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245 - A Highly Sensitive Glucose Electrode Using Glucose Oxidase Collagen Film *

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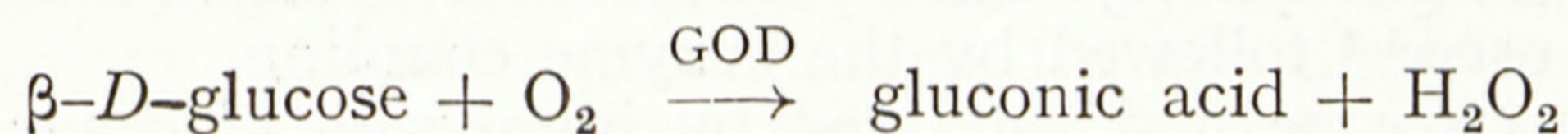
Summary

Industrially reconstituted collagen films have shown excellent properties for β -D-glucose oxidase coupling. Associated with a platinum anode for amperometric detection of hydrogen peroxide, these enzymatic films form a very simple and easy to handle glucose electrode; this device presents a very high sensitivity (*ca.* 10^{-8} M) giving responses proportional to glucose concentration over 5 orders of magnitude.

Introduction

The association of an enzyme in a soluble form with an electrochemical sensor was first reported by CLARK and LYONS, in 1962.¹ Since this date, the obtainment of carrier-bound enzymes has permitted the design of enzyme electrodes. In most cases, enzymes are trapped in gels surrounding the sensor² leading to systems which are difficult to handle.

In the present work, an enzymatic membrane was prepared from reconstituted calf skin collagen after acyl-azide activation and a coupling process giving a surface binding of enzymes of different classes.³⁻⁵ A glucose electrode using β -D-glucose oxidase (GOD) membranes was developed, using an amperometric method with a platinum electrode to detect hydrogen peroxide, which is a product of the enzymatic oxidation of glucose according to the reaction:



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The potential was fixed at +650 mV *vs* Ag|AgCl, KCl 0.1 *M* and the anodic current was recorded. A stationary response was then obtained allowing the measurement of very low glucose concentrations. In a more sophisticated device, a second electrode involving a non-enzymatic membrane was used to compensate for the detection of other electroactive molecules and to enhance the selectivity of the glucose electrode.^{6,7}

Experimental

Instrumentation

The glucose electrode consisted of a modified gas electrode in which the pH detector was replaced by a platinum disk and the usual teflon film by a collagen membrane. In a differential device, electrode 1 was mounted with a β -D-glucose oxidase collagen membrane and electrode 2 with a non-enzymatic one.

Electronics were made by SOLEA TACUSSEL: current outputs of both working electrodes were first subtracted (Deltapol) and then twice differentiated (Derivol) with a time-base of one second (GSTP); thus different current *vs* time curves were available and usually recorded after a glucose pulse (SOLEA TACUSSEL EPL 2 with TV 11 GD plug-in unit, and three traces LINEAR 395 recorders).

Unless otherwise mentioned, the temperature of the solutions were carefully thermostated to 30.0 ± 0.1 °C (COLORA cryothermostat WK 5 DS).

Solutions and reagents

Insoluble films of highly polymerized reconstituted collagen (20 cm wide) were a gift of the Centre Technique du Cuir, Lyon (France); their thickness is about 0.1 mm in a dry state and 0.3–0.5 mm when swollen. They do not need to be tanned and can be stored several years without damage.³⁻⁵

Unless otherwise mentioned, all chemicals were reagent grade. The stock solutions of 0.1 *M* glucose were allowed to mutarotate at room temperature at least 3 hours before using and were stored at 4 °C. Both electrodes were filled with and dipped into 0.2 *M* acetate buffer, 0.1 *M* KCl solutions, pH 5.6.

Glucose oxidase binding on collagen membranes

The mild general acyl azide procedure for collagen membranes activation was used³⁻⁴ followed by the enzyme coupling.

Carboxyls were first esterified by immersion of crude membranes in a methanol/0.2 *M* hydrochloric acid solution for at least 72 hours, then treated overnight by 1% hydrazine and soaked at 4 °C for 3 minutes in 0.5 *M* NaNO₂—0.3 *M* HCl mixture just before coupling. Thorough washings were performed between each step and at the end of the

activation process avoiding contact between reagents and enzyme solutions.

Activated films of $2 \times 1.5 \text{ cm}^2$ were dipped (2 hours at 4°C) into 1.5 cm^3 of borate buffer at pH 9 containing 2.5 mg β -D-glucose oxydase (GOD, E.C. 1.1.3.4, BOEHRINGER lyophilizate, grade I). Excess of soluble glucose oxidase was washed away for about 100 minutes in 1 M KCl and surface activity was in the range $30\text{--}60 \text{ nmol. min}^{-1} \text{ cm}^{-2}$. Enzyme collagen membranes were stored in 0.2 M acetate buffer, 0.1 M KCl (pH 5.6).

Procedure

Both electrodes were allowed to equilibrate in the buffer solution for 15 to 30 minutes after stepping the potential of the platinum disks to $+650 \text{ mV vs Ag|AgCl, } 0.1 \text{ M KCl}$. This potential corresponds to a diffusion-limited current for H_2O_2 oxidation. Calibrations were performed by successive microadditions (10 to 50 mm^3) of stock solutions of 10^{-6} to 10^{-1} M glucose to 10 to 20 cm^3 of buffer. The stationary response was the variation of the steady state values of $I_1 - I_2$ when a sample of a glucose containing solution was added and the dynamic response was the height of the peak of the first derivative $d(I_1 - I_2)/dt$. When successive additions were performed in the same solution, a current offset was used.

Results and discussion

When the enzyme electrode is immersed in a medium in which a pulse of glucose is added a steady state takes place after 2–3 minutes, as shown on Fig. 1, and the value of the anodic current reaches a plateau. The variation of intensity is directly dependent on the glucose concentration in the assay: this is the *stationary response*. On the other hand, the *dynamic response* is measured by the height of the peak obtained after 30–50 s by recording the first derivative of this current.

The lowest glucose concentration detected under these conditions is less than 10^{-8} M (Fig. 2). For values higher than 10^{-2} M , the responses become independent of the glucose concentration. The concentrations, which can be determined, range between *ca.* 10^{-8} and 10^{-2} M *i.e.* over 6 orders of magnitude and the linearity of the calibration curve is obtained between $(3\text{--}5) \times 10^{-8}$ and $(3\text{--}5) \times 10^{-3} \text{ M}$ *i.e.* over 5 orders of magnitude (Fig. 2).

In a typical experiment, the calibration curves remained linear even after 40 hours operation at 30°C and 250 days storage at 4°C , allowing accurate repeated determinations for 160 micro-assays tested. However, a daily calibration was necessary, because of the slight decreasing of the calibration curve slope: *ca.* 20% after 20 hours operation (cumulated time of measurements) and 20–50% after 40 hours. The accuracy of glucose determinations during a set of experiments was tested by successive additions of the same aliquot of glucose (Fig. 1);

2.5.1

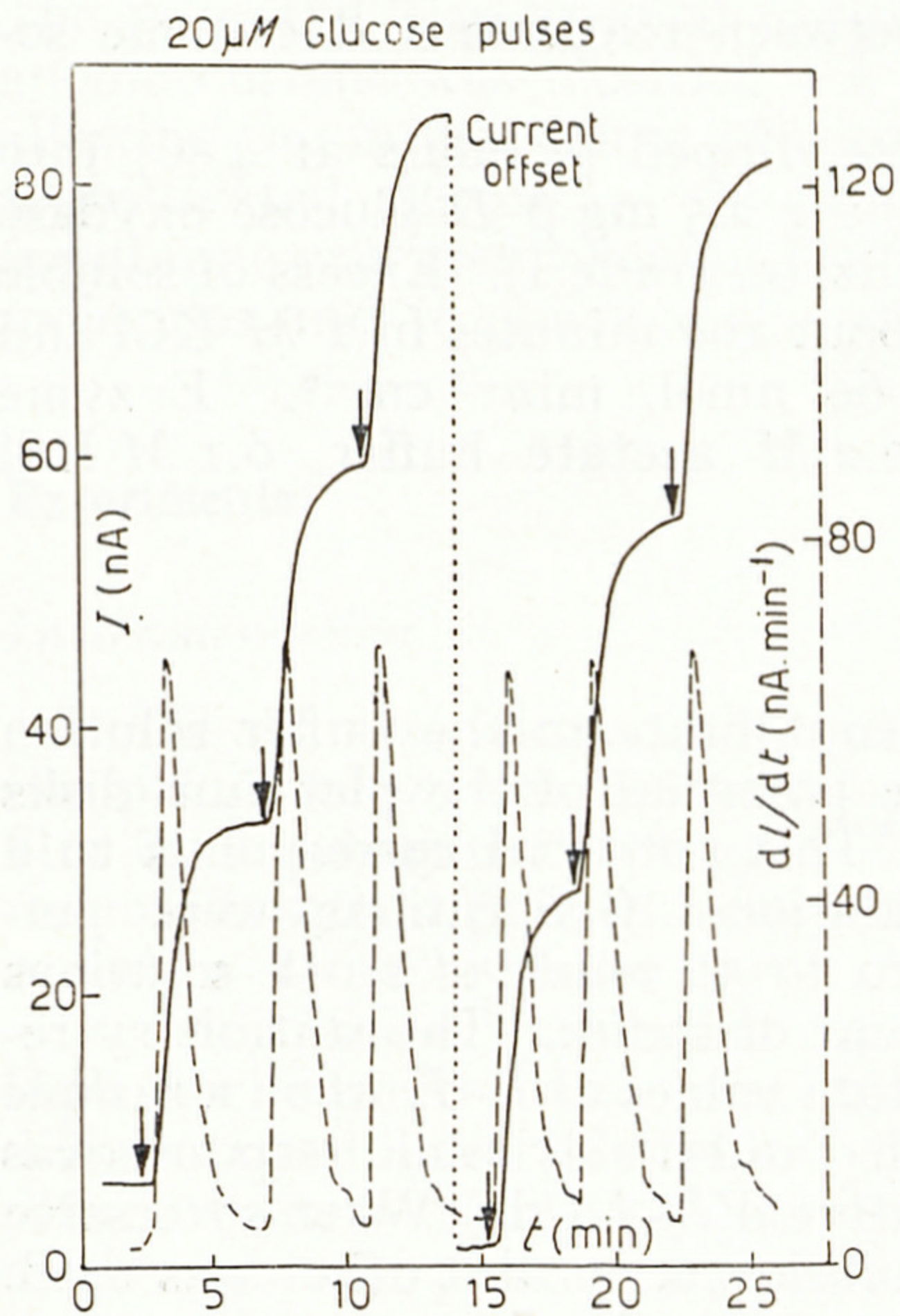
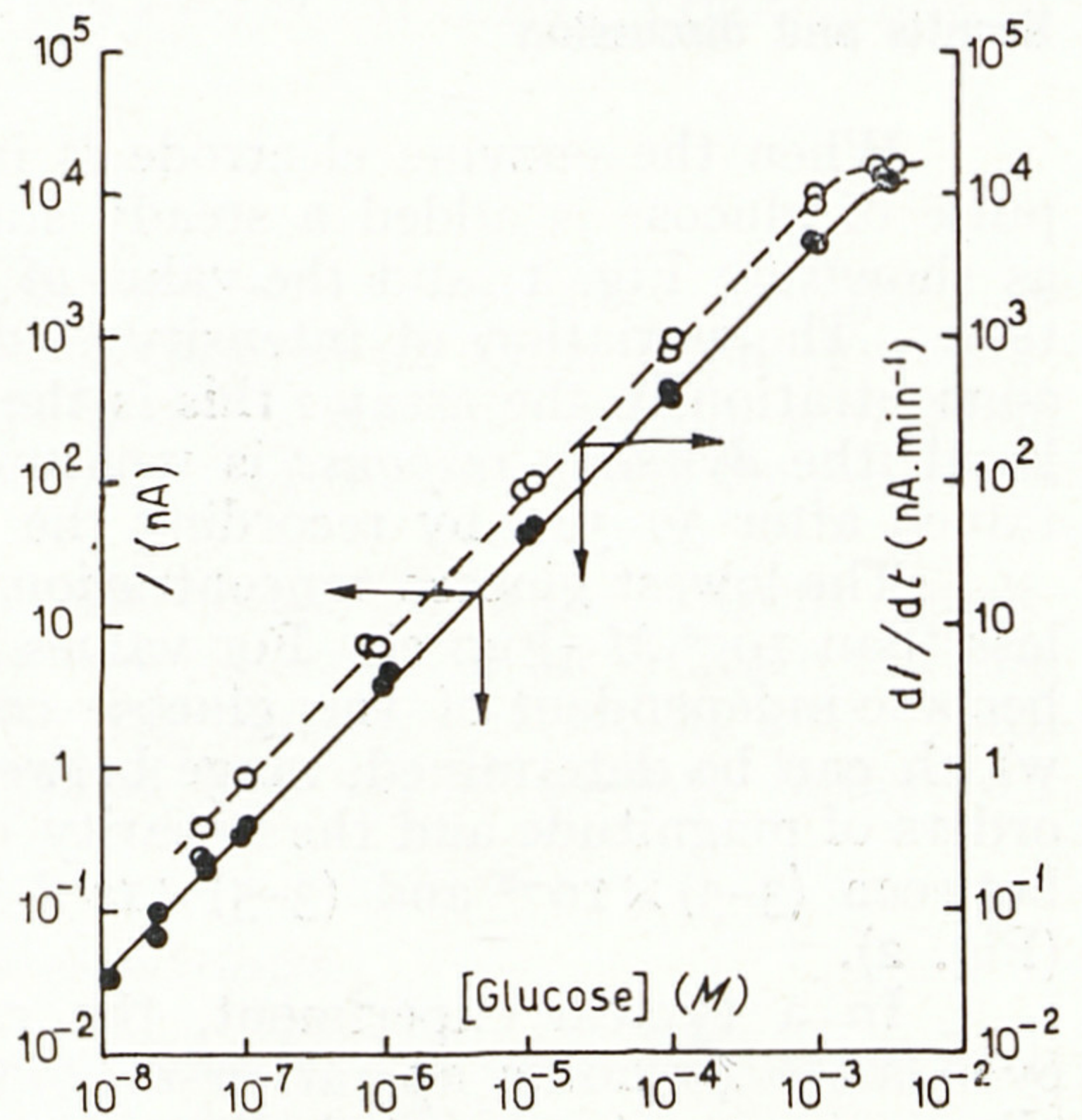


Fig. 1. Electrode responses when successive $20 \mu M$ glucose pulses are added to the buffer solution. Addition of $40 \text{ mm}^3 (\mu l)$ aliquots of 10 mM glucose solution into 20 cm^3 of 0.2 M acetate buffer, 0.1 M KCl (pH 5.6) solution. (—) direct current giving stationary responses, (---) first derivative of this current giving dynamic responses.

2.5.1

Fig. 2. Calibration curves of the glucose electrode. (—) stationary response; (---) dynamic response.



for 13 successive additions of 10^{-5} M glucose, the standard deviation from the mean is usually lower than 2 %.

Glucose oxidase itself is very selective for $\beta\text{-D}$ -glucose; thus enzymatic electrode 1 presents a high selectivity for glucose compared

with usual sugars : selectivity ratios are higher than 2 000 / 1 for fructose, lactose and sucrose. As other species may diffuse through collagen membranes and may be oxidized on platinum at $+0.94$ V (N.H.E.), the use of a compensating non enzymatic electrode 2 eliminates possible interferences of species such as ascorbate, urate, tyrosine or hydrogen peroxide; the selectivity ratio for hydrogen peroxide ranges between 80/1 and 250/1 depending upon experimental parameters (accuracy of the balance of both electrodes and GOD activity of the film).

The use of non-enzymatic electrode 2 is specifically of greatest interest for blood glycemia determinations, I_2 usually reaching 10–50 % of I_1-I_2 . Fig. 3 presents the typical analysis of blood plasma samples after induced glycemia : in this case, a glucose calibration is necessary for each set of 3–4 blood additions. The dotted line, representing the second derivative of the current I_1-I_2 , has been successfully used for monitoring a printing device of the peak of the first derivative, *i.e.* the dynamic response.⁷

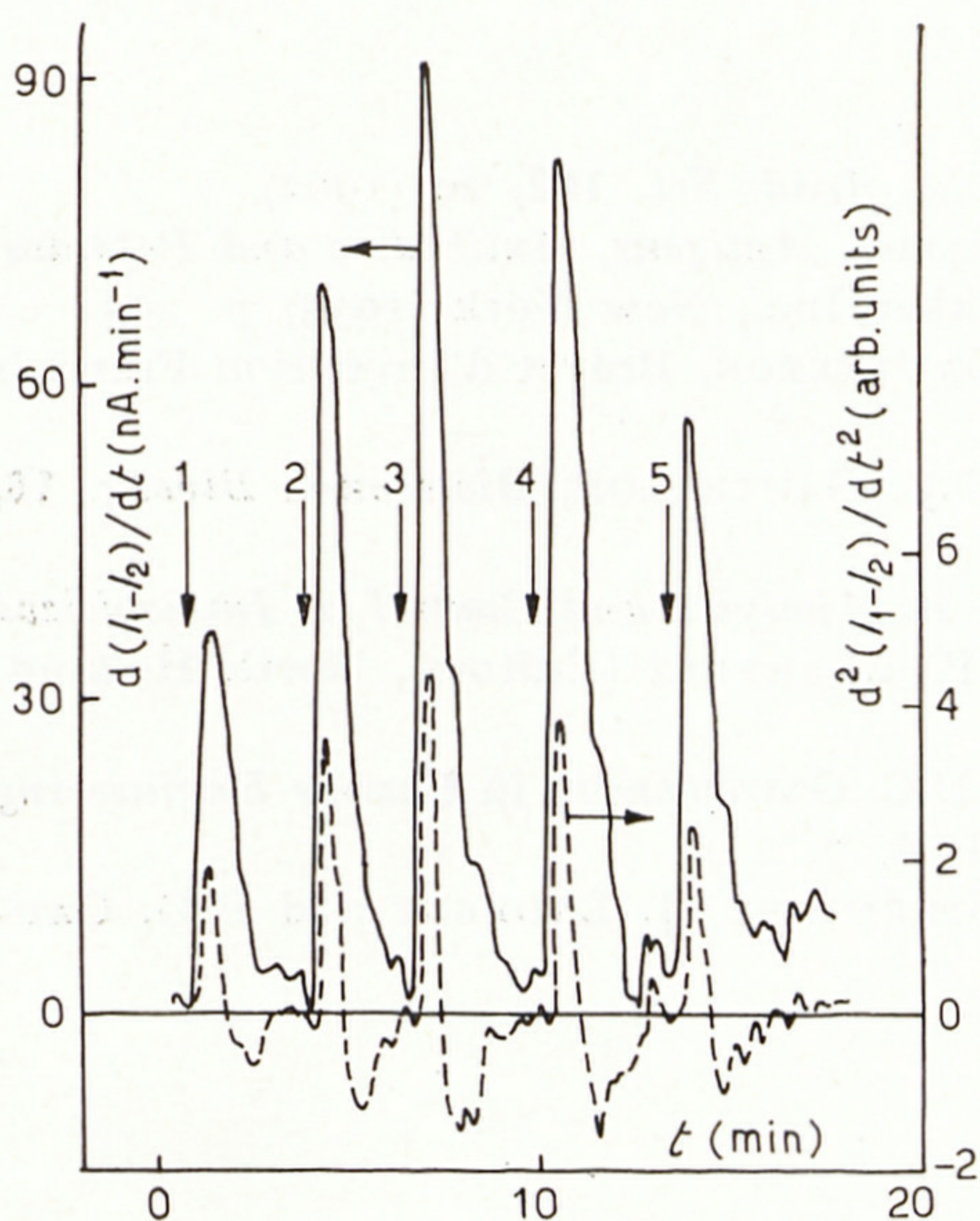


Fig. 3. Induced glycemia. (—) first and (---) second derivative of the current after injection of 50 mm^3 (μl) non deproteinized plasma samples. to 20 cm^3 buffer (1) calibration of the electrode with a $10 \mu\text{M}$ glucose pulse, (2–5) glucose determination in plasma samples taken after 0 (2), 30 (3), 60 (4) and 120 minutes (5) after induced glycemia. Corresponding glucose content were respectively 1.1, 1.7, 1.4 and 1.0 g/l.

The glucose electrode may be used in a large temperature range, from 15 to 40°C . As both responses are very sensitive to temperature (about 4–5 % / $^\circ\text{C}$ at 30°C) it is necessary to carefully thermostate the solutions in contact with both electrodes.

Conclusions

Industrially reconstituted collagen films were found suitable for glucose oxidase immobilization; the enzymatic activity was maintained and its stability enhanced. A membrane loading of 50 to 100 mU per membrane (1 cm diameter) was sufficient to obtain a very sensitive glucose electrode when associated with amperometric hydrogen peroxide detection. Stationary and dynamic responses of this glucose electrode was proportionnal to glucose concentration from $3-5 \times 10^{-8}$ to $3-5 \times 10^{-3}$ M. Furthermore, this sensor was used for whole blood samples analysis in induced glycemia.

Acknowledgements

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