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Effect of Eu substitution on the crystallographic and magnetic properties of the BiMn_2O_5 oxide obtained by urea combustion

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Abstract

Polycrystalline oxide materials with nominal compositions BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ were obtained by urea combustion. Crystallographic studies using X-ray diffraction and Rietveld Refinement techniques showed the formation of single-phase samples for both compositions, crystallizing in a mullite-type orthorhombic perovskite structure, space group $Pbam$ ($Z=4$). Replacement of Bi^{+3} by Eu^{+3} promoted a slight distortion and a unit cell contraction due to a decrease of the c-parameter for $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$. The SEM/EDX techniques confirmed the formation of single-phase materials with excellent mapping distribution. Magnetic measurements showed different behaviors: BiMn_2O_5 is an antiferromagnet (AFM) with a Néel temperature (T_N) of 41 K and a Curie-Weiss temperature Θ_{CW} of -277 K, while $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ displayed a ferrimagnetic behavior with a T_C of 44K and Θ_{CW} of -270 K. The role of Eu^{+3} and the mechanisms responsible for the magnetic transition are discussed on the basis of chemical bonds and exchange interactions.

Keywords

A. Powders: chemical preparation, B. X-ray methods, C. Magnetic properties ;
D. Perovskites.

1. Introduction

Multiferroics is a material class which simultaneously presents two or more primary ferroic orders such as ferroelectricity, ferroelasticity and (anti-)ferromagnetism when an electric field or a magnetic field are applied. The capacity of a material to couple two or more ferroic properties is due to the presence of two different types of cations: i) cations with electronic configuration ns^2 for which the lone pair is responsible for ferroelectricity; ii) cations with d^n or f^n partially-filled subshells, accountable for the magnetic ordering. Thus, these materials may present various properties such as magnetoelectricity, magneto-optical response, magneto-elasticity and ferroelasticity. Research on these materials attracts a great interest from the scientific and technological points of view since they can be used in spintronic and magneto-optical devices or as low-field magnetic sensors [1-7].

Although these materials are subject of fascination in basic physics and of great interest in applied research, only few oxides can be classified as multiferroics with potential possibilities in technological applications. Manganite oxides of the RMn_2O_5 family (R a lanthanide, Yttrium or Bismuth), crystallizing with a mullite-type perovskite structure and space group $Pbam$, have attracted much attention due to their multiferroic properties such as magnetoelectricity and magnetodielectricity [8-10].

Among this family, $BiMn_2O_5$ has been extensively investigated. The two manganese ions have different oxidation states: Mn^{4+} occupies the octahedral sites (4f site) coordinated by six oxygen anions, while Mn^{3+} is located in a

distorted tetragonal pyramidal site ($4h$ site) coordinated by five oxygen anions. The Bi^{3+} ions ($4g$ site) are coordinated to eight oxygen atoms and form BiO_8 polyhedra. These polyhedra are responsible for the interconnection between both manganese cations localized at the octahedral and tetragonal pyramidal sites^[11]. The mixed-valent manganese ions occupying different crystallographic sites lead to the formation of a magnetic structure with two different magnetic sublattices. BiMn_2O_5 orders below 39 K in a noncollinear commensurate magnetic structure with a propagation vector $\mathbf{k} = (\frac{1}{2}, 0, \frac{1}{2})$, in contrast to the RMn_2O_5 series, incommensurate with $\mathbf{k} = (\frac{1}{2}, 0, k_z)$ ^[12,13]. The magnetization curves for all members of the RMn_2O_5 family present a typical antiferromagnetic behavior with Néel temperatures close to 40 K^[11-15].

This paper aims to demonstrate, first, that the urea combustion is an adequate synthesis method to obtain multiferroic materials and, secondly, to investigate the role of the Eu^{3+} ions as improving dopants for the BiMn_2O_5 properties. In particular, the X-ray diffraction analysis and Rietveld refinement should allow correlating the modifications of the magnetic behavior of the RMn_2O_5 family when the Bi element is partially substituted by Eu.

2. Materials and Methods

The oxide materials BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ were obtained by the urea combustion reaction method^[16,17]. In a first step, MnO , Bi_2O_3 and Eu_2O_3 precursors (99,99% purity Aldrich) were weighted and dissolved separately in a concentrated solution of nitric acid. Then, these solutions were mixed under agitation with appropriate amounts of urea dissolved in water. The ratio of the

urea amount (R) to the corresponding nitrates was obtained using equation 1 [16].

$$\varphi = -R \frac{[+4_{(C)} - 2_{(O)} + 2(0_{(N)} + 2 \cdot 1_{(H)})]}{[+3_{(Bi)} + 3(0_{(N)} + 3 \cdot -2_{(O)})] + 2[+3_{(Mn)} + 2(0_{(N)} + 3 \cdot -2_{(O)})]} \quad (1)$$

where φ is a parameter which defines the oxidize/reduction reaction conditions and, for this work, taken as 3. Numbers in numerator and denominator are the oxidation state of elements which for nitrogen was taken as zero.

The obtained products were then heated to 500°C until complete combustion. Powders were progressively calcined in air at 600 °C for 4 hours and then at 700 °C and 800 °C for 4 hours in each temperature with intermediate grinding after each calcination temperature. Similar procedure was adopted for the europium-doped sample with a final calcination at 900°C for 4 hours. The final products were grinded in an agate mortar and characterized by different techniques.

The homogeneity of the samples was confirmed by energy dispersive X-ray mapping (EDX) and scanning electronic microscopy SEM observations (Philips model FEI Inspect S50). X-ray diffraction (XRD) patterns were registered with a Rigaku Rint DMAX-2100/PC diffractometer using Cu K_{α} radiation (1.5406 Å). The 2θ scans were carried out at room temperature from 10° to 90°, with fixed time mode of 1.6 s and a step of 0.02°. Structural refinements were performed by the Rietveld method [18,19] using the GSAS program [20], the JCPDS phase identification and the ICSD (2003) structural database. The diffraction profiles were modeled using the pseudo-Voigt Thompson-Cox-Hastings (TCH) function [21]. Magnetization measurements were

performed using a PPMS Quantum Design 6000 magnetometer with zero field cooling (ZFC) and field cooling (FC) data collected between 2K and 300K under an applied field of 1 kOe.

3. Results and Discussions

3.1. SEM Analysis and Elements Mapping

Figures 1(a) and 1(c) present SEM images of BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ powder specimens. A homogeneous distribution of grains can be observed with average sizes close to 200 nm and 300 nm, for the non-doped and the doped samples, respectively. The difference in the particle sizes can be attributed to the calcination procedure: $T_{(\text{max})} = 800^\circ\text{C}$ (calcination total time = 12h) for BiMn_2O_5 and $T_{(\text{max})} = 900^\circ\text{C}$ (calcination total time = 16h) for $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$. These two parameters (temperature and time) are known to enhance the sintering phenomena and induce an increase of particles size, as it occurred in the case of the $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ material

Heat-treatments performed at 900°C for the Eu doped sample was made considering it crystallized in a monophasic solid solution at this temperature. Otherwise, BiMn_2O_5 sample was crystallized like a monophasic solid solution in 800°C . This were observed using the XRD associated with Rietveld Refinement and will be discussed on the section 3.2.

The elements distribution was checked by EDX analysis by mapping a large surface of compacted powders. Figures 1(b) and 1(d) show a uniform distribution of all elements (Bi, Eu, Mn and O). Quantitative analyses of both

samples by SEM-EDX confirmed the expected stoichiometric ratios of the constituent elements and the results are given in Table 1.

3.2. X-Ray Powder diffraction and Rietveld Refinement

The X-ray diffraction data (XRD) of both samples were refined by the Rietveld method. Figures 2(a) and 3 show a good agreement between the calculated and observed data for BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$, respectively. Both materials crystallize in a mullite-type orthorhombic structure, space group $Pbam$ and $Z=4$. In this structure, Mn ions are found in two different crystallographic sites: Mn^{4+} ions occupy the octahedral pyramidal site ($4f$) coordinated to six oxygen atoms, whereas the Mn^{3+} ions occupy the tetragonal pyramidal site ($4h$) coordinated to five oxygen atoms. The Bi^{3+} ions sit at the $4g$ sites and are coordinated to eight oxygen atoms forming BiO_8 polyhedra [11-14]. Based on our refined data, we may suppose that the Eu^{3+} ions replace Bi^{3+} at the $4g$ site and do not modify the crystal structure of the host material.

Figure 2(b) is a schematic representation of the crystallographic structure along the c -axis. In this representation distorted Mn^{4+}O_6 octahedral are environments of the Mn^{4+} (Mn1 atoms) which form infinite linear chains along the c -axis, sharing edges via the O_3 and O_4 oxygen. Mn^{3+} (Mn2 atoms) occupy Mn^{3+}O_5 distorted tetragonal pyramids and have as neighbors five oxygen atoms. These units present four oxygen atoms in different positions (two O_1 and two O_2) are in a square planar configuration and fifth oxygen (O_4) is in an axial position at a longer distance (2.111 Å). The connections between pyramids and octahedrons are along the c -axis via O_4 and O_2 .

Table 2 indicates the structural parameters obtained by the Rietveld refinement. Tables 3 and 4 list the most important interatomic distances and angles concerning the non-doped BiMn_2O_5 and doped $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ materials, respectively, making emphasis on the different polyhedra Mn^{4+}O_6 , Mn^{3+}O_5 , Bi^{3+}O_8 and Eu^{3+}O_8 .

The average Mn-O distances obtained in this work ($\langle\text{Mn}^{4+}\text{-O}\rangle = 1.897 \text{ \AA}$ and $\langle\text{Mn}^{3+}\text{-O}\rangle = 1.920 \text{ \AA}$ (Table 3)) are slightly lower than those reported by Muñoz *et al* from Neutron Powder Diffraction data ($\langle\text{Mn}^{4+}\text{-O}\rangle = 1.916 \text{ \AA}$ and $\langle\text{Mn}^{3+}\text{-O}\rangle = 1.948 \text{ \AA}$) [11]. This fact is associated to a contraction of the octahedral and pyramidal sites. Reported results of structural parameters for samples synthesized by other methods, such as hydrothermal [22] and solid state reactions [8,10] are in agreement with the data found in this work. However, we should recall the fact that the heat-treatment parameters (T° and total time) adopted in this work (800-900°C and 12h and 16h) are lower than those mentioned by others authors, and it confirms that urea stress the fact that the heat-treatment adopted in this work is lower than all others mentioned above, and it confirms that urea combustion is an excellent method to obtain single-phase oxide materials at low temperatures.

The Eu^{3+} substitution at the Bi^{3+} site leads to a slight contraction of the unit cell (Table 2) due to the smaller ionic radius of Eu^{3+} (1.06 Å) compared to Bi^{3+} (1.17 Å). The lattice contraction occurs predominantly on the lattice parameters a and c , whereas no significant variation is observed for the b -parameter. Slight variations in the positions of the (001), (002), (003) and (004) planes explain the modifications of the octahedral sites occupied by Mn^{4+} ions that form chains along the c axis [23-25].

The average distance Mn1-O for sample BiMn₂O₅ ($\langle \text{Mn1-O} \rangle = 1.897 \text{ \AA}$) is larger than the ($\langle \text{Mn1-O} \rangle = 1.887 \text{ \AA}$), value obtained for the Eu-doped sample suggesting that the insertion of the Eu³⁺ ions in the *4g* position promotes a contraction of this site. The distorted pyramidal polyhedra Mn³⁺O₅ of the Bi_{0.9}Eu_{0.1}Mn₂O₅ suffers a contraction due to the Mn³⁺ located at the *4h* position, the Mn2-O₄ distance changing from 2.111 Å to 2.088 Å, for the non-doped and Eu-doped compounds, respectively (Tables 3 and 4). Such stretching of the Mn2-O distance in the Mn³⁺O₅ pyramid tetrahedral is probably associated to the Jahn-Teller effect of the Mn³⁺ ion.

Other distances related to the chemical bonds, such as $\langle \text{Bi-O} \rangle$ and $\langle \text{Mn1-Mn1} \rangle$, slightly decrease in the doped Bi_{0.9}Eu_{0.1}Mn₂O₅ sample, whereas the $\langle \text{Mn2-Mn2} \rangle$ average distance stays invariant. This fact confirms the larger distortion of the tetragonal pyramidal site, since the average distances $\langle \text{Mn1-O} \rangle$ and $\langle \text{Mn2-O} \rangle$ also decrease with doping.

3.3. Magnetic Measurements

Figure 4 shows the ZFC/FC cycles of BiMn₂O₅ and Bi_{0.9}Eu_{0.1}Mn₂O₅ measured under an applied magnetic field of 1 kOe. Both samples present an antiferromagnetic-type ZFC magnetization, with a maximum at $T = T_N$ of 41 K and 37 K, respectively. During the FC mode, the magnetization increases abruptly at the ferromagnetic ordering temperature T_C , defined in this work as the intersection of two straight lines sketched just above and just below T_C . The T_C values for both systems are identical, 44.3(0.2)K, meaning that Mn³⁺-Mn⁴⁺

ferromagnetic interactions have been optimized through the coupling between the tetrahedral and octahedral sublattices.

Two experimental facts are worth to be noticed in Figure 4: firstly, the large difference in magnetization values between both samples, in the ordered state; secondly, the quite constant evolution of the FC magnetization below 20-30 K for the Eu-doped sample. Knowing that Eu^{3+} is a VanVleck ion, no intrinsic magnetism may be expected from it at these temperatures [26]; in consequence, the large M_{FC} magnetization in the $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ system must be attributed to a better ferromagnetic coupling between the manganese moments, probably due to the contraction of the unit cell when replacing Bi by Eu. The observed phenomena for Eu substitution is attributed to a Morin-type transition, usually observed in Eu doped systems [27].

Modifications of the lattice parameters and probably the accompanying distortion of the lattice due to a Jahn-Teller effect of the Mn^{3+} ion [9], may also create some AFM canting of the magnetic moments, bringing about an additional ferromagnetic component in the Eu-doped case, thus increasing the ferromagnetic contribution at low temperatures.

The second fact pointed above, that is the quite constant thermal evolution of M_{FC} magnetization at low temperature in the case of the Eu-doped material, may be attributed to anisotropic features which are quite important in the case of the full-europium mullite-type perovskite structure, EuMn_2O_5 , in contrast to BiMn_2O_5 which presents none [15]. Our experimental conditions for powdered samples may provoke some reorientation by the applied field, of grains and/or particles due to the magnetic anisotropy, contrary to the reports made by Golovenchits *et al* in single crystal samples [15]. This anisotropy may

also be responsible of the much smaller hysteresis between the ZFC and FC modes in the case of BiMn_2O_5 , even though this system was reported by Golovenchits *et al* as isotropical from the magnetic point of view. Other authors have also found weak hysteretic behavior in BiMn_2O_5 samples ^[11]. Magnetic anisotropy may also be the explanation of small differences in the values of T_N (39 K in ref. [11], ~40 K in ref. [15], 41 K in this work).

The inset, Figure 4, shows the inverse magnetic susceptibility for both samples, at the paramagnetic state. Analysis of $1/\chi$ -vs-T by linear regression in the range [100-300K] allows finding the Curie-Weiss temperature Θ_{WC} and the magnetic effective moment μ_{eff} for both samples, with values of $\Theta_{WC} = -277$ K and -270 K, and $\mu_{\text{eff}} = 6.69 \mu_B$ and $6.90 \mu_B$, for BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$, respectively. These values are quite similar to those reported in the literature for a BiMn_2O_5 perovskite obtained by hydrothermal, solid-state and citrate techniques, $\Theta_{WC} \sim -270$ K and $\mu_{\text{eff}} = 6.9 \mu_B$ ^[8-11], but somewhat higher than the magnetic moment of $6.24 \mu_B$ expected from the contributions of non-interacting Mn^{3+} and Mn^{4+} ions in the paramagnetic state.

Some considerations must be forwarded regarding the coexistence of an AFM and a FM ordering behaviors. It is quite evident that the quite large negative Curie-Weiss temperature (~-270 and -277 K found in this work) is indicative of a ferrimagnetic nature of both materials, which may be extrapolated to other components of the RMn_2O_5 family ^[8-14].

4. Conclusions

In this work we demonstrated that single phase materials BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ were successfully obtained by the urea combustion method, excellent technique to obtain complex oxide materials at intermediate temperatures with good particle size and uniform element distribution. Rietveld refinement of the XRD data agrees with a substitution of Bi^{+3} by Eu^{+3} in the same crystallographic sites, both cations having similar ionic radius. The magnetic behavior changes significantly when Bi^{+3} ions are substituted by Eu^{+3} . Both compositions order below 40 K, with BiMn_2O_5 in an antiferromagnetic state while $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ presents a ferrimagnetic behavior. This difference seems to be related to the commensurate magnetic structure of BiMn_2O_5 eventually becoming incommensurate with the substitution of Bi^{+3} by Eu^{+3} . One of the key mechanisms is the lattice distortion, which modifies the strength of the exchange chains responsible of the ferrimagnetic behavior.

Acknowledgments

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Tables Captions:

Table 1. Stoichiometric Compositions of BiMn_2O_5 and $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ samples obtained by EDS-analysis.

Table 2. Rietveld XPD analysis result data for the samples obtained by combustion urea method. Data were collected at room temperature and the space group used was *Pbam*.

Table3. BiMn_2O_5 Specimen interatomic distances (in Å) and select bonding angles (in degree) obtained by Rietveld Refinement Method at 298 K.

Table4. $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ Specimen interatomic distances (in Å) and select bonding angles (in degree) obtained by Rietveld Refinement Method at 298 K.

Figures Captions:

Figure 1: a) Scanning electron micrograph of a representative samples of BiMn_2O_5 manganite powders obtained by the urea combustion method, followed by heat treatment in air at 800°C for 4 hours. b) Scanning elements analysis of the BiMn_2O_5 sample showing a homogenous elements distribution for Bismuth, Manganese and Oxygen. c) Scanning electron micrograph of representative samples of $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ powders obtained by the urea combustion method, followed by heat treatment in air at 900°C for 4 hours. d) Scanning elements analysis of the $\text{Bi}_{0.9}\text{Eu}_{0.1}\text{Mn}_2\text{O}_5$ sample showing a homogenous elements distribution for Europium, Bismuth, Manganese and Oxygen.

Figure 2. a) XRD Rietveld refinement for the BiMn_2O_5 sample heat-treated in air at 800°C for 4h. b) The structural models of BiMn_2O_5 with chain of Mn atoms along the c-axis.

Figure 3. XRD Rietveld refinement for the $\text{Bi}_{0.90}\text{Eu}_{0.10}\text{Mn}_2\text{O}_5$ sample heat-treated in air at 800°C for 4h.

Figure 4. Magnetic susceptibility versus temperature at a constant applied field of 1 kOe. (Inset) Inverse of magnetic susceptibility vs temperature for an applied field of 1 kOe for the BiMn_2O_5 $\text{Bi}_{0.90}\text{Eu}_{0.10}\text{Mn}_2\text{O}_5$ samples.

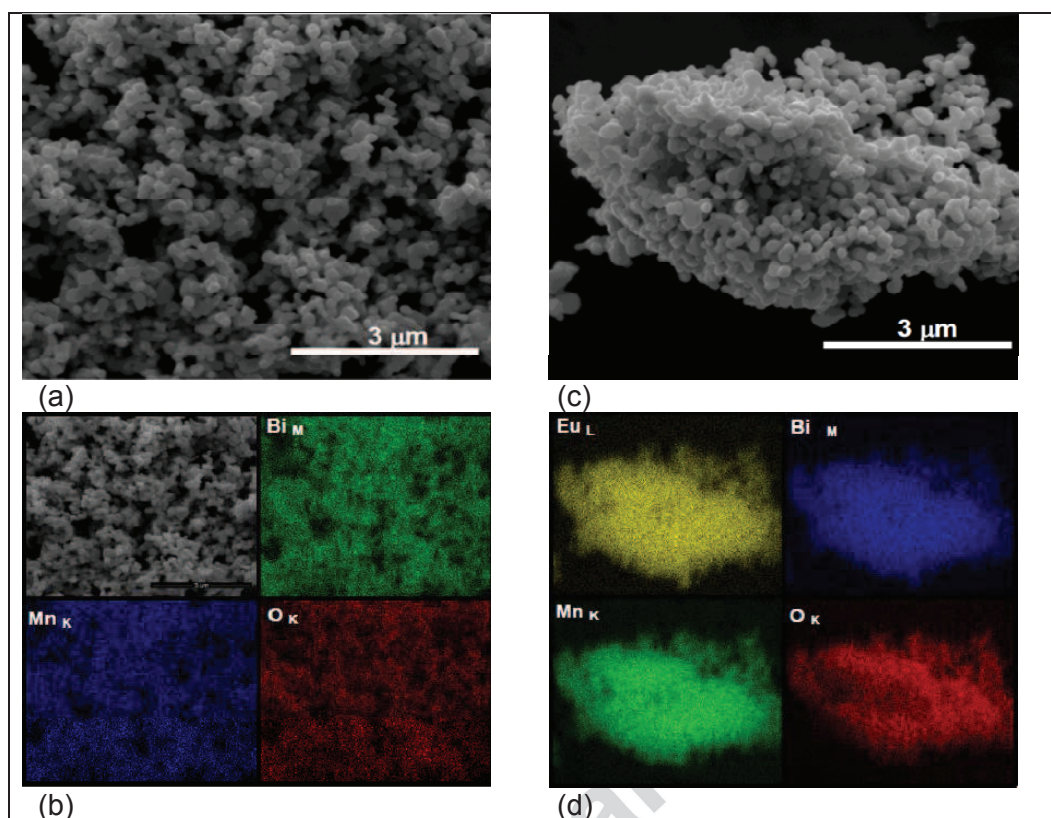
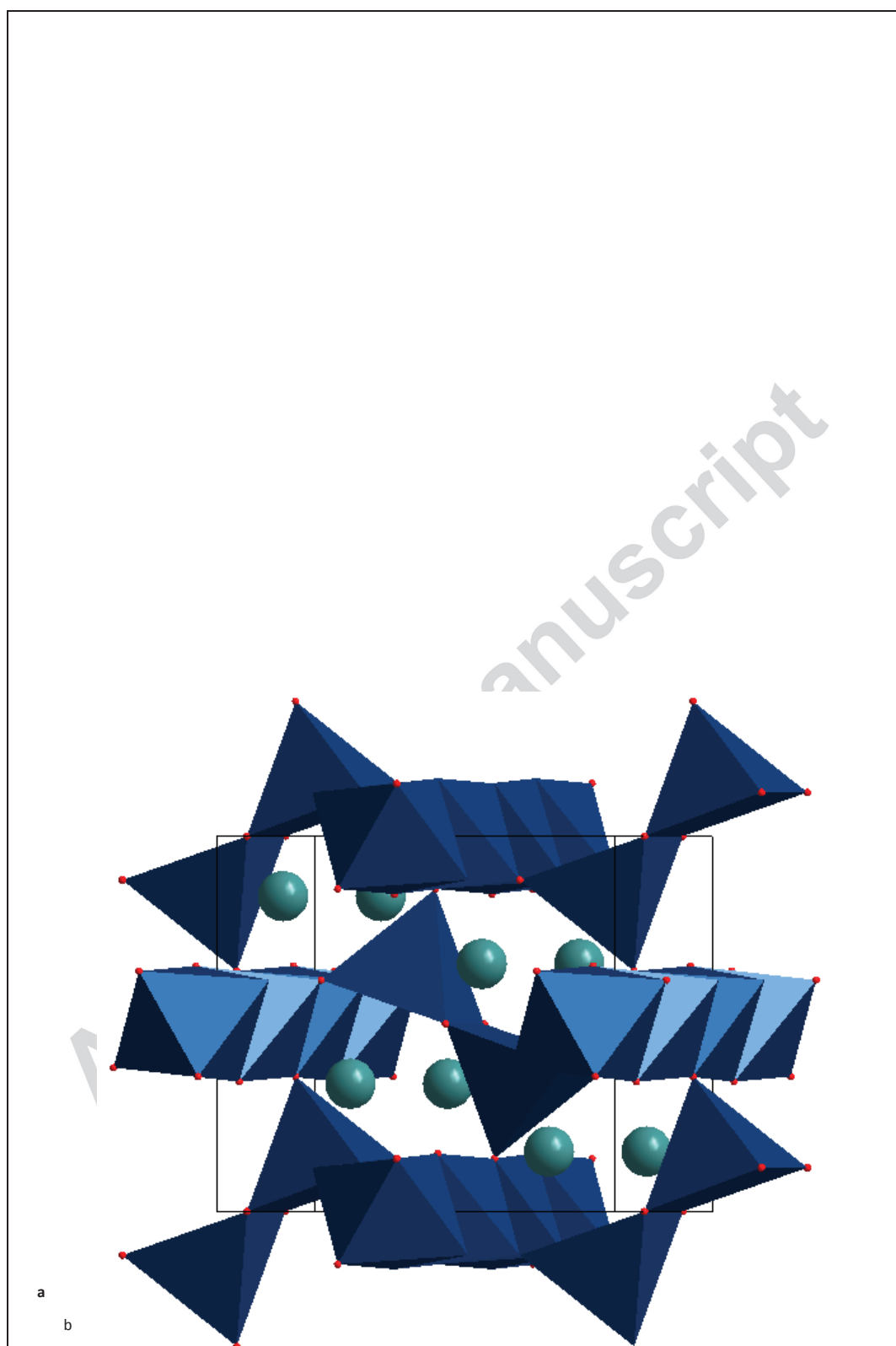


Figure 1.

**Figure 2.**

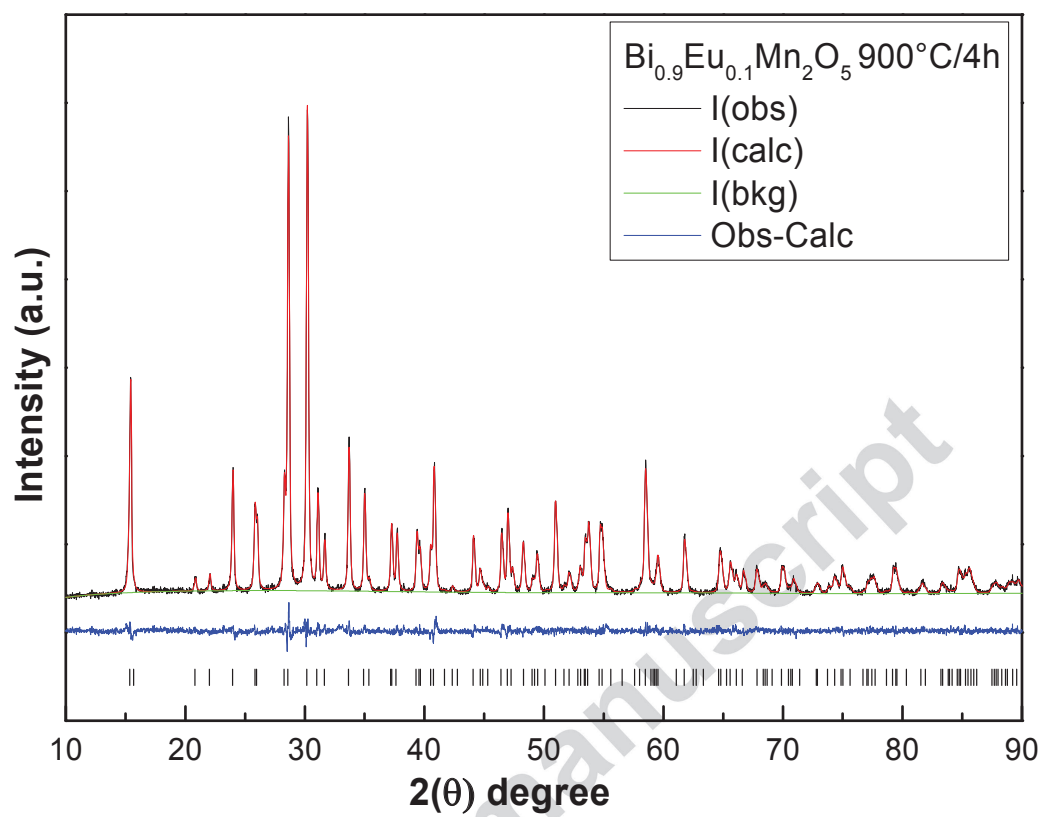


Figure 3.

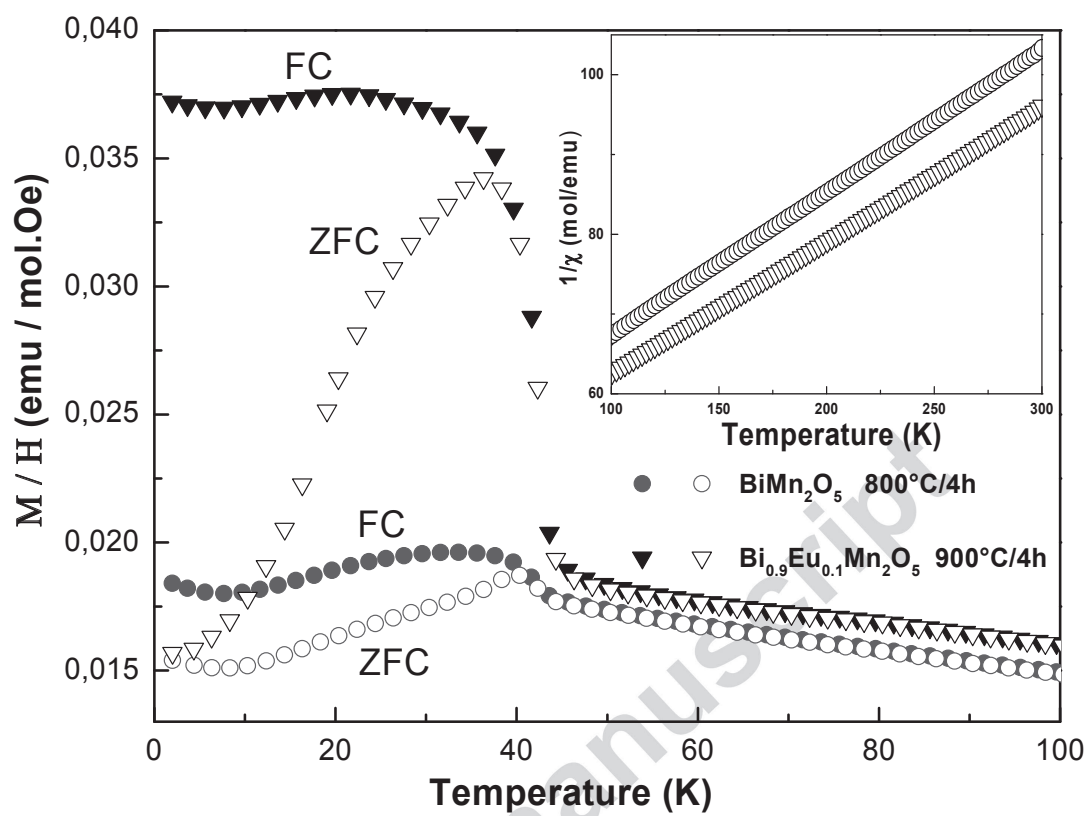


Figure 4.

Table 1:

Sample	Stoichiometric Values (% weight / %atomic)					
	Bi		Mn		Eu	
Nominal Composition BiMn_2O_5	65.54	33.33	34.46	66.67	-	-
Experimental Composition $\text{Bi}_{0.96}\text{Mn}_{2.04}\text{O}_5$	64.20	31.90	35.80	68.10	-	-
Percentage Difference (%)	2.04	4.02	3.88	2.09	-	-
Nominal Composition $\text{Bi}_{0.90}\text{Eu}_{0.10}\text{Mn}_2\text{O}_5$	60.06	30.00	35.09	66.67	4.85	3.33
Experimental Composition $\text{Bi}_{0.88}\text{Eu}_{0.09}\text{Mn}_{2.04}\text{O}_5$	59.28	28.73	36.31	68.16	4.41	3.11
Percentage Difference (%)	1.31	4.23	3.47	3.00	9.97	6.66

 $\text{Mn}^{4+}\text{-O6}$ $\text{Bi}^{3+}\text{-O8}$

Table 2.

Nominal Composition BiMn_2O_5	$R_F = 0.00365$	$\chi^2 = 1.083$	Orthorhombic System
Space Group	<i>Pbam</i>	$\alpha = \beta = \gamma = 90^\circ$	
Lattice parameter (Å)	$a = 7.5597(2)$	$b = 8.5342(4)$	$c = 5.7608(4)$
Volume (Å ³)	371.664		
Nominal Composition $\text{Bi}_{0.90}\text{Eu}_{0.10}\text{Mn}_2\text{O}_5$	$R_F = 0.00356$	$\chi^2 = 1.091$	Orthorhombic System
Space Group	<i>Pbam</i>	$\alpha = \beta = \gamma = 90^\circ$	
Lattice parameter (Å)	$a = 7.5522(8)$	$b = 8.5352(3)$	$c = 5.7577(0)$
Volume (Å ³)	371.143		

Mn1-O2 (x2)	1.863	Bi-O1 (x2)	2.558
Mn1-O3 (x2)	1.956	Bi-O2 (x2)	2.309
Mn1-O4 (x2)	1.873	Bi-O2 (x2)	2.847
<Mn1-O> = 1.897		Bi-O3	2.368
Mn ³⁺ -O5		Bi-O3	2.360
Mn2-O1 (x2)	1.863	<Bi-O>	2.519
Mn2-O2 (x2)	1.883		
Mn2-O4	2.111	Mn-Mn	
<Mn2-O> = 1.920		Mn1-Mn1	2.785
		Mn1-Mn1	2.976
		Mn2-Mn2	2.975
Selected bonding angles			
O3-Mn1-O3	81.0(2)	O1-Mn2-O4	98.9(2)
O3-Mn1-O4	173.9(2)	O1-Mn2-O2	95.2(1)
O3-Mn1-O4	97.2(4)	O1-Mn2-O2	161.9(2)
O3-Mn1-O2	94.5(1)	O4-Mn2-O2	97.7(3)
O3-Mn1-O2	87.8(6)	O2-Mn2-O2	91.1(4)
O4-Mn1-O4	83.9(2)	Mn2-O1-Mn2	106.0(1)
O4-Mn1-O2	86.3(1)	Mn1-O3-Mn1	99.0(0)
O4-Mn1-O2	91.5(4)	Mn1-O4-Mn1	96.1(0)
O2-Mn1-O2	177.5(2)	Mn1-O4-Mn2	130.5(5)
O1-Mn2-O1	74.0(0)	Mn1-O2-Mn2	126.5(5)

Table 3.

Table 4.

$\text{Mn}^{4+}\text{-O6}$		$\text{Bi}^{+3}\text{-O8}$	
Mn1-O2 (x2)	1.854	Bi-O1 (x2)	2.498
Mn1-O3 (x2)	1.941	Bi-O2 (x2)	2.332
Mn1-O4 (x2)	1.861	Bi-O2 (x2)	2.814
$\langle \text{Mn1-O} \rangle = 1.887$		Bi-O3	2.380
$\text{Mn}^{3+}\text{-O5}$		Bi-O3	2.380
Mn2-O1 (x2)	1.914	$\langle \text{Bi-O} \rangle = 2.506$	
Mn2-O2 (x2)	1.887	$\text{Eu}^{+3}\text{-O8}$	
Mn2-O4	2.088	Eu-O1 (x2)	2.554
$\langle \text{Mn2-O} \rangle = 1.938$		Eu-O2 (x2)	2.164
Mn-Mn		Eu-O2 (x2)	3.094
Mn1-Mn1	2.763	Eu-O3	2.575
Mn1-Mn1	2.995	Eu-O3	2.085
Mn2-Mn2	2.976	$\langle \text{Eu-O} \rangle =$	2.535
Selected bonding angles			
O3-Mn1-O3	79.0(9)	O1-Mn2-O4	97.8(5)
O3-Mn1-O4	173.8(8)	O1-Mn2-O2	93.2(1)
O3-Mn1-O4	98.5(6)	O1-Mn2-O2	161.9(2)
O3-Mn1-O2	94.5(1)	O4-Mn2-O2	99.0(6)
O3-Mn1-O2	87.8(6)	O2-Mn2-O2	90.8(4)
O4-Mn1-O4	83.9(2)	Mn2-O1-Mn2	102.0(9)
O4-Mn1-O2	87.1(7)	Mn1-O3-Mn1	101.9(0)
O4-Mn1-O2	91.1(8)	Mn1-O4-Mn1	95.6(1)
O2-Mn1-O2	177.6(1)	Mn1-O4-Mn2	131.2(5)
O1-Mn2-O1	78.0(9)	Mn1-O2-Mn2	127.0(9)