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LaMgX and CeMgX ($X = \text{Ga, In, Tl, Pd, Ag, Pt, Au}$) with ZrNiAl type structure – A systematic view on electronic structure and chemical bonding

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ABSTRACT

The intermetallic magnesium compounds LaMgX and CeMgX ($X = \text{Ga, In, Tl, Pd, Ag, Pt, Au}$) crystallize with the hexagonal ZrNiAl type structure, space group $P\bar{6}2m$, with full Mg–X ordering. From density functional theory calculations carried out exemplarily on four representative compounds: LaMgX and CeMgX with $X = \text{Ga, Pd}$, significant differences were traced out as to the magnetism arising only for the Ce series leading to identify CeMgGa as an antiferromagnet in its ground state, in agreement with experiment. The bulk module magnitudes show the trend of harder transition metal based ternaries and the cohesive energies favor the $X = \text{Pd}$ compounds versus $X = \text{Ga}$ ones. Such features were clarified by examining the properties of chemical bonding which exhibit more directional bonds thanks to the Pd d states. Rationalizing the trends of charge transfers, negatively charged triel and transition element atoms are observed. The resulting chemical pictures assign these compounds as *gallides* and *palladides*.

Keywords: A. functional alloys (magnetic, electrical, biomedical); B. Magnetic properties; B. Density functional theory; E. Ab initio calculations

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1. Introduction

The equiatomic CeTX (T = electron-rich transition metal; X = element of the 3rd, 4th, or 5th main group) phases [1] are one of the broadest investigated families of intermetallic rare earth compounds due to their largely differing magnetic and electric properties. Some outstanding examples are the heavy-fermion silicide CePtSi [2], the 10 K ferromagnet CeAuGe [3], the formation of an anisotropic energy gap in the valence-fluctuating stannide CeNiSn [4], the Kondo semiconductor CeRhSb [5, 6], the non-Fermi liquid system CeRhSn [7], the valence modulations in CeRuSn [8, 9], or the heavy-fermion ($\gamma = 300 \text{ mJ}/(\text{mol K}^2)$) antiferromagnet ($T_N = 0.9 \text{ K}$) CePtPb [10]. Several hundred manuscripts on CeTX intermetallics have been published. Most data are summarized in review articles [11, 12].

The peculiar bonding characteristics and magnetic ground states of the CeTX intermetallics are often a consequence of the $\text{Ce}-T$ hybridization. Substitution of the X component or the formation of solid solutions on both the T and X sites leads to substantial influence on the cerium electronic situation and thus changes in the magnetic behavior, e. g. $\text{CePt}_{1-x}\text{Ni}_x\text{Sn}$ with a transition from an antiferromagnetic Kondo lattice to a mixed valence semiconductor [13], $\text{CeNi}_{1-x}\text{Rh}_x\text{Sn}$ with an interplay between spin-glass-like and non-Fermi-liquid behavior [14] or $\text{CeRhSb}_{1-x}\text{Sn}_x$ with a quantum criticality at the border of Kondo insulator / non-Fermi liquid [7].

Other parameters that influence the magnetic behavior are temperature, high-pressure or hydrogenation. Recent examples are the low- and high-temperature modifications of CePdZn [15] with stronger $\text{Ce}(4f)\text{-Pd}(4d)$ hybridization in $\alpha\text{-CePdZn}$, the temperature induced structural and cerium valence change in CeRhGe [16] and the

high-pressure modification of CePtSn [17]. Hydrogenation of CeZnSn to CeZnSnH_{1.5} increases the Curie temperature from 4.8 to 7.3 K [18].

Modification of the magnetic properties is also possible through substitution of the transition metal by another p element, thus suppressing d - f hybridization. First investigations in such phase diagrams led to the magnetically ordering phases CeGaGe [19, 20], CeAlSi and CeAlGe [21] and CeGaSi [22, 23]. Systematic studies have been performed for the $REMgX$ ($X = \text{Ga, In, Tl}$) phases [24-28] with ZrNiAl type structure, including the cerium intermetallics CeMgGa, CeMgIn, and CeMgTl of which CeMgGa orders antiferromagnetically at $T_N = 3.1$ K. These phases can be considered as the p element analogues of the transition metal counterparts CeMg T ($T = \text{Pd, Ag, Pt, Au}$) [29-31]. These substitutions offer an interesting situation for changes in chemical bonding while keeping the structure type. Herein we present a detailed study of the electronic structures and the chemical bonding situation of these two families of intermetallics.

2. Computational details

Within the framework of the density functional theory (DFT) [32, 33] two computational methods were used in a complementary manner. For the purpose of firstly obtaining energy (total and cohesive) and charge transfer quantities, the Vienna *ab initio* simulation package (VASP) code [34, 35] was used with the projector augmented wave (PAW) method [35, 36]. DFT exchange and correlation effects were accounted for in the framework of the generalized gradient approximation (GGA) scheme following Perdew, Burke and Ernzerhof (PBE) [37]. Atoms were relaxed and the conjugate-gradient algorithm [38] is used in this computational scheme. The tetrahedron method with Blöchl corrections [39] as well as Methfessel-Paxton [40] Brillouin-zone (BZ) integrals were approximated using the special k -point sampling of Monkhorst and Pack [41]. A large number of 150 k -points was used in the irreducible wedges of the hexagonal Brillouin zone to perform total energy calculations. The optimization of the structural parameters was performed until the forces on the atoms were less than $0.02 \text{ eV}\text{\AA}^{-1}$ and all stress components less than $0.003 \text{ eV}\text{\AA}^{-3}$.

Then the full investigation of the electronic and magnetic structures and the chemical bonding properties was performed using the full potential augmented spherical wave (ASW) method [42, 43] in a scalar relativistic implementation [44] with exchange-

correlation effects accounted for within the generalized gradient approximation GGA [37]. In the ASW method, the wave function is expanded in atom-centered augmented spherical waves, which are Hankel functions and numerical solutions of Schrödinger's equation, respectively, outside and inside the so-called augmentation spheres. In the minimal ASW basis set, we chose the outermost shells to represent the valence states and the matrix elements were constructed using partial waves up to $l_{max} + 1 = 4$ for La and Ce, $l_{max} + 1 = 3$ for Pd and, $l_{max} + 1 = 2$ for Mg and Ga. Self-consistency was achieved when charge transfers and energy changes between two successive cycles were such as: $\Delta Q < 10^{-8}$ and $\Delta E < 10^{-6}$ eV, respectively. The BZ integrations were performed using the linear tetrahedron method [39] within the irreducible hexagonal wedge.

The calculations were carried out assuming firstly spin degenerate, non spin-polarized (NSP), then spin polarized (SP) magnetic calculations for both ferromagnetic and antiferromagnetic long range orders for the cerium compounds. The relative magnitude of the chemical bonding is obtained based on the overlap population analysis: S_{ij} , i and j being two chemical species. The crystal orbital overlap population (COOP) criterion is used [45]. In the plots positive, negative and zero COOP magnitudes indicate bonding, anti-bonding and non-bonding interactions respectively.

3. Results and discussion

3.1. Structural chemistry

The CeMgX ($X = \text{Ga, In, Tl, Pd, Ag, Pt, Au}$) phases crystallize with the hexagonal ZrNiAl type structure [46-48], space group $P\bar{6}2m$ Pearson code hP9, and Wyckoff sequence *gfda*. A projection of the CeMgX structure onto the xy plane is presented in Fig. 1. The striking structural motif is the trigonal prismatic coordination of the X atoms. The $X1$ atoms are coordinated by six cerium atoms and six of these $X1@Ce_6$ prisms condense via common edges forming hexagonal rings. The latter are filled by the $X2@Mg_6$ prisms. Both prismatic motifs are shifted with respect to each other by half the c lattice parameter, emphasized by thin and thick lines in Fig. 1. Condensation along the c axis proceeds through common triangular faces.

The cell parameters of the CeMgX phases are listed in Table 1 along with their lanthanum counterparts. The latter have been used for additional calculations in order to avoid problems with the f states. For all pairs of compounds the cell volume of the cerium phase is smaller, a consequence of the lanthanide contraction. From comparison with the complete series [25-27, 29], all cerium phases are trivalent, in agreement with the magnetic susceptibility data. An interesting situation occurs for the compounds with the heaviest X components. In the pairs CeMgPd/CeMgPt, CeMgAg/CeMgAu, and CeMgIn/CeMgTl, the platinum, gold, and thallium compound show the smaller cell volumes, a consequence of the relativistic contraction.

A drastic difference between the CeMgX phases with a p and a d element is the much larger c/a ratio for those with the p elements (Table 1). This behavior is directly related to the magnesium triangles forming the $X_2@Mg_6$ trigonal prisms around the origin of the unit cells. The Mg–Mg distances for CeMgIn (325.9 pm), CeMgTl (327.0 pm), CeMgAg (325.6 pm), and CeMgAu (323.1 pm) are all close to the average Mg–Mg distance of 320 pm in *hcp* magnesium [49]. Thus, changes in Mg– X bonding are managed through increase or decrease of the height of the trigonal prism.

The shortest interatomic distances in the CeMgX structures occur for Mg– X . The average Mg–In (302 pm), Mg–Tl (301 pm), Mg–Ag (295 pm), and Mg–Au (291 pm) distances (Table 2) are all longer than the sums of the covalent radii [50] of 286, 291, 270, and 270 pm, respectively, indicating medium bond strength. It is evident from Table 2, that the cerium atoms have five indium, thallium, silver, or gold atoms as nearest neighbors. Thus, the electropositive cerium atoms bind to the more electronegative triel or transition metal atoms. Accordingly we can ascribe thallide and auride character to CeMgTl and CeMgAu. These features of chemical bonding are addressed in more detail in the following sections.

3.2. *Electronic structure and chemical bonding*

In view of the large CeMgX and LaMgX equiatomic family of X = triel and X = transition metal intermetallics, we have elected to show the results representative of X = Ga and X = Pd respectively.

3.2.1 *Energies and derived quantities*

The experimental crystal data (Table 1) were firstly used for carrying out unconstrained geometry optimization runs with VASP-PAW-GGA in order to get the minimum equilibrium energies and energy derived quantities as the bulk modulus B_0 based on Birch energy-volume equation of state [51] for the four selected compounds, as well as an analysis of trends of charge transfers. Table 3 provides the calculated parameters of total electronic energy, cohesive energy, volume and B_0 . According to the nature of X , the compounds are regrouped in close values of all four quantities with smaller magnitudes for $X = \text{Ga}$. The calculated volumes are found in relative good agreement with experiment with the trend of larger values due to the use of the GGA XC functional known to be under-binding. The cohesive energies obtained by subtracting the energies of the atomic constituents in their ground states from the total energies. The $X = \text{Pd}$ compounds are found more cohesive than $X = \text{Ga}$ ones and the cerium based compounds are found with larger cohesive energy than the lanthanum compounds. Similar trends are found for the bulk modules which are found larger for $X = \text{Pd}$ than for $X = \text{Ga}$ with larger magnitudes for the Ce based ternaries. With one more electron brought in by Ce versus La and the presence of d states in Pd versus p Ga states, directional bonding should be at the origin of the observed differences. These aspects can be reflected in the analysis of the chemical bonding discussed in the next section. Lastly the B_0 magnitudes are within the range of ternary equiatomic intermetallic compounds [11].

We also discuss the trends of charge transfers obtained from the analysis of the charge density within the atoms in molecules AIM theory of Bader [52]. Table 4 provides the charges on the respective atoms in the four studied compounds. The trend of negatively charged X is observed in all four compounds leading to assign the chemical labeling of gallides and palladides. Positively charged La, Ce and Mg are identified with larger electropositive values in the Pd based intermetallics likely due to the larger electronegativity: $\chi_{\text{Pd}} > \chi_{\text{Ga}}$ (2.2 versus 1.81) while Ce and La have close χ of 1.1. The difference in magnitudes between $X1$ and $X2$ arises from their respective crystal chemistry, i.e. with respect to $X1 @ \text{Ce}_6$ prisms versus $X2 @ \text{Mg}_6$ prisms.

3.2.2 Total-spins and spin-polarized magnetic calculations (NSP, SP)

With the experimental structure parameters (Table 1) the complete site and spin projected density of states (DOS) were computed in both non spin polarized (NSP) and spin polarized (SP) configurations using the full potential scalar relativistic ASW method.

Figs. 2 and 3 regroup the NSP DOS for the two series. Along the x axis the energy reference is with respect to the Fermi level (E_F) which is crossed by a large Ce(DOS) while a deep minimum is observed for LaMgGa and a small DOS magnitude can be seen for LaMgPd. The difference is due to the electron configuration of the rare earth elements Ce and La which differ by one $4f$ electron. This point is further stressed in the developments related with the magnetic instability.

The valence band –VB– shows large differences due to the Pd d states at ~ -3.5 eV in Fig. 3 whereas small magnitude broad Ga p states are observed in the VB of Fig. 2. In so far that the bonding is mainly ensured within the VB, one may expect different bonding behaviors as already suggested in the preceding section.

Empty La $4f$ states are mainly found above E_F within the conduction band –CB– where the low occupied Ce $4f$ band is centered. The large magnitude of the Ce DOS at E_F signals instability of the Ce compounds in such a NSP configuration, meaning that stabilization should incur if a spin polarized SP configuration is considered. Subsequent SP calculations lead, assuming implicitly a ferromagnetic configuration (SPF), indeed to the stabilization of the Ce based compounds and the onset of finite magnetic moments.

Table 5 provides the values of energy differences between the considered magnetic configurations and the spin only magnetic moment on Ce. For CeMgGa the energy difference between the two magnetic configurations is $\Delta E_{\text{SPF-NSP}} = -0.172$ eV/cell. This value decreases down to -0.046 eV for $X = \text{Pd}$. The SPF configuration is stabilized with respect to NSP due to magnetic exchange which is then larger magnetic $X = \text{Ga}$ case leading to larger magnitude of the Ce moment. Small moments are obtained for the other constituents resulting from some residual polarization due to the mixing of the valence states with those of magnetically polarized Ce. In view of the fact that both CeMgX order antiferromagnetically at $T_N = 3.1$ 2.1 K respectively [30], further SP antiferromagnetic (SPAF) calculations were carried out by doubling the cell along c so that the atoms belonging to the first half of the double cell are assigned SPIN UP magnetic

subcell and the other half SPIN DOWN magnetic subcell. For CeMgGa the calculations lead to further stabilization $\Delta E_{\text{SPAF-SPF}} = -0.04$ eV/cell whereas similar SPAF calculations carried on CeMgPd led to a vanishing of the magnetization. This can be assigned to its smaller T_N value together with the lower magnitudes of the magnetization in the SPF configuration with respect to CeMgGa. This likely leads to instability of the heavier SPAF calculations with subsequent vanishing magnetic polarization on the atomic constituents, mainly Ce.

The SPF and SPAF results are further illustrated with the site and spin projected DOS shown in Fig. 4.

In the first two panels corresponding to SPF calculations, the majority spin (\uparrow) are shifted to lower energy versus minority spin states (\downarrow). This energy difference, related to magnetic exchange, leads to the onset of the magnetic moment which is, as described above, carried only by Ce. Within the valence bands of both compounds, very small intensity Ga, Pd d states and itinerant Mg and Ce states mix in a similar way as in the NSP state without any energy shift between the two spin channels (\uparrow and \downarrow). For Ce, it can be noted the larger energy shift in CeMgGa as with respect to CeMgPd, proportional to the magnitudes of the respective magnetic moments and total magnetizations (Table 5). The third panel of Fig. 4 shows the fully compensated magnetizations between the two magnetic subcells ‘SPIN UP’ and ‘SPIN DOWN’, signature of the SPAF configuration of CeMgGa.

3.2.3 Properties of chemical bonding (COOP)

Through exploring the energy quantities (Table 3) it was found out that the Pd based intermetallics were significantly more cohesive versus the Ga ones. This is also accompanied by larger hardness. These observations let suggest a role played by the chemical bonding accounting for pair interactions. Such properties can be qualitatively described based on the analysis of the overlap integrals S_{ij} within the crystal orbital overlap population (COOP) introduced by Hoffmann in quantum chemical methods [45] and implemented into the ASW method [44].

Figure 5 shows the COOP for $X = \text{Ga}$ (top panels) and $X = \text{Pd}$ (bottom panels). The calculations are for NSP configuration. Whereas in all panels the VB is of positive CO-

OP intensities, i.e. of bonding nature, the antibonding states starts developing above E_F within the CB. This comforts the cohesive nature of all compounds (Table 3).

Differences can be observed at two different levels, pertaining:

- 1) to the nature of X d which binds either with p states (Ga) in the close neighborhood of E_F or with more intense and localized (directional) d states (Pd) at ~ -4 eV below E_F ; and
- 2) to the nature of the $4f$ element La or Ce whereby the latter has bonding sharp COOP at E_F whereas La shows broader COOP arising mainly from d states, a character missing with the Ga itinerant states.

The major bonding contribution arising from (La,Ce)-Ga and (La,Ce)-Pd and the latter being more directional one may understand the larger bulk modulus of the Pd based compounds (Table 1).

4. Conclusions

In this work aiming at better rationalizing the properties of a rich family of Mg based intermetallic compounds: LaMgX and CeMgX (X = triel or transition metal), DFT-built methods were applied to four representatives based on Ga and Pd. From charge transfer trends, the compounds are found as ‘gallides’ and ‘palladides’ with stronger cohesion and larger hardness characterizing the latter family. Magnetic polarization is found only in the Ce based compounds with a calculated antiferromagnetic ground state in agreement with experiment. Qualitative chemical bonding analysis shows different behaviors between the Ga and Pd families with a role played by Pd assigned to directional d states versus Ga broad s,p and subsequent stronger bonding for the former. This further illustrates the calculated cohesive energy and bulk module trends.

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Table 1

Lattice parameters of equiatomic compounds LaMgX and CeMgX ($X = \text{Ga, In, Tl, Pd, Ag, Pt, Au}$) with hexagonal ZrNiAl type structure.

Compound	a / pm	c / pm	V / nm ³	c/a	Reference
trial intermetallics					
LaMgGa	760.2(2)	459.6(1)	0.2300	0.605	[25]
LaMgIn	782.6(2)	481.08(9)	0.2552	0.615	[26]
LaMgTl	781.3(1)	477.84(8)	0.2526	0.612	[27]
CeMgGa	752.7(2)	454.8(1)	0.2232	0.604	[25]
CeMgIn	774.9(3)	477.7(2)	0.2484	0.616	[26]
CeMgTl	774.1(1)	473.75(7)	0.2458	0.612	[27]
transition metal intermetallics					
LaMgPd	773.6(1)	413.79(4)	0.2145	0.535	[30]
LaMgAg	788.8(3)	437.4(1)	0.2357	0.555	[29]
LaMgPt	762.0(1)	417.48(5)	0.2099	0.548	[30]
LaMgAu	781.0(1)	425.49(9)	0.2248	0.545	[30]
CeMgPd	767.3(1)	410.37(4)	0.2092	0.535	[30]
CeMgAg	782.5(3)	432.8(1)	0.2295	0.553	[29]
CeMgPt	755.02(7)	413.82(4)	0.2043	0.548	[30]
CeMgAu	774.1(3)	421.6(1)	0.2188	0.545	[30]

Table 2

Interatomic distances (pm), within the first coordination spheres of CeMgIn, CeMgTl, CeMgAg, and CeAuMg.

CeMgIn [26]			CeMgTl [27]			CeMgAg [29]			CeAuMg [30]		
Ce:	4	In1 331.5	Ce:	1	Tl2 328.5	Ce:	4	Ag1 319.7	Ce:	4	Au1 314.2
	1	In2 334.6		4	Tl1 330.9		1	Ag2 325.4		1	Au2 321.2
	2	Mg 347.3		2	Mg 349.4		2	Mg 345.4		2	Mg 339.7
	4	Mg 376.1		4	Mg 371.0		4	Mg 356.2		4	Mg 350.0
	4	Ce 398.1		4	Ce 400.1		4	Ce 407.6		4	Ce 403.5
	2	Ce 477.7		2	Ce 473.8		2	Ce 432.8		2	Ce 421.6
In1:	3	Mg 299.6	Tl1:	3	Mg 298.7	Ag1:	3	Mg 303.9	Au1:	3	Mg 300.2
	6	Ce 331.5		6	Ce 330.9		6	Ce 319.7		6	Ce 314.2
In2:	6	Mg 304.1	Tl2:	6	Mg 302.9	Ag2:	6	Mg 286.6	Au2:	6	Mg 281.5
	3	Ce 334.6		3	Ce 328.5		3	Ce 325.4		3	Ce 321.2
Mg:	2	In1 299.6	Mg:	2	Tl1 298.7	Mg:	2	Ag2 286.6	Mg:	2	Au2 281.5
	2	In2 304.1		2	Tl2 302.9		2	Ag1 303.9		2	Au1 300.2
	2	Mg 325.9		2	Mg 327.0		2	Mg 325.6		2	Mg 323.1
	2	Ce 347.3		2	Ce 349.4		2	Ce 345.4		2	Ce 339.7
	4	Ce 376.1		4	Ce 371.0		4	Ce 356.2		4	Ce 350.0

Table 3

Calculated parameters of equiatomic compounds LaMgX and CeMgX ($X = \text{Ga}, \text{Pd}$): Total and cohesive energies per unit cell, volumes and bulk moduli are in eV, nm³ and GPa units respectively.

Compound	E_{total}	$E_{cohesive}$	$V_{exp.}(calc.)$	B_0
LaMgGa	-32.36	-1.56	0.230 (0.232)	49.8
CeMgGa	-31.77	-1.86	0.223 (0.229)	51.4
LaMgPd	-40.87	-2.13	0.214 (0.218)	60.5
CeMgPd	-40.11	-2.38	0.209 (0.215)	62.3

Energies per atom of constituents: $E(\text{Ce}) = -4.43$ eV; $E(\text{La}) = -4.92$ eV; $E(\text{Ga}) = -2.91$ eV; $E(\text{Mg}) = -1.39$ eV; $E(\text{Pd}) = -5.17$ eV

Table 4

Calculated parameters of equiatomic compounds LaMgX and CeMgX ($X = \text{Ga}, \text{Pd}$): Bader charges per atom in e units.

Compound	$Q(\text{La/Ce})$	$Q(\text{Mg})$	$Q(X1)$	$Q(X2)$
LaMgGa	1.05	1.13	-2.53	-2.01
CeMgGa	1.07	1.14	-2.03	-2.56
LaMgPd	1.16	1.14	-2.86	-2.02
CeMgPd	1.19	1.15	-2.87	-2.08

Table 5

Calculated parameters of equiatomic CeMgGa and CeMgPd in ferromagnetic configurations (not ground state). Energies are in eV/cell and Ce magnetic moments (spin-only) in μ_B .

CeMgX	$\Delta E_{\text{SPF-NSP}}$	m_{Ce}	$m_{\text{MgX1/X2}}$	$M(\text{cell})$
$X = \text{Ga}$	-0.176	0.841	0.04/0.004/0.01	2.674
$X = \text{Pd}$	-0.046	0.595	0.02/0.019/0.02	1.922

FIGURE CAPTIONS

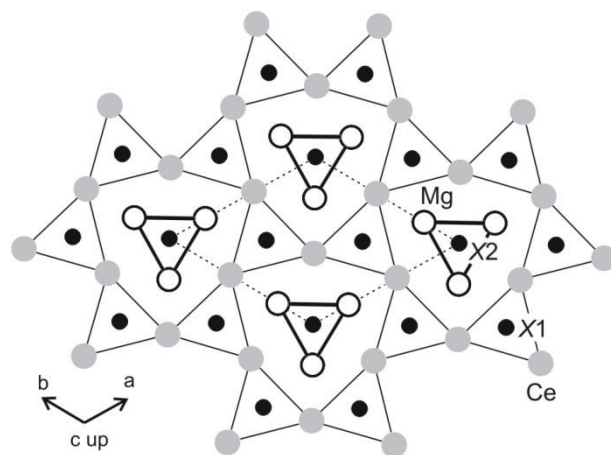
Fig. 1. Projection of the CeMgX structure onto the xy plane. Cerium, magnesium, and X atoms are drawn as medium gray, black open, and filled circles, respectively. All atoms lie on mirror planes at $z = 0$ (thin lines) and $z = 1/2$ (thick lines). The trigonal prisms around the X atoms are emphasized.

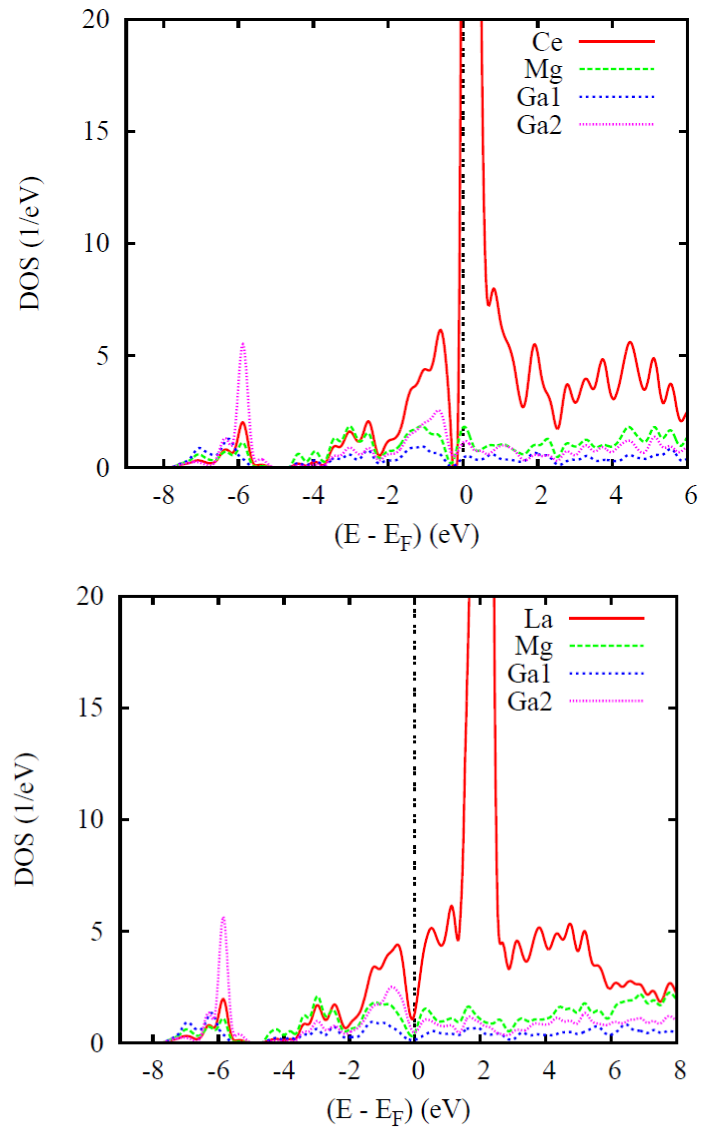
Fig. 2. Triel intermetallics with $X = \text{Ga}$ representatives in non magnetic total spin configuration (NSP). CeMgGa (top) and LaMgGa (bottom).

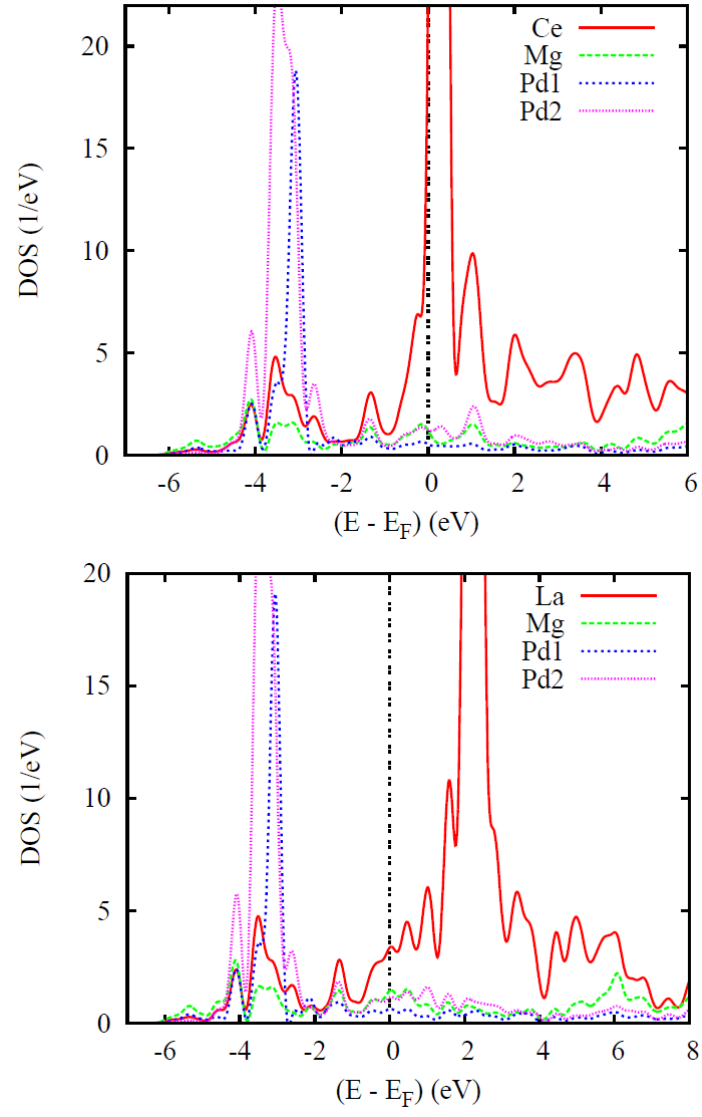
Fig. 3. Transition metal intermetallics with $X = \text{Pd}$ representatives in non magnetic total spin configuration (NSP). CeMgPd (top) and LaMgPd (bottom).

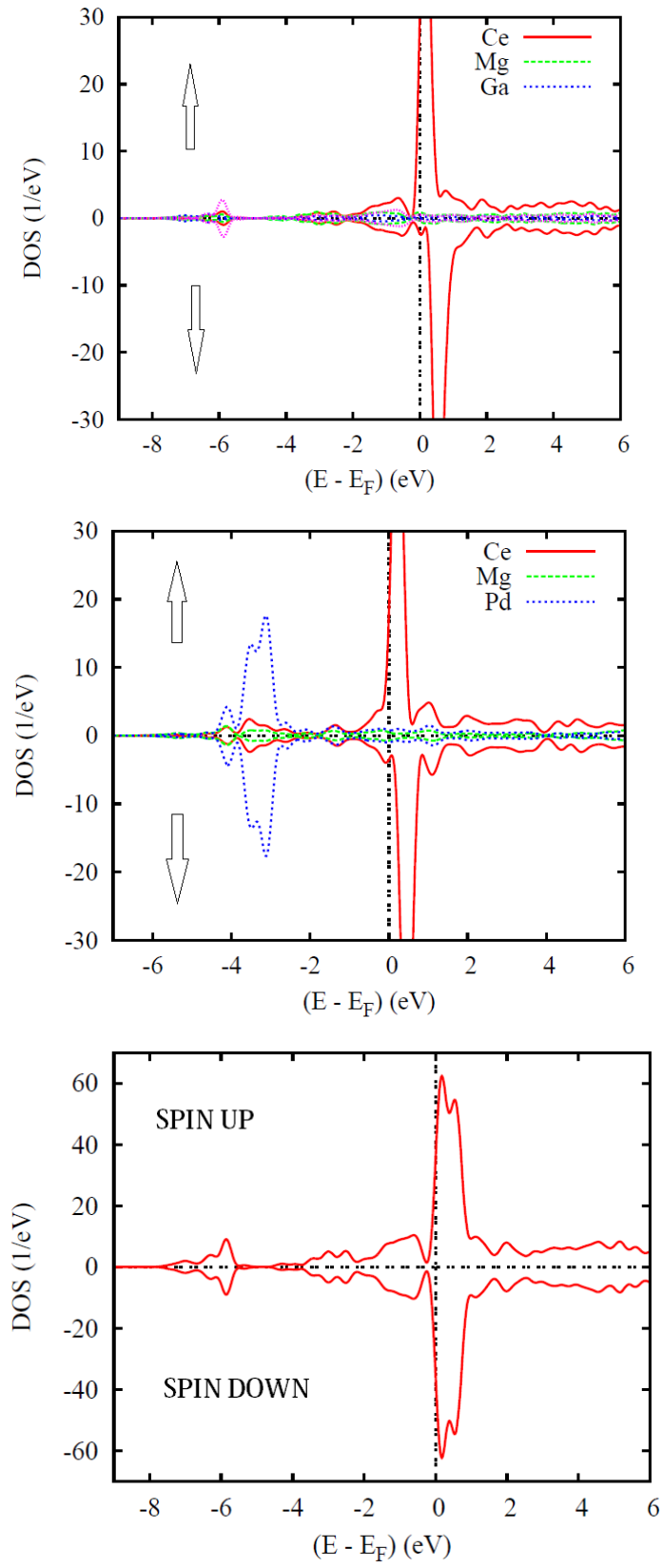
Fig. 4. Site and spin projected density of states (DOS) in SPF configuration of CeMgGa (top) and CeMgPd (middle) and SPAF total DOS for SPAF CeMgGa (bottom).

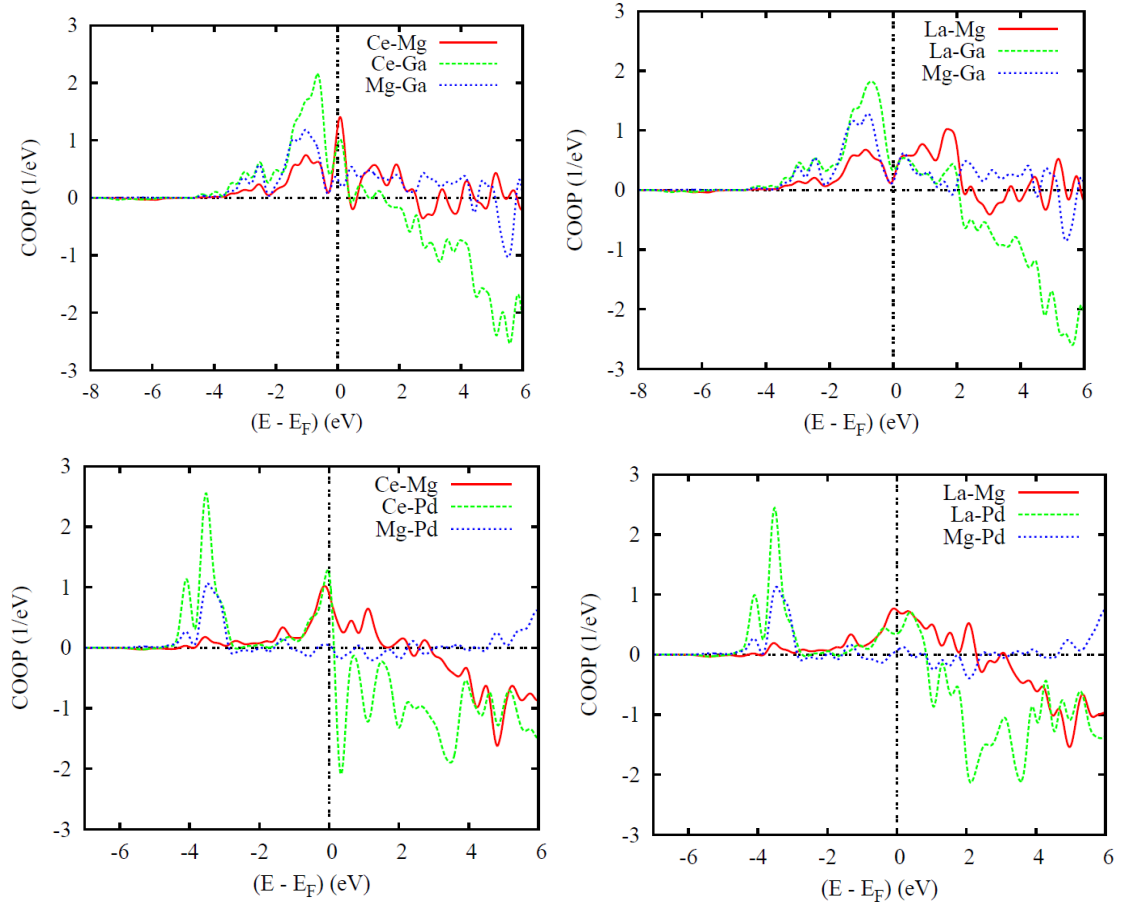
Fig. 5. Chemical bonding for $X = \text{Ga}$ (top) and $X = \text{Pd}$ (bottom) in CeMgX and LaMgX.

**Fig. 1.**

**Fig. 2.**

**Fig. 3.**

**Fig. 4.**

**Fig. 5**