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# Experimental Investigation in the Quaternary Systems Ti–Ni–Al–N and Ti–Ni–Al–O

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## EXPERIMENTAL

The experimental evaluation of phase equilibria in the Ti–Ni–Al–N and Ti–Ni–Al–O phase diagrams are based on alloy samples, which were prepared of elemental powder blends by argon-levitation melting in a *Hukin* crucible. The experimental investigation employed X-ray powder diffraction, metallography, SEM, and EMPA techniques in the as-cast state as well as after annealing at 900°C. Two quaternary compounds  $\text{Ti}_3\text{NiAl}_2\text{N}$  and  $\text{Ti}_3\text{NiAl}_2\text{O}$  deriving from the filled  $\text{Ti}_2\text{Ni}$  type ( $\eta$  phase) were observed. The novel phases are in equilibrium with the  $\text{Ti}_2\text{Ni}$ -type solid solution phase  $(\text{Ti}_{1-x}\text{Al}_x)_2\text{Ni}$ , which exhibits a maximum solubility of 14 at.% Al in binary  $\text{Ti}_2\text{Ni}$ . Atom order in all these phases was monitored by quantitative X-ray powder diffraction (Rietveld analyses). The difference of X-ray spectra among the various phases deriving from parent  $\text{Ti}_2\text{Ni}$  type was analyzed and the complex atom site occupation mode was discussed in terms of the general classification scheme for  $\eta$  phases.

## INTRODUCTION

High-strength, heat-, and oxidation-resistant Ti–Ni–Al alloys have recently attracted widespread attention, competing with nickel-based superalloys for high-temperature applications (1). Nothing, however, is known so far about the performance of these alloys under nitrogen/oxygen or using  $\text{TiN}_{1-x}/\text{TiO}_{2-x}$  and nitride/oxide precipitates as strengtheners. The present report in continuation of our systematic studies of phase equilibria and crystal data within the quaternary Ti–Ni–Al–N and Ti–Ni–Al–O presents experimental details on the crystal structure of  $\text{Ti}_3\text{NiAl}_2\text{N}$  and some of its phase relations within Ti–Ni–Al–N. This work also covers phase equilibria in the Ti–Ni–Al–O system as well as the formation of an isotypic quaternary oxide.

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Nine quaternary samples in each of the Ti–Ni–Al–N and Ti–Ni–Al–O systems have been prepared by high-frequency levitation melting starting from high-purity powders (purity > 99.9 mass%) of Ti (Johnson-Matthey & Co., UK), Ni (Research Chemicals, Princeton, NJ), and Al, TiN, and  $\text{TiO}_2$  (Alfa, Karlsruhe, Germany), which served to introduce the properly calculated amount of nitrogen/oxygen to the alloys. The well-blended powder mixtures with a total amount of about 3 g were compacted in  $\varnothing = 8$  mm steel dies without the use of lubricants and were out-gassed prior to melting under Ti/Zr-gettered high-purity argon atmosphere. For better homogeneity the samples were remelted twice or thrice, and a part of each alloy was successively subjected to a long-term annealing at 900°C for 240 h in evacuated and sealed silica tubes whose internal surface was lined with Mo foil to protect the sample alloy from attack by the hot quartz walls. After heat treatment in a thermocouple controlled wire-wound tube furnace, the alloys were quenched by rapidly submerging the quartz capsules in water. Experimental details of sample preparation and for X-ray powder diffraction experiments, LOM, and quantitative microprobe measurements have in part been given in one of our foregoing manuscripts (2). Full-matrix full-profile Rietveld refinements were carried out with a PC version of the program by Rodriguez (3) on X-ray intensities that were recorded from a flat specimen in the D-5000 diffractometer or from an image-plate Guinier–Huber system.

Quantitative composition analyses for the Ti–Ni–Al–N samples were performed on as-cast and annealed alloys on a CAMEBAX SX50 wavelength dispersive spectrograph comparing the  $K\alpha$  emissions of the four elements in the alloys with those from elemental standards and a VN standard and applying a deconvolution and ZAF correction procedure (4). The experimental parameters employed were acceleration voltage of 15 kV, sample current of 20 nA, and

spectrometer crystals such as PET for TiK $\alpha$ , TAP for AlK $\alpha$ , LiF for NiK $\alpha$ , and PC1 for the NK $\alpha$  radiation. Microstructure examination of Ti–Ni–Al–O samples annealed at 900°C was performed on a Zeiss DSM 962 with a Link EDX system measuring the K $\alpha$  emissions of the metals under 25 kV and 0.57  $\mu$ A sample current.

## RESULTS AND DISCUSSION

### a. Phase Relations in Ti–Ni–Al–N and Ti–Ni–Al–O Systems

Crystallographic data from X-ray powder diffraction and EMPA results for the systems Ti–Ni–Al–N and Ti–Ni–Al–O are summarized in Tables 1 and 2, revealing partial phase relations in the quaternary involving new type quaternary phases with composition  $Ti_3NiAl_2N$  and  $Ti_3NiAl_2O$  ( $Ti_2Ni$ -derivative type, see below). Variations of the lattice parameters and of the composition in EMPA (see Table 1) indicate a small but significant homogeneous region. The EMPA data also reveal a small deficiency in nitrogen ( $12.4 \pm 7$  at.% N in contrast to 14.3 at.% N in the ideal formula). It is important to note that in all the quaternary alloys investigated there was practically no solubility for nitrogen for the binary and ternary phases encountered from the Ti–Ni–Al system. Refining these EMPA data of the nitride phase taking oxygen as a fifth element being the difference weight to 100 mass%, we obtain 550 ppm O, which may account for some further filling of the nonmetal lattice site (see below). Figure 1 shows the microstructure of alloy  $Ti_{3.7}Ni_{28.6}Al_{20.1}N_{14.3}$  after annealing at 900°C, revealing the quaternary phase  $Ti_3NiAl_2N$  in equilibrium with TiN and several nitrogen-free phases from the Ti–Ni–Al ternary system. The microstructures of the alloys  $Ti_{28.57}Ni_{28.57}Al_{28.57}O_{14.29}$  and  $Ti_{42.86}Ni_{14.28}Al_{28.57}O_{14.29}$  show the coexistence of the solid solution  $(Ti, Al)_2Ni$  (expanding from binary  $Ti_2Ni$ ) with the  $Ti_3NiAl_2O$  phase (see Fig. 2). Practically isotypic both these phases derive from the  $Ti_2Ni$ -parent type; however, they differ with respect to atom order and nonmetal interstitials (for structural details see below). A closer inspection of the oxygen content in the  $Ti_3NiAl_2O$  phase by quantitative EMPA revealed a composition with  $13.2 \pm 6$  at.% O, rather close to the nominal values.

### b. X-Ray Diffraction

*b.1. Rietveld refinement of the crystal structures of  $(Ti_{1-x}Al_x)_2Ni$ ,  $Ti_3NiAl_2N$  and  $Ti_3NiAl_2O$ .* Table 3 summarizes the results of the Rietveld X-ray refinement of the quaternary alloys  $Ti_{4.7}Ni_{28.6}Al_{10.14}N_{14.3}$ ,  $Ti_3NiAl_2N$ , and  $Ti_{44.5}Ni_{14.8}Al_{29.6}N_{11.1}$  in comparison with the ternary alloy  $Ti_{56.8}Ni_{29}Al_{14.2}$ , which exhibits the maximum Al solubility in binary  $Ti_2Ni$ . In all these alloys the main phase crystallizes with the  $Ti_2Ni$  type with various atom site occupations. In agreement with the extension of the phase

field  $(Ti_{1-x}Al_x)_2Ni$  at constant Ni content aluminium atoms substitute for Ti atoms in the 48f position ( $x, \frac{1}{8}, \frac{1}{8}$ ) of space group  $Fd\bar{3}m$  up to 75%. It is interesting to note that titanium atoms in the 16c position (0,0,0) at the centers of metal icosahedra show no preference for Ti/Al exchange; Ni atoms are in the 32e sites ( $x, x, x$ ). One may assume that nitrogen atoms, upon entering the quaternary system, will simply occupy the center positions of the octahedral voids provided by the metal host lattice. The new quaternary compound,  $Ti_3NiAl_2N$ , however, shows a completely different metal atom occupation scheme with Ti atoms in the 48f positions, Ni atoms in the 16c sites, and Al atoms in the 32e positions. The nitrogen atoms are localized in the 16d sites ( $\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$ ) with octahedral coordination  $[Ti_6]N$  formed by six Ti atoms. From EMPA and the Rietveld analyses we only observe a small homogeneity range with minor Ni/Al exchange in the 32e and 16c sites (one atom per site; see Table 3). A similar occupation scheme is obtained from Rietveld analysis of the corresponding oxygen-containing alloy  $Ti_3NiAl_2O$ .

With respect to the phase equilibria listed in Tables 1 and 2, and in combination with the results of the Rietveld refinement, we thus may safely assume that neither the ternary solid solution  $(Ti, Al)_2Ni$  nor the ternary phases  $Ti_4Ni_2(N, O)$  show any connectivity with the quaternary compounds  $Ti_3NiAl_2(N, O)$  via atom site exchange accompanied by successive nitrogen/oxygen incorporation. These results are then consistent with the formation of truly quaternary compounds  $Ti_3NiAl_2N$  and  $Ti_3NiAl_2O$  deriving from the  $Ti_2Ni$ -type parent structure.

From the viewpoint of structural chemistry the fully ordered crystal structure of  $Ti_3NiAl_2N$  corresponds to  $Ta_3CrAl_2C$  first reported by Reiffenstein *et al.* (5). The mode of site occupation and filling the octahedral voids in the so-called  $\eta$  phases deriving from the parent  $Ti_2Ni$  type was earlier extensively discussed by Rogl and Nowotny (6). Referring to this analysis  $(Ti_{1-x}Al_x)_2Ni$  must be classified as a filled metal host lattice structure where the titanium atoms (in 16c) at the centers of the metal icosahedra (each formed by six Ti/Al) atoms in 48f and six Ni atoms in 32e do not allow for Ti/Al substitution. For  $\eta$  carbides with the stoichiometry  $M_3M''M_2C$  such as  $Ta_3CrAl_2C$  (5) or  $Ti_3NiAl_2C$  (7), however, we observe the small transition metal such as chromium or nickel at the centers of the metal icosahedra, the latter now formed by Ti (Ta) in 48f and Al in 32e positions. This occupation scheme essentially corresponds with the atom site distribution in  $Ti_3NiAl_2N$  and  $Ti_3NiAl_2O$ . Nonmetal atoms such as C, N, or O are assumed to occupy the octahedral voids in 16d but may also enter the octahedral voids in the 8a sites ( $\frac{1}{8}, \frac{1}{8}, \frac{1}{8}$ ). Due to the rather small X-ray scattering power of the nitrogen/oxygen atoms a further precision in the distribution of these light atoms can only be awaited from neutron diffraction experiments.

**TABLE 1**  
**Crystallographic Data and EMPA Results for Quaternary Ti–Ni–Al–N Alloys**

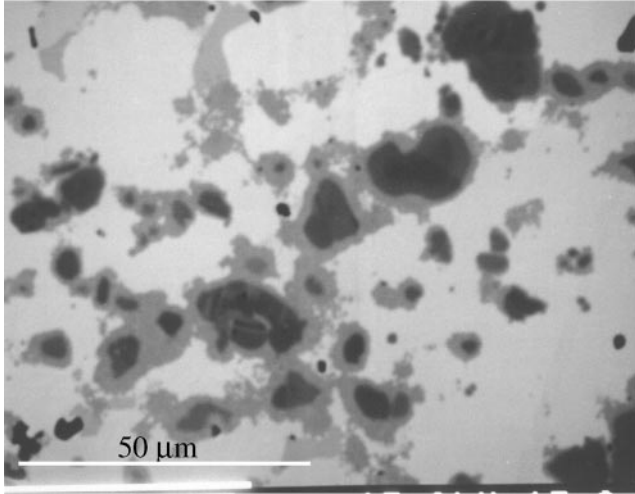
Nominal composition in at. %				Heat treatment	X-ray phase analysis	Space group	Pearson symbol	Structure type	Lattice parameters (pm)		Volume (10 <sup>6</sup> pm <sup>3</sup> )	Results of EMPA in at. %				
Ti	Ni	Al	N						<i>a</i>	<i>c</i>		Ti	Ni	Al	N	
47	28.6	10.1	14.3	Arc	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	422.96(2)		75.665(8)					
					(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1124.91(8)		1423.5(2)					
				900°C	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	422.61(2)		75.477(8)	58.9(6)	0.2(1)	0	40.9(6)	
					(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1128.15(10)		1435.8(3)	59.3(7)	32.8(4)	7.9(5)	0	
					$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>				28.9(5)	48.8(4)	22.3(5)	0	
				TiNi	<i>Pm</i> $\bar{3}m$	cP2	CsCl				47.1(3)	47.3(2)	5.6(1)	0		
37	28.6	20.1	14.3	Arc	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.29(2)		75.845(9)					
					Ti <sub>3</sub> NiAl <sub>2</sub> N	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni <sup>a</sup>	1137.01(4)		1469.9(1)					
				900°C	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.32(3)		75.86(1)	50.6(5)	0.1	—	49.3(5)	
					Ti <sub>3</sub> NiAl <sub>2</sub> N	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1137.60(6)		1472.2(1)	44.2(4)	14.8(2)	28.6(2)	12.4(7)	
					$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>				35.3(1)	25.9(2)	38.8(1)	0	
				$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>				25.4(1)	48.9(2)	25.7(1)	0		
28.6	28.6	28.6	14.3	Arc	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.70(1)		76.062(4)					
				900°C	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.22(3)		75.80(1)	53.7(4)	0.5(1)	0.2(2)	45.6(5)	
					NiAl	<i>Fm</i> $\bar{3}m$	cP2	CsCl	289.31(4)		24.22(1)	7.9(4)	49.0(2)	43.1(5)	0	
42.9	14.3	28.6	14.3	Arc	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.13(6)		75.76(2)					
					Ti <sub>3</sub> NiAl <sub>2</sub> N	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1137.02(6)		1469.9(1)					
				900°C	TiN	<i>Fm</i> $\bar{3}m$	cF8	NaCl	423.18(4)		75.78(1)					
				Ti <sub>3</sub> NiAl <sub>2</sub> N	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1137.37(4)		1471.3(1)						

<sup>a</sup>Derivative of Ti<sub>2</sub>Ni type.

**TABLE 2**  
**Crystallographic Data and EMPA Results for Quaternary Ti–Ni–Al–O Alloys at 900°C**

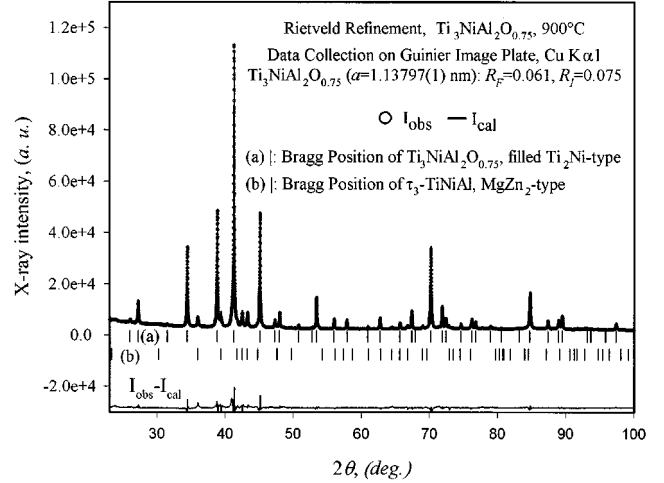
Nominal composition in at. %				Heat treatment	X-ray phase analysis	Space group	Pearson symbol	Structure type	Lattice parameters (pm)		Volume (10 <sup>6</sup> pm <sup>3</sup> )	Results of EMPA in at. %								
Ti	Ni	Al	O						<i>a</i>	<i>c</i>		Ti	Ni	Al	O					
47	28.6	10.1	14.3	Arc	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1123.48(11)		1418.1(2)									
					$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	590.83(2)		206.25(1)									
				900°C	Ti <sub>2</sub> O	<i>Fm</i> $\bar{3}m$	cF8	NaCl	296.64(4)	481.71(24)	42.388(2)									
					(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1124.23(12)		1420.9(1)	54.7	35.2	10.1 <sup>a</sup>						
37	28.6	20.1	14.3	Arc	$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	590.02(-)		205.41(3)	28.5	54.9	16.6 <sup>a</sup>						
					(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1124.05(17)		1420.2(1)									
					$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>	499.95(12)	805.10(61)	201.23(4)									
					$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	589.94(19)		205.32(1)									
				900°C	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1123.48(15)		1418.1(2)	51.5	34.2	14.1	0.2					
					$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>	500.40(11)	809.52(32)	202.71(2)	38.3	32.2	29.5	0					
					$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	589.75(6)		205.12(1)	26.5	56.0	17.5	0					
					Ti <sub>3</sub> NiAl <sub>2</sub> O	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1137.71(16)		1472.6(2)	44.0	14.9	27.9	13.2					
28.6	28.6	28.6	14.3	Arc	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1124.35(22)		1421.4(2)									
					$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>	498.43(32)	802.62(18)	199.41(4)									
					$\tau_4$ -TiNi <sub>2</sub> Al	<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	589.58(22)		204.94(3)									
					900°C	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1123.61(15)		1418.6(2)								
				$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>		<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>	499.86(50)	808.99(58)	202.13(5)	53.2	20.5	26.3 <sup>a</sup>						
				$\tau_4$ -TiNi <sub>2</sub> Al		<i>Fm</i> $\bar{3}m$	cF16	BiF <sub>3</sub>	589.47(14)		204.83(2)									
				Ti <sub>3</sub> NiAl <sub>2</sub> O		<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1137.91(11)		1473.4(1)									
				42.9	14.3	28.6	14.3	Arc	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96	Ti <sub>2</sub> Ni	1130.90(-)		1446.3(1)					
$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>						501.73(18)	816.21(47)	205.47(2)									
Ti <sub>3</sub> NiAl <sub>2</sub> O	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni						1137.97(27)		1473.6(2)									
900°C	(Ti <sub>1-x</sub> Al <sub>x</sub> ) <sub>2</sub> Ni	<i>Fd</i> $\bar{3}m$	cF96						Ti <sub>2</sub> Ni	1125.24(24)		1424.7(2)	53.2	27.9	18.9 <sup>a</sup>					
	$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>	<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12					MgZn <sub>2</sub>	501.87(36)	813.19(51)	204.82(3)	37.3	30.6	32.1 <sup>a</sup>						
	Ti <sub>3</sub> NiAl <sub>2</sub> O	<i>Fd</i> $\bar{3}m$	cF112					der. Ti <sub>2</sub> Ni	1137.71(9)		1472.6(2)									
	44.5	14.8	29.6					11.1	900°C	Ti <sub>3</sub> NiAl <sub>2</sub> O	<i>Fd</i> $\bar{3}m$	cF112	der. Ti <sub>2</sub> Ni	1138.76(6)		1476.7(1)				
$\tau_3$ -Ti <sub>2</sub> NiAl <sub>3</sub>										<i>P6</i> <sub>3</sub> / <i>mmc</i>	hP12	MgZn <sub>2</sub>	501.28(31)	815.19(36)	204.84(3)	53.4	19.9	26.7 <sup>a</sup>		

<sup>a</sup>Oxygen not determined.



**FIG. 1.** Alloy  $\text{Ti}_{37}\text{Ni}_{28.6}\text{Al}_{20.1}\text{N}_{14.3}$  after annealing at  $900^\circ\text{C}$  and quenched. BSE, image. Dark phase TiN ( $\text{Ti}_{50.6}\text{Ni}_{0.1}\text{N}_{49.3}$ ), dark gray phase  $\text{Ti}_3\text{NiAl}_2\text{N}$  ( $\text{Ti}_{44.2}\text{Ni}_{14.8}\text{Al}_{28.6}\text{N}_{12.4}$ ), light gray phase ( $\text{Ti}_{35.3}\text{Ni}_{25.9}\text{Al}_{38.8}$ ), and bright phase ( $\text{Ti}_{25.4}\text{Ni}_{48.9}\text{Al}_{25.7}$ ).

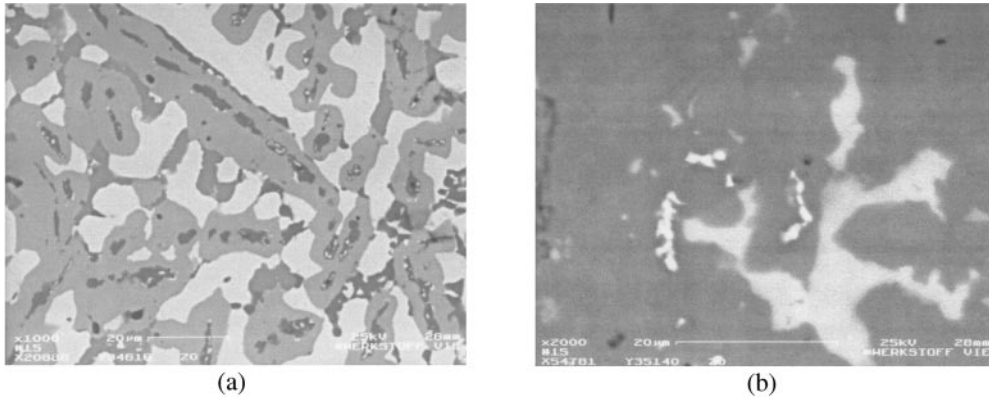
*b.2. On the correspondence of the  $x$  phase with  $\text{Ti}_3\text{NiAl}_2\text{O}_{1-x}$ .* Umemoto and co-workers (8–10) reported an unknown phase with the  $\text{Ti}_2\text{Ni}$  type in the Ti–Ni–Al system. These authors suggested stabilization of this phase, labeled as  $x$ , by small amounts of oxygen. In the homologous systems Ti–{Pd, Pt}–Al–O we indeed found a quaternary oxide,  $\text{Ti}_3\{\text{Pd, Pt}\}\text{Al}_2\text{O}$ , deriving from the filled  $\text{Ti}_2\text{Ni}$  type (11). Speculation on isotopic behavior for the analogous Ni-containing compound and the above analysis on the Ti–Ni–Al–N alloys motivated us to investigate the Ti–Ni–Al–O system (Table 2). X-ray powder diffraction of an annealed alloy  $\text{Ti}_{44.5}\text{Ni}_{14.8}\text{Al}_{29.6}\text{O}_{11.1}$  (at.%) proved a nearly single-phase oxide with a small amount of  $\tau_3$ -TiNiAl (MgZn<sub>2</sub> type). The Rietveld refinement shows that the calculated diffraction pattern on the basis of the



**FIG. 3.** Rietveld refinement of  $\text{Ti}_{44.5}\text{Ni}_{14.8}\text{Al}_{29.6}\text{O}_{11.1}$  (at.%) annealed at  $900^\circ\text{C}$ .

filled  $\text{Ti}_2\text{Ni}$  type and with atom order corresponding to  $\text{Ti}_3\text{NiAl}_2\text{N}$  perfectly matches the observation and thus reveals a formula  $\text{Ti}_3\text{NiAl}_2\text{O}_{0.75}$  with a lattice parameter  $a = 1.13797(1)$  nm (Fig. 3). The metal atoms in the quaternary oxide are completely ordered as shown in Table 3 but the site  $16d$  ( $\frac{1}{2}, \frac{1}{2}, \frac{1}{2}$ ) is not fully occupied by O atoms (only about 75%).

To reveal the characteristic differences in the X-ray intensities depending on the atom site occupation, we calculated model X-ray spectra (Fig. 4) for the various phases that crystallize with the  $\text{Ti}_2\text{Ni}$ -parent type:  $\text{Ti}_2\text{Ni}$ ,  $\text{Ti}_4\text{Ni}_2\text{O}$ ,  $(\text{Ti,Al})_4(\text{Ni,Al})_2\text{O}$  with random Al distribution in Ti and Ni sites and ordered  $\text{Ti}_3\text{NiAl}_2\text{O}$ . Comparing these X-ray spectra in Fig. 4, the major features of the  $\text{Ti}_3\text{NiAl}_2\text{O}$  spectrum are easily observed: a weak X-ray intensity for the crystallographic planes (111) at  $2\theta = 13.4^\circ$  and (442) at  $2\theta = 47.9^\circ$ , but, a strong intensity for the reflections (331) at



**FIG. 2.** Alloy (a)  $\text{Ti}_{28.6}\text{Ni}_{28.6}\text{Al}_{28.6}\text{O}_{14.3}$  (white,  $\tau_4$ ; gray,  $(\text{Ti,Al})_2\text{Ni}$ ; dark gray,  $\tau_3$ ; dark,  $\text{Ti}_3\text{NiAl}_2\text{O}$ ) and (b)  $\text{Ti}_{42.9}\text{Ni}_{14.3}\text{Al}_{28.6}\text{O}_{14.3}$  (gray matrix:  $\text{Ti}_3\text{NiAl}_2\text{O}$ ; white,  $(\text{Ti,Al})_2\text{Ni}$ ; white gray,  $\tau_3$ ) annealed at  $900^\circ\text{C}$ , primary electron image.

**TABLE 3**  
**Rietveld Refinements for  $(\text{Ti}_{1-x}\text{Al}_x)_2\text{Ni}$ ,  $\text{Ti}_3\text{NiAl}_2\text{N}$ , and  $\text{Ti}_3\text{NiAl}_2\text{O}$**

Parameter/compound	$(\text{Ti}_{1-x}\text{Al}_x)_2\text{Ni}$	$\text{Ti}_3(\text{Ni}, \text{Al})_3\text{N}$	$\text{Ti}_3\text{NiAl}_2\text{N}$	$\text{Ti}_3\text{NiAl}_2\text{O}$
Nominal composition in at. %	$\text{Ti}_{56.8}\text{Ni}_{29}\text{Al}_{14.2}$	$\text{Ti}_{3.7}\text{Ni}_{28.6}\text{Al}_{20.1}\text{N}_{14.3}$	$\text{Ti}_{42.8}\text{Ni}_{14.3}\text{Al}_{28.6}\text{N}_{14.3}$	$\text{Ti}_{44.5}\text{Ni}_{14.8}\text{Al}_{29.6}\text{O}_{11.1}$
$a$ (nm)	1.12316(5)	1.13684(2)	1.13685(2)	1.13797(1)
$V$ (nm <sup>3</sup> )	1.41687	1.46926	1.46929	1.47369
Data collection	D5000, CuK $\alpha$	D5000, CuK $\alpha$	D5000, CuK $\alpha$	Image plate CuK $\alpha_1$
$2\theta$ range	10°–110°	10°–110°	10°–110°	8°–100°
Number of variables	25	24	24	24
$R_F = \sum  F_o - F_c  / \sum F_o$	0.035	0.027	0.038	0.061
$R_I = \sum  I_o - I_c  / \sum I_o$	0.052	0.034	0.049	0.075
$R_{wP} = [\sum w_i  y_{oi} - y_{ci} ^2 / \sum w_i  y_{oi} ^2]^{1/2}$	0.139	0.128	0.166	0.067
$R_P = \sum  y_{oi} - y_{ci}  / \sum  y_{oi} $	0.110	0.099	0.118	0.046
$R_e = \{(N - P + C) / (\sum w_i  y_{oi} ^2)\}^{1/2}$	0.076	0.069	0.059	0.036
$\chi^2 = (R_{wP}/R_e)^2$	3.35	3.38	8.04	17.8
<b>Atom parameters</b>				
M1 in 48f (x, 1/8, 1/8)	36Ti + 12Al	48Ti	48Ti	48Ti
x	0.4392(2)	0.4391(2)	0.4390(2)	0.4369(1)
$B_{\text{eq}}(B_{\text{iso}}) 10^2$ (nm <sup>2</sup> )	1.73(8)	1.39(6)	2.27(9)	1.6(1)
Occ.	1.0	1.0	1.0	1.0
M2 in 32e (x, x, x)	32Ni	31Al + 1Ni	32Al	32Al
x	0.2136(1)	0.2077(2)	0.2079(2)	0.2087(1)
$B_{\text{eq}}(B_{\text{iso}}) 10^2$ (nm <sup>2</sup> )	2.05(1)	1.46(9)	2.1(2)	1.2(1)
Occ.	1.0	1.0	1.0	1.0
M3 in 16d (1/2, 1/2, 1/2)	—	16N	16N	16O
$B_{\text{eq}}(B_{\text{iso}}) 10^2$ (nm <sup>2</sup> )	—	1.1(3)	2.6(5)	1.8(1)
Occ.	—	1.0	1.0	0.75
M4 in 16c (0, 0, 0)	16Ti	15Ni + 1Al	16Ni	16Ni
$B_{\text{eq}}(B_{\text{iso}}) 10^2$ (nm <sup>2</sup> )	1.9(2)	1.31(8)	2.7(2)	1.7(1)
Occ.	1.0	1.0	1.0	1.0
Distances (nm) within the first nearest neighbor co-ordination; standard deviations are < 0.001 nm				
M1 2M3	—	0.2125	0.2126	0.2137
2M2	0.2579	0.2699	0.2695	0.2688
2M2	0.2899	0.2948	0.2946	0.2925
2M4	0.2908	0.2943	0.2942	0.2927
4M1	0.2951	0.2988	0.2990	0.3014
4M1	0.2988	0.3023	0.3023	0.3028
M2 3M4	0.2468	0.2457	0.2458	0.2466
3M1	0.2579	0.2699	0.2695	0.2688
3M2	0.2814	0.2658	0.2665	0.2694
3M1	0.2899	0.2948	0.2946	0.2925
3M3	—	0.3392	0.3389	0.3381
M3 6M1	—	0.2125	0.2126	0.2137
M4 6M2	0.2468	0.2457	0.2458	0.2466
6M1	0.2908	0.2943	0.2942	0.2927

Note. Space group  $Fd\bar{3}m-O_h^7$  (No. 227); origin at center,  $Z = 2$ .  $\text{Ti}_2\text{Ni}$  type. Standard deviations in parentheses.

$2\theta = 34.1^\circ$ , (222) at  $2\theta = 27.1^\circ$  and (622) at  $2\theta = 53.4^\circ$ . The atom arrangement in  $\text{Ti}_3\text{NiAl}_2\text{O}_{1-x}$ , being significantly different from that in  $\text{Ti}_2\text{Ni}$  or  $\text{Ti}_4\text{Ni}_2\text{O}$ , reveals a larger unit cell dimension when compared to  $\text{Ti}_2\text{Ni}$  or  $\text{Ti}_4\text{Ni}_2\text{O}$  (see for instance Fig. 4). Although our lattice parameter for  $\text{Ti}_3\text{NiAl}_2\text{O}$  is slightly lower than that for the reported  $x$  phase, the dependency on atom sites in combination with the characteristic intensity relations would suggest that the  $x$  phase reported by Umemoto and co-workers (8–10) may

well be labeled as  $\text{Ti}_3\text{NiAl}_2\text{O}_{1-x}$  rather than a ternary compound with  $(\text{Ti},\text{Al})_2(\text{Ni},\text{Al})$  stoichiometry.

### CONCLUSION

The phase equilibria in the Ti–Ni–Al–N and Ti–Ni–Al–O systems were determined from bulk alloys. Two quaternary compounds  $\text{Ti}_3\text{NiAl}_2\text{N}$  and  $\text{Ti}_3\text{NiAl}_2\text{O}_{1-x}$  with the filled  $\text{Ti}_2\text{Ni}$  type were found to exist. The characteristic

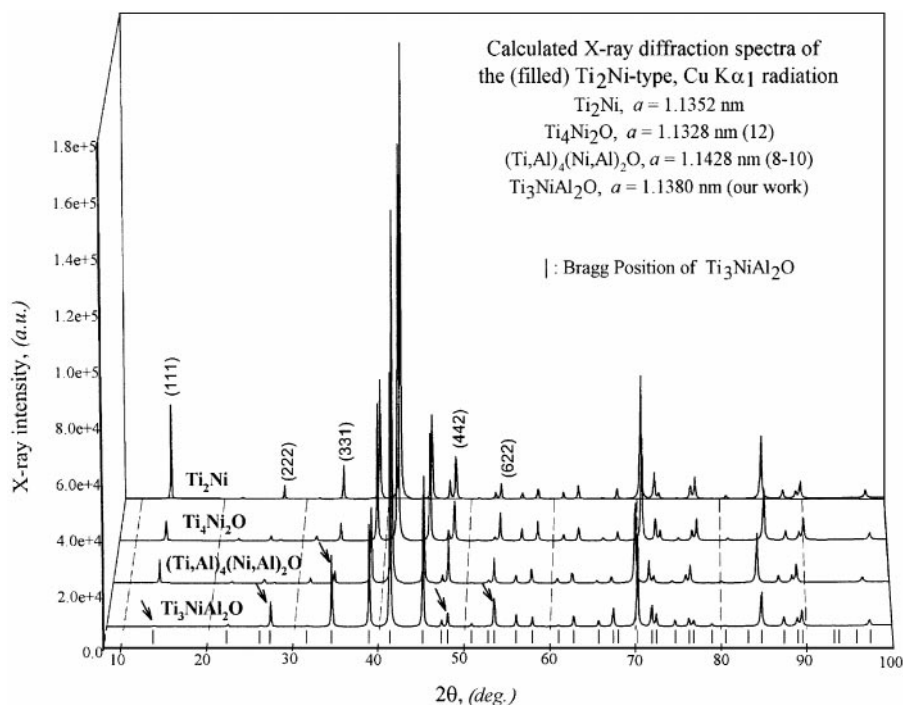


FIG. 4. Comparison of calculated X-ray spectra of  $Ti_2Ni$ -type phases:  $Ti_2Ni$ ,  $Ti_4Ni_2O$ ,  $(Ti,Al)_4(Ni,Al)_2O$ , and  $Ti_3NiAl_2O$ .

differences of X-ray spectra among the various compounds with  $Ti_2Ni$ -type parent structure,  $Ti_2Ni$ ,  $Ti_4Ni_2O$ , and  $Ti_3NiAl_2O$ , were analyzed. Rietveld refinements were presented for  $Ti_3NiAl_2N$  and  $Ti_3NiAl_2O_{1-x}$ .

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