



HAL
open science

Role of evaporation rate on the particle organization and crack patterns obtained by drying a colloidal layer

Keyvan Piroird, Veronique Lazarus, Georges Gauthier, Arnaud Lesaine,
Daniel Bonamy, Cindy Lynn Rountree

► **To cite this version:**

Keyvan Piroird, Veronique Lazarus, Georges Gauthier, Arnaud Lesaine, Daniel Bonamy, et al.. Role of evaporation rate on the particle organization and crack patterns obtained by drying a colloidal layer. EPL - Europhysics Letters, 2016, 113 (3), pp.38002. 10.1209/0295-5075/113/38002 . hal-01279544

HAL Id: hal-01279544

<https://hal.science/hal-01279544>

Submitted on 26 Feb 2016

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Role of evaporation rate on the particle organization and crack patterns obtained by drying a colloidal layer

K. PIROIRD^{1,2}, V. LAZARUS¹ ^(a), G. GAUTHIER¹, A. LESAINE^{1,2}, D. BONAMY² and C. L. ROUNTREE²

¹ *Laboratoire FAST, Univ. Paris-Sud, CNRS, Université Paris-Saclay, F-91405, Orsay, France.*

² *SPEC, CEA, CNRS, Université Paris-Saclay, 91191 Gif-sur-Yvette, France.*

PACS 81.16.Dn – Self-assembly

PACS 82.70 Dd – Colloids

PACS 62.20.mt – Cracks

PACS 68.37.Ps – Surface analysis via Atomic force microscopy (AFM)

Abstract – A scientific hurdle in manufacturing solid films by drying colloidal layers is preventing them from fracturing. This paper examines how the drying rate of colloidal liquids influences the particle packing at the nanoscale in correlation with the crack patterns observed at the macroscale. Increasing the drying rate results in more ordered, denser solid structures, and the dried samples have more cracks. Yet, introducing a holding period (at a prescribed point) during the drying protocol results in a more disordered solid structure with significantly less cracks. To interpret these observations, this paper conjectures that a longer drying protocol favors the formation of aggregates. It is further argued that the number and size of the aggregates increase as the drying rate decreases. This results in the formation of a more disordered, porous film from the viewpoint of the particle packing, and a more resistant film, i.e. less cracks, from the macroscale viewpoint.

Obtaining solid layers via drying of colloidal suspensions is central to many technological fields [1]: Printing and painting [2], manufacturing protective or decorative coatings [3], designing materials at the nanoscale by low-cost processes [4, 5]... From a fundamental viewpoint, colloidal systems provide model systems to mimic the atomic behavior at a larger, more accessible length-scale [6, 7]. Understanding, predicting and controlling the mechanisms driving their self-assembly represents a major challenge. Obtaining this control will aid in ensuring the resistance to failure while staying eco-friendly [3], especially for applications requiring thick and/or hard coatings. Amongst several other parameters (thickness [8], ionic strength [9], substrate adhesion [10]...), the drying rate significantly impacts the cracking patterns [11–13]. This observation remains poorly understood for several reasons. During drying, mass loss plays simultaneously [12] on the evolution of both the loading applying externally to the layer [14, 15] and the intrinsic material properties of the forming solid [16]. From a macroscopic point of view, these two effects can hardly be deconvoluted. An important question to

clarify is whether the drying rate plays only on the formation kinetics yielding the same material or if it changes the way the material self-structures during evaporation leading to variation in the final state.

This Letter experimentally investigates how the evaporation kinetics influences the material properties at the particle (nano)scale and the fracture patterns at the continuum scale. For this purpose, a colloidal suspension of hard monodisperse nanospheres was dried in a controlled atmosphere of tunable humidity. An atomic force microscope (AFM) captures the particle arrangement of the final layer topography. Bulk measurements of the packing density complement these surface analyses. Counter-intuitively [4], these systems exhibit a progressive transition from disorder to order when increasing the drying rate. This transition goes along with an increase of density and more cracks. The balance between particle convection and diffusion alone cannot explain this transition, and particle agglomeration should be considered. Longer drying protocols enhance fracture resistance and favors particle agglomeration, and consequently disordered packing structure and porosity.

Setup– Drying experiments herein invoke model col-

^(a)E-mail: veronique.lazarus@u-psud.fr

loidal suspensions of monodisperse silica nanospheres (Ludox HS-40: $\phi_{SiO_2}^m = 40 \pm 1\%$ in mass of SiO_2 colloids with a spec nominal diameter of 12 nm; $\phi_{Na_2O}^m = 0.42 \pm 0.04\%$ in mass of free alkalinity as Na_2O ; and spec density $\rho_{Ludox} = 1.30 \pm 0.01 \text{ g/cm}^3$). Experiments herein use the same bottle of suspension [17] and occur over a restricted time period (~ 9 months). This ensures a minimal variation and aging of the initial suspension. A glass Petri dish (inner radius $R = 3.5$ cm) initially contains 25 g (m_0) of Ludox. To capture the in-situ mass loss, the Petri dish rests on a precision scale (Sartorius Cubis series, accuracy 10^{-5} g). The layer dries uniformly and forms a flat surface, except in the vicinity of the dish edges where an upwards meniscus prevents the drop singularity (hence the coffee-ring type lateral drying [18]). The layer's initial and final thicknesses are $h_0 \simeq 5$ mm and $h_1 \simeq 1.8$ mm, respectively. An external humidity control system driven by Labview ensures a prescribed constant relative humidity (RH) in the scale housing [19], which is sealed with a gloverbox quality putty to minimize leaks.

Labview records the suspension's mass (m), the temperature (T), and the relative humidity (RH) during the whole drying process. For all the experiments, the temperature was $T = 25 \pm 2 \text{ }^\circ\text{C}$. Hence, RH controls the drying rate. Initially, the mass decreases linearly with time (i.e. $dm/dt \sim \text{constant}$) for a prescribed RH, as reported in the literature [20]. Transforming the mass loss into the evaporation rate, \dot{E}_0 ($\dot{E}_0 \equiv (dm/dt)/(\rho_{water}\pi R^2)$), one finds that \dot{E}_0 decreases linearly with the applied RH. Experiments verified that \dot{E}_0 for pure water and Ludox are equal barring the same drying conditions. Modulating RH from 10% to 95% makes \dot{E}_0 vary over a decade from $\dot{E}_0 = 36.1 \pm 0.7 \text{ nm/s}$ to $\dot{E}_0 = 3.4 \pm 0.2 \text{ nm/s}$.

After the constant evaporation rate regime, a falling-rate regime occurs where the evaporation slows down and eventually stops. Once dm/dt is negligible (after ~ 50 h for RH = 10% and after ~ 700 h for RH = 95%), the sample is brought back to ambient humidity. Postmortem observations reported hereafter are independent of the returning protocol: An abrupt return (obtained by opening the housing) or progressive one (obtained by bringing the RH inside the box back to ambient humidity in a stepwise process as slow as 10% humidity steps every 12h) does not alter the solid porosity nor the particles surface arrangement. It has also been checked that increasing the Petri dish radius from 2.5 to 6.8cm, keeping the same initial thickness, does not alter experimental observations reporter henceforth.

A camera (USB2 uEye from IDS imaging), located above the scale's enclosure, images the evolution of the sample. As the solvent evaporates, the particles move closer to one another forming a solid network which progressively retracts. The rigid substrate (i.e. the Petri dish) hampers the retraction leading to tensile stresses which cause the layer to fracture (Fig. 1). These cracks begin to appear just as the evaporation rate enters the

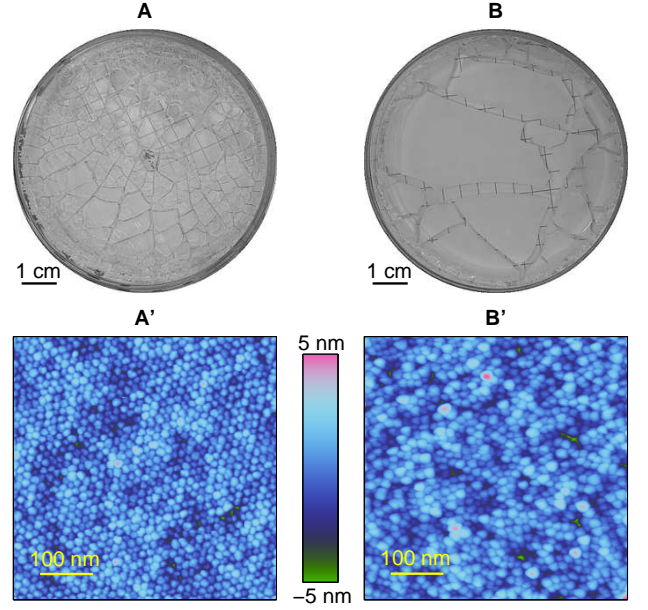


Fig. 1: Effect of the drying rate at the macro and nanoscales. Top: Final crack patterns observed after evaporation for RH = 10% (A) and RH = 95% (B). Bottom: Topographical AFM images depicting the colloid arrangement at the surface of the layers for RH = 10% (A') and RH = 95% (B'). The scan size is $500 \times 500 \text{ nm}^2$. The out-of-plane height ranges over 10 nm which is less than a colloid diameter.

falling-rate regime. This point also coincides with the time at which the meniscus forms at the top of the particles layer, yielding a decreasing capillary pressure in the pores [8, 15, 21]. A Bruker Dimension Icon Atomic Force Microscope (AFM) images the particle arrangement on the evaporation surface of the fragmented morsels. The AFM records topographical images in Tapping mode using a Bruker RTESPA tip (radius of curvature ~ 8 nm). The scan sizes are $500 \times 500 \text{ nm}^2$ with a resolution of 512×512 pixels².

Determining the packing fraction, $\phi = m_{\text{dry}}/(\rho_{\text{silica}}V_T)$, sheds light on the bulk ordering of the solids. The mass of a dried sample (m_{dry}), its total volume (V_T) and the silica density (ρ_{silica}) are estimated as follows. To obtain m_{dry} , the sample are heated at $200 \text{ }^\circ\text{C}$ for more than 3 hours to remove remaining water [22]. To obtain V_T , imbibition and Archimedes' principle are used: After immersion, the pores soak up water and the sample's total volume is obtained via hydrostatic weighing [23]. This process was repeated using ethanol rather than water; the two values of V_T agree within less than 2%. To obtain an accurate estimation of ρ_{silica} , the dilution method is invoked [24, 25]: Several diluted suspensions of Ludox, of densities ρ_{diluted} , are prepared at prescribed mass fractions (ϕ_{diluted}^m). The initial mass fraction ϕ_0^m is estimated by weighing dry residues at $200 \text{ }^\circ\text{C}$ and ρ_{silica} is inferred by linear regression using the rela-

relationship $1 - (\rho_{\text{water}}/\rho_{\text{diluted}}) = \phi_{\text{diluted}}^m (1 - (\rho_{\text{water}}/\rho_{\text{silica}}))$. We found $\rho_{\text{silica}} = 2.26 \text{ g/cm}^3$. Chemically adsorbed water on the nano-particles leads to a possible overestimation of m_{dry} and underestimation of ρ_{silica} , and thus, an overestimation of the packing fraction ϕ . Silica dried at 200°C can retain up to 5 silanol groups per nm^2 [26]. This corresponds to as much as 1.5% residual water in mass, leading to atmost 3.6% overestimation of ϕ .

Results— Figures 1A and B show snapshots of the cracking pattern observed in the Petri dish after drying (i.e. when the mass loss is negligible) for the two extremes. Drying quickly (RH = 10%) leads to a large quantity of small fragments (~ 50 visible fragments in the center of the Petri dish, within a disk of radius $\sim 1.75 \text{ cm}$). On the contrary, the slowest drying rate (RH = 95%) leads to a small quantity of large fragments (5 fragments visible over the same area). Generally, the typical fragment size decreases as \dot{E}_0 increases (or equivalently as RH decreases).

The drying rate also alters the nanoscale arrangement of the colloids at the layer surface as observed via the AFM. An almost 2D crystalline structure forms for the fastest rate (RH = 10%; Fig. 1A'), and an amorphous one for the lowest rate (RH = 95%; Fig. 1B').

Two-dimensional Fourier spectrum conducted on the AFM topographical images shown in Figs. 2A (RH = 10%) and B (RH = 95%) access the differences in particle arrangements on the drying surface. For the fastest evaporation rate, the pattern exhibits a sixfold symmetry characteristic of a hexagonal 2D lattice. On the contrary, no discrete wavenumber can be identified for RH = 95% and only a central circular halo occurs. This is characteristic of an amorphous structure.

The pair correlation function, $g(r)$, represents the probability of finding the center of a particle at a distance r away from a given reference particle. Moreover, it aids in characterizing the long range order of the images. As in standard practice, $g(r)$ is normalized by that of an ideal gas (i.e. non-correlated particle positions) [29]. Fig. 2C depicts the two extreme cases: RH = 10% (top curve with circular points) and RH = 95% (bottom curve with triangular points). For RH = 10%, $g(r)$ presents a well defined sequence of peaks extending more than $12d$ where $d \simeq 15 \text{ nm}$ is the particle diameter acquired from Fig. 1A' and 1B' (note that this value is slightly higher than the spec one). The first several $g(r)$ peaks coincide well with theoretical values for a perfect hexagonal lattice (indicated by dash vertical lines), even if they widen as r increases. This is the signature of a quasi-long range translational order. Conversely, the RH = 95% peaks become barely visible for $r \geq 2d$ indicating a loss of translational order, as expected for amorphous structure.

The orientational order evolution can be investigated via the bond angle order parameter ψ_n defined by

$$\psi_n = \left| \frac{1}{M} \sum_{k=1}^M \frac{1}{N_k} \sum_{l=1}^{N_k} \exp(i n \theta_{kl}) \right|, \quad (1)$$

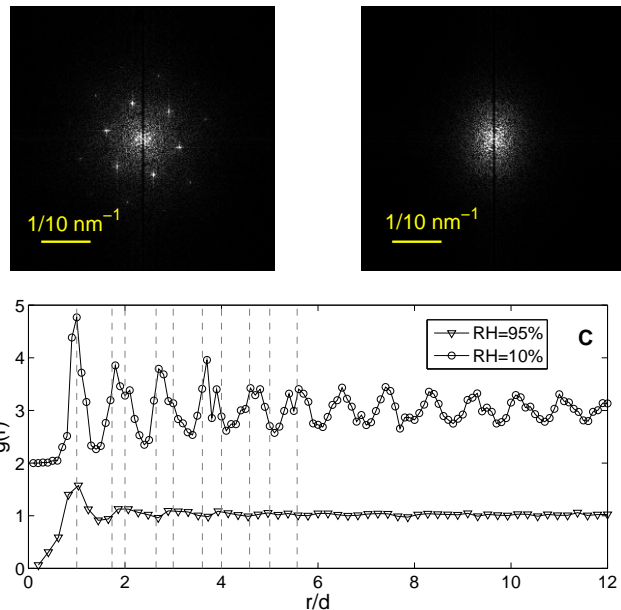


Fig. 2: Fourier spectrum for RH = 10% (A) and RH = 95% (B), obtained by Fast Fourier transforms of the two topographical AFM images presented in Fig. 1. Note the sixfold spot symmetry for RH = 10%, which is the signature of an hexagonal particle arrangement. The spots vanish for RH = 95%, and the Fourier pattern become an isotropic disk, which is the signature of amorphous-like arrangement. Panel C shows the pair correlation function $g(r)$ for these two images. The radius has been normalized by the colloid diameter $d \simeq 15 \text{ nm}$ (as measured from Fig. 1). The vertical dashes lines correspond to the peaks of a perfect hexagonal lattice: $r_1/d = 1$, $r_1/d = \sqrt{3}$, $r_1/d = 2$, $r_1/d = \sqrt{7}$, $r_1/d = 3$, $r_1/d = \sqrt{13}$, $r_1/d = 4$, $r_1/d = \sqrt{21}$, $r_1/d = 5$, $r_1/d = \sqrt{31}$. For sake of clarity, the $g(r)$ obtained for RH = 10% is shifted upwards by two units.

where n is the number of nearest neighbors, M is the number of particles in the AFM image, N_k corresponds to the number of nearest neighbors of particle k , and θ_{kl} is the angle between a fixed direction and the line linking the centers of particles k and l [30]. For the six-fold symmetry visualized in Fig 2A, n should be set to 6. For a perfect hexagonal arrangement $\psi_6 = 1$ while for an amorphous phase $\psi_6 = 0$. The evolution of ψ_6 with \dot{E}_0 (shown in Fig. 3A) exhibits two regimes with a crossover at $\dot{E}_c \sim 10 \text{ nm/s}$. For $\dot{E}_0 \leq \dot{E}_c$, $\psi_6 \simeq 0.05 \ll 1$. This reflects a disordered surface for low evaporation rates. On the other hand when $\dot{E}_0 > \dot{E}_c$, ψ_6 increases linearly with \dot{E}_0 . This is consistent with the observations of the translation order (fig. 2C): Increasing the drying rate yields increasing hexagonal order at the surface. The fact that ψ_6 remains significantly lower than one reveals that even for the fastest drying rate the surfaces are not perfect 2D hexagonal lattices.

The above AFM analysis is restricted to *surface* characterizations of the particle arrangement. In a complementary way, the particle volume fraction ϕ provides in-

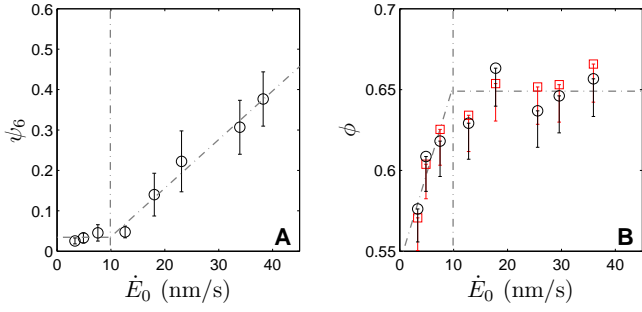


Fig. 3: A: Bond orientational order parameter (ψ_6) of the particle arrangement as a function of the evaporation rate \dot{E}_0 . Error bars indicate the standard deviation over measurements performed on 10 different AFM images. The dashed horizontal and inclined lines correspond to linear fits over the two rate regimes and the vertical line in between indicates the crossover value $\dot{E}_c \simeq 10$ nm/s. B: Overall packing fraction ϕ as a function of \dot{E}_0 measured by immersion of the samples in water (black circles) and ethanol (red squares). The error bars take into account the overestimation of ϕ due to residual water.

formation on the *bulk* particle arrangement. Fig. 3B displays the evolution of ϕ as a function of \dot{E}_0 . Two regimes occur with a crossover $\dot{E}_c \simeq 10$ nm/s. This crossover coincides with that observed in Fig. 3A which is related to the surface orientational order. In the first regime $\dot{E}_0 \leq \dot{E}_c$, ϕ increases linearly with \dot{E}_0 . In the second regime $\dot{E}_0 > \dot{E}_c$, ϕ increase rapidly slows down and even saturates to $\phi \simeq 0.65 \pm 0.02$, which is just above the random close packing value ($\phi_{RCP} = 0.64$ for monodisperse packing) but below the value of a compact structure ($\phi_{HCC/FCC} = 0.74$). This shows that the bulk is, to a large extent, disordered and raises the question to what extent does the order extend into the bulk.

Discussion– The packing structure results from the way the particles bundle. Its dependency on the drying rate is usually rationalized [31,32] via the dimensionless Péclet number: $Pe \equiv h_0 \dot{E}_0 / D_0$ where D_0 is the diffusion coefficient for silica particles. Using Stokes-Einstein relation, one gets $D_0 \simeq 2.04 \times 10^{-11}$ m²/s at $T = 25^\circ\text{C}$. Pe represents the ratio between the convection time for the particles toward the suspension surface and their Brownian diffusion time. A high Pe indicates directional packing implying the solid forms layer by layer. A low Pe means uniform and isotropic bulk compaction.

Checking whether or not Pe controls the packing structure requires two additional experimental runs. The experiments invoke two parameter sets $\{h_0, \text{RH}\}$ at a similar Péclet number well above $Pe = 1.5$ (i.e. $\dot{E}_0 > \dot{E}_c = 10$ nm/s) which should result in a similar surface crystalline arrangement. The first experimental run uses $h_0 = 5$ mm and $\dot{E}_0 = 36.1$ nm/s (RH=10 %) and yields $Pe = 4.4$. The second run uses $h_0 = 10$ mm and $\dot{E}_0 = 26$ nm/s (RH = 50 %) and yields $Pe = 6.4$ (slightly larger). The first run exhibits an ordered surface arrange-

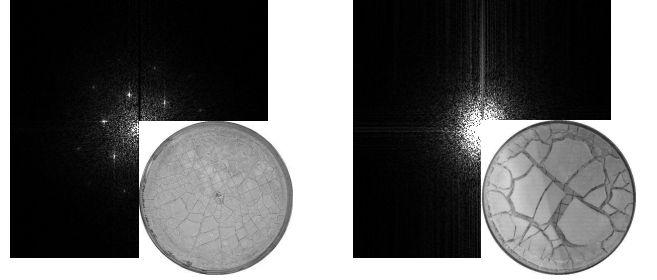
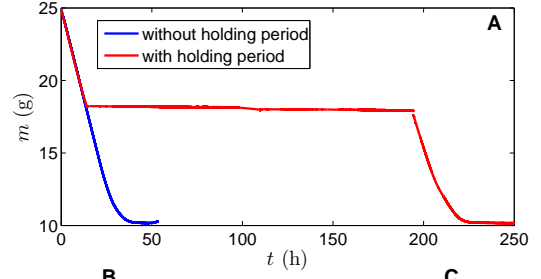


Fig. 4: Influence of the drying protocol on the dried layer's nanostructure and ability to crack. A: Time evolution of the Ludox mass m within the Petri dish in a drying experiment performed at RH = 10% with and without a holding period. This holding period occurs when the solid fraction is $\phi = 0.335$ and lasts 180 h. B: Fourier spectrum of a typical topographical AFM images (main panel) and crack pattern (inset) observed after evaporation without the holding period (this figure is a composite of Fig. 2A and Fig. 1A, and facilitates the comparison with panel C). C: Fourier spectrum of a typical topographical AFM images (main panel) and crack pattern (inset) observed after evaporation with the holding period.

ment, while a disordered arrangement occurs for the second one. This demonstrates that Pe is not the relevant parameter driving the particle packing.

Beyond the competition between convection and diffusion, the formation of aggregates within the suspension could be an important mechanism driving the crystalline-to-amorphous arrangement in the solid. If this is the scenario, then slower drying favors the formation of aggregates and subsequently the formation of an amorphous solid. Testing this hypothesis calls for two new experimental runs. Both runs invoke RH = 10% and $h_0 = 5$ mm (Fig. 4). The first experimental run uses, as previously, a continuous drying scheme and gives a crystalline surface arrangement. The second experimental run invokes an 180 hours holding period (i.e. $\dot{E}_0 \sim 0$ g/h for 180 hours) launched when $m = 18.2$ g. During this holding period $\phi = 0.335$, the suspension is concentrated enough to favor agglomerate formation. After the holding period, the sample resumes drying with RH = 10% until $\dot{E}_0 \sim 0$. The solid layer exhibits an amorphous arrangement on the drying surface. Thus, it is conjectured that long drying protocols (including holding periods) favor the agglomerate formation which subsequently alters the nanostructure packing (amorphous instead of crystalline as imaged with

the AFM). At the macroscale, the crack patterns of the second run are larger than the ones obtained, without a holding period, although they appear under the same external drying conditions. This implies that slowing down the drying enhances the macroscopic resistance to fracture by changing the way the particles bundle.

Conjecturing agglomerate formation sheds new light on the ψ_6 vs. \dot{E}_0 and ϕ vs. \dot{E}_0 curves displayed in fig. 3. At first glance, one expects a correlated increase of ψ_6 and ϕ with \dot{E}_0 as they both increase with the packing order, yet this does not occur. Still, the curves can be understood assuming that \dot{E}_c is a critical value below which every particle of the suspension is part of an aggregate before solification. Subsequently:

- For $\dot{E}_0 < \dot{E}_c$, the packing has to be amorphous and $\Psi_6 \sim 0$. Increasing \dot{E}_0 causes the mean size of the aggregates to decrease, hence the mean size of the pores decreases, thus increasing the packing fraction (i.e. ϕ increases).
- For $\dot{E}_0 > \dot{E}_c$, the number of isolated particles increases with \dot{E}_0 , making it more likely to have crystalline zones, at least at the surface, where surface tension possibly drives the ordering [33].

As a result, ψ_6 increases with \dot{E}_0 . Additionally, isolated particles fill the space between the remaining aggregates. This explains the slower increase, and even the saturation of ϕ .

In conclusion, this paper demonstrates that the evaporation rate, during drying of a colloidal layer, modifies not only the drying kinetics but also the final dried material: Increasing the rate favors more ordered arrangements and denser packing. Explaining this scenario requires more than the Peclet number and its simple balance between advection and diffusion. Studies herein suggest that aggregation as a result of the drying protocol plays a significant role in the formation of the solid. This may be of practical interest in the design of colloidal drying processes to obtain tunable and well-controlled 3D nanoparticle self-assemblies (e.g. crystals, modulated porosity solids...) over large dimensions, with innovative photonics and biotechnology applications [34]. This rate dependent solid formation has a consequence on the fracture behavior: The resistance to fracture increases by decreasing the drying rate or by introducing a holding period at a wisely chosen time. Work in progress aims at quantifying how the material properties at the macroscale, notably its elastic modulus and toughness, emerge from the nanostructure and its history of formation.

The co-authors would like to thank F. Doumenc, B. Guerrier, L.-T. Lee, P. Reis, J.-B. Salmon for fruitful discussions, A. Aubertin for the controlled drying chamber and for C. Wiertel-Gasquet for technical support in

Labview. This research is supported by Triangle de la Physique (RTRA), Ile-de-France (C’Nano and ISC-PIF) and Investissements d’Avenir of LabEx PALM (ANR-10-LABX-0039-PALM) and LabEx LaSIPS (ANR-10-LABX-0040-LaSIPS).

REFERENCES

- [1] XU P., MUJUMDAR A. S. and YU B., *Drying technology*, **27** (2009) 636.
- [2] ROUTH A. F., *Reports on Progress in Physics*, **76** (2013) 046603.
- [3] KEDDIE J. and ROUTH A. F., *Fundamentals of Latex Film Formation* (Springer) 2010.
- [4] MARÍN A. G., GELDERBLOM H., LOHSE D. and SNOEIJER J. H., *Phys. Rev. Lett.*, **107** (2011) 085502.
- [5] MERLIN A., SALMON J.-B. and LENG J., *Soft Matter*, **8** (2012) 3526.
- [6] YETHIRAJ A. and VAN BLAADEREN A., *Nature*, **421** (2003) 513.
- [7] EDITORIAL, *Nature Materials*, **14** (2015) 1.
- [8] LAZARUS V. and PAUCHARD L., *Soft Matter*, **7** (2011) 2552.
- [9] PAUCHARD L., PARISSÉ F. and ALLAIN C., *Phys. Rev. E*, **59** (1999) 3737.
- [10] GROISMAN A. and KAPLAN E., *Europhysics Letters*, **25** (1994) 415.
- [11] CADDOCK B. and HULL D., *Journal of materials science*, **37** (2002) 825.
- [12] GAUTHIER G., LAZARUS V. and PAUCHARD L., *EPL*, **89** (2010) 26002.
- [13] BOULOGNE, F., GIORGIUTTI-DAUPHINÉ, F. and PAUCHARD, L., *Oil Gas Sci. Technol. – Rev. IFP Energies nouvelles*, **69** (2014) 397.
- [14] SCHERER G. W., *Journal of Non-Crystalline Solids*, **109** (1989) 171.
- [15] CHEKCHAKI M. and LAZARUS V., *Transport in Porous Media*, **100** (2013) 143.
- [16] GIORGIUTTI-DAUPHINÉ F. and PAUCHARD L., *The European Physical Journal E*, **37** (2014) 1.
- [17] Ludox HS-40, Batch number BCBK7778V, <http://www.sigmaaldrich.com/Graphics/COFAInfo/flu ka/pdf/PDF474093.pdf>, accessed : November 12, 2015.
- [18] DEEGAN R. D., BAKAJIN O., DUPONT T. F., HUBER G., NAGEL S. R. and WITTEN T. A., *Nature*, **389** (1997) 827.
- [19] BODIGUEL H., DOUMENC F. and GUERRIER B., *Langmuir*, **26** (2010) 10758.
- [20] BRINKER C. J. and SCHERER G. W., *Sol-gel science : the physics and chemistry of sol-gel processing* (Academic Press, Boston) 1990.
- [21] DUFRESNE E. R., CORWIN E. I., GREENBLATT N. A., ASHMORE J., WANG D. Y., DINSMORE A. D., CHENG J. X., XIE X. S., HUTCHINSON J. W. and WEITZ D. A., *Physical Review Letters*, **91** (2003) 224501.
- [22] ILER R. K., *The Chemistry of Silica: Solubility, Polymerization, Colloid and Surface Properties and Biochemistry of Silica* (Wiley) 1979.
- [23] BARLET M., KERRACHE A., DELAYE J.-M. and ROUNTREE C. L., *Journal of Non-Crystalline Solids*, **382** (2013) 32 .

- [24] FINSY R., MOREELS E., BOTTGER A. and LEKKERKERKER H., *Journal of Chemical Physics*, **82** (1985) 3812.
- [25] ADAMCZYK Z., JACHIMSKA B. and KOLASISKA M., *Journal of Colloid and Interface Science*, **273** (2004) 668 .
- [26] ZHURAVLEV L., *Colloids and Surfaces A*, **74** (1993) 71 .
- [27] GIUSEPPE E., DAVAILLE A., MITTELSTAEDT E. and FRANÇOIS M., *Rheologica Acta*, **51** (2012) 451.
- [28] GAUTHIER G., LAZARUS V. and PAUCHARD L., *Langmuir*, **23** (2007) 4715.
- [29] HANSEN J. P. and MCDONALD I. R., *Theory of Simple Liquids (3rd ed.)* (Academic Press) 2005.
- [30] REIS P. M., INGALE R. A. and SHATTUCK M. D., *Physical Review Letters*, **96** (2006) 258001.
- [31] ROUTH A. F. and ZIMMERMAN W. B., *Chem. Engineering Science*, **59** (2004) 2961.
- [32] EKANAYAKE P., MCDONALD P. J. and KEDDIE J. L., *The European Physical Journal Special Topics*, **166** (2009) 21.
- [33] BIGIONI T. P., LIN X.-M., NGUYEN T. T., CORWIN E. I., WITTEN THOMAS. A. and JAEGER H. M., *Nature Materials*, **5** (2006) 265.
- [34] DUTTA J. and HOFMANN H., *Self organization of colloidal nanoparticles* in *Encyclopedia of Nanoscience & Nanotechnology*, edited by NALWA H. S., Vol. 9 (American Scientific Publishers, USA) 2004 pp. 617–640.