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Photochemistry of glycolaldehyde in cryogenic matrices

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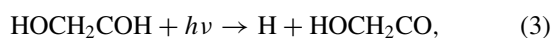
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The photochemistry of glycolaldehyde (GA) upon irradiation at 266 nm is investigated in argon, nitrogen, neon, and para-hydrogen matrices by IR spectroscopy. Isomerization and fragmentation processes are found to compete. The hydrogen-bonded Cis-Cis form of GA is transformed mainly to the open Trans-Trans conformer and to CO and CH₃OH fragments and their mixed complexes. Different photo-induced behaviours appear depending on the matrix. In nitrogen, small amounts of Trans-Gauche and Trans-Trans conformers are detected after deposition and grow together upon irradiation. The Trans-Gauche conformer is characterized for the first time. In para-hydrogen due to a weaker cage effect additional H₂CO and HCO fragments are seen. Calculations of the potential energy surfaces of S₀, S₁, and T₁ states – to analyse the torsional deformations which are involved in the isomerization process – and a kinetic analysis are presented to investigate the different relaxation pathways of GA. Fragmentation of GA under UV irradiation through the CO+CH₃OH molecular channel is a minor process, as in the gas phase. © 2014 AIP Publishing LLC. [<http://dx.doi.org/10.1063/1.4881605>]

I. INTRODUCTION

Because they are widely present in the atmosphere, the fate of carbonyl compounds upon ultraviolet radiation has attracted numerous studies.¹ When subject to electronic excitation, carbonyl compounds can dissociate to highly reactive species and undergo many cascade reactions by combining to radicals. Glycolaldehyde (HOCH₂CHO, GA) is one such product (Fig. 1). Its photochemistry in the gas phase was extensively studied. Photo-dissociation in atmospheric conditions was found to yield many end-products such as formaldehyde (H₂CO), methanol (CH₃OH), carbon monoxide (CO), glyoxal (CHOCHO), formic acid (HCOOH), and carbon dioxide (CO₂).^{2,3} It is only recently that the UV photolysis of pure GA in the gas phase was studied.⁴ Photolysis end-products are similar to those previously observed in air except that CO₂ is missing and intermediate products, i.e., hydroxyl (OH) and formyl (HCO) radicals were detected using cavity ring down spectroscopy. Four channels were considered in these previous works as first steps of photo-dissociation, of which three involve the formation of radicals:



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In the gas phase, the major relaxation pathway is the α -CC bond cleavage (Norrish type I reaction) (channel 1). In atmospheric conditions and depending on the wavelengths of the irradiation,³ it accounts for 65%–80% of the end-products. The formation of CH₃OH molecules reveals that the direct molecular elimination channel (4) also occurs (~10%) while channel 3 accounts for less than 15%. In Zhu *et al.* experiments where GA is photolysed without air, the molecular channel (4) is also observed but is much less efficient than channel 1 which is predominant.⁴ The ratios of end-product molecules relative to the GA photolysed at 308 nm are estimated to 41%, 60%, 10%, 7%, and 7% for H₂CO, CO, (CHO)₂, HCOOH, and CH₃OH, respectively. The radical pathways are supported by the recent theoretical study of Cui *et al.* who calculated the potential energy profiles in the S₀ ground state, the excited S₁ singlet, and T₁ triplet states of GA considering the three C–C, C–H, and C–O bond cleavage channels upon excitation of the most stable form.⁵ They propose that after excitation to S₁, GA relaxes to T₁ through intersystem crossing due to a high S₁/T₁ spin-orbit coupling (23.4 cm⁻¹). Because the T₁ state is adiabatically correlated with the radicals' ground-state along these three cleavage channels, GA dissociates to radicals. They also find that the predominant channel is the production of HCO (channel 1), and that channel 3 is nearly closed, supporting previous studies.²⁻⁴

GA molecule mainly appears as a pseudo-cyclic form named CC (Cis-Cis) with an intramolecular H-bond. According to theory, this CC form is the most stable conformer but three “open” conformers, i.e., with no H-bond, are also stable, lying at higher energies: TT, TG, and CT forms⁵⁻⁸ (Fig. 1). The CC geometry refers to the *cis* configurations of the CO/C=O and OH/CC groups with respect to the C–C

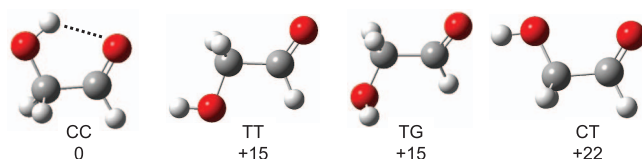


FIG. 1. The four stable conformers of GA: CC, TT, TG, and CT. Relative energies are given in kJ/mol (B3LYP/def2-TZVPP method, this study). Carbon atoms are in grey, oxygen in red, and hydrogen in white. The dotted line represents the internal hydrogen-bond of CC.

and C–O bonds, respectively. The TT and TG forms display Trans-Trans and Trans-Gauche configurations and are found nearly isoenergetic, TT being slightly lower in energy.⁶ The CT (Cis-Trans) conformer is the highest in energy and differs from CC by the rotation of the OH group.

This study aims at investigating the photochemistry of GA in different cryogenic solids, spanning from the classical matrices (nitrogen, argon, and neon) to the softer parahydrogen medium. The behaviour of GA upon irradiation at 266 nm may differ from that in the gas phase. The photochemistry of carbonyl compounds in cryogenic solids often leads to different end-products or complexes due to the matrix cage effect and the recombination of geminate precursors. Furthermore, an earlier study evidenced the formation of the TT conformer from the CC form upon broadband IR irradiation in rare gas matrices.⁹ Because such environments make it possible to trap metastable species such as intermediate products or higher energy conformers, we seek to document the elementary steps of the photolysis of GA and investigate the effect in a confined environment.

The article is organised as follows: we describe the experimental and theoretical methods used in this work; results are presented in Sec. III which reports spectroscopic data before and after UV irradiation; a kinetic analysis of the photo-products and excited state calculations considering the photoisomerization channel for the different conformers of GA are also described; the discussion in Sec. IV focuses on the matrix host effects and on relaxation channels in matrices.

II. METHODOLOGIES

A. Experimental

Three different cryogenic setups were used for matrix isolation experiments, of which two cryostats based at ISMO (Orsay, France) were dedicated to spectroscopy in the mid-IR region for argon (Ar) and nitrogen (N₂) for one setup, and for neon (Ne) and para-hydrogen (pH₂) matrices for another setup. Details about the setups at ISMO are described in earlier works.^{10,11} Matrices were deposited onto CaF₂ window kept at 17 K for Ar and 20 K for N₂ in a Displex closed-cycle cryostat. Experiments in pH₂ were run in a helium bath cryostat coupled with a converter filled with a catalyst for ortho/para conversion.¹¹ The pH₂ matrices were deposited directly from the converter onto a diamond window kept at 4.2 K. For Ne matrices, the same setup as for pH₂ was used; the carrier gas was first passed through the converter at 50 K then condensed onto the deposition window. A small immer-

sion helium cryostat (IHC-3) hosted at the MAX-Lab (Lund University, Sweden) was used for experiments in the far-IR region for Ne and pH₂ solids. Matrices were deposited on a gold plated OFHC copper mirror cooled to 2.8 K. Due to the heat load from incoming sample, gas temperature increased to 3.6 K during the deposition. When the deposition was stopped, the temperature immediately dropped to 2.8 K. Spectra were obtained at 2.8 K, directly after deposition.

GA powder (Fluka, >98%) is a covalent dimer crystal that dissociates to monomeric species upon heating at 60 °C. GA molecules are entrained by the flow of the carrier gas. The behaviour of GA upon UV excitation is monitored by Fourier-Transform IR spectroscopy. The setups used did not allow precise estimation of GA absolute concentration. A rough estimate for GA/host ratio was ~2/1000. IR spectra were measured with a Nicolet Nexus 670/870 FTIR-spectrometer at 0.25 cm⁻¹ resolution in the 600–5000 cm⁻¹ spectral region and a LN₂-cooled MCT detector in Orsay. A Bruker IFS 120 spectrometer at 1 and 0.1 cm⁻¹ resolutions in the 30–600 cm⁻¹ range using a globar light source equipped with a 6 μm Mylar beamsplitter and LHe-cooled Si bolometer was used for experiments in Lund. An optical 650 cm⁻¹ cutoff filter was placed in front of the bolometer in order to minimise noise in the spectra.

UV irradiation studies were performed at ISMO. The UV sources consisted of two 266 nm-quadrupled Nd:YAG laser (Quantel), one operating at 10 Hz and the other (Brio Quantel) at 20 Hz. The latter was used in experiments in Ne and pH₂. The laser bandwidth was ~0.01 nm and typical laser power used for irradiation was ~10 mW. The size of the laser spot was larger than the deposition window so that the whole matrix sample was irradiated by the UV light. Electronic absorption spectra were recorded with a spectrometer (Jobin-Yvon) equipped with a broadband photomultiplier (Hamamatsu) using a deuterium discharge lamp as light source.

B. Theoretical methods

To explore the S₀, S₁, and T₁ potential energy surfaces (PES), the TURBOMOLE package was used.¹² Geometries of S₀ electronic state for the four conformers of GA were fully optimized using density functional theory (DFT), while time-dependent (TD)-DFT was used to study S₁ and T₁ electronic states. All calculations were performed in C₁ symmetry with the B3LYP exchange-correlation functional and the triple-zeta doubly-polarised def2-TZVPP basis set to ensure a fine convergence of the excited states properties. In addition, the two-dimensional PES defined by the two OCCO and CCOH dihedral angles of S₀, S₁, and T₁ states were computed on a grid with 15° intervals. About 325 calculations were, thus, carried out per PES, from –180° to 180° for OCCO and from –180° to 0° for CCOH. The whole PESs were then reconstructed taking advantage of molecular symmetry. Each point of the grid corresponds to geometries where all other degrees of freedom of the molecule were optimized. As mentioned above a more precise exploration was performed in the regions of energy minima.

Full geometry optimisation and frequencies calculations were run for the structures obtained at energy minima on the

three PESs. Vibrational frequency calculations give no imaginary value and indicate real minima for all structures. In addition, such calculations were performed for complexes of GA with a N_2 molecule in the ground state at the B3LYP/def2-TZVPP level of theory including the D3 dispersion correction with Becke-Johnson damping proposed by Grimme.¹³

III. RESULTS

A. Spectroscopy of as-deposited samples

1. IR spectra

IR spectra of GA in pH_2 , Ne, Ar, and N_2 matrices in selected regions are displayed in Fig. 2 and Fig. S1 in the supplementary material.²⁸ In all matrices, the H-bonded CC conformer is observed in accord with previous studies in rare gas and H_2 solids.^{9,11} Experimental frequencies in pH_2 , Ne, Ar, and N_2 are listed in Table I. They are compared to gas phase values¹⁴ and theoretical calculations.⁷ Gas-to-matrix shifts are mode- and host-dependent; the weakest spectral shifts are in Ne. The OH stretch (νOH) mode is red-shifted ($\sim 3540\text{ cm}^{-1}$) compared to a free OH group due to the intramolecular H-bond. The largest red-shift is measured in

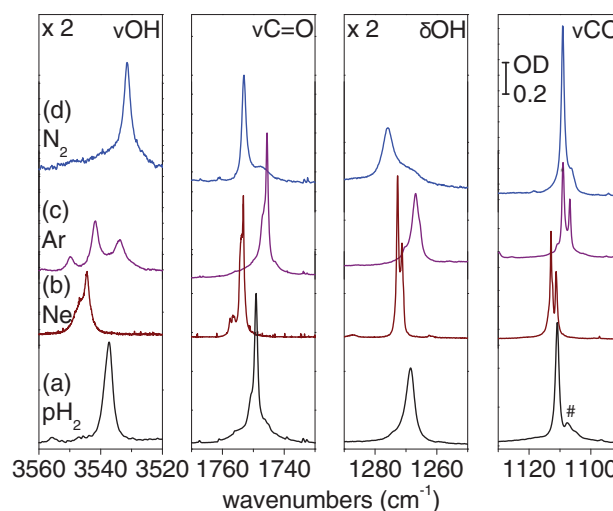


FIG. 2. Vibrational bands selected from IR spectra before irradiation of GA isolated in (from bottom to top): (a) pH_2 , (b) Ne, (c) Ar, (d) N_2 . The # symbol indicates band of dimers of GA.

N_2 and can be correlated with the blue-shift and the relative broadening of the in-plane OH bending mode ($\delta OH \sim 1275\text{ cm}^{-1}$) (Fig. 2(d)): these features are in favour of a

TABLE I. Experimental vibrational frequencies (cm^{-1}) of the CC conformer of GA in para-hydrogen, neon, argon, and nitrogen matrices.^a Only the main absorption is reported in the case of structured bands due to site effects. Gas-phase frequencies are taken from Ref. 14. Calculated anharmonic frequencies (cm^{-1}) and intensities indicated in brackets (km/mol) are taken from Ref. 7. Additional calculated low frequencies are from Ref. 8. Mode numbering is the same as in Ref. 7. The mode description is simplified, especially in the spectral ranges where anharmonic coupling is important.

Mode	CC conformer					CC/B3	Description ^c	
	Gas phase ^b	Experimental						Theoretical
		pH_2	Ne	Ar	N_2			
ν 18	3549	3538	3544.4	3542	3532	3570 (54)	ν OH	
$2^* \nu$ 13		2901		2910	2904.5	2913 (14)		
ν 17	2881	2882	2883.2	2895	2894	2898 (13)	ν_{as} CH_2	
ν 12 + ν 13			2864	2880		2883 (9)		
$2^* \nu$ 12			2854			2816 (14)		
ν 16	2840	2850	2838.6	2854	2868	2874 (11) ^d	ν_s CH_2	
ν 15	2820	2837	2828	2846	2863.4	2854 (19)	ν CH	
$2^* \nu$ 11	2710	2707.3	2707.7	2713.3	2720.9	2707 (6)		
ν 14	1754	1749.7	1753.3	1746.5	1753.7	1761 (108)	ν C=O	
$2^* \nu$ 6	1706	1702.5	1706.4	1697	1714.5	1716 (33)		
ν 13	1456	1424	1431.5	1443	1435.5	1454 (11)	sc. CH_2	
ν 12	1425	1409	1408.4	1400	1414	1421 (27)	w. CH_2 , δ OH	
$2^* \nu$ 4	1378	1398	1393.0	1400	1400.5	1412 (10)		
ν 11	1356	1363.7	1364.1	1365.6	1369.9	1363 (31)	δ CH	
ν 10	1275	1269.6	1272.1	1267.8	1276.8	1273 (41)	δ OH, w. CH_2	
ν 9		1229		1228	1233	1225 (3)	tw. CH_2	
ν 8	1112	1111.1	1112.9	1110.0	1110.1	1120 (80)	ν CO	
ν 6	859	859.4	860.9	855.6	867.6	865 (49)	ν CC	
ν 5	752	750.4	752	750.9	755.4	758 (9)	Backbone	
ν 4						713 (0)	r. CH_2	
ν 3	360.3 ^e	365	360.9	341 ^f		356 ^g	γ OH	
ν 2	282.4 ^e	281.9	287.0	274 ^f		276 ^g	H bond	
ν 1	207.6 ^e	216.5				210 ^g	γ CC	

^aWeak differences appear between this work and our previous results in Ref. 11. The accuracy is better in the present work as obtained by assignment after photolysis.

^bTaken from Ref. 14.

^c δ : bend, γ : torsion, ν : stretch, r: rock, sc: scissor, tw: twist, w: wag, as: asymmetric, s: symmetric.

^dThis theoretical value is assigned to ($2^* \nu$ 4 + ν 13) in Ref. 7 where ν_s CH_2 is found to be involved in more than five vibrational bands.

^eT. Huet, private communication.

^fTaken from Ref. 9.

^gTaken from Ref. 8.

strengthening of the internal H-bond of CC in N₂. GA is too large a molecule to occupy a single substitutional site of the host lattices and site effects can be seen as multiplets in Ar or as doublet structures in Ne (traces c and b of Fig. 2). Strong anharmonic couplings occur between vibrational modes in the CC conformer as illustrated in the 1400 and 2800 cm⁻¹ regions (see Fig. S1²⁸), where band intensity distributions and spectral shifts depend on the nature of the matrix. Assignment in these two regions as given in Table I is based on calculations from Ref. 7. Weak features due to dimeric species or small aggregates of GA can be observed in all matrices (marked with a hash sign in Fig. 2 and Fig. S1²⁸) but their amount varies with the deposition conditions. They are easily distinguished from CC conformer as they do not react with the excitation laser and hence do not perturb the irradiation process.

CC is the only conformer detected in pH₂, Ne, and Ar right after deposition but unexpected bands are also observed in N₂ at 3668, 3644, ~1078, and 1028 cm⁻¹. We carefully checked that they were not induced by the global source of the FTIR spectrometer. Bands at 3668 and 3644 cm⁻¹ are the νOH signatures of two additional conformers of GA adopting an open structure (Fig. S2 in the supplementary material²⁸). They are present at a similar but very weak amount (<2%) and grow upon UV irradiation. These bands belong to the TT and TG conformers. Positions and attributions of these lines were confirmed after UV irradiation of the sample which led to the formation of these conformers.

2. UV absorption spectra

In order to check the electronic excitation wavelength of GA in matrices, we measured its electronic spectrum in Ar and N₂. UV spectra of CC in both matrices strongly resemble that in the gas phase²⁻⁴ or in solution¹⁵ for which the maximum absorption appears near 280 nm. In matrices, the spectrum consists of a broadband (40 nm fwhm) centred at 275 nm (Fig. 3). This suggests that the cryogenic environment does not significantly affect the electronic properties of

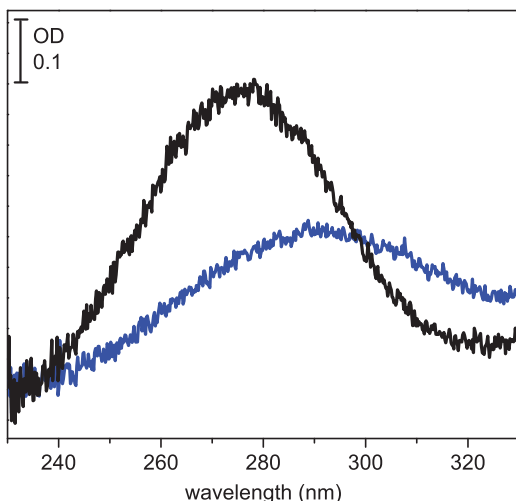


FIG. 3. UV absorption spectra of GA in nitrogen before (black) and after irradiation (blue) at 266 nm.

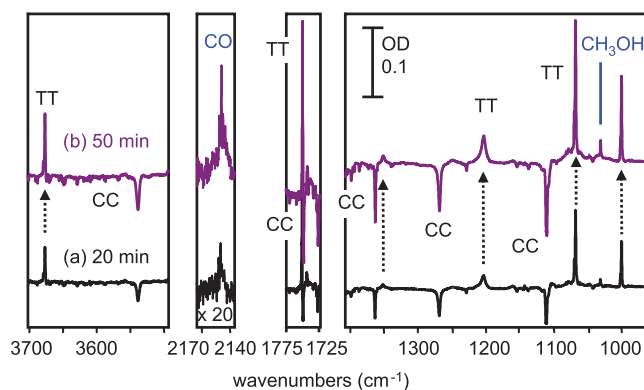


FIG. 4. Difference spectra of GA in pH₂ after (a) 20 min and (b) 50 min of 266 nm irradiation at ~10 mW laser power. Dotted arrows refer to TT (black).

the molecule. Hence, the 266 nm wavelength used for the irradiation experiments is strongly resonant with the absorption of CC in matrices and should efficiently excite the CC conformer.

B. Photochemistry of glycolaldehyde: Products

Products formed upon 266 nm irradiation in cryogenic solids consist of fragments and isomers. The overall behaviour of IR spectra measured at different irradiation times is similar in all solids: the CC bands decrease and new bands due to the photo-products appear. Figures 4–7 show difference spectra in pH₂, Ne, Ar, and N₂, respectively. Negative bands refer to CC. Two different groups were distinguished based on the change in the bands intensity. Bands of open conformers increase upon irradiation then decrease – except in Ne – and bands of dissociation products grow more slowly and continuously along irradiation.

1. Open conformers

Open conformers are observed in all matrices. In pH₂, Ne, and Ar, only the TT conformer is detected (Figs. 4–6). All weak bands that follow the same kinetics as the νCO and νOH TT bands (~1070 and ~3670 cm⁻¹) are attributed to TT.

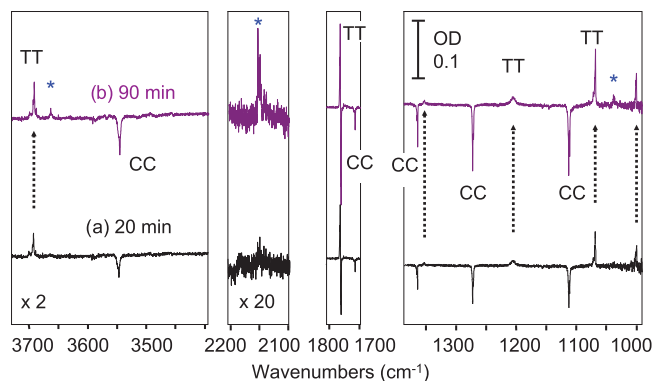


FIG. 5. Difference spectra of GA in Ne after (a) 20 min and (b) 90 min of 266 nm irradiation at ~10 mW laser power. The * sign represents CH₃OH–CO complex. Dotted arrows refer to TT (black).

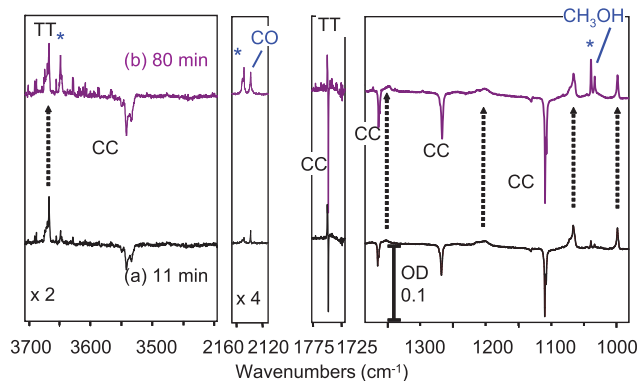


FIG. 6. Difference spectra of GA in Ar after (a) 11 min and (b) 80 min of 266 nm irradiation at ~ 10 mW laser power. The * sign represents $\text{CH}_3\text{OH-CO}$ complex. Dotted arrows refer to TT (black).

Their experimental frequencies are given in Table II. As expected, the νOH frequency is higher than that of CC in agreement with the opening of the internal H-bond. Our assignment of TT agrees well with calculated anharmonic frequencies^{7,8} and earlier studies.⁹

TT conformer is also observed in N_2 (Table II) but several absorptions which do not come from TT grow concomitantly with TT (Fig. 7). They belong to the TG conformer. Bands at 1076, 1051.5, and 1028 cm^{-1} give a satisfactory agreement with the νCO , coupling of CH_2 twist with δCH , and νCC calculated frequencies of TG as reported in Ref. 7 (1079, 1049, and 1023 cm^{-1} respectively, see also Table III). The examination of experimental and calculated νOH spectral shifts between CC and TG ($\Delta\nu_{\text{exp}} = 112 \text{ cm}^{-1}$, $\Delta\nu_{\text{th}} = 133 \text{ cm}^{-1}$) and between TG and TT ($\Delta\nu_{\text{exp}} = 24 \text{ cm}^{-1}$, $\Delta\nu_{\text{th}} = 37 \text{ cm}^{-1}$) allows attributing the 3644 cm^{-1} band to TG νOH mode. Absorptions in the 1400 and 2800 cm^{-1} regions (CH and CH_2 stretches and bends) are not well resolved and are very close to TT bands. Our assignment, also based on Ref. 7, should therefore be considered tentative for these bands (Table III).

Taking advantage of the 266 nm irradiation to increase the concentration of TT conformer in the matrix, we were able to measure its electronic absorption spectrum. In Ar, it consists of a broad band centred at ~ 307 nm. The 266 nm

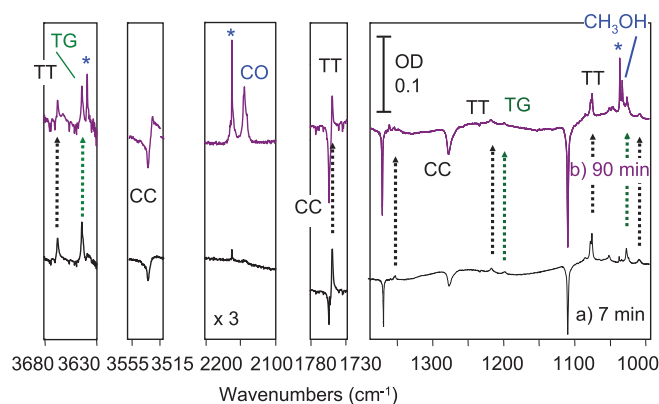


FIG. 7. Difference spectra of GA in N_2 after (a) 7 min and (b) 90 min of 266 nm irradiation at ~ 10 mW laser power. The * sign represents $\text{CH}_3\text{OH-CO}$ complex. Dotted arrows refer to TT (black), TG (green).

TABLE II. Experimental vibrational frequencies (cm^{-1}) of the TT conformer of GA in para-hydrogen, neon, argon, and nitrogen matrices. Only the main absorption is reported in case of structured bands due to site effects. Calculated anharmonic frequencies (cm^{-1}) and intensities indicated in brackets (km/mol) are taken from Ref. 7. Additional calculated low frequency is from Ref. 8. Mode numbering is the same as in Ref. 7.

Mode	TT conformer				CC/B3	Description ^a
	Experimental			Theoretical		
	pH_2	Ne	Ar	N_2		
ν 18	3677.4	3690.7	3668	3668	3740 (36)	ν OH
ν 17	2919	2920	2925	2919	2921 (15)	ν_{as} CH_2
ν 16	2826	2828	2835	2834	2863 (10)	ν_{s} CH_2
ν 15	2815	2809	2810	2822	2823 (35)	ν CH
ν 14	1752.1	1756.2	1750.7	1753.5	1770 (145)	ν C=O
ν 13		1442	1438	1441.5	1458 (12)	sc. CH_2
ν 11	1352.2	1353	1350.4	1352	1354 (4)	δ CH
ν 9	1203.8	1204	1203	1217	1203 (68)	δ OH, w. CH_2
ν 7	1067.7	1068.4	1066.6	1078	1083 (63)	ν CO
ν 6	1000.1	1000.3	998.8	1010	994 (46)	ν CC
ν 4	540 ^b	539.6 ^b	538 ^c		540 (6)	Backbone
ν 3	342.9 ^b	334 ^b	331 ^c		323 ^d	Backbone

^a δ : bend, ν : stretch, sc: scissor, w: wag, as: asymmetric, s: symmetric.

^bRecorded at the MAX-Lab, where $\text{CC} \rightarrow \text{TT}$ isomerization was observed under IR irradiation.

^cTaken from Ref. 9.

^dTaken from Ref. 8.

laser should, thus, weakly excite TT. The same measurement was performed in N_2 . A maximum centred at 307 nm is also observed, suggesting that the electronic transition energy is similar in TT and TG (as predicted by calculations, see below). The electronic spectrum after irradiation in N_2 is displayed in Fig. 3 (blue), also showing the UV spectrum before irradiation (black). The absorption band is clearly red-shifted, corresponding to the decrease of CC concomitantly with the appearance of TT and TG. The UV absorption of open conformers is, thus, obtained from the blue curve after subtraction of the contribution of the CC conformer.

TABLE III. Experimental vibrational frequencies (cm^{-1}) of the TG conformer of GA in nitrogen matrices. Calculated anharmonic frequencies (cm^{-1}) and intensities indicated in brackets (km/mol) are taken from Ref. 7. Mode numbering is the same as in Ref. 7.

Mode	TG conformer		Description ^a
	Experimental	Theoretical	
	N_2	CC/B3	
ν 18	3644	3703 (27)	ν OH
ν 16	2919	2918 (16)	ν_{s} CH_2
ν 15	2822	2794 (23)	ν CH
ν 14	1749	1764 (137)	ν C=O
ν 13	1352	1361 (16)	δ CH
ν 11	1199	1182 (15)	tw. CH_2
ν 9	1076	1079 (66)	ν CO
ν 8	1051.5	1049 (32)	tw. CH_2 , δ CH
ν 6	1028	1023 (29)	ν CC

^a δ : bend, ν : stretch, tw: twist, as: asymmetric, s: symmetric.

TABLE IV. Vibrational frequencies (cm^{-1}) of the photo-dissociation products (most intense bands) formed upon photolysis of GA at 266 nm. Frequencies with no decimal values indicate the centre of broad absorptions. The – sign indicates that bands are not observed in the matrix.

	pH ₂	Ne	Ar	N ₂
CO	2142.9	–	2138.5	2139.8
	3679.5		3670.7	3679.4
	3000		3004	3001.6
CH ₃ OH	2951.7	–	2954	2954
	2840.8		2844.7	2844.5
	1031.1		1033.7	1033.7
CO-CH ₃ OH	–	3662.5	3649.8	3638.3
	2154	2154	2148.7	2152.9
	1036.3	1038.4	1039.4	1040.4
			1037.3	
H ₂ CO	2782.6	–	–	–
	1742.6			
HCO	2449.2	–	–	–
	1864.5			

2. Dissociation products

The fragments formed during the photolysis are mainly composed of isolated carbon monoxide (CO) and methanol (CH₃OH) molecules, and of their molecular complexes and aggregates. Frequencies are summarised in Table IV.

In Ar and N₂ isolated CO and CH₃OH molecules and 1:1 molecular complexes appear simultaneously. More intense CO:CH₃OH bands indicate that complexes are predominant ($I(\text{CH}_3\text{OH})/I(\text{CH}_3\text{OH}:\text{CO}) = 0.6$ for the νCO band of CH₃OH) (Figs. 6 and 7). In our experiments in Ne there is only a very small amount of fragments consisting essentially of CO:CH₃OH complexes showing a singlet band at 1038.4 cm^{-1} and a multiplet near 2154 cm^{-1} (Fig. 5). In pH₂, in contrast with the other matrices, isolated CO and CH₃OH molecules are by far more abundant than complexes as evidenced in Fig. 4(b). The multiplet structure of CO bands is in accord with the rotational structure of the molecule in pH₂¹⁶ but some weak bands also belong to complexes or aggregates. Figure S3 in the supplementary material²⁸ displays a zoom in the CO band region near 2140 cm^{-1} . Other striking differences observed in pH₂ are given by traces of H₂CO formaldehyde at 2782.6 cm^{-1} (symmetric νCH) and 1742.6 cm^{-1} ($\nu\text{C}=\text{O}$), and by very weak bands of HCO formyl radical at 1864.5 cm^{-1} ($\nu\text{C}=\text{O}$) and 2449.2 cm^{-1} (νCH) (Fig. S4 in the supplementary material²⁸). The presence of H₂CO in pH₂ matrix was confirmed by comparing frequencies and relative intensities after GA photolysis with a reference spectrum acquired for plain H₂CO in pH₂.

C. Photochemistry of glycolaldehyde: Kinetics

To quantify the different relaxation channels upon excitation of GA, a kinetic analysis was done in the different matrices. IR spectra were measured at different irradiation times to obtain the time evolution of the photo-products. Laser intensity was kept constant in all experiments but the total initial concentration of GA could have varied from one experiment to the other.

Irradiation effect was monitored by measuring integrated intensities of some bands as a function of irradiation time. The whole spectrum was considered for CC as well as the νCO mode. The same mode at $\sim 1068 \text{ cm}^{-1}$ was used for TT conformer and absorptions of monomeric CO and CH₃OH molecules and their complexes were taken in the 2139–2150 cm^{-1} and 1033–1040 cm^{-1} regions. Relative concentrations were obtained from integrated intensities using the Beer's law. Preliminary vibrational excitation experiments on CC in Ar enabled us to relate the molar absorption coefficients ε_{TT} to ε_{CC} , i.e., $\varepsilon_{TT}/\varepsilon_{CC} = 0.87$ for νCO . This value is close to the 0.78 value issued from calculated intensities within the anharmonic approach.⁷ For CO or CH₃OH, monomeric species and complexes were considered and a mean value was adjusted imposing that the total concentration of GA and products must stay constant. This procedure was used in classical matrices but, in pH₂, partial evaporation of the matrix did not allow us to perform complete kinetics. In N₂, TT and TG display similar behaviour upon irradiation, as checked on their νOH bands and they were treated as a single species arbitrarily referred to as TT. Their calculated band intensities for νCO mode are very close⁷ (63 and 66 km mol^{-1} in Tables II and III).

GA kinetics shows similar trends in Ar and N₂: CC decreases while TT grows up to a maximum and then decreases (Figs. 8(a) and 8(b)). In both matrices and in pH₂ the depletion of CC leads to completion. It is a consequence of a minor but irreversible mechanism, i.e., fragmentation. Fragments are produced from the very beginning of the photolysis at a significant amount in Ar and N₂ (Figs. 6–8). In Ne, CC and TT seem to reach equilibrium in our experimental conditions and only small amounts of fragments are produced (Figs. 5 and 8(c)). In fact, no concentration can be given for the fragments in Ne because of their very low band intensities in the IR spectra. Isomerization and fragmentation processes for CC and TT are taken into account, respectively characterized by (k_i , k_{-i})

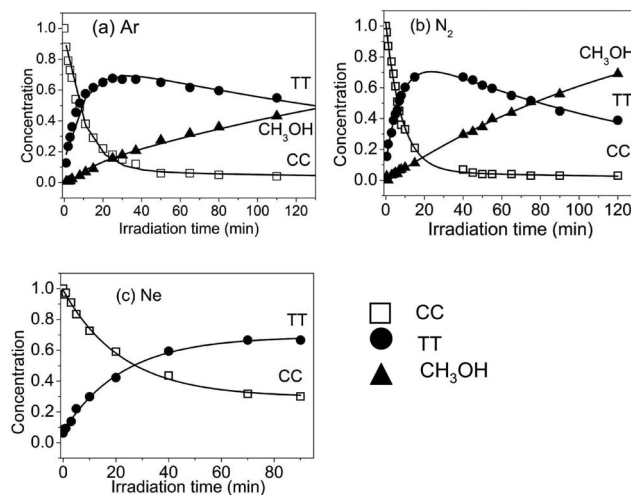
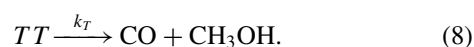
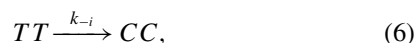


FIG. 8. Time evolution of the relative concentrations of CC (squares) and TT-plus TG in N₂-(circles) conformers of GA and CH₃OH (triangles) in (a) Ar, (b) N₂, and (c) Ne. Solid lines are fits obtained from Eq. (9), keeping the same k_+ and k_- parameters for the three species in each host. The 266 nm laser power was kept constant at $\sim 10 \text{ mW}$.

TABLE V. Rate constants (min^{-1}) obtained from the kinetic analysis of the photolysis of GA, as described in Eqs. (5)–(8), upon 266 nm irradiation at ~ 10 mW laser power.

	Ar	N ₂	Ne
k_i	0.070	0.090	0.030
k_{-i}	0.010	0.005	0.014
k_C	0.010	0.006	0.001
k_T	0.003	0.006	0.001

and (k_C , k_T) rate constants. The kinetic equations are



Each step is considered as elementary. Solving the kinetic equations leads to Eq. (9) which expresses the concentration $C_i(t)$ of each i species as a sum of two exponential functions involving the same k_+ and k_- rate constants for all species plus a constant value,

$$C_i(t) = A_{i1} \exp(-k_+ t) + A_{i2} \exp(-k_- t) + A_{i3}, \quad (9)$$

where A_{i1} , A_{i2} , and A_{i3} parameters are characteristic of each species i . The k_+ and k_- constants are defined as follows:

$$k_+ = (k_i + k_C + k_{-i} + k_T + \Delta^{1/2}) / 2,$$

$$k_- = (k_i + k_C + k_{-i} + k_T - \Delta^{1/2}) / 2,$$

$$\begin{aligned} \Delta &= (k_i + k_C + k_{-i} + k_T)^2 - 4[(k_i + k_C)(k_{-i} + k_T) - k_i k_{-i}] \\ &= [(k_i + k_C) - (k_{-i} + k_T)]^2 + 4k_i k_{-i}. \end{aligned} \quad (10)$$

Figure 8 displays the corresponding fitting curves as solid lines. The rate constants obtained with ~ 10 mW laser power are listed in Table V. From our measurements, determination of k_i is more accurate. Estimated uncertainties are about 10% for k_i and are at least 20% for the other rate constants that are one or two orders of magnitude lower.

The rate constants reflect a combination of diverse and complex processes: photolysis of isolated GA, relaxation of electronically and/or vibrationally excited GA, host effects on GA and photoproducts, etc. Two main effects can be underlined from Table V: isomerization is the major process following CC excitation ($k_i \gg k_C$), and CC displays a higher photo-reactivity at 266 nm than the open conformers ($k_i + k_C \gg k_{-i} + k_T$). In fact, the latter effect simply results from the less efficient excitation of open conformers, as shown by the UV absorption spectra (Fig. 3). The k_C and k_T rate constants are comparable in N₂ but in Ar, $k_C > k_T$. In these matrices, fragmentation occurs from both CC and TT conformers. Only qualitative trends can be given in pH₂. TT appears at early irradiation times but CO and CH₃OH appear later. In fact, they grow very slowly and seem to follow the increase of TT, suggesting that they are essentially produced from TT.

D. Theoretical results

We here focus on the isomerization channels accessible with a 266 nm excitation. Aldehydic compounds easily undergo intersystem crossing (ISC) from S_1 to T_1 upon electronic excitation.¹⁷ This is also the case of GA as estimated by the spin-orbit coupling (23.4 cm^{-1}) calculated in Ref. 5 for the structure at the minimum energy on S_1 . As a consequence, calculations were performed on both S_1 and T_1 states. Since isomerization mainly consists of rotations around the backbone of GA, CCOH and OCCO dihedrals were varied to construct the two-dimensional PES of S_0 , S_1 , and T_1 states.

1. S_0 ground-state calculations

The most stable CC conformer is stabilised by a hydrogen-bond with a distance of 2.13 Å between the hydroxyl hydrogen and the carbonyl oxygen. TT and TG conformers are very close in energy (15 kJ/mol from CC) and the highest CT conformer is 22 kJ/mol above CC (Fig. 1). These values are in excellent agreement with those obtained with the CASSCF/cc-pVDZ method,⁵ with the MP4(SDTQ)/cc-pVQZ method,⁶ and with the CCSD(T)/cc-pVTZ//B3LYP/6-31G(d,p) hybrid quartic force field.⁷ Energies and dihedrals of the four conformers are listed in Table VI.

Our method, therefore, succeeds in capturing the structural and energetic properties of GA conformers at a limited computational cost, which is required in order to perform extensive calculations of PES. The torsional barriers between the conformers were calculated along the minimum energy paths and agree within 2 kJ/mol with the values obtained with the MP4(SDTQ)/cc-pVQZ method.⁶ TT and TG are separated by a small barrier (2 kJ/mol), but it requires 9 kJ/mol to convert to CC. The barrier for the CT \rightarrow CC conversion is 2 kJ/mol. The resulting surface is shown in the bottom part of Fig. 9. Note that this small CT-CC barrier can explain the non-observation of CT in our experiments: even if they are formed upon irradiation, CT isomers can easily transform in CC by proton tunnelling.

2. S_1 and T_1 excited state calculations

a. Potential energy surfaces. The lowest singlet transition is of HOMO-LUMO character. The electron excitation proceeds from the n orbital localised on the carbonyl oxygen to the π^* orbital of the C=O group. This leads to important changes in charge distributions and strengths of bonds, giving very different geometries for the energy minima in S_1 with respect to S_0 . Scan of S_1 -PES shows three energy basins. Each basin is also reached by full geometry optimisation on S_1 after vertical excitation from CC, TT, and TG geometries. The corresponding optimised structures are called minCC, minTT, and minTG, respectively, for the sake of simplicity (Fig. 9). Since the PES of S_1 and T_1 states share similar topologies (Table VI and Fig. 10) we refer in this paragraph to the same “minx” notation to describe the minima on S_1 and T_1 . In particular, the S_1 -minima adopt very close geometries to their T_1 counterparts: dihedrals differ by less than 3° and C=O bond lengths are similar.

TABLE VI. Dihedral angles and calculated energies (kJ/mol) of GA conformers in the ground state and in the S_1 and T_1 excited states. Wavelengths in nm are given in *italic*.

	\angle OCCO ($^\circ$)	\angle CCOH ($^\circ$)	ΔE (kJ/mol)	Vert. energy kJ/mol (nm)	
				$S_1 \leftarrow S_0$	$S_1/T_1 \rightarrow S_0$
CC	0	0	0	441 (271)	
TT	180	180	15	411 (291)	
TG	-163.3	75.7	15	395 (303)	
CT	0	180	22	424 (282)	
minCC S_1/T_1	-73.0 / -71.5	72.5 / 71.8	370 / 295		290 / 180 (413 / 665)
minTT S_1/T_1	76.2 / 79.9	59.1 / 56.6	370 / 298		282 / 174 (424 / 688)
minTG S_1/T_1	-161.9 / -163.8	64.5 / 62.6	373 / 300		283 / 178 (423 / 672)

Energies of minCC, minTT, and minTG in S_1 and T_1 with respect to CC conformer are reported in Table VI. MinCC and minTT are more stable than minTG on S_1 and T_1 and are quasi-isoenergetic. None of the minima presents intramolecular H-bond; details of their structures are given in the supplementary material.²⁸ MinCC and minTT are highly distorted compared to CC and TT, i.e., twisted conformations with respect to the C–C bond are favoured instead of a coplanar configuration. In contrast, minTG is structurally similar to the TG conformer (see Fig. 9 and Table VI). Note that the structures of minCC are similar to the minima obtained by Cui *et al.* in the S_1 and T_1 excited states⁵ (Table S1 in the supplementary material²⁸).

b. Transition energies. Vertical and adiabatic energies between S_0 and S_1 (respectively, T_1) were calculated. Relevant data relative to CC, TT, and TG are summarised in Table VI. Vertical energies for the three conformers fall under the laser excitation energy (450 kJ/mol, corresponding to 266 nm). The S_0 - S_2 PES is found at much higher energy (550 kJ/mol) and cannot be reached with a single photon absorption. The predicted S_0 - S_1 value for CC (441 kJ/mol, 271 nm) is in agreement with the electronic spectra (\sim 427 kJ/mol, or 280 nm in the gas phase;⁴ 435 kJ/mol, 275 nm in this work in matrices). It is slightly blue-shifted compared to the 426 kJ/mol (B3LYP) and 421 kJ/mol (CASPT2) values previously reported.⁵ The calculated

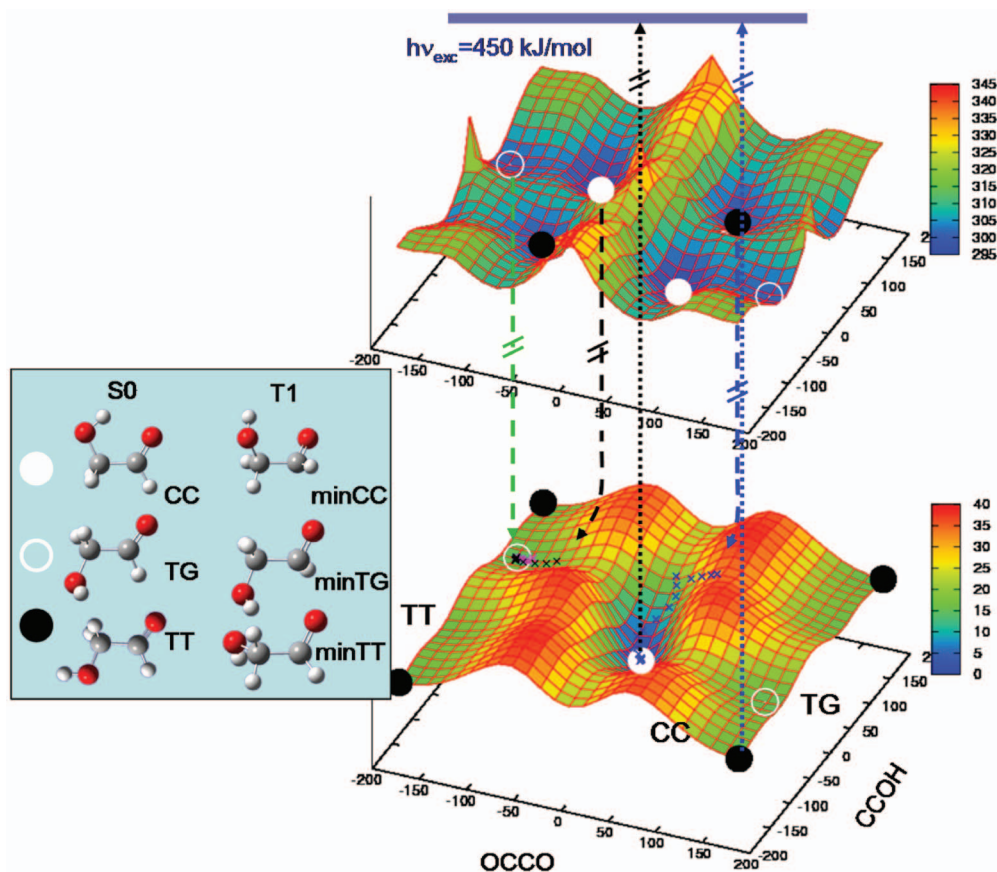


FIG. 9. Potential energy surfaces along OCCO and CCOH dihedrals on the S_0 ground state (bottom) and T_1 excited state (top) and their corresponding minima (CC and minCC are displayed in plain white circles, TG and minTG open white, and TT and minTT plain black, respectively) calculated at the B3LYP/def2-TZVPP level of theory. Energy scales are in kJ/mol. Dotted arrows refer to vertical excitation of CC (black) and TT (blue). Dashed arrows refer to relaxation from T_1 -minCC (black), T_1 -minTT (blue), and T_1 -minTG (green), respectively. Crosses indicate relaxation pathways on the S_0 surface.

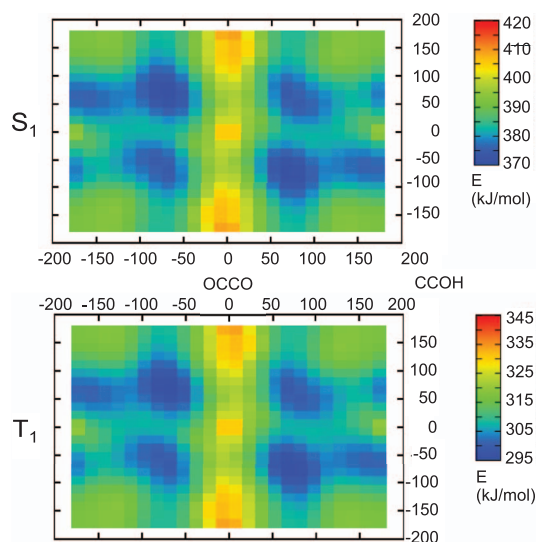


FIG. 10. Contours of S_1 (top) and T_1 (bottom) two-dimensional PESs along the OCCO and CCOH dihedrals showing similar topologies of both surfaces. Energies are given in kJ/mol.

transition for TT (411 kJ/mol, 291 nm) is not far from the electronic absorption measured in matrices, approximately centred at 307 nm. The TG conformer also absorbs in this region (395 kJ/mol, 303 nm). Given the accuracy of the TDDFT method,¹⁸ experimental and theoretical values, therefore, appear in good agreement. Vertical emission energies from the S_1 and T_1 minima to the ground state are also listed in Table VI. The $S_1 \rightarrow S_0$ (respectively, $T_1 \rightarrow S_0$) transitions are predicted around 285 kJ/mol, i.e., near 420 nm (respectively, around 177 kJ/mol, ~ 676 nm), in the typical range of fluorescence spectra of many aldehydes.¹⁷ Note that the values found for S_1 -minCC, minTT, or minTG (from 282 to 290 kJ/mol, 413–424 nm) are very close to the fluorescence band measured in solution by Beeby *et al.* (288 kJ/mol, 415 nm).¹⁵

c. Relaxation pathways. Upon laser excitation, GA reaches the S_1 state with a large amount of internal energy (~ 80 kJ/mol). In the matrix, rapid vibrational relaxation in the electronic excited states can occur by energy transfer to the phonons. GA could, thus, relax to the three minCC, minTT, and minTG S_1 -basins, these minima being very close in energy. Taking into account ISC, the system can also be brought to the corresponding T_1 minima. We calculated the relaxation pathways to the ground state following radiative relaxation from the three minima in the S_1 and T_1 -PESs (see Fig. 9). Since the geometries of GA at the three minima are very similar in both excited states, the relaxation channels are the same coming from S_1 or T_1 states. From minCC basin the system decays to a vibrational level of the ground state down to the well of TG conformer. TG is also obtained by relaxation from minTG with minor geometry changes. Decaying from minTT leads to CC conformer in the ground state. Our calculations show that only TG and CC conformers should be formed directly.

The highly distorted structures on S_1 and T_1 excited states are, therefore, possible routes to the formation of open conformers upon excitation of CC.

IV. DISCUSSION

A. Host effects

1. Specificity in N_2

Matrix-isolation of GA in N_2 brings interesting results in terms of spectroscopy and stability of the open conformers. Only in N_2 are TT and TG visible in freshly deposited matrices. It turns out that their amount varies from zero to very few percent depending on the injection conditions. This leads us to suppose that TT and TG are already present in the gas-phase mixture (GA and carrier gas) as low amount of GA- N_2 complexes since formation of complexes in a jet depends on the mixture and pressure of the gases. Following this hypothesis, we have performed calculations on GA- N_2 complexes in order to evaluate the stabilisation brought by complexation. Our results on isolated complexes for CC, TT, and TG show larger stabilisation energies for TT- N_2 and TG- N_2 complexes (-7.78 kJ/mol and -9.13 kJ/mol) than for CC- N_2 (-5.89 kJ/mol). These findings strongly suggest that the presence of N_2 preferentially stabilises the open conformers.

Comparing the IR spectra in all media also gives us some hints in this direction. For instance, the ν OH TT frequency is very sensitive to the environment: it is the lowest in N_2 (also in Ar) and the highest in Ne (Table II). Actually, matrix shifts of OH group with respect to the gas phase are usually the weakest in Ne and the largest in N_2 and similar behaviour was reported for several organic molecules.¹⁹ These differences can be explained by stabilising interactions of the GA molecules with the matrices, resulting from different polarisabilities of the hosts and also quadrupolar interactions for N_2 . We, thus, suggest that TT and TG conformers are more stabilised in N_2 , making it possible to observe both conformers in this specific matrix.

The low TT-TG torsional barrier (~ 2 kJ/mol) and the slightly higher stability of TT over TG in the free GA molecule⁶ can possibly induce an open pathway from TG to TT, explaining why we do not observe it in pH_2 , Ne, and Ar. In N_2 , interaction with the matrix may either result in a higher barrier that would close the TG \rightarrow TT pathway or in a smaller energy-difference between TT and TG so that both conformers would co-exist in thermal equilibrium. The relative ν OH band intensities of TT and TG suggest that conformers are roughly equally populated in N_2 in agreement with the latter hypothesis. It should be noted that, in contrast to CT which is also separated from CC by a 2 kJ/mol barrier, the TG \rightarrow TT conversion requires more than rotation of the hydroxyl group, which can, in particular, occur through tunnelling processes as reported in matrices.²⁰

2. Radicals in pH_2

In the gas phase, the formation of HCO and CH_2OH radicals (C–C bond cleavage) dominates the photolysis of GA (channel 1). No radicals were seen in classical Ar, Ne, and N_2 matrices. Their absence can be explained by recombination of primary fragments in the matrix cage. The cage effect in the different matrices can be seen through the observation of CH_3OH -CO complexes, CO and CH_3OH being

photo-produced in the same matrix cage. The prominence of such complexes in Ne illustrates a very efficient process in this solid. In contrast in pH_2 , isolated CO and CH_3OH molecules are mainly formed, the cage exit is made easier due to the softness of the lattice.²¹ Minimised geminate recombination, thus, allows detecting HCO radical – even weakly – in this peculiar solid, thereby showing that radical channels are not entirely closed. The CH_2OH counterpart is not observed in pH_2 . According to literature, this radical can dissociate to $\text{HCO}+\text{CO}$,²² or $\text{H}_2\text{CO}+\text{H}^{23}$ upon UV irradiation. We suggest, thus, that this photo-dissociation occurs under our experimental conditions, which could explain the absence of CH_2OH and the presence of H_2CO .

B. Relaxation channels in matrices

In the gas phase, four channels (Eqs. (1)–(4)) are opened as first steps of photo-dissociation upon UV irradiation.^{2–4} channels 1–3 lead to radical fragments, channel 4 to CO and CH_3OH molecular fragments. Following the results obtained by Zhu and Zhu,⁴ most of GA undergo photo-fragmentation upon UV irradiation; one can estimate that channels (1–4) represent at least 70% of the relaxation pathways. In the matrix, two types of products are detected, isomers and molecular fragments as expressed in Eqs. (5) and (7). An additional minor reaction is observed in pH_2 only, leading to the production of HCO radical. The absence of radicals in Ne, Ar, and N_2 may simply reflect the cage effect in the solid inducing geminate recombination of neighbouring radicals to form stable molecules, GA or $\text{CO}+\text{CH}_3\text{OH}$. However, our results allow a deeper insight into the processes at play, as discussed in the following paragraphs.

1. Molecular elimination

The $\text{HOCH}_2\text{CHO}(\text{GA}) \rightarrow \text{CH}_3\text{OH}+\text{CO}$ molecular channel is observed in all media. It accounts at most for 10% of the gas-phase processes when considering excited GA molecules, as based on Zhu and Zhu studies, where the quantum yields of photo-fragmentation and the amount of photoproducts were measured.⁴ In matrices, according to the k_C/k_i ratio (Table V), CH_3OH and CO that come from CC only represent 3 to 14% of the products. Data relative to the return to CC is not available; these numbers, thus, overestimate the efficiency of the molecular channel. Yields in matrices are not much larger than in the gas phase, thereby implying that there is no enhancement of this channel in the solids. Furthermore, our calculations indicate that TT and TG conformers are more stable than the $\text{CO}+\text{CH}_3\text{OH}$ dissociation limit by 2 kJ/mol. It is in favour of recombination of radicals to GA conformers at the expense of formation of CO and CH_3OH molecules in a confined medium.

Several mechanisms may be responsible for the “direct” formation of CO and CH_3OH . An earlier theoretical study on methyl formate, a structural isomer of GA, showed that a transition state with the H atom bridging the carbonyl carbon and the ester oxygen could lead to $\text{CO}+\text{CH}_3\text{OH}$.²⁴ The analogue in the case of GA could then involve a transition state having the H atom shared between the C atoms, struc-

turally similar to the hydrogen-transfer transition state of Ref. 5 (referred to as “S1-TS-PT”). Still for methyl formate, the so-called roaming mechanism implying dissociation on vibrationally hot ground state also yielded $\text{CO}+\text{CH}_3\text{OH}$ but it was only observed in specific conditions and was found to be weak.²⁵ So, we assume that molecular elimination in the matrix and in the gas phase occurs in a similar manner and with similar efficiency: the system passes through a transition state rather than through a dissociation-recombination process.

2. Isomerization

Isomerization occurs in matrices only. It is the dominant process as illustrated in Figs. 4–7 and by the constant rates of Table V. It can result from two competitive pathways: (i) direct isomerization that proceeds on the PESs explored by the system or (ii) recombination of dissociation products due to the cage effect in the matrix.

Direct isomerization is due to the coupling with phonons which allows thermalisation of the system in the excited state, followed by relaxation in the ground state. This mechanism could entirely or partly quench the channels to dissociation. Similar hypothesis was proposed for acetylacetone to rationalise the differences between its photo-reactivity in the gas phase and in matrices, where similar behaviours in rare gas matrices, N_2 , and pH_2 were observed.²⁶ Assuming that relaxation of excited GA occurs in the electronic states by bath thermalisation and radiative transitions between the states, isomerization is a direct consequence of irradiating the molecule with UV light. Our calculations find that only CC and TG conformers should be formed after excitation of GA. The observation of TT, thus, likely originates from a secondary $\text{TG} \rightarrow \text{TT}$ process in the ground state due to the low TG-TT barrier. Reference 5 indicates that the system can also relax through non-radiative transitions via S_1/S_0 and T_1/S_0 conical intersections. In the gas phase, all these pathways lead to relaxation toward the most stable CC conformer but matrices stabilize the TT form (and TG form in N_2). Direct isomerization may be seen as a step in radiative de-excitation and internal conversion in the isolated GA molecule.

Considering that fragmentation proceeds through C–C bond cleavage, the CH_2OH and HCO radicals can easily recombine in the matrix cage (process *ii*). According to theory, the corresponding transition state (415 kJ/mol) is easily accessible under our UV excitation.⁵ Recombination due to the matrix cage is host-dependent: it should be the strongest in Ne and the weakest in pH_2 , on the basis of the previous discussion on the host effects. Although the complete kinetic study is not available in pH_2 , TT is clearly detected meaning that isomerization is at play in pH_2 and has a similar efficiency as in Ne (see Figs. 4 and 5). This suggests that direct isomerization (i) of GA is more efficient than the process through recombination (ii) in all matrices.

Because of its high branching ratio, direct isomerization cannot be the simple matrix analogue of the radiative and non-radiative relaxation processes back to CC ground state that occur in the gas phase, where these “internal” relaxation processes represent less than 30% of the channels according to Ref. 4. Moreover, the authors clearly show that the quantum

yield associated to HCO production is $\sim 60\%$. Comparison with our observations in pH_2 , where TT conformer is much more abundant than HCO, highlights a strong solvent effect quenching – although not completely – the radical dissociation channels that occur in the electronic excited states. This conclusion is in agreement with the recent work on the photodissociation of N-methylformamide in pH_2 ,²⁷ where the authors suggest that the pH_2 surrounding enhances internal conversion to the ground state to explain the high ratio of the molecular channel compared to the radical pathways.

3. Reactivity of TT and TG conformers

As shown in the experiments, isomerization and molecular channels are also opened for the TT (and TG) conformer of GA in matrices (Eqs. (6) and (8)).

A surprising finding is the relative efficiency of the backward $\text{TT} \rightarrow \text{CC}$ conversion in Ne compared to Ar and N_2 ($k_i/k_{-i} = 2, 7, \text{ and } 18$, respectively). In fact, k_i and k_{-i} reflect a global behaviour of the sample during laser irradiation. In their study of GA in rare gas matrices, Aspiala *et al.*⁹ observed thermal conversion from TT to CC in Xe and Kr at temperatures over 30 K, whereas the barrier cannot be classically overcome in these conditions. With the setup used for experiments in Ne and pH_2 , thermal processes do occur, as evidenced by the partial evaporation of the pH_2 matrices upon UV irradiation. Thus, the quite efficient backward $\text{TT} \rightarrow \text{CC}$ isomerization may be partly due to a thermal effect in Ne. Such thermal effect upon irradiation cannot be completely discarded in Ar and N_2 . Nevertheless, the stronger stabilisation of TT and TG in N_2 could prevent it in N_2 explaining differences between Ar and N_2 .

The branching ratio between backward isomerization and molecular elimination is not known accurately, because of the weakness of these two reactions since TT is not efficiently excited at 266 nm. Nevertheless, the k_{-i}/k_T ratio (Table V) should more properly reflect the effect of UV irradiation on open conformers in N_2 than in Ar and Ne, where thermal effects may perturb the observations. It indicates that the molecular elimination is a stronger process in the photolysis of open TT and TG conformers than that of CC. The efficiency of the molecular channel depends on the relaxation pathways followed by the system in the excited and ground states. These pathways depend on the initial geometry of GA and could converge more easily to the transition state of the molecular channel when starting from an initial TT geometry than from CC. This remark is in agreement with the fact that most fragments in pH_2 seem to come from TT irradiation.

V. CONCLUSIONS

We have studied the UV photolysis of GA in cryogenic matrices at 266 nm accompanied with excited state calculations on GA. Four different hosts have been explored in order to extract matrix effects. The IR spectroscopy of as deposited samples highlights the specificity of nitrogen matrices, where more conformers have been detected.

Upon irradiation of the CC H-bonded conformer, the open TT species is formed and stabilised in argon, neon, nitrogen, and para-hydrogen matrices. The TG conformer is also observed in nitrogen. To date, this is the first experimental evidence of the TG conformer of GA. According to calculations on the isolated molecule, TT and TG possess the same energy and are separated by a very low torsional barrier. The photo-induced formation of TG in nitrogen only is unlikely and we postulate that TG probably isomerises to TT in the other matrices studied. In addition, both conformers display similar behaviour upon deposition and irradiation in nitrogen. Following that, we conclude that TT is more stabilised in argon, neon, and para-hydrogen, while in nitrogen both TT and TG are in thermal equilibrium.

Dissociation products such as CH_3OH , CO, and their complexes are detected in all matrices but additional H_2CO and HCO radicals are detected in para-hydrogen. Despite its non-observation in classical matrices, the presence of HCO in the quantum solid provides clear evidence that the predominant gas-phase dissociation channel, i.e., fragmentation to HCO and CH_2OH (CC-bond cleavage) also occurs in matrices. However, the very low concentration of HCO shows that it is nearly quenched in these environments. Dissociation to $\text{CO} + \text{CH}_3\text{OH}$ is observed in all matrices but it is minor as in the gas phase. We conclude that the molecular fragments formed in matrix result from the same mechanism as in the gas phase and that cage effect is not (or only weakly) involved.

Cage effect can neither explain the efficiency of isomerization, as shown by the comparison between para-hydrogen and the other hosts. We think that isomerization is due to relaxation of GA in the excited and ground states by coupling with matrix phonons. This conclusion is supported by theoretical calculations of the excited S_1 and T_1 PES which exhibit non-planar geometries for minimum energy structures, inducing the opening of the internal H-bond. Relaxation of GA, firstly through the excited states and secondly through the ground state, can then lead to the formation of the open conformers.

The UV photo-reactivity of open TT (and TG) conformers has also been investigated. Our results show that fragmentation and isomerization also compete in the relaxation of excited open conformers. Fragmentation is not anymore a minor process, probably because of differences in the relaxation pathways through the electronic states after UV excitation. This process being irreversible, all GA is consumed at very long irradiation times and we observe a relatively high concentration of products.

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