the ' Ti^{3+} ' doublet and presence of a majority of TiO_2 -like environments (89%) that is quite surprising. The latter point and the increase of the ' TiO_2 ' contribution in discharged R1 and R2 electrodes (Figure 3) suggest that extra ' TiO_2 ' species are spontaneously formed when Ti^{3+} species are in contact with the liquid electrolyte. The same reaction is likely to occur when the pristine R3 film is immersed in the liquid electrolyte, causing a spontaneous ' TiO_2 ' enrichment at the surface that remains once the electrode is charged. Indeed, as it has been already evidenced for reduced $Li_4Ti_5O_{12}$ electrodes, Ti^{3+} is prone to react with water traces and with ethylene carbonate.^{32,33} Therefore, it is actually difficult to follow the course of the redox processes involving Ti species. Besides, a quite reversible process occurs also below 1.5 V (reduction peak at 0.4 V/ Li^+/Li , then oxidation peak at 1.3 V/ Li^+/Li) for all samples. The corresponding coulombic balance and other results related to the reduction of TiS_2 indicate that the latter corresponds to the Ti^{3+}/Ti^{2+} redox reaction.^{34,35}

The first galvanostatic cycle in the 1.5-3.2 V/ Li^+/Li potential window is presented in Figure 2d for the three types of film. The voltage curve of the $Li_{1.2}TiO_{0.5}S_{2.1}$ film (R1 target) displays a long plateau at 2.6-2.7 V on which proceeds the extraction of almost all the lithium contained in the pristine material ($\Delta xLi = -1.16$) and the concomitant oxidation of S^{2-} species. The full delithiation of $Li_{0.6}TiO_{0.6}S_{1.8}$ is also achieved during the first charge at 3.2 V. Nevertheless, whereas the same electrochemical process occurs, a lower polarization is measured here. Concerning the film from the R3 target, its voltage curve is quite identical to the one of $Li_{0.6}TiO_{0.6}S_{1.8}$, except that the starting point corresponds to a more reduced and lithiated material. Delithiation of this material proceeds with the oxidation of Ti^{3+} , then of S^{2-} species, exhibiting a 40:60 capacity ratio, which is close to the

one expected from the initial $\text{Li}_{1.0}\text{TiO}_{0.3}\text{S}_{2.0}$ composition. The use of a denser 'LiTiS₂' target actually allows enhancing the kinetics of erosion of the target surface at the expense of its oxygenation (poisoning) during the sputtering process, hence is a means of moving to a less reactive sputtering mode and preserving reduced titanium species.³⁶

On the basis of the initial composition of the films and the charge balance, and assuming that Ti^{3+} and S_2^{2-} species do not coexist in the pristine materials, a formal distribution of the $\text{Ti}^{3+}/\text{Ti}^{4+}$ and $\text{S}^{2-}/\text{S}_2^{2-}$ redox processes involved during the first charge and the corresponding specific capacities can be proposed:



These values are actually in good agreement with the experimental ones, both concerning the full charge capacity and the distribution of the cation- and anion-based redox processes in the R3 film.

Finally, first charge volumetric capacities of thin films prepared from R1 and R3 targets are quite identical (220mAh g^{-1} or $64 \mu\text{Ah cm}^{-2} \mu\text{m}^{-1}$) and significantly higher than the one of the films prepared from the R2. They are close to the capacity reported for LiCoO_2 thin films.¹³ As only this initial charge capacity can be actually used in a Li-ion system, additional electrochemical characterizations in solid state microbatteries were only performed on thin films having the highest capacities (i.e. R1 and R3 thin films). Moreover, these thin film electrodes illustrate respectively the case of electrode materials involving either the larger or the lower amount of sulfide species in the redox process.

1
2
3
4
5 All-solid-state Li/LiPON/Li_xTiO_yS_z thin film batteries were then manufactured to study the
6
7 electrochemical behavior of both Li_xTiO_yS_z materials prepared from R1 and R3 targets (Figure 4).
8
9
10 The stability of the voltage curves (shape and capacity value) versus cycles during constant current
11
12 cycling in the 1.5 – 3.2 V/Li⁺/Li range (Figure 5) highlights the perfect reversibility of the
13
14 electrochemical processes in any case, whatever the current density. From both sample curves, it
15
16 appears that the polarization is very low in the part of the curve involving the Ti³⁺/Ti⁴⁺ cation-based
17
18 process, whereas it is greatly increased in the part involving mainly the S²⁻/S₂²⁻ anion-based process.
19
20 This evolution is likely to originate from a marked decrease in the electronic conductivity in the
21
22 material as the mixed-valence character of the material tends to disappear (i.e. full oxidation of Ti³⁺
23
24 species). When increasing the current density up to 130 μA cm⁻² (~ 2C rate), the processes remain
25
26 perfectly reversible in both materials. The slight increase of the polarization, visible along the main
27
28 part of the charge/discharge curve, is mainly related to the ohmic drop in the LiPON electrolyte (~ 50
29
30 mV). Nevertheless, the polarization rises more significantly at the end of the charge, especially for
31
32 the sample R1, inducing a slight decrease of the capacity for both materials. Then, this capacity is
33
34 perfectly stable (-0.0015% / cycle) for hundreds of cycles.
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

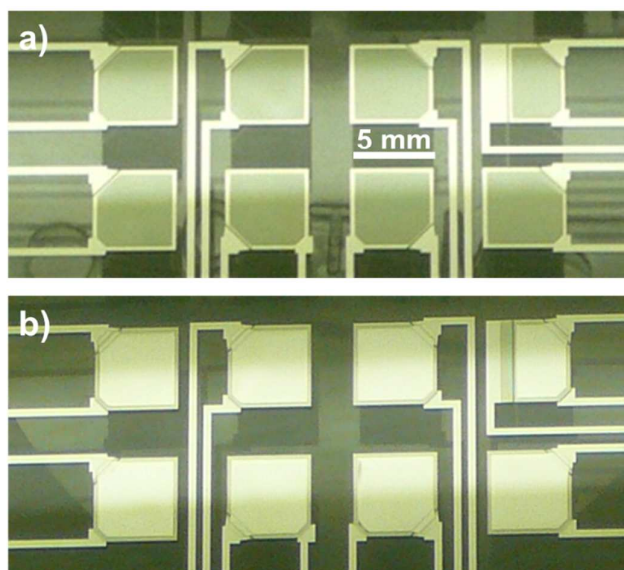


Figure 4. (a) W/LiTiOS thin film electrodes (25 mm^2) deposited on a glass substrate, prior to the deposition of the LiPON electrolyte and the lithium negative electrode. (b) Complete W/LiTiOS/LiPON/Li microbatteries.

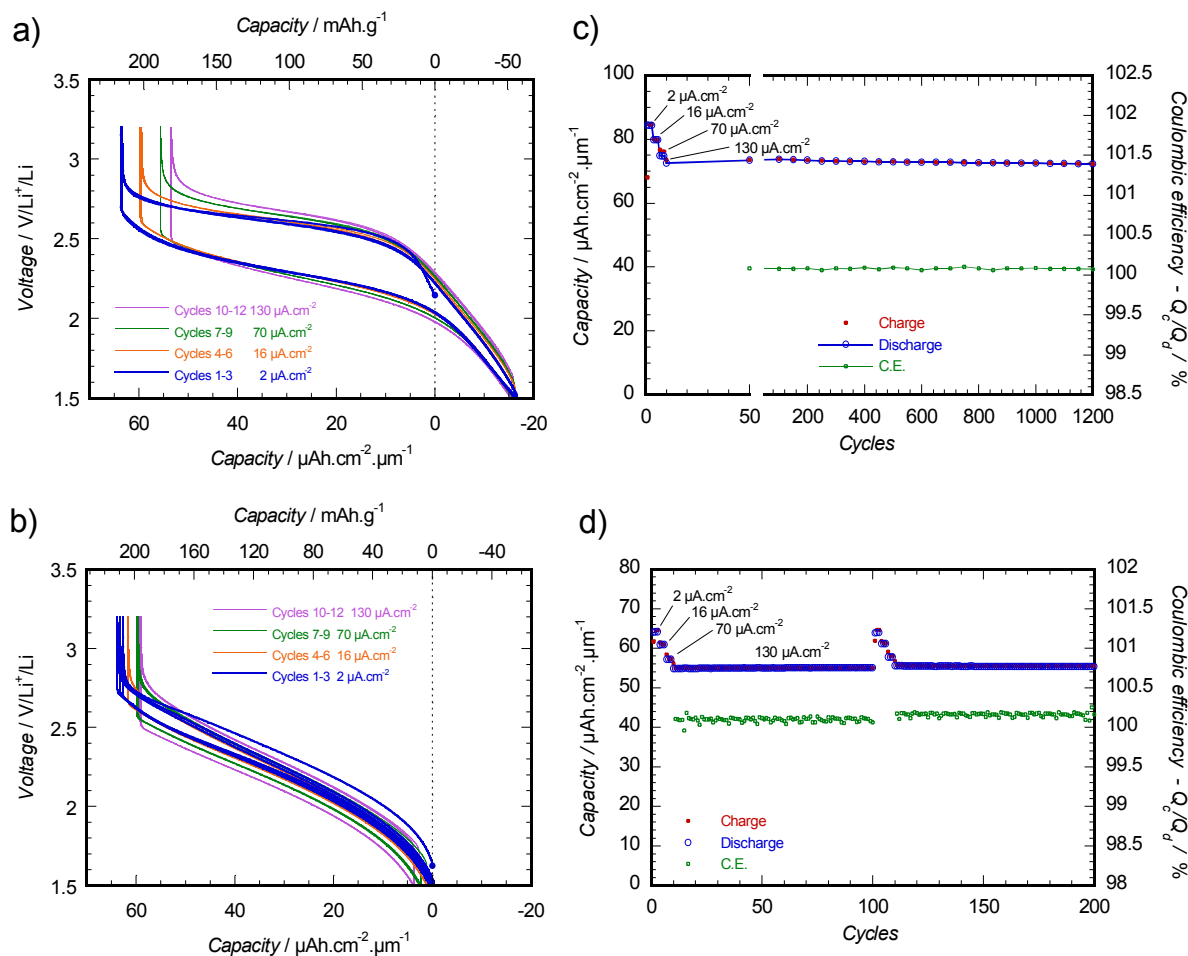


Figure 5. Galvanostatic cycling of (a) $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ (R1) and (b) $\text{Li}_{1.0}\text{TiO}_{0.3}\text{S}_{2.0}$ (R3) thin films in all-solid-state Li/LiPON/LiTiOS batteries at various current densities ($130 \mu\text{Ah cm}^{-2} \sim 2\text{C}$ rate), and (c,d) the respective evolution of both their volumetric capacity and their coulombic efficiency. Insert shows $\text{W}/\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ positive electrodes deposited on a glass wafer prior the achievement of the full all-solid-state cells.

The evolution of the impedance of a Li/LiPON/ $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ (R1) microbattery was studied as a function of the state-of-charge (SOC) during a typical cycle between 1.5 V and 3.2 V (Figure 6), using a GITT protocol (Figure 7a). Between 0% and 90% SOC, the stable contribution of the LiPON electrolyte (first $R_{\text{LiPON}}//\text{CPE}_{\text{LiPON}}$ semi-circle) represents the major part of the impedance and only slight modifications of the medium-low frequency part related to the diffusion process and the charge

accumulation occur. Then, from 90% to 100% SOC, a new contribution appears and grows to become rapidly far larger than the electrolyte contribution. When discharging the cell, a reverse evolution is observed with the rapid fade of this additional contribution. Besides, as revealed in Figure 6a and Figure 6d, the characteristic frequency of the additional component is decreasing when the SOC is increasing (and *vice versa*).

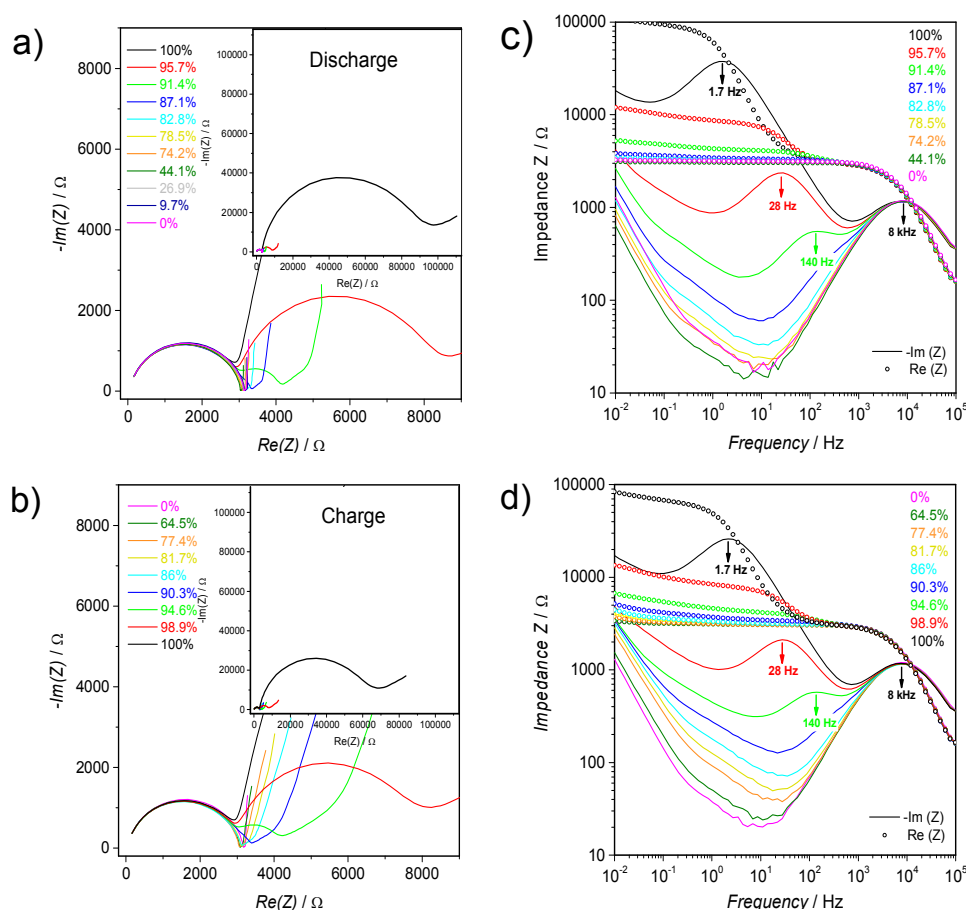


Figure 6. Two representations of the evolution of EIS spectra as a function of the state-of-charge of a Li/LiPON/ Li_{1.2}TiO_{0.5}S_{2.1} (R1) microbattery during a discharge and the subsequent charge: (a,b) Nyquist diagrams and (c,d) the corresponding pseudo-Bode plots.

1
2
3 This contribution that could have been related to the formation of an interphase at the
4 LiPON/Li_xTiOS interface at the highest voltage values, is therefore more likely induced by a sharp
5 increase of the charge transfer resistance at this interface when the electrode material is almost fully
6 oxidized.³⁷ This is the phenomenon that actually hinders the complete delithiation of the material
7 during charges carried out at high current rates (Figure 5a). A similar trend was also observed for
8 Li/LiPON/LiTiOS (R3) cells.
9

10
11 In addition, calculation of the lithium diffusion coefficient (\tilde{D}_{Li^+}) in Li_xTiO_{0.5}S_{2.1} ($0 < x \leq 1.6$)
12 materials was carried out by processing the Warburg part of these EIS spectra, using the method
13 proposed by Huggins and *al.* and a fixed value of the molar volume (V_M) corresponding to 1T-
14 TiS₂.³⁸ Thus, \tilde{D}_{Li^+} was found to increase regularly from 10⁻¹² to 10⁻⁹ cm².s⁻¹ from the fully de-
15 lithiated state to the Li_{1.6}TiO_{0.5}S_{2.1} composition (Figure 7b). Examining these EIS results and other
16 ones previously obtained for TiO_{0.6}S_{1.6} sputtered films (i.e. containing ~ 0.2 S₂²⁻ in the starting
17 material),¹⁸ it appears that in both cases a large decrease of \tilde{D}_{Li^+} (from 10⁻¹⁰ to 10⁻¹² cm² s⁻¹) is
18 observed concomitantly with the increase of the S₂²⁻ pairs concentration. The shortening of some S-
19 S distances (from 3.407-3.46 Å for S²⁻ - S²⁻ in TiS₂ to 2.04 Å for (S-S)²⁻ in TiS₃ or amorphous WS₃
20 sulfides),³⁹⁻⁴¹ inducing a marked distortion of the anion network, is likely to be at the origin of this
21 phenomenon. Finally, \tilde{D}_{Li^+} values measured in Li_xTiO_{0.5}S_{2.1} materials are similar or slightly higher
22 than the ones for LiCoO₂ and LiMn₂O₄ materials.⁴²⁻⁴³
23
24
25
26
27
28
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

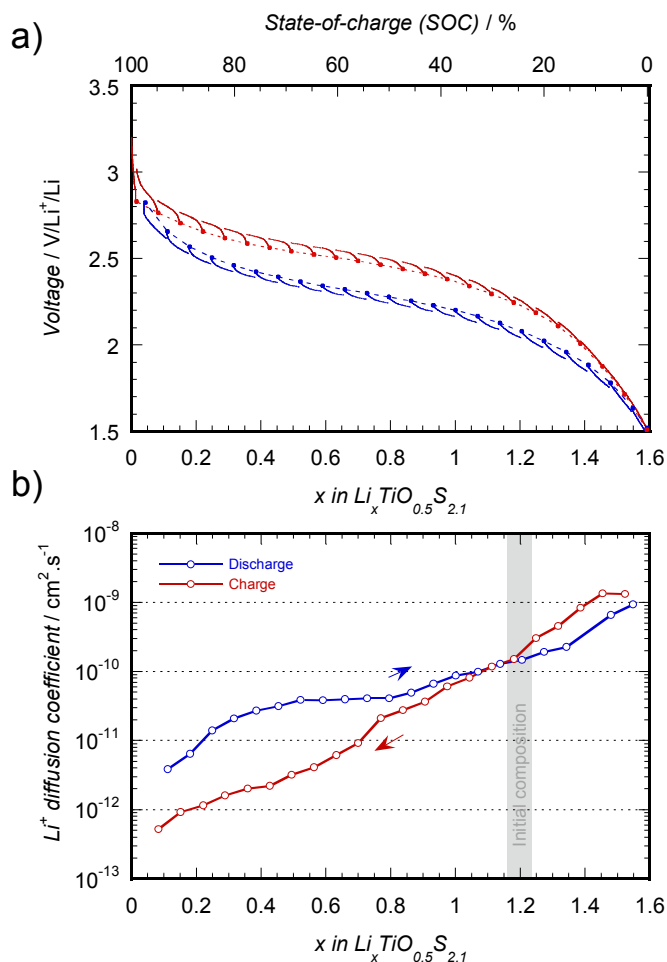
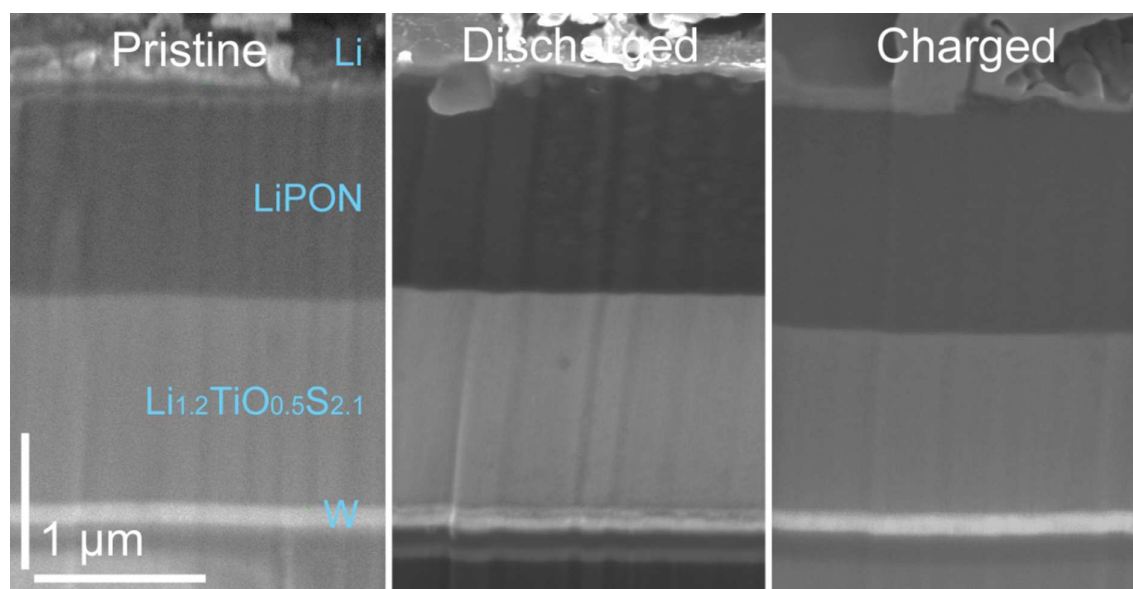


Figure 7. (a) Voltage profile of a Li/LiPON/ $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ (R1) microbattery measured during a GITT experiment. Dots correspond to the OCV measured after 4 hours of relaxation (blue: discharge, red: charge). EIS measurements were carried at the end of every rest period. (b) Lithium diffusion coefficient in $\text{Li}_x\text{TiO}_{0.5}\text{S}_{2.1}$ ($0 < x \leq 1.6$) calculated from EIS spectra.

Cross sections of all-solid-state Li/LiPON/ $\text{Li}_{1.2}\text{TiO}_{0.5}\text{S}_{2.1}$ (R1) microbatteries prepared by Focus Ion Beam etching (Figure 8) were observed by SEM either in the charged or in the discharged state. They reveal a dense all-solid-state stack with well-defined interfaces, with no evidence of the formation of an eventual interphase between LiPON and the fully charged cathode. Besides, a noticeable volume change of the electrode ($\sim 20\%$) is generated between the charged and the

1
2
3 discharged states. The latter has obviously no detrimental effect on the cycling behavior of these
4
5 microbatteries.
6



29
30
31
32
33
34
35
36
37

Figure 8. SEM-FIB image of Li/LiPON/Li_{1.2}TiO_{0.5}S_{2.1} (R1) cells either in the pristine state, or stopped after hundreds of cycles either in the charged (3.2 V/Li⁺/Li) and the discharged state (1.5 V/Li⁺/Li).

38 39 40 41 42 43 44 45 46 47 48 49 50 51 52 53 54 55 56 57 58 59 60

4. Conclusion

Lithiated titanium oxysulphide thin film electrodes were prepared by sputtering from home-made targets having a LiTiS₂ composition. Target precursors, manufacturing process and the resulting compactness of the target were all found to clearly influence the composition of the films, hence their electrochemical behavior in all-solid-state microbatteries. Variations in the composition of the thin films are mainly related to the lithium content and the oxygen uptake during the sputtering process. While the first one directly influences the practical capacity of the electrode in a Li-ion system, the latter mainly determines the nature of the electrochemical processes occurring during the delithiation of these materials. Indeed, oxygen incorporation in the film proceeds with the oxidation

1
2
3 of Ti species originally present in the target into Ti^{4+} . As a consequence, it is observed a tendency to
4
5 diverge from the LiTiS_2 targeted film composition towards $\text{LiTiO}_{0.5}\text{S}_2$. That initiates the participation
6
7 of an increasing amount of anion species to the redox process ($\text{S}^{2-}/\text{S}_2^{2-}$ couple) at the expense of the
8
9 $\text{Ti}^{3+}/\text{Ti}^{4+}$ process, inducing a progressive increase of both the mean operating voltage and the
10
11 polarization. The use of targets having high compactness, as the ones prepared by SPS, is actually a
12
13 promising means to achieve a better control of the process, hence of the composition of the films.

14
15
16 Finally, amorphous lithiated titanium oxysulphide thin film electrodes embedded in all-solid-state
17
18 lithium microbatteries were able to deliver capacities as high as $85 \mu\text{Ah cm}^{-2} \mu\text{m}^{-1}$ (or 270mAh g^{-1}),
19
20 of which a useful part of $65 \mu\text{Ah cm}^{-2} \mu\text{m}^{-1}$, i.e. the same as for LiCoO_2 , is actually usable in the
21
22 prospect of an integration in a Li-ion cell.¹¹ Operation involving cation and/or anion-based redox
23
24 processes revealed in any case the perfect reversibility of the latter and an excellent capacity
25
26 retention (-0.0015% cycle⁻¹) during several hundreds of cycles. Contrary to LiCoO_2 , the synthesis of
27
28 these high capacity lithiated thin film electrodes does not require any annealing nor careful control of
29
30 the preferred crystal orientation,^{13,44,45} hence is simple and is interestingly compatible with the
31
32 deposition on flexible polymer substrates. Finally, taking into account the facile oxidation of Ti^{3+}
33
34 during the sputtering process and the perfect reversibility of the anion-based redox process, the
35
36 preparation of thin electrodes from $\text{Li}_2\text{S-TiS}_2$ target compositions appears as a promising prospect.⁴⁶
37
38
39
40
41
42

43 **Supporting Information**

44
45 Electrochemical behavior of R1 'LiTiOS' films prepared with different total pressures, XRD patterns
46
47 of LiTiS_2 powders prepared by chemical lithiation of TiS_2 (R1) and by solid state reaction of
48
49 $2\text{Li}_2\text{S}:\text{Ti}:\text{TiS}_2$ (R2 and R3 target), SEM images of R1 and R2 thin films, RBS analysis of the R1
50
51 thin film and quantitative XPS analyses of 'LiTiOS' thin films performed before and after
52
53 electrochemical cycling.
54
55
56
57
58
59
60

1
2
3
4
5
6
7
8
9
10
11
12
13
14
15
16
17
18
19
20
21
22
23
24
25
26
27
28
29
30
31
32
33
34
35
36
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

Acknowledgements

This work was partly supported by the French Ministry of industry and STMicroelectronics in the framework of the Investissements d'Avenir program (Tours2015 project). Jean-Marc Boissel (CEA) is acknowledged for having completed the manufacture of microbatteries at CEA Grenoble. The authors are also grateful to Stéphanie Sorieul from the AIFIRA platform (CENBG-CNRS/IN2P3 - Bordeaux University) for her help in acquiring RBS spectra and Michel Lahaye (ICMCB) for Auger Electron Spectroscopy depth profiling.

References

- (1) MacKenzie, J. D.; Ho, C. Perspectives on Energy Storage for Flexible Electronic Systems. *Proc. IEEE* **2015**, *103*, 535-553.
- (2) Wang, Y.; Liu, B.; Li, Q.; Cartmell, S.; Ferrara, S.; Deng, Z. D.; Xiao, J. Lithium and Lithium-ion Batteries for Applications in Microelectronic Devices: A Review. *J. Power Sources* **2015**, *286*, 330.
- (3) Cymbet Home Page: <http://www.cymbet.com> (accessed Dec 15, 2016).
- (4) Front Edge Technology Home Page: <http://www.frontedgetechnology.com> (accessed Dec 15, 2016).
- (5) Prologium Home Page: <http://www.prologium.com> (accessed Dec 15, 2016).
- (6) ST Microelectronics EFL700 Datasheet: http://www.st.com/content/st_com/en/products/power-management/battery-management-ics/enfilm-thin-film-batteries/efl700a39.html (accessed Dec 15, 2016).
- (7) Kanehori, K.; Matsumoto, K.; Miyauchi, K.; Kudo, T. Thin Film Solid Electrolyte and its Application to Secondary Lithium Cell. *Solid State Ionics* **1983**, *9-10*, 1445-1448.
- (8) Jourdain, L.; Souquet, J. L.; Delord, V.; Ribes, M. Lithium Solid State Glass-based Microgenerators. *Solid State Ionics* **1988**, *20-30*, 1490-1494.

- 1
2
3 (9) Meunier, G.; Dormoy, R.; Levasseur, A. New Positive Electrode Materials for Lithium Thin Film
4
5 Secondary Batteries. *Mater. Sci. Eng. B* **1989**, *83*, 19-23.
6
7 (10) Bates, J. B.; Dudney, N. J.; Gruzalski, G. R.; Zuhr, R. A.; Choudhury, A.; Luck, C. F.;
8
9 Robertson, J. D. Fabrication and Characterization of Amorphous Lithium Electrolyte Thin Films and
10
11 Rechargeable Thin-film Batteries. *J. Power Sources*, **1993**, *43*, 103-110.
12
13 (11) Le Cras, F.; Pecquenard, B.; Dubois, V.; Phan, V. P.; Guy-Bouyssou, D. All-solid-state
14
15 Lithium-ion Microbatteries using Silicon Nanofilm Anodes: High Performance and Memory Effect.
16
17 *Adv. Energy Mater.* **2015**, *5*, 1501061.
18
19 (12) Neudecker, B. J.; Dudney, N. J.; Bates, J. B. "Lithium-free" Thin-film Battery with In Situ
20
21 Plated Li Anode. *J. Electrochem. Soc.* **2000**, *147*, 517-523.
22
23 (13) Bates, J. B.; Dudney, N. J.; Neudecker, B. J.; Hart, F. X.; Jun, H. P.; Hackney, S. A. Preferred
24
25 Orientation of Polycrystalline LiCoO₂ Films. *J. Electrochem. Soc.* **2000**, *147*, 59-70.
26
27 (14) Dudney, N. J.; Jang, Y. I. Analysis of Thin-film Lithium Batteries with Cathodes of 50 nm to 4
28
29 μm Thick LiCoO₂. *J. Power Sources* **2003**, *119-121*, 300-304.
30
31 (15) Liao, C. L.; Fung, K. Z. Lithium Cobalt Oxide Cathode Film Prepared by RF Sputtering. *J.*
32
33 *Power Sources* **2004**, *128*, 263-269.
34
35 (16) Park, H. Y.; Nam, S. C.; Lim, Y. C.; Choi, K. G.; Lee, K. C.; Park, G. B.; Kim, J. B.; Kim, H.
36
37 P.; Cho, S. B. LiCoO₂ Thin Film Cathode Fabrication by Rapid Thermal Annealing for Micro Power
38
39 Sources. *Electrochimica Acta* **2007**, *52*, 2062-2067.
40
41 (17) Shiraki, S.; Oki, H.; Takagi, Y.; Suzuki, T.; Kumatani, A.; Shimizu, R.; Haruta, M.; Ohsawa, T.;
42
43 Sato, Y.; Ikuhara, Y.; Hitosugi, T. Fabrication of All-solid-state Battery using Epitaxial LiCoO₂ Thin
44
45 Films. *J. Power Sources* **2014**, *267*, 881-887.
46
47 (18) Fleutot, B.; Pecquenard, B.; Le Cras, F.; Delis, B.; Martinez, H.; Dupont, L.; Guy-Bouyssou, D.
48
49 Characterization of All-solid-state Li/LiPONB/TiOS Microbatteries Produced at the Pilot Scale. *J.*
50
51 *Power Sources* **2011**, *196*, 10289-10296.
52
53
54
55
56
57
58
59
60

- 1
2
3 (19) Dines, M. B. Lithium Intercalation via n-butyl lithium of the Layered Transition Metal
4 Dichalcogenides. *Mater. Res. Bull.* **1975**, *10*, 287-292.
5
6
7 (20) Wiedemann, D.; Nakhal, S.; Senyshyn, A.; Bredow, T.; Lerch, M. The High-temperature
8 Transformation from 1T- to 3R-Li_xTiS₂ (x = 0.7, 0.9) as Observed In Situ with Neutron Powder
9 Diffraction. *Z. Phys. Chem.* **2015**, *229*, 1275-1288.
10
11
12 (21) Lindic, M. H.; Pecquenard, B.; Vinatier, P.; Levasseur, A.; Martinez, H. ; Gonbeau, D.; Petit, P.
13 E.; Ouvrard, G. Characterization of RF Sputtered TiO_yS_z Thin Films. *Thin Solid Films* **2005**, *484*,
14 113-123.
15
16
17 (22) Phan, V. P.; Pecquenard, B.; Le Cras, F. High-performance All-solid-state Cells Fabricated with
18 Silicon Electrodes. *Adv. Funct. Mater.* **2012**, *22*, 2580-2584.
19
20
21 (23) Eckstein, W.; Mayer, M. Rutherford Backscattering from Layered Structures Beyond the Single
22 Scattering Model. *Nucl. Instrum. Methods Phys. Res., Sect. B* **1999**, *153*, 337-344.
23
24
25 (24) Shirley, D. A. High-resolution X-ray Photoemission Spectrum of the Valence Bands of Gold.
26 *Phys. Rev. B* **1972**, *5*, 4709.
27
28
29 (25) Scofield, J. H. Photoionization Cross Sections at 1254 and 1487 eV. *J. Electron. Spectrosc.*
30 **1976**, *8*, 129-137.
31
32
33 (26) Whittingham, M. S. The Role of Ternary Phases in Cathode Reactions. *J. Electrochem. Soc.*
34 **1976**, *123*, 315-320.
35
36
37 (27) Colbow, J.; Dahn, J. R.; Haering, R. R. The 3R Phase of Li_xTiS₂. *J. Power Sources* **1989**, *26*,
38 301-307.
39
40
41 (28) Donaghey, L. F.; Geraghty, K. G. Effect of Target Oxidation on Reactive Sputtering Rates of
42 Titanium in Argon-Oxygen Plasmas. *Thin Solid Films* **1976**, *38*, 271-280.
43
44
45 (29) Blandeau, L.; Ouvrard, G.; Calage, Y.; Brec, R.; Rouxel, J. Transition-metal Dichalcogenides
46 from Disintercalation Processes. Crystal Structure Determination and Mössbauer Study of Li₂FeS₂
47 and its Disintercalates Li_xFeS₂ (0.2 ≤ x ≤ 2). *J. Phys. C: Solid State Phys.* **1987**, *20*, 4271-4281.
48
49
50
51
52
53
54
55
56
57
58
59
60

- 1
2
3 (30) Barker, J.; Kendrick, E. The Electrochemical Insertion and Safety Properties of the Low-cost Li-
4 ion Active Material, Li_2FeS_2 . *J. Power Sources* **2011**, *196*, 6960-6963.
- 5
6
7 (31) Pelé, V.; Flamary, F.; Bourgeois, L.; Pecquenard, B.; Le Cras, F. Perfect Reversibility of the
8 Lithium Insertion in FeS_2 : The Combined Effects of All-solid-state and Thin Film Cell
9 Configurations. *Electrochem. Comm.* **2015**, *51*, 81-84.
- 10
11
12 (32) Lu, X.; Gu, L.; Hu, Y. S.; Chiu, H. C.; Li, H.; Demopoulos, G. P.; Chen, L. New Insight into the
13 Atomic-scale Bulk and Surface Structure Evolution of $\text{Li}_4\text{Ti}_5\text{O}_{12}$ Anode. *J. Am. Chem. Soc.* **2015**,
14 *137*, 1581-1586.
- 15
16
17 (33) He, M.; Castel, E.; Laumann, A.; Nuspl, G.; Novák, P.; Berg, E. J. In Situ Gas Analysis of
18 $\text{Li}_4\text{Ti}_5\text{O}_{12}$ Based Electrodes at Elevated Temperatures. *J. Electrochem. Soc.* **2015**, *162*, A870-A876.
- 19
20
21 (34) Murphy, D. W.; Carides, J. N. Low Voltage Behavior of Lithium/Metal Dichalcogenide
22 Topochemical Cells. *J. Electrochem. Soc.* **1979**, *126*, 349-351.
- 23
24
25 (35) Kim, Y.; Park, K.; Song, S. H.; Han, J.; Goodenough, J. B. Access to $\text{M}^{3+}/\text{M}^{2+}$ Redox Couples
26 in Layered LiMS_2 Sulfides (M = Ti, V, Cr) as Anodes for Li-ion Battery. *J. Electrochem. Soc.* **2009**,
27 *156*, A703-A708.
- 28
29
30 (36) Schiller, S.; Beister, G.; Sieber, W. Reactive High Rate D.C. Sputtering: Deposition Rate,
31 Stoichiometry and Features of TiO_x and TiN_x Films with Respect to the Target Mode. *Thin Solid*
32 *Films* **1984**, *111*, 259-268.
- 33
34
35 (37) Iriyama, Y.; Kako, T.; Yada, C.; Abe, T.; Ogumi, K. Charge Transfer Reaction at the Lithium
36 Phosphorus Oxynitride Glass Electrolyte/Lithium Cobalt Oxide Thin Film Interface. *Solid State*
37 *Ionics* **2005**, *176*, 2371-2376.
- 38
39
40 (38) Ho, C.; Raistrick, I. D.; Huggins, R. A. Application of A-C Techniques to the Study of Lithium
41 Diffusion in Tungsten Trioxide Thin Films. *J. Electrochem. Soc.* **1980**, *127*, 343-350.
- 42
43
44 (39) Chianelli, R. R.; Scanlon, J. C.; Thompson, A. H. Structure Refinement of Stoichiometric TiS_2 .
45 *Mater. Res. Bull.* **1975**, *10*, 1379-1382.
- 46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

1
2
3 (40) Fureseth, S.; Brattås, L.; Kjekshus, A. On the Crystal Structures of TiS_3 , ZrS_3 , ZrSe_3 , ZrTe_3 ,
4 HfS_3 and HfSe_3 . *Acta Chem. Scand., Ser. A* **1975**, *29*, 623-631.

7 (41) Hibble, S. J.; Walton, R. I.; Feaviour, M. R.; Smith, A. D. Sulfur–Sulfur Bonding in the
8 Amorphous Sulfides WS_3 , WS_5 , and Re_2S_7 from sulfur K-edge EXAFS studies. *J. Chem. Soc.,*
9 *Dalton Trans.* **1999**, *16*, 2877-2883.

14 (42) Xia, H.; Lu, L.; Ceder, G. Li Diffusion in LiCoO_2 Thin Films Prepared by Pulsed Laser
15 Deposition. *J. Power Sources* **2006**, *159*, 1422-1427.

18 (43) Xie, J.; Tanaka, T.; Imanishi, N.; Matsumura, T.; Hirano, A.; Takeda, Y.; Yamamoto, O. Li-ion
19 Transport Kinetics in LiMn_2O_4 Thin Films Prepared by Radio Frequency Magnetron Sputtering. *J.*
20 *Power Sources* **2008**, *180*, 576-581.

25 (44) Bouwman, P. J.; Boukamp, B. A.; Bouwmeester, H. J. M.; Wondergem, H. J.; Notten, P. H. L.
26 Structural Analysis of Submicrometer LiCoO_2 Films *J. Electrochem. Soc.* **2001**, *148*, A311-A317.

29 (45) Yoon, Y.; Park, C.; Kim, J.; Shin, D. Lattice Orientation Control of Lithium Cobalt Oxide
30 Cathode Film for All-solid-state Thin Film Batteries. *J. Power Sources* **2013**, *226*, 186-190.

34 (46) Sakuda, A.; Takeuchi, T.; Okamura, K.; Kobayashi, H.; Sakaebe, H.; Tatsumi, K.; Ogumi, Z.
35 Rock-salt-type Lithium Metal Sulphides as Novel Positive-electrode Materials. *Sci. Rep.* **2014**, *4*,
36 4883.
37
38
39
40
41
42
43
44
45
46
47
48
49
50
51
52
53
54
55
56
57
58
59
60

TOC/Abstract Graphic

