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Mechanism and kinetics of the α – β transition in San Carlos olivine $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$

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[1] The mechanism and kinetics of the α – β transformation in San Carlos olivine α - $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$ containing ca. 526 wt.ppm H_2O were studied in situ by time-resolved synchrotron X-ray diffraction (XRD) in the range 12.5 – 14.5 GPa and 870 – 1150 K. Time series of XRD spectra reveal the early formation of an intermediate phase that further transforms to wadsleyite (β -phase). This intermediate phase may be either a metastable ringwoodite (γ -phase) or the disordered spinelloid phase proposed in the pseudomartensitic reaction model. Both microstructural and kinetic data suggest that the transformation rates are controlled by growth processes after the early saturation of nucleation sites along olivine grain boundaries. Growth rates in the range 1.7×10^{-12} – 2.1×10^{-9} m.s^{-1} are estimated by fitting the transformation–time data to the Cahn rate equation and define an activation enthalpy for growth $\Delta H_a = 204(55)$ kJ/mol. Compared to the previous studies performed on pure forsterite Mg_2SiO_4 , these new kinetic results point out an increase of transformation rates with the iron content. Hence, the persistence of a metastable olivine wedge to depths > 660 km might be restricted to even colder and/or dryer subduction zones than previously estimated.

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1. Introduction

[2] Olivine $(\text{Mg,Fe})_2\text{SiO}_4$ (α -phase) is the major constituent of the Earth’s upper mantle. With increasing pressure, it transforms sequentially to wadsleyite (β -phase), with a modified spinel structure, then to ringwoodite (γ -spinel phase), and finally to perovskite + magnesio-wustite [e.g., *Ringwood and Major*, 1966; *Akimoto and Fujisawa*, 1968]. These phase changes have long been associated with the seismic discontinuities observed at 410, 520, and 660 km depth in the mantle, respectively [e.g., *Ringwood*, 1975]. In the cold interior of subducting plates, these phase transformations can be kinetically delayed [e.g., *Stein and Stein*, 1996], and olivine could persist metastably well beyond its equilibrium limit. Such a kinetic hindrance is of significant importance for the buoyancy and ultimate fate of subducting slabs. Because subduction is driven by the thermal structure and density increase related

to phase transformations, the existence of a metastable wedge of olivine would profoundly influence the plate velocity, and ultimately the dynamics and morphology of slabs [e.g., *Bina et al.*, 2001]. The mechanical properties of lithospheric materials would also be altered, as the stress state and the strength of slabs are strongly dependent on mineralogy and grain sizes [e.g., *Karato et al.*, 2001]. Finally, shear instabilities postulated to occur during the transformation of metastable olivine to its high pressure polymorphs have been proposed for the origin of deep focus earthquakes [*Kirby*, 1987; *Green and Burnley*, 1989], as well as “outboard” earthquakes that occur away from Wadatti–Benioff zones [*Chen and Brudzinski*, 2001].

[3] Direct seismic detections of a metastable olivine wedge are so far scarce. Examples of such observations include the Japan [*Kawakatsu and Yoshioka*, 2011] and the Mariana subduction regions [*Kaneshima et al.*, 2007]. An alternative approach is to build thermo-kinetic models relying on accurate experimental data on the mechanism and kinetics of the α – β – γ transitions. The extensive work conducted on the transformation mechanism led however to conflicting results. A martensitic transition was proposed theoretically by *Poirier* [1981] based on the coherent shear of the olivine oxygen sublattice accompanied by “synchroshear” of the cations. The topotaxial relationship between the olivine and spinel phases $(\text{Mg,Fe})_2\text{SiO}_4$ observed in either experiments [*Lacam et al.*, 1980; *Madon et al.*, 1989; *Green and Burnley*, 1989] or shocked meteorites [*Price et al.*, 1982] support this shear-induced mechanism that would be favored by high differential stress and/or a large overstepping

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of the equilibrium conditions. In contrast, numerous experimental studies have documented microstructural evidences for a reconstructive transition involving grain-boundary or intracrystalline nucleation followed by interface-controlled growth [Sung and Burns, 1976; Brearley et al., 1992; Kerschhofer et al., 1996; Kubo et al., 1998, Mosenfelder et al., 2001]. This mechanism is thought to be effective under near-equilibrium and high-temperature conditions. Finally, an intermediate model is the pseudomartensitic mechanism that involves two stages: a shear-induced anionic re-arrangement followed by short-range diffusion of cations. This mechanism was first proposed by Furnish and Basset [1983] who observed during in situ X-ray diffraction (XRD) experiments that the coherent restacking of oxygen atoms appreciably precedes cations reordering during the α - γ transition in Fe_2SiO_4 . It is consistent with the transmission electron microscopy (TEM) observations of a disordered spinelloid phase that subsequently orders to produce the β -phase in Mg_2SiO_4 quench products [Guyot et al., 1991; Brearley et al., 1992]. From the above literature data, it appears that the transformation mechanism might change with pressure, temperature, and stress conditions. However, the data are not yet comprehensive to infer the mechanism that operates under subduction environments.

[4] Early experimental kinetic studies have been carried out at low pressure on analogue materials Mg_2GeO_4 , Ni_2SiO_4 , and Co_2SiO_4 (see Rubie and Ross [1994] for a review). However, the application to mantle olivine requires the extrapolation of kinetic parameters assuming that composition and absolute pressure do not affect transformation rates. More recently, kinetics data have been obtained on the α - β and α - γ transformations in synthetic Mg_2SiO_4 and natural $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$ compositions and conditions relevant to the upper mantle [Kubo et al., 1998, 2004a; Liu et al., 1998; Kerschhofer et al., 2000; Mosenfelder et al., 2001; Hosoya et al., 2005; Diedrich et al., 2009]. Most of these studies relied on quenched experiments performed in multi-anvil devices, aside from Kubo et al. [2004a] and Hosoya et al. [2005] who studied the transition in situ by XRD. These works concluded that overall transformation rates are controlled by growth kinetics rather than nucleation processes and suggested that the persistence of olivine to a depth of ~ 600 – 700 km is restricted to the cold interior of rapidly subducting slabs. However, they also pointed out the effect of transformation stress [Liu et al., 1998; Mosenfelder et al., 2001] and water [Hosoya et al., 2005; Diedrich et al., 2009] that could profoundly affect the extent of olivine metastability. This last point could be exceedingly important in hydrous subducted slabs as both olivine and its high-pressure polymorphs can contain significant H_2O in the form of hydroxyl point defects [e.g., Bolfan-Casanova et al., 2000].

[5] In this paper, we provide new kinetic data on the olivine-wadsleyite transition in a San Carlos mantle composition acquired by the in situ XRD method. The experiments were performed at relatively low-temperature conditions (870–1150 K) and small pressure oversteps of the transition boundary in order to reproduce the transformation in subduction zone environments. Time series of XRD spectra reveal the early formation of an intermediate metastable phase that further transforms to wadsleyite. Kinetic parameters derived from the experimental transformation–time data are compared with literature data and highlight an increase of transformation

rates with the iron content. We discuss the implications of these new kinetic results on the metastability of olivine in subducting slabs.

2. Experimental Methods

[6] The starting material consisted of natural olivine crystals from the San Carlos peridotite (Arizona), ground in an agate mortar down to 2–3 μm in size. The composition of these crystals was $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$ as determined by electron probe measurements. The water content was measured by infrared spectroscopy to 1.8 ± 0.4 wt.ppm H_2O using the calibration coefficient of Libowitzky and Rossman [1997]. Prior to the experiments, powders were dried at 150°C in air for few days to avoid absorption of water. Angle-dispersive XRD experiments were carried out in a Paris-Edinburgh large volume press installed at beamline ID27 of European Synchrotron Radiation Facility (Grenoble, France) [Mezouar et al., 2005; Perrillat, 2008]. Sintered diamond anvils [Morard et al., 2007] were used to generate pressures up to 14.5 GPa. The sample assembly (Figure 1) was composed of a boron-epoxy pressure medium, a resistive heater, and an MgO sample capsule. The heater consisted of two LaCrO_3 ceramic disks in electrical contact through two rhenium strips 50 μm thick, oriented parallel to the X-ray beam. The olivine and a thin layer of Au powder were packed in the MgO capsule. The maximal thermal gradient across the sample was estimated from numerical modelling to ± 80 K at 873 K (Figure 1).

[7] Pressure and temperature conditions were determined from the diffraction lines of MgO and Au using the PVT equations of state of Martinez-Garcia et al. [2000] and Shim et al. [2002], respectively. Uncertainty on P and T depends on both the error in fitting the diffraction data and the uncertainty on the thermoelastic parameters of the calibrants. It was estimated for each P–T conditions from the parallelogram area delineated by the intersection between the Au and MgO isochors (Figure 2). Taking into account a ± 0.002 Å error on the distance of the Au and MgO diffraction lines, the precision on P and T was in all cases better than ± 0.4 GPa and ± 80 K.

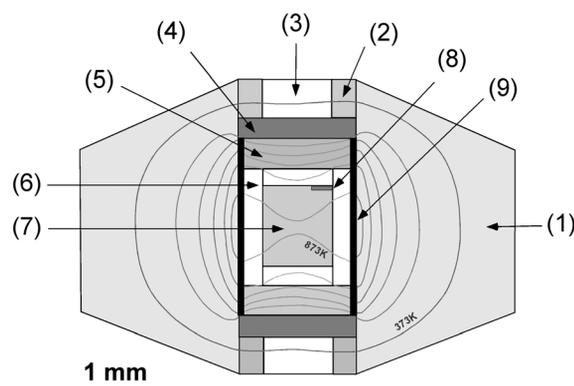


Figure 1. Details of the cell assembly and temperature profile at 873 K calculated using the program “CellAssembly” [Hernlund et al., 2006]. The interval between isotherms is 100 K. (1) boron-epoxy gasket, (2) steel ring, (3) MgO ceramic, (4) molybdenum disk, (5) LaCrO_3 ceramic, (6) MgO capsule, (7) sample, (8) gold foil, (9) rhenium heater.

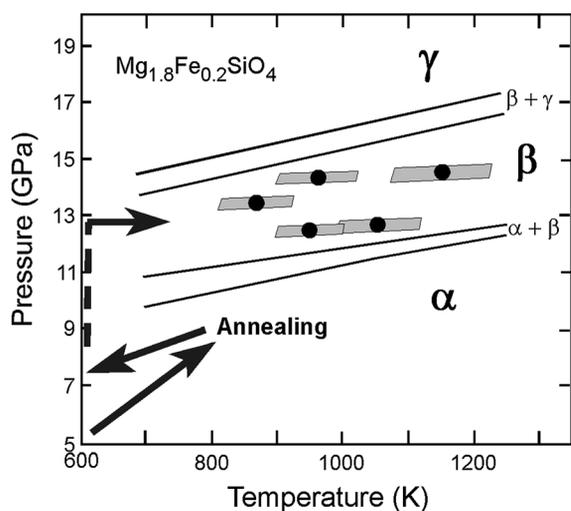


Figure 2. Pressure–temperature diagram showing the experimental conditions with relations between olivine (α), wadsleyite (β), and spinel (γ) phases (after *Akaogi et al.* [1989] and *Katsura et al.* [2004]). P and T were determined from the diffraction lines of Au and MgO using a cross-calibration of their PVT equation of state; uncertainties are displayed as gray-shaded area.

[8] Samples were first compressed at room temperature to about 9 GPa and then annealed at 1100 K for 2–3 h within the stability field of the α -phase. The goal of this annealing is to relax the strain and reduce the density of defects resulting from the cold compression stage. The sharpening of diffraction lines on the XRD images testifies to these recrystallization processes. The achievement of equilibrium microtexture is clearly visible on the scanning electron microscopy (SEM) picture of a sample quenched after annealing (Figure 3a). It exhibits sharp olivine grain boundaries and 120° triple junctions. Following the annealing stage, temperature was reduced to ambient, and pressure was increased to the stability field of the β -phase. Then, temperature was increased again at a heating rate of 50 K/min up to the target value. To avoid a decrease in pressure related to the volume reduction of the reaction, a constant load was applied on the sample thanks to an automatic oil pressure drive system (Sanchez Technologies[®]). To monitor the extent of transformation as a function of time, a sequence of XRD spectra was acquired every 30 s to 4 min, with a collection time of 15 to 60 s. The incident monochromatic X-ray beam $\lambda=0.6199 \text{ \AA}$ (Mo absorption K-edge) was reduced down to $50 \times 50 \mu\text{m}^2$ by two sets of carbide slits placed before the Paris-Edinburgh press. The diffracted X-rays from the sample were collimated through Soller’s slits [*Mezouar et al.*, 2002] and collected on a MAR[®]345 image plate over a 2θ interval from 4° to 18° . The two-dimensional XRD images were integrated after spatial distortion corrections using Fit2D [*Hammersley et al.*, 1996]. Finally, transformation–time plots $\zeta(t)$ were constructed by Rietveld refinement of phase proportions using the GSAS package [*Larson and von Dreele*, 1994]. After completion of the reaction, i.e., disappearance of the diffraction lines of the α -phase, the samples were quenched to room temperature by switching off the power of the heater and then slowly decompressed. The recovered samples were cut and mounted on epoxy resin and polished for microscopic observations of texture and grain

sizes. SEM images were acquired on a JEOL 840 microscope at the MATEIS laboratory (Lyon, France).

[9] Considering the potentially strong catalytic effects of even small amounts of H_2O in the sample charges, the water concentrations were determined using Fourier transform infrared spectroscopy (FTIR) measurements. For this purpose, recovered samples were ground and doubly polished to $\sim 130 \mu\text{m}$ thin sections. The samples thickness was determined with a precision of approximately $2 \mu\text{m}$ with a Mitutoyo electronic micrometer. Unpolarized infrared spectra were obtained in air using a Vertex70 Bruker spectrometer coupled to a Hyperion microscope equipped with $\times 15$ objective and condenser (at Laboratoire Magmas et Volcans, Clermont-Ferrand – France, and Institut de Planétologie et Astrophysique, Grenoble – France). Beam size ranged between 30 and $100 \mu\text{m}$. The spectra were measured through a CaF_2 plate with a resolution of 2 cm^{-1} and with up to 4000 scans between 2500 and 4000 cm^{-1} in the O–H stretching region.

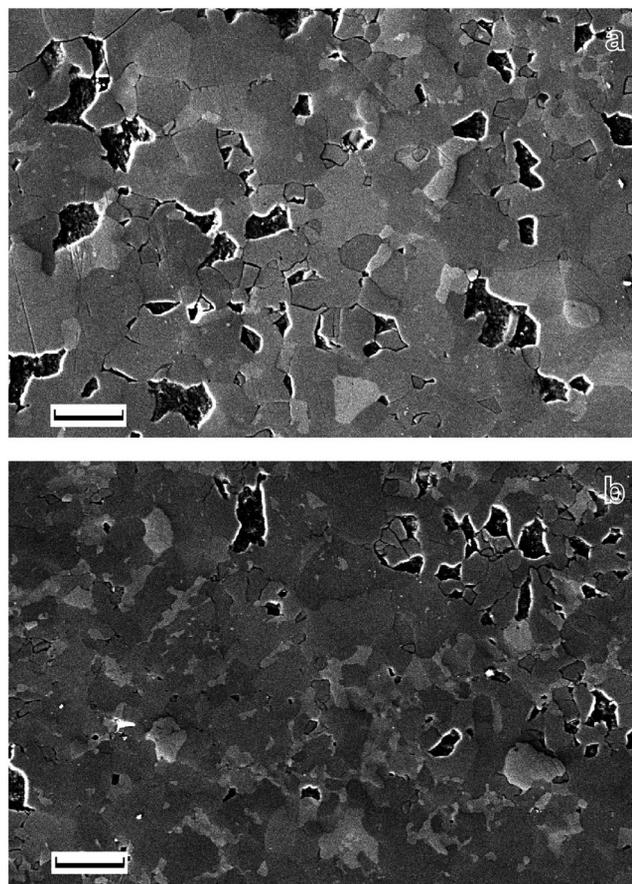


Figure 3. Electron backscattered images showing the characteristic microtextures of samples (a) after the annealing stage at 9.6 GPa, 1100 K during 3 h and (b) after the complete transformation to the β -phase after 6 h at 14.5 GPa, 980 K. For a better display, the gray scale has been stretched, and thereby contrasts reflect crystal orientation. The scale bar is 4 microns. The annealed sample shows a relatively homogeneous grain size of $2.5 \pm 1.1 \mu\text{m}$, well-developed triple junctions, and straight grain boundaries, whereas after reaction, it exhibits scalloped grain boundaries and a smaller grain size of $1.4 \pm 0.6 \mu\text{m}$.

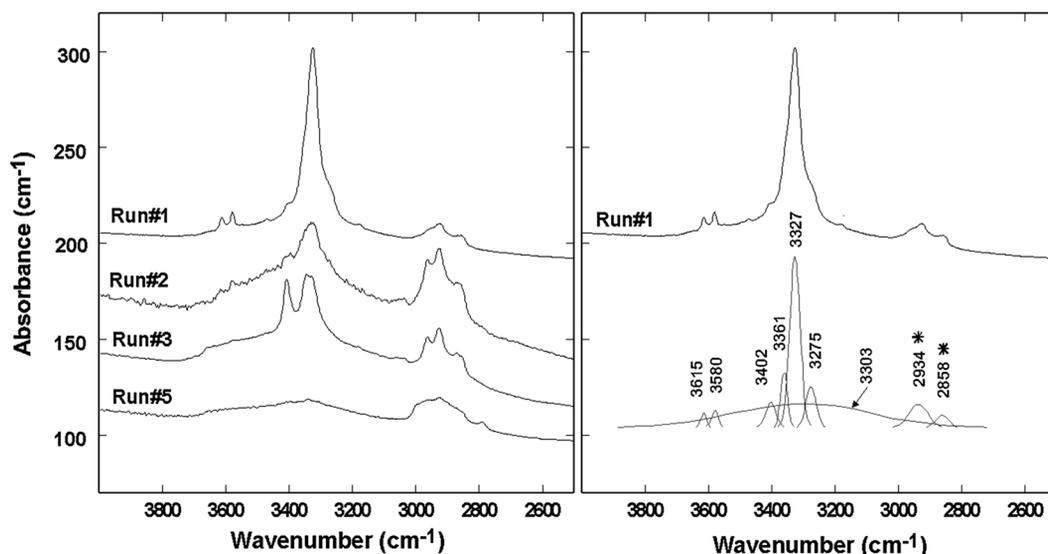


Figure 4. Unpolarized IR spectra of recovered samples in the region of the OH stretching modes. Representative absorption bands are labeled by their wavenumber obtained by peak fitting. See section 2 for the assignment of IR peaks and quantification of hydroxyl groups. Peaks from the epoxy glue and acetone treatment are marked with asterisks.

[10] All FTIR spectra (Figure 4) exhibit a broad absorption band in the range 3000 to 4000 cm^{-1} , attributable to molecular H_2O , which likely reflects the presence of water at grain boundaries. A group of four sharp absorption peaks near 3300 cm^{-1} (i.e., 3275, 3326, 3360, and 3402 cm^{-1}) and a second group of two smaller peaks near 3600 cm^{-1} (i.e., 3580 and 3615 cm^{-1}) are characteristic of the structural hydroxyls in wadsleyite as reported by *Kohlstedt et al.* [1996] and *Kohn et al.* [2002]. Finally, absorption bands near 2900 cm^{-1} , which are characteristic of C–H bending, eventually derive from the epoxy glue and acetone treatment and therefore were not considered for the determination of OH contents.

[11] The hydroxyl contents were calculated from the integrated area of both sharp and broad absorption bands after background baseline correction and thickness normalization to 1 cm. The molar absorption coefficient $\epsilon_i = 73,000 \text{ L. mol}^{-1} \text{ cm}^{-2}$ [Deon *et al.*, 2010] was used for the quantification of wadsleyite IR peaks, while the general calibration of *Libowitzky and Rossman* [1997] was used for the broad band. The OH content in each sample (Table 1) varies from

347 to 596 wt.ppm H_2O with an average value of 526 ± 119 wt.ppm H_2O . Though still low, the amount of water in wadsleyite is significantly higher than in the original olivine. It probably arises from water adsorbed in the cell assembly parts and powdered sample, and/or from the boron-epoxy medium.

3. Results

3.1. Mechanism of the Olivine–Wadsleyite Transition in $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$

[12] Time-resolved XRD data were acquired in five runs over the 12.5–14.5 GPa and 870–1150 K P–T range (Figure 2 and Table 1). A typical sequence of diffraction patterns is shown in Figure 5. At the beginning of the transition, the decrease in intensity of olivine diffraction peaks is accompanied by the early appearance of a metastable intermediate phase whose diffraction lines are successfully indexed with the (220) (311) (400) (511) and (440) reflections of $Fd\bar{3}m$ spinel. This phase transforms progressively into wadsleyite, $Imma$ space group, and at the end of the reaction, only the diffraction

Table 1. Experimental Conditions, Water Contents, and Estimated Values of Grain-Boundary Nucleation and Growth Rates

	V_{Au} (Å)	V_{MgO} (Å)	Pressure (GPa)	Temperature (K)	Free Energy Change ΔG_r (J/mol)	Water Content (wt ppm H_2O)	Nucleation Rate (m^{-2}/s)	Growth Rate (m/s)
Run #1	64.55	69.94	14.5	980	−7017	571 (111)	−	$1.9 (0.2) \times 10^{-11}$
Run #2	65.07	70.59	12.5	950	−2093	590 (29)	$4.83 (0.5) \times 10^{13}$	$6.9 (0.6) \times 10^{-12}$
Run #3	64.68	70.22	13.2	870	−4401	596 (105)	−	$1.7 (0.8) \times 10^{-12}$
Run #4	65.24	70.62	12.8	1050	−2361	−	−	$1.2 (0.4) \times 10^{-10}$
Run #5	64.97	70.12	14.5	1150	−6983	347 (106)	−	$2.1 (0.3) \times 10^{-09}$

[†]Pressure and temperature conditions were determined from the diffraction lines of MgO and Au using the PVT equations of state of *Martinez-Garcia et al.* [2000] and *Shim et al.* [2002], respectively.

[‡]The free energy change ΔG_r at P–T is calculated from the thermodynamic data of *Akaogi et al.* [1989] and *Jacobs and de Jong* [2005].

[§]The FTIR measurements of OH contents were repeated at five to nine locations in each sample. The uncertainty in the water concentration is one standard deviation. Sample Run#4 was lost in the grinding process.

^{||}Nucleation and growth rates were obtained by a fit of the experimental $\xi(t)$ data to Cahn’s equation (1). Estimated uncertainties are listed in parentheses. When a solution for N was not achieved, an arbitrary value of $N = 10^{14} \text{ m}^{-2}/\text{s}$ was set.

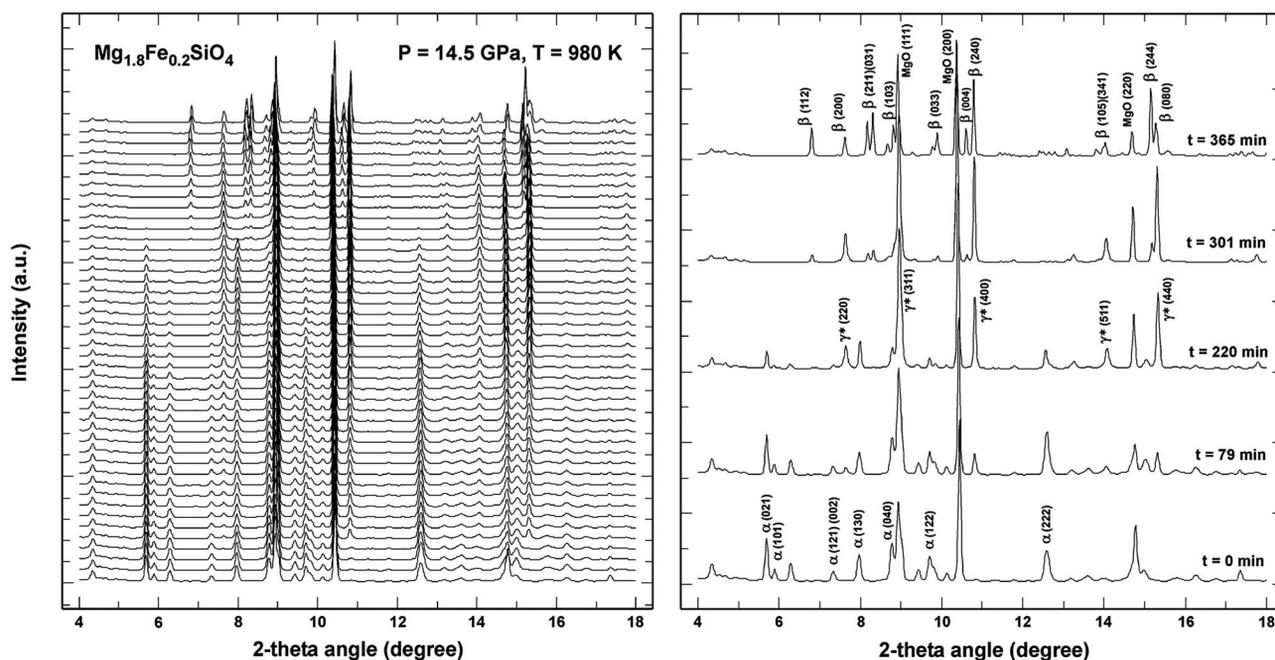


Figure 5. (a) Time series of XRD spectra acquired at 14.5 GPa – 980 K during the α - β transformation and (b) selected patterns at 0, 79, 220, 301, and 365 min. (hkl) diffraction lines are labeled as MgO: magnesium oxide, α : olivine, β : wadsleyite, γ^* : intermediate phase.

lines of the β -phase and MgO capsule are present. This succession of phases was observed in all experiments, independently of the overpressure relative to equilibrium.

[13] The origin of this metastable phase may be twofold. On one hand, it may correspond to ringwoodite that nucleated and grew metastably according to the Ostwald rule. This could have occurred if the free energy sequence is $G_\alpha > G_\gamma > G_\beta$, so that ringwoodite is stable relative to olivine. On the other hand, this phase may be the intermediate spinelloid phase involved in the pseudomartensitic reaction mechanism [Chen *et al.*, 2001]. In the latter model, the restacking of the anions transforms the hcp oxygen arrangement of olivine into the fcc oxygen sublattice of spinel and modified spinel. It results in a spinelloid phase with a disordered cation structure that further rearranges by short-range diffusion in a second step, to produce the β -phase. A detailed description of this intermediate phase can be found in Guyot *et al.* [1991]. Because of the random position of the cations, the spinelloid phase should exhibit more intense diffraction lines for the oxygen array than for the cations one [Furnish and Basset, 1983; Chen *et al.*, 2001]. Unfortunately, this feature cannot be detected on the present XRD patterns, because of peak overlaps in the experimental phase mixture, i.e., the most intense (311) line of spinel is superimposed with the strong α -phase (112) and MgO (111) peaks. However, the intermediate spinelloid phase has been clearly identified in the α - γ transition. Furnish and Basset [1983] in Fe_2SiO_4 reported the initial preferential growth of the only γ phase peaks compatible with the oxygen sublattice, namely the (400) and (440) diffraction lines. Lauterjung and Will [1986] observed that at the beginning of the α - γ transition, in the Mg_2GeO_4 analogue, only diffraction peaks from a spinel-type oxygen array were present. Moreover, Chen *et al.* [2001] performed structure refinements of XRD spectra during the α - γ transition in fayalite and concluded to a delay in the

reordering of cations relative to anions during the phase transformation. We therefore believe, though there is no direct evidence in the present XRD pattern, that the observed intermediate phase can be regarded as a disordered spinelloid and that the pseudomartensitic mechanism operates throughout the Mg_2SiO_4 - Fe_2SiO_4 binary.

[14] The pseudomartensitic mechanism is also consistent with many textural features observed in quenched experimental products and shocked meteorites. Guyot *et al.* [1991] studied by TEM partially transformed Mg_2SiO_4 samples and evidenced a structurally disordered spinelloid phase as an intermediate of the α - β reaction. The lamellar intergrowths of β and γ phases commonly observed in Mg_2SiO_4 [Brearley *et al.*, 1992], in analogue compounds [Burnley *et al.*, 1995], and in shocked meteorites [Price *et al.*, 1982; Madon and Poirier, 1983] might also be interpreted as a result of the incomplete transformation of the metastable spinelloid into the β -phase.

[15] The pseudomartensitic transition has often been considered as favored under high-stress conditions [e.g., Burnley *et al.*, 1995]. In the present experiment, however, PT conditions were close to equilibrium, and the amount of differential stress, estimated from the method of Singh and Kennedy [1974] on MgO diffraction lines, was not detectable above 600 K, i.e., $\sigma < 0.05 \text{ GPa}$. This is in agreement with the conclusions by Raterron *et al.* [2002] that the pseudomartensitic transition operates also under quasi-hydrostatic conditions and is driven by the local stress generated by the volume reduction of the reaction.

[16] Microstructural SEM observations of the transformed samples (Figure 3b) give insights into the geometry of the nucleation and growth processes. In contrast to the texture of olivine after the annealing stage, the reacted samples show irregularly shaped and scalloped grain boundaries. The mean grain size decreased from 2.5(1.1) down to 1.4

(0.6) microns after the transition, providing evidence for a high nucleation rate relative to growth. No significant variation in grain size was observed among samples transformed at different PT conditions. Lamellar structures within grains are not observed. This suggests that the intermediate phase nucleates at olivine grain boundaries and triple junctions, and subsequently coarsens by phase-boundary migration. This conclusion is in good agreement with the interface-controlled nucleation and growth mechanism reported in many previous studies [e.g., *Brearley et al.*, 1992, *Kubo et al.*, 2004a; *Mosenfelder et al.*, 2001].

3.2. Transformation Kinetics

[17] The extent of transformation as a function of time $\zeta(t)$ was determined from the isothermal–isobaric sequence of XRD spectra by Rietveld refinement of phase proportions. The transformation–time plots presented in Figure 6 show the advance of reaction estimated from the decrease of the relative amount of olivine from $\zeta=0$ at the beginning of reaction to $\zeta=1$ after complete transformation into wadsleyite. The kinetic data were analyzed using the general rate equation describing the transformation by the grain-boundary nucleation and growth mechanisms [*Cahn*, 1956]:

$$\zeta(t) = 1 - \exp\left[-2S \int_0^{y'} \{1 - \exp(-z)dy\}\right] \quad (1)$$

where

$$z = \pi \int_0^{t-t'} N \left[G^2(t-\tau)^2 - y^2\right] dt \quad (2)$$

and N is the nucleation rate, G the growth rate of the product phase, $y'=Gt$ is the growth distance after time t , τ is the time at which a nucleus form, and t' is the time necessary for a nucleus to grow to radius y . S is the grain boundary area that was approximated by $S=3.35/d$, with d the mean grain size estimated by SEM on quenched samples. A fit of the experimental $\zeta(t)$ data to equation (1) provided a direct estimation of growth rates (Figure 6 and Table 1). In most runs, the solution for N did not converge because of extremely rapid nucleation, and only one high value of $N=4.83 \times 10^{13} \text{ m}^{-2} \cdot \text{s}^{-1}$ could be obtained, for run#2. The fast increase in ζ at the onset of reaction is actually consistent with such a high nucleation rate, as well as with microstructural data that revealed that the overall transformation rates were dominated by interface-controlled growth processes after the early saturation of nucleation sites. In all runs, the reaction went to completion, i.e., $\zeta > 0.9$ (excepted for run#2 that was stopped at $\zeta=0.7$), and the fit to $\zeta(t)$ data does not deviate significantly from the Cahn's relation. This is an evidence for steady-state growth rates during the transformation and for the absence of grain impingement or transformation stress [*Liu et al.*, 1998] that could potentially decrease the growth rates at the late stages of the reaction. It is in agreement with the conclusion of *Liu et al.* [1998] that transformation stress is minor when the reactant olivine is fine grained.

[18] The integration of the present kinetic data within a thermo-kinetic model requires the extrapolation of measured growth rates at any P–T conditions, assuming the transformation mechanism does not change. An expression of the

growth rate for the product phase in an interface-controlled transformation has been proposed by *Turnbull* [1956]:

$$G = k_0 T \exp\left(-\frac{\Delta H_a + PV^*}{RT}\right) \cdot \left[1 - \exp\left(-\frac{\Delta G_r}{RT}\right)\right] \quad (3)$$

where k_0 is a constant, T is temperature, ΔH_a is the activation enthalpy for growth, P is pressure, V^* is the activation volume for growth, R is the gas constant, and ΔG_r the free energy change of reaction at given P–T. Considering the limited pressure range investigated in the present study, the activation volume is not refined and is assumed to be zero. This assumption is justified by the low activation volume $V^*=1.7$ (4.5) cm^3/mol determined by *Kubo et al.* [2004a] for the α - β transition in Mg_2SiO_4 and $V^*=0$ (2) cm^3/mol estimated for the reconstructive growth of γ - Mg_2GeO_4 [*Burnley*, 1995]. The ΔG_r values are calculated from the thermodynamic data of *Akaogi et al.* [1989] and *Jacobs and de Jong* [2005] (Table 1). Values of k_0 and ΔH_a are estimated from the slope and intercept of a least squares linear fit to the growth rate data on an Arrhenius plot $\ln\{G/T[1-\exp(-\Delta G_r/RT)]\}$ vs $1000/T$ (Figure 7). The best fit yields values for $\ln k_0 = -5.2$ (8.1) m/s/K and $\Delta H_a = 204$ (55) kJ/mol , where uncertainties reported in parentheses are estimated from the propagation of the error on P–T conditions.

4. Discussion

4.1. Effect of the Iron and Water Contents on Kinetics

[19] The present kinetic results are compared to the parameters obtained by *Kubo et al.* [2004a] for the α - β transition in Mg_2SiO_4 synthetic samples. The latter experiments were conducted under nominally anhydrous conditions though low amounts of hydroxyl $< 750 \pm 100 \text{ wt.ppm H}_2\text{O}$ may have contaminated the sample like in the present study. *Kubo et al.* [2004a] starting material was finely grained, and experiments were performed at small pressure oversteps relative to the α - β equilibrium. Transformation–time data were acquired by in situ XRD and analyzed using the Cahn's rate equation for interface-controlled reaction. These experimental and analytical methods are almost identical to the present study; in particular, the water contents are close. However, *Kubo et al.* [2004a] reported growth rates more than one order of magnitude slower than in the present study. For example our predicted growth rate at 13.7 GPa – 1273 K using equation (3) and the above kinetic parameters ($\ln k_0 = -5.2$, $\Delta H_a = 204 \text{ kJ/mol}$ and $V^* = 0 \text{ cm}^3/\text{mol}$) is $G = 5.7 \times 10^{-9} \text{ m}\cdot\text{s}^{-1}$, while *Kubo et al.* [2004a] observed $G = 8.5 \times 10^{-10} \text{ m}\cdot\text{s}^{-1}$. In addition, the activation enthalpy for growth $\Delta H_a = 204$ (55) kJ/mol is also significantly lower than the 391 (73) kJ/mol value (when $V^* = 0 \text{ cm}^3/\text{mol}$) in *Kubo et al.* [2004a] and indicates a lower-temperature dependence of transformations rates. The comparison with the quench experiments is more tricky since the latter are performed at higher-pressure oversteps of the reaction boundary and suffer from a lower constraint in reaction duration. *Liu et al.* [1998] studied the transformation of $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$ polycrystalline samples and reported some transformation rates, but not all, significantly faster than those measured by *Kubo et al.* [2004a]. For example, *Liu et al.* [1998] estimated a growth rate $G = 1.4 \times 10^{-9} \text{ ms}^{-1}$ at 1273 K – 18 GPa, while *Kubo et al.* [2004a] measured $G = 8.5 \times 10^{-10} \text{ m}\cdot\text{s}^{-1}$ at

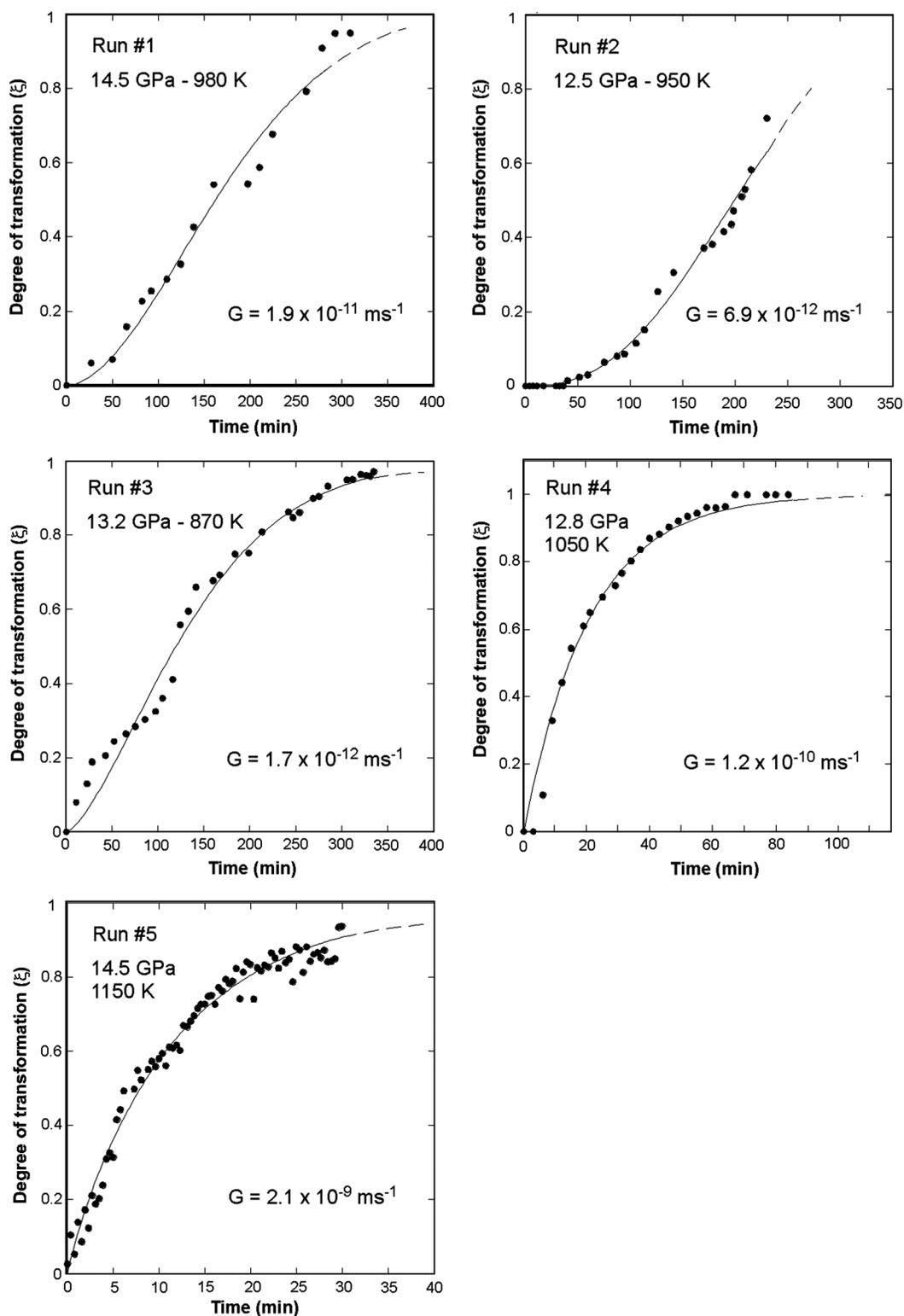


Figure 6. Transformation–time plots at specific temperatures and pressures. The solid curves represent best fits of experimental data to the Cahn rate equation (1) with optimized values of growth rate (G).

similar temperature. The activation enthalpy $\Delta H_a = 334 \text{ kJ/mol}$ estimated by *Liu et al.* [1998] is between the present value $\Delta H_a = 204 \text{ kJ/mol}$ and the one of *Kubo et al.* [2004a] $\Delta H_a = 391 \text{ kJ/mol}$. The large errors on kinetic parameters cannot account for the discrepancy. The difference may originate from the different composition of the starting materials, in

particular on the iron content of the San Carlos olivine that might enhance the α - β - γ transition rates.

[20] The pseudomartensitic reaction involves short-range Fe–Mg interdiffusion, which is a process controlled by composition [*Chakraborty et al.*, 1999; *Kubo et al.*, 2004b; *Holzappel et al.*, 2009]. For instance, *Kubo et al.* [2004b]

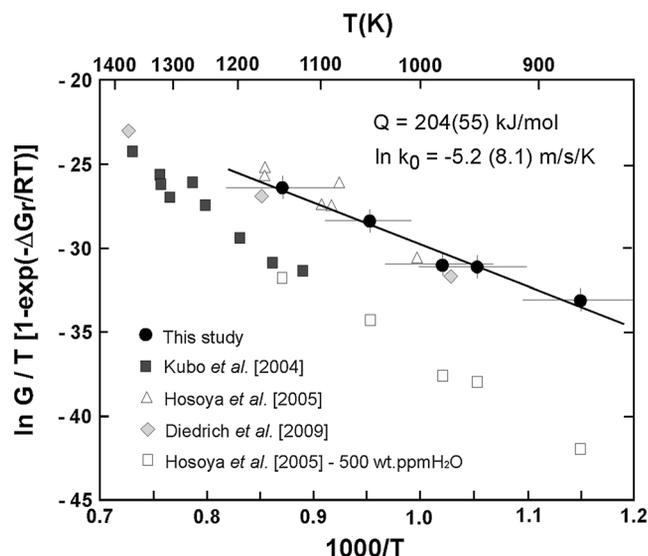


Figure 7. Arrhenius plot showing the temperature dependence of α to β growth rates in $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4 \sim 526$ wt. ppm H_2O (full circles, present study). The slope of the linear relation yields the activation energy for growth $\Delta H_a = 204$ (55) kJ/mol and the intercept a value for $\ln k_0 = -5.2$ (8.1). Kinetic data for synthetic $\text{Mg}_2\text{SiO}_4 \sim 750$ wt.ppm H_2O (gray square, nine points from Figure 8 in Kubo *et al.* [2004a]), synthetic $\text{Mg}_2\text{SiO}_4 \sim 2176$ –5000 wt.ppm H_2O (triangles, six points from Table 1 in Hosoya *et al.*, [2005]), and natural San Carlos $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4 \sim 300$ ppm H_2O (diamonds, three points from Figure 10 in Diedrich *et al.*, [2009]) are shown for comparison. Empty squares show the growth rates predicted for a 526 wt.ppm H_2O bearing Mg_2SiO_4 olivine at PT conditions of the present experiments using the kinetic parameters from Hosoya *et al.* [2005].

showed that the Mg–Fe interdiffusion rate in wadsleyite increases by almost one order of magnitude with increasing the iron concentration from $X_{\text{Mg}}=0.97$ to $X_{\text{Mg}}=0.90$ and decreasing grain size from 60 to 6 μm . Similarly, Yamazaki *et al.* [2005] studied the grain growth kinetics of ringwoodite by the quench method and observed faster growth rates in the iron-bearing samples than in the Mg_2SiO_4 end-member. Besides the effect of iron, the influence of water on the α - β - γ kinetics has been investigated owing to its relevance for olivine metastability in subducting slabs. However, experimental studies lead to conflicting results. Hosoya *et al.* [2005] investigated the α - β kinetics on 750–5000 wt.ppm H_2O synthetic Mg_2SiO_4 samples and inferred a power-law dependence of growth rate on the water content (Figure 7). In contrast, Diedrich *et al.* [2009] reported α - γ transformation rates on 300 wt.ppm H_2O hydrated San Carlos olivines up to four orders of magnitude faster than those predicted by Hosoya *et al.* [2005], suggesting a stronger catalytic effect at low water content. The present results showing a significant enhancement of kinetics of the α - β transition in the iron-bearing San Carlos olivine might well reconcile the previous studies and constrain more tightly the role of water on the phase transition. Indeed, Diedrich *et al.* [2009] determined kinetic parameters $\Delta H_a = 237$ kJ/mol and $\ln k_0 = -2.6$ (with $V^* = 0$ cm^3/mol) close to the present kinetic results for

$\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4$ containing 526 wt.ppm H_2O (Figure 7). Using these parameters, a growth rate $G = 9.0 \times 10^{-12}$ $\text{m} \cdot \text{s}^{-1}$ is estimated at 14.5 GPa – 980 K slightly slower than our measured value $G = 1.9 \times 10^{-11}$ $\text{m} \cdot \text{s}^{-1}$. Thus, it is likely that part of the discrepancy between the results of Hosoya *et al.* [2005] and Diedrich *et al.* [2009] could be explained by the role of iron, independently of the transformation type α to β or α to γ -phase.

4.2. Implications for the Metastability of Olivine in Subducting Slabs

[21] In this section, the experimentally derived kinetic parameters of the olivine α - β transformation are used in a one-dimensional thermo-kinetic model to evaluate the metastable persistence of olivine in a subducting slab. The model assumes a grain-boundary reaction controlled by growth kinetics, as observed in our experiments. It ignores the effects of intracrystalline transformation and transformation stress that may occur in a subducting slab. The extent of transformation is calculated at each time step from the Cahn’s relation (1) and the growth rate determined from equation (3) using the kinetic parameters $\ln k_0 = -5.2$, $\Delta H_a = 204$ kJ/mol, $V^* = 0$ cm^3/mol , and the thermodynamic data of Akaogi *et al.* [1989] and Jacobs and de Jong [2005] for the ΔG_r value. We investigated the transformation in a cold slab with a vertical subduction rate of 12 cm/yr and a thermal gradient of 0.6 K/km, as representative of the West Pacific subduction zones. The temperature of the lithospheric mantle at 300 km was set to 800 K, based on thermal model of Tetzlaff and Schmeling [2009]. The grain size in subducting plate is unconstrained and may lie in the range 0.01 – 10 mm [Riedel and Karato, 1997], a mean value of 5 mm was assumed in the calculation. This model is likely to be oversimplified as it does not consider the effect of latent heat released upon both kinetics and the thermal structure of the slab. For a more realistic model, the reader is referred to the recent work of Tetzlaff and Schmeling [2009]. Though schematic, the model provides an estimate of the effect of iron and water content on the metastable persistence of olivine at depth.

[22] The result of this thermo-kinetic model is presented in Figure 8. It is consistent with previous studies in predicting the presence of a metastable olivine wedge. However, the depth of the tip of the metastable wedge reaches a maximum of 740 km, which is almost 70 km above the maximal depth of 810 km estimated from the kinetic parameters of Kubo *et al.* [2004a] for Mg_2SiO_4 . Although the water content even more drastically reduces the persistence of metastable olivine (see Figure 8, calculations for 500–5000 ppm H_2O hydrous Mg_2SiO_4 , Hosoya *et al.* [2005]), the influence of the iron content cannot be neglected. These new results do not preclude the viability of the transformational faulting hypothesis for the deep focus earthquakes. Indeed, considering the inhibition of growth by transformation stress would widen the metastable olivine wedge to greater depths, while the effect of latent heat release is to lower the α - β transition boundary and speed up reaction kinetics. Intracrystalline nucleation and/or grain size reduction would also strongly enhance transformation rate and reduce the extent of metastable olivine. However, the increase of reaction rate with the iron content evidenced in the present study implies that the development of a metastable olivine wedge to depths > 660 km, corresponding to the deepest earthquakes, would

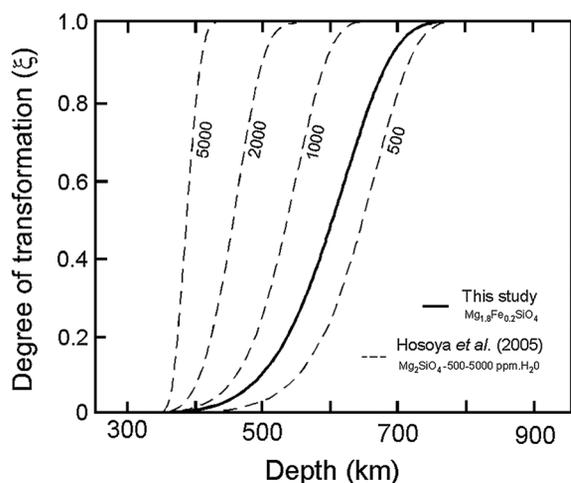


Figure 8. Fraction of olivine transformation as a function of depth for a cold slab with vertical subduction speed of 12 cm/yr and a thermal gradient of 0.6 K/km. Thermo-kinetic calculations were performed using the present kinetic results for $\text{Mg}_{1.8}\text{Fe}_{0.2}\text{SiO}_4 \sim 500$ wt.ppm H_2O (solid line) and the power-law dependence of growth rate on OH content with 500–5000 wt.ppm H_2O in Mg_2SiO_4 from Hosoya et al. [2005] (dash lines). For comparison purpose, in both models, the activation volume was set to a similar value of $V^* = 0 \text{ cm}^3/\text{mol}$.

be restricted to even colder and/or dryer subduction zones than previously estimated.

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